$$AH + H_2O \underset{}{\overset{Ka}{\rightleftharpoons}} H_3O^+ + A^-$$ Réaction

$$Ka \equiv \frac{[H_3O^+][A^-]}{[AH] \times C^\circ}$$

$$PH = PKa \Rightarrow \frac{[AH]}{[A^-]} = 1$$

AH

1

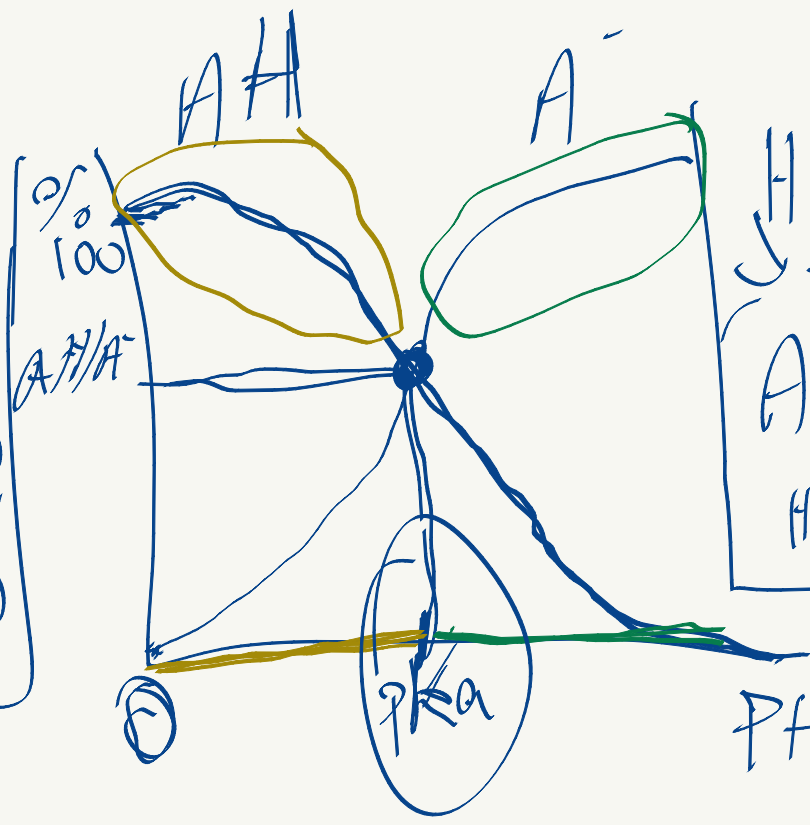
H_2O	H_2O
H_2O	H_2O

2

H_2O	H_3O^+
H_3O^+	H_3O^+

1

AH	AH
H_3O^+	H_3O^+
A^-	A^-
	H_2O



2

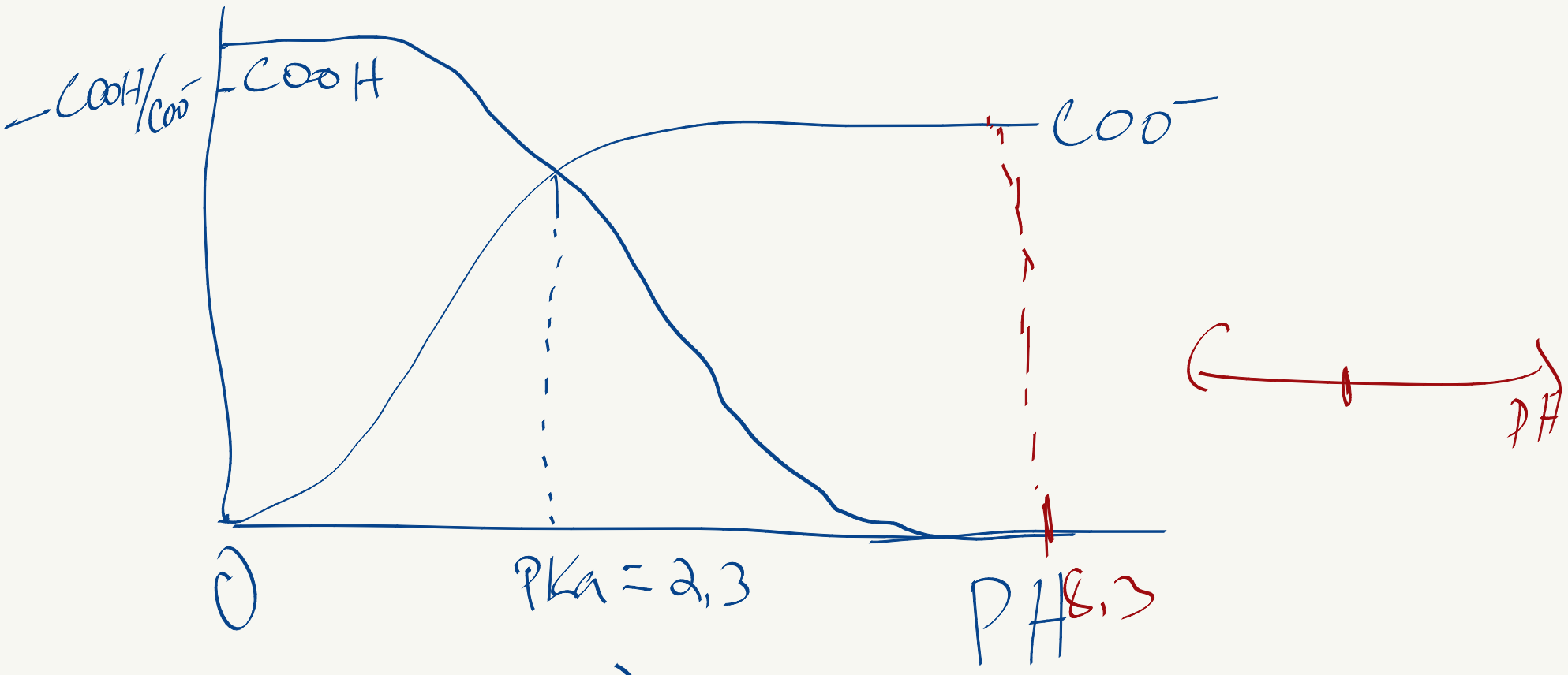
H_3O^+	H_3O^+
A^-	H_3O^+
H_2O	H_3O^+

AH

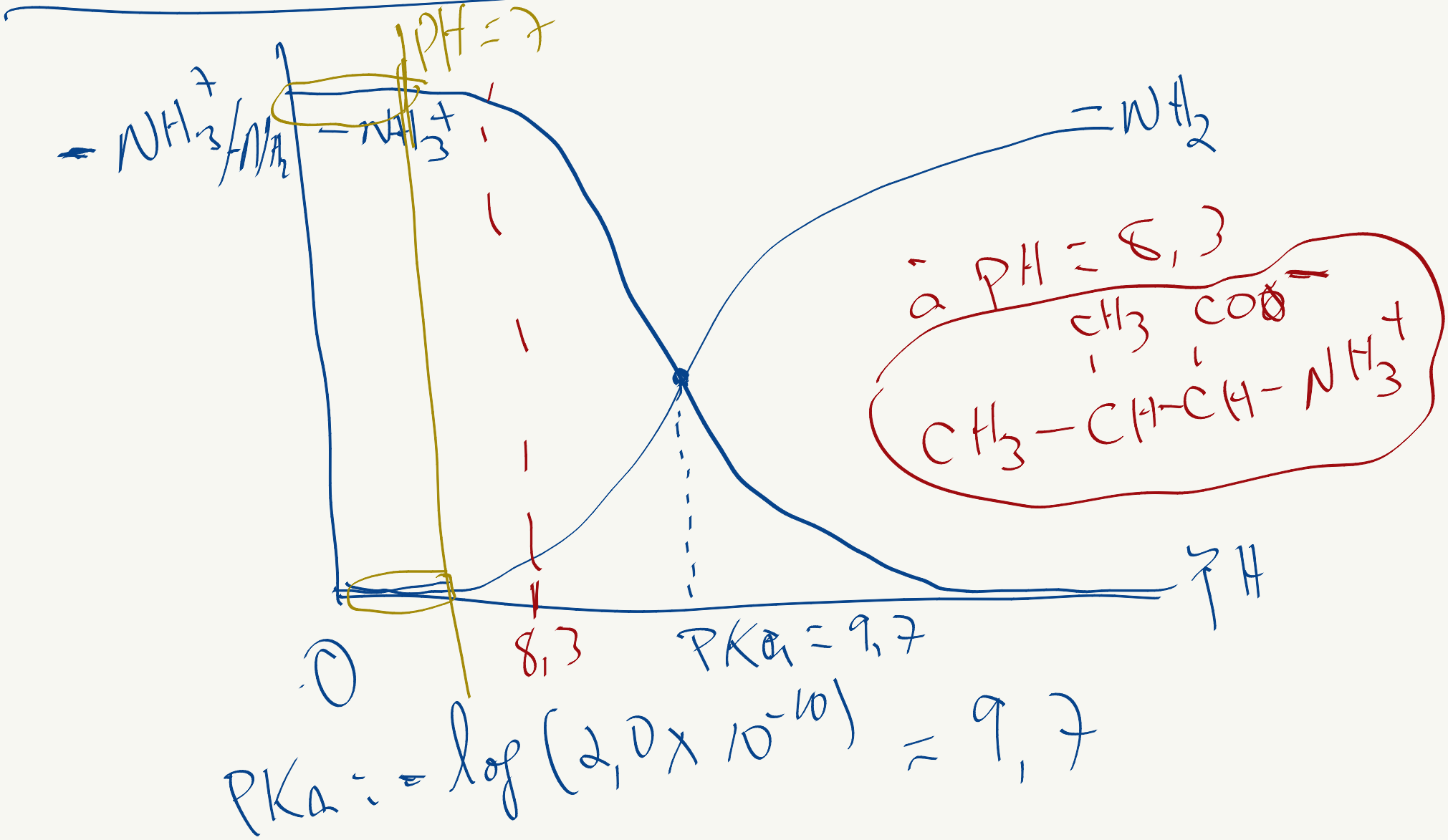
1'

2'

Diagr. de predominance distribuição



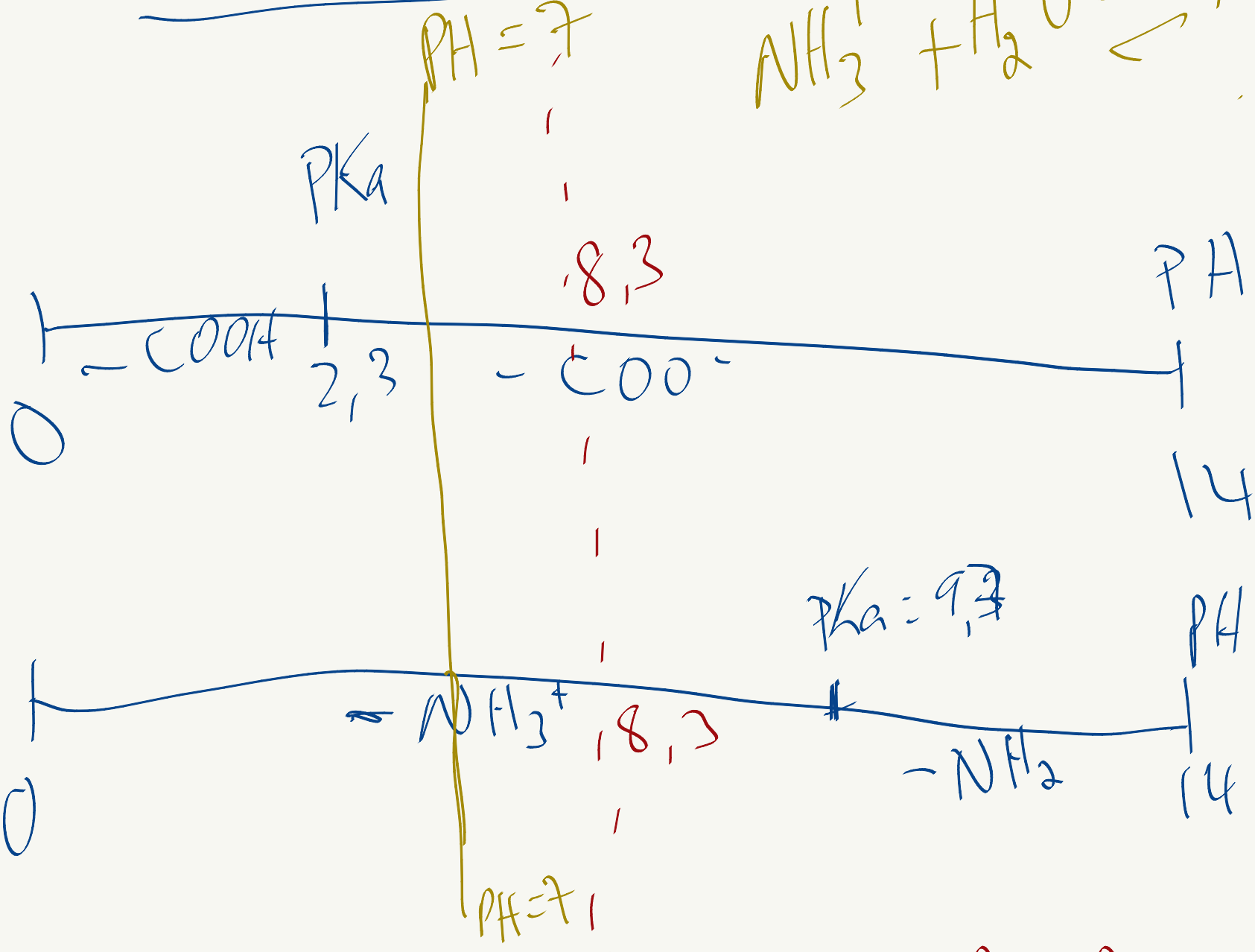
$$-\log(5{,}0\times10^{-3}) = 2{,}3$$



$$PKa = -\log(2{,}0\times10^{-10}) \approx 9{,}7$$

Predominance

$NH_3^+ + H_2O \rightleftharpoons NH_2 + H_3O^+$



Donc, valine sera sous la forme