

# Reaction-Advection-Diffusion Models of Air Pollution

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## 1 Introduction

Airborne pollutants are the single most impactful environmental risk to human health, with about one in nine deaths worldwide attributed to air pollution (Cohen et al., 2017). Poor air quality can cause many negative implications to human health; exposure to more air pollution is directly correlated to an increased rate of heart attacks and asthma. There are a few different types of air pollution that can cause adverse affects on human health. First, fine particulate matter (PM<sub>2.5</sub>) causes approximately 95% of the deaths due to air pollution. (PM<sub>2.5</sub>) is defined to be any microscopic particles that are of diameter of less than 2.5 $\mu$ m, about 20 times smaller than the diameter of a human hair.

Due to the minute scale of these particles, particulate matter is capable of traveling deep into vital organs such as the lungs or the bloodstream. Exposure to increased rates of fine particulate matter can lead to coughing and difficulty breathing. Particulate matter is typically produced by the combustion of gasoline, oil, or wood, but has recently been discovered to occur at smaller rates from natural processes such as volcanic eruptions (Thangavel et al., 2022). In addition to particulate matter, surface-level ozone (O<sub>3</sub>) and nitrogen dioxide (NO<sub>2</sub>) also contribute to the damages to human health from air pollution. In this paper, we will validate recent diffusion models of air pollution and the models to estimate damages to human health caused by pollution.

## 2 Types of Models

Before beginning on the mathematics behind a differential equation model of air pollution, we will discuss how to form a model of air pollution. One of the main problems with making a diffusion model for air pollution is estimating the flux term of the diffusion equation. We apply Fick's first and second laws of diffusion; recall that Fick's first law relates the flux,  $J$  to the diffusion constant  $k$ , and the change in the concentration gradient  $c$ :

$$J = -k \frac{dc}{dx} \tag{1}$$

Typically, chemical transport models (CTMs), programs that simulate and predict atmospheric chemistry, are used to find the fluxes needed in a diffusion equation. There are two classes of CTMs: Eulerian and Lagrangian. Eulerian models have an observer standing in one location who is watching the air pass by, whereas the Lagrangian models have the observer moving along with the parcel of air as it takes a random walk through the atmosphere (Vallero, 2021).

However, executing CTMs can be computationally difficult. We will first look at a simple model for representing the concentrations of chemicals in the atmosphere.

## 2.1 A Simple Model

The model we will first investigate is called the one-box model (Jacob, 1999).

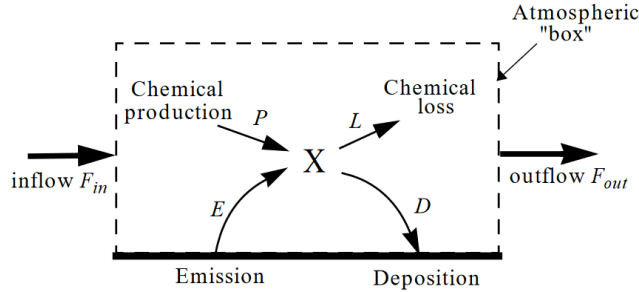


Figure 1: One-box model for an atmospheric species X

The one-box model for an atmospheric species X is illustrated in Figure 1. The model describes the concentration of X inside of some subset of the global atmosphere (ex: the United States). The transport rate is described using the flow in,  $F_{in}$ , and the flow out of,  $F_{out}$  of the box. By conservation, the global atmosphere must obey

$$F_{in} = F_{out} = 0$$

The other parameters of the model are:

- Chemical production ( $P$ )
- Emissions ( $E$ )
- Chemical loss ( $L$ )
- Deposition ( $D$ )

We will now define the lifetime,  $\tau$ , of a molecule, to be the average time a molecule of X remains in our box. This will be the mass of the particle divided by the outflow of the box:

$$\tau = \frac{m}{F_{out} + L + D} \quad (2)$$

We also can set up a mass balance differential equation. The change of the abundance of a species must be equal to the sources less the sinks:

$$\frac{dm}{dt} = F_{in} + E + P - F_{out} - L - D \quad (3)$$

We need to represent the sinks using some loss rate. Let's assume that the losses are proportional to the amount of the species in the box, so the losses are equal to  $km$ . (It turns out trivially that  $k = \frac{1}{\tau}$ ). Let's also assume that the sources are independent of  $m$  and are equal to some constant  $S$ . Then, we yield the ODE

$$\frac{dm}{dt} = S - km \quad (4)$$

which can be solved quite simply via separation of variables:

$$\begin{aligned} \frac{dm}{dt} &= S - km \\ \Rightarrow \int_0^t \frac{dm}{S - km} &= \int_0^t dt \\ \Rightarrow \frac{-1}{k} \ln \left| \frac{S - km(t)}{S - km(0)} \right| &= t \\ \Rightarrow S - km(t) &= Se^{-kt} - km(0)e^{-kt} \\ \Rightarrow m(t) &= \frac{S}{k}(1 - e^{-kt}) + m(0)e^{-kt} \end{aligned} \quad (5)$$

Finding the steady state of this equation yields  $S_\infty = \frac{S}{k}$  (via taking the limit as  $t \rightarrow \infty$ ).

This steady state assumption is reasonable because the rate of change of the mass,  $\frac{dm}{dt}$ , tends to be very small in comparison to the production rate. Because of this, we classify this steady state as a quasi steady state.

## 2.2 A More Realistic Model: the Puff Model

We will now investigate a puff model; such a model describes fluid elements moving in space. We define a fluid element to be a volume of air in which all particles move with the same velocity. The most common application of a puff model is diffusion from a smokestack, something that is essential for understanding how air pollution diffusion. We define the mass balance equation for this model to be

$$\frac{dX_c}{dt} = E + P - L - D \quad (6)$$

where  $X_c$  is the concentration of element X. Assume that the puff takes a column shape and that the height of the column is  $h$ . Also assume that

there is no chemical production ( $P$ ) and that the losses are proportional to the concentration at some rate  $L = kX_c$ . We then form the mass-balance equation:

$$\frac{dX_c}{dt} = \frac{E}{h} - kX_c \quad (7)$$

This is another separable ODE which is trivial to solve. The puff model's steady states are only valid in low-wind conditions, however (Zannetti, 1985). Many computer systems used by the federal government to predict air pollution use variations of this model with more complicated parameters.

### 3 A first diffusion model

We will first review the diffusion model established by Tirabassi (1989). Analytical models such as this one use solutions to the advection-diffusion equation below:

$$u \frac{\partial C}{\partial x} = \frac{\partial}{\partial y} (K_y \frac{\partial C}{\partial y}) + \frac{\partial}{\partial z} (K_z \frac{\partial C}{\partial z}) + S \quad (8)$$

where the parameters are:

- Mean velocity ( $u$ )
- Concentration ( $C$ )
- Source term ( $S$ )
- Lateral eddy exchange constant ( $K_y$ )
- Vertical eddy exchange constant ( $K_z$ )

### 4 InMAP: A more computationally efficient model for finding emissions

One of the recent developments in the study of air pollution was the creation of computationally efficient models for computing the harm to human health from diffusion. A prevalent model that has been implemented in industry by companies such as WattTime is InMAP (<http://spatialmodel.com/>). InMAP is a model that evaluates pollution by simplifying typical CTMs via using an annual-average basis for its results.