Nirandika Jayawardena

-Advanced Level Chemestry-

Answer the Following questions in the space provided.

Submit your answers through the link provided.

1. Define the term "ideal gas" and list three assumptions made about ideal gases in the kinetic molecular theory.

(10 marks)

2. Explain Boyle's Law and provide an example illustrating its application in real-life situations.

(10 marks)

3. Describe the factors that affect the rate of effusion of gases according to Graham's Law. Provide an example to demonstrate the concept.

(10 marks)

4. Discuss the concept of partial pressure in a gas mixture. How is the total pressure of a gas mixture calculated?

(10 marks)

5. Explain why real gases deviate from ideal behavior at high pressures and low temperatures. Discuss two factors that contribute to these deviations and provide examples of gases that exhibit such behavior.

(10 marks)