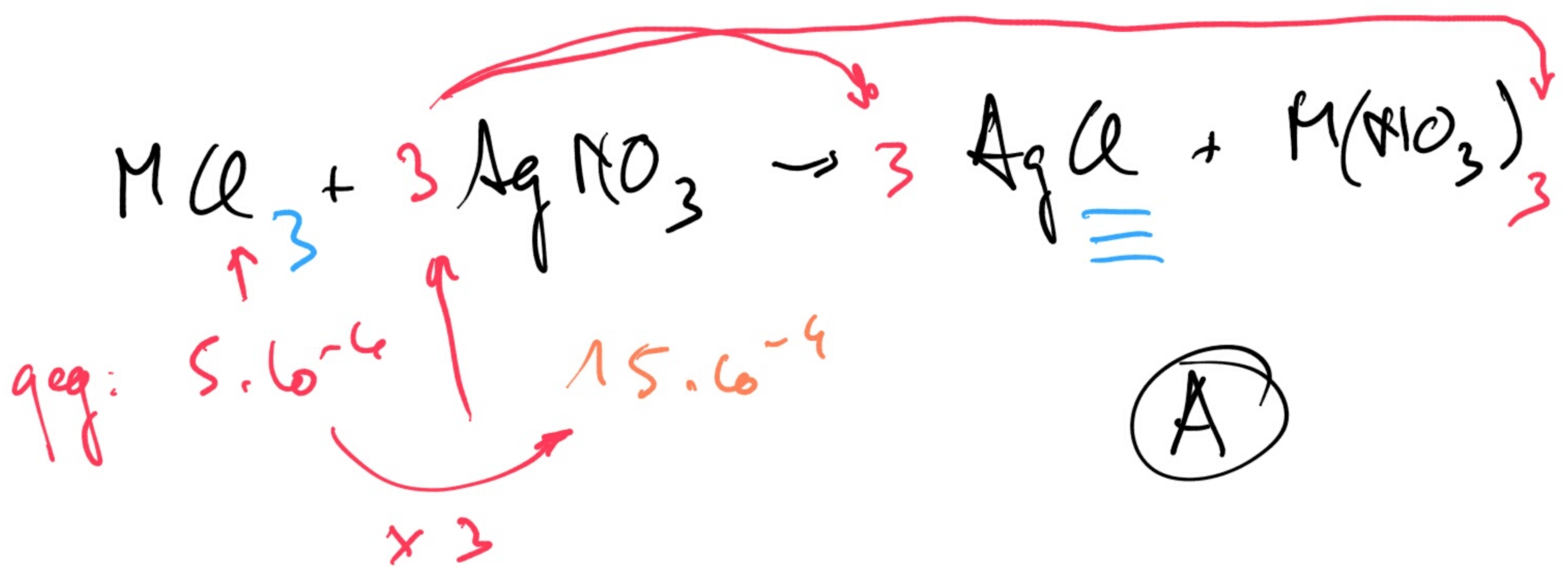


① 60 ml $2,5 \cdot 10^{-2} \text{ mol/l AgNO}_3$

$\Rightarrow 60 \cdot 10^{-3} \cdot \frac{5}{2} \cdot 10^{-2} = 150 \cdot 10^{-5} \text{ mol AgNO}_3$



② Antimon Sb : atoom w.: 51

$\Rightarrow \# p^+ = 51$

$p^+ + n = 51 + 72 = 123$ $\begin{matrix} \beta \\ \gamma \\ \alpha \end{matrix}$

neutraal: $p^+ = e^- = 51$

$51 - 48 = 3 e^-$ tekort!

$\Rightarrow 123 \text{ Sb}^{3+}$

(B)

③ $\text{K}_2\text{O} \rightarrow \# e^-$ $2 \cdot 8 (\text{H}) \leftrightarrow 2 \cdot 5 (\text{H})$

$+ 1 \cdot 8 (\text{O}) \leftrightarrow 1 \cdot 6 (\text{O})$

$\Rightarrow \text{H is III} \rightarrow \text{geen octet!}$

24

16 = 8

(D)

1×11

\leftarrow

4 bindingen

\leftarrow

$$\textcircled{4} \quad K_P = 174^\circ \text{C}$$

$$227^\circ \text{C} \quad 1000 \text{ hPa} \rightarrow \rho = 3,42 \text{ g/l}$$

$$\downarrow$$

$$+ 273 = 500 \text{ K}$$

$$= 3,42 \text{ kg/m}^3$$

$$\rho \cdot V = n \cdot R \cdot T \quad \text{en} \quad n = \frac{m}{M}$$

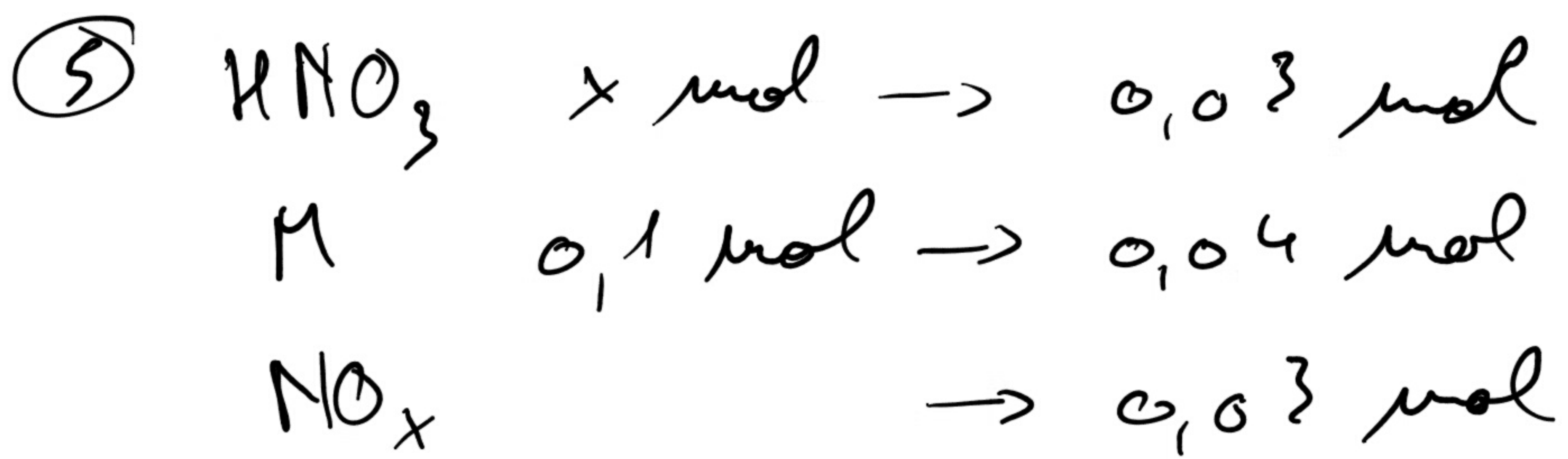
$$\rho \cdot V = \frac{m}{M} \cdot R \cdot T$$

$$\Rightarrow M = \frac{m \cdot R \cdot T}{\rho \cdot V} = \rho \cdot \frac{R \cdot T}{\rho}$$

$$= \frac{3,42 \cdot 8,31 \cdot 500}{1000 \cdot 10^2}$$

\textcircled{B}

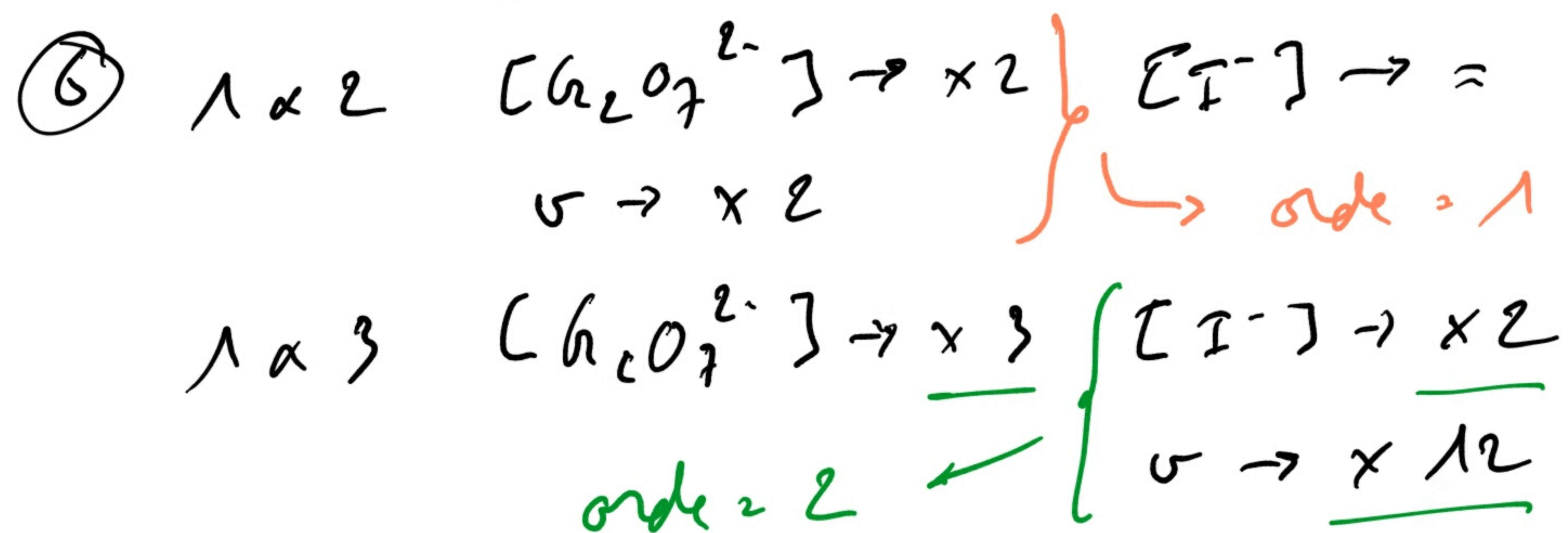
$$= 0,1421 \text{ kg/mol}$$



\Rightarrow mol M gebildet

$$= 0,1 - 0,04 = \underline{\underline{0,06}}$$

$$\frac{M}{NO_x} \approx \frac{2}{1} \text{ enkel in } \textcircled{D}$$



A



V_2 (st) en $T \uparrow \rightarrow$ reactie naar links!

\Rightarrow # mol blijft gelijk!

$A \uparrow$ en $C \downarrow$ $\Delta n_{\text{gas}} = 0$!

grafiek 4 A

⑧ ? Waer is $pH > 3,5$?

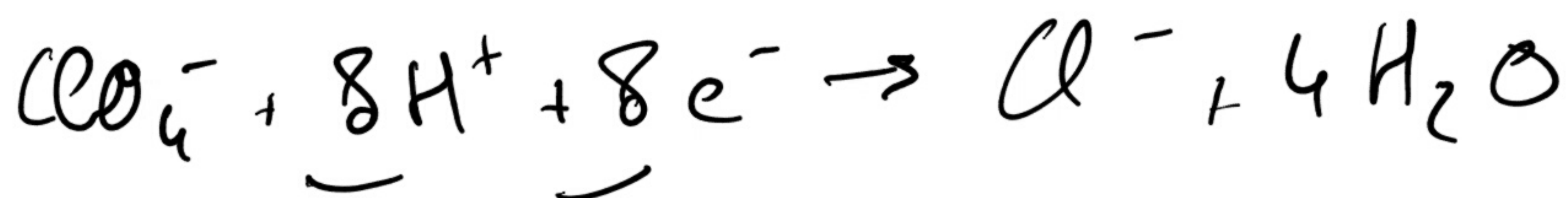
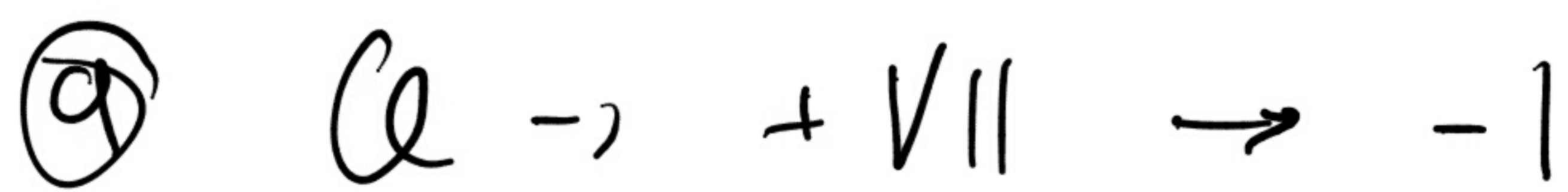
concentratie
 \downarrow

$HCl \rightarrow$ zuur \rightarrow splitst volledig: $pH = -\log(0,1) = 1$

$NaCl \rightarrow$ neutraal zout $\rightarrow pH = 7 \rightarrow$ rood

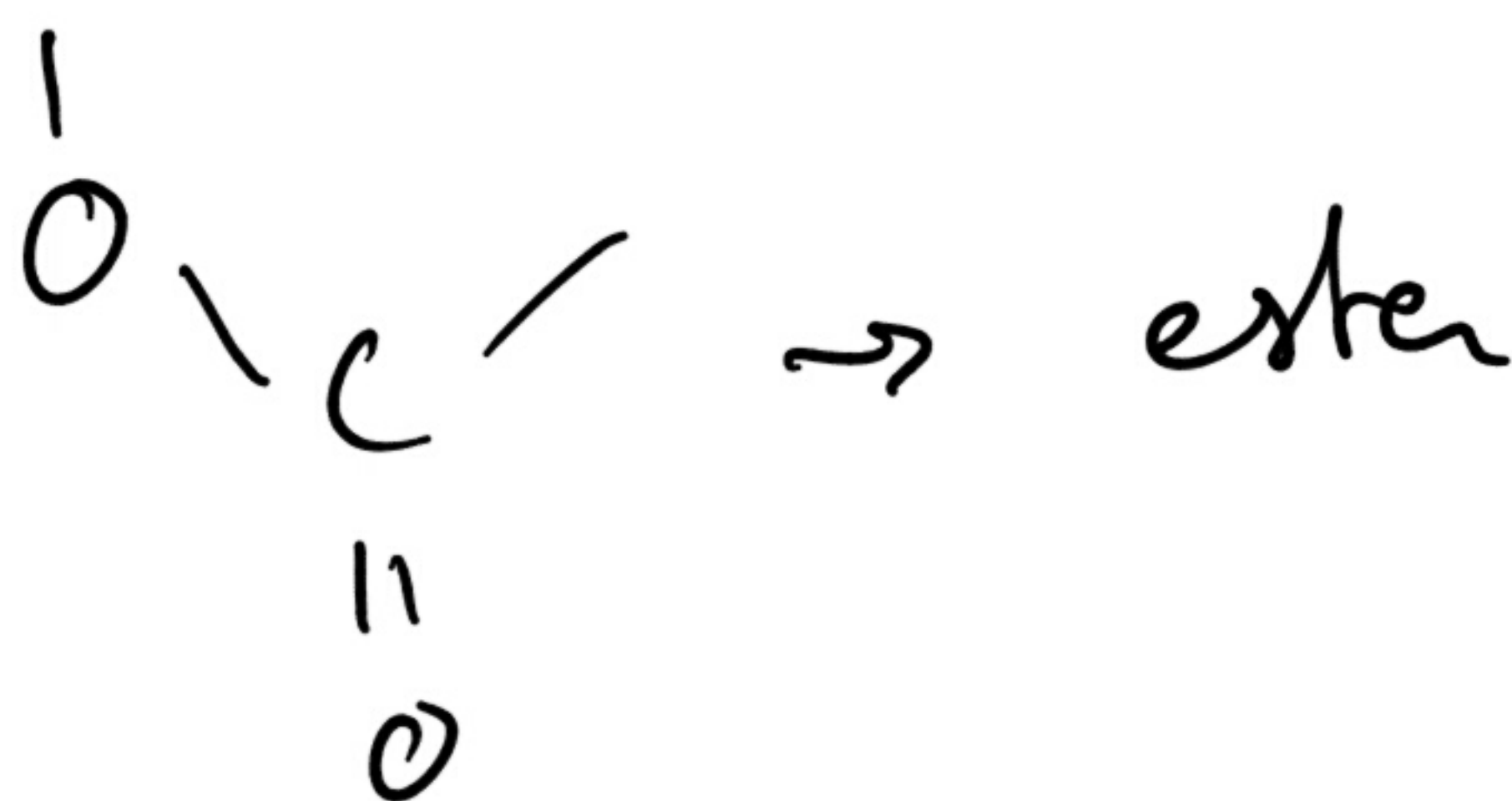
$NaOH \rightarrow$ base \rightarrow splitst volledig: $pOH = -\log(0,1) = 1$
 $\hookrightarrow pH = 13 \rightarrow$ rood

$\Rightarrow 2 \text{ \& } 3 \rightarrow$ B



③

reductre: cythane car $8e^-$



⑪