

OMEN-SED 1.0: A new, numerically efficient sediment module for the coupling to Earth System Models

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Abstract. Here we describe the first version of a new, analytical early diagenetic model resolving organic matter cycling and associated biogeochemical dynamics in marine sediments called OMEN-SED (Organic Matter ENabled SEDiment model). Most biogeochemical cycles and reactions in the surface sediments can be related either directly or indirectly to the degradation of organic matter.

- 5 Despite its fundamental importance, an appropriate Earth System model of the coupled atmosphere-ocean-sediment system which is able to model all relevant processes and feedbacks over geological time-scales currently does not exist. The major problem is the high computational cost of simulating the essential redox reactions in marine sediments which are important to calculate burial of organic matter and benthic recycling fluxes of chemical compounds. In most Earth System models sediment-
10 water dynamics are either neglected or treated in a very simplistic way. To provide a more realistic description of organic matter degradation and nutrient cycles in marine sediments we have developed OMEN-SED, a new, one-dimensional, numerically efficient reactive transport model. OMEN-SED is the first analytical model to explicitly describe organic matter cycling as well as associated dynamics of the most important terminal electron acceptors (i.e. O₂, NO₃, SO₄), related reduced substances
15 (NH₄, H₂S), the full suite of secondary-redox reactions, macronutrients (PO₄) and associated pore water quantities (ALK, DIC). To represent a redox-dependent sedimentary P cycle we consider the formation and burial of Fe-bound P and authigenic Ca-P minerals. Thus, OMEN-SED captures most of the features of a complex, numerical diagenetic model, however, its computational efficiency allows the coupling to global Earth System models and therefore the investigation of coupled global
20 biogeochemical dynamics over different timescales. This paper provides a detailed description of

the new sediment model, **tested with observations, SA and global observations** and describes it's coupling to the Earth System model cGENIE.

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1 Introduction

DH: Will delete the sub-headings!

Role of marine sediments for climate and global biogeochemical cycles:

- 60 Marine surface sediments are key components in the Earth system. They host the largest carbon reservoir within the surficial Earth system, provide the only long term sink for atmospheric CO₂, recycle nutrients and represent the most important geochemical archive used for deciphering past changes in biogeochemical cycles and climate (e.g. Berner, 1991; Archer and Maier-Reimer, 1994; Ridgwell and Zeebe, 2005; Arndt et al., 2013). Physical and chemical processes in sediments (i.e.
65 diagenetic processes) depend on the water column and vice versa: Diagenesis is mainly donor controlled, as it is fuelled by the external supply of solid material (e.g. organic matter, calcium carbonate, opal) from the water column and is affected by overlying bottom water concentrations of solutes. At the same time, sediments impact the water column directly either by short- and long-term storage of deposited material or diagenesis transforms deposited material and returns some of the resulting
70 products (e.g. nutrients, DIC) to the bottom waters. This so-called benthic-pelagic coupling is essential for understanding global biogeochemical cycles and climate (e.g. Archer and Maier-Reimer, 1994; Archer et al., 2000; Soetaert et al., 2000; Mackenzie, 2005).

Biological primary production of organic matter (OM, CH₂O in equation R1) and the reverse process of degradation can be written in a greatly simplified reaction as:



On geological timescales production of OM is generally greater than degradation which results in some organic matter being buried in marine sediments and oxygen accumulating in the atmosphere. Thus burial of OM leads to net oxygen input to, and CO₂ removal from the atmosphere (Berner, 2004). On shorter timescales, the upper few meters of the sediments (i.e. early diagenesis) are specifically important as it is decided here if a substance is recycled to the water column or buried for a longer period of time in the deeper sediments (Hensen et al., 2006). Most biogeochemical cycles and reactions in this part of marine sediments can be related either directly or indirectly to the degradation of organic matter (e.g Boudreau and Ruddick, 1991; Arndt et al., 2013). Organic matter degradation releases metabolic CO₂ to the pore water, causing it to have a lower pH and provoking
80 the dissolution of CaCO₃ (Emerson and Bender, 1981). Oxygen and nitrate for instance, the most powerful electron acceptors, are consumed in the course of the degradation of organic matter, resulting in the release of ammonium and phosphorous to the pore water. As such, degradation of OM in the sediments can profoundly affect the oxygen and nutrient inventory of the ocean and thus primary productivity (Van Cappellen and Ingall, 1994; Lenton and Watson, 2000).

90 Nutrient recycling from marine sediments has been suggested to play a key role for climate and ocean biogeochemistry for different events during Earth history. For example, feedbacks between phosphorous storage and erosion from shelf sediments and marine productivity have been hypothesised to play an important role for glacial/interglacial atmospheric CO₂ changes (Broecker, 1982;

Ruttenberg, 1993). Furthermore, nutrient recycling from anoxic sediments has been invoked to explain the occurrence of more extreme events in Earth history, for instance Oceanic Anoxic Events (OAEs, e.g. Mort et al., 2007; Tsandev and Slomp, 2009). OAEs represent severe disturbances of the global carbon, oxygen and nutrient cycles of the ocean and are usually characterized by widespread bottom water anoxia and photic zone euxinia (Jenkyns, 2010). One way to explain the genesis and persistence of OAEs is increased oxygen demand due to enhanced primary productivity. Increased nutrient inputs to fuel primary productivity may have come from marine sediments as the burial efficiency of phosphorus declines when bottom waters become anoxic (Ingall and Jahnke, 1994; Van Cappellen and Ingall, 1994). The recovery from OAE like conditions is thought to involve the permanent removal of excess CO₂ from the atmosphere and ocean by burying carbon in the form of organic matter in marine sediments (e.g. Arthur et al., 1988; Jarvis et al., 2011), which is consistent with the geological record of widespread black shale formation (Stein et al., 1986). However, the overall amount, exact timing and the rate of organic matter burial remain a topic of an ongoing debate.

Diagenetic Models:

Therefore, globally quantifying the burial and degradation of organic matter in marine sediments and related biogeochemical dynamics is important for understanding climate and the cycling of many chemical elements on various timescales. Such studies and quantifications are possible through the application of idealised mathematical representations of diagenesis, or so-called diagenetic models (see e.g. Berner, 1980; Boudreau, 1997). The number of research questions that can be addressed with diagenetic models is infinite and a plethora of different approaches have been developed, mainly following two distinct directions (Arndt et al., 2013).

First, state-of-the art vertically resolved numerical models simulating the entire suite of essential coupled redox and equilibrium reactions within marine sediments (e.g. BRNS, Aguilera et al., 2005; CANDI, Boudreau, 1996; MEDIA, Meysman et al., 2003; STEADYSED, Van Cappellen and Wang, 1996). These “complete”, non-steady-state models, thus resolve the resulting characteristic redox-zonation of marine sediments through explicitly including oxic OM degradation, denitrification, oxidation by manganese and iron (hydr)oxides, sulfate reduction and methanogenesis as well as the reoxidation of reduced byproducts (i.e. NH₄, Mn²⁺, Fe²⁺, H₂S, CH₄). Furthermore, they incorporate various mineral dissolution and precipitation reactions, as well as fast equilibrium sorption processes for example of NH₄, PO₄ and metal ions (i.e. Mn²⁺, Fe²⁺ and Mg²⁺, compare Van Cappellen and Wang, 1996; Meysman et al., 2003). Modelled, depth-dependent, transport processes usually comprise advection, diffusion, bioturbation and bio-irrigation. This group of diagenetic models generally uses a so-called multi-G approach (Jørgensen, 1978; Berner, 1980), thus dividing the bulk organic matter pool into a number of compound classes that are characterised by different degradabilities k_i , which are generally dependent on the type and concentration of the specific terminal electron acceptor (TEA). These complex models have a great potential for quantifying OM degra-

dation dynamics for sites where enough observations are available to constrain its model parameters (see e.g. Boudreau et al., 1998; Wang and Van Cappellen, 1996; Thullner et al., 2009, for applications). However, due to the high degree of coupled processes and depth-varying parameters the diagenetic equation needs to be solved numerically, thus resulting in a very high computational demand and consequently rendering their application in an Earth system model framework prohibitive. Additionally, their global applicability is limited by the restricted transferability of model parameters from one site to the global scale (Arndt et al., 2013).

The second group of models solves the diagenetic equation analytically or semi-analytically, thus providing an alternative and computational more efficient approach. However, finding an analytical solution, especially when complex reaction networks are to be considered, is not straightforward and generally requires the assumption of steady state. The complexity of the reaction network can be reduced by dividing the sediment column into distinct zones and accounting for the most pertinent biogeochemical processes within each zone, thus increasing the likelihood of finding an analytical solution. In general, analytical diagenetic models are less sophisticated and comprehensive than numerical models and are used for the coupling to global ESMs (e.g. HAMOCC and NorESM use the model of Heinze et al. (1999) or box models (e.g. DCESS, Shaffer et al., 2008 or MBM using MEDUSA, Munhoven, 2007). These analytic or semi-analytical models account for the most important transport processes (i.e advection, bioturbation and molecular diffusion) through basic parametrizations and include fewer biogeochemical reactions which are generally restricted to the upper, bioturbated 10 cm of the sediments. They assume that the sedimentary organic matter pool is composed of just a single compound class which is either degraded with a globally invariant degradation rate constant (Munhoven, 2007) or a fixed rate constant depending on local oxygen concentrations (Shaffer et al., 2008; Palastanga et al., 2011). Pore water tracers explicitly represented in DCESS (Shaffer et al., 2008) and the HAMOCC model of Heinze et al. (1999) and Palastanga et al. (2011) are restricted to DIC, TA, PO₄ and O₂. The MEDUSA model (Munhoven, 2007) considers CO₂, HCO₃⁻, CO₃²⁻ and O₂. Other species produced or consumed during OM degradation are neglected. Thus, with oxygen being the only TEA explicitly modelled the influence of reduced species is only implicitly included in the boundary conditions for O₂. A newer versions of the HAMOCC model, being a notable exception, as Ilyina et al. (2013) include NO₃ and denitrification explicitly. Furthermore, the version of Palastanga et al. (2011) represents an redox-dependent explicit sedimentary phosphorus cycle. Yet, secondary redox reactions involving reduced substances or sorption processes are not included in any of the discussed models.

How are sediments resolved in Earth system models:

Earth system models generally track the biogeochemical dynamics of organic and inorganic carbon, essential nutrients (nitrogen, phosphorus) and oxygen with the aim of investigating the evolution of the ocean's redox structure and carbonate system and its feedbacks on global climate. This general aim thus defines a minimum set of state variables and reaction processes that need to be resolved

for an efficient representation of the benthic-pelagic coupling in Earth system models. A suitable sediment model has to provide a robust quantification of organic (and inorganic) carbon burial fluxes,
170 as well as benthic uptake/return fluxes of oxygen, growth-limiting nutrients and reduced species. As a consequence, the reaction network must account for the most important primary and secondary redox reactions, equilibrium reactions, mineral precipitation/dissolution and adsorption/desorption, resulting in a complex set of coupled reaction-transport equations.

Even though there are more appropriate sediment representations, in most current ESMs sediment-
175 water dynamics are either neglected or treated in a very simplistic way (Soetaert et al., 2000; Hülse et al., 2017). Most Earth system Models of Intermediate Complexity (EMICs) and also some of the higher resolution global carbon cycle models represent the sediment-water interface either as a reflective or a conservative/semi-reflective boundary (Hülse et al., 2017). Thus, all particulate material deposited on the seafloor is either instantaneously consumed (reflective boundary), or a fixed
180 fraction is buried in the sediments (conservative/semi-reflective boundary). Both highly simplified approaches furthermore completely neglect the exchange of solute species through the sediment-water interface and, therefore, cannot resolve the complex benthic-pelagic coupling. However, due to their computational efficiency, both representations are often used in global biogeochemical models (e.g. Najjar et al., 2007; Ridgwell et al., 2007; Goosse et al., 2010). A superior approach is
185 the vertically integrated dynamic model, which represents the whole sediment column as a single box (Hülse et al., 2017). Here, OM deposited on the seafloor is added to the sediment box where it gets degraded and dissolved species diffuse through the sediment-water interface in accordance with these transformations. This approach thus ignores the vertical extent of the sediments and the temporary storage of dissolved species (Soetaert et al., 2000). Yet, it is computationally efficient
190 and allows differentiating between various fractions of organic matter. Most EMICs incorporate a vertically integrated dynamic model for particulate inorganic carbon only (i.e. mainly CaCO_3) and just a few consider oxic-only sediment degradation of organic matter (Hülse et al., 2017).

The most complex description of diagenetic organic matter degradation in Earth system models is the second group of vertically resolved diagenetic models as discussed above (e.g. Heinze et al.,
195 1999; Munhoven, 2007; Shaffer et al., 2008). These models solve the one-dimensional reaction-transport equation for a number of solid and dissolved species for the upper, bioturbated 10 cm of the sediments. Examples of global ESMs employing a vertically resolved diagenetic model are NorESM (Tjiputra et al., 2013) and HAMOCC (Palastanga et al., 2011; Ilyina et al., 2013), both using a version of Heinze et al. (1999). None of the EMICs reviewed by Hülse et al. (2017) use such a
200 sediment representation. DCESS (Shaffer et al., 2008) and MBM (Munhoven, 2007) are box models employing a vertically resolved diagenetic model. However, in general oxygen is the only TEA explicitly modelled and secondary redox reactions and reduced species are completely neglected in these approaches. Furthermore, all models represent the bulk OM pool as a single fraction with a fixed degradation rate constant.

205 **Problem with that:**

Obviously, such a simplification of the OM pool can neither account for the observed vast structural complexity in natural organic matter and its resulting different degradation rates nor for the rapid decrease in OM degradability in the uppermost centimetres of the sediments (Arndt et al., 2013). It has been suggested that at least a 3G approach is necessary to accurately represent organic matter dynamics in this part of the sediments where most OM is degraded (e.g. Soetaert et al., 1996b). Even more restrictive is the use of O₂ as the only TEA and the complete absence of reduced substances and related secondary redox reactions. Even though for the majority of the modern sediments (i.e. in the deep-ocean) O₂ is the primary electron acceptor and Archer et al. (2002) suggested that aerobic degradation accounts for 66% of total organic matter respiration more recent model and data studies have reported that sulfate reduction is the dominant degradation pathway on a global average (with contributions of 55-76% Canfield et al., 2005; Jørgensen and Kasten, 2006; Thullner et al., 2009). O₂ becomes progressively less important as TEA with decreasing seafloor depth and in shallow waters most of it is used to reoxidise reduced substances produced during anaerobic degradation (Canfield et al., 2005; Thullner et al., 2009). Thus, the in situ production of e.g. NO₃ and SO₄ through oxidation of NH₄ and H₂S forms an important sink for O₂ which is entirely neglected in current sediment representations in global models. In addition, due to the lack of an appropriate sedimentary P cycle (with the exception of the HAMOCC version of Palastanga et al. (2011), no current global ESM is able to model the redox dependent P release from marine sediments and its implications for primary productivity, global biogeochemical cycles and climate.

225 **Solution presented here:**

DH: Include this discussion here: Analytical approaches with distinct layers were implemented and used in the eighties to describe observed pore water profiles (e.g. Billen, 1982; Goloway and Bender, 1982) and later for inclusion into global Earth system models (Tromp et al., 1995).

However, in addition to the oxic zone these models only describe one anoxic zone, either a denitrification layer (Billen, 1982; Goloway and Bender, 1982) or a sulfate reduction layer (Tromp et al., 1995). In addition, they ignore reoxidation of reduced species produced in the anoxic layers.

In order to provide a more realistic description of organic matter degradation and nutrient cycles in marine sediments we have developed the OrganicMatter ENabled SEDiment model (OMEN-SED), a new, one-dimensional, numerically efficient reactive transport model. OMEN-SED is the first analytical model to explicitly describe OM cycling as well as associated dynamics of the most important TEAs (i.e. O₂, NO₃, SO₄), related reduced substances (NH₄, H₂S), the full suite of secondary-redox reactions, macronutrients (PO₄) and associated pore water quantities (ALK, DIC). To represent a redox-dependent sedimentary P cycle we consider the formation and burial of Fe-bound P and authigenic Ca-P minerals. Thus, OMEN-SED captures most of the features of a complex, numerical diagenetic model, however, its computational efficiency allows the coupling to global Earth system

DH: which
Vandenborgh papers? &
with ERSEM you mean
the Plymouth model?
Their model is vertically
integrated...

models and therefore the investigation of coupled global biogeochemical dynamics over different timescales. Here, the model is presented as a 2G-approach, however, a third, non-degradable OM fraction can easily be added and OMEN-SED can be further extended to a Multi-G approach.

245 **SA: started here**

general notes:

-name zones I and II bio and non-bio

-name redox zones according to Canfield paper

- be consistent in the paper and always use the same terminology e.g. zone, boundary,...

250 **- don't use I and II for the bioturb, non-bioturb zones and the redox zones. For the redox zones, you actually do not need to use I, II, and III at all, because you name the zone explicitly, so this should be clear**

DH: Not 100% sure about redox zones: I don't use suboxic anymore (Canfield and Thamdrup, 2009), and just use anoxic when referring to all zones $z < z_{\text{ox}}$ together, defined in Sec. 2.2.2)??

2 Model Description

OMEN-SED a new, one-dimensional, computationally efficient reaction-transport model is designed 255 for the coupling to regional/global biogeochemical and Earth system models. OMEN-SED is implemented as a Fortran version that can be easily coupled to a pelagic model via the coupling routine **OMEN_SED_main**. In addition, OMEN-SED exists as a stand-alone version implemented in MATLAB and the entire model can be executed on a standard personal computer in less than 0.1 seconds. The source code of both, the Fortran and the MATLAB stand-alone version, as well as instructions 260 for executing OMEN-SED and for plotting model results are available as a supplement to this paper. The following section provides a detailed description of OMEN-SED. Tables 1 and 14 summarise the biogeochemical reaction network and Tables 9 and 10 provide a glossary of model parameters along with their respective units.

SA: maybe include some examples and test figures in supplement?

2.1 General Model Approach

265 The calculation of benthic uptake, recycling and burial fluxes is based on the vertically resolved conservation equation for solid and dissolved species in porous media (e.g. Berner, 1980; Boudreau, 1997):

$$\frac{\partial \xi C_i}{\partial t} = -\frac{\partial F}{\partial z} + \xi \sum_j R_i^j \quad (1)$$

where C_i is the concentration of biogeochemical species i , ξ equals the porosity ϕ for solute species 270 and $(1 - \phi)$ for solid species. The term z is the sediment depth, t denotes the time, F summarises the transport fluxes and $\sum_j R_i^j$ represents the sum of all biogeochemical rates j affecting species i .

OMEN-SED accounts for both the advective, as well as the diffusive transport of dissolved and solid species. Solid and dissolved species are buried in the sediment according to a constant burial rate w , thus neglecting the effect of sediment compaction (i.e. $\frac{\partial \phi}{\partial z} = 0$) due to mathematical con-

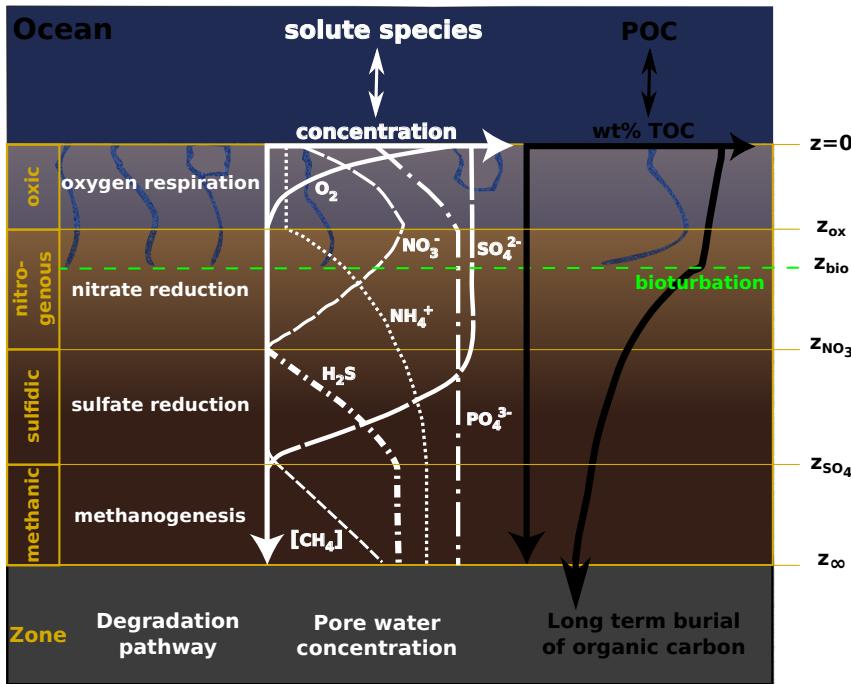


Figure 1. Schematic of the different modelled species and zones in OMEN-SED. Here showing the case $z_{\text{ox}} < z_{\text{bio}} < z_{\text{NO}_3} < z_{\text{SO}_4}$.

275 strains. The molecular diffusion of dissolved species is described by Fick's law applying a species-specific apparent diffusion coefficient, $D_{\text{mol},i}$. In addition, the activity of infaunal organisms in the bioturbated zone is simulated using a diffusive term (e.g. Boudreau, 1986), with a constant bioturbation coefficient D_{bio} in the bioturbated zone, while D_{bio} is set to zero below z_{bio} . The pumping activity by burrow-dwelling animals and the resulting ventilation of tubes, the so-called bioirrigation, is encapsulated in a factor f_{ir} that enhances the molecular diffusion coefficient (hence, $D_{i,0} = D_{\text{mol},i} \cdot f_{\text{ir}}$, Soetaert et al., 1996a). The flux divergence is thus given by:

$$\frac{\partial F}{\partial z} = -\frac{\partial}{\partial z} \left(-\xi D_i \frac{\partial C_i}{\partial z} + \xi w C_i \right) \quad (2)$$

where D_i is the apparent diffusion coefficient of species i ($D_i = D_{i,0} + D_{\text{bio}} = D_{\text{mol},i} \cdot f_{\text{ir}} + D_{\text{bio}}$ for dissolved species and $D_i = D_{\text{bio}}$ for solid species) and w is the burial rate.

285 The reaction network of OMEN-SED accounts for the most important primary and secondary redox reactions, equilibrium reactions, mineral dissolution and precipitation, as well as adsorption and desorption processes that affect the dissolved and solid species explicitly resolved in OMEN-SED. Tables 1 and 14 provide a summary of the reactions and biogeochemical tracers considered in OMEN-SED together with their respective reaction stoichiometry.

290 Because OMEN-SED is designed for the coupling to Earth system models it requires a computationally efficient solution of Eq. (1) and, thus, solves Eq. (1) analytically, by assuming steady state

SA: Tab 1 is maybe a bit redundant? **DH:** I still think it's a good summary/overview!? and the other one could go into the appendix!?

Table 1. Reactions and biogeochemical tracers implemented in the reaction network of OMEN-SED. The primary and secondary redox reactions are listed in the sequence they occur with increasing sediment depth.

Description	
Primary redox reactions	Degradation of organic matter via aerobic degradation, denitrification, sulfate reduction, methanogenesis (implicit)
Secondary redox reactions	Oxidation of ammonium and sulfide by oxygen, anaerobic oxidation of methane by sulfate
Adsorption/Desorption	Ad-/Desorption of P on/from Fe(OH) ₃ , NH ₄ adsorption, PO ₄ adsorption
Mineral precipitation	Formation of authigenic P
Biogeochemical tracers	Organic matter (2-G), oxygen, nitrate, ammonium, sulfate, sulfide (hydrogen sulfide), phosphate, Fe-bound P, DIC, ALK

conditions (i.e. $\frac{\partial C_i}{\partial t} = 0$). By describing the reactive term R_i^j in Eq. (1) more explicitly and applying Eq. (2), the general steady-state transport-reaction equation for a generic tracer C_i can be formulated as:

$$\frac{\partial C_i}{\partial t} = 0 = D_i \frac{\partial^2 C_i}{\partial z^2} - w \frac{\partial C_i}{\partial z} - k \cdot C_i - \sum_j \alpha_j \exp(-\beta_j z) + Q \quad (3)$$

DH: What about ξ in Eq. 3? Also include it here? Is this correct?

where k is a generic reaction rate constant and $g(z) = -\sum_j \alpha_j \exp(-\beta_j z) + Q$ describes the inhomogeneous reaction (production/consumption) of tracer C_i , where Q is a zeroth-order (or constant) rate of C_i reaction, $1/\beta_j$ can be interpreted as the length scale and α_j as the relative importance (or magnitude at $z = 0$) of reaction j (Boudreau, 1997).

The analytical solution of Eq. (3) is of the general form:

$$C_i(z) = A \exp(az) + B \exp(bz) + \sum_j \frac{\alpha_j}{D\beta_j^2 - w\beta_j - k} \cdot \exp(-\beta_j z) + \frac{Q}{k} \quad (4)$$

DH: Should it be $D\beta_i^2 + w\beta_i - k$ in the denom. - see Boudreau (1997) Exp 6-13 Eq. 6.121

where A and B are integration constants that are determined with the boundary conditions (see Section 2.3). In addition to assuming steady-state conditions, the sediment column is divided into a number of distinct geochemical zones which do not allow an overlap of degradation pathways (Fig. 1). Each of these zones is characterised by a different set of equations that encapsulate the most pertinent reaction and transport processes within the respective zone in a linear form (see Section 2.2), thus allowing a simplification of the coupled set of equations in order to find the analytical solutions (compare e.g. Billen, 1982; Goloway and Bender, 1982; Tromp et al., 1995; Gypens et al., 2008, for similar solutions). More specifically, OMEN-SED divides the sediment column into I) a bioturbated and II) a non-bioturbated zone defined by an imposed, constant bioturbation depth z_{bio} (Fig. 1). Furthermore, it resolves the dynamic redox stratification of marine sediments by dividing the sediment into 1) an oxic zone delineated by the oxygen penetration depth z_{ox} ; 2) a denitrification (or nitrogenous) zone situated between z_{ox} and the nitrate penetration depth z_{NO_3} ; 3) a sulfate reduction

DH: More Info on these models given in Intro! which Vandeborght papers? - also ERSEM unclear (see Intro)

zone situated between z_{NO_3} and the sulfate penetration depth z_{SO_4} ; and 4) a methanogenic zone situated below z_{SO_4} (Fig. 1). All penetration depths are dynamically calculated by the model and their depth thus varies in response to changing boundary conditions and forcings. At each dynamic, internal boundary ($z_b \in \{z_{\text{bio}}, z_{\text{ox}}, z_{\text{NO}_3}, z_{\text{SO}_4}\}$) the concentration and flux for the two solutions of
 320 the respective reaction-transport equation above and below z_b is matched (compare e.g. Billen, 1982; Gypens et al., 2008, and Section 2.3.1). The steady-state solution for OM, which follows a simple exponential decay, is calculated first (Section 2.2.1) and drives the oxygen (Section 2.2.2) and other porewater dynamics (Sections 2.2.3 - 2.2.7). The reoxidation of reduced substances is accounted for implicitly by adding a (consumption/production) flux to the internal boundary condition at z_{ox} and
 325 z_{SO_4} for oxygen and the respective tracers (Section 2.2). This simplification has been used previously by Gypens et al. (2008) for nitrate and ammonium and can be justified as it has been shown that the reoxidation mainly occurs at an infinitesimally thin layer at the oxic/anoxic interface (Soetaert et al., 1996b).

Based on Eq. (3) and its analytical solution Eq. (4), OMEN-SED calculates the fraction of particulate organic carbon (POC) buried in the sediment, f_{POC} , as well as the benthic uptake/return fluxes F_{C_i} of dissolved species C_i (in $\text{mol cm}^{-2} \text{ year}^{-1}$) in response to changing boundary conditions and
 330 forcings:

$$f_{\text{POC}} = \frac{\text{POC}(z_\infty)}{\text{POC}(0)} \quad (5)$$

$$F_{C_i} = \phi(0) \left(D_i \frac{\partial C_i(z)}{\partial z} \Big|_{z=0} - w [C_i(0) - C_i(z_\infty)] \right) \quad (6)$$

where w is the deposition rate, D_i is the diffusion coefficient and $\text{POC}(0)$, $\text{POC}(z_\infty)$, $C_i(0)$, $C_i(z_\infty)$ denotes the concentration of POC and dissolved species i at the SWI and at the lower sediment boundary, respectively.

2.2 Conservation Equations and Analytical Solution

340 2.2.1 Organic matter or Particulate Organic Carbon (POC)

In marine sediments, particulate organic carbon (POC) is degraded by heterotrophic activity coupled to the sequential utilisation of terminal electron acceptors according to the free energy gain of the half-reaction $(\text{O}_2 > \text{NO}_3^- > \text{MnO}_2 > \text{Fe(OH)}_3 > \text{SO}_4^{2-}$, e.g. Stumm and Morgan, 2012). Here, organic matter degradation is described via a multi G-model approach (Arndt et al., 2013,
 345 and references therein), dividing the bulk OM into a number i of discrete compound classes POC_i characterised by class-specific first-order degradation rate constants k_i . The conservation equation for organic matter dynamics is thus given by:

$$\frac{\partial \text{POC}_i}{\partial t} = 0 = D_{\text{POC}_i} \frac{\partial^2 \text{POC}_i}{\partial z^2} - w \frac{\partial \text{POC}_i}{\partial z} - k_i \cdot \text{POC}_i \quad (7)$$

DH: @ Sandra: Not sure what Bernie confused (flux of reduced sp. or general boundary matching) & which paper it is: (Boudreau et al., 1992; Boudreau, 1992)!!??

Table 2. OM Boundary conditions applied in OMEN-SED. For the boundaries we define: $z_{\text{bio}}^- := \lim_{h \rightarrow 0} (z_{\text{bio}} - h)$ and $z_{\text{bio}}^+ := \lim_{h \rightarrow 0} (z_{\text{bio}} + h)$. **Awesome! Du bist ja doch ein Mathematiker!**

DH: A very poor one though ;)

Boundary	Condition	
$z = 0$	known concentration	1) $\text{POC}_i(0) = \text{POC}_{0i}$
$z = z_{\text{bio}}$	continuity	2) $\text{POC}_i(z_{\text{bio}}^-) = \text{POC}_i(z_{\text{bio}}^+)$ 3) $-D_{\text{bio}} \cdot \frac{\partial \text{POC}_i}{\partial z} \Big _{z_{\text{bio}}^-} = 0$

with $D_{\text{POC}_i} = D_{\text{bio}}$ for $z \leq z_{\text{bio}}$ and $D_{\text{POC}_i} = 0$ for $z > z_{\text{bio}}$. Integration of equations (7) yields the
350 following general solutions:

I. Bioturbated zone ($z \leq z_{\text{bio}}$)

$$\begin{aligned} \text{POC}_i^I(z) &= A_{1i} \cdot \exp(a_{1i}z) + B_{1i} \cdot \exp(b_{1i}z) \\ &\stackrel{\text{BC1)}{=} A_{1i} \cdot [\exp(a_{1i}z) - \exp(b_{1i}z)] + \text{POC}_{0i} \cdot \exp(b_{1i}z) \end{aligned} \quad (8)$$

II. Non-bioturbated zone ($z_{\text{bio}} < z < z_{\text{ox}}$)

$$\text{POC}_i^{II}(z) = A_{2i} \cdot \exp(a_{2i}z) \quad (9)$$

where

$$a_{1i} = \frac{w - \sqrt{w^2 + 4D_{\text{POC}_i} \cdot k_i}}{2D_{\text{POC}_i}}, \quad b_{1i} = \frac{w + \sqrt{w^2 + 4D_{\text{POC}_i} \cdot k_i}}{2D_{\text{POC}_i}}, \quad a_{2i} = -\frac{k_i}{w} \quad (10)$$

Determining the integration constants ($A_{1,i}$, $B_{1,i}$, $A_{2,i}$) requires the definition of a set of boundary conditions (Table 2). For organic matter, OMEN-SED applies a known concentration/flux at the sediment-water interface and assumes continuity across the bottom of the bioturbated zone, z_{bio} . See Section 2.3.1 for further details on how to find the analytical solution.

365 2.2.2 Oxygen

OMEN-SED explicitly accounts for oxygen consumption by the aerobic degradation of organic matter within the oxic zone, as well as the oxidation of reduced species (e.g. NH_4 , H_2S , CH_4) produced in the anoxic zones of the sediment (here all zones $z < z_{\text{ox}}$ combined). In the oxic zone ($z < z_{\text{ox}}$), the aerobic degradation consumes oxygen with a fixed $\text{O}_2 : \text{C}$ ratio (O_2C , Tab. 10). A predefined fraction, γ_{NH_4} , of the ammonium produced during the aerobic degradation of OM is nitrified to nitrate, consuming two moles of oxygen per mole of ammonium produced. In addition, OMEN-SED implicitly accounts for the oxygen consumption due to oxidation of reduced species (NH_4 , H_2S) produced below the oxic zone through the flux boundary condition at the dynamically calculated (SA: see section ?? for details) oxygen penetration depth z_{ox} . This simplification can be justified
370 as it has been shown that these secondary redox reactions mainly occur at the oxic/anoxic interface

DH: comment Sandra,
which Section do you
mean here?

(Soetaert et al., 1996b). All oxygen consumption processes can thus be formulated as a function of organic matter degradation. The conservation equation for oxygen is given by: **SA: I'd show the POC substitution in the equations below:**

$$380 \quad \frac{\partial O_2}{\partial t} = 0 = D_{O_2} \frac{\partial^2 O_2}{\partial z^2} - w \frac{\partial O_2}{\partial z} - \frac{1-\phi}{\phi} \sum_i k_i \cdot [O_2 C + 2\gamma_{NH_4} NC_i] \cdot POC_i(z) \quad (11)$$

I Bioturbated zone ($z \leq z_{bio}$)

$$\frac{\partial O_2^I}{\partial t} \stackrel{s}{=} D_{O_2}^I \frac{\partial^2 O_2}{\partial z^2} - w \frac{\partial O_2}{\partial z} - \frac{1-\phi}{\phi} \sum_i k_i \cdot [O_2 C + 2\gamma_{NH_4} NC_i] \cdot \left(A_{1i} \cdot [exp(a_{1i}z) - exp(b_{1i}z)] + POC_{0i} \cdot exp(b_{1i}z) \right)$$

385 II Non-bioturbated zone ($z_{bio} < z < z_{ox}$)

$$\frac{\partial O_2^{II}}{\partial t} \stackrel{g}{=} D_{O_2}^{II} \frac{\partial^2 O_2}{\partial z^2} - w \frac{\partial O_2}{\partial z} - \frac{1-\phi}{\phi} \sum_i k_i \cdot [O_2 C + 2\gamma_{NH_4} NC_i] \cdot \left(A_{2i} \cdot exp(a_{2i}z) \right)$$

where $D_{O_2}^I$ and $D_{O_2}^{II}$ denote the O_2 diffusion coefficient for the bioturbated and non-bioturbated zone, respectively. The term $\frac{1-\phi}{\phi}$ accounts for the volume conversion from solid to dissolved phase

390 and NC_i is the nitrogen to carbon ratio in OM. **SA: explain all terms**

Integration yields the following analytical solution for each zone:

I Bioturbated zone ($z \leq z_{bio}$):

$$395 \quad O_2^I(z) = A_{O_2}^1 + B_{O_2}^1 \cdot exp(b_{O_2}^1 z) + \sum_i \Phi_{1,i}^I \cdot exp(a_{1i}z) + \sum_i \Phi_{1,i}^{II} \cdot exp(b_{1i}z) + \sum_i \Phi_{1,i}^{III} \cdot exp(b_{1i}z) \quad (12)$$

II Non-bioturbated zone ($z_{bio} < z < z_{ox}$)

$$O_2^{II}(z) = A_{O_2}^2 + B_{O_2}^2 \cdot exp(b_{O_2}^2 z) + \sum_i \Phi_{i,2}^I \cdot exp(a_{2i}z) \quad (13)$$

with

$$400 \quad b_{O_2}^1 = \frac{w}{D_{O_2}^I}, \quad b_{O_2}^2 = \frac{w}{D_{O_2}^{II}}$$

$$\Phi_{1,i}^I = \frac{1-\phi}{\phi} \cdot \frac{k_i \cdot (O_2 C + 2\gamma_{NH_4} NC_i) \cdot A_{1i}}{D_{O_2}^I (-a_{1i})^2 - w \cdot (-a_{1i})}, \quad \Phi_{1,i}^{II} = -\frac{1-\phi}{\phi} \cdot \frac{k_i \cdot (O_2 C + 2\gamma_{NH_4} NC_i) \cdot A_{1i}}{D_{O_2}^I (-b_{1i})^2 - w \cdot (-b_{1i})}$$

$$\Phi_{1,i}^{III} = \frac{1-\phi}{\phi} \cdot \frac{k_i \cdot (O_2 C + 2\gamma_{NH_4} NC_i) \cdot POC_{0i}}{D_{O_2}^I (-b_{1i})^2 - w \cdot (-b_{1i})}$$

$$\Phi_{i,2}^I := \frac{1-\phi}{\phi} \cdot \frac{k_i \cdot (O_2 C + 2\gamma_{NH_4} NC_i) \cdot A_{2i}}{D_{O_2}^{II} (-a_{2i})^2 - w \cdot (-a_{2i})}$$

405 Determining the four integration constants ($A_{O_2}^1, B_{O_2}^1, A_{O_2}^2, B_{O_2}^2$, see Section 2.3 for details), as well as the *a priori* unknown oxygen penetration depth requires the definition of five boundary

DH: you mean the way I did it or in the solution?
I think it's easier to understand (also how to get the solution) if it's done in the ODE

DH: Other terms are explained earlier in 2.2.2 or 2.2.1

Table 3. Boundary conditions for oxygen. For the boundaries we define: $z_{\text{bio}}^- := \lim_{h \rightarrow 0} (z_{\text{bio}} - h)$ and $z_{\text{bio}}^+ := \lim_{h \rightarrow 0} (z_{\text{bio}} + h)$.

Boundary	Condition	
$z = 0$	known concentration	1) $O_2(0) = O_{20}$
$z = z_{\text{bio}}$	continuity	2) $O_2(z_{\text{bio}}^-) = O_2(z_{\text{bio}}^+)$
$z = z_{\text{ox}}$	O_2 consumption ($z_{\text{ox}} = z_\infty$)	3) $-(D_{O_2,0} + D_{\text{bio}}) \cdot \frac{\partial O_2}{\partial z} \Big _{z_{\text{bio}}^-} = -D_{O_2,0} \cdot \frac{\partial O_2}{\partial z} \Big _{z_{\text{bio}}^+}$
	$(z_{\text{ox}} < z_\infty)$	4) IF ($O_2(z_\infty) > 0$) 4.1) $\frac{\partial O_2}{\partial z} \Big _{z_{\text{ox}}} = 0$ ELSE
	with flux from below	4.2) $O_2(z_{\text{ox}}) = 0 \quad \text{and} \quad -D_{O_2} \cdot \frac{\partial O_2}{\partial z} \Big _{z_{\text{ox}}} = F_{\text{red}}(z_{\text{ox}})$ $F_{\text{red}}(z_{\text{ox}}) = \frac{1-\phi}{\phi} \cdot \int_{z_{\text{ox}}}^\infty \sum_i (2\gamma_{\text{NH}_4} \text{NC}_i + 2\gamma_{\text{H}_2\text{S}} \text{SO}_4 \text{C}) k_i \text{POC}_i dz$

conditions (see Table 3). At the sediment-water interface, OMEN-SED applies a Dirichlet condition (i.e. known concentration) and assumes concentration and flux continuity across the bottom of the bioturbated zone, z_{bio} . The oxygen penetration depth z_{ox} marks the lower boundary and is dynamically calculated as the depth at which $O_2(z) = 0$. Therefore, OMEN-SED applies a Dirichlet boundary condition $O_2(z_{\text{ox}}) = 0$. In addition, a flux boundary is applied that implicitly accounts for the oxygen consumption by the partial oxidation of NH_4 and H_2S diffusing into the oxic zone from below (BC 4.2, Table 3). It is assumed that respective fractions (γ_{NH_4} and $\gamma_{\text{H}_2\text{S}}$) are directly reoxidised at the oxic/anoxic interface. OMEN-SED iteratively solves for z_{ox} by first testing if there is oxygen left at z_∞ (i.e. $O_2(z_\infty) > 0$) and, otherwise, by finding the root for the flux boundary condition 4.2 (Table 3). If $z_{\text{ox}} = z_\infty$, a zero flux boundary condition is applied as lower boundary condition.

2.2.3 Nitrate and Ammonium

Nitrogen dynamics in OMEN-SED are controlled by the metabolic production of ammonium, nitrification, denitrification as well as ammonium adsorption. Ammonium is produced by organic matter degradation in both the oxic and anoxic zones, while denitrification consumes nitrate in the denitrification zone with a fixed $\text{NO}_3^- : \text{C}$ ratio ($\text{NO}_3^- \text{C}$, Tab. 10) **SA: need explanation**. The adsorption of ammonium to sediment particles is formulated as an equilibrium process with constant equilibrium adsorption coefficient K_{NH_4} , thus assuming that the adsorption is fast compared with the characteristic time scales of transport processes (Wang and Van Cappellen, 1996). In addition, a defined fraction, γ_{NH_4} , of metabolically produced ammonium is directly nitrified to nitrate in the oxic zone, while the nitrification of upward diffusing ammonium produced in the sulfidic and methanic zones is implicitly accounted for in the boundary conditions. The conservation equations for ammonium and nitrate are thus given by:

430

DH: Edited explanation of finding z_{ox}

DH: explanation sufficient?

1. Oxic zone ($z \leq z_{\text{ox}}$)

$$\frac{\partial \text{NO}_3^I}{\partial t} = 0 = D_{\text{NO}_3} \frac{\partial^2 \text{NO}_3^I}{\partial z^2} - w \frac{\partial \text{NO}_3^I}{\partial z} + \gamma_{\text{NH}_4} \frac{1-\phi}{\phi} \cdot \sum_i \text{NC}_i \cdot k_i \cdot \text{POC}_i(z) \quad (14)$$

$$\frac{\partial \text{NH}_4^I}{\partial t} = 0 = \frac{D_{\text{NH}_4}}{1+K_{\text{NH}_4}} \frac{\partial^2 \text{NH}_4^I}{\partial z^2} - w \frac{\partial \text{NH}_4^I}{\partial z} + \frac{1-\gamma_{\text{NH}_4}}{1+K_{\text{NH}_4}} \cdot \frac{1-\phi}{\phi} \cdot \sum_i \text{NC}_i \cdot k_i \cdot \text{POC}_i(z) \quad (15)$$

435

2. Denitrification (or nitrogenous) zone ($z_{\text{ox}} < z \leq z_{\text{NO}_3}$)

$$\frac{\partial \text{NO}_3^{II}}{\partial t} = 0 = D_{\text{NO}_3} \frac{\partial^2 \text{NO}_3^{II}}{\partial z^2} - w \frac{\partial \text{NO}_3^{II}}{\partial z} - \frac{1-\phi}{\phi} \text{NO}_3 \text{C} \cdot \sum_i k_i \cdot \text{POC}_i(z) \quad (16)$$

$$\frac{\partial \text{NH}_4^{II}}{\partial t} = 0 = \frac{D_{\text{NH}_4}}{1+K_{\text{NH}_4}} \frac{\partial^2 \text{NH}_4^{II}}{\partial z^2} - w \frac{\partial \text{NH}_4^{II}}{\partial z} \quad (17)$$

440 3. Sulfidic and methanic zone ($z_{\text{NO}_3} < z \leq z_{\infty}$)

$$\frac{\partial \text{NH}_4^{III}}{\partial t} = 0 = \frac{D_{\text{NH}_4}}{1+K_{\text{NH}_4}} \frac{\partial^2 \text{NH}_4^{III}}{\partial z^2} - w \frac{\partial \text{NH}_4^{III}}{\partial z} + \frac{1}{1+K_{\text{NH}_4}} \cdot \frac{1-\phi}{\phi} \cdot \sum_i \text{NC}_i \cdot k_i \cdot \text{POC}_i(z) \quad (18)$$

where D_{NO_3} and D_{NH_4} denote the diffusion coefficients for NO_3 and NH_4 which depend on the bioturbation status of the respective geochemical zone (compare Section 2.3.1). Integration of Eq. (14) - (18) yields the analytical solutions, which are not further developed here but follow the procedure outlined in Section 2.2.2 for oxygen (also see Section 2.3.1 for more details on how to find the analytical solution). Table 4 summarises the boundary conditions applied in OMEN-SED to solve Eq. (14) - (18) and to find the *a priori* unknown nitrate penetration depth, z_{NO_3} . The model assumes known bottom water concentrations for both species, the complete consumption of nitrate at the nitrate penetration depth (in case $z_{\text{NO}_3} < z_{\infty}$) and no change in nitrate and ammonium flux at z_{∞} . In addition, concentration and diffusive flux continuity across z_{bio} and z_{ox} is considered for NO_3 and NH_4 . Furthermore, the reoxidation of upward-diffusing reduced ammonium is accounted for in the oxic-anoxic boundary condition for nitrate and ammonium. OMEN-SED iteratively solves for z_{NO_3} by first testing if there is nitrate left at z_{∞} (i.e. $\text{NO}_3(z_{\infty}) > 0$) and, otherwise, by finding the root for the flux boundary condition 6.2 (Table 4).

DH: all other params explained earlier

2.2.4 Sulfate and Sulfide

Below the denitrification zone ($z > z_{\text{NO}_3}$), organic matter degradation is coupled to sulfate reduction, consuming sulfate and producing hydrogen sulfide with a fixed $\text{SO}_4 : \text{C}$ ratio ($\text{SO}_4 \text{C}$, Tab. 10). In addition, the anaerobic oxidation of upward diffusing methane (AOM) produced below the sulfate penetration and the associated consumption of sulfate and production of sulfide; as well as the production of sulfate and consumption of sulfide through sulfide oxidation are implicitly accounted

Table 4. Boundary conditions for nitrate and ammonium. For the boundaries we define: $z_-^- := \lim_{h \rightarrow 0} (z_- - h)$ and $z_-^+ := \lim_{h \rightarrow 0} (z_- + h)$.

Boundary	Condition	
$z = 0$	known concentration	1) $\text{NO}_3(0) = \text{NO}_{30}$
$z = z_{\text{bio}}$	continuity	2) $\text{NO}_3(z_{\text{bio}}^-) = \text{NO}_3(z_{\text{bio}}^+)$
$z = z_{\text{ox}}$	continuity	3) $-(D_{\text{NO}_3,0} + D_{\text{bio}}) \cdot \frac{\partial \text{NO}_3}{\partial z} _{z_{\text{bio}}^-} = -D_{\text{NO}_3,0} \cdot \frac{\partial \text{NO}_3}{\partial z} _{z_{\text{bio}}^+}$
	where:	4) $\text{NO}_3(z_{\text{ox}}^-) = \text{NO}_3(z_{\text{ox}}^+)$
$z = z_{\text{NO}_3}$	NO_3 consumption ($z_{\text{NO}_3} = z_\infty$)	5) $-D_{\text{NO}_3} \cdot \frac{\partial \text{NO}_3}{\partial z} _{z_{\text{ox}}^-} + \gamma_{\text{NH}_4} \cdot F_{\text{NH}_4}(z_{\text{ox}}) = -D_{\text{NO}_3} \cdot \frac{\partial \text{NO}_3}{\partial z} _{z_{\text{ox}}^+}$ $F_{\text{NH}_4}(z_{\text{ox}}) = \frac{1}{1+K_{\text{NH}_4}} \cdot \frac{1-\phi}{\phi} \cdot \int_{z_{\text{NO}_3}}^\infty \sum_i k_i \cdot \text{NC}_i \cdot \text{POC}_i dz$
		6) IF ($\text{NO}_3(z_\infty) > 0$)
		6.1) $\frac{\partial \text{NO}_3}{\partial z} _{z_{\text{NO}_3}} = 0$
		ELSE
		6.2) $\text{NO}_3(z_{\text{NO}_3}) = 0 \quad \text{and} \quad \frac{\partial \text{NO}_3}{\partial z} _{z_{\text{NO}_3}} = 0$
$z = 0$	known concentration	1) $\text{NH}_4(0) = \text{NH}_{40}$
$z = z_{\text{bio}}$	continuity	2) $\text{NH}_4(z_{\text{bio}}^-) = \text{NH}_4(z_{\text{bio}}^+)$
$z = z_{\text{ox}}$	continuity	3) $-\frac{D_{\text{NH}_4,0} + D_{\text{bio}}}{1+K_{\text{NH}_4}} \cdot \frac{\partial \text{NH}_4}{\partial z} _{z_{\text{bio}}^-} = -\frac{D_{\text{NH}_4,0}}{1+K_{\text{NH}_4}} \cdot \frac{\partial \text{NH}_4}{\partial z} _{z_{\text{bio}}^+}$
	where:	4) $\text{NH}_4(z_{\text{ox}}^-) = \text{NH}_4(z_{\text{ox}}^+)$
$z = z_{\text{NO}_3}$	continuity flux	5) $-\frac{D_{\text{NH}_4}}{1+K_{\text{NH}_4}} \cdot \frac{\partial \text{NH}_4}{\partial z} _{z_{\text{ox}}^-} - \gamma_{\text{NH}_4} \cdot F_{\text{NH}_4}(z_{\text{ox}}) = -\frac{D_{\text{NH}_4}}{1+K_{\text{NH}_4}} \cdot \frac{\partial \text{NH}_4}{\partial z} _{z_{\text{ox}}^+}$ $F_{\text{NH}_4}(z_{\text{ox}}) = \frac{1}{1+K_{\text{NH}_4}} \cdot \frac{1-\phi}{\phi} \cdot \int_{z_{\text{NO}_3}}^\infty \sum_i k_i \cdot \text{NC}_i \cdot \text{POC}_i dz$
$z = z_\infty$	zero NH_4 flux	6) $\text{NH}_4(z_{\text{NO}_3}^-) = \text{NH}_4(z_{\text{NO}_3}^+)$
		7) $-\frac{D_{\text{NH}_4}}{1+K_{\text{NH}_4}} \cdot \frac{\partial \text{NH}_4}{\partial z} _{z_{\text{NO}_3}^-} = -\frac{D_{\text{NH}_4}}{1+K_{\text{NH}_4}} \cdot \frac{\partial \text{NH}_4}{\partial z} _{z_{\text{NO}_3}^+}$
		8) $\frac{\partial \text{NH}_4}{\partial z} _{z_\infty} = 0$

for through the boundary conditions (Table 5). The conservation equations for sulfate and sulfide are thus given by:

465

1. Oxic and nitrogenous zone ($z \leq z_{\text{NO}_3}$)

$$\frac{\partial \text{SO}_4^{I}}{\partial t} = 0 = D_{\text{SO}_4} \frac{\partial^2 \text{SO}_4^{I}}{\partial z^2} - w \frac{\partial \text{SO}_4^{I}}{\partial z} \quad (19)$$

$$\frac{\partial \text{H}_2\text{S}^I}{\partial t} = 0 = D_{\text{H}_2\text{S}} \frac{\partial^2 \text{H}_2\text{S}^I}{\partial z^2} - w \frac{\partial \text{H}_2\text{S}^I}{\partial z} \quad (20)$$

470 2. Sulfidic zone ($z_{\text{NO}_3} < z \leq z_{\text{SO}_4}$)

$$\frac{\partial \text{SO}_4^{II}}{\partial t} = 0 = D_{\text{SO}_4} \frac{\partial^2 \text{SO}_4^{II}}{\partial z^2} - w \frac{\partial \text{SO}_4^{II}}{\partial z} - \frac{1-\phi}{\phi} \cdot \sum_i \text{SO}_4 \text{C} \cdot k_i \cdot \text{POC}_i(z) \quad (21)$$

$$\frac{\partial \text{H}_2\text{S}^{II}}{\partial t} = 0 = D_{\text{H}_2\text{S}} \frac{\partial^2 \text{H}_2\text{S}^{II}}{\partial z^2} - w \frac{\partial \text{H}_2\text{S}^{II}}{\partial z} + \frac{1-\phi}{\phi} \cdot \sum_i \text{SO}_4 \text{C} \cdot k_i \cdot \text{POC}_i(z) \quad (22)$$

3. Methanic zone ($z_{SO_4} < z \leq z_\infty$)

475
$$\frac{\partial H_2S^{III}}{\partial t} = 0 = D_{H_2S} \frac{\partial^2 H_2S^{III}}{\partial z^2} - w \frac{\partial H_2S^{III}}{\partial z} \quad (23)$$

where D_{SO_4} and D_{H_2S} denote the diffusion coefficients for SO_4 and H_2S which depend on the bioturbation status of the respective geochemical zone (compare Section 2.3.1). Integration of Eq. (19) - (23) yields the analytical solution and Table 5 summarises the boundary conditions applied.

480 OMEN-SED assumes known concentrations at the sediment-water interface and continuity across the bioturbation depth and the nitrate penetration depth. The reoxidation of reduced H_2S to SO_4 is accounted for implicitly via the oxic-anoxic boundary condition for both species, while reduction of SO_4 and the associated production of H_2S via AOM is accounted for through the respective boundary conditions at z_{SO_4} . In case $z_{SO_4} < z_\infty$, OMEN-SED assumes zero sulfate concentration at 485 z_{SO_4} and its diffusive flux must equal the amount of methane produced below; or, in case $z_{SO_4} = z_\infty$, a zero flux condition for sulfate is considered. OMEN-SED iteratively solves for z_{SO_4} by first testing if there is sulfate left at z_∞ (i.e. $SO_4(z_\infty) > 0$) and, otherwise, by finding the root for the flux boundary condition 8.2 (Table 5). At the lower boundary z_∞ zero flux of H_2S is considered.

2.2.5 Phosphate

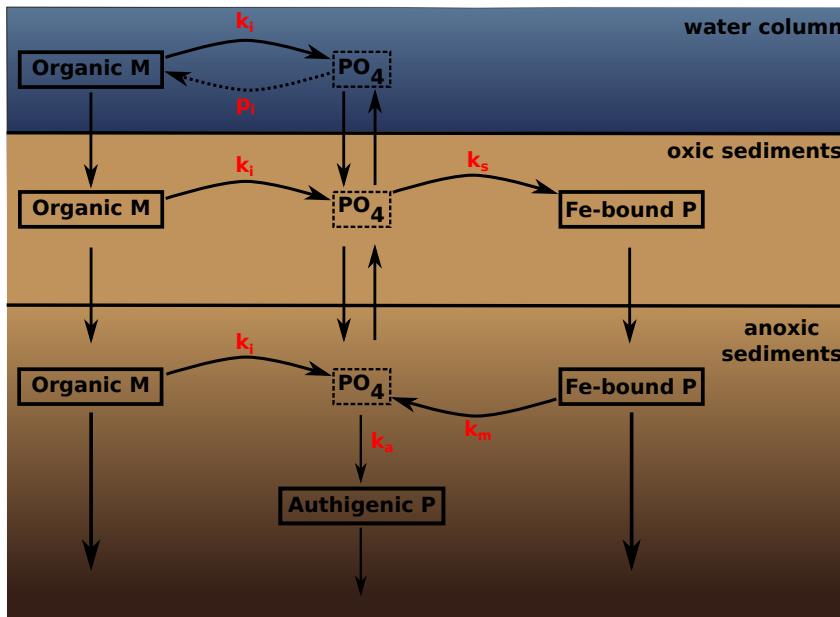


Figure 2. A schematic of the sedimentary P cycle in OMEN-SED. Red numbers represent kinetic rate constants for phosphorus dynamics (compare Table 10; p_i represents uptake rate of PO_4 via primary production in shallow environments). Adapted from Slomp et al. (1996).

Table 5. Boundary conditions for sulfate and sulfide. For the boundaries we define: $z_-^- := \lim_{h \rightarrow 0} (z_- - h)$ and $z_-^+ := \lim_{h \rightarrow 0} (z_- + h)$.

Boundary	Condition	
$z = 0$	known concentration	1) $\text{SO}_4(0) = \text{SO}_{40}$
$z = z_{\text{bio}}$	continuity	2) $\text{SO}_4(z_{\text{bio}}^-) = \text{SO}_4(z_{\text{bio}}^+)$
	flux	3) $-(D_{\text{SO}_4,0} + D_{\text{bio}}) \cdot \frac{\partial \text{SO}_4}{\partial z} _{z_{\text{bio}}^-} = -D_{\text{SO}_4,0} \cdot \frac{\partial \text{SO}_4}{\partial z} _{z_{\text{bio}}^+}$
$z = z_{\text{ox}}$	continuity	4) $\text{SO}_4(z_{\text{ox}}^-) = \text{SO}_4(z_{\text{ox}}^+)$
	flux	5) $-D_{\text{SO}_4} \cdot \frac{\partial \text{SO}_4}{\partial z} _{z_{\text{ox}}^-} + \gamma_{\text{H}_2\text{S}} \cdot F_{\text{H}_2\text{S}}(z_{\text{ox}}) = -D_{\text{SO}_4} \cdot \frac{\partial \text{SO}_4}{\partial z} _{z_{\text{ox}}^+}$
	where:	$F_{\text{H}_2\text{S}}(z_{\text{ox}}) = \frac{1-\phi}{\phi} \cdot \left(\int_{z_{\text{NO}_3}}^{\text{SO}_4} \sum_i \text{SO}_4 \cdot k_i \cdot \text{POC}_i dz + \gamma_{\text{CH}_4} \cdot \int_{z_{\text{SO}_4}}^{\infty} \sum_i \text{MC} \cdot k_i \cdot \text{POC}_i dz \right)$
$z = z_{\text{NO}_3}$	continuity	6) $\text{SO}_4(z_{\text{NO}_3}^-) = \text{SO}_4(z_{\text{NO}_3}^+)$
	flux	7) $-D_{\text{SO}_4} \cdot \frac{\partial \text{SO}_4}{\partial z} _{z_{\text{NO}_3}^-} = -D_{\text{SO}_4} \cdot \frac{\partial \text{SO}_4}{\partial z} _{z_{\text{NO}_3}^+}$
$z = z_{\text{SO}_4}$	SO ₄ consumption ($z_{\text{SO}_4} = z_\infty$)	8) IF ($\text{SO}_4(z_\infty) > 0$)
	$(z_{\text{SO}_4} < z_\infty)$	8.1) $\frac{\partial \text{SO}_4}{\partial z} _{z_{\text{SO}_4}} = 0$
	with flux from below:	ELSE
		8.2) $\text{SO}_4(z_{\text{SO}_4}) = 0 \quad \text{and} \quad -D_{\text{SO}_4} \cdot \frac{\partial \text{SO}_4}{\partial z} _{z_{\text{SO}_4}} = \gamma_{\text{CH}_4} \cdot F_{\text{CH}_4}(z_{\text{SO}_4})$
		$F_{\text{CH}_4}(z_{\text{SO}_4}) = \frac{1-\phi}{\phi} \cdot \int_{z_{\text{SO}_4}}^{\infty} \sum_i \text{MC} \cdot k_i \cdot \text{POC}_i dz$
$z = 0$	known concentration	1) $\text{H}_2\text{S}(0) = \text{H}_2\text{S}_0$
$z = z_{\text{bio}}$	continuity	2) $\text{H}_2\text{S}(z_{\text{bio}}^-) = \text{H}_2\text{S}(z_{\text{bio}}^+)$
	flux	3) $-(D_{\text{H}_2\text{S},0} + D_{\text{bio}}) \cdot \frac{\partial \text{H}_2\text{S}}{\partial z} _{z_{\text{bio}}^-} = -D_{\text{H}_2\text{S},0} \cdot \frac{\partial \text{H}_2\text{S}}{\partial z} _{z_{\text{bio}}^+}$
$z = z_{\text{ox}}$	continuity	4) $\text{H}_2\text{S}(z_{\text{ox}}^-) = \text{H}_2\text{S}(z_{\text{ox}}^+)$
	flux	5) $-D_{\text{H}_2\text{S}} \cdot \frac{\partial \text{H}_2\text{S}}{\partial z} _{z_{\text{ox}}^-} - \gamma_{\text{H}_2\text{S}} F_{\text{H}_2\text{S}}(z_{\text{ox}}) = -D_{\text{H}_2\text{S}} \cdot \frac{\partial \text{H}_2\text{S}}{\partial z} _{z_{\text{ox}}^+}$
	where:	$F_{\text{H}_2\text{S}}(z_{\text{ox}}) = \frac{1-\phi}{\phi} \cdot \left(\int_{z_{\text{NO}_3}}^{\text{SO}_4} \sum_i \text{SO}_4 \cdot k_i \cdot \text{POC}_i dz + \gamma_{\text{CH}_4} \cdot \int_{z_{\text{SO}_4}}^{\infty} \sum_i \text{MC} \cdot k_i \cdot \text{POC}_i dz \right)$
$z = z_{\text{NO}_3}$	continuity	6) $\text{H}_2\text{S}(z_{\text{NO}_3}^-) = \text{H}_2\text{S}(z_{\text{NO}_3}^+)$
	flux	7) $-D_{\text{H}_2\text{S}} \cdot \frac{\partial \text{H}_2\text{S}}{\partial z} _{z_{\text{NO}_3}^-} = -D_{\text{H}_2\text{S}} \cdot \frac{\partial \text{H}_2\text{S}}{\partial z} _{z_{\text{NO}_3}^+}$
$z = z_{\text{SO}_4}$	continuity	8) $\text{H}_2\text{S}(z_{\text{SO}_4}^-) = \text{H}_2\text{S}(z_{\text{SO}_4}^+)$
	flux (with AOM)	9) $-D_{\text{H}_2\text{S}} \cdot \frac{\partial \text{H}_2\text{S}}{\partial z} _{z_{\text{SO}_4}^-} + \gamma_{\text{CH}_4} \cdot F_{\text{CH}_4}(z_{\text{SO}_4}) = -D_{\text{H}_2\text{S}} \cdot \frac{\partial \text{H}_2\text{S}}{\partial z} _{z_{\text{SO}_4}^+}$
	where:	$F_{\text{CH}_4}(z_{\text{SO}_4}) = \frac{1-\phi}{\phi} \cdot \int_{z_{\text{SO}_4}}^{\infty} \sum_i \text{MC} \cdot k_i \cdot \text{POC}_i dz$
$z = z_\infty$	zero H ₂ S flux	10) $\frac{\partial \text{H}_2\text{S}}{\partial z} _{z_\infty} = 0$

490 The biogeochemical description of phosphorus (P) dynamics builds on the work of Slomp et al.
 (1996) and accounts for phosphorous recycling through organic matter degradation, adsorption onto
 sediments and iron(III) hydroxides (Fe-bound P), as well as carbonate fluorapatite (CFA or authi-
 genic P) formation (see Figure 2 for a schematic overview of the sedimentary P cycle). In the oxic
 zone of the sediment, PO_4 liberated through organic matter degradation can adsorb to iron(III) hy-
 495 droxides forming Fe-bound P (or M, Slomp et al., 1998). Below the oxic zone, PO_4 is not only
 produced via organic matter degradation but can also be released from the Fe-bound P pool due to
 the reduction of iron(III) hydroxides under anoxic conditions. Furthermore, in these zones phosphate
 concentrations build up and pore waters can thus become supersaturated with respect to carbonate
 fluorapatite, thus triggering the authigenic formation of CFA (Van Cappellen and Berner, 1988).
 500 Phosphorus bound in these authigenic minerals represents a permanent sink for reactive phosphorus
 (Slomp et al., 1996). As for ammonium, the adsorption of P to the sediment matrix is treated as
 an equilibrium processes, parameterised with dimensionless adsorption coefficients for the oxic and
 anoxic zone, respectively ($K_{\text{PO}_4}^{\text{ox}}$, $K_{\text{PO}_4}^{\text{anox}}$ Slomp et al., 1998). The sorption and desorption of P to
 iron(III) hydroxides as well as the authigenic fluorapatite formation are described as first-order re-
 505 actions with rate constants k_s , k_m and k_a , respectively (Table 10). The rate of the respective process
 is calculated as the product of the rate constant and the difference between the current concentration
 (of PO_4 and M) and an equilibrium or asymptotic concentration Slomp et al. (1996). The asymptotic
 Fe-bound P concentration is M^∞ and the equilibrium concentration for P sorption and authigenic
 fluorapatite formation are PO_4^s and PO_4^a , respectively (Table 10).
 510

DH: enough Info on
Fe-adsorption and M?

1. Oxic zone ($z \leq z_{\text{ox}}$)

$$\frac{\partial \text{PO}_4^I}{\partial t} = \frac{D_{\text{PO}_4}}{1 + K_{\text{PO}_4}^{\text{ox}}} \frac{\partial^2 \text{PO}_4^I}{\partial z^2} - w \frac{\partial \text{PO}_4^I}{\partial z} + \frac{1-\phi}{\phi} \frac{1}{1 + K_{\text{PO}_4}^{\text{ox}}} \sum_i (\text{PC}_i \cdot k_i \cdot \text{POC}_i(z)) - \frac{k_s}{1 + K_{\text{PO}_4}^{\text{ox}}} (\text{PO}_4^I - \text{PO}_4^s) \quad (24)$$

$$\frac{\partial M^I}{\partial t} = D_M \frac{\partial^2 M^I}{\partial z^2} - w \frac{\partial M^I}{\partial z} + \frac{\phi}{1-\phi} k_s (\text{PO}_4^I - \text{PO}_4^s) \quad (25)$$

2. Anoxic zones ($z_{\text{ox}} < z \leq z_\infty$)

$$\frac{\partial M^{II}}{\partial t} = D_M \frac{\partial^2 M^{II}}{\partial z^2} - w \frac{\partial M^{II}}{\partial z} - k_m (M^{II} - M^\infty) \quad (26)$$

$$\frac{\partial \text{PO}_4^{II}}{\partial t} = \frac{D_{\text{PO}_4}}{1 + K_{\text{PO}_4}^{\text{anox}}} \frac{\partial^2 \text{PO}_4^{II}}{\partial z^2} - w \frac{\partial \text{PO}_4^{II}}{\partial z} + \frac{1-\phi}{\phi} \frac{1}{1 + K_{\text{PO}_4}^{\text{anox}}} \sum_i (\text{PC}_i \cdot k_i \cdot \text{POC}_i(z)) - \frac{k_a}{1 + K_{\text{PO}_4}^{\text{anox}}} (\text{PO}_4^{II} - \text{PO}_4^a) + \frac{(1-\phi)}{\phi} \frac{k_m}{1 + K_{\text{PO}_4}^{\text{anox}}} (M^{II} - M^\infty) \quad (27)$$

520 where D_{PO_4} denotes the diffusion coefficient for PO_4 which depends on the bioturbation status of
 the respective geochemical zone and $D_M = D_{\text{bio}}$ for $z \leq z_{\text{bio}}$ and $D_M = 0$ for $z > z_{\text{bio}}$ (compare
 Section 2.3.1). Integration of Eq. (24) - (27) yields the analytical solution and Table 6 summarises

Table 6. Boundary conditions for phosphate and Fe-bound P (M). For the boundaries we define: $z_-^- := \lim_{h \rightarrow 0} (z_- - h)$ and $z_+^+ := \lim_{h \rightarrow 0} (z_- + h)$.

Boundary	Condition	
$z = 0$	known concentration	1) $\text{PO}_4(0) = \text{PO}_{40}$
$z = z_{\text{bio}}$	continuity	2) $\text{PO}_4(z_{\text{bio}}^-) = \text{PO}_4(z_{\text{bio}}^+)$
	flux	3) $(D_{\text{PO}_4,0} + D_{\text{bio}}) \cdot \frac{\partial \text{PO}_4}{\partial z} _{z_{\text{bio}}^-} = D_{\text{PO}_4,0} \cdot \frac{\partial \text{PO}_4}{\partial z} _{z_{\text{bio}}^+}$
$z = z_{\text{ox}}$	continuity	4) $\text{PO}_4(z_{\text{ox}}^-) = \text{PO}_4(z_{\text{ox}}^+)$
	flux	5) $-\frac{D_{\text{PO}_4}}{1 + K_{\text{PO}_4}^{2z}} \cdot \frac{\partial \text{PO}_4}{\partial z} _{z_{\text{ox}}^-} = -\frac{D_{\text{PO}_4}}{1 + K_{\text{PO}_4}^{\text{anox}}} \cdot \frac{\partial \text{PO}_4}{\partial z} _{z_{\text{ox}}^+}$
$z = z_\infty$	flux	10) $\frac{\partial \text{PO}_4}{\partial z} _{z_\infty} = 0$
$z = 0$	known concentration	1) $M(0) = M_0$
$z = z_{\text{bio}}$	continuity	2) $M(z_{\text{bio}}^-) = M(z_{\text{bio}}^+)$
	flux	3) $\frac{\partial M}{\partial z} _{z_{\text{bio}}^-} = \frac{\partial M}{\partial z} _{z_{\text{bio}}^+}$
$z = z_{\text{ox}}$	continuity	4) $M(z_{\text{ox}}^-) = M(z_{\text{ox}}^+)$
	flux	5) $\frac{\partial M}{\partial z} _{z_{\text{ox}}^-} = \frac{\partial M}{\partial z} _{z_{\text{ox}}^+}$
$z = z_\infty$	asymptotic concentration	10) $M(z_\infty) = M_\infty$

the boundary conditions applied in OMEN-SED. The model assumes known bottom water concentrations and equal concentrations and diffusive fluxes at z_{bio} and z_{ox} for both species. A no flux boundary condition is applied at z_∞ .

2.2.6 Dissolved Inorganic Carbon (DIC)

OMEN-SED accounts for the production of dissolved inorganic carbon (DIC) through organic matter degradation, as well as methane oxidation. Organic matter degradation produces dissolved inorganic carbon with a stoichiometric DIC : C ratio of 1:2 in the methanic zone and 1:1 in the rest of the sediment column (DICC^{II} and DICC^{I} respectively). DIC production through methane oxidation is implicitly taken into account through the boundary condition at z_{SO_4} . The conservation equations for DIC are thus given by:

1. Oxic, nitrogenous and sulfidic zone ($z \leq z_{\text{SO}_4}$)

$$\frac{\partial \text{DIC}^{\text{I}}}{\partial t} = 0 = D_{\text{DIC}} \frac{\partial^2 \text{DIC}^{\text{I}}}{\partial z^2} - w \frac{\partial \text{DIC}^{\text{I}}}{\partial z} + \frac{1 - \phi}{\phi} \cdot \sum_i \text{DICC}^{\text{I}} \cdot k_i \cdot \text{POC}_i(z) \quad (28)$$

2. Methanic zone ($z_{\text{SO}_4} < z \leq z_\infty$)

$$\frac{\partial \text{DIC}^{\text{II}}}{\partial t} = 0 = D_{\text{DIC}} \frac{\partial^2 \text{DIC}^{\text{II}}}{\partial z^2} - w \frac{\partial \text{DIC}^{\text{II}}}{\partial z} + \frac{1 - \phi}{\phi} \cdot \sum_i \text{DICC}^{\text{II}} \cdot k_i \cdot \text{POC}_i(z) \quad (29)$$

where D_{DIC} denotes the diffusion coefficient for DIC which depends on the bioturbation status of the respective geochemical zone. Integration of Eq. (28) and (29) yields the analytical solution and

DH: I think what I wrote before also applies: a asymptotic Fe-bound P concentration at z_∞ is assumed! (but is the same as “no flux condition” as it is now?)

SA: need to say something about carbonates here

Table 7. Boundary conditions for DIC. For the boundaries we define: $z_-^- := \lim_{h \rightarrow 0} (z_- - h)$ and $z_-^+ := \lim_{h \rightarrow 0} (z_- + h)$.

Boundary	Condition	
$z = 0$	known concentration	1) $\text{DIC}(0) = \text{DIC}_0$
$z = z_{\text{bio}}$	continuity	2) $\text{DIC}(z_{\text{bio}}^-) = \text{DIC}(z_{\text{bio}}^+)$
	flux	3) $-(D_{\text{DIC},0} + D_{\text{bio}}) \cdot \frac{\partial \text{DIC}}{\partial z} \Big _{z_{\text{bio}}^-} = -D_{\text{DIC},0} \cdot \frac{\partial \text{DIC}}{\partial z} \Big _{z_{\text{bio}}^+}$
$z = z_{\text{SO}_4}$	continuity	4) $\text{DIC}(z_{\text{SO}_4}^-) = \text{DIC}(z_{\text{SO}_4}^+)$
	flux (with AOM)	5) $-D_{\text{DIC}} \cdot \frac{\partial \text{DIC}}{\partial z} \Big _{z_{\text{SO}_4}^-} + \gamma_{\text{CH}_4} \cdot F_{\text{CH}_4}(z_{\text{SO}_4}) = -D_{\text{DIC}} \cdot \frac{\partial \text{DIC}}{\partial z} \Big _{z_{\text{SO}_4}^+}$
	where:	$F_{\text{CH}_4}(z_{\text{SO}_4}) = \frac{1-\phi}{\phi} \cdot \int_{z_{\text{SO}_4}}^{\infty} \sum_i \text{MC} \cdot k_i \cdot \text{POC}_i dz$
$z = z_\infty$	zero DIC flux	6) $\frac{\partial \text{DIC}}{\partial z} \Big _{z_\infty} = 0$

Table 7 summarises the boundary conditions applied in OMEN-SED. A Dirichlet condition is applied at the sediment-water interface. In addition, the model assumes a zero flux through the lower boundary z_∞ and continuity across the bottom of the bioturbated zone, as well as the sulfate penetration depth. An additional flux boundary condition at z_{SO_4} , implicitly accounts for DIC production through anaerobic oxidation of methane (Table 7 Eq. 5).

2.2.7 Alkalinity

Organic matter degradation and secondary redox reactions exert a complex influence on alkalinity (Wolf-Gladrow et al., 2007). To model alkalinity, OMEN-SED divides the sediment column into four geochemical zones, where different equations describe the biogeochemical processes using variable stoichiometric coefficients (compare Tables 10). Above z_{ox} , the combined effects of NH_4 and P release due to aerobic OM degradation increases alkalinity according to ALK^{OX} whereas nitrification decreases alkalinity with stoichiometry ALK^{NIT} . In the remaining three zones anaerobic OM degradation generally results in an increase in alkalinity, with the exact magnitude depending on the nature of the terminal electron acceptor used (i.e. ALK^{DEN} , ALK^{SUL} , ALK^{MET}). In addition, the effect of secondary redox reactions, such as nitrification, methane and sulfide oxidation are implicitly accounted for in the boundary conditions.

In OMEN-SED, the conservation equations for alkalinity is thus given by:

1. Oxic zone ($z \leq z_{\text{ox}}$)

$$\begin{aligned} \frac{\partial \text{ALK}^I}{\partial t} = 0 &= D_{\text{ALK}} \frac{\partial^2 \text{ALK}^I}{\partial z^2} - w \frac{\partial \text{ALK}^I}{\partial z} \\ &+ \frac{1-\phi}{\phi} \cdot \sum_i \left(\text{ALK}^{\text{NIT}} \cdot \frac{\gamma_{\text{NH}_4}}{1 + K_{\text{NH}_4}} \text{NC}_i + \text{ALK}^{\text{OX}} \right) \cdot k_i \cdot \text{POC}_i(z) \end{aligned} \quad (30)$$

SA: TODO: again need to mention carboantes

565 2. Denitrification or nitrogenous zone ($z_{\text{ox}} < z \leq z_{\text{NO}_3}$)

$$\frac{\partial \text{ALK}^{II}}{\partial t} = 0 = D_{\text{ALK}} \frac{\partial^2 \text{ALK}^{II}}{\partial z^2} - w \frac{\partial \text{ALK}^{II}}{\partial z} + \frac{1-\phi}{\phi} \cdot \sum_i \text{ALK}^{\text{DEN}} \cdot k_i \cdot \text{POC}_i(z) \quad (31)$$

3. Sulfidic zone ($z_{\text{NO}_3} < z \leq z_{\text{SO}_4}$)

$$\frac{\partial \text{ALK}^{III}}{\partial t} = 0 = D_{\text{ALK}} \frac{\partial^2 \text{ALK}^{III}}{\partial z^2} - w \frac{\partial \text{ALK}^{III}}{\partial z} + \frac{1-\phi}{\phi} \cdot \sum_i \text{ALK}^{\text{SUL}} \cdot k_i \cdot \text{POC}_i(z) \quad (32)$$

570 4. Methanic zone ($z_{\text{SO}_4} < z \leq z_{\infty}$)

$$\frac{\partial \text{ALK}^{IV}}{\partial t} = 0 = D_{\text{ALK}} \frac{\partial^2 \text{ALK}^{IV}}{\partial z^2} - w \frac{\partial \text{ALK}^{IV}}{\partial z} + \frac{1-\phi}{\phi} \cdot \sum_i \text{ALK}^{\text{MET}} \cdot k_i \cdot \text{POC}_i(z) \quad (33)$$

575 where D_{ALK} denotes the diffusion coefficient for alkalinity which depends on the bioturbation status of the respective geochemical zone. Integration of Eq. (30) - (33) yields the analytical solution and Table 8 summarises the boundary conditions applied in OMEN-SED. A Dirichlet boundary condition is applied at the sediment-water interface. The decrease of alkalinity due to oxidation of reduced species produced in the anoxic zones (with stoichiometry ALK^{NIT} and $\text{ALK}^{\text{H}_2\text{S}}$) is implicitly taken into account through the flux boundary condition at z_{ox} (Table 8 Eq. 5). Furthermore, the 580 oxidation of methane by sulfate reduction increases alkalinity with stoichiometry ALK^{AOM} which is accounted for through the flux boundary condition at z_{SO_4} (Table 8 Eq. 9). At the lower boundary z_{∞} a zero flux condition is applied.

2.3 Determination of Integration Constants

585 The integration constants of all general analytical solutions derived above change in response to changing boundary conditions. Thus, OMEN-SED has to re-determine integration constants for each dynamic zone (i.e. $z_{\text{ox}}, z_{\text{bio}}, z_{\text{NO}_3}$ and z_{SO_4}) at every time step. The bioturbation boundary poses a particular challenge as it can theoretically occur in any of the dynamic geochemical zones (compare Fig. 3). Therefore, in order to generalise and simplify this recurring boundary matching problem, an independent algorithm is implemented (rather than using a fully-worked-out algebraic solution).

590 The algorithm only ever has to solve a two-simultaneous-equation problem.

2.3.1 Generic Boundary Condition Matching (GBCM)

As discussed in Section 2.1, the solution of the general steady-state transport-reaction equation (Eq. (3)) for a generic tracer C is of the general form:

$$C(z) = A \exp(az) + B \exp(bz) + \sum_j \frac{\alpha_j}{D\beta_j^2 - w\beta_j - k} \cdot \exp(-\beta_j z) + \frac{Q}{k} \quad (34)$$

DH: Should it be $D\beta_i^2 + w\beta_i - k$ in the denom. - see Boudreau (1997) Exp 6-13 Eq.

595 and can therefore be expressed as:

$$C(z) = A \cdot E(z) + B \cdot F(z) + G(z) \quad (35) \quad 6.121$$

Table 8. Boundary conditions for alkalinity. For the boundaries we define: $z_-^- := \lim_{h \rightarrow 0} (z_- - h)$ and $z_-^+ := \lim_{h \rightarrow 0} (z_- + h)$.

Boundary	Condition	
$z = 0$	known concentration	1) $\text{ALK}(0) = \text{ALK}_0$
$z = z_{\text{bio}}$	continuity	2) $\text{ALK}(z_{\text{bio}}^-) = \text{ALK}(z_{\text{bio}}^+)$
	flux	3) $-(D_{\text{ALK},0} + D_{\text{bio}}) \cdot \frac{\partial \text{ALK}}{\partial z} _{z_{\text{bio}}^-} = -D_{\text{ALK},0} \cdot \frac{\partial \text{ALK}}{\partial z} _{z_{\text{bio}}^+}$
$z = z_{\text{ox}}$	continuity	4) $\text{ALK}(z_{\text{ox}}^-) = \text{ALK}(z_{\text{ox}}^+)$
	flux	5) $-D_{\text{ALK}} \cdot \frac{\partial \text{ALK}}{\partial z} _{z_{\text{ox}}^-} + F_{\text{ALK}}(z_{\text{ox}}) = -D_{\text{ALK}} \cdot \frac{\partial \text{ALK}}{\partial z} _{z_{\text{ox}}^+}$ $F_{\text{ALK}}(z_{\text{ox}}) = \frac{1-\phi}{\phi} \cdot \left(\text{ALK}^{\text{H}_2\text{S}} \cdot \gamma_{\text{H}_2\text{S}} \int_{z_{\text{NO}_3}}^{\text{SO}_4} \sum_i \text{SO}_4 \cdot k_i \cdot \text{POC}_i dz \right) + \frac{1-\phi}{\phi} \cdot \left(\text{ALK}^{\text{NIT}} \frac{\gamma_{\text{NH}_4}}{1+k_{\text{NH}_4}} \int_{z_{\text{NO}_3}}^{\infty} \sum_i \text{NC}_i \cdot k_i \cdot \text{POC}_i dz \right)$
	where:	
$z = z_{\text{NO}_3}$	continuity	6) $\text{ALK}(z_{\text{NO}_3}^-) = \text{ALK}(z_{\text{NO}_3}^+)$
	flux	7) $-D_{\text{ALK}} \cdot \frac{\partial \text{ALK}}{\partial z} _{z_{\text{NO}_3}^-} = -D_{\text{ALK}} \cdot \frac{\partial \text{ALK}}{\partial z} _{z_{\text{NO}_3}^+}$
$z = z_{\text{SO}_4}$	continuity	8) $\text{ALK}(z_{\text{SO}_4}^-) = \text{ALK}(z_{\text{SO}_4}^+)$
	flux (with AOM)	9) $-D_{\text{ALK}} \cdot \frac{\partial \text{ALK}}{\partial z} _{z_{\text{SO}_4}^-} + F_{\text{ALK}}(z_{\text{SO}_4}) = -D_{\text{ALK}} \cdot \frac{\partial \text{ALK}}{\partial z} _{z_{\text{SO}_4}^+}$ $F_{\text{ALK}}(z_{\text{SO}_4}) = \frac{1-\phi}{\phi} \cdot \left(\text{ALK}^{\text{AOM}} \gamma_{\text{CH}_4} \cdot \int_{z_{\text{SO}_4}}^{\infty} \sum_i k_i \cdot \text{POC}_i dz \right)$
	where:	
$z = z_\infty$	zero ALK flux	10) $\frac{\partial \text{ALK}}{\partial z} _{z_\infty} = 0$

where $E(z)$, $F(z)$ are the homogeneous solutions of the ODE, $G(z)$ the particular integral (collectively called the basis functions), and A , B are the integration constants that must be determined with the boundary conditions (compare Fig. 3 for the whole sediment column).

Each boundary matching problem involves matching continuity and flux for the two solutions of the respective reaction-transport equation above, $C_U(z)$ (= 'upper'), and below, $C_L(z)$ (= 'lower'), the dynamic boundary at $z = z_b$:

$$605 \quad C_U(z) = A_U \cdot E_U(z) + B_U \cdot F_U(z) + G_U(z) \quad (36)$$

$$C_L(z) = A_L \cdot E_L(z) + B_L \cdot F_L(z) + G_L(z). \quad (37)$$

OMEN-SED generally applies concentration continuity and flux boundary conditions at its internal, dynamic boundaries:

610 Continuity (where for generality we allow a discontinuity V_b)

$$C_U(z_b) = C_L(z_b) + V_b \quad (38)$$

Flux

$$615 \quad D_U C'_U(z_b) + w C_U(z_b) = D_L C'_L(z_b) + w C_L(z_b) + F_b \quad (39)$$

where w is advection, D are the diffusion coefficients and F_b is any flux discontinuity.

Substituting the general ODE solutions (36), (37), the boundary conditions can be represented as

two equations connecting the four integration constants:

$$\begin{pmatrix} E_U & F_U \\ D_U E'_U & D_U F'_U \end{pmatrix} \begin{pmatrix} A_U \\ B_U \end{pmatrix} = \begin{pmatrix} E_L & F_L \\ D_L E'_L & D_L F'_L \end{pmatrix} \begin{pmatrix} A_L \\ B_L \end{pmatrix} + \begin{pmatrix} G_L - G_U + V_b \\ D_L G'_L - D_U G'_U + F_b - wV_b \end{pmatrix} \quad (40)$$

where the ODE solutions E, F, G are all evaluated at z_b .

620

Equation (40) can now be solved to give A_U and B_U as a function of the integration constants from the layer below (A_L and B_L), thereby constructing a piecewise solution for both layers, with just two integration constants (this is implemented in the function **benthic_utils.matchsoln** of OMEN-SED):

$$625 \quad \begin{pmatrix} A_U \\ B_U \end{pmatrix} = \begin{pmatrix} c_1 & c_2 \\ c_3 & c_4 \end{pmatrix} \begin{pmatrix} A_L \\ B_L \end{pmatrix} + \begin{pmatrix} d_1 \\ d_2 \end{pmatrix}. \quad (41)$$

Using Eq. (41), $C_U(z)$ in (36) can now be rewritten as a function of A_L and B_L (implemented in **benthic_utils.xformsoln**):

$$C_U(z) = (c_1 A_L + c_2 B_L + d_1) \cdot E_U(z) + (c_3 A_L + c_4 B_L + d_2) \cdot F_U(z) + G_U(z) \quad (42)$$

and hence define the “transformed” basis functions $E_U^*(z), F_U^*(z), G_U^*(z)$ such that:

$$630 \quad C_U(z) = A_L \cdot E_U^*(z) + B_L \cdot F_U^*(z) + G_U^*(z) \quad (43)$$

where

$$\begin{aligned} E_U^*(z) &= c_1 E_U(z) + c_3 F_U(z) \\ F_U^*(z) &= c_2 E_U(z) + c_4 F_U(z) \\ 635 \quad G_U^*(z) &= G_U(z) + d_1 E_U(z) + d_2 F_U(z) \end{aligned} \quad (44)$$

Equations (41), (43) and (44) can now be consecutively applied for each of the dynamic biogeochemical zone boundaries, starting at the bottom of the sediment column. The net result is a piecewise solution of the whole sediment column with just two integration constants (coming from the lowest layer), which can then be solved for by applying the boundary conditions at the sediment-water interface and the bottom of the sediments .

2.3.2 Abstracting out the bioturbation boundary

The bioturbation boundary affects the diffusion coefficient of the modelled solutes, as well as the conservation equation of organic matter (and thereby the exact form of each reaction-transport equation). This boundary is particularly inconvenient as it can, in principle, occur in the middle of any of the dynamically shifting biogeochemical zones and therefore generates multiple cases (compare

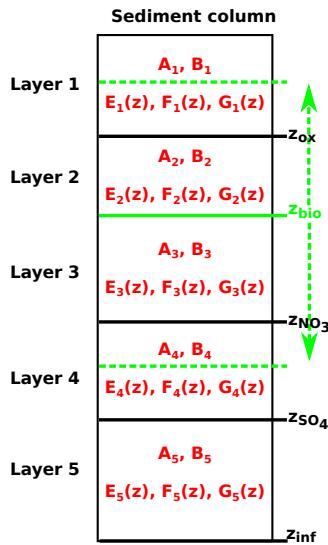


Figure 3. Schematic of the generic boundary condition matching (GBCM) problem. Showing the resulting integration constants (A_i, B_i) and ODE solutions (E_i, F_i, G_i) for the different sediment layers and the variable bioturbation boundary.

Fig. 3). The GBCM algorithm described above is thus not only used to construct a piecewise solution of the whole sediment column, but also to abstract out the bioturbation boundary. For each biogeochemical zone the "bioturbation-status" is initially tested (i.e. fully bioturbated, fully non-bioturbated, or crossing the bioturbation boundary). Therefore, the upper and lower boundaries for 650 the different zones (e.g. for the nitrogenous zone: $z_U = z_{\text{Ox}}$, $z_L = z_{\text{NO}_3}$), as well as the respective reactive terms and diffusion coefficients (bioturbated and non-bioturbated) are passed over to the routine **zTOC.prepfg_I12** where the bioturbations-status is determined. In case the bioturbation depth is located within this zone (i.e. $z_U < z_{\text{bio}} < z_L$) a piecewise solution for this layer is constructed. Therefore, the reactive terms and diffusion coefficients are handed over to the routines 655 **zTOC.calcfg_I1** and **zTOC.calcfg_I2** which calculate the basis functions (E_U, F_U, G_U and E_L, F_L, G_L) and their derivatives for the bioturbated and the non-bioturbated part of this specific geochemical zone. The concentration and flux for both solutions at z_{bio} are matched and the coefficients $c_1, c_2, c_3, c_4, d_1, d_2$ (as in Eq. (41)) are calculated by the routine **benthic_utils.matchsoln**. These coefficients and the "bioturbation-status" of the layer are passed back to the main GBCM 660 algorithm where they can be used by the routine **benthic_utils.xformsoln** to calculate the "transformed" basis functions ($E_U^*(z), F_U^*(z), G_U^*(z)$) such that both layers are expressed in the same basis (compare Eq. (42 - 44)).

For instance, in the case of sulfate, **zTOC.prepfg_I12** is called three times before the actual profile is calculated (once per zone: oxic, nitrogenous, sulfidic) and hands back the information 665 about the "bioturbation-status" of the three layers and the coefficients $c_1, c_2, c_3, c_4, d_1, d_2$ for the

biogeochemical zone including the bioturbation depth. When calculating the complete piecewise solution for the sediment column, this information is passed to the function **zTOC.calcfg_I12** which sorts out the correct solution type to use. The main GBCM algorithm therefore never needs to know whether it is dealing with a piecewise solution (i.e. matched across the bioturbation boundary) or a
670 “simple” solution (i.e. the layer is fully bioturbated or fully non-bioturbated).

2.4 Model Parameters

The following section provides a summary of global relationships used to constrain reaction and transport parameters in OMEN-SED. Table 9 synthesises sediment and transport parameters, while table 10 provides an overview of all biogeochemical parameters used in OMEN-SED.

675 2.4.1 Transport Parameters

The burial of sediments and porewater is directly related to the accumulation of new material on the seafloor (i.e. sedimentation, Burdige, 2006). This results in a downward advective flux of older sediment material and porewater in relation to the sediment-water interface. When coupled to an ocean model, its sedimentation flux can be readily used in OMEN-SED. The stand-alone version of
680 OMEN-SED uses the empirical global relationship between sediment accumulation rate (cm yr^{-1}) and seafloor depth (m) of Middelburg et al. (1997):

$$w = 3.3 \cdot 10^{-0.87478367 - 0.00043512 \cdot \text{depth}}, \quad (45)$$

As mentioned before (Sec. 2.1), the diffusion coefficient of species i is calculated as $D_i = D_{i,0} + D_{\text{bio}} = D_{\text{mol},i} \cdot f_{ir} + D_{\text{bio}}$ for dissolved species and $D_i = D_{\text{bio}}$ for solid species. The bioturbation coefficient D_{bio} ($\text{cm}^2 \text{ yr}^{-1}$) is constant in the bioturbated zone and also follows the empirical relationship by Middelburg et al. (1997):
685

$$D_{\text{bio}} = 5.2 \cdot 10^{0.76241122 - 0.00039724 \cdot \text{depth}} \quad (46)$$

Observations indicate that bioturbation is largely restricted to the upper 10 cm of the sediments and is only marginally related to seafloor depth (e.g. Boudreau, 1998; Teal et al., 2010). Therefore, OMEN-
690 SED imposes a globally invariant bioturbation depth z_{bio} of 10 cm. In case the bottom water oxygen concentration is below 5 nanomole cm^{-3} infaunal activity is assumed to cease and $z_{\text{bio}} = 0.01 \text{ cm}$. We choose a low value unequal to zero in order to simplify the implementation of the model as this ensures the same method to solve the conservation equations can be used.

SA: need to explain why not 0; **DH:** Decent explanation?

Bioirrigation can enhance the molecular diffusion coefficient $D_{i,0} = D_{\text{mol},i} \cdot f_{ir}$ (Soetaert et al.,
695 1996a). However, here we do not consider this effect as it is difficult to constrain on geological timescales and set f_{ir} to a constant value of 1 . The specific molecular diffusion coefficients $D_{\text{mol},i}$ are corrected for sediment porosity ϕ , tortuosity F and are linearly interpolated for an ambient temperature T using zero-degree coefficients D_i^0 and temperature-dependent diffusion coefficients D_i^T

(compare Gypens et al., 2008):

700
$$D_{\text{mol},i} = (D_i^0 + D_i^T \cdot T) \cdot \frac{1}{\phi \cdot F}.$$

Tortuosity can be expressed in terms of porosity as $F = \frac{1}{\phi^m}$ (Ullman and Aller, 1982) with the exponent m varying according to the type of sediment (here we use $m=3$). Values for D_i^T and D_i^0 are summarised in Table 9 and are adapted from Li and Gregory (1974), Schulz (2006) and Gypens et al. (2008).

705 **2.4.2 Reaction Parameters and Stoichiometries**

difficult to find a structure, I'd adapt to the order in the table and groups -as much as possible- according to chemical species, Also ALL parameters must be described and justified. Let me take care of this section: The first-order organic matter degradation constants of compound class i , k_i (yr^{-1}), are assumed invariant along the sediment column and therefore independent of the nature of the terminal electron acceptor. The rate constants can be altered manually to fit observed sediment profiles (compare Section 5) or related to a master variable (e.g. sedimentation rate or POC-flux) provided by a coupled Earth system model (compare Section ?? and 5.2). **until here, also need to mention f_i** Organic matter degradation consumes oxygen with a 1:1 stoichiometry.

N-related parameters: N/C, NO₃C, NH4, adsorption coefficient, gamma- explain nitrification
715 **SO₄-H₂S related parameters: SO₄C , gamma, AOM** reduction of 1 mol organic matter additionally produces SO₄C = $\frac{138}{212}$ mol of hydrogen sulfide
methane: MC, AOM MC = 0.5 mol of methane.

The rate constants sorption of PO₄ to Fe oxides, k_s , release of PO₄ from Fe-bound P due to Fe-oxide reduction k_m and authigenic CFA precipitation k_a are constrained on the basis of ???. The pore water equilibrium concentrations for P sorption and CFA precipitation (PO₄^s, PO₄^a) and the asymptotic concentration for Fe-bound P (M[∞]) are taken from Palastanga et al. (2011). The phosphorus to carbon ratio represents the composition of organic matter and is chosen as ?? based on ??.

See Table 10 for a complete summary of the parameters and their values.

725 **3 Methodology for sensitivity and case studies**

3.1 Sensitivity study

To validate our model approach an extensive sensitivity analysis for the most important model parameters is performed with the stand-alone version of OMEN-SED and resulting sediment-water interface fluxes are compared with a global database (Section 4). Model parameters implicitly account for processes that are not explicitly resolved. Therefore, model parameters are notoriously difficult to constrain and a source of uncertainty for numerical and analytical models. A comprehen-

Table 9. Sediment characteristics and transport parameters. **TODO: PO4 adsorption coefficients okay?**

Parameter	Unit	Value	Description/Source
ρ_{sed}	g cm^{-3}	2.6	Sediment density
w	cm yr^{-1}	Fct. of seafloor depth or from ESM	Advection/Sediment accumulation rate (Middelburg et al., 1997)
z_{bio}	cm	10 or 0.01	Bioturbation depth (Boudreau, 1998; Teal et al., 2010)
D_{bio}	$\text{cm}^2 \text{yr}^{-1}$	Fct. of seafloor depth	Bioturbation coefficient (Middelburg et al., 1997)
ϕ	-	0.85	Porosity
F	-	$\frac{1}{\phi^m}$	Tortuosity, here m=3
f_{ir}	-	1	Irrigation factor
Adsorption coefficients			
K_{NH_4}	-	1.3	NH_4 adsorption coefficient, (see Wang and Van Cappellen, 1996)
$K_{\text{PO}_4}^{\text{ox}}$	-	200.0	PO_4 adsorption coefficient (oxic), (see Slomp et al., 1998)
$K_{\text{PO}_4}^{\text{anox}}$	-	1.3	PO_4 adsorption coefficient (anoxic), (see Slomp et al., 1998)
Diffusion coefficients (Li and Gregory, 1974; Schulz, 2006; Gypens et al., 2008)			
$D_{\text{O}_2}^0$	$\text{cm}^2 \text{yr}^{-1}$	348.62	Molecular diffusion coefficient of oxygen at 0°C
$D_{\text{O}_2}^T$	$\text{cm}^2 \text{yr}^{-1} \text{°C}^{-1}$	14.09	Diffusion coefficient for linear temp. dependence of oxygen
$D_{\text{NO}_3}^0$	$\text{cm}^2 \text{yr}^{-1}$	308.42	Molecular diffusion coefficient of nitrate at 0°C
$D_{\text{NO}_3}^T$	$\text{cm}^2 \text{yr}^{-1} \text{°C}^{-1}$	12.26	Diffusion coefficient for linear temp. dependence of nitrate
$D_{\text{NH}_4}^0$	$\text{cm}^2 \text{yr}^{-1}$	309.05	Molecular diffusion coefficient of ammonium at 0°C
$D_{\text{NH}_4}^T$	$\text{cm}^2 \text{yr}^{-1} \text{°C}^{-1}$	12.26	Diffusion coefficient for linear temp. dependence of ammonium
$D_{\text{SO}_4}^0$	$\text{cm}^2 \text{yr}^{-1}$	157.68	Molecular diffusion coefficient of sulfate at 0°C
$D_{\text{SO}_4}^T$	$\text{cm}^2 \text{yr}^{-1} \text{°C}^{-1}$	7.88	Diffusion coefficient for linear temp. dependence of sulfate
$D_{\text{H}_2\text{S}}^0$	$\text{cm}^2 \text{yr}^{-1}$	307.48	Molecular diffusion coefficient of sulfide at 0°C
$D_{\text{H}_2\text{S}}^T$	$\text{cm}^2 \text{yr}^{-1} \text{°C}^{-1}$	9.64	Diffusion coefficient for linear temp. dependence of sulfide
$D_{\text{PO}_4}^0$	$\text{cm}^2 \text{yr}^{-1}$	112.91	Molecular diffusion coefficient of phosphate at 0°C
$D_{\text{PO}_4}^T$	$\text{cm}^2 \text{yr}^{-1} \text{°C}^{-1}$	5.59	Diffusion coefficient for linear temp. dependence of phosphate
D_{DIC}^0	$\text{cm}^2 \text{yr}^{-1}$	151.69	Molecular diffusion coefficient of DIC at 0°C
D_{DIC}^T	$\text{cm}^2 \text{yr}^{-1} \text{°C}^{-1}$	7.93	Diffusion coefficient for linear temp. dependence of DIC
D_{ALK}^0	$\text{cm}^2 \text{yr}^{-1}$	151.69	Molecular diffusion coefficient of ALK at 0°C
D_{ALK}^T	$\text{cm}^2 \text{yr}^{-1} \text{°C}^{-1}$	7.93	Diffusion coefficient for linear temp. dependence of ALK
Note: DIC and ALK coefficients are the values of HCO_3^- from Schulz (2006).			

Table 10. Values for biogeochemical parameters used in OMEN-SED. The variables x , y and z denote the atomic ratio of carbon, nitrogen and phosphorous of the degrading organic matter (here set to $C : N : P = 106 : 16 : 1$). **P related coefficients okay?**

Parameter/Variable	Unit	Value	Description
Stoichiometric factors and molecular ratios			
NC _i	mol/mol	$\frac{y}{x} = \frac{16}{106}$	nitrogen to carbon ratio
PC _i	mol/mol	$\frac{z}{x} = \frac{1}{106}$	phosphorus to carbon ratio
MC	mol/mol	0.5	methane to carbon ratio
			produced during methanogenesis
DICC ^I	mol/mol	1.0	DIC to carbon ratio until z _{SO₄}
DICC ^{II}	mol/mol	0.5	DIC to carbon ratio below z _{SO₄}
O ₂ C	mol/mol	$\frac{x+2y}{x} = \frac{138}{106}$	oxygen to carbon ratio
NO ₃ C	mol/mol	$\frac{4x+3y}{5x} = \frac{94.4}{106}$	nitrate to carbon ratio
SO ₄ C	mol/mol	$\frac{1}{2}O_2C = \frac{138}{212}$	sulfate to carbon ratio
ALK ^{OX}	mol/mol	$\frac{y-2z}{x} = \frac{14}{106}$	ALK from aerobic degradation
ALK ^{NIT}	mol/mol	-2	ALK from nitrification
ALK ^{DEN}	mol/mol	$\frac{4x+3y-10z}{5x} = \frac{92.4}{106}$	ALK from denitrification
ALK ^{SUL}	mol/mol	$\frac{x+y-2z}{x} = \frac{120}{106}$	ALK from sulfate reduction
ALK ^{MET}	mol/mol	$\frac{y-2z}{x} = \frac{14}{106}$	ALK from methanogenesis
ALK ^{H2S}	mol/mol	-2	ALK from H ₂ S oxidation
ALK ^{AOM}	mol/mol	2	ALK from AOM
Secondary reaction parameters			
γ _{NH₄}	-	0.9	fraction of NH ₄ that is oxidised in oxic zone
γ _{H₂S}	-	0.95	fraction of H ₂ S that is oxidised in oxic zone
γ _{CH₄}	-	0.99	fraction of CH ₄ that is oxidised at z _{SO₄}
P related parameters			
k _s	yr ⁻¹	1.0	Rate constant for PO ₄ sorption
k _m	yr ⁻¹	0.02	Rate constant for Fe-bound P release
k _a	yr ⁻¹	10.0	Rate constant for authigenic CFA precipitation
PO ₄ ^s	mol cm ⁻³	1 · 10 ⁻⁹	equilibrium conc. for P sorption (Slomp et al., 1996)
PO ₄ ^a	mol cm ⁻³	3.7 · 10 ⁻⁹	equilibrium conc. for authigenic P precipitation (Slomp et al., 1996)
M [∞]	mol cm ⁻³	1.99 · 10 ⁻¹⁰	asymptotic concentration for Fe-bound P (Slomp et al., 1996)

hensive sensitivity analysis (SA) can help quantify this uncertainty and identify the most sensitive parameters. More specifically, sensitivity analysis is used to investigate how the variations in the outputs (y_1, \dots, y_N) of a model can be attributed to variations in the different input parameters
735 (x_1, \dots, x_M , Pianosi et al., 2016). Different types of sensitivity indices, which quantify this relative influence with a scalar S_i , can be calculated, ranging from simple one-at-a-time methods to statistical evaluations of the output distribution (e.g. variance-based or density-based approaches Pianosi et al., 2016). The latter indices take values between zero and one ($S_i \in [0, 1]$), where zero indicates a non-influential parameter and a higher value a more influential parameter. Here, we use SA mainly
740 to identify which parameters have the largest impact on the different model outputs and therefore require careful calibration. As the probability density functions of our model outputs (i.e. the resulting SWI-fluxes) are generally highly-skewed towards extreme organic matter degradation rates (not shown) variance-based sensitivity indices may not be a suitable proxy for output uncertainty (Pianosi et al., 2016). Hence we employ the novel density-based PAWN method by Pianosi and
745 Wagener (2015) which considers the entire conditional and unconditional Cumulative Distribution Function (CDF) of the model output rather than its variance only. The sensitivity index of parameter i is measured by the distance between the two CDFs, i.e.

$$S_i = \max_{x_i} \max_y |F_y(y) - F_{y|x_i}(y)| \quad (47)$$

where $F_y(y)$ is the unconditional CDF of the output y and $F_{y|x_i}(y)$ represents the conditional CDF
750 when the i -th parameter is fixed to the conditioning value x_i . For a more detailed description of the method we refer the interested reader to Pianosi and Wagener (2015).

Due to the model complexity it is impossible to compute the sensitivity indices analytically therefore they are approximated from a Latin-Hypercube sampling of parameter inputs and calculated outputs. The PAWN method, as implemented within the Sensitivity Analysis for Everyone (SAFE)
755 matlab toolbox (Pianosi et al., 2015), is used to investigate 11 model parameters for ranges as specified in Table 11. Sensitivity indices for all resulting SWI-fluxes for two idealised sediment conditions (i.e. anoxic at 400 m and oxic at 4000 m, see Table 12) are calculated. We use 200 samples to estimate the unconditional CDF, 100 samples to estimate the conditional CDFs and 10 conditioning points, thus $200 + 100 \cdot 10 \cdot 11 = 11,200$ evaluations are performed for each sediment condition. The
760 resulting indices are then translated into a color code and summarised in a pattern plot to simplify comparison (Fig. 6).

3.2 Case study: Stand-alone simulations of sediment cores

In order to illustrate the capabilities of OMEN-SED, comprehensive datasets from the Santa Barbara Basin (Reimers et al., 1996), as well as from the Iberian margin and the Nazaré Canyon (Epping
765 et al., 2002) are modelled. Section 5 compares modelled profiles with measured pore water data from different depths including the continental shelf (108 m) and the lower slope (2213 m) located at the

SA: needs polishing

DH: Better like this?

Table 11. Range of model parameters used for sensitivity analysis of model predicted output.

Parameter	Description	Units	Minimum	Maximum	Source
k_1	labile OM degradation constant	yr^{-1}	$1e^{-4}$	5.0	(1)
\tilde{k}_2	order of refractory OM degradation $(k_2 = \tilde{k}_2 \cdot k_1)$	-	$1e^{-4}$	$1e^{-1}$	(1)
f_1	fraction of labile OM	-	0.02	0.98	-
K_{NH_4}	Adsorption coefficient	-	0.8	1.7	(2)
γ_{NH_4}	NH_4 fraction oxidised		0.5	1.0	-
$\gamma_{\text{H}_2\text{S}}$	H_2S fraction oxidised		0.5	1.0	-
$K_{\text{PO}_4}^{\text{ox}}$	Adsorption coeff. oxic	-	100.0	400.0	(3)
$K_{\text{PO}_4}^{\text{anox}}$	Adsorption coeff. anoxic	-	1.3	2.0	(3)
k_s	kinetic P sorption	yr^{-1}	0.1	100.0	(4, 5)
k_m	Fe-bound P release	yr^{-1}	0.015	0.02	(4, 5)
k_a	authigenic P formation	yr^{-1}	0.001	10.0	(4, 6)

Sources: (1) Arndt et al. (2013); (2): Van Cappellen and Wang (1996); (3): Krom and Berner (1980)
(4): Gypens et al. (2008); (5): Slomp et al. (1996); (6): Van Cappellen and Berner (1988)

Table 12. Model boundary conditions for the two idealised sediment conditions used for the sensitivity analysis (Fig. 6 and 7). All solute concentrations are in nanomole cm^{-3} .

Depth (m)	Temp. ($^{\circ}\text{C}$)	OC (wt%)	O_2	NO_3	SO_4	PO_4	z_{bio} (cm)
400	8.0	2.0	0.0	40.0	28,000	40.0	0.001
4000	1.5	1.0	300.0	20.0	28,000	40.0	10.0

Iberian margin, the upper slope (585 m) from the Santa Barbara Basin, and a deep sea site (4298 m) in the Nazaré Canyon. The Santa Barbara Basin is characterised by anoxic bottom waters, high POC concentrations and varved sediments (Reimers et al., 1990), therefore the depth of bioturbation in
770 OMEN-SED is restricted to the upper 0.01 cm. In the uppermost sediments iron(III) hydroxides are reduced, releasing Fe^{2+} which reacts with sulfide to form iron sulfides. Thus, the Fe cycle exerts a strong control on sulfide concentrations in the sediments of this basin (Reimers et al., 1996). In addition, the sediments are generally supersaturated with respect to carbonate fluorapatite by and below 2 cm (Reimers et al., 1996). The Iberian margin, situated in the northeastern Atlantic, generally belongs to the more productive regions of the global ocean (Longhurst et al., 1995), however, seasonal changes in upwelling creates a strong temporal variability in primary productivity and organic carbon deposition and submarine canyons in this area (like the Nazaré Canyon) may deliver organic carbon from the shelf to the ocean interior (van Weering et al., 2002; Epping et al., 2002). For a more detailed description of the study areas and the experimental work, the interested reader
775 is referred to the Reimers et al. (1996) and Epping et al. (2002).

Table 13. Model boundary conditions for the sampling stations in Figure 8. (For all sites DIC bottom water concentration of 2,400 nanomole cm⁻³ is assumed.)

Sediment characteristics:							
Depth (m)	Temp. (°C)	z_{bio} (cm)	D_{bio} (cm ² yr ⁻¹)	POC ₁ (wt%)	POC ₂ (wt%)	k ₁ (yr ⁻¹)	k ₂ (yr ⁻¹)
108	12.5	1.0	0.02	2.64	1.8	0.65	1.0e ⁻⁵
585	5.85	0.01	0.02	2.0	3.5	0.2	8.0e ⁻⁴
2213	3.2	10.0	0.17	0.45	0.5	0.1	4.0e ⁻⁴
4298	2.5	4.2	0.18	0.83	1.2	0.052	1e ⁻⁵

Bottom water concentrations of solutes (all in nanomole cm ⁻³):								
Depth	O ₂	NO ₃	SO ₄	NH ₄	H ₂ S	PO ₄	PO ₄ ^a	Alkalinity
108	210	9.6	28,000	0.4	0.0	0.0	15.0	2,400
585	10	25.0	28,000	0.0	0.0	50.0	90.0	2,480
2213	250	25.0	28,000	0.6	0.0	0.0	5.0	2,400
4298	243	30.1	28,000	0.22	0.0	0.0	5.0	2,400

In OMEN-SED sediment characteristics and boundary conditions are set to the observed values where available (Table 13). Other sediment characteristics (e.g. sedimentation rate, porosity, density), stoichiometric factors and secondary reaction parameters are set to the default value (see Table 9 and 10). Organic matter is modelled as two fractions, with different first-order degradation rate constants. The POC and pore water profiles were fitted by optimizing the POC partitioning into the fast and slow degrading pool and their respective first-order degradation rate constants (priority is given to reproduce the POC and O₂ profiles). For phosphorus the equilibrium concentration for authigenic P formation (PO₄^a) was adjusted to fit the PO₄ concentration at z_∞ .

3.3 Case study: Stand-alone simulations of global ocean transect

In section 5.1 it is tested to which degree OMEN-SED is capable of capturing the dynamics of organic matter degradation pathways and related TEA-fluxes as simulated with a complete, numerical diagenetic model. Therefore, we reproduce the simulations of typical conditions along a global ocean hypsometry of Thullner et al. (2009) and compare our modeled TEA-fluxes with observations from Middelburg et al. (1996). To explore the global remineralisation of OM in the seafloor Thullner et al. (2009) quantified various diagenetic processes using the Biogeochemical Reaction Network Simulator (BRNS, Aguilera et al., 2005), a flexible simulation environment suitable for reactive transport simulations of complex biogeochemical problems (e.g. Jourabchi et al., 2005; Thullner et al., 2005). Thullner et al. (2009) used seafloor depth (SFD) as the master variable and calculated model parameters, such as w , D_{bio} and ϕ , from existing empirical relationships (e.g. Van Cappellen and Wang, 1995; Middelburg et al., 1997). Organic matter degradation is described with a 1-G approach, thus

SA: needs polishing

DH: Better like this?

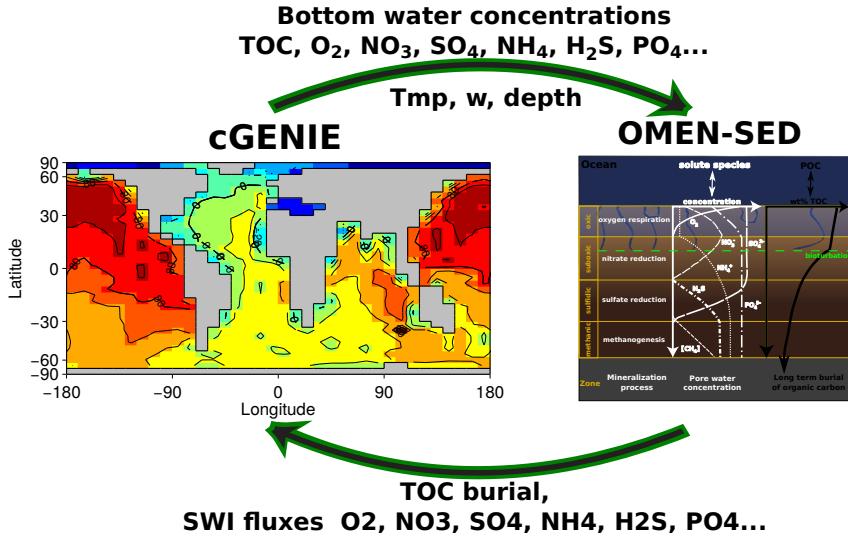


Figure 4. Schematic of the relationship between OMEN-SED and cGENIE. Arrows and accompanied text represent the information transferred between models.

assuming a single pool of organic matter of uniform reactivity. The first order rate constant follows the empirical relationship of Boudreau (1997):

$$k = 0.38 \cdot w^{0.59}. \quad (48)$$

This rate constant can be assumed as the mean reactivity of the organic matter fractions which are degraded in the upper, bioturbated 10 – 20 cm of the sediments. Thus, more reactive fractions (degraded during days/weeks close to the SWI) and more refractory fractions are not captured by this relationship (Boudreau, 1997). BRNS simulations were performed using boundary conditions and parameters for depths representative for shelf, slope and deep sea sediments (i.e. SFD of 100m, 200m, 500m, 1000m, 2000m, 3500m and 5000m). In order to reproduce these results, OMEN-SED is configured as a 1-G model and boundary conditions and model parameters are defined as given in Table 3 of (Thullner et al., 2009). As OMEN-SED assumes a fixed fraction of reduced substances to be reoxidised which exerts a large impact on the resulting SWI-fluxes (compare Section 4) we perform two sets of simulations in order to show the range of possible model outputs. In the first setup 95% of the reduced substances are reoxidised (i.e. $\gamma_{NH_4} = \gamma_{H_2S} = 0.95$) and in the second only 5% are reoxidised (all other model parameters and boundary conditions are equal).

3.4 Case Study: Coupled pre-industrial Earth system model simulations

3.4.1 Coupling to an Earth System Model

OMEN-SED is coupled to the carbon-centric version of the “GENIE” Earth system model (cGENIE Ridgwell et al., 2007) to illustrate the abilities of the newly developed model. The following

820 section provides a brief description of cGENIE and the coupling procedure (Fig. 4). cGENIE is a model of Intermediate Complexity based on the efficient climate model “C-GOLDSTEIN” of Edwards and Marsh (2005), featuring a frictional-geostrophic 3D-ocean circulation model coupled to a fast Energy-Moisture Balance 2D-atmosphere together with a dynamic-thermodynamic sea-ice component. The version of cGENIE used here includes the marine geochemical cycling of carbon, 825 oxygen, phosphorus and sulfur (Ridgwell et al., 2007), preservation of carbonates in deep-sea sediments (Ridgwell and Hargreaves, 2007) and terrestrial weathering (Colbourn et al., 2013). The ocean model is implemented on a 36×36 equal-area horizontal grid with 16 vertical levels using the pre-industrial continental configuration and bathymetry as in Archer et al. (2009). A finer grid (72×72) is used for the sediments. Instead of completely degrading POC at the seafloor, OMEN-SED is called 830 by SEDGEM for each wet ocean grid point. Depending on the overlying biogeochemical ocean model, processes can be included or excluded in OMEN-SED and stoichiometric factors need to be adjusted to ensure preservation of mass. As nitrogen is not modelled explicitly in the employed cGENIE configuration the related stoichiometries in OMEN-SED are set to zero (i.e. NC_i , ALK^{NIT} and ALK^{DEN}). cGENIE, however, implicitly includes the effects of NH_4 release and its nitrification 835 on Alkalinity and neglects the impact of P release; therefore, related stoichiometries are changed to $ALK^{OX} = -16/106$ and $ALK^{SUL} = 122/106$.

Several biogeochemical tracers and parameters are transferred from SEDGEM to OMEN-SED and have to be converted into the required units. Bottom water concentrations of solutes are converted from $mol\ kg^{-1}$ to $mol\ cm^{-3}$ and the depositional flux of POC (POC_{flux}) is converted from 840 $cm^3\ cm^{-2}\ yr^{-1}$ to $mol\ cm^{-2}\ yr^{-1}$ assuming an average density of POC of $1.0\ cm^3\ g^{-1}$. Other parameters used from cGENIE are seafloor depth, local temperature and the partitioning of bulk POC into the slower and faster degrading pool (as cGENIE represents a labile and a refractory POC fraction, see Ridgwell et al., 2007). The advection/sediment accumulation rate is generally taken from cGENIE, however, a minimum value of $w = 0.5\ cm\ kyr^{-1}$ is imposed as OMEN-SED tends to be 845 unstable for lower values. The bulk POC_{flux} is separated into the labile and refractory component and the routine to calculate the sedimentary POC profiles is called. Here, the two POC depositional fluxes are first converted into SWI concentrations (in $mol\ cm^{-3}$) by solving Eq. (2) for $z=0$. OMEN-SED computes the resulting SWI-fluxes of solutes (in $mol\ cm^{-2}\ yr^{-1}$) and the fraction of POC preserved in the sediment at a depth of 100 cm (POC_{pres}) and returns the results to cGENIE. In case no 850 POC is deposited on the seafloor (i.e. $POC_{flux} = 0$), OMEN-SED is not called and the SWI-fluxes of solutes and POC_{pres} are set to zero. In order to reduce memory requirements the sediment profiles (e.g. as shown in Fig. 8) are not calculated in the FORTRAN version of OMEN-SED, however, the boundary conditions are saved at the end of the experiment and sediment profiles for specific grid-cells, ocean basins and/or ocean transects can be plotted using the stand-alone MATLAB version of 855 OMEN-SED

3.4.2 Coupled pre-industrial simulations

All simulations presented in section 5.2 are run for 20,000 years to steady-state and OMEN-SED is called for each grid-cell in every time step, feeding back the resulting SWI-fluxes and the fraction of POC preserved in the sediments to cGENIE. As shown in our sensitivity analysis (Section 4) and 860 discussed by Arndt et al. (2013) the degradation rate constants for OM (k_i) are the most influential parameters and strongly determine the SWI-flux of redox-sensitive elements as well as the preservation of organic matter. Yet, their spatial variability is unknown at the global scale and reported rate constants can vary by almost 10 orders of magnitude (Arndt et al., 2013), thus defining appropriate OM degradation rate constants is a major challenge and source of uncertainty for diagenetic models. 865 The rate constants in models are either determined through profile fitting for a specific site or, for global applications, they are related to a single, readily available characteristic (or master variable) of the local environmental conditions. For instance, considerable effort has been expended to relate the apparent rate constant for oxic and anoxic OM degradation to sedimentation rate (w) and various empirical relations have been proposed (Toth and Lerman, 1977; Tromp et al., 1995; Boudreau, 870 1997; Stolpovsky et al., 2015). Nevertheless, these relationships are generally based on limited data sets and their global applicability is questionable (Arndt et al., 2013).

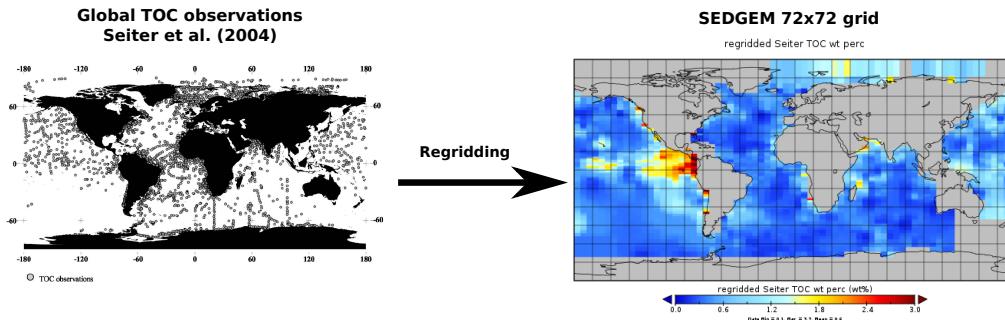


Figure 5. TODO: Update with new regridded observations using MATLAB. Global distribution of TOC observations (N=5500) from the database of Seiter et al. (2004, left) and regridded observations onto the SEDGEM grid (right). Grid points without any observations are left blank (white).

In section 5.2, we test globally invariant values for the OM degradation rate constants (k_1, k_2) as well as the empirical relation of Boudreau (1997) which is also used in Section 5.1. The modelled mean POC content in the upper 5cm of the sediments is compared to the global POC observations 875 in surface sediments (< 5cm sediment depth) of Seiter et al. (2004). The original data field of Seiter et al. (2004) has been transformed onto the 72×72 SEDGEM grid (Figure 5). The relationship proposed by Boudreau (1997) relates the apparent OM degradation rate constant in the bioturbated

layer to the burial velocity, w (cm year $^{-1}$):

$$k_{\text{app}} = 0.38 \cdot w^{0.59}. \quad (49)$$

- 880 We assume that k_{app} represents the mean OM reactivity in the bioturbated layer and make the following assumptions in order to calculate the two degradation rate constants for OMEN-SED:

$$k_{\text{app}} = f_1 \cdot k_1 + f_2 \cdot k_2 \quad (50)$$

$$k_2 = k_1/x \quad (51)$$

- 885 where x describes the relation between k_1 and k_2 which is subject to sensitivity experiments in Section 5.2 (with values of $x \in \{2, 4, 5, 8, 10, 25, 50, 100, 125\}$). As the fractions of labile and refractory OM reaching the sediments (f_1, f_2) is known from cGENIE we can calculate k_1 and k_2 independently for each grid-cell. For the globally invariant degradation rate experiment in Section 5.2, k_2 is systematically varied between 0.0002 and 0.005 year $^{-1}$ and the more labile OM component, de-
890 scribed by k_1 , is assumed to degrade 2, 4, 5, 8, 10, 20 and 50 times faster, respectively.

4 Sensitivity Analysis

- Fig. 6 summarises results of the sensitivity analysis (compare Section 3.1) as a colour map. Results indicate that the most significant parameters for all model outputs are the degradation rate constant for the labile OM part (k_1) and the fraction of the labile pool to the total OM pool (f_1). **also need to
895 be careful how we formulate it, as you made k2 dependant on k1. TODO: -frst shortly describe the main results, e.g. fluxes are most sensitivity to variations in OM degradation parameters, other parameters play a secondary role
the describe which other parameters are red and why
then describe difference odic-anoxic**
- 900 **Overall: needs more critical analysis!** For the anoxic setup the secondary redox parameters (i.e. γ_{NH_4} , $\gamma_{\text{H}_2\text{S}}$) are essentially non-influential. Whereas in the oxic scenario, SWI-fluxes of NH_4 , SO_4 and H_2S are very sensitive to changes in the secondary redox parameters. The PO_4 SWI-flux appears to be insensitive (TODO: to which parameters specifically) especially for the oxic condition as the majority is absorbed to Fe-oxides. The sensitivities change if other PO_4 related equilibrium
905 concentrations PO_4^s , PO_4^a and M^∞ are used (not shown).

- We further explore the sensitivity of simulated sediment-water exchange fluxes to variations in organic matter degradation parameters by varying k_1 , f_1 and \tilde{k}_2 while all other model parameters are set to their default values (Tables 9 and 10). For the deep sea condition we account for the presence of more refractory OM by sampling $f_1 \in [0.02, 0.3]$. Minimum and maximum values for k_1 , \tilde{k}_2 and
910 f_1 in the shallow ocean are as in Table 11. The parameter space is sampled using another Latin-Hypercube approach with sample sizes of $N = 3500$ for each idealised sediment condition. Figure

DH: explanation for small P sensitivity correct?

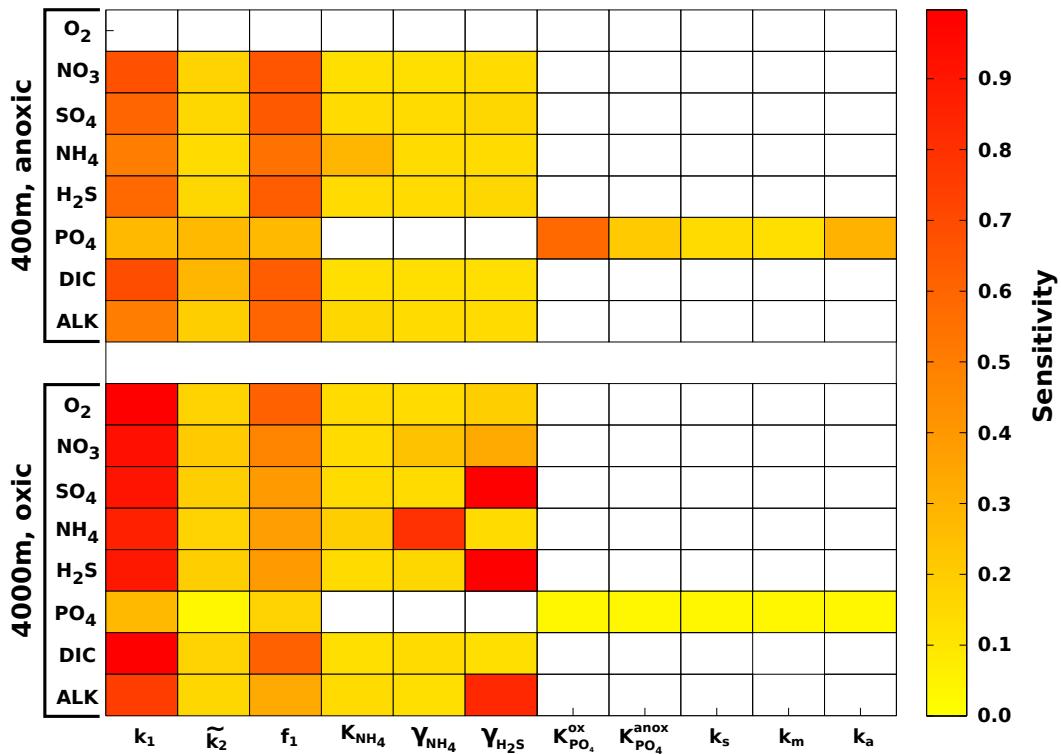


Figure 6. Pattern plot, showing the output sensitivity for each SWI flux (i.e. the chemical compounds on the vertical axis) and each input factor (i.e. the model parameters on the horizontal axis) for two idealised sediment cores. White patterns are assigned where the SWI flux is independent of the specific parameter.

DH: Rather low PO_4 sensitivity - bc of Equil. concentr.?

7 summarises the results of the sensitivity study. In addition, the range of observed O_2 and NO_3 sediment-water interface fluxes extracted from a global database (Bohlen et al., 2012) is indicated.

TODO: needs a more critical analysis, what is the message you want the reader to understand

915 **here?** Figure 7 shows that the observed fluxes fall well in the range of SWI-fluxes calculated with OMEN-SED. Also highlighted by the emergence of colour patterns in Figure 7 A+B are the strong interactions between the amount of labile OM f_1 and its degradation rate k_1 for the resulting SWI-fluxes of the most powerful TEA available. In general, a higher degradation rate in combination with more labile OM available leads to a higher SWI-flux.

920 5 Sediment profiles

needs more critical analysis

state that the model is able to capture main features across a wide range of different environments

-quickly explain the dynamics at each location- e.g. what are the dominant processes and how

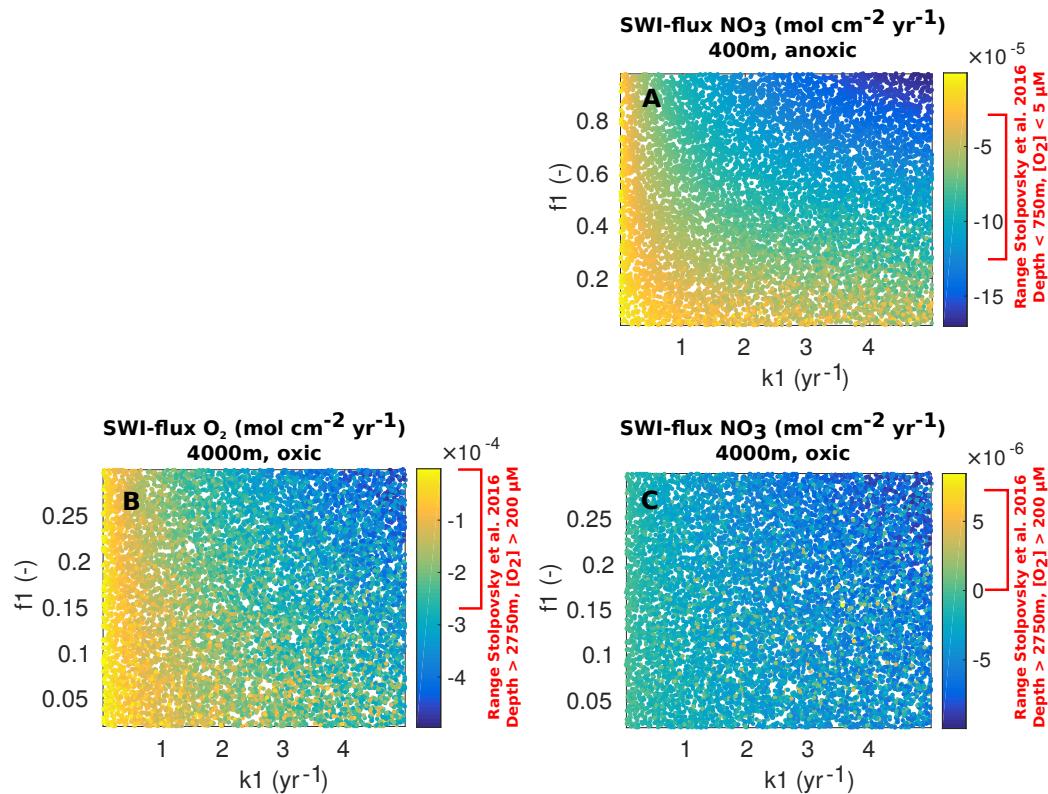


Figure 7. Coloured scatter plots (k_1 vs f_1) of resulting OMEN-SED SWI-fluxes for the 400m anoxic (A: NO_3) and 4000m oxic (B: O_2 , C: NO_3) scenario. Negative values representing a flux from the water column into the sediments. Indicated area in red at the respective colour scale represents the range of benthic fluxes in the global database of Bohlen et al. (2012).

DH: Why mainly negative NO_3 -flux 4000m in contrast to database!?... might change when using diff. gamma

925 **does OMEN do in capturing these? -point out where it does well**

-then point out where it does not sooo well and explain why

For the two open Iberian margin stations (108 and 2213 m) OMEN-SED fits all observations well. OMEN-SED does especially well at depth 2213m by reproducing the deep O_2 penetration and the subsurface maximum in NO_3 concentration due to the nitrification of NH_4 . For the Santa Barbara basin (585 m) a misfit is observed for H_2S and PO_4 in the upper 20 cm of the sediment. This can be explained by the presence of Mn^{2+} , Fe^{2+} and dissolved Fe at this site which are either reduced to degrade POC and/or react with H_2S to form iron sulfides, therefore inhibiting the rise in concentration of H_2S (Reimers et al., 1996). Phosphorus is adsorbed to Fe oxides and incorporated into carbonate fluorapatite (CFA) which is highly parameterised in OMEN-SED and not modelled explicitly.

Explanation later for high H₂S in OMEN from Reimers (also check abstract): this maybe in discussion section:

Fe cycling exerts a strong control on the dissolved sulfide concentrations in the uppermost sediments of the SBB

940 *for PO₄: sediments become saturated with respect to carbonate fluorapatite and calcite within the first 0.25 mm and are highly supersaturated by and below 2 cm*

from conclusion: for H₂S and PO₄: SBB conforms to a model of mineral paragenesis wherein CFA accumulation is restricted to the near-surface of organic-rich sediments. In such surface sediments, a critical factor is that iron oxyhydroxides are reduced, which raises pH and re-leases sorbed phosphate, fluoride, and Fe^{+}. The Fe^{*+} is then available to sequester dissolved sulfide into FeS, leading to pyrite formation. Also, colloidal precipitates of iron or sulfur, or fresh bacterial products are suggested to supply surfaces which may initiate CFA precipitation.*

Generally, sulfide is undetectable in the uppermost 2-4 cm of sediment (detection limit 0.5 microM), and sulfide concentrations do not begin to increase rapidly until depths where dissolved Fe is absent

950 *(Reimers et al., 1990). Porewater sulfide concentrations are below detection in spite of the presence of high rates of sulfate reduction in the first 0-2, and 2-4 cm of sediment (Fig. 5). In this zone, Fe is clearly being reduced and reacting with sulfide (primarily in the 2-4 cm interval) to form iron sulfides represented as acid-volatile S and pyrite...*

For the PO₄ minium - from Reimers: Ruttenberg and Bemer (1993) suggest diagenetic redistribution of organic P into CFA may be a nearly continuous process at least to 50 cm in terrigenous dominated coastal sediments.

The NW Iberian Margin ranks among the more productive areas of the global ocean (Longhurst, Sathyendranath, Platt, & Caverhill, 1995)... The equatorward slope/shelf current during summer 960 is generally associated with the development of a band of cool upwelled North Atlantic Central Water (Fraga, 1981; Van Aken, 2001), which results there being high primary production along the Iberian Margin.... The pattern of upwelling along the Iberian Margin is intermittent generating strong temporal variability in primary productivity and benthic deposition of carbon.

Site 108m - 99-06:

965 *The concentrations of reduced iron and manganese at the sandy station were less, possibly the consequence of lower deposition of metal-oxides at this site*

Site 2213m - 99-04:

could have been improved with a change of γ_{NH_4} *The pore water profiles show that there was a much deeper penetration of oxygen and nitrate compared to the shelf stations, whereas ammonium and dissolved iron and manganese did not accumulate above a depth of 12 cm. Conceivably, carbon oxidation at this station is dominated by aerobic oxidation so that at this site denitrification and*

anaerobic oxidation appear to be of minor importance. Ammonium, released during organic matter mineralization is effectively nitrified to nitrate as evidenced by the subsurface maximum in nitrate concentration, indicating a tight coupling between nitrification and denitrification at this station.

975 Site 4298m (canyon) - 99-17:

At 4298 m, however, sandy layers are intercalated with layers of high silt content, which is reflected in the erratic, but analytically highly reproducible distribution of sediment organic carbon. These intercalations hint at the possibility of episodic deposition events occurring over the fan of the canyon.

980 For the canyon, a problem arises because the oxygen microprofiles were probably not being influenced by non-local exchanges because of the extremely small volume of pore water that is sensed by the tip of the electrode, whereas the profiles of ammonium and nitrate are. Secondly, OMEXDIA does not include a non-local transport routine

probably not important: The accumulation of reduced manganese and iron in the pore water as well as their declining concentrations down core, point to an intensive cycling of metal-oxides and at a significant precipitation of sulfides and carbonates in these sediments

better fits could have been potentially acquired by using a sensitivity study over the C/N used in OMEN-SED: Epping: of showed C/N ratios exceeding Redfield stoichiometry

Canyon: These intercalations hint at the possibility of episodic deposition events occurring over the fan of the canyon... For the canyon, a problem arises because the oxygen microprofiles were probably not being influenced by non-local exchanges because of the extremely small volume of pore water that is sensed by the tip of the electrode, whereas the profiles of ammonium and nitrate are. Secondly, OMEXDIA does not include a non-local transport routine For the Nazaé Canyon station (4298 m) satisfactory fits could be realised apart from NH₄. However, also the original study (Epping et al., 2002) had the same problem using a more complex diagenetic model and suggested non-local solute exchange being responsible for the higher NH₄ concentrations at this site. ... this is bioirrigation (see Epping) ...

DH: P explanation
correct/okay?

Also, say something
about [ALK] at 585m?

5.1 Global Hypsometric Analysis of TEA Fluxes

Figure 9 compares simulated SWI-fluxes of TEAs (i.e. O₂, NO₃ and SO₄) along the global hypsometry using OMEN-SED (black lines) with the results of Thullner et al. (2009) (red lines). Observations for O₂ and NO₃ fluxes are taken from Middelburg et al. (1996). Also plotted in Fig. 9A are the total oxygen uptake (TOU) estimates presented in Thullner et al. (2009) (filled red symbols), who assumed the organic matter flux to be equivalent to TOU.

quickly describe what we see, i.e. fluxes decrease along the global hypsometry and mention 1005 why

also describe and explain the differences between the OMEN min-max simulations

Results show that OMEN-SED captures the main trends in observed TEA fluxes well. In general, fluxes calculated with BRNS fall within the range of possible OMEN-SED results.

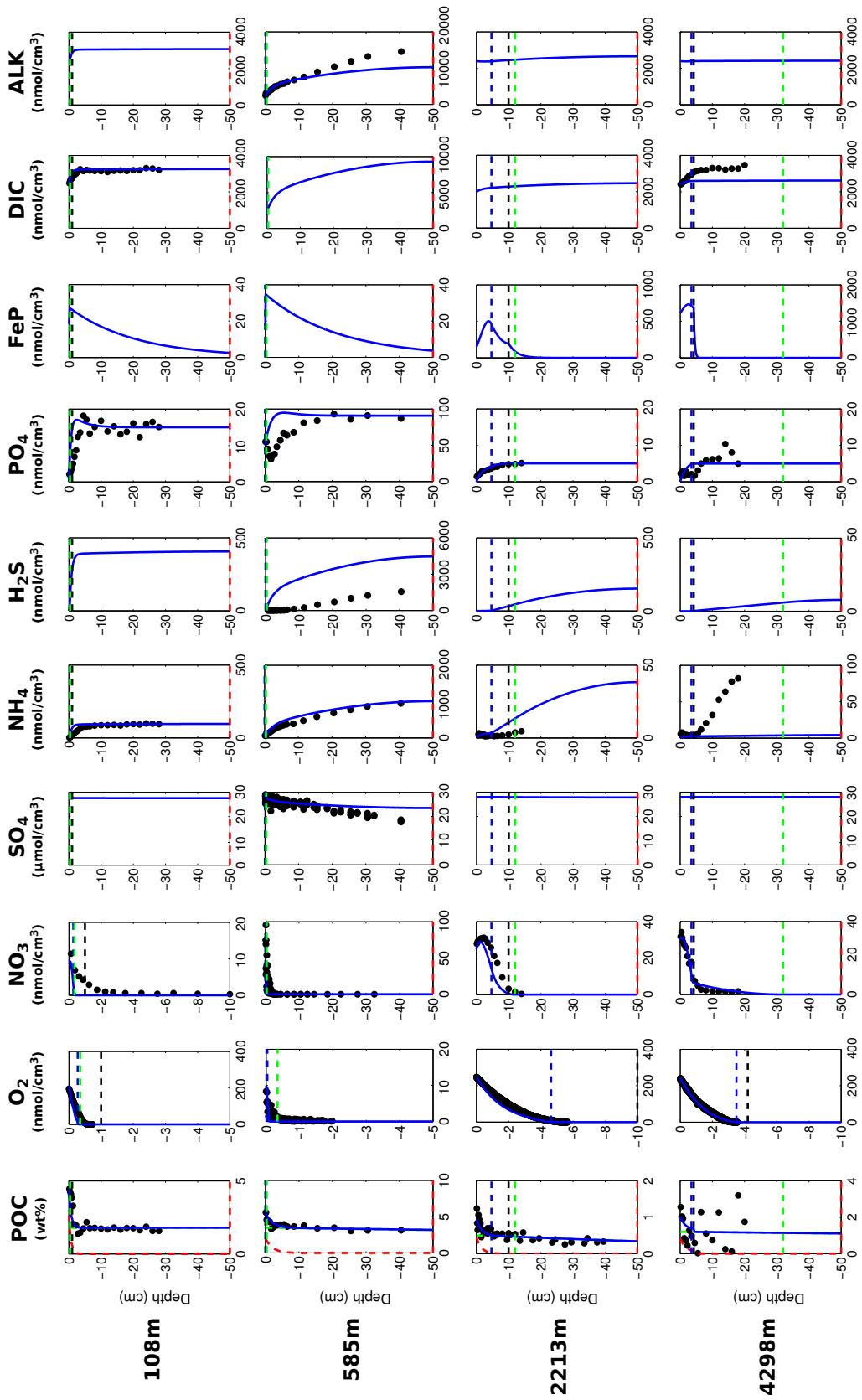


Figure 8. Modelled (curves) and measured (filled dots) dissolved and solid phase pore water profiles for four different sediment cores. Note that different scales are used for different stations. The blue POC curve represents the sum of the refractory (green) and labile (red) POC fraction.

In particular, the observed O₂ fluxes in the upper 2000m are well predicted by the two OMEN-
1010 SED simulations (Fig. 9A). Oxygen fluxes for the deep-sea sediments, however, are slightly under-
estimated. These deviations can presumably be related to the assumed 1-G description of organic
matter degradation, which neglects the more labile OM pool. This highly reactive pool is degraded
close to the sediment surface, thus promoting higher aerobic degradation rates and higher O₂ fluxes.
Nitrate fluxes in the upper 500m of the Atlantic Ocean are well predicted. However, as in Middel-
1015 burg et al. (1996) the direction of calculated nitrate fluxes in the upper 1000m of the Pacific Ocean
differ from the observations. Middelburg et al. (1996) related these discrepancies to the globally
averaged model parameters and the applied boundary conditions. They could reduce the disagree-
ments significantly by using more representative bottom water concentrations for the eastern Pacific
1020 and a higher flux of labile organic matter for their 2-G model. By changing the boundary conditions
and the NC-atomic ratio of organic matter for the whole hypsometry, it is possible to obtain a better
model-data fit with OMEN-SED for the shallow Pacific Ocean (green line in Fig. 9B). Following
Bohlen et al. (2012), who could show that the atomic NC-ratio strongly deviates from Redfield sto-
chiometry (0.151) with specifically lower values for the East Pacific Ocean, we adopt their globally
1025 averaged value of 0.067. In addition, bottom water conditions are changed to low oxygen/high nitrate
levels more likely to be found in the shallow Pacific Ocean (O₂ = 10 nanomole cm⁻³ and NO₃ = 80
nanomole cm⁻³).

5.2 Pre-industrial cGENIE coupling and the OM degradation rate

OMEN-SED has been coupled to the global Earth system model cGENIE as described in Section
??.

1030 Describe Figure 13 and compare with what we know. E.g. Thullner et al. 2009!

Schulz and Zabel: The basic mechanism inducing microbial activity is the supply of organic matter
to the seafloor and this is generally coupled to surface water produc- tivity. Most of the highly
productive areas in the global ocean are adjacent to the continents, so that we can expect a decrease
1035 of degradation inten- sity from the coastal marine environments over the continental shelves and
slopes into the deep- sea. This becomes evident when we look at the data compiled by Middelburg
et al. (1993) which indicate that 83% mineralization and 87% burial in marine sediments occurs in
the coastal zone occupying only 9% of the total ocean area. This means that the sediments with
the highest respi- ration rates also have the highest burial efficiency in marine environments. Fluxes
1040 of oxygen and nitrate, therefore, vary over several orders of magnitude between oligo- trophic open
ocean areas and continental shelf and slope areas. This is about 50 to 6,000 mmol m⁻² yr⁻¹ for
oxygen and -600 to 380 mmol m⁻² yr⁻¹ for nitrate ...

(SHOW O₂, NO₃ AND E.G. PO₄ SWI-fluxes).

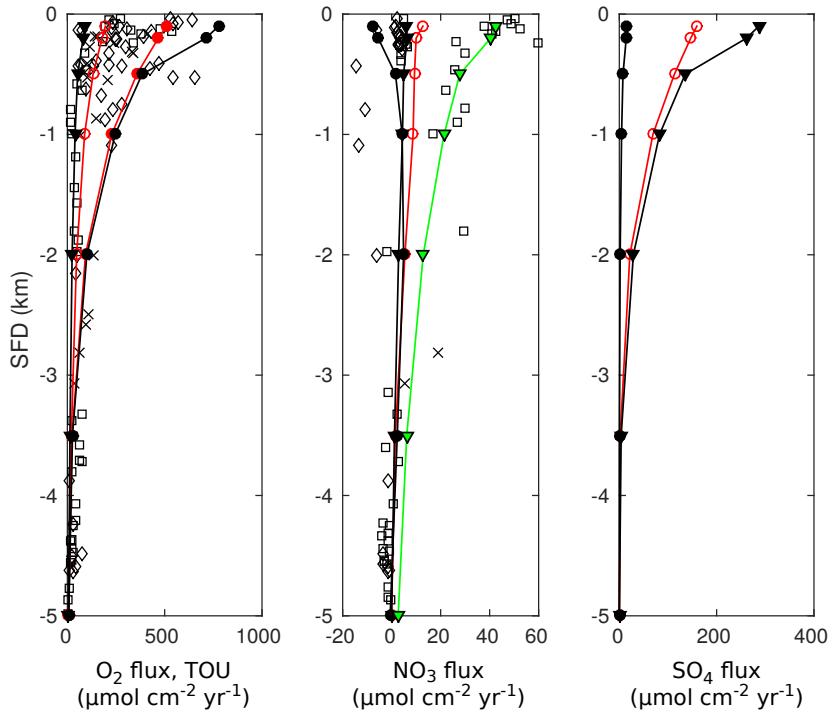


Figure 9. Fluxes of O₂, NO₃ and SO₄ to the sediment along the global hypsometry. Red lines (with open symbols) are modelled fluxes from Thullner et al. (2009) using BRNS; black lines are results from OMEN-SED (● : $\gamma_{NH_4} = \gamma_{H_2S} = 0.95$; ▼: $\gamma_{NH_4} = \gamma_{H_2S} = 0.05$). Observations of TEA fluxes are taken from Middelburg et al. (1996) (◊: Atlantic, □: Pacific, ×: Arctic/Indian Ocean). Also plotted in Figure (A) are the total oxygen uptake (TOU) estimates of Thullner et al. (2009) (filled red symbols). The green line indicates OMEN-SED results for low oxygen/high nitrate levels and the lower NC-ratio. Negative values are directed out of the sediments.

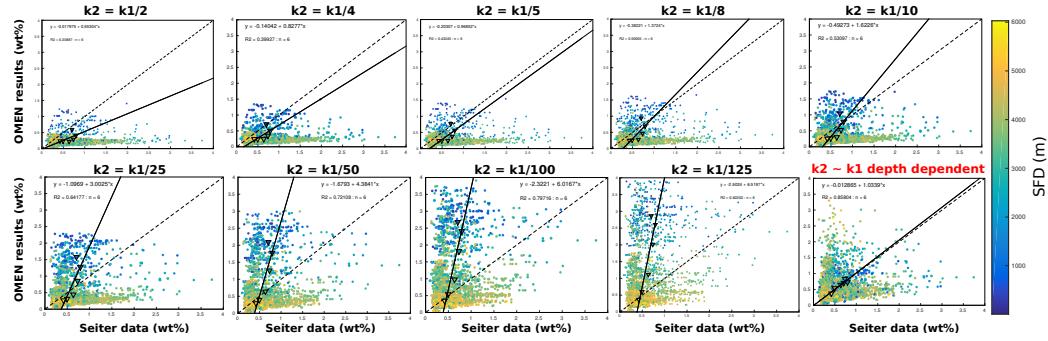


Figure 10. TEXT TEXT TEXT TEXT

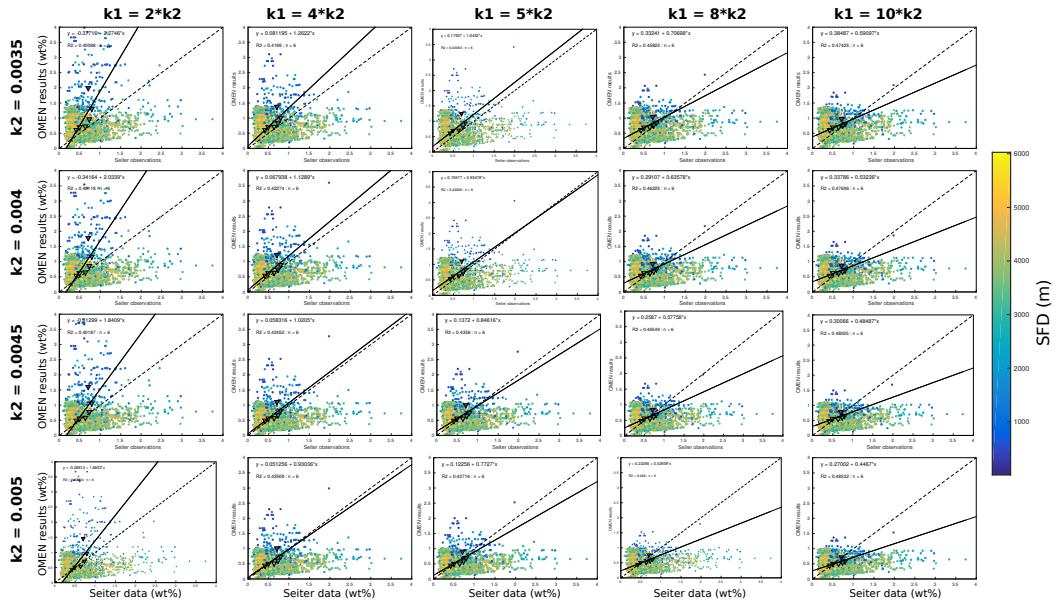


Figure 11. TEXT TEXT TEXT TEXT

Could also plot Figure 4 from Palastanga et al. (2011) with POC wt% for different sites in the
1045 ocean! UNITS of Fe-P! This is a solid! Therefore not nmol/cm³ rather mol/g as in Palastanga...

Also check what water column features to show and maybe sediment P or Fe-P concentrations
(see PALASTANGA et al. 2011, 2013).

Stolpovsky: bioturbated sediments deposited at continental margins quickly become anoxic within
1050 a few millimeters [Wenzhöfer and Glud, 2002].

up until here

6 Scope of applicability and model limitations

State-of-the art numerical models representing the full complexity of the diagenetic processes typically perform adequately at reproducing site-specific biogeochemical dynamics, however, tuning
1055 model parameters is laborious, the computational demand is high and, thus, their transferability to the global scale is limited. On the other hand, analytical models are very efficient, but existing approaches coupled to global models generally use highly simplified reaction networks, often restricted to oxic degradation with a limited number of explicit pore water tracers. However, our ability to assess the role of organic matter dynamics for global biogeochemical cycles and climate requires
1060 tools that resolve the most important biogeochemical processes and tracers explicitly, while at the same time are computationally efficient and have a degree of predictive capability to extrapolate

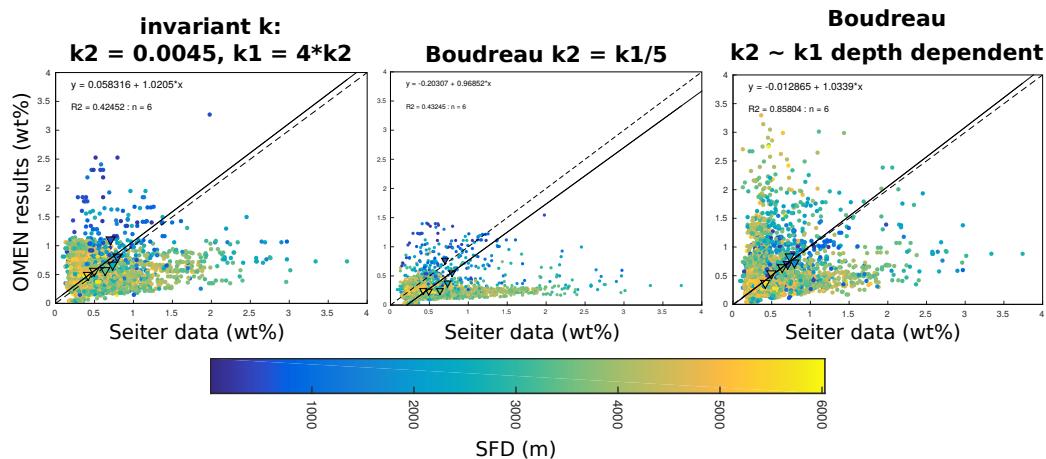


Figure 12. TEXT TEXT TEXT TEXT

knowledge to data poor areas. The new model OMEN-SED presented here is a legitimate compromise between complexity of biogeochemical processes and computational efficiency. Its scope of applicability covers the entire range from regional to global scales. OMEN-SED's computational 1065 efficiency facilitates its use in two very different ways. Firstly, it can be coupled to global Earth system models and therefore allows the investigation of coupled global biogeochemical dynamics over different timescales. Secondly, it can be used to calculate quantitative sensitivity indices requiring large sample sizes such as variance- or density-based approaches. Therefore, OMEN-SED can help 1070 to quantitatively investigate how systematic variations in model parameters impact the model output, for instance when the model has been tuned to a site-specific problem. Due to the represented anaerobic processes and secondary-redox reactions, OMEN-SED is also useful to investigate the role of benthic-pelagic coupling on the development of ocean anoxia and euxinia for instance during extreme climate events such as OAEs. On more regional scales it can be applied to systems like continental margins or estuaries which are characterised by complex interactions between different 1075 pathways of organic matter degradation and redox reactions. Here, the model can help to disentangle the complex process interplay that drives the biogeochemical dynamics and give quantifications for upper and lower constraints of carbon and nutrient budgets for these dynamic systems. In addition, OMEN-SED can be used to model eutrophication processes in shallow coastal waters as sediment-water oxygen and nutrient exchange fluxes are explicitly modelled and depend on reoxidation of 1080 reduced substances which causes a substantial part of oxygen consumption in these environments.

However, the model presented here, even more complete than previous analytical models, is still associated with a certain degree of simplifications. In order to solve the diagenetic equation analytically important assumptions have been made, which limit the general applicability of the model. One of the most important simplifications is assuming steady-state. When coupled to an Earth sys-

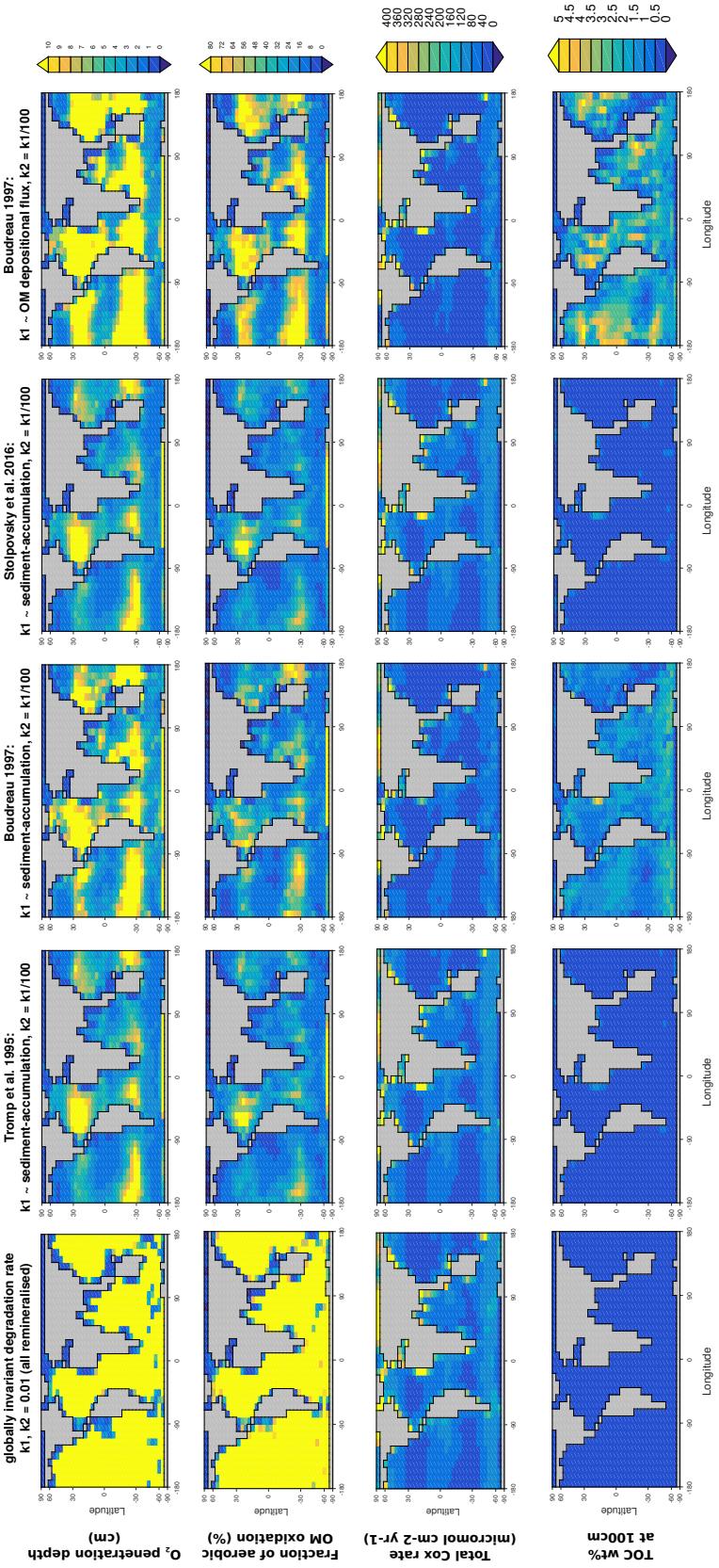


Figure 13. Results of OMEN-cGENIE coupling using different, published parameterisations for the OM degradation rates (k_1 , k_2 , compare Table ??) coupled to the same cGENIE ocean setup. All results shown are sediment characteristics calculated by OMEN-SED. **NOTE:** Preliminary results (here frac2 still constant!) **TODO:** plot wt% just until 3%.

1085 tem model this assumption is only valid if the relevant variability in boundary conditions and fluxes
is generally longer than the characteristic timescales of the reaction-transport processes. In that case
the sediment column can be described by a series of pseudo steady-states as it is done in OMEN-
SED. Consequently, the model can be used for investigating the long term effects of changes in
boundary conditions such as the the input of OM or bottom water oxidation state on degradation
1090 and burial dynamics, as for instance during OAEs. Yet, it is not able to predict the system response
to short-term or seasonal variations of boundary conditions. The separation of the sediment column
into distinct biogeochemical zones and the resulting lack of overlap in degradation pathways may
cause distorted organic matter degradation rates for the different TEAs. For instance, in OMEN-SED
denitrification does not occur in the oxic zone, while in reality, although inhibited by the presence of
1095 oxygen, denitrification can still occur in the oxic zone, even at shallow sediments depths with high
OM contents. Manganese and iron are not represented and as such OMEN-SED is not able to model
all processes important in coastal marine environments and highly accumulating upwelling regions.
This can cause problems when modelling H₂S and PO₄ profiles in anoxic environments as their
concentrations are affected by these metal ions (compare Section 5). In addition the depth invariant
1100 porosity limits the correct calculation of the sediment-water interface flux of dissolved species as in
reality porosity decreases with sediment depth.

7 Conclusions

In this paper we have described and tested a new, analytical early diagenetic model resolving organic
matter cycling and associated biogeochemical dynamics called OMEN-SED. Our new model is the
1105 first of this class of analytical approaches to explicitly represent oxic degradation, denitrification,
sulfate reduction and implicitly methanogenesis, as well as the reoxidation of reduced substances
produced during organic matter degradation. Pore water tracers include O₂, NO₃, NH₄, SO₄, H₂S,
DIC and ALK and the solid phase includes two fractions of organic matter, Fe-bound P and authi-
genic Ca-P minerals. We have shown that the new analytical model is able to reproduce observed
1110 pore water profiles from different ocean depths when organic matter partitioning and degradation
rate constants are tuned to site specific conditions. An extensive sensitive analysis, based on the
novel density-based PAWN method (Pianosi and Wagener, 2015), has been performed to asses the
importance of 11 internal model parameters for all resulting SWI-fluxes. The results reveal that the
degradation rate constant for labile organic matter is the most influential parameter for all model
1115 outputs. Under anoxic conditions secondary redox parameters exert an important control on related
SWI-fluxes of SO₄, H₂S, NH₄ and alkalinity. In addition, the sensitivity analysis showed that glob-
ally observed benthic O₂ and NO₃ fluxes fall well into the range of produced model results.

Furthermore, the coupling of OMEN-SED to the Earth system model cGENIE is described and
various published parameterisations for the apparent first order degradation rate constant are tested

1120 to fit the model to observed global organic matter concentrations. **TODO: write about results.** ...in
order to investigate the role of sediments for global biogeochemical cycles and climate.

8 Code Availability

The OMEN-SED source code (Fortran and Matlab) related to this article is provided as a supplementary package together with a ReadMe file, where hardware and software requirements, source
1125 code files and model output file management are fully described.

Appendix A: Reaction Network

Appendix B: Sensitivity Analysis

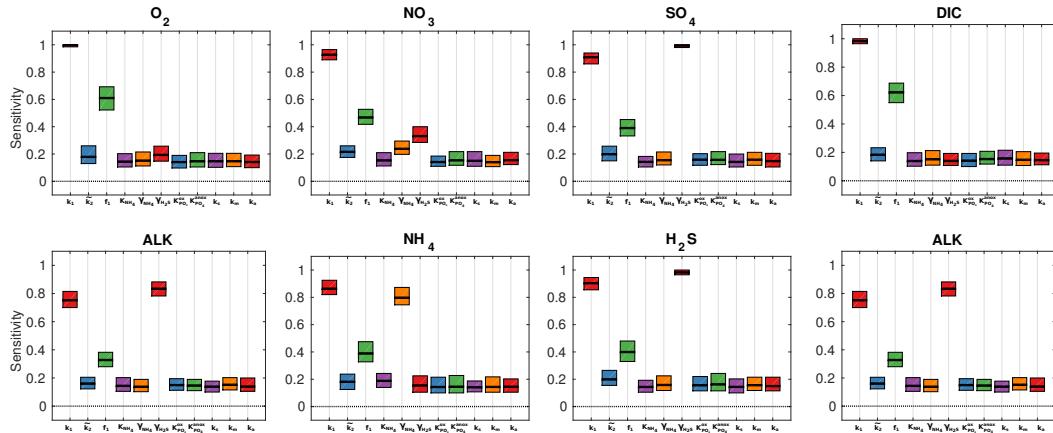


Figure 14. Move to Appendix Box plot of parameter sensitivities for the calculated SWI-fluxes for the 4000m oxic condition. Average sensitivities (black lines) and 90% confidence intervals using $N = 11200$ model evaluations and $N_{boot} = 100$ bootstrap resamples.

B1

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Table 14. Primary pathways of organic matter degradation, secondary redox reactions and stoichiometries implemented in the reaction network.

Pathway	Stoichiometry
Primary Redox reactions	
Aerobic degradation	$(\text{CH}_2\text{O})_x(\text{NH}_3)_y(\text{H}_3\text{PO}_4)_z + (\text{x}+2\text{y})\text{O}_2 + (\text{y}+2\text{z})\text{HCO}_3^- \rightarrow (\text{x}+\text{y}+2\text{z})\text{CO}_2 + \text{yNO}_3^- + \text{zHPO}_4^{2-} + (\text{x}+2\text{y}+2\text{z})\text{H}_2\text{O}$
Denitrification	$(\text{CH}_2\text{O})_x(\text{NH}_3)_y(\text{H}_3\text{PO}_4)_z + \frac{(\text{4x}+3\text{y})}{5}\text{NO}_3^- \rightarrow \frac{(2\text{x}+\text{3y})}{5}\text{N}_2 + \frac{(\text{x}-3\text{y}+10\text{z})}{5}\text{CO}_2 + \frac{(\text{4x}+3\text{y}-10\text{z})}{5}\text{HCO}_3^- + \text{zHPO}_4^{2-} + \frac{(\text{3x}+6\text{y}+10\text{z})}{5}\text{H}_2\text{O}$
Sulfate reduction	$(\text{CH}_2\text{O})_x(\text{NH}_3)_y(\text{H}_3\text{PO}_4)_z + \frac{\text{x}}{2}\text{SO}_4^{2-} + (\text{y}-2\text{z})\text{CO}_2 + (\text{y}-2\text{z})\text{H}_2\text{O} \rightarrow \frac{\text{x}}{2}\text{H}_2\text{S} + (\text{x}+\text{y}-2\text{z})\text{HCO}_3^- + \text{yNH}_4^+ + \text{zHPO}_4^{2-}$
Methanogenesis	$(\text{CH}_2\text{O})_x(\text{NH}_3)_y(\text{H}_3\text{PO}_4)_z + (\text{y}-2\text{z})\text{H}_2\text{O} \rightarrow \frac{\text{x}}{2}\text{CH}_4 + \frac{\text{x}-2\text{y}+4\text{z}}{2}\text{CO}_2 + (\text{x}-2\text{z})\text{HCO}_3^- + \text{yNH}_4^+ + \text{zHPO}_4^{2-}$
Secondary Redox reactions	
Nitrification	$\text{NH}_4^+ + 2\text{O}_2 + 2\text{HCO}_3^- \rightarrow \text{NO}_3^- + 2\text{CO}_2 + 3\text{H}_2\text{O}$
Sulfide oxidation	$\text{H}_2\text{S} + 2\text{O}_2 + 2\text{HCO}_3^- \rightarrow \text{SO}_4^{2-} + 2\text{CO}_2 + 2\text{H}_2\text{O}$
AOM	$\text{CH}_4 + \text{CO}_2 + \text{SO}_4^{2-} \rightarrow 2\text{HCO}_3^- + \text{H}_2\text{S}$
Adsorption reactions and mineral precipitation	
NH ₄ adsorption	$\text{NH}_4^+ \xrightarrow{K_{\text{NH}_4}} \text{NH}_4^+ \text{ (ads)}$
P ad-/desorption ???	$\text{PO}_4^{2-} \xrightarrow{K_{\text{PO}_4}^{\text{I},\text{II}}} \text{PO}_4^{2-} \text{ (ads)};$
CFA precipitation	$\text{PO}_4^{2-} \xrightarrow{k_a} \text{CFA}$

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