

# OMEN-SED 1.0: A new, numerically efficient sediment module for the coupling to Earth System Models

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**Abstract.** Here we describe the first version of the Organic Matter ENabled SEDiment model (OMEN-SED 1.0).

## 1 Introduction

**DH:** How to include comments.

### Role of marine sediments for climate and global biogeochemical cycles:

- 5 Marine surface sediments are key components in the Earth system. They host the largest carbon reservoir within the surficial Earth system, provide the only long term sink for atmospheric CO<sub>2</sub>, recycle nutrients and represent the most important geochemical archive used for deciphering past changes in biogeochemical cycles and climate. Physical and chemical processes in sediments (i.e. diagenetic processes) depend on the water column and vice versa: Diagenesis is mainly donor controlled, as it is fuelled by the external supply of solid material (e.g. organic matter, calcium carbonate, opal) from the water column and is affected by overlying bottom water concentrations of solutes. At the same time, diagenesis in the sediments transforms the deposited material and returns some of the resulting products (e.g. nutrients, DIC) to the water column. This so-called benthic-pelagic coupling is essential for understanding global biogeochemical cycles and climate (e.g. Archer and
- 10 Maier-Reimer, 1994; Soetaert et al., 2000; Archer et al., 2000; Mackenzie, 2005).
- 15

Biological primary production of organic matter (OM, CH<sub>2</sub>O in equation R1) and the reverse process of degradation can be written in a greatly simplified reaction as:



- On geological timescales production of OM is generally greater than degradation which results in
- 20 some organic matter being buried in marine sediments and oxygen accumulating in the atmosphere.

Thus burial of OM leads to net oxygen input to, and CO<sub>2</sub> removal from the atmosphere (Berner, 2004). On shorter timescales, the upper few meters of the sediments (i.e. early diagenesis) are specifically important as it is decided here if a substance is recycled to the water column or buried for a longer period of time in the deeper sediments (Hensen et al., 2006). Most biogeochemical cycles  
25 and reactions in this part of marine sediments can be related either directly or indirectly to the degradation of organic matter (e.g Boudreau and Ruddick, 1991; Arndt et al., 2013). Organic matter degradation releases metabolic CO<sub>2</sub> to the pore water, causing it to have a lower pH and provoking the dissolution of CaCO<sub>3</sub> (Emerson and Bender, 1981). Oxygen and nitrate for instance, the most powerful electron acceptors, are consumed in the course of the degradation of organic matter, resulting  
30 in the release of ammonium and phosphate to the pore water. As such, changes in the degradation rate of OM can profoundly affect the oxygen and nutrient inventory of the ocean and thus primary productivity (Van Cappellen and Ingall, 1994; Lenton and Watson, 2000).

**TODO:** Something on: Role of marine sediments for another time period , e.g. glacial/interglacial atmospheric pCO<sub>2</sub> cycles

35 Nutrient recycling from the sediments has been invoked by modelling and data studies to explain the occurrence of extreme events in Earth history, for instance Oceanic Anoxic Events (OAEs) (e.g. Mort et al., 2007; Tsandev and Slomp, 2009). OAEs represent severe disturbances of the global carbon, oxygen and nutrient cycles of the ocean and are usually characterized by widespread bottom water anoxia and photic zone euxinia (Jenkyns, 2010). One way to explain the genesis and retention  
40 of OAEs is increased oxygen demand due to enhanced primary productivity (PP). Increased nutrient inputs to fuel PP may have come from marine sediments as the burial efficiency of phosphorus declines when bottom waters become anoxic (Ingall and Jahnke, 1994; Van Cappellen and Ingall, 1994). The recovery from OAE like conditions is thought to involve the permanent removal of excess CO<sub>2</sub> from the atmosphere and ocean by burying carbon in the form of organic matter in marine  
45 sediments (e.g. Arthur et al., 1988; Jarvis et al., 2011), which is consistent with the geological record of widespread black shale formation (Stein et al., 1986). However, the overall amount, exact timing and the rate of organic matter burial remain a topic of an ongoing debate.

### **Diagenetic Models:**

Therefore, globally quantifying the burial and degradation of organic matter in marine sediments and  
50 related biogeochemical dynamics is important for understanding climate and the cycling of many chemical elements on various timescales. Such studies and quantifications are possible through the application of idealised mathematical representations of diagenesis, or so-called diagenetic models (e.g. Berner, 1980; Boudreau, 1997). The number of research questions that can be addressed with diagenetic models is infinite and a plethora of different approaches have been developed, mainly  
55 following two distinct directions (Arndt et al., 2013). First, state-of-the art vertically resolved diagenetic models simulating all of the essential coupled redox and equilibrium reactions within marine sediments that control carbon burial and benthic recycling fluxes (e.g. BRNS, Aguilera et al., 2005;

MEDIA, Meysman et al., 2003; OMEXDIA, Soetaert et al., 1996b). These “complete”, non-steady-state models generally use a so-called multi-G approach, thus dividing the bulk organic carbon pool  
60 into a number of compound classes that are characterised by different degradabilities  $k_i$ . Combining such a complex diagenetic model with an ocean biogeochemical model results in the most realistic benthic-pelagic coupling. However, the global applicability of this set-up is limited by the high computation cost of simulating all biogeochemical reactions in the sediments. The second group of models is less sophisticated and comprehensive than the “complete”, non-steady-state diagenetic  
65 models and is used for the coupling to global Earth System Models (e.g. DCESS, Shaffer et al., 2008; MEDUSA, Munhoven, 2007 or the model of Heinze et al. (1999)). In particular, these analytic models consider fewer biogeochemical reactions and assume that the sedimentary organic matter pool is composed of just a single or two discrete compound classes (1G or 2G approach respectively).

## 70 How are sediments resolved in Earth System models:

Even though there is potential to use more appropriate global sediment representations, in most current Earth System models sediment-water exchange of OM and chemical elements is either neglected or treated in a very simplistic way (Soetaert et al., 2000; Hülse et al., 2016). Most Earth system Models of Intermediate Complexity (EMICs) and also some of the higher resolution global carbon  
75 cycle models represent the sediment-water interface either as a reflective or a conservative/semi-reflective boundary (Hülse et al., 2016). Thus, all particulate material deposited on the seafloor is either instantaneously consumed (reflective boundary), or a fixed fraction is buried in the sediments (conservative/semi-reflective boundary). Both highly simplified approaches furthermore completely neglect the exchange of solute species through the sediment-water interface and, therefore, can-  
80 not resolve the complex benthic-pelagic coupling. However, due to their computational efficiency, both representations are often used in global biogeochemical models (e.g. Najjar et al., 2007; Ridgwell et al., 2007; Goosse et al., 2010). A superior approach is the vertically integrated dynamic model, which represents the whole sediment column as a single box (Hülse et al., 2016). Here, OM deposited on the seafloor is added to the sediment box where it gets degraded and dissolved  
85 species diffuse through the sediment-water interface in accordance with these transformations. This approach thus ignores the vertical extent of the sediments and the temporary storage of dissolved species (Soetaert et al., 2000). Yet, it is computationally efficient and also allows differentiating between various fractions of organic matter. Most EMICs incorporate a vertically integrated dynamic model for particulate inorganic carbon only (i.e. mainly  $\text{CaCO}_3$ ) and just a few consider oxic-only  
90 sediment degradation of organic matter (Hülse et al., 2016). The most complex description of marine sediments in Earth System models are vertically resolved diagenetic models, solving the one-dimensional, fully coupled reaction-transport equation for multiple solid and dissolved species (e.g. Berner, 1980; Boudreau, 1997). These models account for important transport processes (e.g. bioturbation, molecular diffusion and bioirrigation), resolve the coupled biogeochemical dynamics of

95 dissolved species and usually at least a part of the resulting characteristic redox-zonation of marine sediments (Soetaert et al., 2000; Hülse et al., 2016). Examples of global carbon cycle models employing a vertically resolved diagenetic model are PISCES (Gehlen et al., 2006) and CLIMBER 3 $\alpha$  (Ilyina et al., 2013), both using a version of Heinze et al. (1999), however, just utilizing a 1G approach with a constant degradation rate constant. None of the EMICs reviewed by Hülse et al.  
100 (2016) use such a sediment representation.

**Problem with that:**

Obviously, such simplifications can neither account for the observed vast structural complexity in natural organic matter and its resulting different degradation rates nor for the rapid decrease in OM degradability in the uppermost centimeters of the sediments (Arndt et al., 2013). At least a 3G approach is necessary to accurately represent organic matter dynamics in this part of the sediments where most OM is degraded (e.g. Soetaert et al., 1996b). Furthermore, the spatial variability in benthic OM degradation kinetics is unknown at the global scale and reported rate constants can vary by almost 10 orders of magnitude (Arndt et al., 2013). Thus, a major challenge for diagenetic models is defining appropriate OM degradation rate constants which is either achieved through profile fitting for a specific site or, for global applications, the rate constants follow empirical relations with a related, readily available sediment characteristic such as water depth (Middelburg et al., 1997), sedimentation rate (Toth and Lerman, 1977; Tromp et al., 1995) or OM flux (Boudreau, 1997). However, these relationships are mostly based on simple fitting exercises to limited data sets and show at the most a very weak trend and no statistically significant relationship (Arndt et al., 2013). This creates considerable uncertainties when diagenetic models are coupled to global biogeochemical ocean models especially when applying the model for different geological timescales. Other model parameters implicitly accounting for processes not explicitly described are sedimentation and bioturbation rate and are generally related to water depth following Middelburg et al. (1997).

**Solution presented here:**

120 Here, we present the OrganicMatter ENabled SEDiment model (OMEN-SED 1.0), a new, one-dimensional, numerically efficient reactive transport model (RTM) that describes OM cycling as well as the associated dynamics of the most important terminal electron acceptors (TEAs, i.e. O<sub>2</sub>, NO<sub>3</sub>, SO<sub>4</sub>, CH<sub>4</sub>), related reduced substances (NH<sub>4</sub>, H<sub>2</sub>S), macronutrients (PO<sub>4</sub>) and associated pore water quantities (ALK, DIC). OMEN-SED's computational efficiency allows its coupling to  
125 Earth System Models of different complexities and therefore the investigation of coupled global biogeochemical dynamics over geological timescales. Here, the model is presented as a 2G-approach, however, a third, nondegradable OM pool can be added and furthermore OMEN-SED can easily be extended to a Multi-G approach.

## 2 Model Description

130 The following section provides a detailed description of the new model. Table 1 summarizes the biogeochemical reaction network and a glossary of parameters along with their respective units is provided in Tables 9 and 10.

### 2.1 General Model Approach

135 The calculation of benthic return/uptake and burial fluxes is based on the vertically resolved conservation equation for solid and dissolved species in porous media is given by (e.g. Berner, 1980; Boudreau, 1997):

$$\frac{\partial \xi C_i}{\partial t} = -\frac{\partial F}{\partial z} + \xi \sum_j R_i^j \quad (1)$$

140 where  $C_i$  is the concentration of the biogeochemical species  $i$ ,  $\xi$  equals the porosity  $\phi$  for solute species and  $(1-\phi)$  for solid species, hence represents the partitioning of species  $i$  into the solute and dissolved phase. The term  $z$  is the sediment depth,  $t$  denotes the time,  $F$  summarises the transport fluxes and  $\sum_j R_i^j$  represents the sum of production/consumption rates  $j$  that affect species  $i$ . The reaction network accounts for the most important primary and secondary redox reactions, equilibrium reactions, mineral dissolution and precipitation, as well as adsorption and desorption processes that 145 affect the explicitly resolved chemical species.

150 State-of-the-art reaction-transport models generally solve the ordinary differential equation (ODE) (eq. 1) numerically and thus allow to account for transient dynamics, depth-varying parameters or a high degree of coupling between different chemical species (e.g. Soetaert et al., 1996b; Aguilera et al., 2005). Yet, numerical models are computational expensive, thus rendering their application in an Earth System Model framework prohibitive. An analytical solution of equation (1) provides an alternative and computational more efficient approach which enjoyed great popularity in the early days of diagenetic modelling and computer technology. However, early analytical models were often very problem-specific and only considered one or two coupled species (e.g. Berner, 1964). ([e.g. Lehrman, Berner](#)) ?? which other pub? Over the next decades, a number of more complex analytical 155 models describing the coupled dynamics of OM degradation, TEAs and reduced substances in sediments were developed (e.g. Billen, 1982; Goloway and Bender, 1982; Jahnke et al., 1982), before the boost in computing power enabled the development of fully-coupled, multi-species, numerical models (e.g. Van Cappellen and Wang, 1995; Soetaert et al., 1996b).

160 Finding an analytical solution to Eq. (1), especially when complex reaction networks are to be considered, is not straightforward and generally requires the assumption of steady state. In addition, the complexity of the reaction network can be reduced by dividing the sediment into distinct zones and accounting for the most pertinent biogeochemical processes within each zone, thus increasing



**Figure 1.** Schematic of the different modelled species and layers in OMEN-SED (1.0). Here showing the case  $z_{\text{ox}} < z_{\text{bio}} < z_{\text{NO}_3} < z_{\text{SO}_4}$ .

the likelihood of finding an analytical solution to Eq. (1) (see Eq. (30) in section 2.5.1 for the general steady-state solution).

165 Therefore, OMEN-SED assumes that benthic dynamics can be represented by a series of steady-states. Because the Earth system model relevant variability in boundary conditions and fluxes is generally longer than the characteristic timescales of the reaction-transport processes, the sediment can be described by a series of pseudo steady-states. In addition, it divides the sediment into a bioturbated and a non-bioturbated zone defined by the constant bioturbation depth  $z_{\text{bio}}$  (see Fig. 1).  
 170 Furthermore, it accounts for the dynamic redox zonation of marine sediments by dividing the sediment into: 2) an oxic zone delineated by the oxygen penetration depth  $z_{\text{ox}}$ , 3) a denitrification zone situated between  $z_{\text{ox}}$  and the nitrate penetration depth  $z_{\text{NO}_3}$ , 4) a sulfate reduction zone situated between  $z_{\text{NO}_3}$  and the sulfate penetration depth  $z_{\text{SO}_4}$  and 5) a methanogenic zone situated below  $z_{\text{SO}_4}$  (Fig. 1). All penetration depths are dynamically calculated by the model. Each zone is characterised  
 175 by a set of diagenetic equations that encapsulate the most pertinent reaction and transport processes in this zone (see section 2.2 and 2.3 for more details).

OMEN calculates and feeds back to the Earth System model the fraction of POC preserved in the sediments and the sediment-water interface fluxes of the dissolved species  $C_i$  (in  $\text{mol cm}^{-2} \text{ year}^{-1}$ ):

180     $\text{Flux\_SWI}(C_i) = \phi \left( D_i \frac{\partial C_i(z)}{\partial z} \Big|_0 - w [C_i(0) - C_i(z_\infty)] \right)$  (2)

where  $w$  is the deposition rate,  $D_i$  is the diffusion coefficient and  $C_i(0)$ ,  $C_i(z_\infty)$  the concentration of species  $i$  at the SWI and at the lower sediment boundary.

## 2.2 Transport

The model accounts for both the advective, as well as the diffusive transport of dissolved and solid species, assuming that sediment compaction is negligible (i.e.  $\frac{\partial \phi}{\partial z} = 0$ ). The molecular diffusion of dissolved species is described via a species-specific apparent diffusion coefficient,  $D_{\text{mol},i}$ . In addition, the activity of infaunal organisms in the bioturbated zone of the sediment ( $z < z_{\text{bio}}$ ) that causes random displacements of sediments and porewaters is simulated using a diffusive term (e.g. Boudreau, 1986), with a constant bioturbation coefficient  $D_{\text{bio}}$  in the bioturbated zone. The pumping activity by burrow-dwelling animals and the resulting ventilation of tubes, the so-called bioirrigation, is encapsulated in a factor,  $f_{ir}$  that enhances the molecular diffusion coefficient ( hence,  $D_{i,0} = D_{\text{mol},i} \cdot f_{ir}$ , Soetaert et al., 1996a). The flux divergence can thus be formulated as:

$$\frac{\partial F}{\partial z} = - \frac{\partial}{\partial z} \left( -\xi D_i \frac{\partial C_i}{\partial z} + \xi w C_i \right) \quad (3)$$

where  $D_i$  is the diffusion coefficient of species  $i$  ( $D_i = D_{i,0} + D_{\text{bio}} = D_{\text{mol},i} \cdot f_{ir} + D_{\text{bio}}$  for dissolved species and  $D_i = D_{\text{bio}}$  for solid species) and  $w$  is the deposition rate. The bioturbation coefficient  $D_{\text{bio}}$  is set to zero below  $z_{\text{bio}}$ . In addition, infaunal activity ceases ( $D_{\text{bio}} = 0$ ) once bottom waters become anoxic ( $O_2 < ??? \text{ mol cm}^{-3}$ ). **check for good value + add if-query in code!!**

## 2.3 Reaction Network

Earth System models generally track the biogeochemical dynamics of organic and inorganic carbon, essential nutrients (nitrogen, phosphorus) and oxygen with the aim of investigating the evolution of the ocean's redox structure and carbonate system and its feedbacks on global climate. This general aim thus defines a minimum set of state variables and reaction processes that need to be resolved for an efficient representation of the benthic-pelagic coupling in Earth system models. A suitable sediment model must provide a robust quantification of organic (and inorganic) carbon burial fluxes, as well as the benthic return fluxes of growth-limiting nutrients, equilibrium invariant and reduced species, and oxygen uptake fluxes. As a consequence, the reaction network must account for the most important primary and secondary redox reactions, equilibrium reactions, mineral precipitation/dissolution and adsorption/desorption, resulting in a complex set of coupled reaction-transport equations. The following subsections provide a short discussion of the reaction processes included in the model and give an overview of the vertically resolved conservation equations and boundary conditions for solid and dissolved species in each layer. Table 1 provides a summary of the reactions

**Table 1.** Reactions and variables implemented in the Reaction Network of OMEN-SED (1.0). The primary and secondary redox reactions are listed in the sequence they occur with increasing sediment depth.

Description	
Primary redox reactions	Degradation of organic matter via aerobic respiration, denitrification, sulfate reduction, methanogenesis (implicit)
Secondary redox reactions	Oxidation of ammonium and sulfide by oxygen, anaerobic oxidation of methane by sulfate
Adsorption/Desorption	Ad-/Desorption of P on/from Fe(OH) <sub>3</sub> , NH <sub>4</sub> adsorption, PO <sub>4</sub> adsorption
Mineral precipitation	Formation of authigenic P
Variables	Organic matter, oxygen, nitrate, ammonium, sulfate, sulfide (hydrogen sulfide), phosphate, Fe-bound P, DIC, ALK

and variables considered in the reaction network. Table 14 summarises their reaction stoichiometry and Table ?? provides an overview of their description in the model.

### 2.3.1 Organic matter or Particulate Organic Carbon (POC)

- 215 In marine sediments, organic matter (OM) is degraded by heterotrophic activity coupled to the sequential utilisation of terminal electron acceptors, typically in the order of O<sub>2</sub>, NO<sub>3</sub><sup>-</sup>, Mn(VI), Fe(III) and SO<sub>4</sub><sup>2-</sup> followed by methanogenesis and/or fermentation. Here, organic matter degradation is described via a multi-G model approach (Arndt et al., 2013, and references therein), assuming that the bulk OM consists of a number of discrete compound classes  $C_i$  characterised by  
 220 specific degradation rate constants  $k_i$ . Such a multi-G approach allows for selective preservation of compound classes according to their reactivity,  $k_i$  and, thus, accounts for the change in organic matter reactivity during burial. Each compound class is degraded according to first-order kinetics. The conservation equation for organic matter dynamics is thus given by:

$$225 \quad \frac{\partial C_i}{\partial t} = 0 = D_{C_i} \frac{\partial^2 C_i}{\partial z^2} - w \frac{\partial C_i}{\partial z} - k_i \cdot C_i \quad (4)$$

The analytical solution of Eq. (4) (see section 2.5 for details) requires the definition of a set of boundary conditions (Table 2). The model assumes a known concentration/flux at the sediment-water interface and continuity across the bottom of the bioturbated zone,  $z_{\text{bio}}$ .

### 2.3.2 Oxygen

- 230 In marine sediments, oxygen is consumed through the aerobic degradation of organic matter and a number of secondary redox reactions. In the oxic layer ( $z < z_{\text{ox}}$ ), the model explicitly accounts for the aerobic degradation of OM, which consumes oxygen with a fixed O<sub>2</sub> : C ratio (O<sub>2</sub>C, Tab. 10) and produces ammonium, which is partially nitrified to nitrate ( $\gamma_{\text{NH}_4}$ ). In addition, the oxygen

**Table 2.** Boundary conditions for organic matter. For the boundaries we define:  $z_{\text{bio}}^- := \lim_{h \rightarrow 0} (z_{\text{bio}} - h)$  and  $z_{\text{bio}}^+ := \lim_{h \rightarrow 0} (z_{\text{bio}} + h)$ .

Boundary	Condition	
$z = 0$	known concentration	1) $C_i(0) = C_{i0}$
$z = z_{\text{bio}}$	continuity	2) $C_i(z_{\text{bio}}^-) = C_i(z_{\text{bio}}^+)$ 3) $-D_{\text{bio}} \cdot \frac{\partial C_i}{\partial z} \Big _{z_{\text{bio}}^-} = 0$

**Table 3.** Boundary conditions for oxygen. For the boundaries we define:  $z_{\text{bio}}^- := \lim_{h \rightarrow 0} (z_{\text{bio}} - h)$  and  $z_{\text{bio}}^+ := \lim_{h \rightarrow 0} (z_{\text{bio}} + h)$ .

Boundary	Condition	
$z = 0$	known concentration	1) $O_2(0) = O_{20}$
$z = z_{\text{bio}}$	continuity	2) $O_2(z_{\text{bio}}^-) = O_2(z_{\text{bio}}^+)$ 3) $-(D_{O_2,0} + D_{\text{bio}}) \cdot \frac{\partial O_2}{\partial z} \Big _{z_{\text{bio}}^-} = -D_{O_2,0} \cdot \frac{\partial O_2}{\partial z} \Big _{z_{\text{bio}}^+}$
$z = z_{\text{ox}}$	$O_2$ consumption ( $z_{\text{ox}} = z_\infty$ )	4) <b>IF</b> ( $O_2(z_\infty) > 0$ ) $\frac{\partial O_2}{\partial z} \Big _{z_{\text{ox}}} = 0$ <b>ELSE</b> $O_2(z_{\text{ox}}) = 0$ and $-D_{O_2} \cdot \frac{\partial O_2}{\partial z} \Big _{z_{\text{ox}}} = F_{\text{red}}(z_{\text{ox}})$ $F_{\text{red}}(z_{\text{ox}}) = \frac{1-\phi}{\phi} \cdot \int_{z_{\text{ox}}}^\infty \sum_i (2\gamma_{\text{NH}_4} \text{NC}_i + \gamma_{\text{H}_2\text{S}} \text{SO}_4 \text{C}) k_i C_i dz$
	$(z_{\text{ox}} < z_\infty)$ with flux from below	

consumption through the oxidation of reduced species ( $\text{NH}_4$ ,  $\text{H}_2\text{S}$ ) produced in the suboxic and  
235 anoxic layers of the sediment is implicitly taken into account through the flux boundary condition  
at the dynamic oxygen penetration depth  $z_{\text{ox}}$ . This simplification can be justified as it has been  
shown that these secondary redox reactions mainly occur at the oxic/suboxic interface (Soetaert  
et al., 1996b). The factor  $\frac{1-\phi}{\phi}$  accounts for the volume conversion from the solid to the dissolved  
phase. Oxygen dynamics are thus described by:

$$240 \quad \frac{\partial O_2}{\partial t} = 0 = D_{O_2} \frac{\partial^2 O_2}{\partial z^2} - w \frac{\partial O_2}{\partial z} - \frac{1-\phi}{\phi} \sum_i k_i \cdot [O_2 \text{C} + 2\gamma_{\text{NH}_4} \text{NC}_i] \cdot C_i(z) \quad (5)$$

The analytical solution of Eq. (5) requires the definition of boundary conditions (Table 3). The model  
assumes a known bottom water concentration and the complete consumption of oxygen at the oxygen  
penetration depth (or zero flux if  $z_{\text{ox}} = z_\infty$ ). It considers equal oxygen concentration and diffusive  
245 flux above ( $z_{\text{bio}}^-$ ) and below ( $z_{\text{bio}}^+$ ) the bioturbation boundary. In addition, the model imposes a flux  
of reduced species through the bottom of the oxic zone that is calculated as the reduced substances  
produced through anoxic mineralization of organic matter below  $z_{\text{ox}}$ . Thus, assuming that fractions  
( $\gamma_{\text{NH}_4}$  and  $\gamma_{\text{H}_2\text{S}}$ ) of these reduced species are oxidised at the oxic/suboxic interface.

### 2.3.3 Nitrate and Ammonium

250 To model nitrate and ammonium dynamics the sediment is partitioned into two geochemical layers (oxic and suboxic), where different equations describe the biogeochemical processes. Above the oxygen penetration depth organic matter mineralization produces ammonium, which is partly nitrified to nitrate (the fraction  $\gamma_{\text{NH}_4}$ ). In the suboxic zone ( $z > z_{\text{ox}}$ ), oxygen concentration is zero and nitrate serves as the electron acceptor to respire organic matter, thus nitrate is consumed by denitrification and ammonium is produced. Below the nitrate penetration depth  $z_{\text{NO}_3}$ , ammonium is still produced through OM mineralization. Therefore the diagenetic equations for nitrate and ammonium are given by:

1. Layer ( $z \leq z_{\text{ox}}$ )

$$260 \frac{\partial \text{NO}_3^I}{\partial t} = 0 = D_{\text{NO}_3} \frac{\partial^2 \text{NO}_3^I}{\partial z^2} - w \frac{\partial \text{NO}_3^I}{\partial z} + \gamma_{\text{NH}_4} \frac{1-\phi}{\phi} \cdot \sum_i \text{NC}_i \cdot k_i \cdot C_i(z) \quad (6)$$

$$\frac{\partial \text{NH}_4^I}{\partial t} = 0 = \frac{D_{\text{NH}_4}}{1+K_{\text{NH}_4}} \frac{\partial^2 \text{NH}_4^I}{\partial z^2} - w \frac{\partial \text{NH}_4^I}{\partial z} + \frac{1-\gamma_{\text{NH}_4}}{1+K_{\text{NH}_4}} \cdot \frac{1-\phi}{\phi} \cdot \sum_i \text{NC}_i \cdot k_i \cdot C_i(z) \quad (7)$$

2. Layer ( $z_{\text{ox}} < z \leq z_{\text{NO}_3}$ )

$$\frac{\partial \text{NO}_3^{II}}{\partial t} = 0 = D_{\text{NO}_3} \frac{\partial^2 \text{NO}_3^{II}}{\partial z^2} - w \frac{\partial \text{NO}_3^{II}}{\partial z} - \frac{1-\phi}{\phi} \text{NO}_3 C \cdot \sum_i k_i \cdot C_i(z) \quad (8)$$

$$265 \frac{\partial \text{NH}_4^{II}}{\partial t} = 0 = \frac{D_{\text{NH}_4}}{1+K_{\text{NH}_4}} \frac{\partial^2 \text{NH}_4^{II}}{\partial z^2} - w \frac{\partial \text{NH}_4^{II}}{\partial z} \quad (9)$$

3. Layer ( $z_{\text{NO}_3} < z \leq z_{\infty}$ )

$$\frac{\partial \text{NH}_4^{III}}{\partial t} = 0 = \frac{D_{\text{NH}_4}}{1+K_{\text{NH}_4}} \frac{\partial^2 \text{NH}_4^{III}}{\partial z^2} - w \frac{\partial \text{NH}_4^{III}}{\partial z} + \frac{1}{1+K_{\text{NH}_4}} \cdot \frac{1-\phi}{\phi} \cdot \sum_i \text{NC}_i \cdot k_i \cdot C_i(z) \quad (10)$$

270 The boundary conditions to solve Equations 6 - 10 are summarized in Table 4. The model assumes known bottom water concentrations for both species, the complete consumption of nitrate at the nitrate penetration depth (or zero flux if  $z_{\text{NO}_3} = z_{\infty}$ ) and no change in ammonium flux at  $z_{\infty}$ . It considers equal concentrations and diffusive fluxes at  $z_{\text{bio}}$  and  $z_{\text{ox}}$ . In addition, the re-oxidation of upward-diffusing reduced ammonium is considered in the oxic-suboxic boundary condition for 275 nitrate and ammonium.

### 2.3.4 Sulfate and Sulfide

When nitrate is depleted, sulfate reduction is the pathway to mineralize organic matter, thus consuming sulfate ( $\text{SO}_4$ ) and producing hydrogen sulfide ( $\text{H}_2\text{S}$ ) until the sulfate penetration depth ( $z_{\text{SO}_4}$ ).

280 Sulfate and sulfide dynamics are thus described by:

**Table 4.** Boundary conditions for nitrate and ammonium. For the boundaries we define:  $z_-^- := \lim_{h \rightarrow 0} (z_- - h)$  and  $z_-^+ := \lim_{h \rightarrow 0} (z_- + h)$ .

Boundary	Condition	
$z = 0$	known concentration	1) $NO_3(0) = NO_{30}$
$z = z_{\text{bio}}$	continuity	2) $NO_3(z_{\text{bio}}^-) = NO_3(z_{\text{bio}}^+)$
$z = z_{\text{ox}}$	continuity	3) $-(D_{NO_3,0} + D_{\text{bio}}) \cdot \frac{\partial NO_3}{\partial z}  _{z_{\text{bio}}^-} = -D_{NO_3,0} \cdot \frac{\partial NO_3}{\partial z}  _{z_{\text{bio}}^+}$
	where:	4) $NO_3(z_{\text{ox}}^-) = NO_3(z_{\text{ox}}^+)$
$z = z_{NO_3}$	NO <sub>3</sub> consumption ( $z_{NO_3} = z_\infty$ )	5) $-D_{NO_3} \cdot \frac{\partial NO_3}{\partial z}  _{z_{\text{ox}}^-} + \gamma_{NH_4} \cdot F_{NH_4}(z_{\text{ox}}) = -D_{NO_3} \cdot \frac{\partial NO_3}{\partial z}  _{z_{\text{ox}}^+}$ $F_{NH_4}(z_{\text{ox}}) = \frac{1}{1+K_{NH_4}} \cdot \frac{1-\phi}{\phi} \cdot \int_{z_{NO_3}}^\infty \sum_i k_i \cdot NC_i \cdot C_i dz$
		6) <b>IF</b> ( $NO_3(z_\infty) > 0$ ) $\frac{\partial NO_3}{\partial z}  _{z_{NO_3}} = 0$ <b>ELSE</b>
		7) $NO_3(z_{NO_3}) = 0$
$z = 0$	known concentration	1) $NH_4(0) = NH_{40}$
$z = z_{\text{bio}}$	continuity	2) $NH_4(z_{\text{bio}}^-) = NH_4(z_{\text{bio}}^+)$
$z = z_{\text{ox}}$	continuity	3) $-\frac{D_{NH_4,0} + D_{\text{bio}}}{1+K_{NH_4}} \cdot \frac{\partial NH_4}{\partial z}  _{z_{\text{bio}}^-} = -\frac{D_{NH_4,0}}{1+K_{NH_4}} \cdot \frac{\partial NH_4}{\partial z}  _{z_{\text{bio}}^+}$
	where:	4) $NH_4(z_{\text{ox}}^-) = NH_4(z_{\text{ox}}^+)$
$z = z_{NO_3}$	continuity flux	5) $-\frac{D_{NH_4}}{1+K_{NH_4}} \cdot \frac{\partial NH_4}{\partial z}  _{z_{\text{ox}}^-} - \gamma_{NH_4} \cdot F_{NH_4}(z_{\text{ox}}) = -\frac{D_{NH_4}}{1+K_{NH_4}} \cdot \frac{\partial NH_4}{\partial z}  _{z_{\text{ox}}^+}$ $F_{NH_4}(z_{\text{ox}}) = \frac{1}{1+K_{NH_4}} \cdot \frac{1-\phi}{\phi} \cdot \int_{z_{NO_3}}^\infty \sum_i k_i \cdot NC_i \cdot C_i dz$
$z = z_\infty$	zero NH <sub>4</sub> flux	6) $NH_4(z_{NO_3}^-) = NH_4(z_{NO_3}^+)$
		7) $-\frac{D_{NH_4}}{1+K_{NH_4}} \cdot \frac{\partial NH_4}{\partial z}  _{z_{NO_3}^-} = -\frac{D_{NH_4}}{1+K_{NH_4}} \cdot \frac{\partial NH_4}{\partial z}  _{z_{NO_3}^+}$
		8) $\frac{\partial NH_4}{\partial z}  _{z_\infty} = 0$

1. Layer ( $z \leq z_{NO_3}$ )

$$\frac{\partial SO_4^{I}}{\partial t} = 0 = D_{SO_4} \frac{\partial^2 SO_4^{I}}{\partial z^2} - w \frac{\partial SO_4^{I}}{\partial z} \quad (11)$$

$$\frac{\partial H_2S^{I}}{\partial t} = 0 = D_{H_2S} \frac{\partial^2 H_2S^{I}}{\partial z^2} - w \frac{\partial H_2S^{I}}{\partial z} \quad (12)$$

285 2. Layer ( $z_{NO_3} < z \leq z_{SO_4}$ )

$$\frac{\partial SO_4^{II}}{\partial t} = 0 = D_{SO_4} \frac{\partial^2 SO_4^{II}}{\partial z^2} - w \frac{\partial SO_4^{II}}{\partial z} - \frac{1-\phi}{\phi} \cdot \sum_i SO_4C \cdot k_i \cdot C_i(z) \quad (13)$$

$$\frac{\partial H_2S^{II}}{\partial t} = 0 = D_{H_2S} \frac{\partial^2 H_2S^{II}}{\partial z^2} - w \frac{\partial H_2S^{II}}{\partial z} + \frac{1-\phi}{\phi} \cdot \sum_i SO_4C \cdot k_i \cdot C_i(z) \quad (14)$$

3. Layer ( $z_{SO_4} < z \leq z_\infty$ )

$$290 \quad \frac{\partial H_2S^{III}}{\partial t} = 0 = D_{H_2S} \frac{\partial^2 H_2S^{III}}{\partial z^2} - w \frac{\partial H_2S^{III}}{\partial z} \quad (15)$$

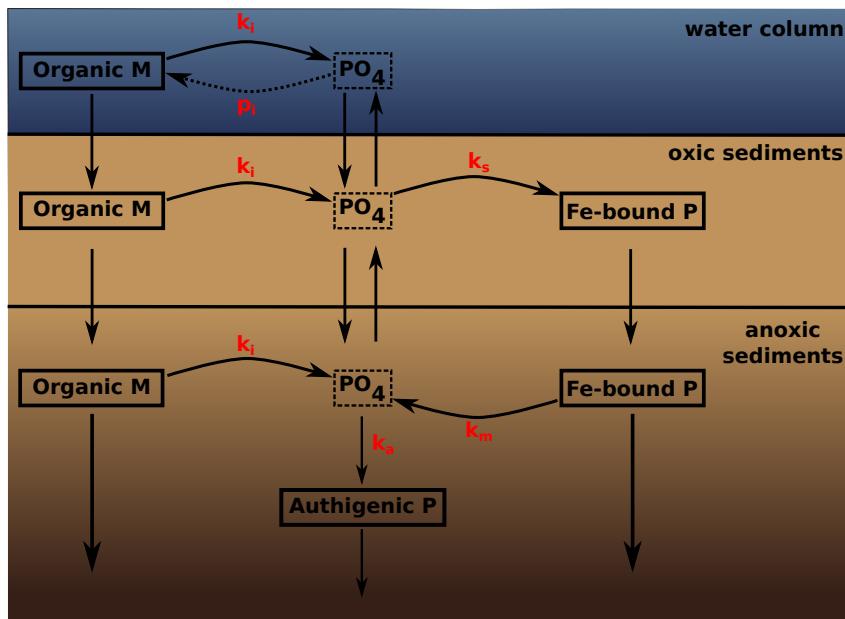
To solve equations 11 - 15 the model assumes known concentrations at the sediment-water interface and continuity across the bioturbation depth and the nitrate penetration depth (see Table 5).

**Table 5.** Boundary conditions for sulfate and sulfide. For the boundaries we define:  $z_-^- := \lim_{h \rightarrow 0} (z_- - h)$  and  $z_-^+ := \lim_{h \rightarrow 0} (z_- + h)$ .

Boundary	Condition	
$z = 0$	known concentration	1) $\text{SO}_4(0) = \text{SO}_{40}$
$z = z_{\text{bio}}$	continuity	2) $\text{SO}_4(z_{\text{bio}}^-) = \text{SO}_4(z_{\text{bio}}^+)$
	flux	3) $-(D_{\text{SO}_4,0} + D_{\text{bio}}) \cdot \frac{\partial \text{SO}_4}{\partial z}  _{z_{\text{bio}}^-} = -D_{\text{SO}_4,0} \cdot \frac{\partial \text{SO}_4}{\partial z}  _{z_{\text{bio}}^+}$
$z = z_{\text{ox}}$	continuity	4) $\text{SO}_4(z_{\text{ox}}^-) = \text{SO}_4(z_{\text{ox}}^+)$
	flux	5) $-D_{\text{SO}_4} \cdot \frac{\partial \text{SO}_4}{\partial z}  _{z_{\text{ox}}^-} + \gamma_{\text{H}_2\text{S}} \cdot F_{\text{H}_2\text{S}}(z_{\text{ox}}) = -D_{\text{SO}_4} \cdot \frac{\partial \text{SO}_4}{\partial z}  _{z_{\text{ox}}^+}$ $F_{\text{H}_2\text{S}}(z_{\text{ox}}) = \frac{1-\phi}{\phi} \cdot \left( \int_{z_{\text{NO}_3}}^{z_{\text{SO}_4}} \sum_i \text{SO}_4 \cdot k_i \cdot C_i dz + \gamma_{\text{CH}_4} \cdot \int_{z_{\text{SO}_4}}^{\infty} \sum_i \text{MC} \cdot k_i \cdot C_i dz \right)$
	where:	
$z = z_{\text{NO}_3}$	continuity	6) $\text{SO}_4(z_{\text{NO}_3}^-) = \text{SO}_4(z_{\text{NO}_3}^+)$
	flux	7) $-D_{\text{SO}_4} \cdot \frac{\partial \text{SO}_4}{\partial z}  _{z_{\text{NO}_3}^-} = -D_{\text{SO}_4} \cdot \frac{\partial \text{SO}_4}{\partial z}  _{z_{\text{NO}_3}^+}$
$z = z_{\text{SO}_4}$	$\text{SO}_4$ consumption ( $z_{\text{SO}_4} = z_\infty$ )	8) <b>IF</b> ( $\text{SO}_4(z_\infty) > 0$ ) $\frac{\partial \text{SO}_4}{\partial z}  _{z_{\text{SO}_4}} = 0$ <b>ELSE</b> $\text{SO}_4(z_{\text{SO}_4}) = 0$ and $-D_{\text{SO}_4} \cdot \frac{\partial \text{SO}_4}{\partial z}  _{z_{\text{SO}_4}} = \gamma_{\text{CH}_4} \cdot F_{\text{CH}_4}(z_{\text{SO}_4})$ $F_{\text{CH}_4}(z_{\text{SO}_4}) = \frac{1-\phi}{\phi} \cdot \int_{z_{\text{SO}_4}}^{\infty} \sum_i \text{MC} \cdot k_i \cdot C_i dz$
	$(z_{\text{SO}_4} < z_\infty)$ with flux from below:	
$z = 0$	known concentration	1) $\text{H}_2\text{S}(0) = \text{H}_2\text{S}_0$
$z = z_{\text{bio}}$	continuity	2) $\text{H}_2\text{S}(z_{\text{bio}}^-) = \text{H}_2\text{S}(z_{\text{bio}}^+)$
	flux	3) $-(D_{\text{H}_2\text{S},0} + D_{\text{bio}}) \cdot \frac{\partial \text{H}_2\text{S}}{\partial z}  _{z_{\text{bio}}^-} = -D_{\text{H}_2\text{S},0} \cdot \frac{\partial \text{H}_2\text{S}}{\partial z}  _{z_{\text{bio}}^+}$
$z = z_{\text{ox}}$	continuity	4) $\text{H}_2\text{S}(z_{\text{ox}}^-) = \text{H}_2\text{S}(z_{\text{ox}}^+)$
	flux	5) $-D_{\text{H}_2\text{S}} \cdot \frac{\partial \text{H}_2\text{S}}{\partial z}  _{z_{\text{ox}}^-} - \gamma_{\text{H}_2\text{S}} F_{\text{H}_2\text{S}}(z_{\text{ox}}) = -D_{\text{H}_2\text{S}} \cdot \frac{\partial \text{H}_2\text{S}}{\partial z}  _{z_{\text{ox}}^+}$ $F_{\text{H}_2\text{S}}(z_{\text{ox}}) = \frac{1-\phi}{\phi} \cdot \left( \int_{z_{\text{NO}_3}}^{z_{\text{SO}_4}} \sum_i \text{SO}_4 \cdot k_i \cdot C_i dz + \gamma_{\text{CH}_4} \cdot \int_{z_{\text{SO}_4}}^{\infty} \sum_i \text{MC} \cdot k_i \cdot C_i dz \right)$
	where:	
$z = z_{\text{NO}_3}$	continuity	6) $\text{H}_2\text{S}(z_{\text{NO}_3}^-) = \text{H}_2\text{S}(z_{\text{NO}_3}^+)$
	flux	7) $-D_{\text{H}_2\text{S}} \cdot \frac{\partial \text{H}_2\text{S}}{\partial z}  _{z_{\text{NO}_3}^-} = -D_{\text{H}_2\text{S}} \cdot \frac{\partial \text{H}_2\text{S}}{\partial z}  _{z_{\text{NO}_3}^+}$
$z = z_{\text{SO}_4}$	continuity flux (with AOM)	8) $\text{H}_2\text{S}(z_{\text{SO}_4}^-) = \text{H}_2\text{S}(z_{\text{SO}_4}^+)$ 9) $-D_{\text{H}_2\text{S}} \cdot \frac{\partial \text{H}_2\text{S}}{\partial z}  _{z_{\text{SO}_4}^-} + \gamma_{\text{CH}_4} \cdot F_{\text{CH}_4}(z_{\text{SO}_4}) = -D_{\text{H}_2\text{S}} \cdot \frac{\partial \text{H}_2\text{S}}{\partial z}  _{z_{\text{SO}_4}^+}$ $F_{\text{CH}_4}(z_{\text{SO}_4}) = \frac{1-\phi}{\phi} \cdot \int_{z_{\text{SO}_4}}^{\infty} \sum_i \text{MC} \cdot k_i \cdot C_i dz$
$z = z_\infty$	where: zero $\text{H}_2\text{S}$ flux	10) $\frac{\partial \text{H}_2\text{S}}{\partial z}  _{z_\infty} = 0$

The re-oxidation of reduced  $\text{H}_2\text{S}$  to  $\text{SO}_4$  is considered in the oxic-suboxic boundary condition for 295 both species, here including the methanic zone, as  $\text{H}_2\text{S}$  is also produced during anaerobic oxidation of methane (AOM). Furthermore, sulfate is used at  $z_{\text{SO}_4}$  to oxidize methane from below and thus producing  $\text{H}_2\text{S}$ . In case  $z_{\text{SO}_4} < z_\infty$ , sulfate concentration is zero at  $z_{\text{SO}_4}$  and its diffusive flux must equal the amount of methane produced below; or, in case  $z_{\text{SO}_4} = z_\infty$ , a zero flux condition for sulfate is considered. At lower boundary ( $z_\infty$ ) zero flux of  $\text{H}_2\text{S}$  is considered. **correct??**

### 2.3.5 Phosphate



**Figure 2.** A schematic of the sedimentary P cycle in OMEN-SED (1.0). Red numbers represent kinetic rate constants for phosphorus dynamics (compare Table 10;  $p_i$  represents uptake rate of  $\text{PO}_4$  via primary production in shallow environments). Adapted from Slomp et al. (1996).

To model phosphorus (P) in the sediments the model takes into account the change with depth of phosphate ( $\text{PO}_4$ ) and iron-bound P, thereby mainly following the description of Slomp et al. (1996) and Gypens et al. (2008). Throughout the sediment column organic matter is mineralized resulting  
305 in a release of phosphate to the pore water. In the oxic part of the sediment, this  $\text{PO}_4$  either diffuses upward to the water column or is adsorbed to Fe oxides forming Fe-bound P (or M) (Slomp et al., 1998). In the suboxic/anoxic zone,  $\text{PO}_4$  is not only produced through organic matter degradation but is also released from the Fe-bound P pool due to the reduction of Fe oxides. Furthermore, phosphate concentrations can become high enough in this layer for authigenic mineral formation to occur  
310 (Van Cappellen and Berner, 1988). This phosphorus bound in authigenic minerals represents a permanent sink for reactive phosphorus (Slomp et al., 1996). See Figure 2 for a schematic overview of the sedimentary P cycle. Therefore the diagenetic equations for phosphorus are written:

1. Layer ( $z \leq z_{\text{ox}}$ )

$$315 \quad \frac{\partial \text{PO}_4^I}{\partial t} = \frac{D_{\text{PO}_4}}{1 + K_{\text{PO}_4}^I} \frac{\partial^2 \text{PO}_4^I}{\partial z^2} - w \frac{\partial \text{PO}_4^I}{\partial z} + \frac{1-\phi}{\phi} \frac{1}{1 + K_{\text{PO}_4}^I} \sum_i (\text{PC}_i \cdot k_i \cdot C_i(z)) \\ - \frac{k_s}{1 + K_{\text{PO}_4}^I} (\text{PO}_4^I - \text{PO}_4^s) \quad (16)$$

$$\frac{\partial M^I}{\partial t} = D_M \frac{\partial^2 M^I}{\partial z^2} - w \frac{\partial M^I}{\partial z} + \frac{\phi}{1-\phi} k_s (\text{PO}_4^I - \text{PO}_4^s) \quad (17)$$

2. Layer ( $z_{\text{ox}} < z$ )

$$320 \quad \frac{\partial M^{II}}{\partial t} = D_M \frac{\partial^2 M^{II}}{\partial z^2} - w \frac{\partial M^{II}}{\partial z} - k_m (M^{II} - M^\infty) \quad (18)$$

$$\frac{\partial \text{PO}_4^{II}}{\partial t} = \frac{D_{\text{PO}_4}}{1 + K_{\text{PO}_4}^{II}} \frac{\partial^2 \text{PO}_4^{II}}{\partial z^2} - w \frac{\partial \text{PO}_4^{II}}{\partial z} + \frac{1-\phi}{\phi} \frac{1}{1 + K_{\text{PO}_4}^{II}} \sum_i (\text{PC}_i \cdot k_i \cdot C_i(z)) \\ - \frac{k_a}{1 + K_{\text{PO}_4}^{II}} (\text{PO}_4^{II} - \text{PO}_4^a) + \frac{(1-\phi)}{\phi} \frac{k_m}{1 + K_{\text{PO}_4}^{II}} (M^{II} - M^\infty) \quad (19)$$

(20)

325 The boundary conditions to solve Equations 16 - 19 are summarized in Table 6. The model assumes known bottom water concentrations and equal concentrations and diffusive fluxes at  $z_{\text{bio}}$  and  $z_{\text{ox}}$  for both species. Additionally it considers no change in phosphate flux and an asymptotic Fe-bound P concentration at  $z_\infty$ .

330 @ Sandra: SWI Flux for M does not exist, right???

### 2.3.6 Dissolved Inorganic Carbon (DIC)

Dissolved inorganic carbon (DIC) is produced when organic matter is degraded. DIC dynamics are described by:

335 1. Layer ( $z \leq z_{\text{SO}_4}$ )

$$\frac{\partial \text{DIC}^I}{\partial t} = 0 = D_{\text{DIC}} \frac{\partial^2 \text{DIC}^I}{\partial z^2} - w \frac{\partial \text{DIC}^I}{\partial z} + \frac{1-\phi}{\phi} \cdot \sum_i \text{DICC}^I \cdot k_i \cdot C_i(z) \quad (21)$$

2. Layer ( $z_{\text{SO}_4} < z \leq z_\infty$ )

$$340 \quad \frac{\partial \text{DIC}^{II}}{\partial t} = 0 = D_{\text{DIC}} \frac{\partial^2 \text{DIC}^{II}}{\partial z^2} - w \frac{\partial \text{DIC}^{II}}{\partial z} + \frac{1-\phi}{\phi} \cdot \sum_i \text{DICC}^{II} \cdot k_i \cdot C_i(z) \quad (22)$$

To solve equations 21 and 22 the model assumes the boundary conditions summarised in Table 7.

**Table 6.** Boundary conditions for phosphate and Fe-bound P (M). For the boundaries we define:  $z_-^- := \lim_{h \rightarrow 0} (z_- - h)$  and  $z_-^+ := \lim_{h \rightarrow 0} (z_- + h)$ .

Boundary	Condition	
$z = 0$	known concentration	1) $\text{PO}_4(0) = \text{PO}_{40}$
$z = z_{\text{bio}}$	continuity	2) $\text{PO}_4(z_{\text{bio}}^-) = \text{PO}_4(z_{\text{bio}}^+)$
	flux	3) $(D_{\text{PO}_4,0} + D_{\text{bio}}) \cdot \frac{\partial \text{PO}_4}{\partial z}  _{z_{\text{bio}}^-} = D_{\text{PO}_4,0} \cdot \frac{\partial \text{PO}_4}{\partial z}  _{z_{\text{bio}}^+}$
$z = z_{\text{ox}}$	continuity	4) $\text{PO}_4(z_{\text{ox}}^-) = \text{PO}_4(z_{\text{ox}}^+)$
	flux	5) $- \frac{D_{\text{PO}_4}}{1 + K_{\text{PO}_4}^I} \cdot \frac{\partial \text{PO}_4}{\partial z}  _{z_{\text{ox}}^-} = - \frac{D_{\text{PO}_4}}{1 + K_{\text{PO}_4}^{II}} \cdot \frac{\partial \text{PO}_4}{\partial z}  _{z_{\text{ox}}^+}$
$z = z_\infty$	flux	10) $\frac{\partial \text{PO}_4}{\partial z}  _{z_\infty} = 0$
$z = 0$	known concentration	1) $M(0) = M_0$
$z = z_{\text{bio}}$	continuity	2) $M(z_{\text{bio}}^-) = M(z_{\text{bio}}^+)$
	flux	3) $\frac{\partial M}{\partial z}  _{z_{\text{bio}}^-} = \frac{\partial M}{\partial z}  _{z_{\text{bio}}^+}$
$z = z_{\text{ox}}$	continuity	4) $M(z_{\text{ox}}^-) = M(z_{\text{ox}}^+)$
	flux	5) $\frac{\partial M}{\partial z}  _{z_{\text{ox}}^-} = \frac{\partial M}{\partial z}  _{z_{\text{ox}}^+}$
$z = z_\infty$	assymptotic concentration	10) $M(z_\infty) = M_\infty$

**Table 7.** Boundary conditions for DIC. For the boundaries we define:  $z_-^- := \lim_{h \rightarrow 0} (z_- - h)$  and  $z_-^+ := \lim_{h \rightarrow 0} (z_- + h)$ .

Boundary	Condition	
$z = 0$	known concentration	1) $\text{DIC}(0) = \text{DIC}_0$
$z = z_{\text{bio}}$	continuity	2) $\text{DIC}(z_{\text{bio}}^-) = \text{DIC}(z_{\text{bio}}^+)$
	flux	3) $-(D_{\text{DIC},0} + D_{\text{bio}}) \cdot \frac{\partial \text{DIC}}{\partial z}  _{z_{\text{bio}}^-} = -D_{\text{DIC},0} \cdot \frac{\partial \text{DIC}}{\partial z}  _{z_{\text{bio}}^+}$
$z = z_{\text{SO}_4}$	continuity	8) $\text{DIC}(z_{\text{SO}_4}^-) = \text{DIC}(z_{\text{SO}_4}^+)$
	flux (with AOM)	9) $-D_{\text{DIC}} \cdot \frac{\partial \text{DIC}}{\partial z}  _{z_{\text{SO}_4}^-} + \gamma_{\text{CH}_4} \cdot F_{\text{CH}_4}(z_{\text{SO}_4}) = -D_{\text{DIC}} \cdot \frac{\partial \text{DIC}}{\partial z}  _{z_{\text{SO}_4}^+}$
	where:	$F_{\text{CH}_4}(z_{\text{SO}_4}) = \frac{1-\phi}{\phi} \cdot \int_{z_{\text{SO}_4}}^\infty \sum_i \text{MC} \cdot k_i \cdot C_i dz$
$z = z_\infty$	zero DIC flux	10) $\frac{\partial \text{DIC}}{\partial z}  _{z_\infty} = 0$

### 2.3.7 Alkalinity

Alkalinity is produced and consumed when organic matter is degraded. **TODO: Explain why it is produced/consumed, where and how in the different layers**  $j \in \{I, II, III, IV\}$  and talk earlier on 345 in the model description why it matters.

#### 1. Layer ( $z \leq z_{\text{ox}}$ )

$$\frac{\partial \text{ALK}^I}{\partial t} = 0 = D_{\text{ALK}} \frac{\partial^2 \text{ALK}^I}{\partial z^2} - w \frac{\partial \text{ALK}^I}{\partial z} + \frac{1-\phi}{\phi} \cdot \sum_i \left( \text{ALK}^{\text{NIT}} \cdot \frac{\gamma_{\text{NH}_4}}{1 + K_{\text{NH}_4}} \text{NC}_i + \text{ALK}^{\text{OX}} \right) \cdot k_i \cdot C_i(z) \quad (23)$$

**Table 8.** Boundary conditions for alkalinity. For the boundaries we define:  $z_-^- := \lim_{h \rightarrow 0} (z_- - h)$  and  $z_-^+ := \lim_{h \rightarrow 0} (z_- + h)$ .

Boundary	Condition	
$z = 0$	known concentration	1) $\text{ALK}(0) = \text{ALK}_0$
$z = z_{\text{bio}}$	continuity	2) $\text{ALK}(z_{\text{bio}}^-) = \text{ALK}(z_{\text{bio}}^+)$
	flux	3) $-(D_{\text{ALK},0} + D_{\text{bio}}) \cdot \frac{\partial \text{ALK}}{\partial z}  _{z_{\text{bio}}^-} = -D_{\text{ALK},0} \cdot \frac{\partial \text{ALK}}{\partial z}  _{z_{\text{bio}}^+}$
$z = z_{\text{ox}}$	continuity	4) $\text{ALK}(z_{\text{ox}}^-) = \text{ALK}(z_{\text{ox}}^+)$
	flux	5) $-D_{\text{ALK}} \cdot \frac{\partial \text{ALK}}{\partial z}  _{z_{\text{ox}}^-} + F_{\text{ALK}}(z_{\text{ox}}) = -D_{\text{ALK}} \cdot \frac{\partial \text{ALK}}{\partial z}  _{z_{\text{ox}}^+}$ $F_{\text{ALK}}(z_{\text{ox}}) = \frac{1-\phi}{\phi} \cdot \left( \text{ALK}^{\text{H}_2\text{S}} \cdot \gamma_{\text{H}_2\text{S}} \int_{z_{\text{NO}_3}}^{\text{SO}_4} \sum_i \text{SO}_4 \text{C} \cdot k_i \cdot C_i dz \right)$ $+ \frac{1-\phi}{\phi} \cdot \left( \text{ALK}^{\text{NIT}} \frac{\gamma_{\text{NH}_4}}{1+k_{\text{NH}_4}} \int_{z_{\text{NO}_3}}^{\infty} \sum_i \text{NC}_i \cdot k_i \cdot C_i dz \right)$
	where:	
$z = z_{\text{NO}_3}$	continuity	6) $\text{ALK}(z_{\text{NO}_3}^-) = \text{ALK}(z_{\text{NO}_3}^+)$
	flux	7) $-D_{\text{ALK}} \cdot \frac{\partial \text{ALK}}{\partial z}  _{z_{\text{NO}_3}^-} = -D_{\text{ALK}} \cdot \frac{\partial \text{ALK}}{\partial z}  _{z_{\text{NO}_3}^+}$
$z = z_{\text{SO}_4}$	continuity	8) $\text{ALK}(z_{\text{SO}_4}^-) = \text{ALK}(z_{\text{SO}_4}^+)$
	flux (with AOM)	9) $-D_{\text{ALK}} \cdot \frac{\partial \text{ALK}}{\partial z}  _{z_{\text{SO}_4}^-} + F_{\text{ALK}}(z_{\text{SO}_4}) = -D_{\text{ALK}} \cdot \frac{\partial \text{ALK}}{\partial z}  _{z_{\text{SO}_4}^+}$ $F_{\text{ALK}}(z_{\text{SO}_4}) = \frac{1-\phi}{\phi} \cdot \left( \text{ALK}^{\text{AOM}} \gamma_{\text{CH}_4} \cdot \int_{z_{\text{SO}_4}}^{\infty} \sum_i k_i \cdot C_i dz \right)$
	where:	
$z = z_{\infty}$	zero ALK flux	10) $\frac{\partial \text{ALK}}{\partial z}  _{z_{\infty}} = 0$

2. Layer ( $z_{\text{ox}} < z \leq z_{\text{NO}_3}$ )

$$\frac{\partial \text{ALK}^{II}}{\partial t} = 0 = D_{\text{ALK}} \frac{\partial^2 \text{ALK}^{II}}{\partial z^2} - w \frac{\partial \text{ALK}^{II}}{\partial z} + \frac{1-\phi}{\phi} \cdot \sum_i \text{ALK}^{\text{DEN}} \cdot k_i \cdot C_i(z) \quad (24)$$

3. Layer ( $z_{\text{NO}_3} < z \leq z_{\text{SO}_4}$ )

$$355 \quad \frac{\partial \text{ALK}^{III}}{\partial t} = 0 = D_{\text{ALK}} \frac{\partial^2 \text{ALK}^{III}}{\partial z^2} - w \frac{\partial \text{ALK}^{III}}{\partial z} + \frac{1-\phi}{\phi} \cdot \sum_i \text{ALK}^{\text{SUL}} \cdot k_i \cdot C_i(z) \quad (25)$$

4. Layer ( $z_{\text{SO}_4} < z \leq z_{\infty}$ )

$$\frac{\partial \text{ALK}^{IV}}{\partial t} = 0 = D_{\text{ALK}} \frac{\partial^2 \text{ALK}^{IV}}{\partial z^2} - w \frac{\partial \text{ALK}^{IV}}{\partial z} + \frac{1-\phi}{\phi} \cdot \sum_i \text{ALK}^{\text{MET}} \cdot k_i \cdot C_i(z) \quad (26)$$

360 To solve equations 23 and 26 the model assumes the boundary conditions summarised in Table 8.

## 2.4 Model Parameters

This section describes the parameters used in OMEN-SED (1.0) to describe sediment transport and biogeochemical reactions related to the burial and mineralization of organic matter under a wide range of environmental conditions. Table 9 states the parameters for sediment characteristics and

365 Table 10 summarizes the stoichiometric factors and secondary reaction parameters used in the model.

## 2.4.1 Transport Parameters

Advection is the bulk flow of sediments and can be directly related to the accumulation of new material on the seafloor (i.e. sedimentation, Burdige, 2006). This results in a downward flux of older sediment material and porewater in relation to the sediment-water interface. When coupled to an ocean model, its sedimentation flux can be readily used in OMEN-SED 1.0. The stand-alone version of OMEN-SED 1.0 uses the empirical global relationship between sediment accumulation rate ( $\text{cm yr}^{-1}$ ) and seafloor depth (m) of Middelburg et al. (1997):

$$w = 3.3 \cdot 10^{-0.87478367 - 0.00043512 \cdot \text{depth}}, \quad (27)$$

As discussed before (Sec. 2.2), the diffusion coefficient of species  $i$  is calculated as  $D_i = D_{i,0} + D_{\text{bio}} = D_{\text{mol},i} \cdot f_{ir} + D_{\text{bio}}$  for dissolved species and  $D_i = D_{\text{bio}}$  for solid species. The bioturbation coefficient  $D_{\text{bio}}$  ( $\text{cm}^2 \text{ yr}^{-1}$ ) is constant in the bioturbated zone and also follows the empirical relationship by Middelburg et al. (1997):

$$D_{\text{bio}} = 5.2 \cdot 10^{0.76241122 - 0.00039724 \cdot \text{depth}} \quad (28)$$

Studies showed that bioturbational effects on a global scale are largely restricted to the upper 10 cm of the sediments and are only marginally related to seafloor depth (e.g. Boudreau, 1998; Teal et al., 2010). Therefore, OMEN-SED (1.0) imposes a globally invariant bioturbation depth of 10 cm. Bioirrigation can enhance the molecular diffusion coefficient  $D_{i,0} = D_{\text{mol},i} \cdot f_{ir}$  (Soetaert et al., 1996a). However, here we do not consider this effect and set  $f_{ir}$  to a constant value of 1. The specific molecular diffusion coefficients  $D_{\text{mol},i}$  are corrected for sediment porosity  $\phi$ , tortuosity  $F$  and are linearly interpolated for an ambient temperature  $T$  using zero-degree coefficients  $D_i^0$  and temperature dependent diffusion coefficients  $D_i^T$  (compare Gypens et al., 2008):

$$D_{\text{mol},i} = (D_i^0 + D_i^T \cdot T) \cdot \frac{1}{\phi \cdot F}$$

Tortuosity can be expressed in terms of porosity as  $F = \frac{1}{\phi^m}$  (Ullman and Aller, 1982) with the exponent  $m$  varying according to the type of sediment (here we use  $m=3$ ). Values for  $D_i^T$  and  $D_i^0$  are summarized in Table 9 and are adapted from Li and Gregory (1974) and Gypens et al. (2008).

## 2.4.2 Reaction Parameters

The applied multi-G approach for organic matter degradation considers specific degradation rate constants  $k_i$  ( $\text{yr}^{-1}$ ) for each compound class. **TODO:** *The degradation constants are generally taken from the coupled Earth System model* and are assumed invariant along the sediment column, therefore independent of the nature of the terminal electron acceptor. The stoichiometry of organic matter is represented by the factors  $\text{NC}_i$  and  $\text{PC}_i$  denoting the molecular nitrogen to carbon and phosphorus to carbon ratio. In the sulfidic and methanic zone the reduction of 1 mol organic matter

**Table 9.** Sediment characteristics and transport parameters. **TODO: Update table!**

Parameter	Unit	Value	Description/Source
$\rho_{\text{sed}}$	$\text{g cm}^{-3}$	2.6	Sediment density
$w$	$\text{cm yr}^{-1}$	From ESM  (Middelburg et al., 1997)	Advection/Sediment accumulation rate
$z_{\text{bio}}$	cm	10 or 0.0	Bioturbation depth  (Boudreau, 1998; Teal et al., 2010)
$D_{\text{bio}}$	$\text{cm}^2 \text{yr}^{-1}$	Fct. of seafloor  depth	Bioturbation coefficient  (Middelburg et al., 1997)
$\phi$	-	0.85	Porosity
$F$	-	$\frac{1}{\phi^m}$	Tortuosity, here m=3
$f_{ir}$	-	1	Irrigation factor
$\text{PO}_4^s$	$\text{mol cm}^{-3}$	$1 \cdot 10^{-9}$	equilibrium conc. for P sorption  (Slomp et al., 1996)
$\text{PO}_4^a$	$\text{mol cm}^{-3}$	$3.7 \cdot 10^{-9}$	equilibrium conc. for authigenic P precipitation  (Slomp et al., 1996)
$M^\infty$	$\text{mol cm}^{-3}$	$1.99 \cdot 10^{-9}$	asymptotic concentration for Fe-bound P  (Slomp et al., 1996)
<b>Adsorption coefficients and rate constants</b>			
$K_{\text{NH}_4}$	-	1.4	NH <sub>4</sub> adsorption coefficient (Wang and Van Cappellen, 1996)
$K_{\text{PO}_4}^I$	-	200.0	PO <sub>4</sub> adsorption coefficient (oxic) (Slomp et al., 1998)
$K_{\text{PO}_4}^{II}$	-	1.3	PO <sub>4</sub> adsorption coefficient (anoxic) (Slomp et al., 1998)
$k_s$	$\text{yr}^{-1}$	???	Rate constant for P sorption
$k_m$	$\text{yr}^{-1}$	???	Rate constant for Fe-bound P release
$k_a$	$\text{yr}^{-1}$	???	Rate constant for authigenic CFA precipitation
<b>Diffusion coefficients</b> (Li and Gregory, 1974; Gypens et al., 2008)			
$D_{\text{O}_2}^0$	$\text{cm}^2 \text{yr}^{-1}$	348.62172	Molecular diffusion coefficient of oxygen at 0°C
$D_{\text{O}_2}^T$	$\text{cm}^2 \text{yr}^{-1} \text{°C}^{-1}$	14.08608	Diffusion coefficient for linear temp. dependence of oxygen
$D_{\text{NO}_3}^0$	$\text{cm}^2 \text{yr}^{-1}$	308.42208	Molecular diffusion coefficient of nitrate at 0°C
$D_{\text{NO}_3}^T$	$\text{cm}^2 \text{yr}^{-1} \text{°C}^{-1}$	12.2640	Diffusion coefficient for linear temp. dependence of nitrate
$D_{\text{NH}_4}^0$	$\text{cm}^2 \text{yr}^{-1}$	308.42208	Molecular diffusion coefficient of ammonium at 0°C
$D_{\text{NH}_4}^T$	$\text{cm}^2 \text{yr}^{-1} \text{°C}^{-1}$	12.2640	Diffusion coefficient for linear temp. dependence of ammonium
$D_{\text{SO}_4}^0$	$\text{cm}^2 \text{yr}^{-1}$	157.68	Molecular diffusion coefficient of sulfate at 0°C
$D_{\text{SO}_4}^T$	$\text{cm}^2 \text{yr}^{-1} \text{°C}^{-1}$	7.884	Diffusion coefficient for linear temp. dependence of sulfate
$D_{\text{H}_2\text{S}}^0$	$\text{cm}^2 \text{yr}^{-1}$	307.476	Molecular diffusion coefficient of sulfide at 0°C
$D_{\text{H}_2\text{S}}^T$	$\text{cm}^2 \text{yr}^{-1} \text{°C}^{-1}$	9.636	Diffusion coefficient for linear temp. dependence of sulfide
$D_{\text{PO}_4}^0$	$\text{cm}^2 \text{yr}^{-1}$	112.90764	Molecular diffusion coefficient of phosphate at 0°C
$D_{\text{PO}_4}^T$	$\text{cm}^2 \text{yr}^{-1} \text{°C}^{-1}$	5.586252	Diffusion coefficient for linear temp. dependence of phosphate

**Table 10.** Values for biogeochemical parameters used in OMEN-SED (1.0). The variables  $x$ ,  $y$  and  $z$  denote the atomic ratio of carbon, nitrogen and phosphorous of the degrading organic matter (here set to  $C : N : P = 106 : 16 : 1$ ).

Parameter/Variable	Unit	Value	Description
<b>Stoichiometric factors and molecular ratios</b>			
NC <sub>1</sub>	mol/mol	$\frac{y}{x} = \frac{16}{106}$	nitrogen to carbon ratio  refractory fraction, was 2 different ones? why?
NC <sub>2</sub>	mol/mol	$\frac{y}{x} = \frac{16}{106}$	nitrogen to carbon ratio  labile fraction
PC <sub>i</sub>	mol/mol	$\frac{z}{x} = \frac{1}{106}$	phosphorus to carbon ratio
MC	mol/mol	0.5	methane to carbon ratio  produced during methanogenesis
DICC <sup>I</sup>	mol/mol	1.0	DIC to carbon ratio until $z_{SO_4}$
DICC <sup>II</sup>	mol/mol	0.5	DIC to carbon ratio below $z_{SO_4}$
O <sub>2</sub> C	mol/mol	$\frac{x+2y}{x} = \frac{138}{106}$	oxygen to carbon ratio
NO <sub>3</sub> C	mol/mol	$\frac{4x+3y}{5x} = \frac{94.4}{106}$	nitrate to carbon ratio
SO <sub>4</sub> C	mol/mol	$\frac{1}{2}O_2C = \frac{138}{212}$	sulfate to carbon ratio
ALK <sup>OX</sup>	mol/mol	$-\frac{y+2z}{x} = -\frac{18}{106}$	ALK from aerobic degradation
ALK <sup>NIT</sup>	mol/mol	-2	ALK from nitrification
ALK <sup>DEN</sup>	mol/mol	$\frac{4x+3y-10z}{5x} = \frac{92.4}{106}$	ALK from denitrification
ALK <sup>SUL</sup>	mol/mol	$\frac{x+y-2z}{x} = \frac{120}{106}$	ALK from sulfate reduction
ALK <sup>MET</sup>	mol/mol	$\frac{y-2z}{x} = \frac{14}{106}$	ALK from methanogenesis
ALK <sup>H2S</sup>	mol/mol	-2	ALK from H <sub>2</sub> S oxidation
ALK <sup>AOM</sup>	mol/mol	2	ALK from AOM
<b>Secondary reaction parameters</b>			
$\gamma_{NH_4}$	-	0.9	fraction of NH <sub>4</sub> that is oxidised in oxic layer
$\gamma_{H_2S}$	-	0.95	fraction of H <sub>2</sub> S that is oxidised in oxic layer
$\gamma_{CH_4}$	-	0.99	fraction of CH <sub>4</sub> that is oxidised at $z_{SO_4}$

additionally produces  $\frac{138}{212}$  mol of hydrogen sulfide (SO<sub>4</sub>C) and 0.5 mol of methane (MC). In the total sediment column organic matter mineralization consumes the specific TEA with a fixed ratio  
400 (O<sub>2</sub>C, NO<sub>3</sub>C and SO<sub>4</sub>C respectively). See Table 10 for a complete summary of the parameters and their values.

## 2.5 Module Structure

**TODO:** An analytical steady-state solution is found for the reaction-transport equation of each chemical species in each layer. At each boundary (i.e.  $z_{\text{ox}}$ ,  $z_{\text{bio}}$ ,  $z_{\text{NO}_3}$  and  $z_{\text{SO}_4}$ ) the model has to

405 match continuity and flux for different ODE solutions of the layer above and below the specific boundary. In particular the bioturbation boundary is problematic as it can theoretically occur in any geochemical layer (compare Fig. 3). In order to simplify this recurring boundary matching problem it is implemented in an independent algorithm which is described in Section 2.5.1. Instructions and requirements for coupling OMEN-SED (1.0) to a global Earth System Model are given in Section

410 2.5.2.

### 2.5.1 Generic boundary condition matching (GBCM)

A general steady-state advection-diffusion-reaction (ADR) diagenetic equation looks like:

$$\frac{\partial C}{\partial t} = 0 = D \frac{\partial^2 C}{\partial z^2} - w \frac{\partial C}{\partial z} - \sum_i \alpha_i \exp(-\beta_i z) - k \cdot C + Q. \quad (29)$$

415 where  $z$  is the sediment depth,  $t$  the time,  $D$  is the diffusion coefficient and  $w$  is the advection rate.

The ODE solution is of the general form:

$$C(z) = A \exp(az) + B \exp(bz) + \sum_i \frac{\alpha_i}{D\beta_i^2 - w\beta_i - k} \cdot \exp(-\beta_i z) + \frac{Q}{k} \quad (30)$$

and can therefore be expressed as:

420  $C(z) = A \cdot E(z) + B \cdot F(z) + G(z) \quad (31)$

where  $E(z)$ ,  $F(z)$  are the homogeneous solutions of the ODE,  $G(z)$  the particular integral, and  $A$ ,  $B$  are the integration constants (compare Fig. 3 for the whole sediment column).

Each boundary matching problem involves matching continuity and flux for the two solutions  $C_U(z)$

425 ( $=$  'upper') and  $C_L(z)$  ( $=$  'lower') across a boundary at  $z = z_b$ . Therefore, we get two ODE solutions of the genral form:

$$C_U(z) = A_U \cdot E_U(z) + B_U \cdot F_U(z) + G_U(z) \quad (32)$$

$$C_L(z) = A_L \cdot E_L(z) + B_L \cdot F_L(z) + G_L(z). \quad (33)$$

430 The two boundary conditions are: for continuity (where for generality we allow a discontinuity  $V_b$ )

$$C_U(z_b) = C_L(z_b) + V_b \quad (34)$$

and for flux

435  $D_U C'_U(z_b) + w C_U(z_b) = D_L C'_L(z_b) + w C_L(z_b) + F_b \quad (35)$

where  $w$  is advection,  $D$  are the diffusion coefficients and  $F_b$  is any flux discontinuity.

In terms of the ODE solutions (32), (33), the boundary conditions represent two equations connecting the four integration constants:

$$440 \quad \begin{pmatrix} E_U & F_U \\ D_U E'_U & D_U F'_U \end{pmatrix} \begin{pmatrix} A_U \\ B_U \end{pmatrix} = \begin{pmatrix} E_L & F_L \\ D_L E'_L & D_L F'_L \end{pmatrix} \begin{pmatrix} A_L \\ B_L \end{pmatrix} + \begin{pmatrix} G_L - G_U + V_b \\ D_L G'_L - D_U G'_U + F_b - wV_b \end{pmatrix} \quad (36)$$

where the ODE solutions  $E$ ,  $F$ ,  $G$  are all evaluated at  $z_b$ .

Equation (36) can be solved to give  $A_U$  and  $B_U$  as a function of the integration constants from the layer below ( $A_L$  and  $B_L$ ), thereby constructing a piecewise solution for the whole region, with now just two integration constants  $A_L$  and  $B_L$ .

445 In the code the function **benthic\_utils.matchsoln** provides this solution in the form:

$$\begin{pmatrix} A_U \\ B_U \end{pmatrix} = \begin{pmatrix} c_1 & c_2 \\ c_3 & c_4 \end{pmatrix} \begin{pmatrix} A_L \\ B_L \end{pmatrix} + \begin{pmatrix} d_1 \\ d_2 \end{pmatrix}. \quad (37)$$

Using (37) we can now rewrite  $C_U(z)$  in (32) as a function of  $A_L$  and  $B_L$ :

$$C_U(z) = (c_1 A_L + c_2 B_L + d_1) \cdot E_U(z) + (c_3 A_L + c_4 B_L + d_2) \cdot F_U(z) + G_U(z)$$

and hence define the “transformed” basis functions  $E_U^*(z)$ ,  $F_U^*(z)$ ,  $G_U^*(z)$  such that:

$$450 \quad C_U(z) = A_L \cdot E_U^*(z) + B_L \cdot F_U^*(z) + G_U^*(z) \quad (38)$$

where

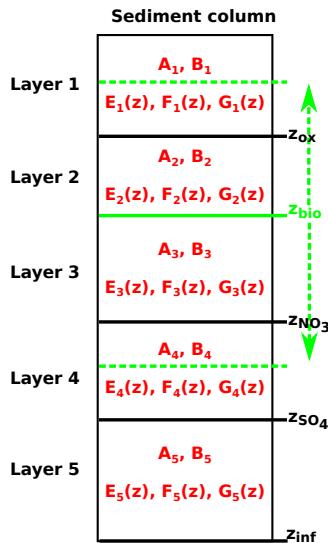
$$\begin{aligned} E_U^*(z) &= c_1 E_U(z) + c_3 F_U(z) \\ F_U^*(z) &= c_2 E_U(z) + c_4 F_U(z) \\ 455 \quad G_U^*(z) &= G_U(z) + d_1 E_U(z) + d_2 F_U(z) \end{aligned} \quad (39)$$

(in the code this is done by **benthic\_utils.xformsoln**).

### Solving the sediment layer stack

Equations (37), (38) and (39) can now be applied for each layer boundary, working up from the bottom of the sediments. The net result is to construct a piecewise solution with just two integration

460 constants (coming from the lowest layer), which can then be solved for by applying one boundary condition for the sediment-water interface and one for the bottom of the sediments (e.g. a concentration condition at the bottom of the sediments, and a flux condition at the SWI).

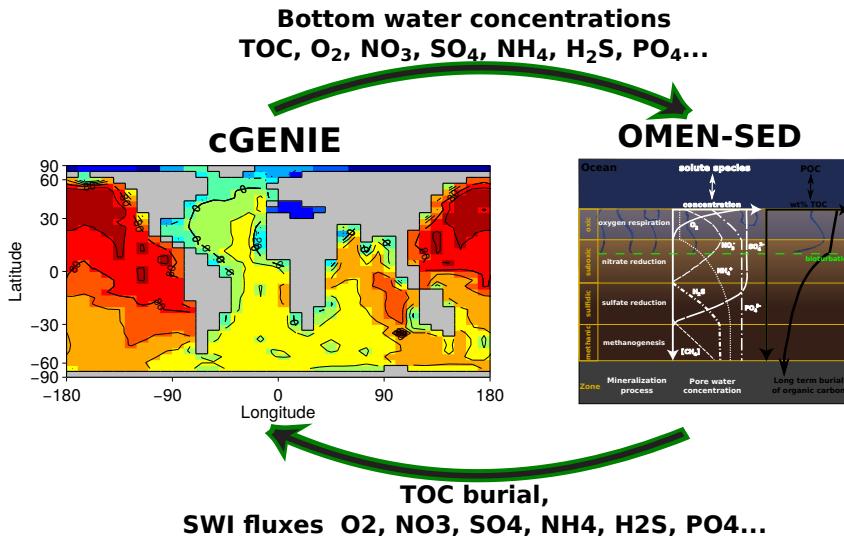


**Figure 3.** Schematic of the generic boundary condition matching (GBCM) problem. Showing the resulting integration constants ( $A_i, B_i$ ) and ODE solutions ( $E_i, F_i, G_i$ ) for the different sediment layers and the variable bioturbation boundary.

### Abstracting out the bioturbation boundary

The bioturbation boundary affects the diffusion coefficient of the modelled solutes and the conservation equation of organic matter which is available for mineralization. The boundary is particularly inconvenient as it can in principle occur in the middle of any “geochemical” layer and therefore generates multiple cases (compare Fig. 3). To simplify this for solutes, the “piecewise solution construction” above is used to abstract out the bioturbation boundary. An initial test for each layer is made to identify its “bioturbation-status” (fully bioturbated, fully non-bioturbated or crossing the bioturbation boundary) and (if needed) a piecewise solution is constructed by matching boundary conditions across the bioturbation boundary. The “outside” code therefore never needs to know whether it is dealing with a piecewise solution (i.e. matched across a bioturbation boundary) or a “simple” solution (i.e. the layer is fully bioturbated or fully non-bioturbated).

In the code, this is performed by **zTOC.prepfg\_I12** which hands back a structure **ls** containing the “bioturbation-status” for each layer and (if needed) the description of the piecewise solution (coefficients  $c_1, c_2, c_3, c_4, d_1, d_2$  as above). So e.g. for sulfate, **zTOC.prepfg\_I12** is called three times at the beginning of **zSO4.calcbc** (one for each “geochemical” layer: oxic, suboxic, sulfidic) handing back three structures **ls** describing the layer’s “bioturbation-status”, abstracting away the bioturbation boundary and all associated conditional logic. When calculating the solutions for the different



**Figure 4.** Schematic of the relationship between OMEN-SED (1.0) and GENIE. Arrows and accompanied text represent the information transferred between models.

480 layers, the pre-calculated structure `ls` is passed to the function `zTOC.calcfg_l12` which sorts out the correct solution type to use.

## 2.5.2 Coupling to an Earth System Model

485 **TODO: Now: use flux, not POC-concentration.** Here the coupling of OMEN-SED (1.0) to the GENIE Earth System model is described, in particular its incorporation into the sediment geochemistry model (SEDGEM, Ridgwell and Hargreaves, 2007). Results from pre-industrial experiments are presented in section 3.4. OMEN-SED (1.0) is called by SEDGEM for each grid point instead of remineralising the settling flux of OM at the seafloor, passing over the bottom water concentrations of solutes (in  $\frac{mol}{kg}$ ) and the flux of OM (POC<sub>flux</sub>, in  $\frac{cm^3}{cm^2 \cdot yr}$ , see Fig. 4). To be used by OMEN-SED (1.0) dissolved species and OM flux need to be converted into  $\frac{mol}{cm^3}$ . For the solutes this is trivial; the 490 OM flux is converted into a concentration (POC<sub>con</sub>, in  $\frac{mol}{cm^3}$ ) using the sediment accumulation rate from GENIE ( $w$ , in  $\frac{cm}{yr}$ ) as:

$$POC_{con} = \frac{1}{12} \cdot (1 - \phi) \cdot \frac{POC_{flux}}{w} \quad (40)$$

where  $\phi$  is the sediment porosity (-) and  $\frac{1}{12}$  is the conversion of POC from  $cm^3$  to  $mol$  assuming the average density of POC as  $1.0 \frac{cm^3}{g}$ . Other parameters needed from GENIE are the seafloor depth, 495 the local temperature and the partitioning of bulk OM into the slower and faster degrading pool (and a third unreactive pool if modelled). OMEN-SED (1.0) then computes the resulting SWI-fluxes of solutes (in  $\frac{mol}{cm^2 \cdot yr^{-1}}$ ) and the fraction of POC preserved in the sediments at a depth of 100 cm which are given back to GENIE. In order to reduce memory requirements the sediment profiles (e.g.

as shown in Fig. 5) are by default not calculated by OMEN-SED (1.0), however, the information  
500 needed to do this once the experiment has terminated is passed over and saved in SEDGEM (i.e.  
the integration constants from the lowest layer  $A_L$ ,  $B_L$  and the ODE solutions  $E_i$ ,  $F_i$ ,  $G_i$  for the  
different sediment layers  $i$ ). **TODO:** check that is correct... e.g. need all ODE solutions but just IC  
from lowest layer? - why not calculated like this in matlab when calc. the profiles...?

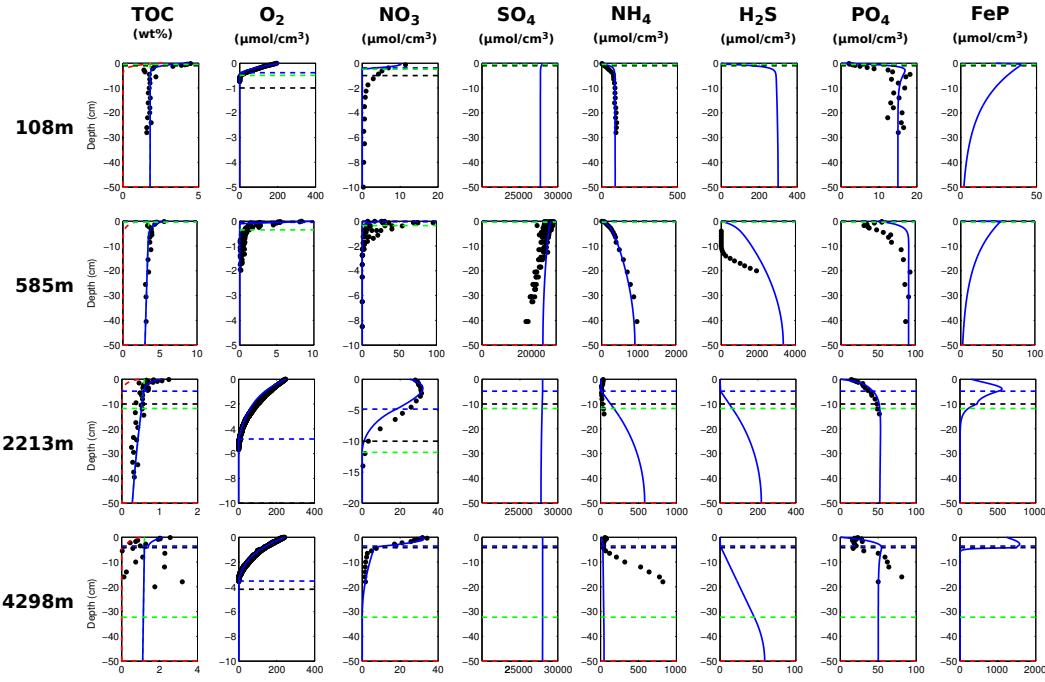
### 3 Test Cases

#### 505 3.1 Sediment profiles

Figure 5 compares profiles calculated with OMEN-SED (1.0) to observations from different ocean depths taken at the Santa Barbara basin (585 m, Reimers et al., 1996), the Iberian margin (108 and 2213 m) and Nazaé Canyon (4298 m, Epping et al., 2002). Sediment-water interface characteristics and concentrations of total organic matter and all dissolved species in OMEN-SED (1.0) were set  
510 to the observed values where available (Table 11). The OM and pore water profiles were fitted by optimizing the OM partitioning into the fast and slow degrading pool and their respective first order degradation rate constants (priority given to reproduce the OM and  $O_2$  profiles). For phosphorus the equilibrium concentration for authigenic P formation ( $PO_4^a$ ) was adjusted to fit the  $PO_4$  concentration at  $z_\infty$ . For the two open Iberian margin stations (108 and 2213 m) OMEN-SED (1.0) fits all  
515 observations well. OMEN-SED (1.0) does especially well at depth 2213m by reproducing the deep  $O_2$  penetration and the subsurface maximum in  $NO_3$  concentration due to the nitrification of  $NH_4$ . For the Santa Barbara basin (585 m) a misfit is observed for  $H_2S$  and  $PO_4$  in the upper 20 cm of the sediment. This can be explained by the presence of  $Mn^{2+}$ ,  $Fe^{2+}$  and dissolved Fe at this site which are either reduced to degrade OM and/or react with  $H_2S$  to form iron sulfides, therefore inhibiting  
520 the rise in concentration of  $H_2S$  (Reimers et al., 1996). Phosphorus is adsorbed to Fe oxides and incorporated into carbonate fluorapatite (CFA) which is not modelled explicitly in OMEN-SED (1.0).  
**above correct/okay....?** For the Nazaé Canyon station (4298 m) satisfactory fits could be realised apart from  $NH_4$ . However, also the original study (Epping et al., 2002) had the same problem using a more complex diagenetic model and suggested non-local solute exchange being responsible for the  
525 higher observed  $NH_4$  concentrations.

#### 3.2 Sensitivity Analysis

Model parameters implicitly account for processes not explicitly described, they are notoriously difficult to constrain and a source of uncertainty for a numerical model. Sensitivity Analysis (SA) is a term used for mathematical techniques to investigate how the variations in the outputs ( $y_1$ , ...,  $y_N$ )  
530 of a numerical model can be attributed to variations in the different input parameters ( $x_1$ , ...,  $x_M$ ) (Pianosi et al., 2016). Different types of sensitivity indices, which quantify this relative influence with a scalar  $S_i$ , can be calculated, ranging from simple one-at-a-time methods to statistical eval-



**Figure 5.** Modelled (curves) and measured (filled dots) dissolved and solid phase pore water profiles for four different sediment cores. Note that different scales are used for different stations. The blue TOC curve represents the sum of the refractory (green) and labile (red) TOC fraction.

**Table 11.** Model boundary conditions for the sampling stations in Figure 5.

Sediment characteristics:							
Depth (m)	Temp. (°C)	$z_{\text{bio}}$ (cm)	$D_{\text{bio}}$ ( $\text{cm}^2 \text{yr}^{-1}$ )	TOC <sub>1</sub> (wt%)	TOC <sub>2</sub> (wt%)	$k_1$ ( $\text{yr}^{-1}$ )	$k_2$ ( $\text{yr}^{-1}$ )
108	12.5	1.0	0.02	2.64	1.8	0.65	$1.0e^{-5}$
585	5.85	0.01	0.02	2.0	3.5	0.2	$8.0e^{-4}$
2213	3.2	10.0	0.17	0.45	0.6	0.15	$9.0e^{-4}$
4298	2.5	4.2	0.18	1.0	1.2	0.055	$1e^{-5}$

#### Bottom water concentrations of solutes (all in $\mu\text{mol cm}^{-3}$ ):

Depth	O <sub>2</sub>	NO <sub>3</sub>	SO <sub>4</sub>	NH <sub>4</sub>	H <sub>2</sub> S	PO <sub>4</sub>	PO <sub>4</sub> <sup>a</sup>
108	210.0	9.6	28,000	0.4	0.0	0.0	15.0
585	10.0	25.0	28,000	0.0	0.0	50.0	90.0
2213	250	25.0	28,000	0.6	0.0	0.0	5.0
4298	243	30.1	28,000	0.22	0.0	5.0	-

535 uations of the output distribution (e.g. variance-based or density-based approaches Pianosi et al.,  
 2016). Especially the latter indices are easy to interpret and to compare across different parameters  
 and/or different model outputs as they generally take values between zero and one ( $S_i \in [0, 1]$ ). An  
 index of zero indicates a non-influential parameter and a higher value a more influential parameter.  
 Here, we use SA mainly to identify which parameters have the largest impact on the different  
 model outputs and therefore require careful calibration. As the probability density functions of our  
 model outputs (i.e. the resulting SWI-fluxes) are generally highly-skewed towards extreme organic  
 540 matter degradation rates (not shown) variance-based sensitivity indices are not very reliable uncer-  
 tainty indicators (Pianosi et al., 2016). Rather than just considering the variance we employ the novel  
 density-based PAWN method by (Pianosi and Wagener, 2015) which considers the entire conditional  
 and unconditional Cumulative Distribution Function (CDF) of the model output. The sensitivity in-  
 dex of parameter  $i$  is calculated as the difference between the two CDFs, i.e.

545 
$$S_i = \max_{x_i} \max_y |F_y(y) - F_{y|x_i}(y)| \quad (41)$$

where  $F_y(y)$  is the unconditional CDF of the output  $y$  and  $F_{y|x_i}(y)$  is the conditional CDF when the  
 $i$ -th parameter is fixed to  $x_i$ . For a more detailed description of the method we refer the interested  
 reader to Pianosi and Wagener (2015). Due to the model complexity it is impossible to compute the  
 sensitivity indices analytically therefore they are approximated from a Latin-Hypercube sampling of  
 550 parameter inputs and calculated outputs. We use PAWN as implemented within the Sensitivity Anal-  
 ysis for Everyone (SAFE) matlab toolbox (Pianosi et al., 2015) to investigate 11 model parameters  
 for ranges specified in Table 12. We calculate sensitivity indices for all resulting SWI-fluxes from  
 OMEN-SED for two idealised sediment conditions (see Table 13). The resulting indices (compare  
 Fig. 6 for two exemplary cases) are then translated into a color code and summarised in a pattern  
 555 plot to simplify comparison (Fig. 7). The most significant parameters for all model outputs are the  
 degradation rate for the labile OM part ( $k_1$ ) and its share in the total OM pool ( $f_1$ ). For the anoxic  
 setup, where no oxidation occurs, the secondary redox parameters (i.e.  $\gamma_{\text{NH}_4}$ ,  $\gamma_{\text{H}_2\text{S}}$ ) are essentially  
 non-influential, however, they are getting more important for the oxic scenario. The  $\text{PO}_4$  SWI-flux  
 appears to be insensitive under oxic conditions as the majority is absorbed to Fe-oxides correct....?.

560 To test if OMEN-SED (1.0) is able to reproduce observed SWI-fluxes, the Latin-Hypercube sam-  
 pling outputs are compared with a global database of benthic fluxes of  $\text{O}_2$  and  $\text{NO}_3$  (Bohlen et al.,  
 2012). The coloured scatter plots (Figure 8) show that the observed fluxes fall well in the range of  
 SWI-fluxes calculated with OMEN-SED (1.0). Also highlighted by the emergence of colour patterns  
 in Figure 8 A+B are the strong interactions between the amount of labile OM  $f_1$  and its degrada-  
 565 tion rate  $k_1$  for the resulting SWI-flux of the most powerful TEA available.

### 3.3 Benthic fluxes on a global scale

Application to Seitert, 2004 OM, burwiczk see rate data and evaluation based on global data (Archer)

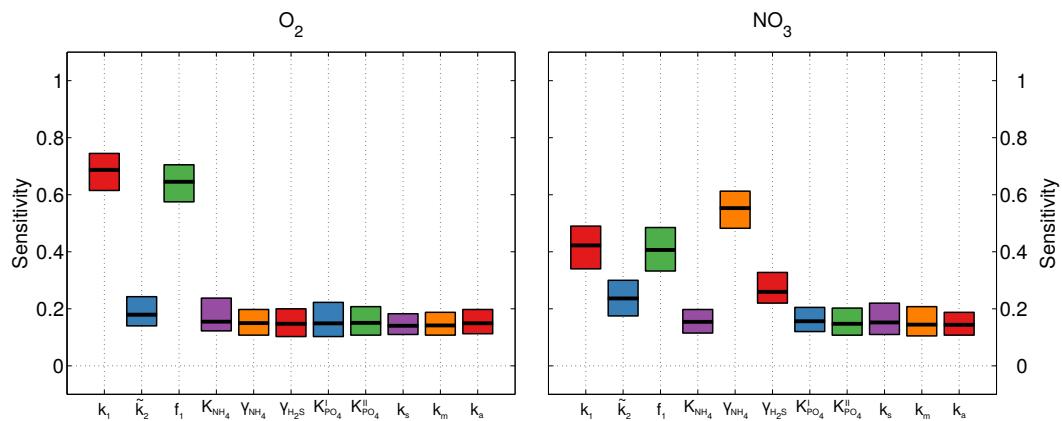
**Table 12.** Range of model parameters used for sensitivity analysis of model predicted output.

Parameter	Description	Units	Minimum	Maximum	Source
$k_1$	labile OM degradation constant	$\text{yr}^{-1}$	$1e^{-4}$	5.0	(1)
$\tilde{k}_2$	order of refractory OM degradation constant ( $k_2 = \tilde{k}_2 \cdot k_1$ )	-	$1e^{-4}$	$1e^{-1}$	(1)
$f_1$	fraction of labile OM	-	0.02	0.98	-
$K_{\text{NH}_4}$	Adsorption coefficient	-	0.8	1.7	(2)
$\gamma_{\text{NH}_4}$	$\text{NH}_4$ fraction oxidised		0.5	1.0	-
$\gamma_{\text{H}_2\text{S}}$	$\text{H}_2\text{S}$ fraction oxidised		0.5	1.0	-
$K_{\text{PO}_4}^I$	Adsorption coeff. oxic	-	100.0	400.0	(3)
$K_{\text{PO}_4}^{II}$	Adsorption coeff. anoxic	-	1.3	2.0	(3)
$k_s$	kinetic P sorption	$\text{yr}^{-1}$	0.1	100.0	(4, 5)
$k_m$	Fe-bound P release	$\text{yr}^{-1}$	0.015	0.02	(4, 5)
$k_a$	authigenic P formation	$\text{yr}^{-1}$	0.001	10.0	(4, 6)

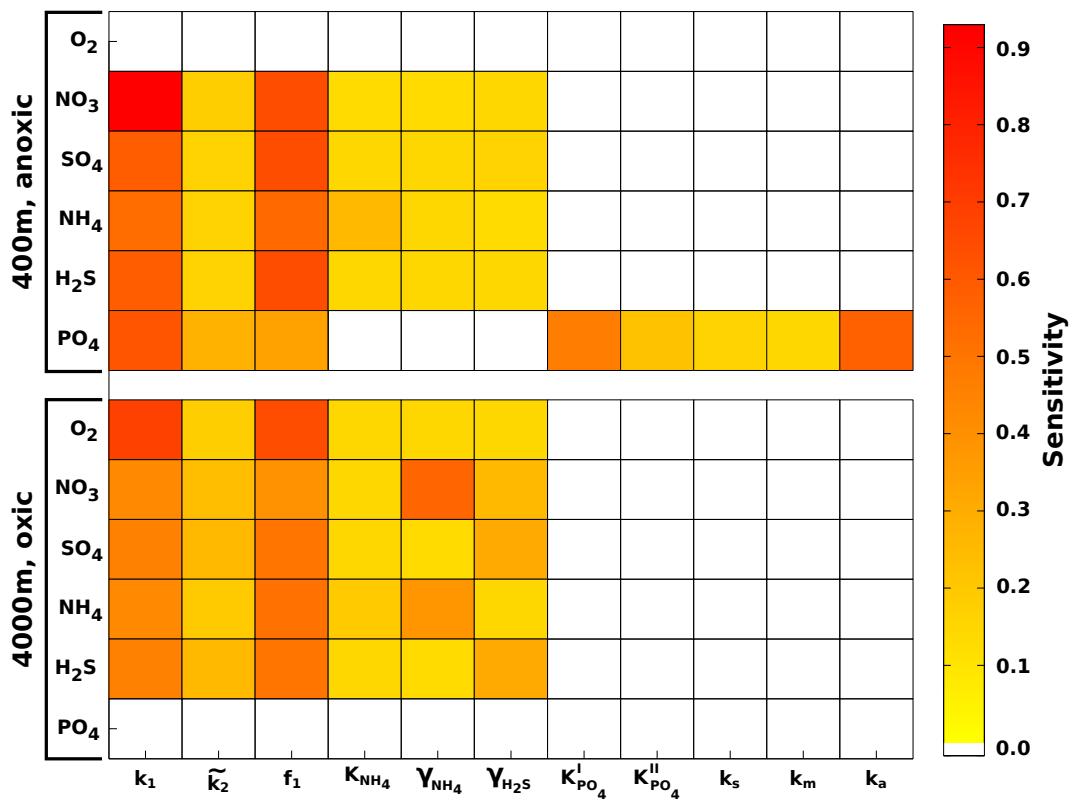
Sources: (1) Arndt et al. (2013); (2): Van Cappellen and Wang (1996); (3): Krom and Berner (1980)  
(4): Gypens et al. (2008); (5): Slomp et al. (1996); (6): Van Cappellen and Berner (1988)

**Table 13.** *Maybe in Suppl.* Model boundary conditions for the two idealised sediment conditions used for the sensitivity analysis (Figure 7). All solute concentrations are in  $\mu\text{mol cm}^{-3}$ .

Depth (m)	Temp. ( $^{\circ}\text{C}$ )	OC (wt%)	$\text{O}_2$	$\text{NO}_3$	$\text{SO}_4$	$\text{PO}_4$	$z_{\text{bio}}$ (cm)
400	8.0	2.0	0.0	40.0	28,000	40.0	0.001
4000	1.5	1.0	300.0	20.0	28,000	40.0	10.0



**Figure 6.** *Maybe in Suppl.* Box plot of parameter sensitivities for the  $\text{O}_2$  and  $\text{NO}_3$  SWI-flux for the 4000m oxic condition. Average sensitivities (black lines) and 90% confidence intervals using  $N = 11200$  model evaluations and  $N_{\text{boot}} = 100$  bootstrap resamples.



**Figure 7.** Pattern plot, showing the output sensitivity for each SWI flux (i.e. the chemical compounds on the horizontal axis) and each input factor (i.e. the model parameters on the vertical axis) for two idealised sediment cores.

### 3.4 GENIE pre-industrial

### 3.5 GENIE-Cretaceous test?

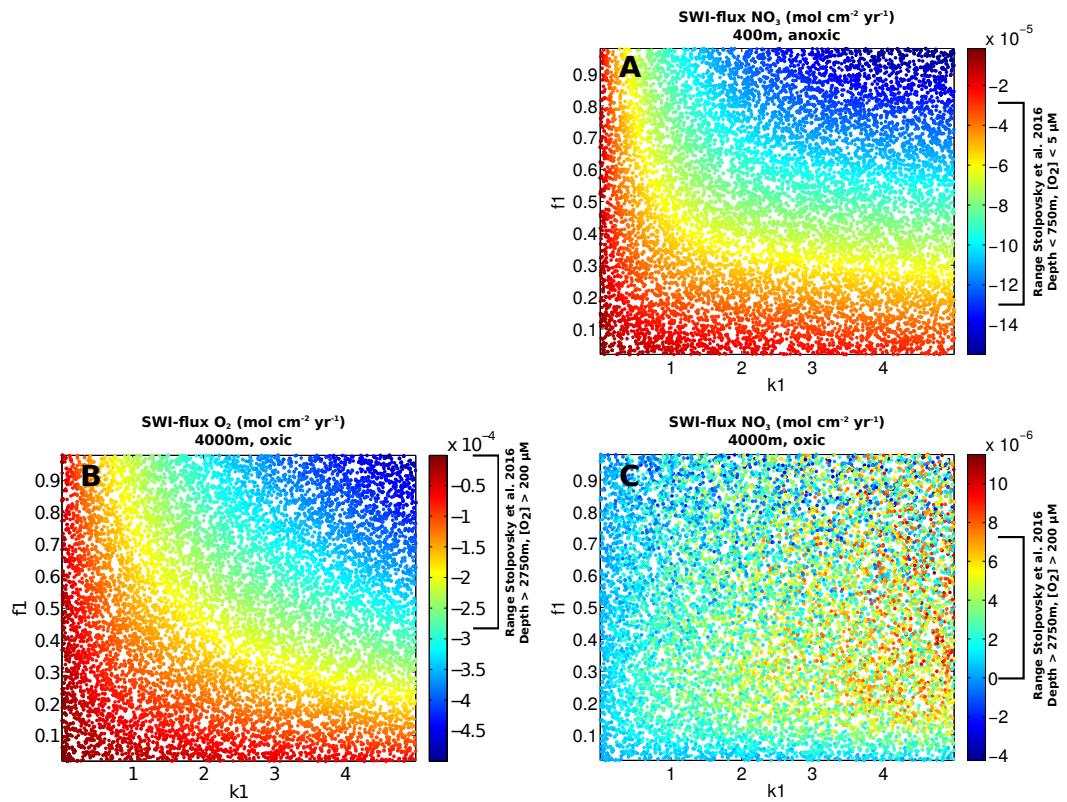
## 570 4 Scope of applicability and model limitations

### 5 Conclusions

TEXT

**Table 14.** Primary pathways of organic matter degradation, secondary redox reactions and stoichiometries implemented in the reaction network.

Pathway	Stoichiometry
Primary Redox reactions	
Aerobic degradation	$(\text{CH}_2\text{O})_x(\text{NH}_3)_y(\text{H}_3\text{PO}_4)_z + (\text{x}+2\text{y})\text{O}_2 + (\text{y}+2\text{z})\text{HCO}_3^- \rightarrow (\text{x}+\text{y}+2\text{z})\text{CO}_2 + \text{yNO}_3^- + \text{zHPO}_4^{2-} + (\text{x}+2\text{y}+2\text{z})\text{H}_2\text{O}$
Denitrification	$(\text{CH}_2\text{O})_x(\text{NH}_3)_y(\text{H}_3\text{PO}_4)_z + \frac{(\text{4x}+3\text{y})}{5}\text{NO}_3^- \rightarrow \frac{(2\text{x}+\text{3y})}{5}\text{N}_2 + \frac{(\text{x}-3\text{y}+10\text{z})}{5}\text{CO}_2 + \frac{(\text{4x}+3\text{y}-10\text{z})}{5}\text{HCO}_3^- + \text{zHPO}_4^{2-} + \frac{(\text{3x}+6\text{y}+10\text{z})}{5}\text{H}_2\text{O}$
Sulfate reduction	$(\text{CH}_2\text{O})_x(\text{NH}_3)_y(\text{H}_3\text{PO}_4)_z + \frac{\text{x}}{2}\text{SO}_4^{2-} + (\text{y}-2\text{z})\text{CO}_2 + (\text{y}-2\text{z})\text{H}_2\text{O} \rightarrow \frac{\text{x}}{2}\text{H}_2\text{S} + (\text{x}+\text{y}-2\text{z})\text{HCO}_3^- + \text{yNH}_4^+ + \text{zHPO}_4^{2-}$
Methanogenesis	$(\text{CH}_2\text{O})_x(\text{NH}_3)_y(\text{H}_3\text{PO}_4)_z + (\text{y}-2\text{z})\text{H}_2\text{O} \rightarrow \frac{\text{x}}{2}\text{CH}_4 + \frac{\text{x}-2\text{y}+4\text{z}}{2}\text{CO}_2 + (\text{x}-2\text{z})\text{HCO}_3^- + \text{yNH}_4^+ + \text{zHPO}_4^{2-}$
Secondary Redox reactions	
Nitrification	$\text{NH}_4^+ + 2\text{O}_2 + 2\text{HCO}_3^- \rightarrow \text{NO}_3^- + 2\text{CO}_2 + 3\text{H}_2\text{O}$
Sulfide oxidation	$\text{H}_2\text{S} + 2\text{O}_2 + 2\text{HCO}_3^- \rightarrow \text{SO}_4^{2-} + 2\text{CO}_2 + 2\text{H}_2\text{O}$
AOM	$\text{CH}_4 + \text{CO}_2 + \text{SO}_4^{2-} \rightarrow 2\text{HCO}_3^- + \text{H}_2\text{S}$
Adsorption reactions and mineral precipitation	
$\text{NH}_4$ adsorption	$\text{NH}_4^+ \xrightarrow{\text{K}_{\text{NH}_4}} \text{NH}_4^+$ (ads)
P ad-/desorption ???	$\text{PO}_4^{2-} \xrightarrow{\text{K}_{\text{PO}_4}^{\text{I},\text{II}}} \text{PO}_4^{2-}$ (ads); $\text{HPO}_4^{2-} \xrightarrow{\text{k}_s} \text{Fe-bound P} \xrightarrow{\text{k}_m} \text{HPO}_4^{2-}$
CFA precipitation	$\text{PO}_4^{2-} \xrightarrow{\text{k}_a} \text{CFA}$



**Figure 8.** Coloured scatter plots ( $k_1$  vs  $f_1$ ) of resulting OMEN-SED (1.0) SWI-fluxes for the 400m anoxic (A:  $\text{NO}_3$ ) and 4000m oxic (B:  $\text{O}_2$ , C:  $\text{NO}_3$ ) scenario. Negative values representing flux from the water column into the sediments. Indicated area at the respective colour scale represents range of benthic fluxes in the global database of Bohlen et al. (2012). **TODO:** Add N and swap colour scale + add text: Low influx, High influx...

## 6 Code Availability

### Appendix A: Reaction Network

575 **A1**

*Acknowledgements.* TEXT

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