CHEMISTRY JULY TEST M.M - 25

MULTIPLE CHOICE QUESTIONS

1) A measured temperature on Fahrenheit scale is 200° F .what will be this reading be on Celsius scale?

a) 40°C b) 94°C

b) 93.3°C d) 30°C

2) What is the mass percent of carbon in carbon dioxide.

a) 0.034 % b) 27.27%

c) 28.7% d) 3.4%

2) Is the density of a solution is 3.1g/ mL the mass of 1.5 mL solution is

a)4.7 g b) 4680 × 10^{-3} g

c) 4.680 g d) 46.80 g

3)The empirical formula and Molecular mass of a compound are CH2O and 180 g respectively .what will we the moleculer formula of the compound?

a)C9H18O9 b) CH20

c)C6H12O6 d) C2H4O2 .

4)one mole of oxygen gas at STP is equal to.

a)6.022×10^{23} molecules of oxygen b) 6.022×10^{23} atom of oxygen.

c)16g of Oxygen. d) 32g of Oxygen.

5)How many moles of nitrogen are needed to produce 8.2 mole of ammonia by reaction with hydrogen?

a) 2.1mol of Hydrogen b) 4.1mol of nitrogen

c)4.1 mol of Hydrogen d) 4.2 mol of Ammonia

6)at what temperature have the Celsius and Fahrenheit reading the same numerical value.

a) +40 b) -40

c) 25 d) none of these.

7)Calculate the mass of one molecule of water

a)6.022×10^{23}g b) 18g

c)2.99×10-23 g d) 108.39 ×10^{23}g

8)In a reaction A+B2 =AB2

Identify the limiting reagent in 2.5 mol of A+5.0 mol of B.

a)A b) B2

c)A2 d) B

9) Calculate the molarity of pure water (density of water =1g/ml

a)55.0M b) 55.55M

c) 1M d) 1000M