

2n Report - *Numerical solution of the Schrödinger's equation*

Simulació de Sistemes Nanomètrics - *Nanociència i Nanotecnologia* - 22/23

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1 Description of the problem

I have chosen to make a 1D Particle in a Box with a normal potential barrier in the middle of the box. For this type of problem I have picked the Atomic Units, which I found a Libre text article [1]. In this problem the important units are the Hartrees, the Bohrs and the electron rest mass. The equivalences can be seen in 1. With this units my Schrödinger equation for a numerical integration resolution with D. Truhlar's method has a μ value of 1. Therefore, my equation is 1:

$$\Psi_{i-1} - 2[1 + h^2 V_i] \Psi_i + \Psi_{i+1} = 2h^2 E \Psi_i \quad (1)$$

Table 1: Equivalences between SI units and AtomicUnits

	SI UNits	More common Units	Atomic Units
Hartrees (H)	$4.35910^{-18} J$	$27.2 eV$	$1H$
Bohr (a_o)	$5.29110^{-11} m$	$0.5291 nm$	$1a_o$
m_e	$9.10910^{31} kg$	————	$1m_e$

2 Results

I based my code in two Libre Text articles [2] [3]. Therefore I chose a Box dimension of 1 a_o , a Barrier Energy of 100 H and my barrier is placed in the middle and has a dimension of 0.1 a_o .

With this units and using a grid of 500 points I have obtained the energies that can be seen in Figure 1, and their Wave Functions can be seen in Figure 2.

State	0	Energy = 15.54
State	1	Energy = 20.46
State	2	Energy = 62.69
State	3	Energy = 81.73
State	4	Energy = 143.30

Figure 1: The Energy levels with its corresponding state number ($n - 1$)

As we can see in Figure 2 even though the barrier dictates how the electron behaves inside the box, because we can clearly see that this wave functions are not the typical Particle in a Box functions [?], the barrier is thin and energetically low enough to let tunneling happen with barely any difficulty. We can also appreciate that the state $n = 5$ is not affected by the barrier, as the calculus in Figure (1), which say that this state has an energy of $143.3H$, therefore, bigger than the $100H$ barrier.

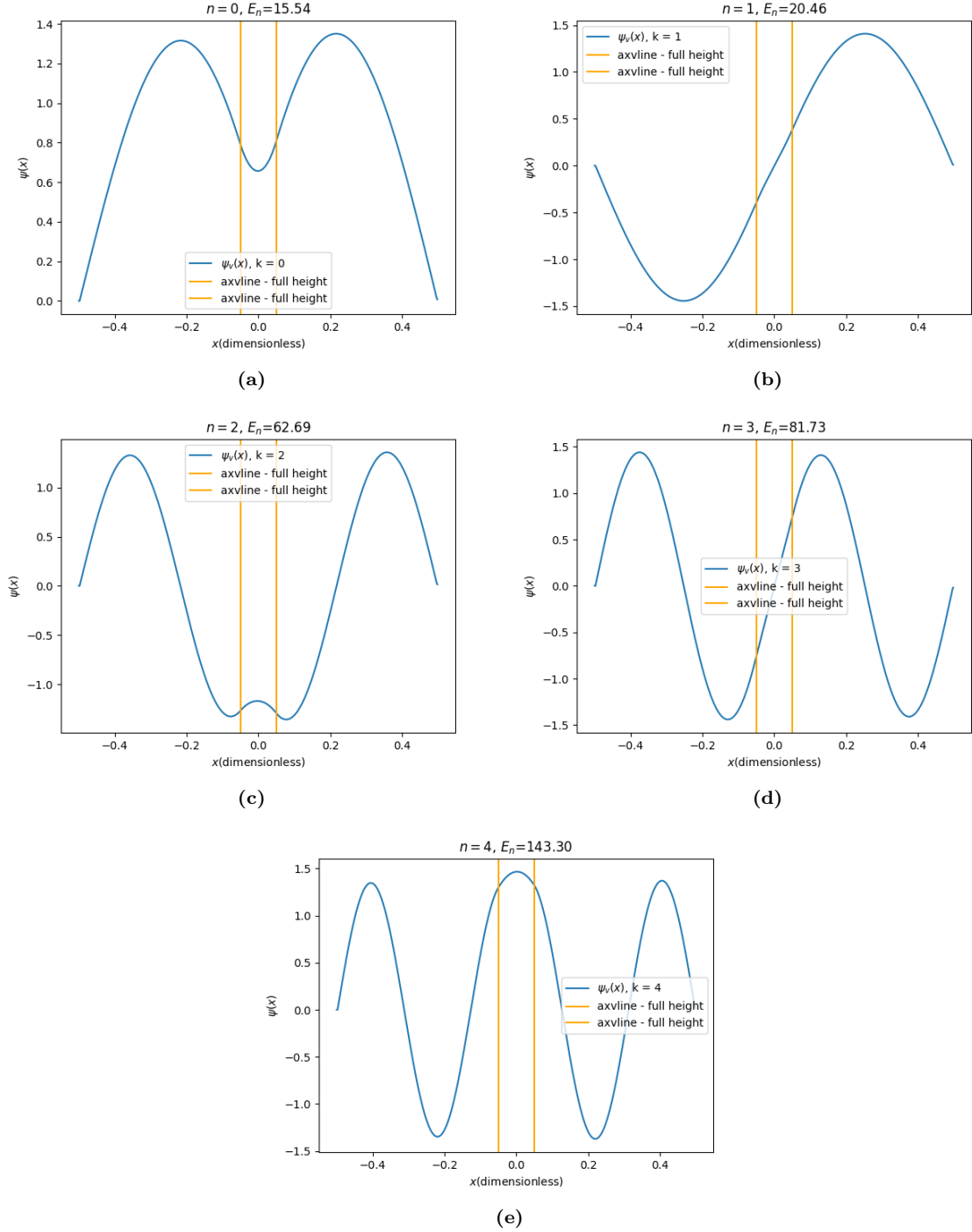


Figure 2: Figures of the different states, in blue we can see the Wave Functions, the vertical orange lines mark where the Energy Barrier is placed.

3 The code

Because of the convenience of the units used the only things I had to modify were the potential and the plots so the vertical lines appeared. This parts can be seen in Figure ?? . The entire code can be found in my GitHub repository [4].

```
#Potential as a function of position
def getV(x,Ebar):
    if (x==xlower) or (x==xupper):
        potvalue = 10000000000
    if (x>-0.05) and (x<0.05):
        potvalue = Ebar
    else:
        potvalue = 0
    return potvalue
```

(a)

```
#v = int(input("\n Quantum Number (enter 0 for ground state):\n>"))
for v in range(0,5):
    plt.plot(x,psi[v],label=r'$\psi_v(x)$', k = ' + ' + str(v))
    plt.axvline(x = 0.05, color = 'orange', label = 'axvline - full height')
    plt.axvline(x = -0.05, color = 'orange', label = 'axvline - full height')
    plt.title(r'$n$'+ str(v) + r', $E_n$=' + '{:.2f}'.format(E[v]))
    #plt.axhline(x=-0.05, color = 'orange')
    #plt.axhline(x=0.05, color = 'orange')
    plt.legend()
    plt.xlabel(r'$x$(dimensionless)')
    plt.ylabel(r'$\psi(x)$')
    plt.show()
```

(b)

References

- [1] “Atomic and molecular calculations are expressed in atomic units, libre text.” [https://chem.libretexts.org/Bookshelves/Physical_and_Theoretical_Chemistry_Textbook_Maps/Physical_Chemistry_\(LibreTexts\)/08%3AMultielectron_Atoms/8.01%3AAtomic_and_Molecular_Calculations_are_Expressed_in_Atomic_Units](https://chem.libretexts.org/Bookshelves/Physical_and_Theoretical_Chemistry_Textbook_Maps/Physical_Chemistry_(LibreTexts)/08%3AMultielectron_Atoms/8.01%3AAtomic_and_Molecular_Calculations_are_Expressed_in_Atomic_Units).
- [2] “Atomic and molecular calculations are expressed in atomic units, libre text.” [https://chem.libretexts.org/Bookshelves/Physical_and_Theoretical_Chemistry_Textbook_Maps/Quantum_Tutorials_\(Rioux\)/09%3ANumerical_Solutions_for_Schrodinger's_Equation/9.15%3AParticle_in_a_Box_with_an_Internal_Barrier](https://chem.libretexts.org/Bookshelves/Physical_and_Theoretical_Chemistry_Textbook_Maps/Quantum_Tutorials_(Rioux)/09%3ANumerical_Solutions_for_Schrodinger's_Equation/9.15%3AParticle_in_a_Box_with_an_Internal_Barrier).
- [3] “Another look at the in a box with an internal barrier, libre text.” [https://chem.libretexts.org/Bookshelves/Physical_and_Theoretical_Chemistry_Textbook_Maps/Quantum_Tutorials_\(Rioux\)/09%3ANumerical_Solutions_for_Schrodinger's_Equation/9.16%3AAnother_Look_at_the_in_a_Box_with_an_Internal_Barrier](https://chem.libretexts.org/Bookshelves/Physical_and_Theoretical_Chemistry_Textbook_Maps/Quantum_Tutorials_(Rioux)/09%3ANumerical_Solutions_for_Schrodinger's_Equation/9.16%3AAnother_Look_at_the_in_a_Box_with_an_Internal_Barrier).
- [4] “Github repository with the code to solve this problem.”