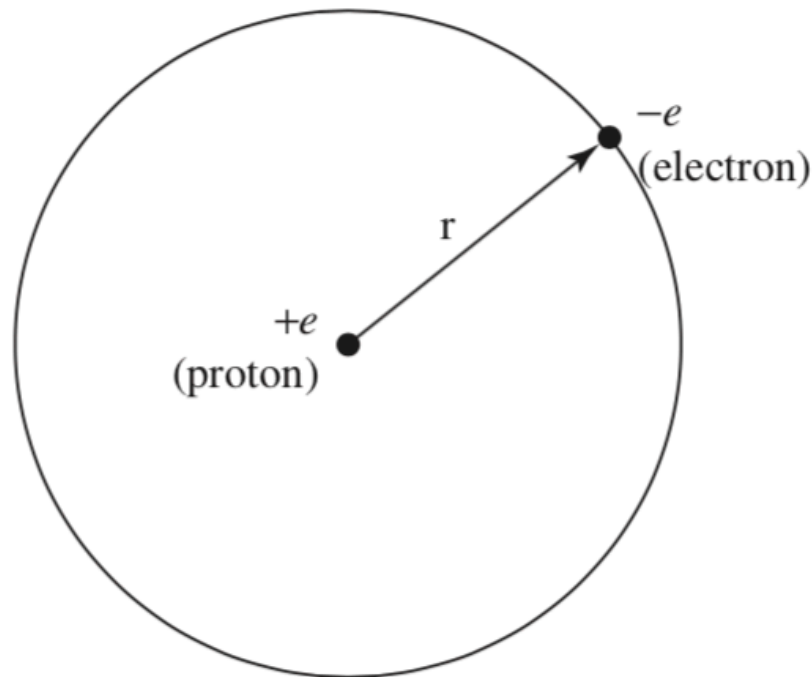


The hydrogen atom



The hydrogen atom consists of proton of charge e , together with a much lighter electron charge $-e$.

$$m_p = 1.67 \times 10^{-27} \text{ kg}$$

$$m_e = 9.11 \times 10^{-31} \text{ kg}$$

$$e = 1.60 \times 10^{-19} \text{ C}$$

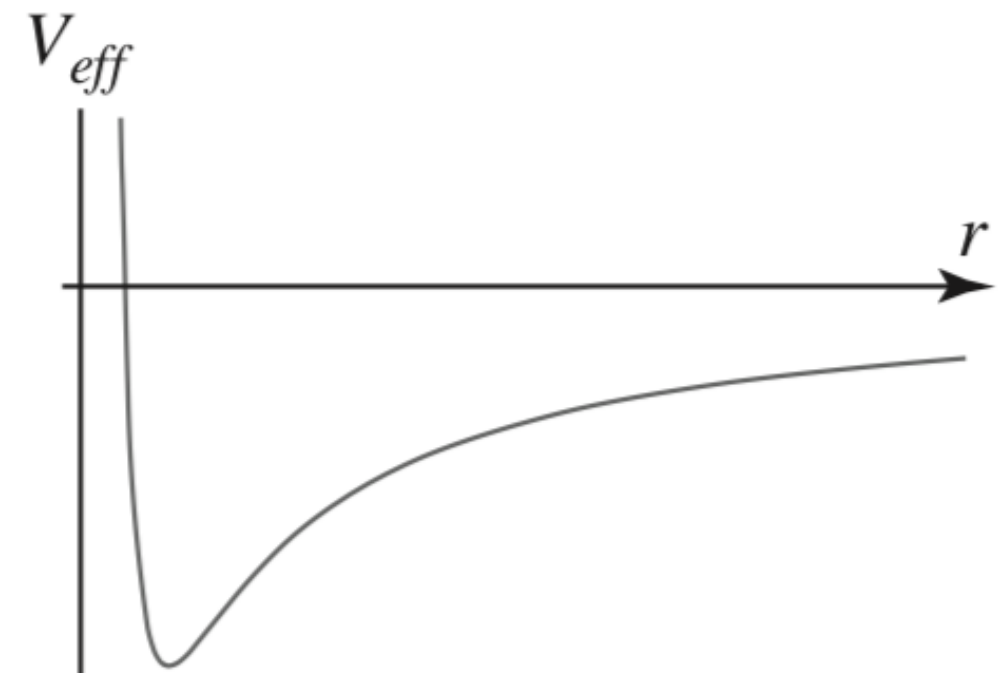
From Coulomb's law: $F = \frac{1}{4\pi\epsilon_0} \frac{e^2}{r^2}$ the potential in SI units is: $V(r) = -\frac{e^2}{4\pi\epsilon_0} \frac{1}{r}$

And the radial equation becomes:

$$-\frac{\hbar^2}{2m} \frac{d^2 u}{dr^2} + \left[V + \frac{\hbar^2}{2m} \frac{\ell(\ell+1)}{r^2} \right] u = Eu$$

$$-\frac{\hbar^2}{2m_e} \frac{d^2 u}{dr^2} + \left[-\frac{e^2}{4\pi\epsilon_0} \frac{1}{r} + \frac{\hbar^2}{2m_e} \frac{\ell(\ell+1)}{r^2} \right] u = Eu.$$

Effective potential (V_{eff})



The hydrogen atom

We need to solve this equation for $u(r)$, and determine the allowed energies.

The Coulomb potential admits:

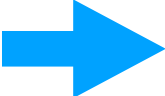
- Scattering states ($E > 0$) \rightarrow electron-proton scattering
- Bound states ($E < 0$) \rightarrow hydrogen atom

The Radial Wave Function

We are interested in finding bound states ($E < 0$) of:

$$-\frac{\hbar^2}{2m_e} \frac{d^2 u}{dr^2} + \left[-\frac{e^2}{4\pi\epsilon_0} \frac{1}{r} + \frac{\hbar^2}{2m_e} \frac{\ell(\ell+1)}{r^2} \right] u = Eu$$

Let's divide this equation by E and define: $\kappa \equiv \frac{\sqrt{-2m_e E}}{\hbar}$

 $\frac{1}{\kappa^2} \frac{d^2 u}{dr^2} = \left[1 - \frac{m_e e^2}{2\pi\epsilon_0 \hbar^2 \kappa} \frac{1}{(\kappa r)} + \frac{\ell(\ell+1)}{(\kappa r)^2} \right] u$

The hydrogen atom

The Radial Wave Function

$$\frac{1}{\kappa^2} \frac{d^2 u}{dr^2} = \left[1 - \frac{m_e e^2}{2\pi \epsilon_0 \hbar^2 \kappa} \frac{1}{(\kappa r)} + \frac{\ell(\ell+1)}{(\kappa r)^2} \right] u$$

We introduce:

$$\begin{aligned} \rho &\equiv \kappa r, \\ \rho_0 &\equiv \frac{m_e e^2}{2\pi \epsilon_0 \hbar^2 \kappa}, \end{aligned} \quad \Rightarrow \quad \frac{d^2 u}{d\rho^2} = \left[1 - \frac{\rho_0}{\rho} + \frac{\ell(\ell+1)}{\rho^2} \right] u.$$

Let's analyse the asymptotic behaviour of this equation.

$$\begin{aligned} \rho \rightarrow \infty &\Rightarrow \frac{d^2 u}{d\rho^2} = u \Rightarrow u(\rho) = Ae^{-\rho} + Be^{\rho} \xrightarrow{\text{blows up}} u(\rho) \sim Ae^{-\rho} \\ &\hspace{15em} \text{(for large } \rho) \end{aligned}$$

$$\begin{aligned} \rho \rightarrow 0 &\Rightarrow \frac{d^2 u}{d\rho^2} = \frac{\ell(\ell+1)}{\rho^2} u \Rightarrow u(\rho) = C\rho^{\ell+1} + D\rho^{-\ell} \xrightarrow{\text{blows up}} u(\rho) \sim C\rho^{\ell+1} \\ &\hspace{15em} \text{(for small } \rho) \end{aligned}$$

The hydrogen atom

The Radial Wave Function

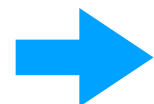
We introduce the new function $v(\rho)$:

$$u(\rho) \sim Ae^{-\rho}$$

(for large ρ)

$$u(\rho) \sim C\rho^{\ell+1}$$

(for small ρ)


$$u(\rho) = \rho^{\ell+1}e^{-\rho}v(\rho)$$

$$\frac{du}{d\rho} = \rho^{\ell}e^{-\rho} \left[(\ell + 1 - \rho)v + \rho \frac{dv}{d\rho} \right]$$

$$\frac{d^2u}{d\rho^2} = \rho^{\ell}e^{-\rho} \left\{ \left[-2\ell - 2 + \rho + \frac{\ell(\ell+1)}{\rho} \right] v + 2(\ell+1-\rho) \frac{dv}{d\rho} + \rho \frac{d^2v}{d\rho^2} \right\}$$

Therefore:

$$\frac{d^2u}{d\rho^2} = \left[1 - \frac{\rho_0}{\rho} + \frac{\ell(\ell+1)}{\rho^2} \right] u \quad \Rightarrow \quad \rho \frac{d^2v}{d\rho^2} + 2(\ell+1-\rho) \frac{dv}{d\rho} + [\rho_0 - 2(\ell+1)]v = 0$$

We assume the solution, $v(\rho)$, can be expressed as a power series in ρ :

$$v(\rho) = \sum_{j=0}^{\infty} c_j \rho^j$$

The hydrogen atom

The Radial Wave Function $\rho \frac{d^2 v}{d\rho^2} + 2(\ell + 1 - \rho) \frac{dv}{d\rho} + [\rho_0 - 2(\ell + 1)] v = 0$

We assume the solution, $v(\rho)$, can be expressed as a power series in ρ , for which we need to determine the coefficients (c_0, c_1, c_2, \dots).

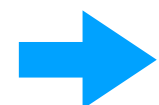
$$v(\rho) = \sum_{j=0}^{\infty} c_j \rho^j$$

$$\frac{dv}{d\rho} = \sum_{j=0}^{\infty} j c_j \rho^{j-1} = \sum_{j=0}^{\infty} (j+1) c_{j+1} \rho^j.$$

$$\frac{d^2 v}{d\rho^2} = \sum_{j=0}^{\infty} j(j+1) c_{j+1} \rho^{j-1}$$

Replacing into the radial equation above, we get:

$$\sum_{j=0}^{\infty} j(j+1) c_{j+1} \rho^j + 2(\ell+1) \sum_{j=0}^{\infty} (j+1) c_{j+1} \rho^j - 2 \sum_{j=0}^{\infty} j c_j \rho^j + [\rho_0 - 2(\ell+1)] \sum_{j=0}^{\infty} c_j \rho^j = 0$$



$$j(j+1) c_{j+1} + 2(\ell+1)(j+1) c_{j+1} - 2j c_j + [\rho_0 - 2(\ell+1)] c_j = 0$$

The hydrogen atom

The Radial Wave Function

$$\rho \frac{d^2 v}{d\rho^2} + 2(\ell + 1 - \rho) \frac{dv}{d\rho} + [\rho_0 - 2(\ell + 1)] v = 0$$

➡ $j(j+1)c_{j+1} + 2(\ell+1)(j+1)c_{j+1} - 2jc_j + [\rho_0 - 2(\ell+1)]c_j = 0$

➡
$$c_{j+1} = \left\{ \frac{2(j + \ell + 1) - \rho_0}{(j+1)(j+2\ell+2)} \right\} c_j$$

This recursion formula determines the coefficients, and hence the function $v(\rho)$.

For large j (this corresponds to large ρ , where the higher powers dominate):

$$c_{j+1} \approx \frac{2j}{j(j+1)} c_j = \frac{2}{j+1} c_j \quad \rightarrow \quad c_j \approx \frac{2^j}{j!} c_0$$

If this were the *exact* result, it blows up at large ρ (so it is not normalisable):

$$v(\rho) = c_0 \sum_{j=0}^{\infty} \frac{2^j}{j!} \rho^j = c_0 e^{2\rho} \quad \rightarrow \quad u(\rho) = c_0 \rho^{l+1} e^{\rho}$$

The hydrogen atom

The Radial Wave Function: the Bohr radius

$$v(\rho) = c_0 \sum_{j=0}^{\infty} \frac{2^j}{j!} \rho^j = c_0 e^{2\rho} \quad \rightarrow \quad u(\rho) = c_0 \rho^{l+1} e^{\rho}$$

Thus, the series must terminate: $c_{N-1} \neq 0$ but $c_N = 0$

$$c_{j+1} = \left\{ \frac{2(j + \ell + 1) - \rho_0}{(j + 1)(j + 2\ell + 2)} \right\} c_j \quad \rightarrow \quad 2(N + \ell) - \rho_0 = 0.$$

which makes $v(\rho)$ a polynomial of order $(N - 1)$, with (therefore) $N - 1$ roots, and hence the radial wave function has $N - 1$ nodes.

Let's define: $n \equiv N + \ell$ \rightarrow $\rho_0 = 2n$

Remember: $\rho_0 \equiv \frac{m_e e^2}{2\pi \epsilon_0 \hbar^2 \kappa},$

$$\rightarrow \quad \kappa = \left(\frac{m_e e^2}{4\pi \epsilon_0 \hbar^2} \right) \frac{1}{n} = \frac{1}{an}$$

$$a \equiv \frac{4\pi \epsilon_0 \hbar^2}{m_e e^2} = 0.529 \times 10^{-10} \text{ m}$$

This is the so-called **Bohr radius**.

The hydrogen atom

The Radial Wave Function: the Bohr formula

Remember: $\rho_0 \equiv \frac{m_e e^2}{2\pi\epsilon_0 \hbar^2 \kappa}, \quad \kappa \equiv \frac{\sqrt{-2m_e E}}{\hbar}$

$$\rightarrow E = -\frac{\hbar^2 \kappa^2}{2m} = -\frac{m_e e^4}{8\pi^2 \epsilon_0^2 \hbar^2 \rho_0^2} \rightarrow \boxed{E_n = -\left[\frac{m_e}{2\hbar^2} \left(\frac{e^2}{4\pi\epsilon_0} \right)^2 \right] \frac{1}{n^2} = \frac{E_1}{n^2}, \quad n = 1, 2, 3, \dots}$$

This is the **Bohr formula**

Therefore: $\rho = \frac{r}{an}$

The **spatial wave functions** are labeled by three quantum numbers (n , ℓ , and m):

$$\psi_{n\ell m}(r, \theta, \phi) = R_{n\ell}(r) Y_{\ell}^m(\theta, \phi)$$

where: $R_{n\ell}(r) = \frac{1}{r} \rho^{\ell+1} e^{-\rho} v(\rho)$

$n \equiv$ principal quantum number
 $\ell \equiv$ azimuthal quantum number
 $m \equiv$ magnetic quantum number

and $v(\rho)$ is a polynomial of degree $n - \ell - 1$ in ρ , whose coefficients are determined (up to an overall normalisation factor) by the recursion formula:

$$c_{j+1} = \frac{2(j + \ell + 1 - n)}{(j + 1)(j + 2\ell + 2)} c_j$$