



Food Science and Culinary Techniques - FSCT

**Notes for the course NFOK13004U, at the University of
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Link to Git repo.: https://github.com/DanishUnicorn/fsct_notes



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Chapter 1

Course Description

1.1 Content

The aim of the course is to use a basic understanding of food chemistry and physics to obtain a scientific approach to cooking when various culinary techniques are applied during processing of foods.

The course includes a series of lectures giving a scientific description of foods as a chemical and physical system. It relates to proteins, lipids, carbohydrates as well as topics within general chemistry (inorganic and organic), acids and bases, and interactions of these components. The course provides an understanding of the culinary techniques used in the production of foods and highlights the effects of food processing on the chemical reactions leading to changes of flavour/taste and colour as well as the physical properties of the food in relation to changes in structure and functionality.

Practical kitchen exercises in preparation of foods will be used as a learning tool to the understanding of culinary techniques. The use of ingredients in various recipes will be evaluated and thereby demonstrate important experimental aspects of food processing and preparation. This will include an introduction to experimental design where recipes and preparations are varied, and the methods of evaluation are identified. The following afterlab discussions will reflect on the outcome of the experiments and correlate it to the scientific principles of the exercise.

The lectures and theoretical exercises will demonstrate how food components contributes to the functional properties in dry systems, crystalline states, emulsions, foams and other real food systems. During the practical kitchen exercises students will evaluate the different preparations in relation to texture, flavour/taste and colour, and explain the outcome according to the theory.

1.2 Learning Outcome

A student who has fulfilled the aim of the course should be able to:

1.2.1 Knowledge

- Describe important chemical reactions and physical changes during processing of foods
- Describe carbohydrates, lipids and proteins and their basic functions and characteristics in food and point out the effects of culinary processes on physical, chemical and sensory conditions of food components
- Describe the effect of physical processes on the structure of food during cooking
- Identify factors of relevance for detection, perception and loss of aroma and flavour compounds in different solvents.

1.2.2 Skills

- Work in a gastronomic laboratory with specific culinary techniques and follow instructions to obtain a well-defined product
- Explain the changes in foods taking place during preparation of food from a chemical and a physical point of view
- Predict the outcome of various preparation methods and recipes based on a simple experimental design
- Evaluate a complex food and communicate the compositional structure of the product
- Evaluate the effect of various culinary techniques on the food structure and flavour
- Ability to perform simple statistical analyses using Excel or other related software
- Ability to give and receive peer feedback
- Ability to communicate scientific topics within food science and culinary techniques in (academic) English.

1.2.3 Competences

- Plan experiments related to the effect of a culinary technique on the sensory properties of food
- Integrate scientific disciplines (food chemistry and food physics) in planning and evaluation of practical experiments
- Cooperate with other students on planning and performing practical exercises including written and oral evaluation of the theoretical outcome through afterlab discussions at plenary sessions.

1.3 Literature

See Absalon for a list of course literature.

1.4 Recommended Academic Qualifications

A basic knowledge of chemical reactions involving carbohydrates, proteins and lipids as well and basic statistics is highly recommended.

Academic qualifications equivalent to a BSc degree is recommended.

Table 1.1: A table with an overview over the workload for the course.

Category	Hours
Lectures	43
Preparation	110
Theory exercises	28
Practical exercises	21
Exam	4
Total	206

1.5 Teaching and Learning Methods

The teaching use a general understanding of food chemistry and physics in combination with practical kitchen exercises in a gastronomical laboratory to examine the influence of various processing methods on the food components. The practical kitchen exercises set the frame for group work and will be evaluated by afterlab discussions, problem-based learning and answering questions from the lecturers. The course also includes mandatory written assignments with peer feedback (student-student) based on the practical exercises. Specific practical exercises might be organised as take-home exercises where ingredients and tools will be provided.

1.6 Remarks

It is recommended to follow the course on the first year of the MSc Programme in Food Innovation and Health.

1.7 Workload

1.8 Feedback Form

- Oral
- Collective
- Continuous feedback during the course of the semester
- Peer feedback (Students give each other feedback)

1.9 Sign Up

Self Service at KUNet

<http://www.science.ku.dk/english/courses-and-programmes/>

<https://www.science.ku.dk/english/continuing-and-professional-education/single-subject-courses/practical/>

1.10 Exam

Table 1.2: The table shows the details of the course exam, as defined from the website of the University of Copenhagen.

Category	Details
Credit	7.5 ECTS
Type of assessment	On-site written exam, 4 hours under invigilation
Type of assessment details	Individual written 4 hour exam on specific topics based on the course curriculum. The on-site written exam is an ITX exam. See important information about ITX-exams at Study Information, menu point: Exams -> Exam types and rules -> Written on-site exams (ITX)
Exam registration requirements	Approval of all assignments for all practical kitchen exercises.
Aid	All aids allowed. The University will make computers available to students at the ITX-exam. Students are not permitted to bring digital aids like computers, tablets, mobile phones etc. Students are, however, allowed to bring a calculator. Books, notes, and similar materials can be brought in paper form or uploaded before the exam and accessed digitally from the ITX computer. Read more about this at Study Information.
Marking scale	7-point grading scale
Censorship form	No external censorship. Several internal examiners.
Re-exam	Same as ordinary exam. Possibility to edit and re-submit all non-approved assignments from the practical kitchen exercises two weeks before the re-exam. If 10 or fewer register for the re-examination the examination form will be oral. The oral exam will be 20 minutes in the course curriculum, no preparation time and all aids allowed.

Criteria for exam assessment See Learning Outcome.

Chapter 2

Basic Knowledge Recap

This section of the course notes is designed to streamline access to the key findings from each reading material (RM), providing a concise and accessible overview of essential information. Created through experimentation with various AI platforms, this chapter also serves to enhance my prompt engineering skills, exploring diverse methods of note-taking for maximum efficiency and clarity. The procedures for creating these summaries have varied, but all methods share a common approach: each RM has been fully read, with summaries and notes prepared after completing each respective subsection. By using these AI-co-op'ed approaches, these notes aim to be both a reliable reference and a resource for continuous improvement in capturing complex microbiology concepts.

2.1 1st RM

Introduction

Chemistry evolved over a millennia, merging with biology to form diverse fields like **food chemistry**, which focuses on sustenance and nutrition. This chapter outlines chemistry's relation to food sciences and its educational context, addressing challenges like increased teaching loads and exploring new opportunities [1].

On the Need for Chemistry

Food chemistry studies the composition, structure, and properties of food and their transformations. It builds on general chemistry, including **organic**, **physical**, and **inorganic** branches. This chapter surveys these areas and introduces key terms for future topics [1].

Organic Chemistry

Organic chemistry focuses on carbon compounds and **covalent bonds** with hydrogen. Hydrocarbon nomenclature (e.g., **alkanes**, **alkenes**, and **alkynes**) introduces systematic naming. Topics include bonding theories, **valence**, and molecular conformation constraints [1].

Functional Group Chemistry

Organic compounds belong to families like **alkanes**, **alcohols** (R-OH), and **carboxylic acids** (R-COOH). Functional groups have consistent behaviour, e.g., **hydroxyl** groups act similarly across compounds. Their reac-

tivity may vary with the R group, as in alcohol vs. phenol [1].

Aromatic Compounds

Aromaticity in compounds like benzene involves **electron delocalisation** and resonance. Conjugated double bonds in unsaturated lipids and fat-soluble vitamins contribute to their color and oxidative sensitivity. Free radicals drive polymerization of natural monomers like **isoprene** [1].

Organic Reaction Mechanisms

Organic reactions include **addition**, **substitution**, and **rearrangement**, often occurring in raw and processed foods. **Free radical** reactions drive polymerization and impact **lipid oxidation**, photochemical processes, and ageing-related deterioration [1].

Stereochemistry

Stereochemistry studies chiral compounds, often with asymmetric carbons attached to four groups. Molecules with N stereogenic centers yield (2^N) stereoisomers. Chirality includes helices lacking stereogenic centers. Chiral molecules rotate plane-polarized light, but **racemic mixtures** show no net optical activity [1].

Physical Chemistry

Physical chemistry explores material transformations, focusing on three key themes: **thermodynamics**, **chemical kinetics**, and **quantum mechanics** [1].

Thermodynamics

Thermodynamics examines energy forms, transformations, and efficiency. The **first law** states energy is conserved, expressed as ($U = Q + W$). The **second law** highlights entropy increase, with **Gibbs free energy** (ΔG) predicting spontaneity. Applications include **reaction calorimetry** and energy content in food [1].

Chemical Kinetics

Chemical kinetics studies reaction rates, influenced by factors like concentration, temperature, and catalysts. Reaction rates follow Equation 2.1.

$$\text{rate} = k[A]^x[B]^y \quad (2.1)$$

With (k) described by the **Arrhenius equation** which can be seen from Equation 2.2

$$\ln k = \ln k_0 - \frac{\Delta E^\#}{RT} \quad (2.2)$$

The **Eyring theory** relates k to $\Delta G^\#$, $\Delta H^\#$, and $\Delta S^\#$, emphasizing transition states [1].

Inorganic Chemistry

Inorganic chemistry covers non-carbon elements, including metals, nonmetals, and metalloids. Periodic trends, such as **atomic radii**, **ionization energies**, and **electronegativity**, arise from effective nuclear charge and electron shielding [1].

Chemical Bonding

Chemical bonding can be described with **Lewis structures** which illustrates electron sharing or transfer. The **octet rule** explains covalent bonds, while differences in electronegativity lead to polarized or ionic bonds. **Resonance** applies to compounds like nitrate, where multiple structures describe electron distribution [1].

The Shapes of Molecules

Molecular shapes are predicted using the **VSEPR model**, based on repulsions between electron pairs. Geometries like **linear**, **trigonal planar**, **tetrahedral**, and **octahedral** depend on bonded atoms and lone pairs, as seen in molecules like H_2O and NH_3 [1].

Valence

Valence-bond theory explains covalent bonds through orbital overlap. Some valencies require **hybrid orbitals** (e.g., sp^1 , sp^2 , sp^3), formed via electron promotion and orbital mixing. Orbital shapes are influenced by electron repulsion [1].

Molecular Orbital Theory

Molecular orbital theory explains bonding via atomic orbital combinations forming equal numbers of bonding σ and anti-bonding σ^* orbitals. Molecular orbitals follow **Pauli's exclusion principle** and fill singly before pairing. Diatomic oxygen illustrates these principles [1].

Food Chemistry

Definition and Scope

Food chemistry applies chemistry principles to food systems, studying macroconstituents (e.g., **water**, **carbohydrates**) and microconstituents (e.g., **vitamins**, **additives**). Post-WWII advances improved **shelf-life**, **packaging**, and analysis of toxicants. Food sciences ensure nutritious, safe, and affordable food [1].

Areas of Expertise Required by the Institute of Food Technology

The **Institute of Food Technology** identifies five core competencies for food scientists: **chemistry**, **analysis**, **nutrition**, **microbiology**, and **engineering** [1].

Chemistry and the Food System

Food chemistry encompasses all levels of the food system, from **soil pH** in agronomy to the effects of **pasteurization**. Its scope includes **harvesting**, **processing**, **packaging**, and **distribution**, as well as studying **ingredient behaviour** during manufacture [1].

2.2 2nd RM

Osmosis and Osmotic Pressure

Osmosis involves water moving to solute-rich areas across membranes. **Osmotic pressure** drives this process, dehydrating microbes in salted foods. Preservation methods like salting meats and sugaring jams rely on this principle, as seen in **beef jerky** and jellies [1].

Carbohydrates

Foods High in Carbohydrates

Carbohydrates, including sugars, starches, and fibers, primarily originate from plants. Sources include **grains**, **legumes**, **fruits**, and **vegetables**. Exceptions are **milk**, containing lactose, and animal muscles, storing glycogen. Table sugar comes from **sugar cane** and **sugar beets**, while honey is floral nectar [1].

Composition of Carbohydrates

Carbohydrates consist of **carbon**, **hydrogen**, and **oxygen**, following the formula $C_n(H_2O)_n$. Synthesized via **photosynthesis**, they form saccharides classified as **monosaccharides**, **disaccharides**, **oligosaccharides**, and **polysaccharides** based on saccharide units [1].

Chapter 3

Laboratory Exercises

Introduction

In this chapter a summary of the laboratory exercises is given. The exercises are designed to give the student a practical understanding of the theoretical concepts discussed in the previous chapters. The exercises are divided into two parts: the first part deals with the basic techniques of microbiology, while the second part deals with the identification of bacteria.

Chapter 4

Lecture Exercises

4.1 02.09.24 - Exercise 1 - CasePCR

Chapter 5

Literature résumés

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5.1 1st lecture

5.1.1 Article 1 - Fermented Foods as Experimentally Tractable Microbial Ecosystems

Introduction

Chapter 6

Exam

Chapter 7

Abbreviations and Explanations

Topic	Abb.	Description
16S ribosomal RNA	16S rRNA	<i>A component of the 30S subunit of prokaryotic ribosomes, commonly used in phylogenetic studies to identify bacteria and archaea.</i>

Bibliography

- [1] Richard Owusu-Apenten. *Introduction to Food Chemistry*. Boca Raton, FL: CRC Press, 2004.

Chapter A

Appendix

A.1 Appendix 1 - Principles for isolation of microorganisms from fermented food and beverages