

OMEN-SED 0.9: A novel, numerically efficient organic matter sediment diagenesis module for coupling to Earth System Models

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Abstract. We present the first version of a new, one-dimensional analytical early diagenetic model resolving organic matter cycling and associated biogeochemical dynamics in marine sediments called OMEN-SED (Organic Matter ENabled SEDiment model) designed to be coupled to Earth system models. OMEN-SED explicitly describes organic matter cycling as well as associated dynamics of the most important terminal electron acceptors (i.e. O₂, NO₃, SO₄), related reduced substances (NH₄, H₂S), the full suite of secondary-redox reactions, macronutrients (PO₄) and associated pore water quantities (ALK, DIC). To represent a redox-dependent sedimentary P cycle we also include a representation of the formation and burial of Fe-bound P and authigenic Ca-P minerals. Thus, OMEN-SED captures most of the features of a complex, numerical diagenetic model. However, its computational efficiency allows the coupling to global Earth System models and therefore the investigation of coupled global biogeochemical dynamics over different timescales. This paper provides a detailed description of the new sediment model, as well as an extensive sensitivity analysis using a novel method which considers the entire cumulative distribution functions of the model outputs. OMEN-SED's performance is evaluated through comprehensive comparisons with observations and against a more complex analytical model. We find solid phase and dissolved porewater profiles for different ocean depths are reproduced with good accuracy and modelled terminal electron acceptor fluxes fall well within the range of globally observed fluxes. Finally, we couple OMEN-SED to the Earth System model cGENIE and tune the organic matter degradation rate constants to optimise the fit of modelled OM concentrations to a global database. We find simulated sediment characteristics of the coupled model framework (such as OM degradation rates, oxygen penetration depths and

sediment-water interface fluxes) are generally in good agreement with observations and in line with what one would expect on a global scale. AR: Final sentence to wrap up, maybe future look...

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1 Introduction

DH: Will delete the sub-headings!

Role of marine sediments for climate and global biogeochemical cycles:

- 65 Marine surface sediments are key components in the Earth system. They host the largest carbon reservoir within the surficial Earth system, provide the primary long term sink for atmospheric CO₂, recycle nutrients, and represent the most important geochemical archive used for deciphering past changes of biogeochemical cycles and climate (e.g. Berner, 1991; Archer and Maier-Reimer, 1994; Ridgwell and Zeebe, 2005; Arndt et al., 2013). Physical and chemical processes in sediments (i.e.
- 70 diagenetic processes) depend on the water column and vice versa: Diagenesis is controlled by the external supply of solid material (e.g. organic matter, calcium carbonate, opal) from the water column and is affected by overlying bottom water concentrations of solutes. At the same time, sediments impact the water column directly either by short- and long-term storage of deposited material or diagenetic processing of deposited material and diffusion of some of the resulting products (e.g.
- 75 nutrients, DIC) to the overlying bottom waters. This so-called benthic-pelagic coupling is essential for understanding global biogeochemical cycles and climate (e.g. Archer and Maier-Reimer, 1994; Archer et al., 2000; Soetaert et al., 2000; Mackenzie, 2005).

Biological primary production of organic matter (OM, generally represented by CH₂O in equation R1) and the reverse process of degradation can be written in a greatly simplified reaction as:



On geological timescales production of OM is generally greater than degradation which results in some organic matter being buried in marine sediments and oxygen accumulating in the atmosphere. Thus burial of OM leads to net oxygen input to, and CO₂ removal from the atmosphere (Berner, 2004). On shorter timescales, the upper few meters of the sediments where early diagenesis occurs are specifically important, as this zone controls whether a substance is recycled to the water column or buried for a longer period of time in the deeper sediments (Hensen et al., 2006). Most biogeochemical cycles and reactions in this part of marine sediments can be related either directly or indirectly to the degradation of organic matter (Middelburg et al., 1993; Arndt et al., 2013). Oxygen and nitrate for instance, the highest energy yielding electron acceptors, are preferentially consumed in the course of the degradation of organic matter, resulting in the release of ammonium and phosphorus to the pore water. As such, degradation of OM in the sediments can profoundly affect the oxygen and nutrient inventory of the ocean and thus primary productivity (Van Cappellen and Ingall, 1994; Lenton and Watson, 2000). Furthermore, organic matter degradation releases metabolic CO₂ to the pore water, causing it to have a lower pH and carbonate ion concentration, thus provoking the dissolution of calcium carbonate CaCO₃ (Emerson and Bender, 1981).

Nutrient recycling from marine sediments has been suggested to play a key role for climate and ocean biogeochemistry throughout Earth history. For example, feedbacks between phosphorus storage and erosion from shelf sediments and marine productivity have been hypothesised to play an im-

portant role for glacial/interglacial atmospheric CO₂ changes (Broecker, 1982; Ruttenberg, 1993).
100 Furthermore, nutrient recycling from anoxic sediments has been invoked to explain the occurrence of more extreme events in Earth history, for instance Oceanic Anoxic Events (OAEs, e.g. Mort et al., 2007; Tsandev and Slomp, 2009). OAEs represent severe disturbances of the global carbon, oxygen and nutrient cycles of the ocean and are usually characterized by widespread bottom water anoxia and photic zone euxinia (Jenkyns, 2010). One way to explain the genesis and persistence of OAEs
105 is increased oxygen demand due to enhanced primary productivity. Increased nutrient inputs to fuel primary productivity may in turn have come from marine sediments as the burial efficiency of phosphorus declines when bottom waters become anoxic (Ingall and Jahnke, 1994; Van Cappellen and Ingall, 1994). The recovery from OAE like conditions is thought to involve the permanent removal of excess CO₂ from the atmosphere and ocean by burying carbon in the form of organic matter in
110 marine sediments (e.g. Arthur et al., 1988; Jarvis et al., 2011), which is consistent with the geological record of widespread black shale formation (Stein et al., 1986). Models capable of simulating the expansion and intensification of oxygen minimum zones and how the underlying sediments interact are hence needed.

Diagenetic Models in general:

115 Quantifications of biogeochemical processes in the sediments are possible through the application of idealised mathematical representations, or so-called diagenetic models (see e.g. Berner, 1980; Boudreau, 1997). A plethora of different approaches have been developed, mainly following two distinct directions (Arndt et al., 2013). First, state-of-the art vertically resolved numerical models simulating the entire suite of essential coupled redox and equilibrium reactions within marine sediments (e.g. BRNS, Aguilera et al., 2005; CANDI, Boudreau, 1996; MEDIA, Meysman et al., 2003;
120 MUDS, Archer et al., 2002; STEADYSED, Van Cappellen and Wang, 1996). These “complete”, non-steady-state models, thus resolve the resulting characteristic redox-zonation of marine sediments through explicitly including oxic OM degradation, denitrification, oxidation by manganese and iron (hydr)oxides, sulfate reduction and methanogenesis as well as the reoxidation of reduced
125 byproducts (i.e. NH₄, Mn²⁺, Fe²⁺, H₂S, CH₄, see e.g. Regnier et al., 2011; Arndt et al., 2013). Furthermore, they incorporate various mineral dissolution and precipitation reactions, as well as fast equilibrium sorption processes for example of NH₄, PO₄ and metal ions (i.e. Mn²⁺, Fe²⁺ and Mg²⁺, compare Van Cappellen and Wang, 1996; Meysman et al., 2003). Modelled, depth-dependent, transport processes usually comprise advection, diffusion, bioturbation and bio-irrigation. This group
130 of diagenetic models generally uses a so-called multi-G approach (Jørgensen, 1978; Berner, 1980), thus dividing the bulk organic matter pool into a number of compound classes that are characterised by different degradabilities k_i , which are generally dependent on the type and concentration of the specific terminal electron acceptor (TEA). Alternative approaches, in particular reactive continuum models (Boudreau and Ruddick, 1991), assume a continuous distribution of reactive types but are
135 much less often used. These complex, “complete” models have a great potential for quantifying OM

degradation dynamics for sites where enough observations are available to constrain its model parameters (see e.g. Boudreau et al., 1998; Wang and Van Cappellen, 1996; Thullner et al., 2009, for applications). However, due to the high degree of coupled processes and depth-varying parameters, the diagenetic equation needs to be solved numerically, thus resulting in a very high computational
140 demand and consequently rendering their application in an Earth system model (ESM) framework with a large number of grid points prohibitive. Additionally, their global applicability is limited by the restricted transferability of model parameters from one site to the global scale (Arndt et al., 2013).

The second group of models solves the diagenetic equation analytically or semi-analytically, thus
145 providing an alternative and computationally more efficient approach. However, finding an analytical solution, especially when complex reaction networks are to be considered, is not straightforward and generally requires the assumption of steady state. The complexity of the reaction network can be reduced by dividing the sediment column into distinct zones and accounting for the most pertinent biogeochemical processes within each zone, thus increasing the likelihood of finding an analytical
150 solution. In general, analytical diagenetic models are less sophisticated and comprehensive than numerical models and are used for the coupling to global ESMs (e.g. HAMOCC and NorESM use the model of Heinze et al. (1999)) or box models (e.g. DCESS, Shaffer et al., 2008 or MBM using MEDUSA, Munhoven, 2007). These analytic or semi-analytical models account for the most important transport processes (i.e advection, bioturbation and molecular diffusion) through basic
155 parametrizations and include fewer biogeochemical reactions which are generally restricted to the upper, bioturbated 10 cm of the sediments. They assume that the sedimentary organic matter pool is composed of just a single compound class which is either degraded with a globally invariant degradation rate constant (Munhoven, 2007) or a fixed rate constant depending on local oxygen concentrations (Shaffer et al., 2008; Palastanga et al., 2011). Pore water tracers explicitly represented in
160 DCESS (Shaffer et al., 2008) and the HAMOCC model of Heinze et al. (1999) and Palastanga et al. (2011) are restricted to DIC, TA, PO₄ and O₂. The MEDUSA model (Munhoven, 2007) considers CO₂, HCO₃⁻, CO₃²⁻ and O₂. Other species produced or consumed during OM degradation are neglected. Thus, with oxygen being the only TEA explicitly modelled the influence of reduced species is only implicitly included in the boundary conditions for O₂. A newer versions of the HAMOCC
165 model, being a notable exception, as Ilyina et al. (2013) include NO₃ and denitrification explicitly. Furthermore, the version of Palastanga et al. (2011) represents an redox-dependent explicit sedimentary phosphorus cycle. Yet, reoxidation of reduced byproducts, so-called secondary redox reactions (e.g. oxidation of NH₄, H₂S or CH₄), or sorption processes are not included in any of the discussed
models.

170 **How are sediments resolved in Earth system models:**

A sediment model suitable for the coupling to an ESM and enabling a wide range of paleo questions to be addressed has to provide a robust quantification of organic (and inorganic) carbon burial fluxes,

DH: I shortened this paragraph a bit. I think, it should be somehow included as it talks specifically about diagenetic models in ESMs!? **AR:** Looks OK to me

as well as benthic uptake/return fluxes of oxygen, growth-limiting nutrients and reduced species. As a consequence, the reaction network must account for the most important primary and secondary
175 redox reactions, equilibrium reactions, mineral precipitation/dissolution and adsorption/desorption, resulting in a complex set of coupled reaction-transport equations. In most current ESMs sediment-water dynamics are either neglected or treated in a very simplistic way (Soetaert et al., 2000; Hülse et al., 2017). Most Earth system Models of Intermediate Complexity (EMICs) and also some of the higher resolution Earth system/climate models represent the sediment-water interface either as
180 a reflective or a conservative/semi-reflective boundary (Hülse et al., 2017). Thus, all particulate material deposited on the seafloor is either instantaneously consumed (reflective boundary), or a fixed fraction is buried in the sediments (conservative/semi-reflective boundary). Both highly simplified approaches furthermore completely neglect the exchange of solute species through the sediment-water interface and, therefore, cannot resolve the complex benthic-pelagic coupling. However, due
185 to their computational efficiency, both representations are often used in global biogeochemical models (e.g. Najjar et al., 2007; Ridgwell et al., 2007; Goosse et al., 2010). The most complex description of diagenetic organic matter degradation in Earth system models is the second group of vertically resolved diagenetic models as discussed above (e.g. Heinze et al., 1999; Munhoven, 2007; Shaffer et al., 2008). These models solve the one-dimensional reaction-transport equation for a number of
190 solid and dissolved species for the upper, bioturbated 10 cm of the sediments. Examples of global ESMs employing a vertically resolved diagenetic model are NorESM (Tjiputra et al., 2013) and HAMOCC (Palastanga et al., 2011; Ilyina et al., 2013), both using a version of Heinze et al. (1999). None of the EMICs reviewed by Hülse et al. (2017) use such a sediment representation. DCESS (Shaffer et al., 2008) and MBM (Munhoven, 2007) are box models employing a vertically resolved
195 diagenetic model. However, in general, oxygen is the only TEA explicitly modelled and secondary redox reactions and reduced species are completely neglected in these approaches. Furthermore, all models represent the bulk OM pool as a single fraction with a fixed degradation rate constant.

Problem with that:

Obviously, such a simplification of the OM pool can neither account for the observed vast structural
200 complexity in natural organic matter and its resulting different degradation rates nor for the rapid decrease in OM degradability in the uppermost centimetres of the sediments (Arndt et al., 2013). It has been suggested that at least a 3G approach is necessary to accurately represent organic matter dynamics in this part of the sediments where most OM is degraded (e.g. Soetaert et al., 1996). Even more restrictive is the use of O₂ as the only TEA and the complete absence of reduced substances
205 and related secondary redox reactions. For the majority of the modern sediments (i.e. in the deep-ocean) O₂ is the primary electron acceptor and Archer et al. (2002) found that aerobic degradation accounts for 66% of total organic matter respiration. However, more recent model and data studies have reported that sulfate reduction is the dominant degradation pathway on a global average (with contributions of 55-76% Canfield et al., 2005; Jørgensen and Kasten, 2006; Thullner et al., 2009).

210 Oxygen becomes progressively less important as TEA with decreasing seafloor depth and sulfate reduction has been shown to account for 83% of OM degradation in coastal sediments (Krumins et al., 2013). In these environments most O₂ is used to reoxidise reduced substances produced during anaerobic degradation (Canfield et al., 2005; Thullner et al., 2009). Thus, the in situ production of e.g. NO₃ and SO₄ through oxidation of NH₄ and H₂S forms an important sink for O₂ which is
215 entirely neglected in current sediment representations in global models. In addition, due to the lack of an appropriate sedimentary P cycle (with the exception of the HAMOCC version of, Palastanga et al., 2011), no current global ESM is able to model the redox dependent P release from marine sediments and its implications for primary productivity, global biogeochemical cycles and climate.

Solution presented here:

220 Analytical approaches with distinct biogeochemical zones were implemented and used in the seventies and eighties to describe observed pore water profiles (e.g. Vanderborght and Billen, 1975; Vanderborght et al., 1977; Billen, 1982; Goloway and Bender, 1982) and later for inclusion into 1-D ecosystem models (e.g. Ruardij and Van Raaphorst, 1995) and global Earth system models (Tromp et al., 1995). However, alongside the oxic zone these models only describe one anoxic zone explicitly, either a denitrification (Vanderborght and Billen, 1975; Vanderborght et al., 1977; Billen, 1982; Goloway and Bender, 1982; Ruardij and Van Raaphorst, 1995) or a sulfate reduction zone (Tromp et al., 1995). Furthermore, the approaches of Vanderborght and Billen (1975), Goloway and Bender (1982) and Tromp et al. (1995) do not explicitly model the produced reduced species (i.e. NH₄ and H₂S, respectively). In addition, the model of Tromp et al. (1995) ignores reoxidation of
225 H₂S produced during sulfate reduction. In order to provide a more realistic description of organic matter degradation and nutrient cycles in marine sediments we have developed the OrganicMatter ENabled SEDiment model (OMEN-SED), a new, one-dimensional, numerically efficient reactive transport model. OMEN-SED is the first analytical model to explicitly describe OM cycling as well as associated dynamics of the most important TEAs (i.e. O₂, NO₃, SO₄), related reduced substances
230 (NH₄, H₂S), the full suite of secondary-redox reactions, macronutrients (PO₄) and associated pore water quantities (ALK, DIC). To represent a redox-dependent sedimentary P cycle we consider the formation and burial of Fe-bound P and authigenic Ca-P minerals. Thus, while OMEN-SED captures most of the features of a complex, numerical diagenetic model, its computational efficiency allows the coupling to global Earth system models and therefore the investigation of coupled global
235 biogeochemical dynamics over different timescales. Here, the model is presented as a 2G-approach, however, a third, non-degradable OM fraction can easily be added and OMEN-SED can be further extended to a Multi-G approach.

The first part of the paper provides a detailed description of OMEN-SED (Section 2). This includes descriptions of the general model approach (Section 2.1), of the conservation equations for
240 all explicitly represented biogeochemical tracers (Section 2.2), as well as a summary of global relationships used to constrain reaction and transport parameters in OMEN-SED (Section 2.4). In

addition, a generic algorithm is described which is used to match internal boundary conditions and to determine the integration constants for the analytical solutions (Section 2.3). In order to validate the stand-alone version of OMEN-SED, the second part of the paper performs an extensive sensitivity analysis for the most important model parameters and resulting sediment-water interface fluxes are compared with a global database (Section 3.1). In addition, results of the stand-alone model are compared with observed pore water profiles from different ocean depths (Section 3.2) and OMEN-SED simulations of TEA-fluxes along a typical ocean transect are compared with observations and results from a complete, numerical diagenetic model (Section 3.3). Thereafter, OMEN-SED is coupled to the carbon-centric version of the “GENIE” Earth system model (cGENIE, Ridgwell et al., 2007, Section 4.1). Sensitivity studies are carried out using this coupled model and modelled organic matter concentrations in the surface sediments are compared to a global database (Seiter et al., 2004, Section 4.2). We finally discuss potential applicabilities of OMEN-SED and critically analyse model limitations (Section 5).

260 2 Model Description

OMEN-SED is implemented as a FORTRAN version that can be easily coupled to a pelagic model via the coupling routine **OMEN_SED_main**. In addition, OMEN-SED exists as a stand-alone version implemented in MATLAB and the entire model can be executed on a standard personal computer in less than 0.1 seconds. The source code of both, the Fortran and the MATLAB stand-alone 265 version, as well as instructions for executing OMEN-SED and for plotting model results are available as a supplement to this paper. The following section provides a detailed description of OMEN-SED and the fundamental equations underlying the model are highlighted. Tables 1 and 15 summarise the biogeochemical reaction network and Tables 9 and 10 provide a glossary of model parameters along with their respective units.

SA: maybe include
some examples and test
figures in supplement?

270 2.1 General Model Approach

In OMEN-SED, the calculation of benthic uptake, recycling and burial fluxes is generally based on the vertically resolved conservation equation for solid and dissolved species in porous media (e.g. Berner, 1980; Boudreau, 1997):

$$\frac{\partial \xi C_i}{\partial t} = -\frac{\partial}{\partial z} \left(-\xi D_i \frac{\partial C_i}{\partial z} + \xi w C_i \right) + \xi \sum_j R_i^j \quad (1)$$

275 where C_i is the concentration of biogeochemical species i , ξ equals the porosity ϕ for solute species and $(1 - \phi)$ for solid species. The term z is the sediment depth, t denotes the time, D_i is the apparent diffusion coefficient of species i ($D_i = D_{i,0} + D_{\text{bio}} = D_{\text{mol},i} \cdot f_{ir} + D_{\text{bio}}$ for dissolved species and $D_i = D_{\text{bio}}$ for solid species), w is the burial rate and $\sum_j R_i^j$ represents the sum of all biogeochemical rates j affecting species i .

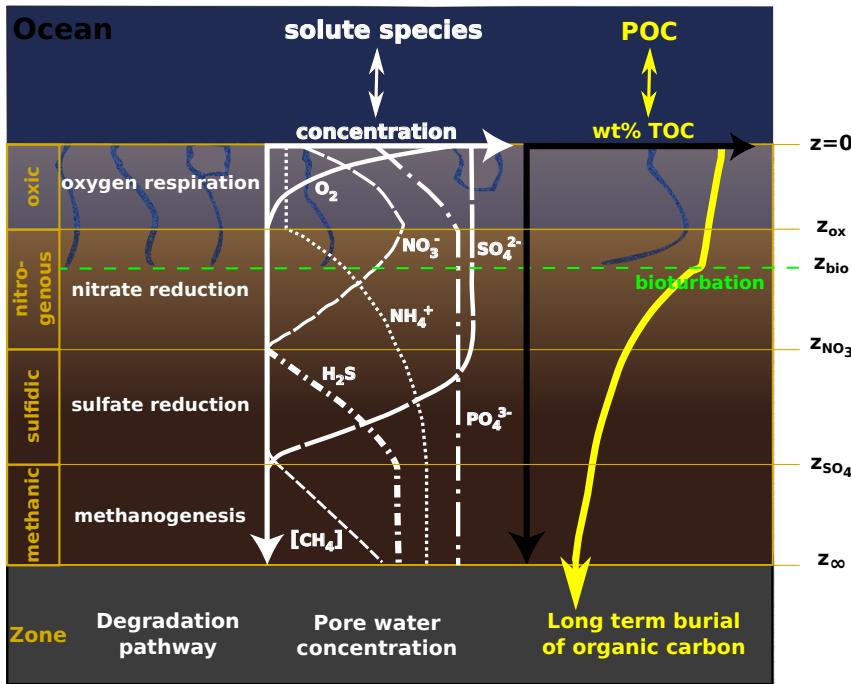


Figure 1. Schematic of the different modelled species and zones in OMEN-SED. Here showing the case $z_{\text{ox}} < z_{\text{bio}} < z_{\text{NO}_3} < z_{\text{SO}_4}$.

280 OMEN-SED accounts for both the advective, as well as the diffusive transport of dissolved and solid species. Solid and dissolved species are buried in the sediment according to a constant burial rate w , thus neglecting the effect of sediment compaction (i.e. $\frac{\partial \phi}{\partial z} = 0$) due to mathematical constraints. The molecular diffusion of dissolved species is described by Fick's law applying a species-specific apparent diffusion coefficient, $D_{\text{mol},i}$. In addition, the activity of infaunal organisms in the
285 bioturbated zone is simulated using a diffusive term (e.g. Boudreau, 1986), with a constant bioturbation coefficient D_{bio} in the bioturbated zone, while D_{bio} is set to zero below the maximum bioturbation depth, z_{bio} . The pumping activity by burrow-dwelling animals and the resulting ventilation of tubes, the so-called bioirrigation, is encapsulated in a factor f_{ir} that enhances the molecular diffusion coefficient (hence, $D_{i,0} = D_{\text{mol},i} \cdot f_{\text{ir}}$, Soetaert et al., 1996). The reaction network of OMEN-SED
290 accounts for the most important primary and secondary redox reactions, equilibrium reactions, mineral dissolution and precipitation, as well as adsorption and desorption processes associated with OM dynamics that affect the dissolved and solid species explicitly resolved in the model. Tables 1 and 15 provide a summary of the reactions and biogeochemical tracers considered in OMEN-SED together with their respective reaction stoichiometries.

295 All parameters in Eq. (1) may vary with depth and many reaction rate expressions depend on the concentration of other species. Expressing Eq. (1) for a set of chemical species thus results in

Table 1. Reactions and biogeochemical tracers implemented in the reaction network of OMEN-SED. The primary and secondary redox reactions are listed in the sequence they occur with increasing sediment depth.

Description	
Primary redox reactions	Degradation of organic matter via aerobic degradation, denitrification, sulfate reduction, methanogenesis (implicit)
Secondary redox reactions	Oxidation of ammonium and sulfide by oxygen, anaerobic oxidation of methane by sulfate
Adsorption/Desorption	Ad-/Desorption of P on/from Fe(OH) ₃ , NH ₄ adsorption, PO ₄ adsorption
Mineral precipitation	Formation of authigenic P
Biogeochemical tracers	Organic matter (2-G), oxygen, nitrate, ammonium, sulfate, sulfide (hydrogen sulfide), phosphate, Fe-bound P, DIC, ALK

a non-linear, coupled set of equations that can only be solved numerically. However, OMEN-SED is designed for the coupling to Earth system models and, therefore, cannot afford a computationally expensive numerical solution. Instead, a computationally efficient analytical solution of Eq. (1) 300 can be derived by 1) assuming steady state conditions (i.e. $\frac{\partial C_i}{\partial t} = 0$) and 2) reducing the vertical variability in parameters and reaction rate expressions by dividing the sediment column into a number of functional biogeochemical zones (Fig. 1, compare e.g. Billen, 1982; Goloway and Bender, 1982; Tromp et al., 1995; Gypens et al., 2008, for similar solutions). More specifically, OMEN-SED divides the sediment column into: I) a bioturbated and II) a non-bioturbated zone defined by 305 an imposed, constant bioturbation depth z_{bio} (Fig. 1). Furthermore, it resolves the dynamic redox stratification of marine sediments by dividing the sediment into 1) an oxic zone delineated by the oxygen penetration depth z_{ox} ; 2) a denitrification (or nitrogenous) zone situated between z_{ox} and the nitrate penetration depth z_{NO_3} ; 3) a sulfate reduction zone situated between z_{NO_3} and the sulfate penetration depth z_{SO_4} ; and 4) a methanogenic zone situated below z_{SO_4} (Fig. 1). Although in each 310 of these zones Eq. (1) is applied with depth invariant parameters, parameter values may differ across zones. The biogeochemical zones are linked by stating continuity in both concentrations and fluxes at the dynamic, internal boundaries ($z_b \in \{z_{\text{bio}}, z_{\text{ox}}, z_{\text{NO}_3}, z_{\text{SO}_4}\}$, compare e.g. Billen, 1982; Ruardij and Van Raaphorst, 1995). These boundaries are dynamic because their depth varies in response to changing ocean boundary conditions and forcings (see Section 2.3.1 for details). Furthermore, the 315 maximum bioturbation depth is not restricted to a specific biogeochemical zone, hence OMEN-SED allows bioturbation to occur in the anoxic zones of the sediment (here all zones $z > z_{\text{ox}}$ combined). In addition, the formulation of the reaction term in Eq. (1) varies between zones and encapsulates the most pertinent reaction processes within the respective zone (see Section 2.2), thus simplifying the mathematical description of the reaction network while retaining most of its biogeochemical 320 complexity.

All consumption or production processes of dissolved species related to the degradation of organic matter are a function of the organic matter concentration and, because first-order decay is assumed in the kinetic expression, can be expressed as a series of exponential terms ($\sum_j \alpha_j \exp(-\beta_j z)$, see Eq. (2)). In addition, slow adsorption/desorption and mineral precipitation processes can be expressed
325 as zero or first order (reversible) reaction (Q or $k \cdot (C_i - \tilde{C})$, in Eq. (2)). Fast adsorption is described as an instantaneous equilibrium reaction using a constant adsorption coefficient K_i . The reoxidation of reduced substances is accounted for implicitly by adding a (consumption/production) flux to the internal boundary conditions (see Sections 2.2.2, 2.2.3 and 2.2.4 **SA: make reference to the section where this is explained in detail**). This simplification has been used previously by Gypens et al.
330 (2008) for nitrate and ammonium and can be justified as it has been shown that the reoxidation mainly occurs within a thin layer at the oxic/anoxic interface (Soetaert et al., 1996). The general reaction-transport equation of OMEN-SED is thus given by:

$$\frac{\partial C_i}{\partial t} = 0 = \frac{D_i}{1 + K_i} \frac{\partial^2 C_i}{\partial z^2} - w \frac{\partial C_i}{\partial z} - \frac{1}{1 + K_i} \left(\sum_j \alpha_j \exp(-\beta_j z) + \sum_l k_l \cdot C_i - \sum_m Q_m \right)$$

(2)

where $1/\beta_j$ can be interpreted as the length scale and α_j as the relative importance (or the magnitude
335 at $z = 0$) of reaction j (Boudreau, 1997), k_l are generic first order reaction rate constants and Q_m are zeroth-order (or constant) reaction rates.

The analytical solution of Eq. (2) is of the general form:

$$C_i(z) = A \exp(az) + B \exp(bz) + \sum_j \frac{\alpha_j}{D\beta_j^2 - w\beta_j - \sum_l k_l} \cdot \exp(-\beta_j z) + \frac{\sum_m Q_m}{\sum_l k_l} \quad (3)$$

340 with

$$a = \frac{w - \sqrt{w^2 + 4 \cdot D \cdot \sum_l k_l}}{2 \cdot D}, \quad b = \frac{w + \sqrt{w^2 + 4 \cdot D \cdot \sum_l k_l}}{2 \cdot D} \quad (4)$$

where A and B are integration constants that can be determined by applying a set of internal boundary conditions (see Section 2.3).

Based on Eq. (2) and its analytical solution Eq. (3), OMEN-SED returns the fraction of particulate
345 organic carbon (POC) buried in the sediment, f_{POC} , as well as the benthic uptake/return fluxes F_{C_i} of dissolved species C_i (in $\text{mol cm}^{-2} \text{ year}^{-1}$) in response to changing boundary conditions and forcings:

$$f_{\text{POC}} = \frac{\text{POC}(z_\infty)}{\text{POC}(0)} \quad (5)$$

$$F_{C_i} = \phi(0) \left(D_i \frac{\partial C_i(z)}{\partial z} \Big|_{z=0} - w \cdot C_i(0) \right) \quad (6)$$

DH: Do you mean these sections? Not discussed in further details anywhere else

where w is the deposition rate, D_i is the diffusion coefficient and $\text{POC}(0)$, $\text{POC}(z_\infty)$, $C_i(0)$ denotes the concentration of POC and dissolved species i at the SWI and at the lower sediment boundary, respectively.

2.2 Conservation Equations and Analytical Solution

355 2.2.1 Organic matter or Particulate Organic Carbon (POC)

In marine sediments, particulate organic carbon (POC) is degraded by heterotrophic activity coupled to the sequential utilisation of terminal electron acceptors according to the free energy gain of the half-reaction ($\text{O}_2 > \text{NO}_3^- > \text{MnO}_2 > \text{Fe(OH)}_3 > \text{SO}_4^{2-}$, e.g. Stumm and Morgan, 2012). Here, organic matter degradation is described via a multi G-model approach (Arndt et al., 2013, 360 and references therein), dividing the bulk OM into a number i of discrete compound classes POC_i characterised by class-specific first-order degradation rate constants k_i . The conservation equation for organic matter dynamics is thus given by:

$$\frac{\partial \text{POC}_i}{\partial t} = 0 = D_{\text{POC}_i} \frac{\partial^2 \text{POC}_i}{\partial z^2} - w \frac{\partial \text{POC}_i}{\partial z} - k_i \cdot \text{POC}_i \quad (7)$$

with $D_{\text{POC}_i} = D_{\text{bio}}$ for $z \leq z_{\text{bio}}$ and $D_{\text{POC}_i} = 0$ for $z > z_{\text{bio}}$. Integration of equations (7) yields the 365 following general solutions:

I. Bioturbated zone ($z \leq z_{\text{bio}}$)

$$\text{POC}_i^I(z) = A_{1i} \cdot \exp(a_{1i}z) + B_{1i} \cdot \exp(b_{1i}z)$$

370 which using the boundary condition at $z = 0$ can be rewritten as:

$$\text{POC}_i^I(z) \stackrel{\text{BC1}}{=} A_{1i} \cdot [\exp(a_{1i}z) - \exp(b_{1i}z)] + \text{POC}_{0i} \cdot \exp(b_{1i}z) \quad (8)$$

II. Non-bioturbated zone ($z_{\text{bio}} < z$)

$$375 \text{POC}_i^{II}(z) = A_{2i} \cdot \exp(a_{2i}z) \quad (9)$$

where

$$a_{1i} = \frac{w - \sqrt{w^2 + 4 \cdot D_{\text{POC}_i} \cdot k_i}}{2 \cdot D_{\text{POC}_i}}, \quad b_{1i} = \frac{w + \sqrt{w^2 + 4 \cdot D_{\text{POC}_i} \cdot k_i}}{2 \cdot D_{\text{POC}_i}}, \quad a_{2i} = -\frac{k_i}{w} \quad (10)$$

Determining the integration constants ($A_{1,i}$, $B_{1,i}$, $A_{2,i}$) requires the definition of a set of boundary conditions (Table 2). For organic matter, OMEN-SED applies a known concentration/flux at the 380 sediment-water interface and assumes continuity across the bottom of the bioturbated zone, z_{bio} . See Section 2.3.1 for further details on how to find the analytical solution.

Table 2. OM Boundary conditions applied in OMEN-SED. For the boundaries we define: $z_{\text{bio}}^- := \lim_{h \rightarrow 0} (z_{\text{bio}} - h)$ and $z_{\text{bio}}^+ := \lim_{h \rightarrow 0} (z_{\text{bio}} + h)$.

Boundary	Condition	
$z = 0$	known concentration	1) $\text{POC}_i(0) = \text{POC}_{0i}$
$z = z_{\text{bio}}$	continuity	2) $\text{POC}_i(z_{\text{bio}}^-) = \text{POC}_i(z_{\text{bio}}^+)$ 3) $-D_{\text{bio}} \cdot \frac{\partial \text{POC}_i}{\partial z} \Big _{z_{\text{bio}}^-} = 0$

2.2.2 Oxygen

OMEN-SED explicitly accounts for oxygen consumption by the aerobic degradation of organic matter within the oxic zone, as well as the oxidation of reduced species (i.e. NH_4 , H_2S) produced in the
 385 anoxic zones of the sediment. In the oxic zone ($z < z_{\text{ox}}$), the aerobic degradation consumes oxygen with a fixed $\text{O}_2 : \text{C}$ ratio (O_2C , Tab. 10). A predefined fraction, γ_{NH_4} , of the ammonium produced during the aerobic degradation of OM is nitrified to nitrate, consuming two moles of oxygen per mole of ammonium produced. In addition, OMEN-SED implicitly accounts for the oxygen consumption due to oxidation of reduced species (NH_4 , H_2S) produced below the oxic zone through the flux
 390 boundary condition at the dynamically calculated (see section 2.4.2 for details) oxygen penetration depth z_{ox} . All oxygen consumption processes can thus be formulated as a function of organic matter degradation. The conservation equation for oxygen is given by: **SA: I'd show the POC substitution in the equations below:**

$$395 \quad \frac{\partial \text{O}_2}{\partial t} = 0 = D_{\text{O}_2} \frac{\partial^2 \text{O}_2}{\partial z^2} - w \frac{\partial \text{O}_2}{\partial z} - \frac{1-\phi}{\phi} \sum_i k_i \cdot [\text{O}_2\text{C} + 2\gamma_{\text{NH}_4} \text{NC}_i] \cdot \text{POC}_i(z) \quad (11)$$

which, using Eq. (8) and (9) for the depth-distribution of $\text{POC}_i(z)$, can be written as:

I Bioturbated zone ($z \leq z_{\text{bio}}$)

$$400 \quad \frac{\partial \text{O}_2^I}{\partial t} = 0 \stackrel{8}{=} D_{\text{O}_2}^I \frac{\partial^2 \text{O}_2}{\partial z^2} - w \frac{\partial \text{O}_2}{\partial z} - \frac{1-\phi}{\phi} \sum_i k_i \cdot [\text{O}_2\text{C} + 2\gamma_{\text{NH}_4} \text{NC}_i] \cdot \left(A_{1i} \cdot [\exp(a_{1i}z) - \exp(b_{1i}z)] + \text{POC}_{0i} \cdot \exp(b_{1i}z) \right)$$

II Non-bioturbated zone ($z_{\text{bio}} < z < z_{\text{ox}}$)

$$405 \quad \frac{\partial \text{O}_2^{II}}{\partial t} = 0 \stackrel{9}{=} D_{\text{O}_2}^{II} \frac{\partial^2 \text{O}_2}{\partial z^2} - w \frac{\partial \text{O}_2}{\partial z} - \frac{1-\phi}{\phi} \sum_i k_i \cdot [\text{O}_2\text{C} + 2\gamma_{\text{NH}_4} \text{NC}_i] \cdot \left(A_{2i} \cdot \exp(a_{2i}z) \right)$$

where $D_{\text{O}_2}^I$ and $D_{\text{O}_2}^{II}$ denote the O_2 diffusion coefficient for the bioturbated and non-bioturbated zone, respectively. The term $\frac{1-\phi}{\phi}$ accounts for the volume conversion from solid to dissolved phase and NC_i is the nitrogen to carbon ratio in OM. Integration yields the following analytical solution for each zone:

DH: you mean the way I did it or in the solution?
 I think it's easier to understand (also how to get the solution) if it's done in the ODE

I Bioturbated zone ($z \leq z_{\text{bio}}$):

$$O_2^I(z) = A_{O_2}^1 + B_{O_2}^1 \cdot \exp(b_{O_2}^1 z) + \sum_i \Phi_{1,i}^I \cdot \exp(a_{1i} z) + \sum_i \Phi_{1,i}^{II} \cdot \exp(b_{1i} z) + \sum_i \Phi_{1,i}^{III} \cdot \exp(b_{1i} z) \quad (12)$$

II Non-bioturbated zone ($z_{\text{bio}} < z < z_{\text{ox}}$)

$$O_2^{II}(z) = A_{O_2}^2 + B_{O_2}^2 \cdot \exp(b_{O_2}^2 z) + \sum_i \Phi_{i,2}^I \cdot \exp(a_{2i} z) \quad (13)$$

with

$$\begin{aligned} b_{O_2}^1 &= \frac{w}{D_{O_2}^I}, \quad b_{O_2}^2 = \frac{w}{D_{O_2}^{II}} \\ \Phi_{1,i}^I &= \frac{1-\phi}{\phi} \cdot \frac{k_i \cdot (O_2 C + 2\gamma_{NH_4} NC_i) \cdot A_{1i}}{D_{O_2}^I (-a_{1i})^2 - w \cdot (-a_{1i})}, \quad \Phi_{1,i}^{II} = -\frac{1-\phi}{\phi} \cdot \frac{k_i \cdot (O_2 C + 2\gamma_{NH_4} NC_i) \cdot A_{1i}}{D_{O_2}^I (-b_{1i})^2 - w \cdot (-b_{1i})} \\ \Phi_{1,i}^{III} &= \frac{1-\phi}{\phi} \cdot \frac{k_i \cdot (O_2 C + 2\gamma_{NH_4} NC_i) \cdot POC_{0i}}{D_{O_2}^I (-b_{1i})^2 - w \cdot (-b_{1i})} \\ \Phi_{i,2}^I &:= \frac{1-\phi}{\phi} \cdot \frac{k_i \cdot (O_2 C + 2\gamma_{NH_4} NC_i) \cdot A_{2i}}{D_{O_2}^{II} (-a_{2i})^2 - w \cdot (-a_{2i})} \end{aligned}$$

Determining the four integration constants ($A_{O_2}^1, B_{O_2}^1, A_{O_2}^2, B_{O_2}^2$, see Section 2.3 for details), as well as the *a priori* unknown oxygen penetration depth requires the definition of five boundary conditions (see Table 3). At the sediment-water interface, OMEN-SED applies a Dirichlet condition (i.e. known concentration) and assumes concentration and flux continuity across the bottom of the bioturbated zone, z_{bio} . The oxygen penetration depth z_{ox} marks the lower boundary and is dynamically calculated as the depth at which $O_2(z) = 0$. Therefore, OMEN-SED applies a Dirichlet boundary condition $O_2(z_{\text{ox}}) = 0$. In addition, a flux boundary is applied that implicitly accounts for the oxygen consumption by the partial oxidation of NH_4 and H_2S diffusing into the oxic zone from below (BC 4.2, Table 3). It is assumed that respective fractions (γ_{NH_4} and γ_{H_2S}) are directly reoxidised at the oxic/anoxic interface and the remaining fraction escapes reoxidation. OMEN-SED iteratively solves for z_{ox} by first testing if there is oxygen left at z_∞ (i.e. $O_2(z_\infty) > 0$) and, otherwise, by finding the root for the flux boundary condition 4.2 (Table 3). If $z_{\text{ox}} = z_\infty$, a zero diffusive flux boundary condition is applied as lower boundary condition.

DH: Edited explanation of finding z_{ox}

2.2.3 Nitrate and Ammonium

Nitrogen dynamics in OMEN-SED are controlled by the metabolic production of ammonium, nitrification, denitrification as well as ammonium adsorption. Ammonium is produced by organic matter degradation in both the oxic and anoxic zones, while denitrification consumes nitrate in the denitrification zone with a fixed $NO_3^- : C$ ratio (NO_3^-C , Tab. 10) **SA: need explanation**. The adsorption of ammonium to sediment particles is formulated as an equilibrium process with constant equilibrium adsorption coefficient K_{NH_4} , thus assuming that the adsorption is fast compared with the characteristic time scales of transport processes (Wang and Van Cappellen, 1996). In addition, a defined

DH: explanation sufficient?

Table 3. Boundary conditions for oxygen. For the boundaries we define: $z_{\text{bio}}^- := \lim_{h \rightarrow 0} (z_{\text{bio}} - h)$ and $z_{\text{bio}}^+ := \lim_{h \rightarrow 0} (z_{\text{bio}} + h)$.

Boundary	Condition	
$z = 0$	known concentration	1) $O_2(0) = O_{20}$
$z = z_{\text{bio}}$	continuity	2) $O_2(z_{\text{bio}}^-) = O_2(z_{\text{bio}}^+)$
$z = z_{\text{ox}}$	O_2 consumption ($z_{\text{ox}} = z_\infty$)	3) $-(D_{O_2,0} + D_{\text{bio}}) \cdot \frac{\partial O_2}{\partial z} \Big _{z_{\text{bio}}^-} = -D_{O_2,0} \cdot \frac{\partial O_2}{\partial z} \Big _{z_{\text{bio}}^+}$ 4) IF ($O_2(z_\infty) > 0$) 4.1) $\frac{\partial O_2}{\partial z} \Big _{z_{\text{ox}}} = 0$ ELSE
	($z_{\text{ox}} < z_\infty$) with	4.2) $O_2(z_{\text{ox}}) = 0 \quad \text{and} \quad -D_{O_2} \cdot \frac{\partial O_2}{\partial z} \Big _{z_{\text{ox}}} = F_{\text{red}}(z_{\text{ox}})$ $F_{\text{red}}(z_{\text{ox}}) = \frac{1-\phi}{\phi} \cdot \int_{\tilde{z}_{\text{NO}_3}}^{\infty} \sum_i (2\gamma_{\text{NH}_4} \text{NC}_i + 2\gamma_{\text{H}_2\text{S}} \text{SO}_4 \text{C}) k_i \text{POC}_i dz$

Note: $\tilde{z}_{\text{NO}_3} = z_{\text{ox}}$ as upper boundary here, as z_{NO_3} is not known at this point.

fraction, γ_{NH_4} , of metabolically produced ammonium is directly nitrified to nitrate in the oxic zone,
445 while the nitrification of upward diffusing ammonium produced in the sulfidic and methanic zones is implicitly accounted for in the boundary conditions. The conservation equations for ammonium and nitrate are thus given by:

450 1. Oxic zone ($z \leq z_{\text{ox}}$)

$$\frac{\partial \text{NO}_3^I}{\partial t} = 0 = D_{\text{NO}_3} \frac{\partial^2 \text{NO}_3^I}{\partial z^2} - w \frac{\partial \text{NO}_3^I}{\partial z} + \gamma_{\text{NH}_4} \frac{1-\phi}{\phi} \cdot \sum_i \text{NC}_i \cdot k_i \cdot \text{POC}_i(z) \quad (14)$$

$$\frac{\partial \text{NH}_4^I}{\partial t} = 0 = \frac{D_{\text{NH}_4}}{1 + K_{\text{NH}_4}} \frac{\partial^2 \text{NH}_4^I}{\partial z^2} - w \frac{\partial \text{NH}_4^I}{\partial z} + \frac{1 - \gamma_{\text{NH}_4}}{1 + K_{\text{NH}_4}} \cdot \frac{1-\phi}{\phi} \cdot \sum_i \text{NC}_i \cdot k_i \cdot \text{POC}_i(z) \quad (15)$$

2. Denitrification (or nitrogenous) zone ($z_{\text{ox}} < z \leq z_{\text{NO}_3}$)

$$\frac{\partial \text{NO}_3^{II}}{\partial t} = 0 = D_{\text{NO}_3} \frac{\partial^2 \text{NO}_3^{II}}{\partial z^2} - w \frac{\partial \text{NO}_3^{II}}{\partial z} - \frac{1-\phi}{\phi} \text{NO}_3 \text{C} \cdot \sum_i k_i \cdot \text{POC}_i(z) \quad (16)$$

$$\frac{\partial \text{NH}_4^{II}}{\partial t} = 0 = \frac{D_{\text{NH}_4}}{1 + K_{\text{NH}_4}} \frac{\partial^2 \text{NH}_4^{II}}{\partial z^2} - w \frac{\partial \text{NH}_4^{II}}{\partial z} \quad (17)$$

3. Sulfidic and methanic zone ($z_{\text{NO}_3} < z \leq z_\infty$)

$$\frac{\partial \text{NH}_4^{III}}{\partial t} = 0 = \frac{D_{\text{NH}_4}}{1 + K_{\text{NH}_4}} \frac{\partial^2 \text{NH}_4^{III}}{\partial z^2} - w \frac{\partial \text{NH}_4^{III}}{\partial z} + \frac{1}{1 + K_{\text{NH}_4}} \cdot \frac{1-\phi}{\phi} \cdot \sum_i \text{NC}_i \cdot k_i \cdot \text{POC}_i(z) \quad (18)$$

where D_{NO_3} and D_{NH_4} denote the diffusion coefficients for NO_3 and NH_4 which depend on the bioturbation status of the respective geochemical zone (compare Section 2.3.1). Integration of

Table 4. Boundary conditions for nitrate and ammonium. For the boundaries we define: $z_-^- := \lim_{h \rightarrow 0} (z_- - h)$ and $z_-^+ := \lim_{h \rightarrow 0} (z_- + h)$.

Boundary	Condition	
$z = 0$	known concentration	1) $\text{NO}_3(0) = \text{NO}_{30}$
$z = z_{\text{bio}}$	continuity	2) $\text{NO}_3(z_{\text{bio}}^-) = \text{NO}_3(z_{\text{bio}}^+)$
$z = z_{\text{ox}}$	continuity	3) $-(D_{\text{NO}_3,0} + D_{\text{bio}}) \cdot \frac{\partial \text{NO}_3}{\partial z} _{z_{\text{bio}}^-} = -D_{\text{NO}_3,0} \cdot \frac{\partial \text{NO}_3}{\partial z} _{z_{\text{bio}}^+}$
	where:	4) $\text{NO}_3(z_{\text{ox}}^-) = \text{NO}_3(z_{\text{ox}}^+)$
$z = z_{\text{NO}_3}$	NO_3 consumption ($z_{\text{NO}_3} = z_\infty$)	5) $-D_{\text{NO}_3} \cdot \frac{\partial \text{NO}_3}{\partial z} _{z_{\text{ox}}^-} + \gamma_{\text{NH}_4} \cdot F_{\text{NH}_4}(z_{\text{ox}}) = -D_{\text{NO}_3} \cdot \frac{\partial \text{NO}_3}{\partial z} _{z_{\text{ox}}^+}$ $F_{\text{NH}_4}(z_{\text{ox}}) = \frac{1}{1+K_{\text{NH}_4}} \cdot \frac{1-\phi}{\phi} \cdot \int_{z_{\text{NO}_3}}^\infty \sum_i \text{NC}_i \cdot k_i \cdot \text{POC}_i dz$
		6) IF ($\text{NO}_3(z_\infty) > 0$)
		6.1) $\frac{\partial \text{NO}_3}{\partial z} _{z_{\text{NO}_3}} = 0$
		ELSE
		6.2) $\text{NO}_3(z_{\text{NO}_3}) = 0 \quad \text{and} \quad \frac{\partial \text{NO}_3}{\partial z} _{z_{\text{NO}_3}} = 0$
$z = 0$	known concentration	1) $\text{NH}_4(0) = \text{NH}_{40}$
$z = z_{\text{bio}}$	continuity	2) $\text{NH}_4(z_{\text{bio}}^-) = \text{NH}_4(z_{\text{bio}}^+)$
$z = z_{\text{ox}}$	continuity	3) $-\frac{D_{\text{NH}_4,0} + D_{\text{bio}}}{1+K_{\text{NH}_4}} \cdot \frac{\partial \text{NH}_4}{\partial z} _{z_{\text{bio}}^-} = -\frac{D_{\text{NH}_4,0}}{1+K_{\text{NH}_4}} \cdot \frac{\partial \text{NH}_4}{\partial z} _{z_{\text{bio}}^+}$
	where:	4) $\text{NH}_4(z_{\text{ox}}^-) = \text{NH}_4(z_{\text{ox}}^+)$
$z = z_{\text{NO}_3}$	continuity flux	5) $-\frac{D_{\text{NH}_4}}{1+K_{\text{NH}_4}} \cdot \frac{\partial \text{NH}_4}{\partial z} _{z_{\text{ox}}^-} - \gamma_{\text{NH}_4} \cdot F_{\text{NH}_4}(z_{\text{ox}}) = -\frac{D_{\text{NH}_4}}{1+K_{\text{NH}_4}} \cdot \frac{\partial \text{NH}_4}{\partial z} _{z_{\text{ox}}^+}$ $F_{\text{NH}_4}(z_{\text{ox}}) = \frac{1}{1+K_{\text{NH}_4}} \cdot \frac{1-\phi}{\phi} \cdot \int_{z_{\text{NO}_3}}^\infty \sum_i \text{NC}_i \cdot k_i \cdot \text{POC}_i dz$
$z = z_\infty$	zero NH_4 flux	6) $\text{NH}_4(z_{\text{NO}_3}^-) = \text{NH}_4(z_{\text{NO}_3}^+)$
		7) $-\frac{D_{\text{NH}_4}}{1+K_{\text{NH}_4}} \cdot \frac{\partial \text{NH}_4}{\partial z} _{z_{\text{NO}_3}^-} = -\frac{D_{\text{NH}_4}}{1+K_{\text{NH}_4}} \cdot \frac{\partial \text{NH}_4}{\partial z} _{z_{\text{NO}_3}^+}$
		8) $\frac{\partial \text{NH}_4}{\partial z} _{z_\infty} = 0$

Eq. (14) - (18) yields the analytical solutions, which are not further developed here but follow the procedure outlined in Section 2.2.2 for oxygen (also see Section 2.3.1 for more details on how to 465 find the analytical solution). Table 4 summarises the boundary conditions applied in OMEN-SED to solve Eq. (14) - (18) and to find the *a priori* unknown nitrate penetration depth, z_{NO_3} . The model assumes known bottom water concentrations for both NO_3 and NH_4 , the complete consumption of nitrate at the nitrate penetration depth (in case $z_{\text{NO}_3} < z_\infty$) and no change in ammonium flux at z_∞ . In addition, concentration and diffusive flux continuity across z_{bio} and z_{ox} is considered for NO_3 470 and NH_4 . Furthermore, the reoxidation of upward-diffusing reduced ammonium is accounted for in the oxic-anoxic boundary condition for nitrate and ammonium. OMEN-SED iteratively solves for z_{NO_3} by first testing if there is nitrate left at z_∞ (i.e. $\text{NO}_3(z_\infty) > 0$) and, otherwise, by finding the root for the flux boundary condition 6.2 (Table 4).

2.2.4 Sulfate and Sulfide

475 Below the denitrification zone ($z > z_{\text{NO}_3}$), organic matter degradation is coupled to sulfate reduction, consuming sulfate and producing hydrogen sulfide with a fixed $\text{SO}_4 : \text{C}$ ratio (SO_4C , Tab. 10). In addition, the anaerobic oxidation of upward diffusing methane (AOM) produced below the sulfate penetration and the associated consumption of sulfate and production of sulfide; as well as the production of sulfate and consumption of sulfide through sulfide oxidation are implicitly accounted
 480 for through the boundary conditions (Table 5). The conservation equations for sulfate and sulfide are thus given by:

1. Oxic and nitrogenous zone ($z \leq z_{\text{NO}_3}$)

$$485 \quad \frac{\partial \text{SO}_4^I}{\partial t} = 0 = D_{\text{SO}_4} \frac{\partial^2 \text{SO}_4^I}{\partial z^2} - w \frac{\partial \text{SO}_4^I}{\partial z} \quad (19)$$

$$\frac{\partial \text{H}_2\text{S}^I}{\partial t} = 0 = D_{\text{H}_2\text{S}} \frac{\partial^2 \text{H}_2\text{S}^I}{\partial z^2} - w \frac{\partial \text{H}_2\text{S}^I}{\partial z} \quad (20)$$

2. Sulfidic zone ($z_{\text{NO}_3} < z \leq z_{\text{SO}_4}$)

$$\frac{\partial \text{SO}_4^{II}}{\partial t} = 0 = D_{\text{SO}_4} \frac{\partial^2 \text{SO}_4^{II}}{\partial z^2} - w \frac{\partial \text{SO}_4^{II}}{\partial z} - \frac{1-\phi}{\phi} \cdot \sum_i \text{SO}_4\text{C} \cdot k_i \cdot \text{POC}_i(z) \quad (21)$$

$$490 \quad \frac{\partial \text{H}_2\text{S}^{II}}{\partial t} = 0 = D_{\text{H}_2\text{S}} \frac{\partial^2 \text{H}_2\text{S}^{II}}{\partial z^2} - w \frac{\partial \text{H}_2\text{S}^{II}}{\partial z} + \frac{1-\phi}{\phi} \cdot \sum_i \text{SO}_4\text{C} \cdot k_i \cdot \text{POC}_i(z) \quad (22)$$

3. Methanic zone ($z_{\text{SO}_4} < z \leq z_\infty$)

$$\frac{\partial \text{H}_2\text{S}^{III}}{\partial t} = 0 = D_{\text{H}_2\text{S}} \frac{\partial^2 \text{H}_2\text{S}^{III}}{\partial z^2} - w \frac{\partial \text{H}_2\text{S}^{III}}{\partial z} \quad (23)$$

495 where D_{SO_4} and $D_{\text{H}_2\text{S}}$ denote the diffusion coefficients for SO_4 and H_2S which depend on the bioturbation status of the respective geochemical zone (compare Section 2.3.1). Integration of Eq.
 (19) - (23) yields the analytical solution and Table 5 summarises the boundary conditions applied.
 OMEN-SED assumes known concentrations at the sediment-water interface and continuity across
 the bioturbation depth and the nitrate penetration depth. The reoxidation of reduced H_2S to SO_4 is
 500 accounted for implicitly via the oxic-anoxic boundary condition for both species, while reduction
 of SO_4 and the associated production of H_2S via AOM is accounted for through the respective
 boundary conditions at z_{SO_4} . In case $z_{\text{SO}_4} < z_\infty$, OMEN-SED assumes zero sulfate concentration
 at z_{SO_4} and its diffusive flux must equal the amount of methane produced below (with a methane to
 carbon ratio of MC); or, in case $z_{\text{SO}_4} = z_\infty$, a zero diffusive flux condition for sulfate is considered.
 505 OMEN-SED iteratively solves for z_{SO_4} by first testing if there is sulfate left at z_∞ (i.e. $\text{SO}_4(z_\infty) >$
 0) and, otherwise, by finding the root for the flux boundary condition 8.2 (Table 5). At the lower
 boundary z_∞ zero diffusive flux of H_2S is considered.

Table 5. Boundary conditions for sulfate and sulfide. For the boundaries we define: $z_-^- := \lim_{h \rightarrow 0} (z_- - h)$ and $z_-^+ := \lim_{h \rightarrow 0} (z_- + h)$.

Boundary	Condition	
$z = 0$	known concentration	1) $\text{SO}_4(0) = \text{SO}_4$
$z = z_{\text{bio}}$	continuity	2) $\text{SO}_4(z_{\text{bio}}^-) = \text{SO}_4(z_{\text{bio}}^+)$
	flux	3) $-(D_{\text{SO}_4,0} + D_{\text{bio}}) \cdot \frac{\partial \text{SO}_4}{\partial z} _{z_{\text{bio}}^-} = -D_{\text{SO}_4,0} \cdot \frac{\partial \text{SO}_4}{\partial z} _{z_{\text{bio}}^+}$
$z = z_{\text{ox}}$	continuity	4) $\text{SO}_4(z_{\text{ox}}^-) = \text{SO}_4(z_{\text{ox}}^+)$
	flux	5) $-D_{\text{SO}_4} \cdot \frac{\partial \text{SO}_4}{\partial z} _{z_{\text{ox}}^-} + \gamma_{\text{H}_2\text{S}} \cdot F_{\text{H}_2\text{S}}(z_{\text{ox}}) = -D_{\text{SO}_4} \cdot \frac{\partial \text{SO}_4}{\partial z} _{z_{\text{ox}}^+}$
	where:	$F_{\text{H}_2\text{S}}(z_{\text{ox}}) = \frac{1-\phi}{\phi} \cdot \left(\int_{z_{\text{NO}_3}}^{\text{SO}_4} \sum_i \text{SO}_4 \cdot k_i \cdot \text{POC}_i dz + \gamma_{\text{CH}_4} \cdot \int_{z_{\text{SO}_4}}^{\infty} \sum_i \text{MC} \cdot k_i \cdot \text{POC}_i dz \right)$
$z = z_{\text{NO}_3}$	continuity	6) $\text{SO}_4(z_{\text{NO}_3}^-) = \text{SO}_4(z_{\text{NO}_3}^+)$
	flux	7) $-D_{\text{SO}_4} \cdot \frac{\partial \text{SO}_4}{\partial z} _{z_{\text{NO}_3}^-} = -D_{\text{SO}_4} \cdot \frac{\partial \text{SO}_4}{\partial z} _{z_{\text{NO}_3}^+}$
$z = z_{\text{SO}_4}$	SO ₄ consumption ($z_{\text{SO}_4} = z_\infty$)	8) IF ($\text{SO}_4(z_\infty) > 0$) 8.1) $\frac{\partial \text{SO}_4}{\partial z} _{z_{\text{SO}_4}} = 0$ ELSE
	($z_{\text{SO}_4} < z_\infty$) with	8.2) $\text{SO}_4(z_{\text{SO}_4}) = 0 \quad \text{and} \quad -D_{\text{SO}_4} \cdot \frac{\partial \text{SO}_4}{\partial z} _{z_{\text{SO}_4}} = \gamma_{\text{CH}_4} \cdot F_{\text{CH}_4}(z_{\text{SO}_4})$ $F_{\text{CH}_4}(z_{\text{SO}_4}) = \frac{1-\phi}{\phi} \cdot \int_{z_{\text{SO}_4}}^{\infty} \sum_i \text{MC} \cdot k_i \cdot \text{POC}_i dz$
$z = 0$	known concentration	1) $\text{H}_2\text{S}(0) = \text{H}_2\text{S}_0$
$z = z_{\text{bio}}$	continuity	2) $\text{H}_2\text{S}(z_{\text{bio}}^-) = \text{H}_2\text{S}(z_{\text{bio}}^+)$
	flux	3) $-(D_{\text{H}_2\text{S},0} + D_{\text{bio}}) \cdot \frac{\partial \text{H}_2\text{S}}{\partial z} _{z_{\text{bio}}^-} = -D_{\text{H}_2\text{S},0} \cdot \frac{\partial \text{H}_2\text{S}}{\partial z} _{z_{\text{bio}}^+}$
$z = z_{\text{ox}}$	continuity	4) $\text{H}_2\text{S}(z_{\text{ox}}^-) = \text{H}_2\text{S}(z_{\text{ox}}^+)$
	flux	5) $-D_{\text{H}_2\text{S}} \cdot \frac{\partial \text{H}_2\text{S}}{\partial z} _{z_{\text{ox}}^-} - \gamma_{\text{H}_2\text{S}} F_{\text{H}_2\text{S}}(z_{\text{ox}}) = -D_{\text{H}_2\text{S}} \cdot \frac{\partial \text{H}_2\text{S}}{\partial z} _{z_{\text{ox}}^+}$
	where:	$F_{\text{H}_2\text{S}}(z_{\text{ox}}) = \frac{1-\phi}{\phi} \cdot \left(\int_{z_{\text{NO}_3}}^{\text{SO}_4} \sum_i \text{SO}_4 \cdot k_i \cdot \text{POC}_i dz + \gamma_{\text{CH}_4} \cdot \int_{z_{\text{SO}_4}}^{\infty} \sum_i \text{MC} \cdot k_i \cdot \text{POC}_i dz \right)$
$z = z_{\text{NO}_3}$	continuity	6) $\text{H}_2\text{S}(z_{\text{NO}_3}^-) = \text{H}_2\text{S}(z_{\text{NO}_3}^+)$
	flux	7) $-D_{\text{H}_2\text{S}} \cdot \frac{\partial \text{H}_2\text{S}}{\partial z} _{z_{\text{NO}_3}^-} = -D_{\text{H}_2\text{S}} \cdot \frac{\partial \text{H}_2\text{S}}{\partial z} _{z_{\text{NO}_3}^+}$
$z = z_{\text{SO}_4}$	continuity flux (with AOM) where:	8) $\text{H}_2\text{S}(z_{\text{SO}_4}^-) = \text{H}_2\text{S}(z_{\text{SO}_4}^+)$ 9) $-D_{\text{H}_2\text{S}} \cdot \frac{\partial \text{H}_2\text{S}}{\partial z} _{z_{\text{SO}_4}^-} + \gamma_{\text{CH}_4} \cdot F_{\text{CH}_4}(z_{\text{SO}_4}) = -D_{\text{H}_2\text{S}} \cdot \frac{\partial \text{H}_2\text{S}}{\partial z} _{z_{\text{SO}_4}^+}$ $F_{\text{CH}_4}(z_{\text{SO}_4}) = \frac{1-\phi}{\phi} \cdot \int_{z_{\text{SO}_4}}^{\infty} \sum_i \text{MC} \cdot k_i \cdot \text{POC}_i dz$
$z = z_\infty$	zero H ₂ S flux	10) $\frac{\partial \text{H}_2\text{S}}{\partial z} _{z_\infty} = 0$

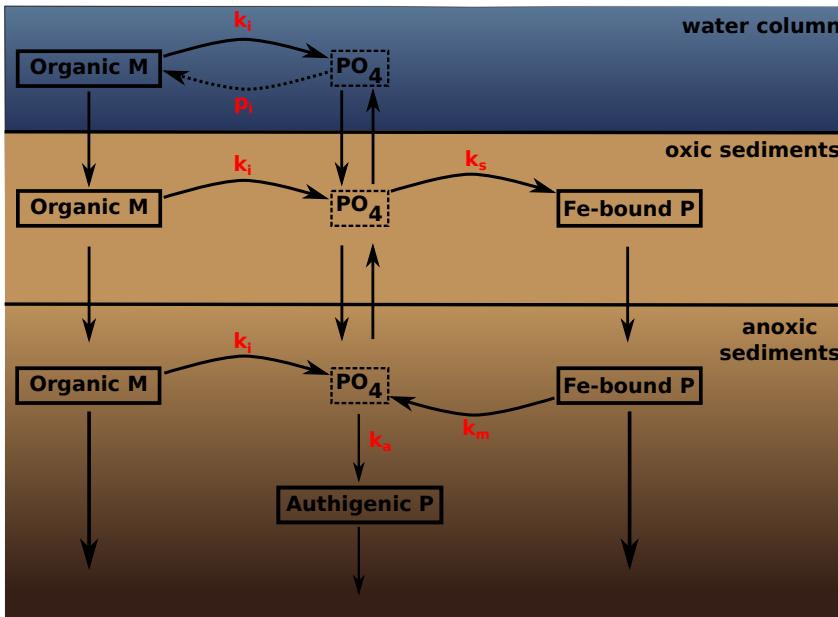


Figure 2. A schematic of the sedimentary P cycle in OMEN-SED. Red numbers represent kinetic rate constants for phosphorus dynamics (compare Table 10; p_i represents uptake rate of PO₄ via primary production in shallow environments). Adapted from Slomp et al. (1996).

2.2.5 Phosphate

The biogeochemical description of phosphorus (P) dynamics builds on the work of Slomp et al. 510 (1996) and accounts for phosphorus recycling through organic matter degradation, adsorption onto sediments and iron(III) hydroxides (Fe-bound P), as well as carbonate fluorapatite (CFA or authigenic P) formation (see Figure 2 for a schematic overview of the sedimentary P cycle). In the oxic zone of the sediment, PO₄ liberated through organic matter degradation can adsorb to iron(III) hydroxides forming Fe-bound P (or FeP, Slomp et al., 1998). Below the oxic zone, PO₄ is not only 515 produced via organic matter degradation but can also be released from the Fe-bound P pool due to the reduction of iron(III) hydroxides under anoxic conditions. Furthermore, in these zones phosphate concentrations build up and pore waters can thus become supersaturated with respect to carbonate fluorapatite, thus triggering the authigenic formation of CFA (Van Cappellen and Berner, 1988). Phosphorus bound in these authigenic minerals represents a permanent sink for reactive phosphorus 520 (Slomp et al., 1996). As for ammonium, the adsorption of P to the sediment matrix is treated as an equilibrium processes, parameterised with dimensionless adsorption coefficients for the oxic and anoxic zone, respectively ($K_{\text{PO}_4}^{\text{ox}}$, $K_{\text{PO}_4}^{\text{anox}}$ Slomp et al., 1998). The sorption and desorption of P to iron(III) hydroxides as well as the authigenic fluorapatite formation are described as first-order reactions with rate constants k_s , k_m and k_a , respectively (Table 10). The rate of the respective process is

525 calculated as the product of the rate constant and the difference between the current concentration (of PO₄ and FeP) and an equilibrium or asymptotic concentration Slomp et al. (1996). The asymptotic Fe-bound P concentration is FeP[∞] and the equilibrium concentration for P sorption and authigenic fluorapatite formation are PO₄^s and PO₄^a, respectively (Table 10). The last term in Eq. (24) and
 530 (25) represents sorption of PO₄ to FeP in the oxic zone, the last term in Eq. (26) and (27) is the release of PO₄ from the FeP pool and the 4th term in Eq. (27) represents the permanent loss of PO₄ to authigenic fluorapatite formation. The conservation equations for phosphate and Fe-bound P are thus given by:

1. Oxic zone ($z \leq z_{\text{ox}}$)

$$535 \frac{\partial \text{PO}_4^I}{\partial t} = \frac{D_{\text{PO}_4}}{1 + K_{\text{PO}_4}^{\text{ox}}} \frac{\partial^2 \text{PO}_4^I}{\partial z^2} - w \frac{\partial \text{PO}_4^I}{\partial z} + \frac{1-\phi}{\phi} \frac{1}{1 + K_{\text{PO}_4}^{\text{ox}}} \sum_i (\text{PC}_i \cdot k_i \cdot \text{POC}_i(z)) \\ - \frac{k_s}{1 + K_{\text{PO}_4}^{\text{ox}}} (\text{PO}_4^I - \text{PO}_4^s) \quad (24)$$

$$\frac{\partial \text{FeP}^I}{\partial t} = D_{\text{FeP}} \frac{\partial^2 \text{FeP}^I}{\partial z^2} - w \frac{\partial \text{FeP}^I}{\partial z} + \frac{\phi}{1-\phi} k_s (\text{PO}_4^I - \text{PO}_4^s) \quad (25)$$

2. Anoxic zones ($z_{\text{ox}} < z \leq z_{\infty}$)

$$540 \frac{\partial \text{FeP}^{II}}{\partial t} = D_{\text{FeP}} \frac{\partial^2 \text{FeP}^{II}}{\partial z^2} - w \frac{\partial \text{FeP}^{II}}{\partial z} - k_m (\text{FeP}^{II} - \text{FeP}^{\infty}) \quad (26)$$

$$\frac{\partial \text{PO}_4^{II}}{\partial t} = \frac{D_{\text{PO}_4}}{1 + K_{\text{PO}_4}^{\text{anox}}} \frac{\partial^2 \text{PO}_4^{II}}{\partial z^2} - w \frac{\partial \text{PO}_4^{II}}{\partial z} + \frac{1-\phi}{\phi} \frac{1}{1 + K_{\text{PO}_4}^{\text{anox}}} \sum_i (\text{PC}_i \cdot k_i \cdot \text{POC}_i(z)) \\ - \frac{k_a}{1 + K_{\text{PO}_4}^{\text{anox}}} (\text{PO}_4^{II} - \text{PO}_4^a) + \frac{(1-\phi)}{\phi} \frac{k_m}{1 + K_{\text{PO}_4}^{\text{anox}}} (\text{FeP}^{II} - \text{FeP}^{\infty}) \quad (27)$$

where D_{PO_4} denotes the diffusion coefficient for PO₄ which depends on the bioturbation status of
 545 the respective geochemical zone and $D_{\text{FeP}} = D_{\text{bio}}$ for $z \leq z_{\text{bio}}$ and $D_{\text{FeP}} = 0$ for $z > z_{\text{bio}}$ (compare Section 2.3.1). Integration of Eq. (24) - (27) yields the analytical solution and Table 6 summarises the boundary conditions applied in OMEN-SED. The model assumes known bottom water concentrations and equal concentrations and diffusive fluxes at z_{bio} and z_{ox} for both species. Additionally OMEN-SED considers no change in phosphate flux and an asymptotic Fe-bound P concentration at
 550 z_{∞} .

2.2.6 Dissolved Inorganic Carbon (DIC)

OMEN-SED accounts for the production of dissolved inorganic carbon (DIC) through organic matter degradation, as well as methane oxidation. Organic matter degradation produces dissolved inorganic carbon with a stoichiometric DIC : C ratio of 1:2 in the methanic zone and 1:1 in the rest of the
 555 sediment column (DIC^{CII} and DIC^{CI} respectively). DIC production through methane oxidation is implicitly taken into account through the boundary condition at z_{SO_4} . A mechanistic description of DIC production from CaCO₃ dissolution would lead to significant mathematical problems and is

DH: I think what I wrote before also applies: a asymptotic Fe-bound P concentration at z_{∞} is assumed! (but is the same as “no flux condition” as it is now?’)

Table 6. Boundary conditions for phosphate and Fe-bound P (FeP). For the boundaries we define: $z_-^- := \lim_{h \rightarrow 0} (z_- - h)$ and $z_-^+ := \lim_{h \rightarrow 0} (z_- + h)$.

Boundary	Condition	
$z = 0$	known concentration	1) $\text{PO}_4(0) = \text{PO}_{40}$
$z = z_{\text{bio}}$	continuity	2) $\text{PO}_4(z_{\text{bio}}^-) = \text{PO}_4(z_{\text{bio}}^+)$
	flux	3) $(D_{\text{PO}_4,0} + D_{\text{bio}}) \cdot \frac{\partial \text{PO}_4}{\partial z} _{z_{\text{bio}}^-} = D_{\text{PO}_4,0} \cdot \frac{\partial \text{PO}_4}{\partial z} _{z_{\text{bio}}^+}$
$z = z_{\text{ox}}$	continuity	4) $\text{PO}_4(z_{\text{ox}}^-) = \text{PO}_4(z_{\text{ox}}^+)$
	flux	5) $-\frac{D_{\text{PO}_4}}{1 + K_{\text{PO}_4}^{\text{ox}}} \cdot \frac{\partial \text{PO}_4}{\partial z} _{z_{\text{ox}}^-} = -\frac{D_{\text{PO}_4}}{1 + K_{\text{PO}_4}^{\text{anox}}} \cdot \frac{\partial \text{PO}_4}{\partial z} _{z_{\text{ox}}^+}$
$z = z_\infty$	flux	10) $\frac{\partial \text{PO}_4}{\partial z} _{z_\infty} = 0$
$z = 0$	known concentration	1) $\text{FeP}(0) = \text{FeP}_0$
$z = z_{\text{bio}}$	continuity	2) $\text{FeP}(z_{\text{bio}}^-) = \text{FeP}(z_{\text{bio}}^+)$
	flux	3) $\frac{\partial \text{FeP}}{\partial z} _{z_{\text{bio}}^-} = \frac{\partial \text{FeP}}{\partial z} _{z_{\text{bio}}^+}$
$z = z_{\text{ox}}$	continuity	4) $\text{FeP}(z_{\text{ox}}^-) = \text{FeP}(z_{\text{ox}}^+)$
	flux	5) $\frac{\partial \text{FeP}}{\partial z} _{z_{\text{ox}}^-} = \frac{\partial \text{FeP}}{\partial z} _{z_{\text{ox}}^+}$
$z = z_\infty$	asymptotic concentration	10) $\text{FeP}(z_\infty) = \text{FeP}_\infty$

therefore not included in the current version of OMEN-SED. The conservation equations for DIC are thus given by:

560

1. Oxic, nitrogenous and sulfidic zone ($z \leq z_{\text{SO}_4}$)

$$\frac{\partial \text{DIC}^I}{\partial t} = 0 = D_{\text{DIC}} \frac{\partial^2 \text{DIC}^I}{\partial z^2} - w \frac{\partial \text{DIC}^I}{\partial z} + \frac{1 - \phi}{\phi} \cdot \sum_i \text{DICC}^I \cdot k_i \cdot \text{POC}_i(z) \quad (28)$$

2. Methanic zone ($z_{\text{SO}_4} < z \leq z_\infty$)

565

$$\frac{\partial \text{DIC}^{II}}{\partial t} = 0 = D_{\text{DIC}} \frac{\partial^2 \text{DIC}^{II}}{\partial z^2} - w \frac{\partial \text{DIC}^{II}}{\partial z} + \frac{1 - \phi}{\phi} \cdot \sum_i \text{DICC}^{II} \cdot k_i \cdot \text{POC}_i(z) \quad (29)$$

where D_{DIC} denotes the diffusion coefficient for DIC which depends on the bioturbation status of the respective geochemical zone. Integration of Eq. (28) and (29) yields the analytical solution and Table 7 summarises the boundary conditions applied in OMEN-SED. A Dirichlet condition is applied at the sediment-water interface. In addition, the model assumes a zero diffusive flux through the lower boundary z_∞ and continuity across the bottom of the bioturbated zone, as well as the sulfate penetration depth. An additional flux boundary condition at z_{SO_4} , implicitly accounts for DIC production through anaerobic oxidation of methane (Table 7 Eq. 5).

2.2.7 Alkalinity

575 Organic matter degradation and secondary redox reactions exert a complex influence on alkalinity (e.g. Jourabchi et al., 2005; Wolf-Gladrow et al., 2007; Krumins et al., 2013). To model alkalinity, OMEN-SED divides the sediment column into four geochemical zones, where different equations

SA: TODO: again need to mention carboantes

Table 7. Boundary conditions for DIC. For the boundaries we define: $z_-^- := \lim_{h \rightarrow 0} (z_- - h)$ and $z_-^+ := \lim_{h \rightarrow 0} (z_- + h)$.

Boundary	Condition	
$z = 0$	known concentration	1) $\text{DIC}(0) = \text{DIC}_0$
$z = z_{\text{bio}}$	continuity	2) $\text{DIC}(z_{\text{bio}}^-) = \text{DIC}(z_{\text{bio}}^+)$
	flux	3) $-(D_{\text{DIC},0} + D_{\text{bio}}) \cdot \frac{\partial \text{DIC}}{\partial z} \Big _{z_{\text{bio}}^-} = -D_{\text{DIC},0} \cdot \frac{\partial \text{DIC}}{\partial z} \Big _{z_{\text{bio}}^+}$
$z = z_{\text{SO}_4}$	continuity	4) $\text{DIC}(z_{\text{SO}_4}^-) = \text{DIC}(z_{\text{SO}_4}^+)$
	flux (with AOM)	5) $-D_{\text{DIC}} \cdot \frac{\partial \text{DIC}}{\partial z} \Big _{z_{\text{SO}_4}^-} + \gamma_{\text{CH}_4} \cdot F_{\text{CH}_4}(z_{\text{SO}_4}) = -D_{\text{DIC}} \cdot \frac{\partial \text{DIC}}{\partial z} \Big _{z_{\text{SO}_4}^+}$
	where:	$F_{\text{CH}_4}(z_{\text{SO}_4}) = \frac{1-\phi}{\phi} \cdot \int_{z_{\text{SO}_4}}^{\infty} \sum_i \text{MC} \cdot k_i \cdot \text{POC}_i dz$
$z = z_\infty$	zero DIC flux	6) $\frac{\partial \text{DIC}}{\partial z} \Big _{z_\infty} = 0$

describe the biogeochemical processes using variable stoichiometric coefficients (compare values in Table 10). Above z_{ox} , the combined effects of NH₄ and P release due to aerobic OM degradation increases alkalinity according to ALK^{OX} whereas nitrification decreases alkalinity with stoichiometry ALK^{NIT}. In the remaining three zones anaerobic OM degradation generally results in an increase in alkalinity, with the exact magnitude depending on the nature of the terminal electron acceptor used (i.e. ALK^{DEN}, ALK^{SUL}, ALK^{MET}). In addition, the effect of secondary redox reactions, such as nitrification, sulfide and methane oxidation are implicitly accounted for in the boundary conditions. Again, a mechanistic description of ALK production from CaCO₃ dissolution would lead to significant mathematical problems and is therefore not included in the current version of OMEN-SED. In OMEN-SED, the conservation equations for alkalinity are thus given by:

1. Oxic zone ($z \leq z_{\text{ox}}$)

$$590 \quad \frac{\partial \text{ALK}^I}{\partial t} = 0 = D_{\text{ALK}} \frac{\partial^2 \text{ALK}^I}{\partial z^2} - w \frac{\partial \text{ALK}^I}{\partial z} + \frac{1-\phi}{\phi} \cdot \sum_i \left(\text{ALK}^{\text{NIT}} \cdot \frac{\gamma_{\text{NH}_4}}{1 + K_{\text{NH}_4}} \text{NC}_i + \text{ALK}^{\text{OX}} \right) \cdot k_i \cdot \text{POC}_i(z) \quad (30)$$

2. Denitrification or nitrogenous zone ($z_{\text{ox}} < z \leq z_{\text{NO}_3}$)

$$595 \quad \frac{\partial \text{ALK}^{II}}{\partial t} = 0 = D_{\text{ALK}} \frac{\partial^2 \text{ALK}^{II}}{\partial z^2} - w \frac{\partial \text{ALK}^{II}}{\partial z} + \frac{1-\phi}{\phi} \cdot \sum_i \text{ALK}^{\text{DEN}} \cdot k_i \cdot \text{POC}_i(z) \quad (31)$$

3. Sulfidic zone ($z_{\text{NO}_3} < z \leq z_{\text{SO}_4}$)

$$\frac{\partial \text{ALK}^{III}}{\partial t} = 0 = D_{\text{ALK}} \frac{\partial^2 \text{ALK}^{III}}{\partial z^2} - w \frac{\partial \text{ALK}^{III}}{\partial z} + \frac{1-\phi}{\phi} \cdot \sum_i \text{ALK}^{\text{SUL}} \cdot k_i \cdot \text{POC}_i(z) \quad (32)$$

4. Methanic zone ($z_{\text{SO}_4} < z \leq z_\infty$)

$$600 \quad \frac{\partial \text{ALK}^{IV}}{\partial t} = 0 = D_{\text{ALK}} \frac{\partial^2 \text{ALK}^{IV}}{\partial z^2} - w \frac{\partial \text{ALK}^{IV}}{\partial z} + \frac{1-\phi}{\phi} \cdot \sum_i \text{ALK}^{\text{MET}} \cdot k_i \cdot \text{POC}_i(z) \quad (33)$$

Table 8. Boundary conditions for alkalinity. For the boundaries we define: $z_-^- := \lim_{h \rightarrow 0} (z_- - h)$ and $z_-^+ := \lim_{h \rightarrow 0} (z_- + h)$.

Boundary	Condition	
$z = 0$	known concentration	1) $\text{ALK}(0) = \text{ALK}_0$
$z = z_{\text{bio}}$	continuity	2) $\text{ALK}(z_{\text{bio}}^-) = \text{ALK}(z_{\text{bio}}^+)$
	flux	3) $-(D_{\text{ALK},0} + D_{\text{bio}}) \cdot \frac{\partial \text{ALK}}{\partial z} _{z_{\text{bio}}^-} = -D_{\text{ALK},0} \cdot \frac{\partial \text{ALK}}{\partial z} _{z_{\text{bio}}^+}$
$z = z_{\text{ox}}$	continuity	4) $\text{ALK}(z_{\text{ox}}^-) = \text{ALK}(z_{\text{ox}}^+)$
	flux	5) $-D_{\text{ALK}} \cdot \frac{\partial \text{ALK}}{\partial z} _{z_{\text{ox}}^-} + F_{\text{ALK}}(z_{\text{ox}}) = -D_{\text{ALK}} \cdot \frac{\partial \text{ALK}}{\partial z} _{z_{\text{ox}}^+}$ $F_{\text{ALK}}(z_{\text{ox}}) = \frac{1-\phi}{\phi} \cdot \left(\text{ALK}^{\text{H}_2\text{S}} \cdot \gamma_{\text{H}_2\text{S}} \int_{z_{\text{NO}_3}}^{\text{SO}_4} \sum_i \text{SO}_4 \cdot k_i \cdot \text{POC}_i dz \right) + \frac{1-\phi}{\phi} \cdot \left(\text{ALK}^{\text{NIT}} \frac{\gamma_{\text{NH}_4}}{1+k_{\text{NH}_4}} \int_{z_{\text{NO}_3}}^{\infty} \sum_i \text{NC}_i \cdot k_i \cdot \text{POC}_i dz \right)$
	where:	
$z = z_{\text{NO}_3}$	continuity	6) $\text{ALK}(z_{\text{NO}_3}^-) = \text{ALK}(z_{\text{NO}_3}^+)$
	flux	7) $-D_{\text{ALK}} \cdot \frac{\partial \text{ALK}}{\partial z} _{z_{\text{NO}_3}^-} = -D_{\text{ALK}} \cdot \frac{\partial \text{ALK}}{\partial z} _{z_{\text{NO}_3}^+}$
$z = z_{\text{SO}_4}$	continuity	8) $\text{ALK}(z_{\text{SO}_4}^-) = \text{ALK}(z_{\text{SO}_4}^+)$
	flux (with AOM)	9) $-D_{\text{ALK}} \cdot \frac{\partial \text{ALK}}{\partial z} _{z_{\text{SO}_4}^-} + F_{\text{ALK}}(z_{\text{SO}_4}) = -D_{\text{ALK}} \cdot \frac{\partial \text{ALK}}{\partial z} _{z_{\text{SO}_4}^+}$ $F_{\text{ALK}}(z_{\text{SO}_4}) = \frac{1-\phi}{\phi} \cdot \left(\text{ALK}^{\text{AOM}} \gamma_{\text{CH}_4} \cdot \int_{z_{\text{SO}_4}}^{\infty} \sum_i k_i \cdot \text{POC}_i dz \right)$
	where:	
$z = z_{\infty}$	zero ALK flux	10) $\frac{\partial \text{ALK}}{\partial z} _{z_{\infty}} = 0$

where D_{ALK} denotes the diffusion coefficient for alkalinity which depends on the bioturbation status of the respective geochemical zone. Integration of Eq. (30) - (33) yields the analytical solution and Table 8 summarises the boundary conditions applied in OMEN-SED. A Dirichlet boundary condition is applied at the sediment-water interface. The decrease of alkalinity due to oxidation of reduced species produced in the anoxic zones (with stoichiometry ALK^{NIT} and $\text{ALK}^{\text{H}_2\text{S}}$) is implicitly taken into account through the flux boundary condition at z_{ox} (Table 8 Eq. 5). Furthermore, the oxidation of methane by sulfate reduction increases alkalinity with stoichiometry ALK^{AOM} which is accounted for through the flux boundary condition at z_{SO_4} (Table 8 Eq. 9). At the lower boundary z_{∞} a zero diffusive flux condition is applied.

2.3 Determination of Integration Constants

The integration constants of all general analytical solutions derived above change in response to changing boundary conditions. Thus, OMEN-SED has to re-determine integration constants for each dynamic zone (i.e. z_{ox} , z_{bio} , z_{NO_3} and z_{SO_4}) at every time step for all biogeochemical tracers. The bioturbation boundary poses a particular challenge as it can theoretically occur in any of the dynamic geochemical zones (Fig. 3). Therefore, in order to generalise and simplify this recurring boundary matching problem, an independent, generic algorithm is implemented (rather than using multiple fully-worked-out algebraic solutions for each possible case and every biogeochemical tracer). The algorithm only has to solve a two-simultaneous-equation problem.

620 **2.3.1 Generic Boundary Condition Matching (GBCM)**

As discussed in Section 2.1, the solution of the general steady-state transport-reaction equation (Eq. (2)) for a generic tracer C is of the general form:

$$C(z) = A \exp(az) + B \exp(bz) + \sum_j \frac{\alpha_j}{D\beta_j^2 - w\beta_j - k} \cdot \exp(-\beta_j z) + \frac{Q}{k} \quad (34)$$

625 and can therefore be expressed as:

$$C(z) = A \cdot E(z) + B \cdot F(z) + G(z) \quad (35)$$

where $E(z)$, $F(z)$ are the homogeneous solutions of the ODE, $G(z)$ the particular integral (collectively called the basis functions), and A , B are the integration constants that must be determined with
630 the boundary conditions (shown in Fig. 3 for the whole sediment column).

Each internal boundary matching problem (i.e. excluding $z = 0$ and $z = z_\infty$) involves matching continuity and flux for the two solutions of the respective reaction-transport equation above, $C_U(z)$ (= 'upper'), and below, $C_L(z)$ (= 'lower'), the dynamic boundary at $z = z_b$:

$$C_U(z) = A_U \cdot E_U(z) + B_U \cdot F_U(z) + G_U(z) \quad (36)$$

635 $C_L(z) = A_L \cdot E_L(z) + B_L \cdot F_L(z) + G_L(z). \quad (37)$

OMEN-SED generally applies concentration continuity and flux boundary conditions at its internal, dynamic boundaries:

Continuity (where for generality we allow a discontinuity V_b)

640 $C_U(z_b) = C_L(z_b) + V_b \quad (38)$

Flux

$$D_U C'_U(z_b) + w C_U(z_b) = D_L C'_L(z_b) + w C_L(z_b) + F_b \quad (39)$$

645 where w is advection, D are the diffusion coefficients and F_b is any flux discontinuity (e.g. resulting from secondary redox reactions).

Considering that the advective flux above and below the boundary is equal (i.e. $w C_U(z_b) = w C_L(z_b)$) and substituting the general ODE solutions (36), (37), the boundary conditions can be represented as two equations connecting the four integration constants:

$$\begin{pmatrix} E_U & F_U \\ D_U E'_U & D_U F'_U \end{pmatrix} \begin{pmatrix} A_U \\ B_U \end{pmatrix} = \begin{pmatrix} E_L & F_L \\ D_L E'_L & D_L F'_L \end{pmatrix} \begin{pmatrix} A_L \\ B_L \end{pmatrix} + \begin{pmatrix} G_L - G_U + V_b \\ D_L G'_L - D_U G'_U + F_b - wV_b \end{pmatrix} \quad (40)$$

650

where the ODE solutions E , F , G are all evaluated at z_b .

Equation (40) can now be solved to give A_U and B_U as a function of the integration constants from the layer below (A_L and B_L), thereby constructing a piecewise solution for both layers, with just two integration constants (this is implemented in the function **benthic_utils.matchsoln** of OMEN-
655 SED):

$$\begin{pmatrix} A_U \\ B_U \end{pmatrix} = \begin{pmatrix} c_1 & c_2 \\ c_3 & c_4 \end{pmatrix} \begin{pmatrix} A_L \\ B_L \end{pmatrix} + \begin{pmatrix} d_1 \\ d_2 \end{pmatrix}. \quad (41)$$

Using Eq. (41), $C_U(z)$ in (36) can now be rewritten as a function of A_L and B_L (implemented in **benthic_utils.xformsoln**):

$$C_U(z) = (c_1 A_L + c_2 B_L + d_1) \cdot E_U(z) + (c_3 A_L + c_4 B_L + d_2) \cdot F_U(z) + G_U(z) \quad (42)$$

660 and hence define the “transformed” basis functions $E_U^*(z)$, $F_U^*(z)$, $G_U^*(z)$ such that:

$$C_U(z) = A_L \cdot E_U^*(z) + B_L \cdot F_U^*(z) + G_U^*(z) \quad (43)$$

where

$$\begin{aligned} E_U^*(z) &= c_1 E_U(z) + c_3 F_U(z) \\ F_U^*(z) &= c_2 E_U(z) + c_4 F_U(z) \\ 665 \quad G_U^*(z) &= G_U(z) + d_1 E_U(z) + d_2 F_U(z) \end{aligned} \quad (44)$$

Equations (41), (43) and (44) can now be consecutively applied for each of the dynamic biogeochemical zone boundaries, starting at the bottom of the sediment column. The net result is a piecewise solution of the whole sediment column with just two integration constants (coming from
670 the lowest layer), which can then be solved for by applying the boundary conditions at the sediment-water interface and the bottom of the sediments.

2.3.2 Abstracting out the bioturbation boundary

The bioturbation boundary affects the diffusion coefficient of the modelled solutes, as well as the conservation equation of organic matter (and thereby the exact form of each reaction-transport equation). This boundary is particularly inconvenient as it can, in principle, occur in the middle of any of the dynamically shifting biogeochemical zones and therefore generate multiple cases (Fig. 3). The GBCM algorithm described above is thus not only used to construct a piecewise solution of the whole sediment column, but also to abstract out the bioturbation boundary. For each biogeochemical zone the “bioturbation-status” is initially tested (i.e. fully bioturbated, fully non-bioturbated, or crossing
680 the bioturbation boundary). Therefore, the upper and lower boundaries for the different zones (e.g. for the nitrogenous zone: $z_U = z_{\text{ox}}$, $z_L = z_{\text{NO}_3}$), as well as the respective reactive terms and diffusion coefficients (bioturbated and non-bioturbated) are passed over to the routine **zTOC.prepfg_l12**

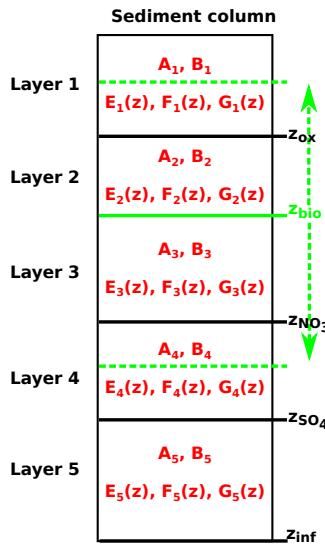


Figure 3. Schematic of the generic boundary condition matching (GBCM) problem. Showing the resulting integration constants (A_i , B_i) and ODE solutions (E_i , F_i , G_i) for the different sediment layers and the variable bioturbation boundary.

where the bioturbation-status is determined. In case the bioturbation depth is located within this zone (i.e. $z_U < z_{\text{bio}} < z_L$) a piecewise solution for this layer is constructed. Therefore, the reactive terms
685 and diffusion coefficients are handed over to the routines **zTOC.calcfg_I1** and **zTOC.calcfg_I2** which calculate the basis functions (E_U, F_U, G_U and E_L, F_L, G_L) and their derivatives for the bioturbated and the non-bioturbated part of this specific geochemical zone. The concentration and flux for both solutions at z_{bio} are matched and the coefficients $c_1, c_2, c_3, c_4, d_1, d_2$ (as in Eq. (41)) are calculated by the routine **benthic_utils.matchsoln**. These coefficients and the "bioturbation-status"
690 of the layer are passed back to the main GBCM algorithm where they can be used by the routine **benthic_utils.xformsoln** to calculate the "transformed" basis functions ($E_U^*(z)$, $F_U^*(z)$, $G_U^*(z)$) such that both layers are expressed in the same basis (compare Eq. (42 - 44)).

For instance, in the case of sulfate, **zTOC.prepfg_I12** is called three times before the actual profile is calculated (once per zone: oxic, nitrogenous, sulfidic) and hands back the information
695 about the "bioturbation-status" of the three layers and the coefficients $c_1, c_2, c_3, c_4, d_1, d_2$ for the biogeochemical zone including the bioturbation depth. When calculating the complete piecewise solution for the sediment column, this information is passed to the function **zTOC.calcfg_I12** which sorts out the correct solution type to use. The main GBCM algorithm therefore never needs to know whether it is dealing with a piecewise solution (i.e. matched across the bioturbation boundary) or a
700 "simple" solution (i.e. the layer is fully bioturbated or fully non-bioturbated).

2.4 Model Parameters

The following section provides a summary of global relationships used to constrain reaction and transport parameters in OMEN-SED. Table 9 synthesises sediment and transport parameters, while table 10 provides an overview of all biogeochemical parameters used in OMEN-SED.

705 2.4.1 Transport Parameters

The burial of sediments and pore water is directly related to the accumulation of new material on the seafloor (i.e. sedimentation, Burdige, 2006). This results in a downward advective flux of older sediment material and pore water in relation to the sediment-water interface. When coupled to an ocean model, its sedimentation flux can be readily used in OMEN-SED. The stand-alone version of
710 OMEN-SED uses the empirical global relationship between sediment accumulation rate (cm yr^{-1}) and seafloor depth (m) of Middelburg et al. (1997):

$$w = 3.3 \cdot 10^{-0.87478367 - 0.00043512 \cdot \text{depth}}, \quad (45)$$

As mentioned before (Section 2.1), the diffusion coefficient of species i is calculated as $D_i = D_{i,0} + D_{\text{bio}} = D_{\text{mol},i} \cdot f_{ir} + D_{\text{bio}}$ for dissolved species and $D_i = D_{\text{bio}}$ for solid species. The bioturbation coefficient D_{bio} ($\text{cm}^2 \text{ yr}^{-1}$) is constant in the bioturbated zone and also follows the empirical relationship by Middelburg et al. (1997):
715

$$D_{\text{bio}} = 5.2 \cdot 10^{0.76241122 - 0.00039724 \cdot \text{depth}} \quad (46)$$

Observations indicate that bioturbation is largely restricted to the upper 10 cm of the sediments and is only marginally related to seafloor depth (e.g. Boudreau, 1998; Teal et al., 2010). Therefore,
720 OMEN-SED imposes a globally invariant bioturbation depth z_{bio} of 10 cm. In case the bottom water oxygen concentration is below 5 nmol cm^{-3} REF for it??? infaunal activity is assumed to cease and $z_{\text{bio}} = 0.01 \text{ cm}$. We choose a low value unequal to zero in order to simplify the implementation of the model. This approach ensures that the sediment column always consists of a bioturbated (even though very small for the low oxygen condition) and a non-bioturbated zone, thus the same
725 GBCM algorithm can be used to solve the conservation equations. Furthermore, when OMEN-SED is coupled to an Earth system model the same method can be used to convert the POC depositional flux into a SWI concentration (i.e. the flux needs to be converted assuming bioturbation, see Section 4.1).

Bioirrigation (i.e. the pumping activity by burrow-dwelling animals) exchanges burrow water
730 with overlying water and may enhance the SWI-flux of solutes (Aller, 1984, 1988). Several approaches exist to incorporate this into a 1-D diagenetic model, for instance as a non-local transport/exchange process (Boudreau, 1984; Emerson et al., 1984) or as an enhancement factor of the molecular diffusion coefficient (Devol and Christensen, 1993; Soetaert et al., 1996). In OMEN-SED the latter approach is applied and the apparent “bio-diffusion” coefficient is calculated as

SA: need to explain why not 0; **DH:** Decent explanation?

735 $D_{i,0} = D_{\text{mol},i} \cdot f_{ir}$. Soetaert et al. (1996) derived an empirical relationship between f_{ir} and seafloor depth ($f_{ir} = \text{Min}\{1; 15.9 \cdot \text{depth}^{-0.43}\}$) based on observations from Archer and Devol (1992) and Devol and Christensen (1993). As this relationship just varies for depth above ~ 623 m (with a maximum value of 3 at ~ 50 m) a constant value of $f_{ir} = 1$ is used in the default OMEN-SED configuration. The specific molecular diffusion coefficients $D_{\text{mol},i}$ are corrected for sediment porosity ϕ ,

740 tortuosity F and are linearly interpolated for an ambient temperature T using zero-degree coefficients D_i^0 and temperature-dependent diffusion coefficients D_i^T (Soetaert et al., 1996):

$$D_{\text{mol},i} = (D_i^0 + D_i^T \cdot T) \cdot \frac{1}{\phi \cdot F}.$$

Tortuosity can be expressed in terms of porosity as $F = \frac{1}{\phi^m}$ (Ullman and Aller, 1982) with the exponent m varying according to the type of sediment (here $m = 3$ is used representing muddy sediments with high porosity). Values for D_i^T and D_i^0 are summarised in Table 9 and are adapted from Li and Gregory (1974), Schulz (2006) and Gypens et al. (2008).

2.4.2 Stoichiometries and reaction parameters

The first-order organic matter degradation constants of compound class i , k_i (yr^{-1}), are assumed invariant along the sediment column and therefore independent of the nature of the terminal electron acceptor. The rate constants can be altered manually to fit observed sediment profiles (compare modelled profiles in Section 3.2) or related to a master variable provided by a coupled Earth system model (e.g. sedimentation rate, see Section 4.2). The partitioning of the bulk OM pool into reactivity classes (f_i) needs to be specified manually in the stand-alone version or can be provided by the ESM. Organic matter degradation releases N, P and DIC to the pore water using Redfield molar ratios (Redfield, 1963) and consumes TEA with specific stoichiometries (O_2C , NO_3C , SO_4C) as summarised in Table 10. Table 15 in the appendix provides a list of reactions and their stoichiometries as implemented in OMEN-SED. The effect of OM degradation and secondary redox reactions on total alkalinity is also accounted for via reaction specific stoichiometries representing the release of NH_4 , H_2S and P and is based on Jourabchi et al. (2005). As the reoxidation of reduced substances produced during OM degradation may be incomplete, we allow part of the NH_4 , H_2S and CH_4 flux to escape reoxidation. The secondary redox parameters (i.e. γ_{NH_4} , $\gamma_{\text{H}_2\text{S}}$, γ_{CH_4}) account for the fraction of reduced substances that are reoxidised and would be ideally parameterised for instance in relation to bottom water oxygen concentration or oxygen penetration depth (z_{ox}). Gypens et al. (2008) for example expressed γ_{NH_4} as a function of oxygen penetration depth ($\gamma_{\text{NH}_4} = 0.243 \cdot \ln(z_{\text{ox}}) + 1.8479$) based on a fitting exercises to a numerical model and showed that the fraction varies between 0.2 for $z_{\text{ox}} = 0.1\text{cm}$ and 1.0 for $z_{\text{ox}} > 3\text{cm}$. Due to mathematical constraints for finding an analytical solution to the model equations these fractions take constant values generally representing oxygenated deep sea conditions. **More on that? REFERENCE where these values come from!???** The instantaneous equilibrium adsorption coefficients of NH_4 and PO_4 (K_{NH_4} , $K_{\text{PO}_4}^{\text{ox}}$, $K_{\text{PO}_4}^{\text{anox}}$) are based on

Table 9. Sediment characteristics and transport parameters.

Parameter	Unit	Value	Description/Source
ρ_{sed}	g cm^{-3}	2.6	Sediment density
w	cm yr^{-1}	Fct. of seafloor depth or from ESM	Advection/Sediment accumulation rate (Middelburg et al., 1997)
z_{bio}	cm	10 or 0.01	Bioturbation depth (Boudreau, 1998; Teal et al., 2010)
D_{bio}	$\text{cm}^2 \text{yr}^{-1}$	Fct. of seafloor depth	Bioturbation coefficient (Middelburg et al., 1997)
ϕ	-	0.85	Porosity
F	-	$\frac{1}{\phi^m}$	Tortuosity, here m=3
f_{ir}	-	1	Irrigation factor
Diffusion coefficients (Li and Gregory, 1974; Schulz, 2006; Gypens et al., 2008)			
$D_{\text{O}_2}^0$	$\text{cm}^2 \text{yr}^{-1}$	348.62	Molecular diffusion coefficient of oxygen at 0°C
$D_{\text{O}_2}^T$	$\text{cm}^2 \text{yr}^{-1} \text{°C}^{-1}$	14.09	Diffusion coefficient for linear temp. dependence of oxygen
$D_{\text{NO}_3}^0$	$\text{cm}^2 \text{yr}^{-1}$	308.42	Molecular diffusion coefficient of nitrate at 0°C
$D_{\text{NO}_3}^T$	$\text{cm}^2 \text{yr}^{-1} \text{°C}^{-1}$	12.26	Diffusion coefficient for linear temp. dependence of nitrate
$D_{\text{NH}_4}^0$	$\text{cm}^2 \text{yr}^{-1}$	309.05	Molecular diffusion coefficient of ammonium at 0°C
$D_{\text{NH}_4}^T$	$\text{cm}^2 \text{yr}^{-1} \text{°C}^{-1}$	12.26	Diffusion coefficient for linear temp. dependence of ammonium
$D_{\text{SO}_4}^0$	$\text{cm}^2 \text{yr}^{-1}$	157.68	Molecular diffusion coefficient of sulfate at 0°C
$D_{\text{SO}_4}^T$	$\text{cm}^2 \text{yr}^{-1} \text{°C}^{-1}$	7.88	Diffusion coefficient for linear temp. dependence of sulfate
$D_{\text{H}_2\text{S}}^0$	$\text{cm}^2 \text{yr}^{-1}$	307.48	Molecular diffusion coefficient of sulfide at 0°C
$D_{\text{H}_2\text{S}}^T$	$\text{cm}^2 \text{yr}^{-1} \text{°C}^{-1}$	9.64	Diffusion coefficient for linear temp. dependence of sulfide
$D_{\text{PO}_4}^0$	$\text{cm}^2 \text{yr}^{-1}$	112.91	Molecular diffusion coefficient of phosphate at 0°C
$D_{\text{PO}_4}^T$	$\text{cm}^2 \text{yr}^{-1} \text{°C}^{-1}$	5.59	Diffusion coefficient for linear temp. dependence of phosphate
D_{DIC}^0	$\text{cm}^2 \text{yr}^{-1}$	151.69	Molecular diffusion coefficient of DIC at 0°C
D_{DIC}^T	$\text{cm}^2 \text{yr}^{-1} \text{°C}^{-1}$	7.93	Diffusion coefficient for linear temp. dependence of DIC
D_{ALK}^0	$\text{cm}^2 \text{yr}^{-1}$	151.69	Molecular diffusion coefficient of ALK at 0°C
D_{ALK}^T	$\text{cm}^2 \text{yr}^{-1} \text{°C}^{-1}$	7.93	Diffusion coefficient for linear temp. dependence of ALK
Note: DIC and ALK coefficients are the values of HCO_3^- from Schulz (2006).			

770 Wang and Van Cappellen (1996) and Slomp et al. (1998), respectively. The first order rate constants
 for sorption of PO_4 to Fe oxides (k_s), release of PO_4 from Fe-bound P due to Fe-oxide reduction
 (k_m) and authigenic CFA precipitation (k_a), as well as the pore water equilibrium concentrations for
 P sorption and CFA precipitation (PO_4^s , PO_4^a) and the asymptotic concentration for Fe-bound P
 (FeP^∞) are taken from Slomp et al. (1996). See Table 10 for a complete summary of the parameters
 775 and their values.

3 Stand-alone sensitivity analysis and case studies

3.1 Sensitivity Analysis

3.1.1 Methodology

Model parameters implicitly account for processes that are not explicitly resolved. They are notoriously difficult to constrain and thus a primary source of uncertainty for numerical and analytical models. A comprehensive sensitivity analysis can help quantify this uncertainty and identify the most sensitive parameters. More specifically, sensitivity analysis is used to investigate how the variations in the outputs (y_1, \dots, y_N) of a model can be attributed to variations in the different input parameters (x_1, \dots, x_M , Pianosi et al., 2016). Different types of sensitivity indices, which quantify the relative influence of parameter x_i on output y_j with a scalar $S_{i,j}$ (for $i \in \{1, \dots, M\}$ and $j \in \{1, \dots, N\}$), can be calculated, ranging from simple one-at-a-time methods to statistical evaluations of the output distribution (e.g. variance-based or density-based approaches Pianosi et al., 2016). The latter indices take values between zero and one ($S_{i,j} \in [0, 1]$), where zero indicates a non-influential parameter and a higher value a more influential parameter. Here, sensitivity analysis is used mainly to identify
 780 which parameters have the largest impact on the different model outputs and therefore require more careful calibration. As the probability density functions of our model outputs (i.e. the resulting SWI-fluxes) are generally highly-skewed towards extreme organic matter degradation rates (not shown)
 785 variance-based sensitivity indices may not be a suitable proxy for output uncertainty (Pianosi et al., 2016). Hence, instead the density-based PAWN method by Pianosi and Wagener (2015) is employed
 790 which considers the entire conditional and unconditional Cumulative Distribution Function (CDF) of the model output rather than its variance only. The unconditional CDF, $F_y(y)$, of output y is obtained when all uncertain parameters (x_1, \dots, x_M) are varied simultaneously, and the conditional CDFs,
 $F_{y|x_i}(y)$, are obtained when all inputs but the i -th parameter are varied (i.e. x_i is fixed to a so-called
 795 conditioning value). The sensitivity index of parameter i is measured by the distance between the
 800 two CDFs using the Kolmogorov-Smirnov statistic (Kolmogorov, 1933; Smirnov, 1939), i.e.:

$$S_i = \max_{x_i} \max_y |F_y(y) - F_{y|x_i}(y)|. \quad (47)$$

Since $F_{y|x_i}(y)$ accounts for what happens when the variability due to x_i is removed, the distance between the two CDFs provides a measure of the effects of x_i on the output y . Due to the model

Table 10. Values for biogeochemical parameters used in OMEN-SED. The variables x , y and z denote the elemental ratio of carbon, nitrogen and phosphorus of the degrading organic matter (here set to $C : N : P = 106 : 16 : 1$).

Parameter/Variable	Unit	Value	Description
Stoichiometric factors and molecular ratios			
NC _i	mol/mol	$\frac{y}{x} = \frac{16}{106}$	Nitrogen to carbon ratio
PC _i	mol/mol	$\frac{z}{x} = \frac{1}{106}$	Phosphorus to carbon ratio
MC	mol/mol	0.5	Methane to carbon ratio
			produced during methanogenesis
DICC ^I	mol/mol	1.0	DIC to carbon ratio until z _{SO₄}
DICC ^{II}	mol/mol	0.5	DIC to carbon ratio below z _{SO₄}
O ₂ C	mol/mol	$\frac{x+2y}{x} = \frac{138}{106}$	Oxygen to carbon ratio
NO ₃ C	mol/mol	$\frac{4x+3y}{5x} = \frac{94.4}{106}$	Nitrate to carbon ratio
SO ₄ C	mol/mol	$\frac{1}{2}O_2C = \frac{138}{212}$	Sulfate to carbon ratio
ALK ^{OX}	mol/mol	$\frac{y-2z}{x} = \frac{14}{106}$	ALK from aerobic degradation
ALK ^{NIT}	mol/mol	-2	ALK from nitrification
ALK ^{DEN}	mol/mol	$\frac{4x+3y-10z}{5x} = \frac{92.4}{106}$	ALK from denitrification
ALK ^{SUL}	mol/mol	$\frac{x+y-2z}{x} = \frac{120}{106}$	ALK from sulfate reduction
ALK ^{MET}	mol/mol	$\frac{y-2z}{x} = \frac{14}{106}$	ALK from methanogenesis
ALK ^{H₂S}	mol/mol	-2	ALK from H ₂ S oxidation
ALK ^{AOM}	mol/mol	2	ALK from AOM
Secondary reaction parameters			
γ _{NH₄}	-	0.9	Fraction of NH ₄ that is nitrified
γ _{H₂S}	-	0.95	Fraction of H ₂ S that is oxidised
γ _{CH₄}	-	0.99	Fraction of CH ₄ that is oxidised
Adsorption coefficients (Wang and Van Cappellen, 1996; Slomp et al., 1998)			
K _{NH₄}	-	1.4	NH ₄ adsorption coefficient
K _{PO₄} ^{OX} , K _{PO₄} ^{ANOX}	-	200.0, 2.0	PO ₄ adsorption coefficient (oxic, anoxic)
P related parameters (Slomp et al., 1996)			
k _s	yr ⁻¹	94.9	Rate constant for PO ₄ sorption
k _m	yr ⁻¹	0.193	Rate constant for Fe-bound P release
k _a	yr ⁻¹	0.365	Rate constant for authigenic CFA precipitation
PO ₄ ^s	mol cm ⁻³	$1 \cdot 10^{-9}$	Equilibrium conc. for P sorption
FeP [∞]	mol cm ⁻³	$1.99 \cdot 10^{-10}$	Asymptotic concentration for Fe-bound P
PO ₄ ^a	mol cm ⁻³	$3.7 \cdot 10^{-9}$	Equilibrium conc. for authigenic P precipitation

DH: ALK^{OX} correct?
y=NH₄ prod.; -2z=P release

Table 11. Range of model parameters used for sensitivity analysis of model predicted output.

Parameter	Description	Units	Minimum	Maximum	Source
k_1	labile OM degradation constant	yr^{-1}	$1e^{-4}$	5.0	(1)
\tilde{k}_2	order of refractory OM degradation constant ($k_2 = \tilde{k}_2 \cdot k_1$)	-	$1e^{-4}$	$1e^{-1}$	(1)
f_1	fraction of labile OM	-	0.02	0.98	-
K_{NH_4}	Adsorption coefficient	-	0.8	1.7	(2)
γ_{NH_4}	NH_4 fraction oxidised		0.5	1.0	-
$\gamma_{\text{H}_2\text{S}}$	H_2S fraction oxidised		0.5	1.0	-
$K_{\text{PO}_4}^{\text{ox}}$	Adsorption coeff. oxic	-	100.0	400.0	(3)
$K_{\text{PO}_4}^{\text{anox}}$	Adsorption coeff. anoxic	-	1.3	2.0	(3)
k_s	kinetic P sorption	yr^{-1}	0.1	100.0	(4, 5)
k_m	Fe-bound P release	yr^{-1}	0.015	0.02	(4, 5)
k_a	authigenic P formation	yr^{-1}	0.001	10.0	(4, 6)

Sources: (1) Arndt et al. (2013); (2): Van Cappellen and Wang (1996); (3): Krom and Berner (1980)
(4): Gypens et al. (2008); (5): Slomp et al. (1996); (6): Van Cappellen and Berner (1988)

Table 12. Model boundary conditions for the two idealised sediment conditions used for the sensitivity analysis (Fig. 5 and 6). All solute concentrations are in nmol cm^{-3} .

Depth (m)	Temp. ($^{\circ}\text{C}$)	OC (wt%)	O_2	NO_3	SO_4	PO_4	z_{bio} (cm)
400	8.0	2.0	0.0	40.0	28,000	40.0	0.001
4000	1.5	1.0	300.0	20.0	28,000	40.0	10.0

complexity it is impossible to compute the sensitivity indices analytically. Therefore, they are approximated from a Latin-Hypercube sampling of parameter inputs and calculated outputs. For a brief description of the methodology see Fig. 4. For more details we refer the interested reader to Pianosi and Wagener (2015).

The PAWN method, as implemented within the Sensitivity Analysis for Everyone (SAFE) matlab toolbox (Pianosi et al., 2015), is used to investigate $M = 11$ model parameters for ranges as specified in Table 11. Sensitivity indices for all resulting SWI-fluxes for two idealised sediment conditions (i.e. anoxic at 400 m and oxic at 4000 m, see Table 12) are calculated. We use $NU = 200$ samples to estimate the unconditional CDF, $NC = 100$ samples to estimate the conditional CDFs and $n = 10$ conditioning points. Thus as $N_{\text{eval}} = 200 + 100 \cdot 10 \cdot 11$, 11200 model evaluations are performed for each sediment condition. The resulting indices are then translated into a color code and summarised in a pattern plot to simplify comparison (Fig. 5).

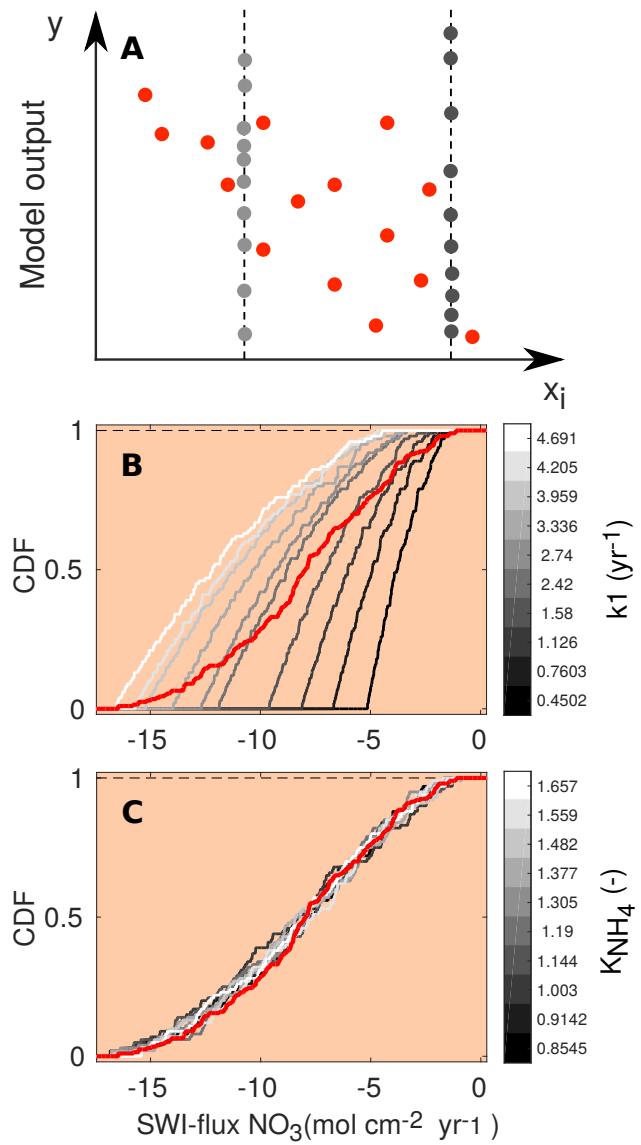


Figure 4. A: Schematic of the PAWN method, plotting an uncertain parameter (x_i) against a generic model output (y). Red dots represent points for calculating the unconditional CDF (NU, here 15), grey dots are points for calculating each conditional CDF (NC, here 10) with $n = 2$ conditioning points as an example. The user can change the values of NU, NC and n . The number of model evaluations equals $N_{\text{eval}} = \text{NU} + n \cdot \text{NC} \cdot M$, where M is the number of uncertain input parameters. B + C: Two examples of CDFs of the model calculated SWI-flux of NO_3 using NU = 200, NC = 100 and $n = 10$. The red lines are the unconditional distribution functions $F_y(\text{NO}_3)$ and the grey lines are the conditional distribution functions $F_{y|x_i}(\text{NO}_3)$ at different fixed values of the input parameters k_1 (B) and K_{NH_4} (C). As the maximal distance between conditional CDFs and unconditional CDF is greater for k_1 , this parameter is more influential for the model output (here SWI-flux of NO_3 , compare Fig. 5).

3.1.2 Results

Fig. 5 summarises results of the sensitivity analysis as a colour map. Results indicate that generally the most significant parameters for all model outputs are the degradation rate constant for the labile OM pool (k_1) and the fraction of this pool to the total OM stock (f_1). Other parameters play a minor role for the SWI-fluxes, with the exception of the secondary redox parameters (i.e. γ_{NH_4} , $\gamma_{\text{H}_2\text{S}}$) in the oxic scenario. Here, NH_4 , SO_4 and H_2S are very sensitive to changes in γ_{NH_4} and $\gamma_{\text{H}_2\text{S}}$, as these parameters determine how much of the respective TEA is produced in situ via reoxidation, thus affecting the resulting SWI-fluxes. For the oxic scenario, the reoxidation of H_2S produced in the sulfidic layer ($\gamma_{\text{H}_2\text{S}}$, Table 8 Eq. 5) also has a strong influence on alkalinity as it decreases alkalinity by 2 moles per mole of S oxidized ($\text{ALK}^{\text{H}_2\text{S}}$, Table 10). For the anoxic scenario the secondary redox parameters are essentially non-influential as no O_2 is available for the reoxidation of reduced substances. Especially for the oxic condition the PO_4 SWI-flux appears to be insensitive to P-related parameters (i.e. $K_{\text{PO}_4}^{\text{ox}}$, $K_{\text{PO}_4}^{\text{anox}}$, k_s , k_m , k_a) as the majority is absorbed to Fe-oxides. The sensitivities change if other PO_4 related equilibrium concentrations PO_4^s , PO_4^a and FeP^∞ are used (not shown). Overall the results of the sensitivity analysis are in line with what one would expect from a diagenetic model and thus provide ground to confirm that OMEN-SED provides sensible results. AR: A little about what we've learned from doing this!?

We further explore the sensitivity of simulated sediment-water exchange fluxes to variations in organic matter degradation parameters by varying k_1 , f_1 and \tilde{k}_2 while all other model parameters are set to their default values (Tables 9 and 10). Minimum and maximum values for k_1 , \tilde{k}_2 and f_1 in the shallow ocean are as in Table 11. For the deep sea condition we account for the presence of more refractory OM by sampling $f_1 \in [0.02, 0.3]$, whereas the variation of k_1 and \tilde{k}_2 is as in the shallow ocean. The parameter space is sampled using another Latin-Hypercube approach with sample sizes of $N = 3500$ for each idealised sediment condition. Figure 6 summarises the results of the sensitivity study and the ranges of observed O_2 and NO_3 sediment-water interface fluxes extracted from a global database (Stolpovsky et al., 2015) are indicated on the colour scale. Figure 6 shows that the ranges of SWI-fluxes simulated with OMEN-SED are comparable to the observed ranges reported by Stolpovsky et al. (2015). The colour patterns in Figure 6 A and B also reveal the complex interplay between the amount of labile OM f_1 and its degradation rate k_1 for the resulting SWI-fluxes of NO_3 in anoxic sediments and O_2 in aerobic sediments. In general, a higher degradation rate in combination with more labile OM available leads to a higher SWI-flux. However, higher fluxes extend over a larger range of k_1 -values when the amount of labile OM f_1 is high. The absence of a colour pattern in Figure 6 C highlights the limited interaction of the two model parameters for NO_3 SWI-fluxes under oxic conditions. AR: Take home message?

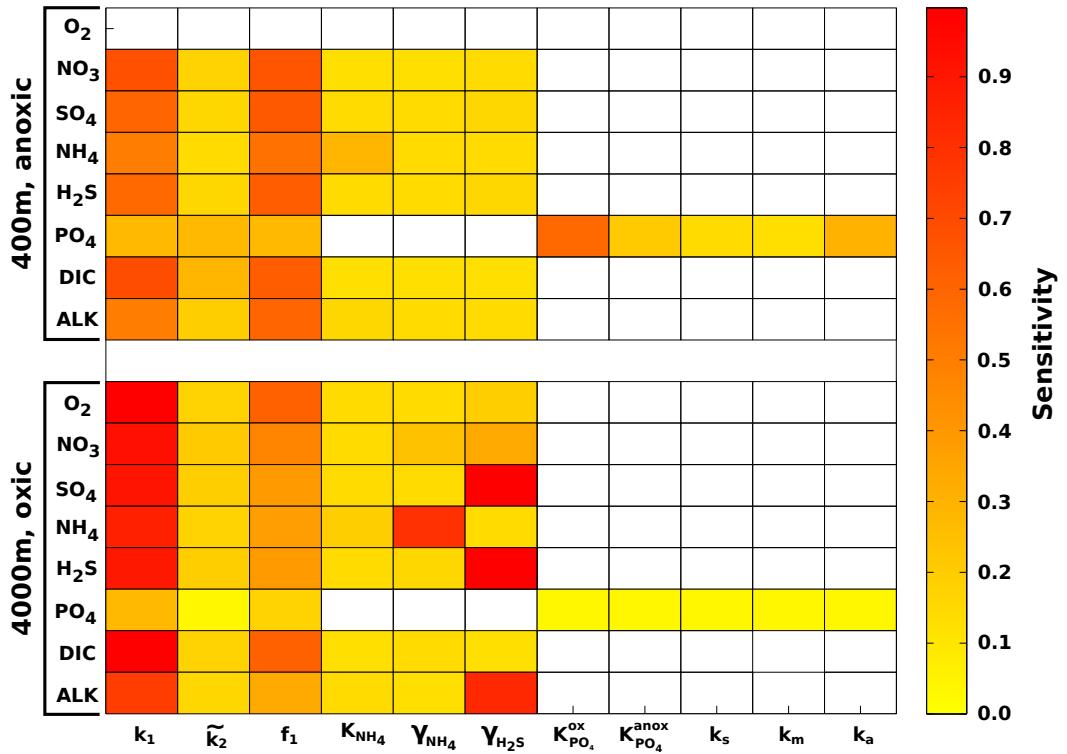


Figure 5. Pattern plot, showing the output sensitivity for each SWI flux (i.e. the chemical compounds on the vertical axis) and each input factor (i.e. the model parameters on the horizontal axis) for two idealised sediment cores. White patterns are assigned where the SWI flux is independent of the specific parameter.

850 3.2 Case study: Simulations of sediment cores

3.2.1 Methodology

In order to illustrate the capabilities of OMEN-SED, comprehensive datasets from the Santa Barbara Basin (Reimers et al., 1996), as well as from the Iberian margin and the Nazaré Canyon (Epping et al., 2002) are modelled. Modelled profiles are compared with measured pore water data from 855 different depths including the continental shelf (108 m) and the lower slope (2213 m) located at the Iberian margin, the upper slope (585 m) from the Santa Barbara Basin, and a deep sea site (4298 m) in the Nazaré Canyon. The Santa Barbara Basin is characterised by anoxic bottom waters, high POC concentrations and varved sediments (Reimers et al., 1990), therefore the depth of bioturbation in OMEN-SED is restricted to the upper 0.01 cm. In the uppermost sediments iron(III) hydroxides are 860 reduced, releasing Fe^{2+} which reacts with sulfide to form iron sulfides. Thus, the Fe cycle exerts a strong control on sulfide concentrations in the sediments of this basin (Reimers et al., 1996). In addition, the sediments are generally supersaturated with respect to carbonate fluorapatite by and be-

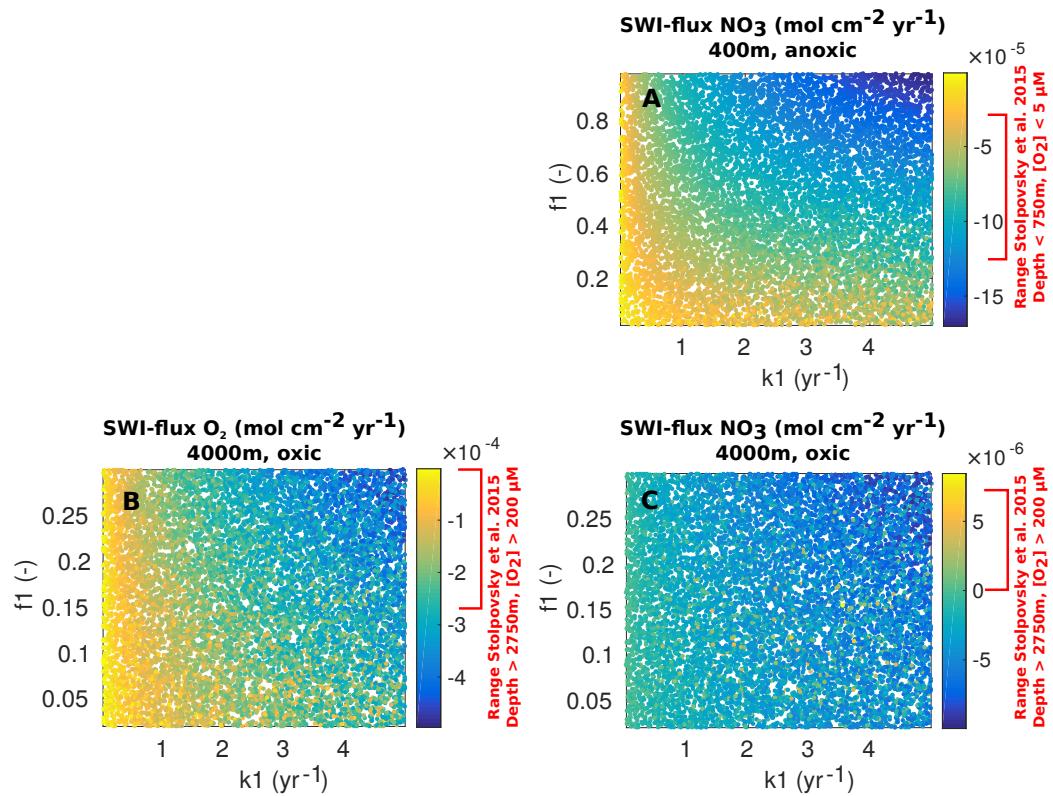


Figure 6. Scatter plots (k_1 vs f_1) of resulting OMEN-SED SWI-fluxes for the 400m anoxic (A: NO_3) and 4000m oxic (B: O_2 , C: NO_3) scenario. Negative values represent a flux from the water column into the sediments. Ranges indicated in red on the colour scale correspond to observed benthic fluxes as reported in the global database of Stolpovsky et al. (2015).

low 2 cm (Reimers et al., 1996). The Iberian margin, situated in the northeastern Atlantic, generally belongs to the more productive regions of the global ocean (Longhurst et al., 1995), however, seasonal changes in upwelling creates a strong temporal variability in primary productivity and organic carbon deposition. Submarine canyons in this area (like the Nazaré Canyon) may deliver organic carbon from the shelf to the ocean interior (van Weering et al., 2002; Epping et al., 2002). For a more detailed description of the study areas and the experimental work, the interested reader is referred to the publications by Reimers et al. (1996) and Epping et al. (2002).

In OMEN-SED sediment characteristics and boundary conditions are set to the observed values where available (Table 13). Other sediment characteristics (e.g. sedimentation rate, porosity, density), stoichiometric factors and secondary reaction parameters are set to the default value (see Tables 9 and 10). Organic matter is modelled as two fractions, with different first-order degradation rate constants. The POC and pore water profiles were manually fitted by optimizing the POC partitioning into the fast and slow degrading pool and their respective first-order degradation rate

Table 13. Model boundary conditions for the simulated sediment profiles in the Santa Barbara basin (108 and 2213 m) and Iberian margin (585 and 4298 m) reported in Figure 7. For all sites a DIC bottom water concentration of 2,400 nmols cm⁻³ is assumed.

Sediment characteristics:							
Depth (m)	Temp. (°C)	z_{bio} (cm)	D_{bio} (cm ² yr ⁻¹)	POC ₁ (wt%)	POC ₂ (wt%)	k ₁ (yr ⁻¹)	k ₂ (yr ⁻¹)
108	12.5	1.0	0.02	2.64	1.8	0.65	$1.0e^{-5}$
585	5.85	0.01	0.02	2.0	3.5	0.2	$8.0e^{-4}$
2213	3.2	10.0	0.17	0.45	0.5	0.1	$4.0e^{-4}$
4298	2.5	4.2	0.18	0.83	1.2	0.052	$1e^{-5}$

Bottom water concentrations of solutes (all in nmol cm ⁻³):								
Depth	O ₂	NO ₃	SO ₄	NH ₄	H ₂ S	PO ₄	PO ₄ ^a	Alkalinity
108	210	9.6	28,000	0.4	0.0	0.0	15.0	2,400
585	10	25.0	28,000	0.0	0.0	50.0	90.0	2,480
2213	250	25.0	28,000	0.6	0.0	0.0	5.0	2,400
4298	243	30.1	28,000	0.22	0.0	0.0	5.0	2,400

constants (priority is given to reproduce the POC and O₂ profiles). For phosphorus the equilibrium concentration for authigenic P formation (PO₄^a) was adjusted to fit the PO₄ concentration at z_{∞} .

3.2.2 Results

Fig. 7 compares modelled and observed sediment profiles for the Santa Barbara Basin and the Iberian margin. Results show that OMEN-SED is able to capture the main diagenetic features across a range of different environments without changing model parameters (other than the 3 we tuned) to site specific conditions. For the two open Iberian margin stations (108 and 2213 m) OMEN-SED fits all observations well. OMEN-SED does especially well at seafloor depth (SFD) 2213 m by reproducing the deep O₂ penetration and the subsurface maximum in NO₃ concentration due to the nitrification of NH₄ (note, that NH₄ is overestimated at this SFD). For the anoxic Santa Barbara Basin (585 m) the decrease in SO₄ and the increase in ALK concentration with sediment depth is well represented, indicating the importance of sulfate reduction as the primary pathway of OM degradation at this site (compare with Meysman et al., 2003). However, a misfit is observed for H₂S and PO₄ in the upper 20 cm of this sediment core. The discrepancy for H₂S can be explained by high iron(III) hydroxide concentrations, which is reduced to degrade organic matter (especially in the 2 – 4 cm depth interval), therefore placing the beginning of the sulfate reduction zone and the production of H₂S to the deeper sediments (Reimers et al., 1996). Iron processes are currently not dynamically represented in OMEN-SED. In addition, produced dissolved Fe reacts with H₂S to form iron sulfides (e.g. pyrite, FeS₂) and thus further inhibits the rise of H₂S (Reimers et al., 1990). The iron cycle also

895 plays a critical role for phosphorus, as the reduction of iron(III) hydroxides in the surface sediments
releases sorbed phosphate, leading to pore waters around and below 2 cm which are supersaturated
with respect to fluorapatite, thus initiating CFA precipitation. Reimers et al. (1996) could even show
that the accumulation of CFA is mainly restricted to the near-surface sediments (~ 5 cm) instead
of throughout the sediment column. As OMEN-SED does not include an iron-cycle, and Fe-bound
900 P and CFA processes are highly parameterised, the model is not able to capture these complex,
non-steady state phosphorus dynamics at this specific site. For the Nazaré Canyon station (4298
m) satisfactory fits could be realised apart from NH_4 . However, also Epping et al. (2002) could
not obtain a better fit using a more complex diagenetic model. They suggested non-local solute
exchange resulting from bioirrigation being responsible for the higher NH_4 concentrations at this site
905 which is neglected in their model, as well as in OMEN-SED. Furthermore, the fractured POC profile
(indicating episodic depositional events through the canyon) could have been approximated using a
different partitioning of the bulk POC into labile and refractory pool with different degradation rate
constants, thus potentially leading to a better fit of the NH_4 profile. In general, better approximations
of the data could have potentially been acquired by applying a sensitivity study using different NC-
910 ratios (e.g. Epping et al., 2002, report different ratios from Redfield stoichiometry) and exploring
the parameter space for the secondary reaction parameters (γ_{NH_4} , $\gamma_{\text{H}_2\text{S}}$). However, considering
these generalisations and our assumption of steady-state, which might not be valid, particularly for
the complex Santa Barbara basin, the shallow core and the Nazaré Canyon, which are affected by
seasonality and biology, OMEN-SED generally reproduces the observed pore water trends and hence
915 captures the main diagenetic processes.

3.3 Case study: Stand-alone simulations of global ocean transect

3.3.1 Methodology

In this section we explore to what degree OMEN-SED is capable of capturing the dynamics of or-
ganic matter degradation pathways and related TEA-fluxes as simulated with a more complete and
920 complex numerical diagenetic model. Therefore, we reproduce the simulations of typical conditions
along a global ocean hypsometry of Thullner et al. (2009) and compare our modelled TEA-fluxes
with the results of the complex model as well as with observations from Middelburg et al. (1996).
To explore the global degradation of OM in the seafloor Thullner et al. (2009) quantified various di-
agenetic processes using the Biogeochemical Reaction Network Simulator (BRNS, Aguilera et al.,
925 2005), a flexible simulation environment suitable for reactive transport simulations of complex bio-
geochemical problems (e.g. Jourabchi et al., 2005; Thullner et al., 2005). Thullner et al. (2009) used
seafloor depth (SFD) as the master variable and calculated model parameters, such as w , D_{bio} and ϕ ,
from existing empirical relationships (e.g. Van Cappellen and Wang, 1995; Middelburg et al., 1997).
Organic matter degradation was described with a 1G approach, thus assuming a single pool of or-

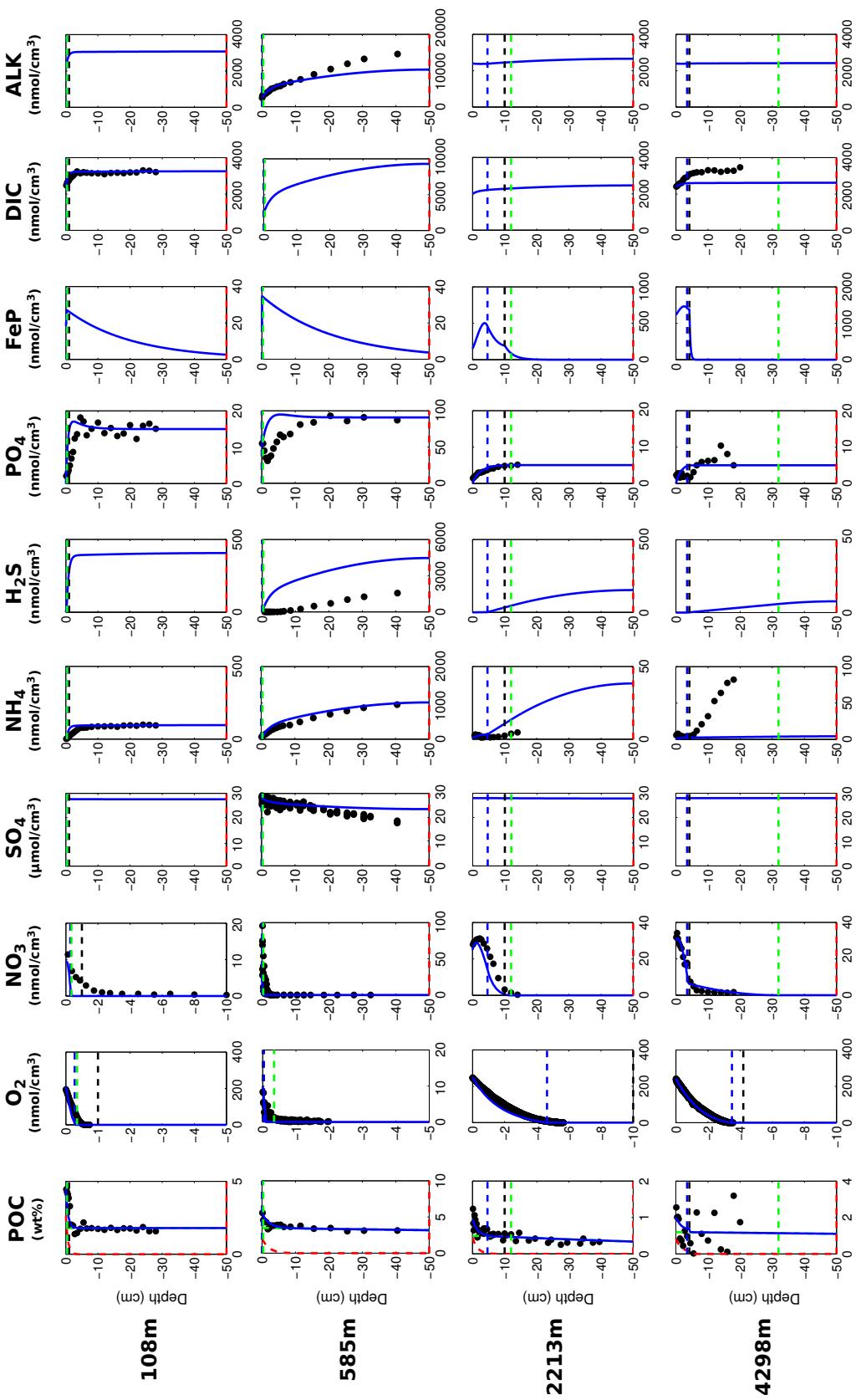


Figure 7. Modelled (curves) and measured (filled dots) solid phase and dissolved pore water profiles for four different sediment cores. Note that different scales are used for different stations. The blue POC curve represents the sum of the refractory (green) and labile (red) POC fraction. The horizontal dashed lines in each panel indicate the bioturbation depth (black) and the penetration depths of oxygen (blue), nitrate (green) and sulfate (red) as calculated by OMEN-SED.

930 ganic matter of uniform reactivity. The first order rate constant was related to the burial velocity, w (cm year $^{-1}$), following the empirical relationship of Boudreau (1997):

$$k = 0.38 \cdot w^{0.59}. \quad (48)$$

This rate constant can be assumed as the mean reactivity of the organic matter fractions which are degraded in the upper, bioturbated 10 – 20 cm of the sediments. Thus, more reactive fractions (de-
935 graded during days/weeks close to the SWI) and more refractory fractions (degraded on longer time scales deeper in the sediments) are not captured by this relationship (Boudreau, 1997). BRNS simulations were performed using boundary conditions and parameters for depths representative for shelf, slope and deep sea sediments (i.e. SFD of 100m, 200m, 500m, 1000m, 2000m, 3500m and 5000m). In order to reproduce these results, OMEN-SED is configured here as a 1G model and boundary
940 conditions and model parameters are defined as in Thullner et al. (2009, see Table 14). As OMEN-SED assumes a fixed fraction (i.e. γ_{NH_4} , γ_{H_2S}) of reduced substances to be reoxidised, which exerts a large impact on the resulting SWI-fluxes (compare Section 3.1), two sets of simulations are performed in order to show the range of possible model outputs. In the first setup 95% of the reduced substances are reoxidised (i.e. $\gamma_{NH_4} = \gamma_{H_2S} = 0.95$) and in the second only 5% are reoxidised (all
945 other model parameters and boundary conditions are equal).

3.3.2 Results

Figure 8 compares simulated SWI-fluxes of TEAs (i.e. O₂, NO₃ and SO₄) along the global hypsometry using OMEN-SED (black lines) with the results of Thullner et al. (2009) (red lines). Observations for O₂ and NO₃ fluxes are taken from Middelburg et al. (1996). Due to the applied empirical
950 relations organic matter flux to the seafloor decreases by 2 orders of magnitude from 100 to 5000 m and its degradation rate constant by 1 order of magnitude (Table 14). Therefore, the rate of organic matter degradation is about 50 times greater at 100 m than at 5000 m (compare Thullner et al., 2009), thus resulting in a decrease of TEA-fluxes along the hypsometry (Figure 8). The 95%-reoxidation experiments (dots) show proportionally higher O₂ in-fluxes than the the 5%-reoxidation experiments
955 (triangles) because more O₂ is utilised for in situ production of NO₃ and SO₄ in the sediments. This is also mirrored by the increased NO₃ out-flux and decreased SO₄ in-flux for shallower SFDs. This is in line with the results of Thullner et al. (2009) which showed that in situ production is an im-
960 portant pathway of SO₄ supply in the sediment, which is responsible for ~80% of the total OM degradation at depths between 100 and 2000 m (in our results SO₄ is not used for OM degrada-
tion in OMEN-SED below 2000m). In general, Figure 8 shows that OMEN-SED captures the main trends in observed TEA fluxes well and fluxes calculated with BRNS fall within the range of possible OMEN-SED results.

The observed O₂ fluxes in the upper 2000m are generally encompassed by our total range in predicted OMEN-SED fluxes. Oxygen fluxes for the deep-sea sediments, however, are slightly un-

Table 14. Seafloor depth dependency of key model parameters and boundary conditions (adapted from Thullner et al., 2009).

	Seafloor depth					5000 m
	100 m	200 m	500 m	1000 m	2000 m	
Model parameters						
w^a (cm yr ⁻¹)	3.98 × 10 ⁻¹	3.60 × 10 ⁻¹	2.67 × 10 ⁻¹	1.62 × 10 ⁻¹	5.94 × 10 ⁻²	1.32 × 10 ⁻²
D_{bio}^a (cm ² yr ⁻¹)	27.5	25.1	19.0	12.1	4.83	1.23
ϕ^b	0.85	0.85	0.80	0.80	0.80	0.80
T ^c (°C)	10.3	9.7	8.1	5.8	3.0	1.5
ρ_{sed}^c (g cm ⁻³)	2.5	2.5	2.5	2.5	2.5	2.5
k^d (yr ⁻¹)	0.221	0.208	0.174	0.130	0.0718	0.0296
Upper boundary conditions						
POC _{flux} ^a ($\mu\text{mol cm}^{-2} \text{yr}^{-1}$)	510	467	357	228	93.0	24.3
POC ^e (wt%)	0.79	0.78	0.55	0.50	0.42	0.32
O _{2,0} ^c (nmol cm ⁻³)	132	129	121	114	116	135
NO _{3,0} ^c (nmol cm ⁻³)	17.3	18.6	22.1	26.5	31.0	31.6
SO _{4,0} ^b (nmol cm ⁻³)	28,000	28,000	28,000	28,000	28,000	28,000

^a Derived from Middelburg et al. (1997).

^b Derived from Van Cappellen and Wang (1995).

^c Derived from Conkright et al. (2002).

^d Calculated with OMEN-SED from POC_{flux}.

^e Derived from Boudreau (1997).

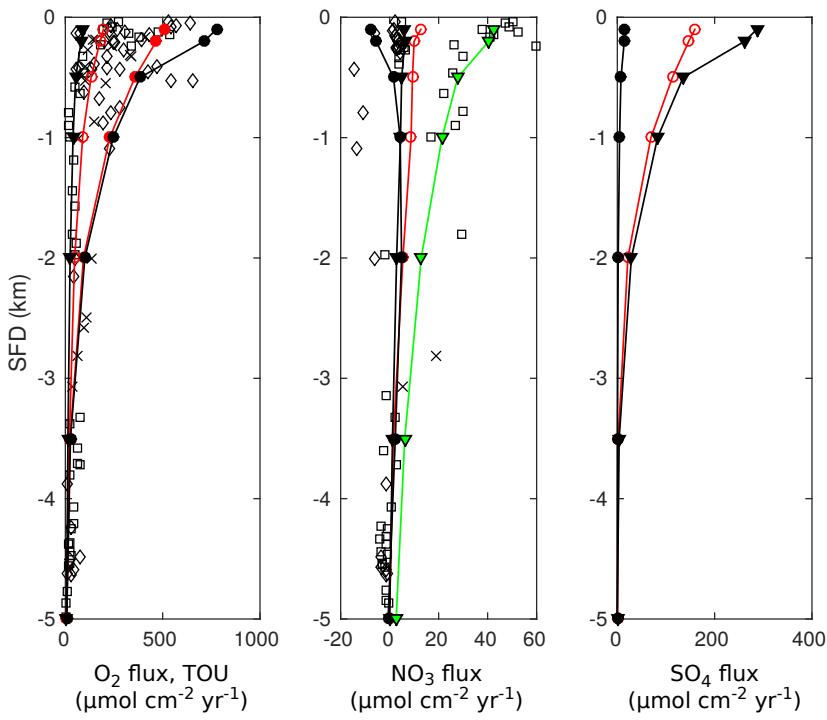


Figure 8. Fluxes of O₂, NO₃ and SO₄ to the sediment along the global hypsometry. Red lines (with open symbols) are modelled fluxes from Thullner et al. (2009) using BRNS; black lines are results from OMEN-SED (● : $\gamma_{NH_4} = \gamma_{H_2S} = 0.95$; ▼: $\gamma_{NH_4} = \gamma_{H_2S} = 0.05$). Observations of TEA fluxes are taken from Middelburg et al. (1996) (◊: Atlantic, □: Pacific, ×: Arctic/Indian Ocean). Also plotted in Figure (A) are the total oxygen uptake (TOU) estimates of Thullner et al. (2009) (filled red symbols). The green line indicates OMEN-SED results for low oxygen/high nitrate levels and the lower NC-ratio. Positive fluxes are directed from the ocean into the sediments.

965 derestimated. These deviations can presumably be related to the assumed 1G description of organic matter degradation, which neglects the more labile OM pool. This highly reactive pool is degraded close to the sediment surface, thus promoting higher aerobic degradation rates and higher O₂ fluxes. Nitrate fluxes in the upper 500m of the Atlantic Ocean are well predicted. However, as in Middelburg et al. (1996) the direction of calculated nitrate fluxes in the upper 1000m of the Pacific Ocean 970 differ from the observations. Middelburg et al. (1996) related these discrepancies to the globally averaged model parameters and the applied boundary conditions. They could reduce the disagreements significantly by using more representative bottom water concentrations for the eastern Pacific and a higher flux of labile organic matter for their 2G model. By changing the boundary conditions and the N:C elemental ratio of organic matter for the whole hypsometry, it is possible to obtain a better 975 model-data fit with OMEN-SED for the shallow Pacific Ocean (green line in Fig. 8B). Bohlen et al. (2012) report that the elemental N:C ratio strongly deviates from Redfield stoichiometry (0.151)

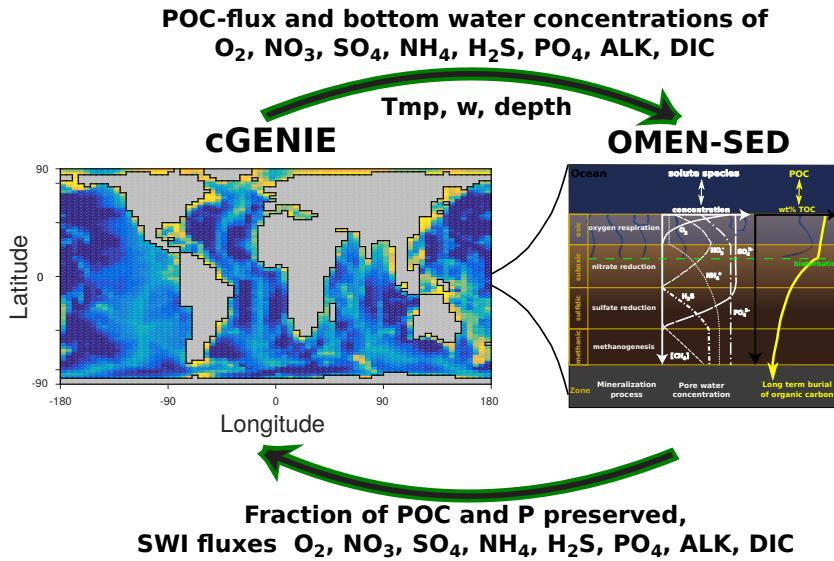


Figure 9. Schematic of the relationship between OMEN-SED and cGENIE. Arrows represent the information transferred between models.

with specifically lower values for the East Pacific Ocean. The use of their globally averaged value of 0.067 allows reconciling modelled and observed values provided that bottom water conditions are also changed to the low oxygen/high nitrate levels more likely to be found in the shallow Pacific
980 Ocean ($O_2 = 10 \text{ nmol cm}^{-3}$ and $NO_3 = 80 \text{ nmol cm}^{-3}$).

4 Coupled pre-industrial Earth system model simulations

4.1 Coupling to the cGENIE Earth system model

In a final step, we couple OMEN-SED to the carbon-centric version of the “GENIE” Earth system model (cGENIE, Ridgwell et al., 2007) in order to illustrate how a fully coupled ocean-sediment
985 system can be configured and applied. We start by providing a brief description of cGENIE and the coupling procedure (Fig. 9).

cGENIE is a model of Intermediate Complexity based on the efficient climate model “C-GOLDSTEIN” of Edwards and Marsh (2005), featuring a frictional-geostrophic 3D-ocean circulation model coupled to a fast Energy-Moisture Balance 2D-atmosphere together with a dynamic-thermodynamic
990 sea-ice component. The version of cGENIE used here includes the marine geochemical cycling of carbon, oxygen, phosphorus and sulfur (Ridgwell et al., 2007), preservation of carbonates in deep-sea sediments (SEDGEM, Ridgwell and Hargreaves, 2007) and terrestrial weathering (Colbourn et al., 2013). The ocean model is implemented on a 36×36 equal-area horizontal grid with 16 vertical levels using the pre-industrial continental configuration and bathymetry as in Archer et al.

995 (2009). A finer grid (72×72) is used for the sediments (see Fig. 11C and Ridgwell and Hargreaves, 2007) and OMEN-SED is called by SEDGEM for each wet ocean grid point. Depending on the configuration of the overlying biogeochemical ocean model, processes can be included or excluded in OMEN-SED and assumed stoichiometric factors need to be matched between models to ensure preservation of mass. As nitrogen is not modelled explicitly in the employed cGENIE configuration, 1000 NC_i , ALK^{NIT} and ALK^{DEN} in OMEN-SED are set to zero. cGENIE, however, implicitly includes the effects of NH_4 release and its complete nitrification on alkalinity but neglects the impact of P release. Therefore, alkalinity stoichiometries for aerobic degradation and sulfate reduction are changed to $\text{ALK}^{\text{OX}} = -16/106$ and $\text{ALK}^{\text{SUL}} = 122/106$, respectively (compare to default in Table 10).

Various biogeochemical tracers and parameters are transferred from SEDGEM to OMEN-SED 1005 (see Fig. 9) and are converted into the required units. Bottom water concentrations of solutes are converted from mol kg^{-1} to mol cm^{-3} and the depositional flux of POC (POC_{flux}) is converted from $\text{cm}^3 \text{cm}^{-2} \text{yr}^{-1}$ to $\text{mol cm}^{-2} \text{yr}^{-1}$ assuming an average density of POC of $1.0 \text{ cm}^3 \text{ g}^{-1}$. Within 1010 the water column in cGENIE, POC is partitioned into two fractions with different degradation length scales. The labile pool degrades while sinking through the water column, whereas the refractory pool is assumed relatively unreactive (Ridgwell et al., 2007). Thus, depending on seafloor depth, the partitioning of bulk POC reaching the sediments is different (Fig. 10A+B). This information is used by OMEN-SED to define the parameters f_1 and f_2 . Other parameters used from cGENIE are seafloor 1015 depth and local temperature. The advection/burial rate (w) is generally taken from cGENIE from the previous time-step, however, it is assured that w is not smaller than the detrital flux (Det_{flux}) to the sediments (e.g. $w < 0$ can occur if initially carbonate rich sediments are eroded during the spin-up of cGENIE). In case ($w \leq \text{Det}_{\text{flux}}$ & $\text{Det}_{\text{flux}} = 0.0$) all POC is remineralised at the ocean floor. Furthermore, a minimum value of $w = 0.4 \text{ cm kyr}^{-1}$ is imposed as OMEN-SED tends to be less 1020 stable for lower values. The bulk POC_{flux} is separated into the labile and refractory component and the routine to find the steady-state solution for POC is called. Here, the two POC depositional fluxes are first converted into SWI concentrations ($\text{POC}_i(z=0)$, in mol cm^{-3}) by solving the flux divergence equation:

$$\frac{\partial F}{\partial z} = -\frac{\partial}{\partial z} \left(-\xi D_i \frac{\partial \text{POC}_i}{\partial z} + \xi w \text{POC}_i \right) \quad (49)$$

for $z=0$. OMEN-SED then computes the fraction of POC preserved in the sediment (f_{POC} , see Eq. 1025 (5)) and subsequently calls the routines to find the steady-state solutions for the solute substances.

Note, that in this initial coupling the calculated benthic uptake/return fluxes F_{C_i} of dissolved species C_i (compare Eq. (6)) are adjusted for the advective loss at the lower sediment boundary ($w \cdot C_i(z_\infty)$) to assure the conservation of mass in the coupled model:

$$F_{C_i} = \phi(0) \left(D_i \frac{\partial C_i(z)}{\partial z} \Big|_{z=0} - w [C_i(0) - C_i(z_\infty)] \right). \quad (50)$$

In case OMEN-SED computes unrealistic results for POC preservation (i.e. $f_{\text{POC}} < 0.0$ or $f_{\text{POC}} > 1.0$) we discard the results of OMEN-SED and all POC is remineralised at the ocean floor. Finally,

f_{POC} and the SWI-fluxes of solutes (F_{Ci} , in $\text{mol cm}^{-2} \text{yr}^{-1}$) are returned to cGENIE. In case no POC is deposited on the seafloor (i.e. $POC_{flux} = 0$), OMEN-SED is not executed and f_{POC} and F_{Ci} for all i are set to zero. In order to reduce memory requirements, the sediment profiles (e.g. as shown in Fig. 7) are not calculated in the FORTRAN version of OMEN-SED, however, the boundary
1035 conditions are saved at the end of the experiment and sediment profiles for specific grid-cells, ocean basins and ocean transects can be plotted using the stand-alone MATLAB version of OMEN-SED.

4.2 Parameterising the OM degradation rate constants in a global model

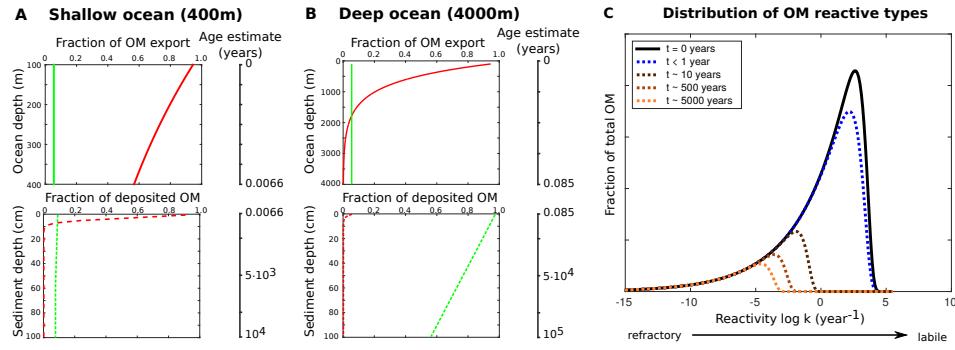


Figure 10. Idealised relationship of organic matter decomposition during remineralisation in the water column and the sediments. **A+B - Upper panels:** Water column development of the two organic matter fractions as represented in cGENIE for two ocean depths (red: labile OM with degradation length scale of 589m; green: refractory OM which is unreactive in the water column). The values are normalised to OM export at 100m. Age estimates for the OM since its export from the euphotic zone are calculated using a sinking velocity of 125m/day. **A+B - Lower panels:** Schematic representation of the development of the two OM fractions in the sediments (normalised to OM deposited on the seafloor). For the age estimates in the sediment column an advection rate of 0.01 and 0.001cm/yr is assumed, respectively. **C:** Idealised distribution functions of OM reactive types during remineralisation for different OM ages assuming a reactive continuum model for OM degradation. The initial distribution (at $t = 0$) represents fresh OM when it is exported from the euphotic zone (characterised by $a = 3e^{-4} \text{ yr}^{-1}$ and $\nu = 0.125$ Boudreau et al., 2008).

As shown in our sensitivity analysis (Section 3.1) and discussed by Arndt et al. (2013), the degradation rate constants for OM (k_i) are the most influential parameters and strongly determine the
1040 SWI-flux of redox-sensitive elements as well as the preservation of organic matter. Yet, their spatial variability is unknown at the global scale and reported rate constants in the sediments can vary by about 10 orders of magnitude or more (Middelburg et al., 1993; Arndt et al., 2013). Furthermore, when OMEN-SED is coupled to cGENIE, very different timescales have to be considered for OM degradation in the sediments compared to the water column (Fig. 10A+B), thus the diagenetic rate
1045 constants cannot be easily implied by the assumed water column POC flux profiles in cGENIE. Also,

microbes tend to degrade the more reactive organic matter compounds first (Emerson and Hedges, 1988; Wakeham et al., 1997; Lee et al., 2000), thus depending on the age of OM (or depth in the sediment and water column) the reactivity distribution of its compounds changes significantly (Fig. 10). For instance, in the water column, represented by the reactivity distribution $t < 1$ year in Fig. 10C, only the most reactive OM compounds are remineralised. This explains why the POC flux in the ocean can be represented with a 1G or pseudo 2G degradation model. In the sediments much longer timescales have to be considered, thus also more unreactive compounds are degraded and the reactivity distribution changes significantly already in the upper mm of the sediments ($t \sim 10$ years, Fig. 10C). Therefore, a broader range of OM reactive types must be represented by the degradation model to capture the reactivity spectrum of OM in surface sediments, explaining why at least a pseudo 3G model is required (including two degradable and one refractory fraction Soetaert et al., 1996; Boudreau, 1997; Stolpovsky et al., 2015). In addition, the advection rate in the sediments determines the age of OM at a specific sediment depth and thus its reactivity. For instance, assuming an advection rate of 0.01 cm/yr for the shallow ocean, OM at 5cm depth is about 500 years old, whereas it is one order of magnitude older for an advection rate of 0.001 cm/year in the deep ocean and thus degraded OM compounds cover a broader range of reactive types (Fig. 10C).

Thus defining appropriate OM degradation rate constants is a major challenge and source of uncertainty for diagenetic models. The rate constants in models are either determined through profile fitting for a specific site or, for global applications, they are related to a single, readily available characteristic (or master variable) of the local environmental conditions. For instance, considerable effort has been expended to relate the apparent rate constant for oxic and anoxic OM degradation to sedimentation rate (w) and various empirical relations have been proposed (Toth and Lerman, 1977; Tromp et al., 1995; Boudreau, 1997; Stolpovsky et al., 2015). Nevertheless, these relationships are generally based on limited data sets and their global applicability is questionable (Arndt et al., 2013). We hence test several alternative schemes in the coupled OMEN-cGENIE model. Our objective is not to perform and discuss a detailed calibration of the coupled models as this is beyond the scope of this sediment model development paper. Rather we want to showcase, that a coupling is possible and that the results show main sediment features one would hope to see on a global scale.

4.2.1 Methodology

In this section we compare modelled mean POC weight percentages (wt%) in the upper 5cm of the sediments ($\text{POC}_{5\text{cm}}$) to the global distribution pattern of POC content in surface sediments (< 5cm sediment depth) of Seiter et al. (2004) using different parameterisations for the degradation rate constants k_1 and k_2 . For our observational target we take the original POC distribution pattern in $1^\circ \times 1^\circ$ grid resolution (interpolated from > 5500 measurements, compare Seiter et al., 2004) and transform it onto the 72×72 SEDGEM grid (Figure 11). The regridding of the original POC distribution obviously affects the resolution of the data, especially for the continental margin, as

some sites with higher POC wt% are lost in the regridding process (compare e.g. maximum values for the East Pacific and upwelling waters of the Namibian shelf, Figure 11A + B). The colour of the points in Figures 12 - 14 indicates the seafloor depth (SFD) of the respective cGENIE grid-cell. As
1085 the individual data-points are highly scattered and in order to see if a certain relation between k_1 and k_2 performs better for specific ocean depths, the data-points are binned into 6 uniform depth-classes of 1000m each (respective mean POC wt% and SFD are represented by the triangles). The regression line (and the corresponding R^2 -value) is calculated for the 6 bin-classes and included in the figures.

**TODO Andy: Section describing the prescribed fields of solids (i.e. detrital, opal, CaCO_3) to the
1090 sediments!**

To parameterise the reactivity of organic matter in OMEN-SED two different schemes are tested and compared. First, spatially uniform degradation rate constants k_1 and k_2 are assumed. By simulating two different pools of POC in the water-column characterised by different degradation length scales (Ridgwell et al., 2007), cGENIE implicitly accounts for the decrease in mean POC reactivity
1095 with water-depth. The rate constants for the more refractory OM pool, k_2 , is systematically varied between 0.004 and 0.006 year⁻¹ and the more labile OM component, described by k_1 , is assumed to degrade a multiple times faster (i.e. $x \in \{1.1, 1.2, 1.3, 1.5, 2\}$). However, although accounting for the decrease in mean POC reactivity with seafloor depth, this approach does not take into account the change in distribution of organic matter reactivity types caused by different burial velocities and
1100 thus different residence time scales in the sediments (Fig. 10). Therefore, the second approach uses the empirical relationship proposed by Boudreau (1997), which relates the apparent OM degradation rate constant in the upper sediments to the burial velocity, w (cm year⁻¹, see also Section 3.3):

$$k_{\text{app}} = 0.38 \cdot w^{0.59}. \quad (51)$$

Following Boudreau (1997) and Stolpovsky et al. (2015) it can be assumed that k_{app} represents the
1105 mean OM reactivity within the upper 10-20cm of the sediments. The following assumptions are made in order to calculate the two degradation rate constants for OMEN-SED:

$$k_{\text{app}} = f_1 \cdot k_1 + f_2 \cdot k_2 \quad (52)$$

$$k_1 = x \cdot k_2 \quad (53)$$

1110 where x describes the relation between k_1 and k_2 and is subject to sensitivity experiments (with values of $x \in \{2, 5, 8, 10, 12, 15, 20, 25\}$). Note that the differences between k_1 and k_2 using this approach is significantly larger as in the globally uniform approach. As the fractions of labile and refractory OM reaching the sediments (f_1, f_2) is known from cGENIE, k_1 and k_2 can be calculated independently for each grid-cell.

1115 To simulate steady state sediment composition we configure the model as a “closed” system, i.e., one in which there is no loss of CaCO_3 through burial. The redox dependent P-cycle in OMEN-SED is not used in these experiments and all organic phosphorus is returned at the seafloor. To speed up

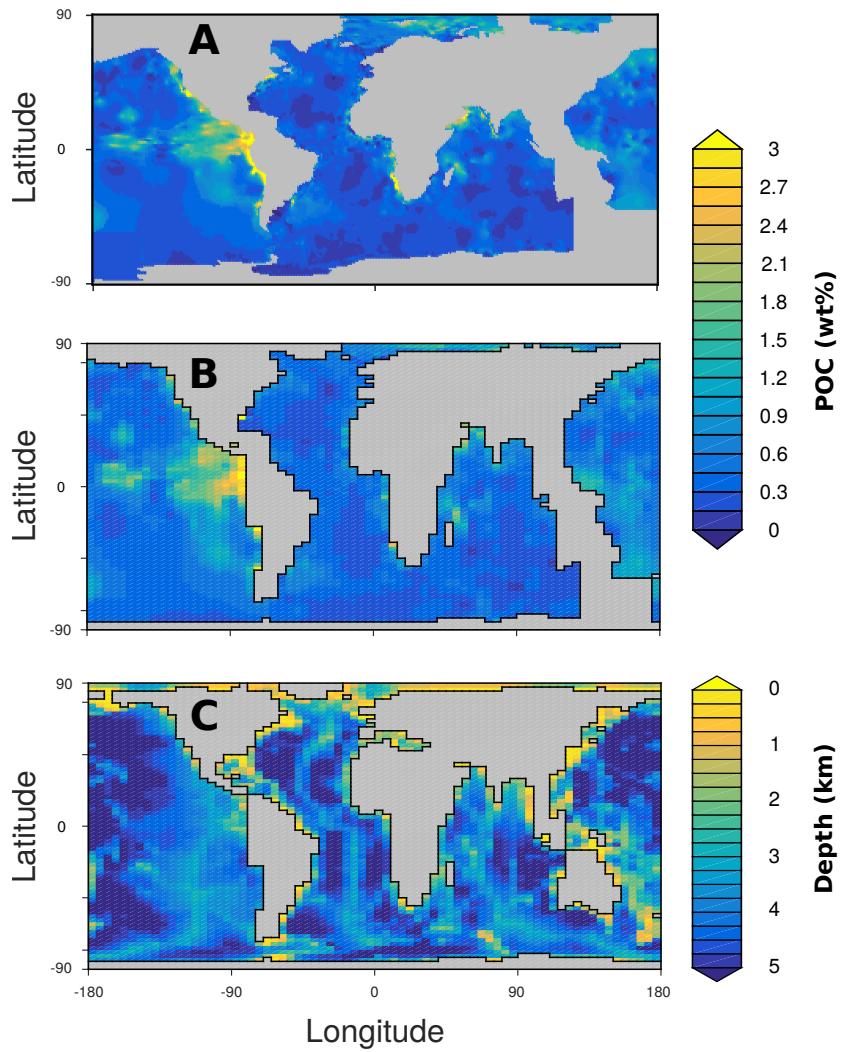


Figure 11. Observed distribution of sediment surface (< 5cm) POC wt% (A, B) and cGENIE bathymetry (C). (A) Original global distribution of POC wt% interpolated on a $1^\circ \times 1^\circ$ grid from more than 5500 individual data points (compare Seiter et al., 2004, for the interpolation procedure). (B) Observed POC wt% data transformed onto the 72 \times 72 SEDGEM grid. Grid points without any observations are left blank (grey). (C) Gridded continental configuration and ocean bathymetry of the 16-level, 72 \times 72 equal-area cGENIE grid.

the calculation and to assure that ocean redox changes caused by OMEN-SED do not impact the sediment composition of CaCO_3 , we use the prescribed solid fields as described earlier. Apart from
1120 the prescribed fields and the 72×72 sediment grid the model is configured as in Archer et al. (2009) and atmospheric CO_2 is restored to a pre-industrial value of 278 ppmv. First a 20,000 year spinup is performed without OMEN-SED being coupled. All presented coupled cGENIE-OMEN simulations are run for 10,000 years to steady state from this spinup. OMEN-SED is called for each grid-cell in every time step, feeding back the resulting SWI-fluxes and the fraction of POC preserved in the
1125 sediments to cGENIE.

4.2.2 Results

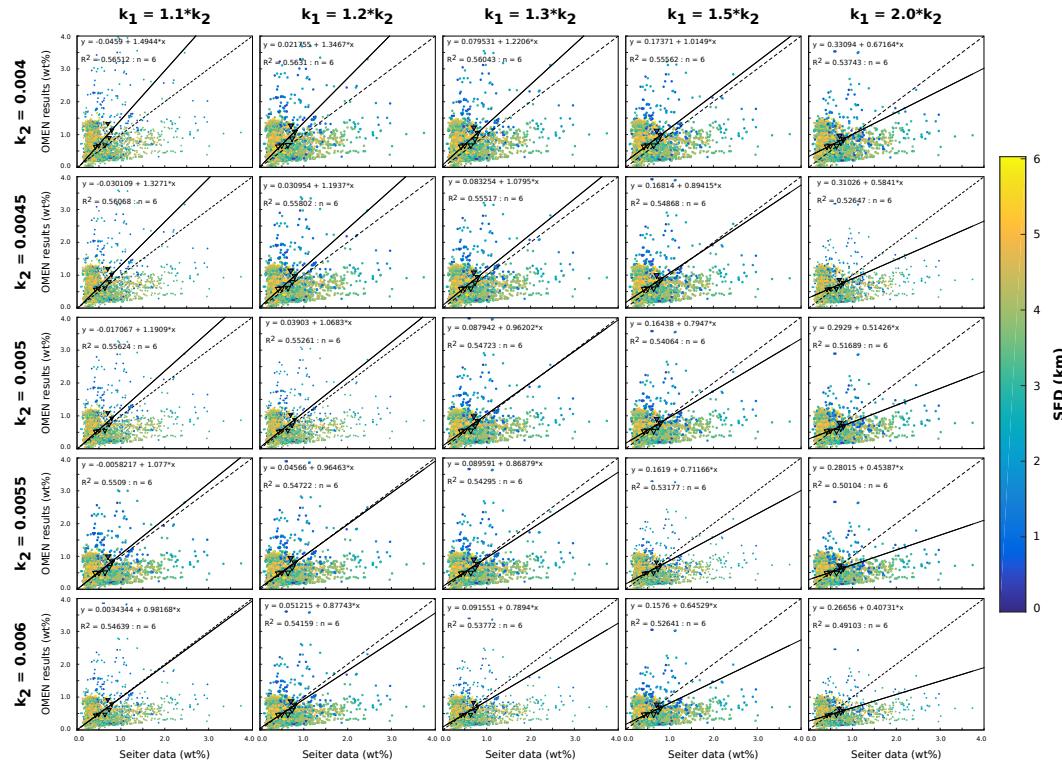


Figure 12. Crossplots comparing modelled and observed mean POC wt% in the upper 5 cm of the sediments using spatially uniform degradation rate constants k_1 and k_2 . Data-points are binned into 6 uniform depth-classes of 1000m as in Fig. 13, each class is represented by a triangle. Grid-points with more than 4.0 POC wt% are not shown.

Figure 12 presents results for the spatially uniform degradation rate experiments. In general, using spatially uniform degradation rate constants 5 of the 6 bin-classes are located closer to the 1:1 line as in the experiments using the Boudreau (1997) relation (Fig. 13). Also the slope of some regression
1130 lines is close to 1.0 (e.g. $(k_2, x) \in \{(0.004, 1.5), (0.0045, 1.3), (0.005, 1.2), (0.005, 1.3), (0.0055, 1.1), (0.0055, 1.2), (0.006, 1.1)\}$),

indicating that the simpler parameterisation adequately captures the relationship between depth and observed POC wt% by bin-class. The reason for this is that BIOGEM provides a depth dependent POC flux and partitioning between the two fractions (Fig. 10). The shallowest bin-class (between 0 and 1000m) represents an exception, as OMEN-SED tends to overestimate POC preservation for this depth class. However, this could also be related to the regridding of the original POC distribution pattern of (Seiter et al., 2004) on to the SEDGEM grid, as some data grid-cells with higher POC wt% on the continental margin are lost due to the restricted SEDGEM resolution (compare Section 4.2). Overall, using this parameterisation, a relationship where the labile POC fraction degrades not more than 1.5 times faster than the refractory fraction fits the Seiter et al. (2004) data better than a larger spread between both POC pools (i.e. $x > 2.0$).

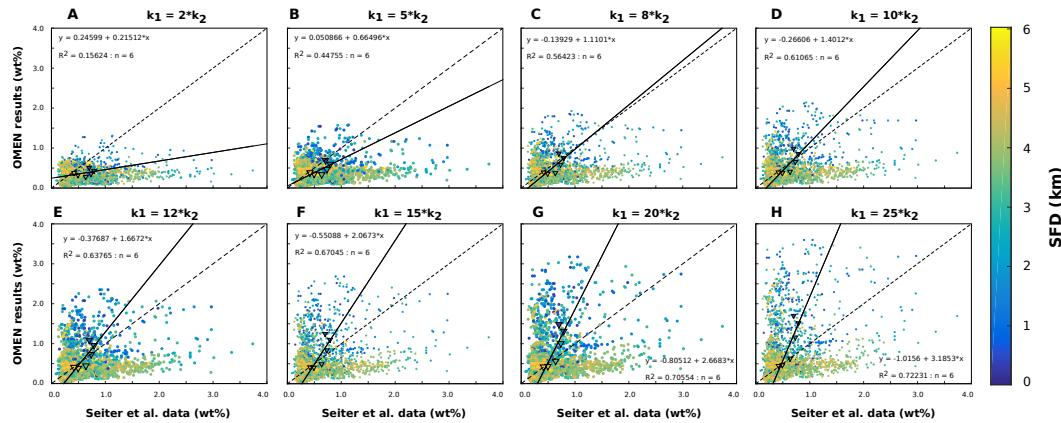


Figure 13. Crossplots comparing modelled and observed mean POC wt% in the upper 5 cm of the sediments using the relationship of Boudreau (1997) and the assumptions of Eq. (52) and (53) to calculate k_1 and k_2 . Data-points are binned into 6 uniform depth-classes of 1000m, each class is represented by a triangle. Grid-points with more than 4.0 POC wt% are not shown.

Next the relationship of Boudreau (1997) and the assumptions of Eq. (52) and (53) are used to calculate k_1 and k_2 . In Figure 13 (A-H) the relation between the two degradation rate constants (Eq. (53)) is changed globally, thus independent of the seafloor depth. The crossplots show that it is not possible to achieve a solution where all bin-classes fall onto, or close to, the 1:1 line. Also, the slope of the regression lines are generally much larger or smaller than 1.0 (with the exception of Figure 13C), indicating that the relationship between depth and observed POC wt% by bin-class is not adequately represented by the model. The R^2 values are strictly monotonically increasing for increasing x because a depth-dependency is artificially imposed for the modelled POC wt% through the relation between k_1 and k_2 . When looking at the individual bin-classes it can be seen that shallow ocean depths are better represented by smaller differences between k_1 and k_2 (e.g. $k_1 = 5 \cdot k_2$ for SFD < 1000m, Figure 13B), and the deep ocean by a larger spread (e.g. $k_1 = 25 \cdot k_2$ for SFD > 3000m, Figure 13H). These results reflect the preferential degradation of more reactive organic matter types

(Wakeham et al., 1997; Lee et al., 2000) and thus the change in the distribution functions of OM reactive types for different OM ages (Fig. 10C). In the shallow ocean bulk POC consists of fresher organic matter types on average and is therefore generally more reactive overall (i.e. higher k_{app} due to higher w in the model) as in the deep ocean. In addition, OM at 5cm sediment depth in the deep ocean is generally older as in the shallow ocean due to lower burial rates, therefore more reactive types are affected by degradation and a larger spread between k-values is needed to capture these dynamics (compare Fig. 10C).

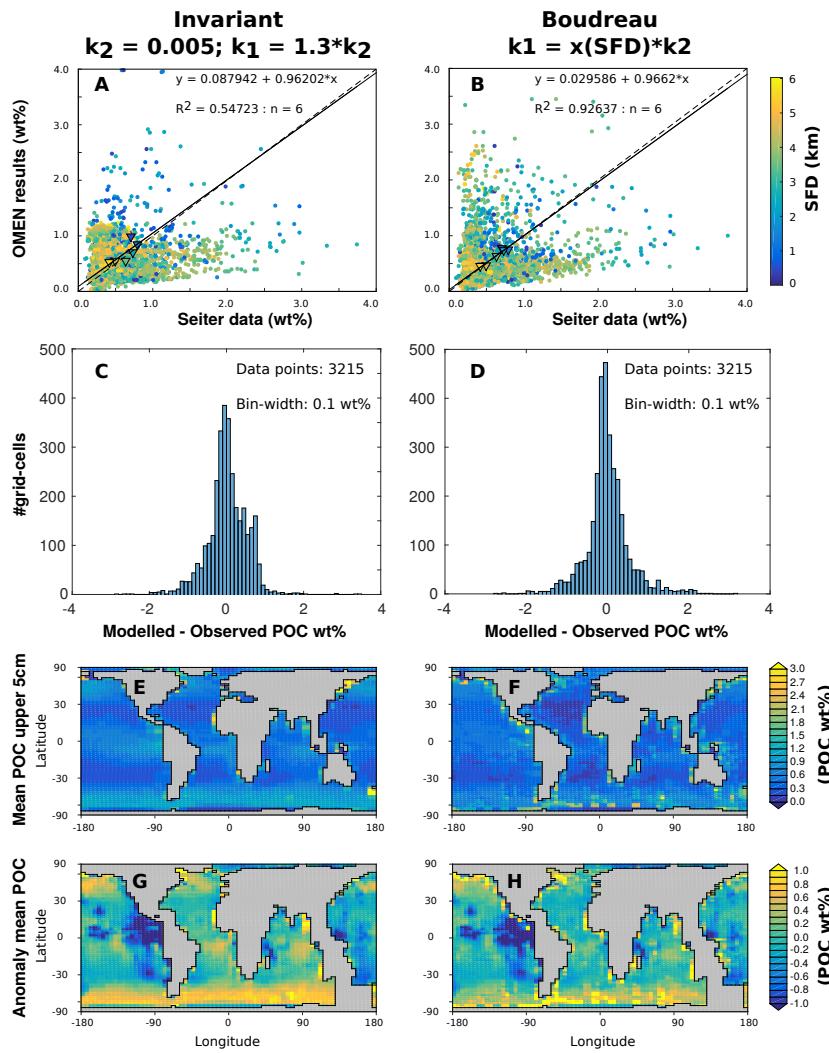


Figure 14. Mean POC concentrations in the upper 5cm of the sediments ($\text{POC}_{5\text{cm}}$) using the globally uniform model ($k_2 = 0.005$, $k_1 = 1.3 \cdot k_2$) and the depth dependent parameterisation $k_1 = x(\text{SFD}) \cdot k_2$ adapted from Boudreau (1997). A+B: Crossplots as shown before in Fig. 12 and 13. C+D: Histograms of the residuals of modelled minus observed $\text{POC}_{5\text{cm}}$. E+F: $\text{POC}_{5\text{cm}}$ as calculated with OMEN-SED. G+H: Difference map of $\text{POC}_{5\text{cm}}$ as calculated with OMEN-SED and interpolated data from Seiter et al. (2004).

1160 We use these observations to create a depth dependent relationship between the two degradation
rate constants, where x in Eq. (53) is a function of SFD and takes values of $x = 5$ for $SFD < 1000\text{m}$,
 $x = 8$ for $1000\text{m} \leq SFD < 2000\text{m}$, $x = 12$ for $2000\text{m} \leq SFD < 3000\text{m}$ and $x = 25$ for $SFD \geq 3000\text{m}$
for the 6 SFD bin-classes, respectively. In this depth dependent approach all bin-classes are close to
the 1:1 line and the resulting regression model accounts for 92.6% of the variance of the modeled
1165 POC wt% around the observed mean of the bin-classes (Figure 14A). Furthermore, the slope of the
regression line (0.9662) indicates that the relationship between depth and observed POC wt% for the
bin-classes is well represented by the model. The histograms (Fig. 14C+D) visualize the difference
between modelled and observed mean POC concentrations and demonstrate the high density of data
points close to the 1:1 line. For the depth dependent approach, 92.5% of the cGENIE grid-cells
1170 show a difference between modelled and observed POC concentration of less than 1.0 POC wt%; in
79.9% of the grid-cells, the difference is less than 0.5 POC wt% (for the globally uniform approach
the percentages are 95.37% and 70.95%, respectively).

Both experiments reproduce minimal POC concentrations in the subtropical gyres and generally
higher concentrations along the continental margins (Fig. 14E+F). Both experiments, however, un-
1175 derestimate mean POC wt% in the surface sediments of the equatorial east Pacific and overestimate
POC concentrations in the North Pacific and Southern Oceans (Fig. 14G+H). The depth dependent
approach of Boudreau (1997) shows more spatial variability in POC preservation than the other pa-
rameterisation. In general, implementing lower, anaerobic degradation rate constants when bottom
water oxygen concentrations fall below a threshold value could potentially improve the simulation of
1180 higher POC concentrations in areas with high POC input to the sediments (Palastanga et al., 2011).

4.3 Modelled fluxes and sediment characteristics

For the spatially uniform experiment A ($(k_2, x) = (0.005, 1.3)$) modelled SWI-fluxes and sediment
characteristics are shown in Figure 15. Modelled total POC degradation (POC_{degr}) rates in the upper
sediments decrease from the shelves to the deep sea by up to 2 orders of magnitude (Fig. 15B). This is
1185 in agreement with data from the literature (e.g. Middelburg et al., 1993, 1997) and other model results
(e.g. Thullner et al., 2009) which indicate that the highest degradation rates in marine sediments are
found in the coastal ocean ($SFD < 200\text{m}$). Oxygen fluxes into the sediments (Fig. 15C) range from
0.0 for the deep ocean and sites without OM deposition to values of about $300 \mu\text{mol cm}^{-2}\text{yr}^{-1}$
for the shallow ocean with the highest POC degradation rates. Influx of SO_4 into the sediments is
1190 rather low (between 0.0 and $23.9 \mu\text{mol cm}^{-2}\text{yr}^{-1}$) because in OMEN-SED 95% of produced H_2S is
reoxidised to SO_4 , therefore sulfate reduction is mainly driven by in situ sulfide oxidation. However,
in general the coupled model fluxes fall well within the ranges predicted by the stand-alone global
hypsometry experiments (O_2 between 0.0 and $800 \mu\text{mol cm}^{-2}\text{yr}^{-1}$ and SO_4 between 0.0 and about
300 $\mu\text{mol cm}^{-2}\text{yr}^{-1}$, compare Section 3.3). In accordance with the total POC degradation rates
1195 the release of PO_4 shows a maximum value of $8.12 \text{ nmol cm}^{-2}\text{yr}^{-1}$ on the shelves (Fig. 15D). The

$$\text{Boudreau } k_1 = x(\text{SFD}) \cdot k_2$$

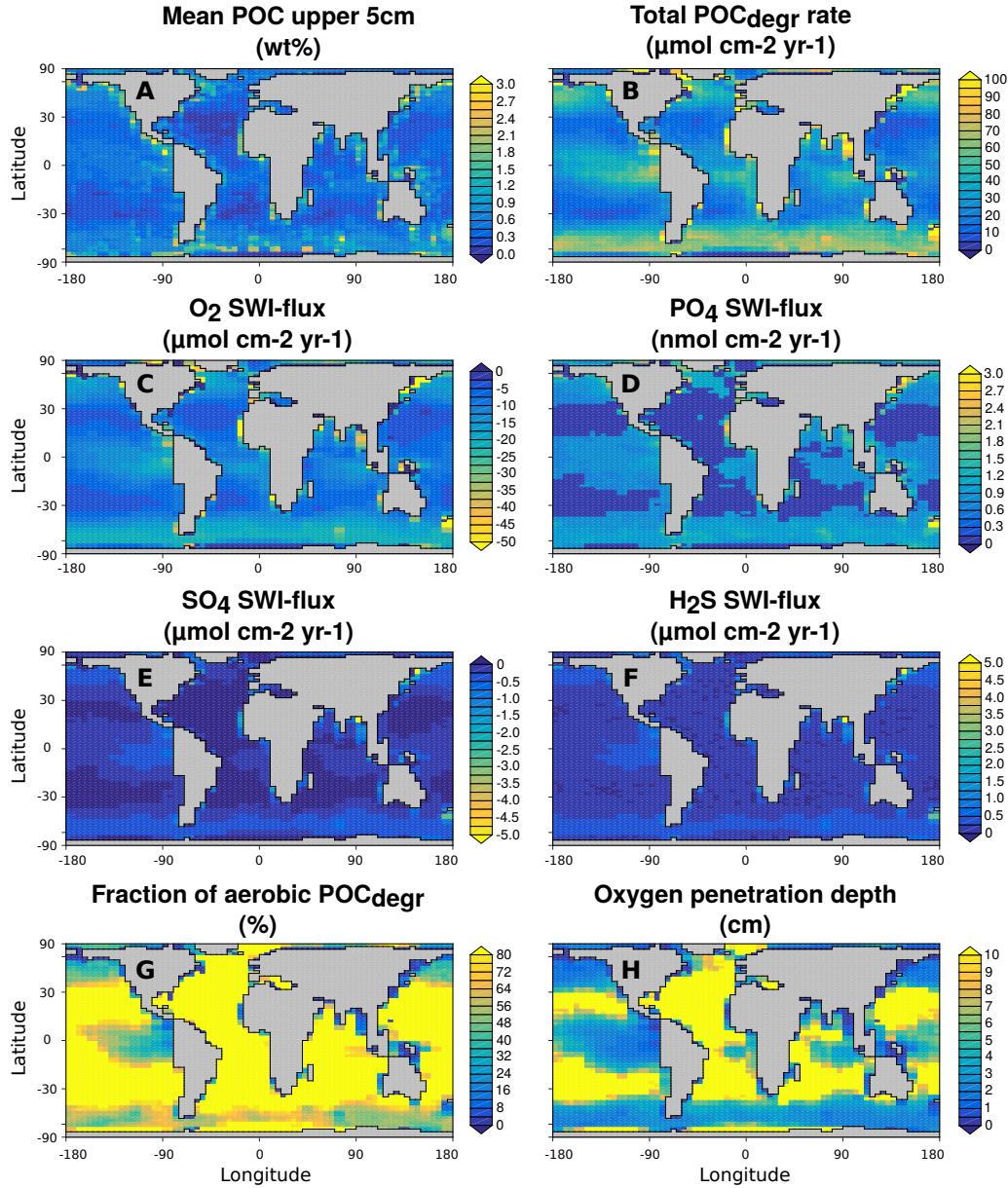


Figure 15. Sediment characteristics related to POC degradation and oxygen consumption for the depth dependent parameterisation after Boudreau (1997) with $k_1 = x(\text{SFD}) \cdot k_2$. Total POC_{degr} rate and fraction of aerobic POC_{degr} are the respective values for the first 5cm in the sediments.

relative contribution of aerobic POC degradation in the upper sediments increases from the shelves to the deep sea (Fig. 15G) which is also consistent with estimates from Thullner et al. (2009) who found that oxygen is responsible for less than 10% of POC_{degr} at 100m SFD and for more than 80% in the deep sea. The oxygen penetration depth in OMEN-SED increases from values below 1cm at the
1200 shelves to more than 10cm in the deep ocean (Fig. 15H and 16). Small oxygen penetration depths of a few millimetres are typical for bioturbated sediments in the coastal ocean (e.g. Wenzhöfer and Glud, 2002) and the oxygen penetration depth has been shown to increase rapidly with SFD to more than 10cm in the deep sea (Meile and Van Cappellen, 2003; Glud, 2008). Fischer et al. (2009) even found cores along a transect in the South Pacific gyre being oxygenated over their entire length (up to 8m)
1205 which is consistent with our model results (not shown). Simulated mean oxygen penetration depths for the 6 depth bin-classes also agree well with observations compiled by Meile and Van Cappellen (2003, Fig. 16).

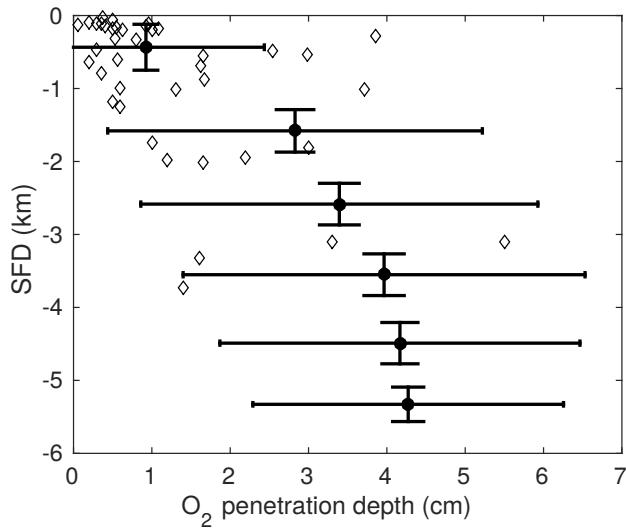


Figure 16. Seafloor depth versus O_2 penetration depth for the depth dependent paramaterisation after Boudreau (1997) with $k_1 = x(\text{SFD}) \cdot k_2$. Diamonds represent the observations compiled by Meile and Van Cappellen (2003). Circles are the mean model results for the 6 SFD bin-classes (with standard deviations). Grid-cells where the entire sediment column is oxic (i.e. $z_{\text{ox}} = 100\text{cm}$) are not considered in these statistics (17, 32, 102, 300, 477 and 307 cells for the 6 bin-classes, respectively).

5 Scope of applicability and model limitations

State-of-the art numerical models representing the full complexity of the diagenetic processes typically perform adequately at reproducing site-specific biogeochemical dynamics. However, tuning
1210 model parameters is laborious, the computational demand is high and, thus, their transferability to

the global scale is limited. On the other hand, analytical models are very efficient, but existing approaches coupled to global models generally use highly simplified reaction networks, often restricted to oxic degradation with a limited number of explicit pore water tracers. Our ability to assess the role
1215 of organic matter dynamics for global biogeochemical cycles and climate requires tools that resolve the most important biogeochemical processes and tracers explicitly, while at the same time are computationally efficient and have a degree of predictive capability to extrapolate knowledge to data poor areas.

The new model OMEN-SED presented here is a practical and legitimate compromise between
1220 complexity of biogeochemical processes and computational efficiency. Its scope of applicability covers the entire range from regional to global scales. OMEN-SED's computational efficiency facilitates its use in two very different ways. Firstly, it can be coupled to global Earth system models and therefore allows the investigation of coupled global biogeochemical dynamics over different timescales. Secondly, it can be used to calculate quantitative sensitivity indices requiring large sample sizes such
1225 as variance- or density-based approaches. Therefore, OMEN-SED can help to quantitatively investigate how systematic variations in model parameters impact the model output, for instance when the model has been tuned to a site-specific problem. Due to the represented anaerobic processes and secondary-redox reactions, OMEN-SED is also useful to investigate the role of benthic-pelagic coupling on the development of ocean anoxia and euxinia for instance during extreme climate events
1230 such as OAEs. On more regional scales it can be coupled to local models of systems like continental margins or estuaries which are characterised by complex interactions between different pathways of organic matter degradation and redox reactions. Here, the model can help to disentangle the complex process interplay that drives the biogeochemical dynamics and give quantifications for upper and lower constraints of carbon and nutrient budgets for these dynamic systems. In addition, OMEN-
1235 SED can be used to model eutrophication processes in shallow coastal waters as sediment-water oxygen and nutrient exchange fluxes are explicitly modelled and depend on reoxidation of reduced substances which causes a substantial part of oxygen consumption in these environments.

However, the model presented here, despite being more complete than previous analytical models, is still associated with a certain degree of simplifications. In order to solve the diagenetic equation
1240 analytically important assumptions have been made, which limit the general applicability of the model. One of the most important simplifications is assuming steady-state. When coupled to an Earth system model this assumption is only valid if the relevant variability in boundary conditions and fluxes is generally longer than the characteristic timescales of the reaction-transport processes. In that case the sediment column can be described by a series of pseudo steady-states as it is done in
1245 OMEN-SED. Consequently, the model can be used for investigating the long term effects of changes in boundary conditions such as input of OM or bottom water oxidation state on degradation and burial dynamics, for instance during OAEs. **Also mention the secondary redox parameters γ here. They are influential for the model output and hence should be formally tuned to a database of SWI-**

1250 fluxes and/or be parameterised depending on a master variable like seafloor depth or oxygenation state.

Yet, OMEN-SED is not able to predict the system response to short-term or seasonal variations of boundary conditions. Furthermore, the separation of the sediment column into distinct biogeochemical zones and the resulting lack of overlap in degradation pathways may cause distorted organic matter degradation rates for the different TEAs. In more complex models an overlap is allowed and
1255 the inhibition of a particular pathway by energetically more favourable oxidants is represented by an inhibition function with a half-saturation constant (see e.g. Boudreau, 1996; Meysman et al., 2003; Aguilera et al., 2005). Manganese and iron are not represented currently and as such OMEN-SED is not able to model all processes important in coastal marine environments and highly accumulating upwelling regions. Furthermore, it limits the scope of applicability of the model, especially
1260 in systems characterized by high Fe oxide inputs and anoxic conditions, where redox driven iron dissolution-precipitation reactions induce a high internal recycling of the sedimentary Fe pool. This can cause problems when modelling H₂S and PO₄ profiles in anoxic environments as their concentrations are affected by these metal ions (compare Section 3.2). The model also simulates DIC and alkalinity production and, thus, has the potential to predict pH profiles within the sediment. A major
1265 limitation at this stage is the lack of representation of CaCO₃ precipitation/dissolution coupled to OM decomposition, which also controls the inorganic carbon system (Krumins et al., 2013). In addition the depth invariant porosity limits the correct calculation of the sediment-water interface flux of dissolved species as in reality porosity decreases with sediment depth.

6 Conclusions

1270 In this paper we have described and tested a new, analytical early diagenetic model resolving organic matter cycling and associated biogeochemical dynamics called OMEN-SED. Our new model is the first analytical model to explicitly represent oxic degradation, denitrification, sulfate reduction and implicitly methanogenesis, as well as the reoxidation of reduced substances produced during organic matter degradation. Pore water tracers include O₂, NO₃, NH₄, SO₄, H₂S, DIC and ALK and the
1275 solid phase includes two degradable fractions of organic matter, Fe-bound P and authigenic Ca-P minerals. We have shown that the new analytical model is able to reproduce observed pore water profiles from different ocean depths when organic matter partitioning and degradation rate constants are tuned to site specific conditions. An extensive sensitive analysis, based on the novel density-based PAWN method (Pianosi and Wagener, 2015), has been performed to asses the importance of
1280 11 internal model parameters for all resulting SWI-fluxes. The results reveal that the degradation rate constant for labile organic matter is the most influential parameter for all model outputs. Under anoxic conditions secondary redox parameters exert an important control on related SWI-fluxes of SO₄, H₂S, NH₄ and alkalinity. In addition, the sensitivity analysis showed that produced model re-

sults fall well into the range of globally observed benthic O₂ and NO₃ fluxes. OMEN-SED is also
1285 used to quantify terminal electron acceptor fluxes across the sediment-water interface associated with
organic matter degradation along a global ocean hysometry. The results demonstrate that OMEN-
SED is capable of capturing most of the dynamics simulated with a complex, numerical diagenetic
model and our model results fall well within the range of observed fluxes. Furthermore, the coupling
1290 of OMEN-SED to the Earth system model cGENIE is described and spatially uniform degradation
rate constants as well as the empirical relationship of Boudreau (1997) for the apparent first order
degradation rate constant are tested to fit modelled to observed global organic matter concentrations.
Generally, large scale patterns of modelled surface sediment organic matter concentrations are in
agreement with observations of Seiter et al. (2004) and calculated SWI-fluxes and sediment char-
acteristics are consistent with estimates from the literature. However, results also show that **smaller**
1295 **scale OM degradation dynamics** in the sediments are too complex in time and space to be adequately
represented using spatially uniform or depth-independent OM rate constant parameterisations. More
work is needed to develop and test mechanistic parameterisations relating degradation rate constants
to available environmental parameters (e.g. bottom water oxygenation, burial rate, seafloor depth) in
order to model the heterogeneous reactivity distribution of OM types due to preferential degradation
1300 and different time scales in the sediments. Due to its computational efficiency the coupled model
presented here can be used to explore these questions using large simulation ensembles and objec-
tive statistical methods for sensitivity analysis. Furthermore, as the major parts of the global carbon
cycle are included in the new model it is well suited to examine the role of sediments for global
biogeochemical cycles and climate and for exploring feedbacks within the Earth system in response
1305 to a wide range of perturbations and over various timescales.

DH: Not sure, how to
express this in a better
way!?

7 Code Availability

The OMEN-SED source code (Fortran and Matlab) related to this article is provided as a supple-
mentary package together with a ReadMe file, where hardware and software requirements, source
code files and model output file management are fully described.

1310 Appendix A: Reaction Network

Appendix B: Sensitivity Analysis

B1

Acknowledgements. We thank Claire Reimers and Filip Meysman for supplying the dataset from the Santa Bar-
bara Basin, as well as Martin Thullner and Jack Middelburg for making the BRNS results and observations in-
1315 cluded in Section 3.3 available. We are also grateful to Andy Dale for providing the global flux database used in

Table 15. Primary pathways of organic matter degradation, secondary redox reactions and stoichiometries implemented in the reaction network.

Pathway	Stoichiometry
Primary Redox reactions	
Aerobic degradation	$(\text{CH}_2\text{O})_x(\text{NH}_3)_y(\text{H}_3\text{PO}_4)_z + (\text{x}+2\text{y})\text{O}_2 + (\text{y}+2\text{z})\text{HCO}_3^- \rightarrow (\text{x}+\text{y}+2\text{z})\text{CO}_2 + \text{yNO}_3^- + \text{zHPO}_4^{2-} + (\text{x}+2\text{y}+2\text{z})\text{H}_2\text{O}$
Denitrification	$(\text{CH}_2\text{O})_x(\text{NH}_3)_y(\text{H}_3\text{PO}_4)_z + \frac{(\text{4x}+3\text{y})}{5}\text{NO}_3^- \rightarrow \frac{(2\text{x}+\text{4y})}{5}\text{N}_2 + \frac{(\text{x}-3\text{y}-10\text{z})}{5}\text{CO}_2 + \frac{(\text{4x}-3\text{y}-10\text{z})}{5}\text{HCO}_3^- + \text{zHPO}_4^{2-} + \frac{(\text{3x}+3\text{y}+10\text{z})}{5}\text{H}_2\text{O}$
Sulfate reduction	$(\text{CH}_2\text{O})_x(\text{NH}_3)_y(\text{H}_3\text{PO}_4)_z + \frac{\text{x}}{2}\text{SO}_4^{2-} + (\text{y}-2\text{z})\text{CO}_2 + (\text{y}-2\text{z})\text{H}_2\text{O} \rightarrow \frac{\text{x}}{2}\text{H}_2\text{S} + (\text{x}+\text{y}-2\text{z})\text{HCO}_3^- + \text{yNH}_4^+ + \text{zHPO}_4^{2-}$
Methanogenesis	$(\text{CH}_2\text{O})_x(\text{NH}_3)_y(\text{H}_3\text{PO}_4)_z + (\text{y}-2\text{z})\text{H}_2\text{O} \rightarrow \frac{\text{x}}{2}\text{CH}_4 + \frac{\text{x}-2\text{y}+4\text{z}}{2}\text{CO}_2 + (\text{x}-2\text{z})\text{HCO}_3^- + \text{yNH}_4^+ + \text{zHPO}_4^{2-}$
Secondary Redox reactions	
Nitrification	$\text{NH}_4^+ + 2\text{O}_2 + 2\text{HCO}_3^- \rightarrow \text{NO}_3^- + 2\text{CO}_2 + 3\text{H}_2\text{O}$
Sulfide oxidation	$\text{H}_2\text{S} + 2\text{O}_2 + 2\text{HCO}_3^- \rightarrow \text{SO}_4^{2-} + 2\text{CO}_2 + 2\text{H}_2\text{O}$
AOM	$\text{CH}_4 + \text{CO}_2 + \text{SO}_4^{2-} \rightarrow 2\text{HCO}_3^- + \text{H}_2\text{S}$
Adsorption reactions and mineral precipitation	
NH ₄ adsorption	$\text{NH}_4^+ \xrightarrow{K_{\text{NH}_4}} \text{NH}_4^+(\text{ads})$
P ad-/desorption	$\text{PO}_4^{2-} \xrightarrow{K_{\text{P}\text{O}_4}^{\text{I},\text{II}}} \text{PO}_4^{2-}(\text{ads}); \quad \text{PO}_4^{2-} \xrightarrow{k_s} \text{Fe-bound P} \xrightarrow{k_m} \text{PO}_4^{2-}$
CFA precipitation	$\text{PO}_4^{2-} \xrightarrow{k_a} \text{CFA}$

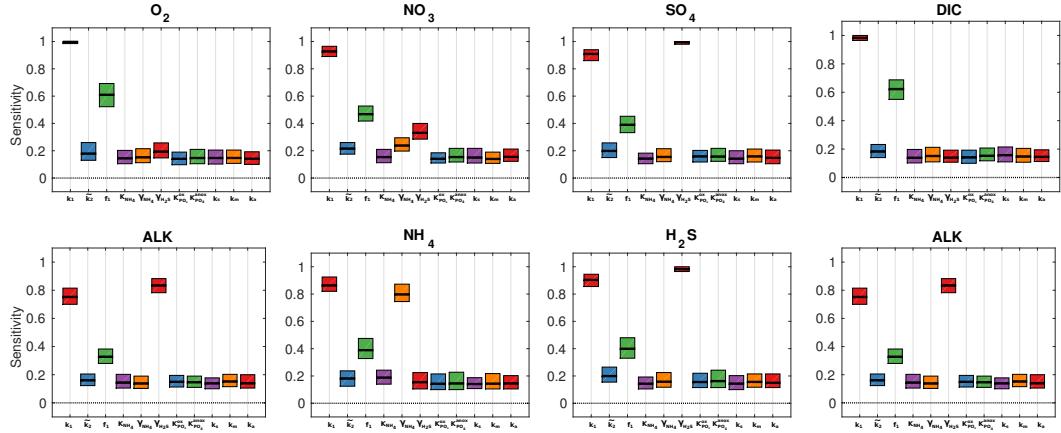


Figure 17. Move to Appendix Box plot of parameter sensitivities for the calculated SWI-fluxes for the 4000m oxic condition. Average sensitivities (black lines) and 90% confidence intervals using $N = 11200$ model evaluations and $Nboot = 100$ bootstrap resamples.

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1320 acknowledge funding from the EU grant ERC-2013-CoG-617313. **TODO: add for Stuart and Pierre!**

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