BIOL 1110- 6.1.1

[Ch.6]
Potential is energy not moving
Kinetic is energy moving

Energy is the capacity to do work

ATK / 1100 DEF / 1200

22699248 ©2017 TRAVIS

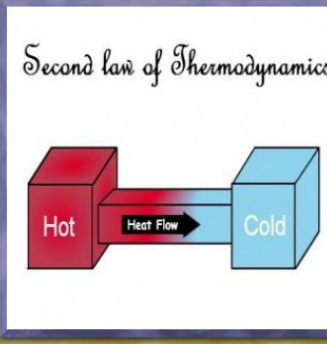
BIOL 1110- 6.2.1

[Ch.6]
First Law of Thermodynamics states that energy cannot be created nor destroyed; Also, energy in the universe is constant

ATK / 2500 DEF / 1900

24410840 ©2017 TRAVIS

BIOL 1110- 6.2.2

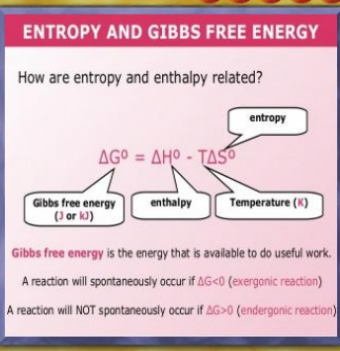


[Ch.6]
Second Law of Thermodynamics states that disorder increases via the more energy transfers from place to another

ATK / 2700 DEF / 1500

89517384 ©2017 TRAVIS

BIOL 1110- 6.2.3 #1

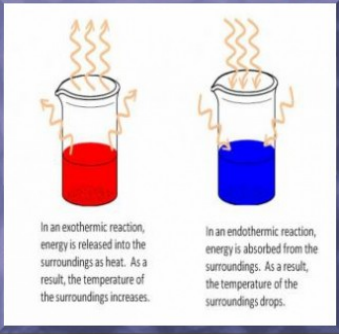


[Ch.6]
Entropy is disorder
Enthalpy is change in heat
Free Energy is energy available to do work
Positive Change in G= More free energy in Products
(-) Change in G= More free energy in Reactants

ATK / 1600 DEF / 1400

50333683 ©2017 TRAVIS

BIOL 1110- 6.2.3 #2



[Ch.6]
Spontaneous means that the reaction occurs completely without outside influence

ATK / 1500 DEF / 1300

99322961 ©2017 TRAVIS

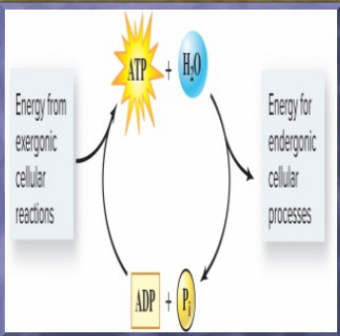
BIOL 1110- 6.7

[Ch.6]
3 works done by cells are: Transport, Chemical Reactions, and Movement

ATK / 1700 DEF / 1600

27222476 ©2017 TRAVIS

BIOL 1110- 6.3.1 #1

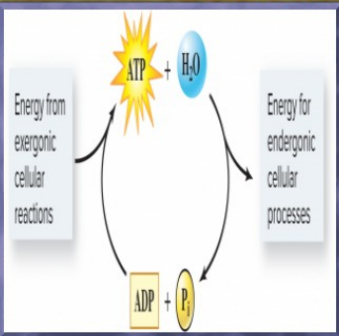


[Ch.6]
ATP Hydrolysis and an endothermic reaction turns into two exergetic reactions because ATP hydrolysis produces more energy than the endothermic consumes

ATK / 1500 DEF / 1800

17944768 ©2017 TRAVIS

BIOL 1110- 6.3.1 #2



[Ch.6]
Energy Coupling is using ATP hydrolysis and endothermic reactions to make two exergetic reactions that produce a lot of energy for cellular needs

ATK / 1800 DEF / 1500

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