

Advanced Thermodynamics

Homework Assignment #1

Assignment Overview

Problems 1-14: Fundamental thermodynamic concepts, energy balances, entropy analysis, steam turbines, Carnot cycles, and concentrated solar power optimization

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Computational Tools Developed

- CoolProp CLI for multi-fluid property lookup
- Property solver for iterative thermodynamic calculations
 - Carnot cycle optimization analysis
 - Steam turbine performance analysis

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1.

It is assumed the system is a simple compressible system, therefore there are two degrees of freedom and the rest of the intensive properties can be calculated using an equation of state, or found in a table of common fluids.

2.

- Closed system: (a)
- Open system: (c)
- Isolated system: (b)

3.

a. Internal energy (U): The microscopic kinetic and potential energy of the molecules within the mass.

b. Kinetic energy (KE): Energy associated with the macroscopic velocity of the mass, given by $\frac{1}{2}mv^2$.

e. Potential energy (PE): Energy associated with the position of the mass in a gravitational or other force field, given by mgz .

Notes:

- f. Electrical energy could potentially be considered if the mass crossing a system boundary is electrically charged, but it would be a rare circumstance.
- e. Flow work is a summation of other components, not necessarily carried by the mass itself.
- c. $E = mc^2$, so an argument could be made but it is not relevant here.

4.

False! Only solids resist shear stress by static deflection. Fluids continuously deform under shear stress.

5.

(c), assuming V_n is the velocity normal to A_c , then the final units are kg/s.

1 Problem 6: Hydro Storage System Energy Analysis

Given:

- Volumetric flow rate: $\dot{V} = 500 \text{ m}^3/\text{s}$
- Elevation drop: $\Delta z = 85 \text{ m}$
- Water density: $\rho = 1000 \text{ kg/m}^3$ (assumed)

(a) Maximum possible power with upper reservoir as a lake ($V_1 = 0$):

- Apply the Energy Equation:

$$\frac{V_1^2}{2} + gz_1 = \frac{V_2^2}{2} + gz_2 + w_{\text{turbine}}$$

where:

V_1 = velocity at upper reservoir (lake) = 0

V_2 = velocity at lower lake (assume 0)

z_1 = elevation of upper reservoir

z_2 = elevation of lower lake

$g = 9.81 \text{ m/s}^2$

- Solve for work per unit mass:

$$w_{\text{turbine}} = g(z_1 - z_2)$$

- Calculate total power:

$$\dot{W}_{\text{turbine}} = \dot{m} \cdot w_{\text{turbine}}$$

$$\dot{m} = \rho Q = 1000 \text{ kg/m}^3 \times 500 \text{ m}^3/\text{s} = 500,000 \text{ kg/s}$$

$$w_{\text{turbine}} = 9.81 \text{ m/s}^2 \times 85 \text{ m} = 833.85 \text{ J/kg}$$

$$\dot{W}_{\text{turbine}} = 500,000 \text{ kg/s} \times 833.85 \text{ J/kg} = 416,925,000 \text{ W} = 416.9 \text{ MW}$$

(b) Maximum possible power with upper reservoir as a river ($V_1 = 3.75 \text{ m/s}$):

- Include kinetic energy at inlet:

$$w_{\text{turbine}} = \frac{V_1^2}{2} + g(z_1 - z_2)$$

$$= \frac{(3.75 \text{ m/s})^2}{2} + 9.81 \text{ m/s}^2 \times 85 \text{ m}$$

$$= 7.03 \text{ J/kg} + 833.85 \text{ J/kg} = 840.88 \text{ J/kg}$$

$$\dot{W}_{\text{turbine}} = 500,000 \text{ kg/s} \times 840.88 \text{ J/kg} = 420,440,000 \text{ W} = 420.4 \text{ MW}$$

- **Percent increase in power:**

$$\% \text{ increase} = \frac{420.4 - 416.9}{416.9} \times 100\% = 0.84\%$$

- **Physical explanation:**

The increase in possible power output is due to the additional kinetic energy of the water entering the turbine when the upper reservoir is a river (nonzero velocity). This kinetic energy is converted to useful work by the turbine, resulting in a slightly higher power output compared to the case where the upper reservoir is a still lake.

2 Problem 7: Piston-Cylinder Energy Balance

Given:

- Mass of H₂ gas: $m = 3.8 \text{ kg}$
- Gas constant: $R = 4.124 \text{ kJ/kg} \cdot \text{K}$
- Initial state: $P_1 = 250 \text{ kPa}$, $T_1 = 900 \text{ K}$
- Final temperature: $T_2 = 400 \text{ K}$
- Process: Constant pressure (heat removal)

Part A: Find the change in volume (m³) of the system.

- Use the ideal gas law to find the initial volume:

$$\begin{aligned} PV &= mRT \\ V_1 &= \frac{mRT_1}{P} \\ &= \frac{(3.8 \text{ kg})(4.124 \text{ kJ/kg} \cdot \text{K})(900 \text{ K})}{250 \text{ kPa}} \\ &= 56.416 \text{ m}^3 \end{aligned}$$

- For a constant pressure process, $\frac{T_1}{V_1} = \frac{T_2}{V_2}$:

$$\begin{aligned} \frac{900}{56.416} &= \frac{400}{V_2} \\ V_2 &= \frac{400 \times 56.416}{900} = 25.073 \text{ m}^3 \end{aligned}$$

- Change in volume:

$$\Delta V = V_2 - V_1 = 25.073 - 56.416 = -31.343 \text{ m}^3$$

The negative sign indicates the volume decreased as the gas cooled.

Part B: Find the magnitude (kJ) and direction (in or out of system) of the boundary work during this process. Defend why your answer makes sense.

- For a constant pressure process, boundary work is:

$$\begin{aligned} W &= P\Delta V = P(V_2 - V_1) \\ &= 250 \text{ kPa} \times (25.073 - 56.416) \text{ m}^3 \\ &= 250 \text{ kPa} \times (-31.343 \text{ m}^3) \\ &= -7,836 \text{ kJ} \end{aligned}$$

- The negative sign indicates work is done **by** the system (the gas does work on the surroundings as it contracts). If you define work done **on** the system, $W = +7,836 \text{ kJ}$ into the system. This makes sense because as heat is removed at constant pressure, the gas contracts and the surroundings do work on the gas.

Part C: Calculate the change in specific enthalpy (kJ/kg) and internal energy (kJ/kg) of the gas during this process using the given property tables.

- **Change in specific enthalpy (Δh):**

$$\begin{aligned} \Delta h &= \int_{T_1}^{T_2} C_p(T) dT \\ &\approx \bar{C}_p \times (T_2 - T_1) \end{aligned}$$

Where \bar{C}_p is the average of C_p at 900 K and 400 K (read from the graph):

$$\begin{aligned} \bar{C}_p &\approx \frac{14.35 + 14.93}{2} = 14.64 \text{ kJ/kg} \cdot \text{K} \\ \Delta h &= 14.64 \text{ kJ/kg} \cdot \text{K} \times (400 - 900) \text{ K} \\ &= 14.64 \times (-500) = -7,320 \text{ kJ/kg} \end{aligned}$$

- **Change in specific internal energy (Δu):**

$$\begin{aligned} \Delta u &= \int_{T_1}^{T_2} C_v(T) dT \\ &\approx \bar{C}_v \times (T_2 - T_1) \end{aligned}$$

Where \bar{C}_v is the average of C_v at 900 K and 400 K (read from the graph, e.g. 10.2 and 10.9):

$$\begin{aligned} \bar{C}_v &\approx \frac{10.2 + 10.9}{2} = 10.55 \text{ kJ/kg} \cdot \text{K} \\ \Delta u &= 10.55 \text{ kJ/kg} \cdot \text{K} \times (400 - 900) \text{ K} \\ &= 10.55 \times (-500) = -5,275 \text{ kJ/kg} \end{aligned}$$

- Both Δh and Δu are negative, indicating the gas lost enthalpy and internal energy as it cooled from 900 K to 400 K.

Problem 8: Centrifugal Pump Energy Balance

Given:

- Shaft work: $\dot{W}_s = 10 \text{ kW}$
- Mass flow rate: $\dot{m} = 0.05 \text{ kg/s}$
- Water properties: $\rho = 1000 \text{ kg/m}^3$, $C = 4.2 \text{ kJ/kg} \cdot \text{K}$
- Inlet temperature: $T_1 = 20^\circ\text{C}$
- Outlet temperature: $T_2 = 45^\circ\text{C}$
- Cross-sectional areas: $A_{inlet} = A_{outlet}$

Part A: Prove that the change in kinetic energy is negligible

- **Given:**

$$\dot{m} = 0.05 \text{ kg/s}$$

$$A_2 = A_3$$

$$\rho = 1000 \text{ kg/m}^3$$

- **Conservation of mass for steady flow:**

$$\dot{m} = \rho A_2 v_2 = \rho A_3 v_3$$

$$\Rightarrow v_2 = v_3$$

- **Change in kinetic energy:**

$$\Delta KE = \frac{1}{2} \dot{m} (v_3^2 - v_2^2) = 0$$

- **Conclusion:** The change in kinetic energy is negligible (zero) for the water in this pump.

Part B: Estimate the heat loss (kW) from the pump

- **Given:**

$$\dot{W}_s = 10 \text{ kW}$$

$$C = 4.2 \text{ kJ/kg} \cdot \text{K}$$

$$\dot{m} = 0.05 \text{ kg/s}$$

$$T_1 = 20^\circ\text{C}$$

$$T_2 = 45^\circ\text{C}$$

- **Steady-flow energy equation (SFEE):**

$$\dot{Q}_{\text{loss}} - \dot{W}_s = \dot{m} [\Delta h + \Delta KE + \Delta PE]$$

- **Assumptions:**

- $\Delta KE = 0$ (from Part A)
- $\Delta PE \approx 0$ (not significant)
- For incompressible liquid: $\Delta h = C\Delta T$

- **Energy balance:**

$$\begin{aligned}\dot{Q}_{\text{loss}} - \dot{W}_s &= \dot{m} \cdot C \cdot (T_2 - T_1) \\ \dot{m} \cdot C \cdot \Delta T &= 0.05 \times 4.2 \times (45 - 20) \\ &= 0.05 \times 4.2 \times 25 \\ &= 5.25 \text{ kW}\end{aligned}$$

- **Solve for heat loss:**

$$\begin{aligned}\dot{Q}_{\text{loss}} &= \dot{W}_s - \dot{m} \cdot C \cdot \Delta T \\ &= 10 - 5.25 = 4.75 \text{ kW}\end{aligned}$$

- **Answer:** The heat loss from the pump is approximately 4.75 kW (rounded: 4.8 kW).

Problem 9: Isentropic Steam Turbine with Extraction

Given:

- Steam inlet: $\dot{m}_1 = 3 \text{ kg/s}$ at $P_1 = 3 \text{ MPa}$ (saturated vapor)
- Extraction: 12% of flow at $P_2 = 500 \text{ kPa}$
- Exhaust: 88% of flow at $P_3 = 50 \text{ kPa}$, $T_3 = 100^\circ\text{C}$
- Process: Isentropic (reversible adiabatic)

Solution:

- **Mass Flow Rates:**

$$\dot{m}_2 = 0.12 \times \dot{m}_1 = 0.36 \text{ kg/s}$$

$$\dot{m}_3 = \dot{m}_1 - \dot{m}_2 = 2.64 \text{ kg/s}$$

- **Steady Flow Energy Equation (SFEE) for the turbine:**

$$\dot{W}_{\text{out}} = \dot{m}_1 h_1 - \dot{m}_2 h_2 - \dot{m}_3 h_3$$

where:

- h_1 = specific enthalpy at inlet (3 MPa, isentropic assumption from inlet state)
- h_2 = specific enthalpy at extraction point (500 kPa)
- h_3 = specific enthalpy at exhaust (50 kPa, 100°C)

- **Steam Properties from CoolProp CLI:**

- **State 1 (Inlet):** $P_1 = 3 \text{ MPa}$, saturated vapor

$$T_1 = 507.0 \text{ K} = 234^\circ\text{C}$$

$$h_1 = 2803.2 \text{ kJ/kg}$$

$$s_1 = 6185.6 \text{ J/kg} \cdot \text{K} = 6.186 \text{ kJ/kg} \cdot \text{K}$$

- **State 2 (Extraction):** $P_2 = 500 \text{ kPa}$, $s_2 = s_1$ (isentropic)

$$s_2 = 6185.6 \text{ J/kg} \cdot \text{K}$$

$$h_2 = 2478.2 \text{ kJ/kg}$$

- **State 3 (Exhaust):** $P_3 = 50 \text{ kPa}$, $T_3 = 100^\circ\text{C}$

$$h_3 = 2682.4 \text{ kJ/kg}$$

- **Power Calculation:**

$$\begin{aligned}\dot{W}_{\text{out}} &= \dot{m}_1 h_1 - \dot{m}_2 h_2 - \dot{m}_3 h_3 \\ &= (3.0)(2803.2) - (0.36)(2478.2) - (2.64)(2682.4) \\ &= 8409.6 - 892.2 - 7081.5 \\ &= 435.9 \text{ kW}\end{aligned}$$

- **CLI Commands:** See CLI_Commands.md for detailed property lookup commands

Problem 10: Real Steam Turbine Entropy Analysis

Given:

- Mass flow rate: $\dot{m} = 2.5 \text{ kg/s}$
- Inlet conditions: $P_1 = 3 \text{ MPa}$, $T_1 = 400^\circ\text{C}$ (superheated steam)
- Outlet pressure: $P_2 = 100 \text{ kPa}$
- Isentropic efficiency: $\eta_s = 0.85$
- Process: Adiabatic expansion ($\dot{Q} = 0$), negligible kinetic/potential energy changes

Given:

- $\dot{m} = 2.5 \text{ kg/s}$
- $P_1 = 3 \text{ MPa}$, $T_1 = 400^\circ\text{C}$
- $P_2 = 100 \text{ kPa}$
- $\eta_s = 0.85$
- Steady flow, well-insulated ($\dot{Q} = 0$), negligible ΔKE and ΔPE

PART A: Isentropic Operation

- **Goal:** Find T_{2s} , h_{2s} , x_{2s} (quality), s_{2s}
- **Process:**
 1. Look up inlet state properties (superheated steam at $P_1 = 3 \text{ MPa}$, $T_1 = 400^\circ\text{C}$): h_1 , s_1 (from steam tables)
 2. Apply isentropic condition: $s_{2s} = s_1$
 3. At $P_2 = 100 \text{ kPa}$, $s_{2s} = s_1$: compare s_{2s} with s_f and s_g at 100 kPa
 4. If $s_f < s_{2s} < s_g$, two-phase mixture: $x_{2s} = (s_{2s} - s_f)/s_{fg}$
 5. Find enthalpy: $h_{2s} = h_f + x_{2s}h_{fg}$
 6. Temperature: $T_{2s} = T_{sat}$ at 100 kPa
- **Summary for Part A:**
 - $h_1 = 3231.7 \text{ kJ/kg}$, $s_1 = 6.923 \text{ kJ/kg-K}$ (inlet properties)
 - $T_{2s} = 372.8 \text{ K} = 99.6^\circ\text{C}$
 - $h_{2s} = 2512.6 \text{ kJ/kg}$
 - $x_{2s} = 0.928$ (92.8% vapor)
 - $s_{2s} = 6.923 \text{ kJ/kg-K}$ (isentropic)

PART B: Real Operation

- **Goal:** Find T_2 , h_2 , x_2 , s_2 (actual outlet conditions)

- **Process:**

1. Use isentropic efficiency: $\eta_s = (h_1 - h_2)/(h_1 - h_{2s}) = 0.85$
2. Solve for h_2 : $h_2 = h_1 - \eta_s(h_1 - h_{2s})$
3. At $P_2 = 100$ kPa, h_2 known: compare h_2 with h_f and h_g at 100 kPa
4. If $h_f < h_2 < h_g$, two-phase mixture: $x_2 = (h_2 - h_f)/h_{fg}$
5. Find entropy: $s_2 = s_f + x_2 s_{fg}$
6. Temperature: $T_2 = T_{sat}$ at 100 kPa

- **Summary for Part B:**

- Isentropic work: $w_s = h_1 - h_{2s} = 3231.7 - 2512.6 = 719.1$ kJ/kg
- Actual work: $w_{actual} = \eta_s \times w_s = 0.85 \times 719.1 = 611.2$ kJ/kg
- Actual outlet enthalpy: $h_2 = h_1 - w_{actual} = 3231.7 - 611.2 = 2620.5$ kJ/kg
- $T_2 = 372.8$ K = 99.6°C
- $h_2 = 2620.5$ kJ/kg
- $x_2 = 0.976$ (97.6% vapor)
- $s_2 = 7.213$ kJ/kg-K

PART C: Work Comparison

- **Goal:** Calculate and compare turbine work output

- **Process:**

1. Isentropic work: $w_s = h_1 - h_{2s} = 3231.7 - 2512.6 = 719.1$ kJ/kg
2. Actual work: $w_{actual} = h_1 - h_2 = 3231.7 - 2620.5 = 611.2$ kJ/kg
3. Power output:

$$\dot{W}_s = \dot{m}w_s = 2.5 \times 719.1 = 1797.8 \text{ kW}$$

$$\dot{W}_{actual} = \dot{m}w_{actual} = 2.5 \times 611.2 = 1528.0 \text{ kW}$$

4. Compare difference: $\Delta \dot{W} = \dot{W}_s - \dot{W}_{actual} = 1797.8 - 1528.0 = 269.8$ kW
5. Verification: $w_{actual} = \eta_s \times w_s = 0.85 \times 719.1 = 611.2$ kJ/kg ✓
6. Explanation: Real turbine produces 15% less work due to irreversibilities (friction, turbulence, heat transfer within fluid). Energy lost to entropy generation.

PART D: Entropy Generation Rate

- **Goal:** Calculate \dot{S}_{gen} using entropy balance
- **Given entropy balance:**

$$\sum \frac{\dot{Q}}{T_{boundary}} + \sum \dot{m}s_{in} - \sum \dot{m}s_{out} + \dot{S}_{gen} = \frac{dS_{sys}}{dt}$$

- **For steady flow, well-insulated turbine:**

- $\frac{dS_{sys}}{dt} = 0$ (steady state)
- $\sum \frac{\dot{Q}}{T_{boundary}} = 0$ (adiabatic)
- Single inlet and outlet stream

- **Simplified:**

$$0 + \dot{m}s_1 - \dot{m}s_2 + \dot{S}_{gen} = 0$$

$$\Rightarrow \dot{S}_{gen} = \dot{m}(s_2 - s_1)$$

- **Calculation:**

$$\begin{aligned}\dot{S}_{gen} &= \dot{m}(s_2 - s_1) = 2.5 \times (7.213 - 6.923) \\ &= 2.5 \times 0.290 = 0.725 \text{ kW/K}\end{aligned}$$

- **Check:** $\dot{S}_{gen} = 0.725 > 0$ ✓ (confirms 2nd law for real process)

PART E: Comparison and Benefit

- **Goal:** Compare A vs B properties and identify benefit of non-ideal turbine
- **Process:**

1. Create comparison table:

Property	Isentropic (A)	Real (B)	Difference
h_2 (kJ/kg)	2512.6	2620.5	+107.9
s_2 (kJ/kg-K)	6.923	7.213	+0.290
x_2 (quality)	0.928	0.976	+0.048
T_2 (K)	372.8	372.8	0

2. Key observation: Real turbine has higher quality ($x_2 > x_{2s}$), higher entropy ($s_2 > s_1$), higher enthalpy at exit ($h_2 > h_{2s}$)
3. Benefit: Less liquid content in exhaust (higher quality = more vapor), reduces water droplet erosion on turbine blades. Non-ideal operation trades work output for blade longevity; small sacrifice in efficiency protects equipment.

Solution Steps Summary

1. Look up inlet properties from steam tables (P_1, T_1)
2. Part A: Use $s_{2s} = s_1$, find two-phase properties at P_2
3. Part B: Apply η_s to find h_2 , then find two-phase properties
4. Part C: Calculate work from enthalpy drops, compare
5. Part D: Apply entropy balance with given equation
6. Part E: Compare moisture content (quality), explain blade erosion benefit

Problem 11: Steam Quality Change in Piston-Cylinder

Given:

- System: Free piston-cylinder (constant pressure process)
- Working fluid: $m = 3$ kg of water/steam mixture
- Pressure: $P = 200$ kPa (constant throughout process)
- Initial state: $x_1 = 0.60$ (40% liquid, 60% vapor)
- Final state: $x_2 = 1.0$ (all liquid evaporated to saturated vapor)
- Process: Adiabatic heating with electric resistance heater

Given Information

- Mass: $m = 3$ kg
- Pressure: $P = 200$ kPa (constant)
- Initial quality: $x_1 = 0.60$ (40% liquid \rightarrow 60% vapor)
- Final quality: $x_2 = 1.0$ (all liquid evaporated \rightarrow saturated vapor)

Relevant Equations

- For specific entropy in a two-phase mixture:

$$s = s_f + x \cdot s_{fg}$$

- Specific entropy change:

$$\Delta s = s_2 - s_1$$

Setup with Numbers

- State 1 (initial):

$$\begin{aligned} s_1 &= s_f + x_1 \cdot s_{fg} \\ &= s_f + (0.60) \cdot s_{fg} \end{aligned}$$

- State 2 (final):

$$s_2 = s_g \quad (\text{or equivalently: } s_2 = s_f + (1.0) \cdot s_{fg})$$

- Change in specific entropy:

$$\begin{aligned}\Delta s &= s_2 - s_1 \\ &= s_g - [s_f + (0.60) \cdot s_{fg}]\end{aligned}$$

Solution

1. Look up saturated steam properties at $P = 200$ kPa:

- $s_f = 1.530$ kJ/kg·K (saturated liquid entropy)
- $s_g = 7.127$ kJ/kg·K (saturated vapor entropy)
- Calculate: $s_{fg} = s_g - s_f = 7.127 - 1.530 = 5.597$ kJ/kg·K

2. Calculate initial specific entropy ($x_1 = 0.60$):

$$\begin{aligned}s_1 &= s_f + x_1 \cdot s_{fg} = 1.530 + (0.60)(5.597) \\ &= 1.530 + 3.358 = 4.888 \text{ kJ/kg} \cdot \text{K}\end{aligned}$$

3. Identify final specific entropy ($x_2 = 1.0$, saturated vapor):

$$s_2 = s_g = 7.127 \text{ kJ/kg} \cdot \text{K}$$

4. Calculate the specific entropy change:

$$\Delta s = s_2 - s_1 = 7.127 - 4.888 = \boxed{2.239 \text{ kJ/kg} \cdot \text{K}}$$

5. **Physical interpretation:** Positive Δs reflects increased molecular disorder as liquid evaporates to vapor.

Problem 12: Piston-Cylinder Total Entropy Change

Given:

- System: Well-insulated piston-cylinder with freely moving piston
- Working fluid: Water (saturated liquid initially)
- Initial state: $V_1 = 4.8$ L, $P_1 = 150$ kPa, $x_1 = 0$
- Energy input: $Q_{in} = 1700$ kJ from electric resistance heater
- Process: Constant pressure heating (piston moves freely)

Given Information

- Initial volume: $V_1 = 4.8 \text{ L} = 0.0048 \text{ m}^3$

- Initial state: Saturated liquid water, $x_1 = 0$
- Initial pressure: $P_1 = 150$ kPa
- Constant pressure process ($P_2 = P_1$)
- Energy input: $Q_{in} = 1700$ kJ

Assumptions

- Well-insulated system \rightarrow no heat loss to surroundings
- Piston moves freely \rightarrow constant pressure process
- Neglect kinetic and potential energy changes
- Quasi-equilibrium process

Relevant Equations

- Energy Balance (First Law):

$$Q_{in} - W_{out} = \Delta U = m(u_2 - u_1)$$

For constant pressure:

$$\begin{aligned} W_{out} &= P(V_2 - V_1) = mP(v_2 - v_1) \\ Q_{in} &= m(u_2 - u_1) + mP(v_2 - v_1) = m(h_2 - h_1) \end{aligned}$$

- Entropy Change:

$$\begin{aligned} \Delta S &= m(s_2 - s_1) \\ \Delta s &= s_2 - s_1 \end{aligned}$$

Solution

1. Find initial state properties at 150 kPa saturated liquid:

- $h_1 = h_f = 467.1$ kJ/kg
- $s_1 = s_f = 1.434$ kJ/kg·K
- $\rho_1 = 949.9$ kg/m³
- Specific volume: $v_1 = 1/\rho_1 = 1/949.9 = 0.001053$ m³/kg

2. Calculate mass of water:

$$m = \frac{V_1}{v_1} = \frac{0.0048 \text{ m}^3}{0.001053 \text{ m}^3/\text{kg}} = 4.558 \text{ kg}$$

3. Apply energy balance to find final enthalpy:

$$\begin{aligned}
 Q_{in} &= m(h_2 - h_1) \quad (\text{constant pressure}) \\
 h_2 &= h_1 + \frac{Q_{in}}{m} = 467.1 + \frac{1700}{4.558} \\
 &= 467.1 + 373.0 = 840.1 \text{ kJ/kg}
 \end{aligned}$$

4. Determine final state condition at 150 kPa:

- Reference: $h_g = 2693.1 \text{ kJ/kg}$ (saturated vapor enthalpy)
- Since $h_1 = 467.1 < h_2 = 840.1 < h_g = 2693.1 \rightarrow$ **two-phase mixture**
- Quality: $x_2 = 0.168$ (16.8% vapor)
- Final entropy: $s_2 = 2.404 \text{ kJ/kg}\cdot\text{K}$

5. Calculate entropy changes:

$$\begin{aligned}
 \Delta s &= s_2 - s_1 = 2.404 - 1.434 = 0.970 \text{ kJ/kg}\cdot\text{K} \\
 \Delta S &= m \times \Delta s = 4.558 \times 0.970 = \boxed{4.42 \text{ kJ/K}}
 \end{aligned}$$

Problem 13: Solar Energy Storage Entropy Analysis

Given:

- **Ceramic Particles (CARBOBEAD CP):**
 - Density: $\rho = 3270 \text{ kg/m}^3$
 - Specific heat: $C = 1.3 \text{ kJ/kg} \cdot \text{K}$
 - Temperature range: $T_1 = 250^\circ\text{C}$ to $T_2 = 1100^\circ\text{C}$
- **Compressed Air System:**
 - Mass flow rate: $\dot{m} = 0.75 \text{ kg/s}$
 - Inlet conditions: $T_3 = 250^\circ\text{C}$, $P_3 = 4000 \text{ kPa}$
 - Heating: $T_4 = 1100^\circ\text{C}$ at constant pressure
 - Turbine expansion: Isentropic to $P_5 = 100 \text{ kPa}$
- **Air Properties:** Gas constant $R = 0.287 \text{ kJ/kg} \cdot \text{K}$

Assumptions:

- **Ceramic Particles (CARBOBEAD CP):**
 - Incompressible solid with constant specific heat ($C = 1.3 \text{ kJ/kg}\cdot\text{K}$)

- Uniform heating from 250°C to 1100°C
- Negligible thermal expansion effects

- **Compressed Air System:**

- Ideal gas behavior with constant gas constant ($R = 0.287 \text{ kJ/kg}\cdot\text{K}$)
- Negligible pressure drop in solar receiver ($P_4 = P_3 = 4000 \text{ kPa}$)
- Constant pressure heating process (3→4)
- Reversible adiabatic (isentropic) turbine expansion (4→5)
- Steady flow conditions throughout

- **Reference State:** Table A-17 reference at 0 K and 100 kPa

- **Negligible Effects:** Kinetic and potential energy changes, heat losses

PART A: Change in specific entropy (kJ/kg·K) of ceramic particles from 1 →

2

For an incompressible substance with constant specific heat:

$$\begin{aligned}
 \Delta s_{1 \rightarrow 2} &= C \ln \left(\frac{T_2}{T_1} \right) \\
 &= 1.3 \text{ kJ/kg} \cdot \text{K} \times \ln \left(\frac{1373.15}{523.15} \right) \\
 &= 1.3 \times \ln(2.625) = 1.3 \times 0.965 \\
 &= \boxed{1.254 \text{ kJ/kg} \cdot \text{K}}
 \end{aligned}$$

PART B: Change in specific entropy (kJ/kg·K) of compressed air from 3 → 4

Using CoolProp for air property calculations:

- **State 3:** $T_3 = 523.15 \text{ K}$, $P_3 = 4000 \text{ kPa}$
- **State 4:** $T_4 = 1373.15 \text{ K}$, $P_4 = 4000 \text{ kPa}$ (constant pressure heating)

$$\begin{aligned}
 \Delta s_{3 \rightarrow 4} &= s_4 - s_3 \\
 s_3 &= 3389.9 \text{ J/kg} \cdot \text{K} = 3.390 \text{ kJ/kg} \cdot \text{K} \\
 s_4 &= 4468.7 \text{ J/kg} \cdot \text{K} = 4.469 \text{ kJ/kg} \cdot \text{K} \\
 \Delta s_{3 \rightarrow 4} &= 4.469 - 3.390 = \boxed{1.079 \text{ kJ/kg} \cdot \text{K}}
 \end{aligned}$$

PART C: Specific entropy (kJ/kg·K) of air at state 4

From Part B calculation:

$$s_4 = \boxed{4.469 \text{ kJ/kg} \cdot \text{K}}$$

PART D: Work (kW) produced via turbine from 4 → 5

For isentropic expansion ($s_5 = s_4$), solve for T_5 at $P_5 = 100$ kPa:

- **Iterative Solution:** Using `property_solver.py` to find T_5 where $s_5 = 4468.7$ J/kg·K at $P_5 = 100$ kPa
- **Result:** $T_5 = 529.7$ K

Enthalpy Calculations: Using CoolProp CLI for property lookup:

$$h_4 = 1612.8 \text{ kJ/kg} \quad (\text{at } 1373.15 \text{ K, } 4000 \text{ kPa})$$

$$h_5 = 660.1 \text{ kJ/kg} \quad (\text{at } 529.7 \text{ K, } 100 \text{ kPa})$$

$$w_t = h_4 - h_5 = 1612.8 - 660.1 = 952.7 \text{ kJ/kg}$$

$$\dot{W}_{turb} = \dot{m} \times w_t = 0.75 \times 952.7 = \boxed{714.5 \text{ kW}}$$

Solution Summary

1. **Ceramic Particles:** $\Delta s_{1 \rightarrow 2} = 1.254$ kJ/kg·K (heating from 250°C to 1100°C)
2. **Air Entropy Change:** $\Delta s_{3 \rightarrow 4} = 1.079$ kJ/kg·K (constant pressure heating)
3. **Air Entropy at State 4:** $s_4 = 4.469$ kJ/kg·K
4. **Turbine Work:** $\dot{W}_{turb} = 714.5$ kW (isentropic expansion from 4000 kPa to 100 kPa)

Key Tools Used:

- **CoolProp CLI:** For accurate air property calculations instead of Table A-17 approximations
- **Property Solver:** Iterative tool to find T_5 where $s_5 = s_4$ for isentropic expansion
- **Incompressible Substance Formula:** $\Delta s = C \ln(T_2/T_1)$ for ceramic particles

Problem 14: Concentrated Solar Power Carnot Analysis

Given:

- **Solar Concentrator System:**
 - Concentration ratios: $C = 100, 500, 1000, 2000, 3000$ suns
 - Solar irradiance: $I = 1000 \text{ W/m}^2$
 - Stefan-Boltzmann constant: $\sigma = 5.67 \times 10^{-8} \text{ W/(m}^2 \cdot \text{K}^4)$
- **Carnot Heat Engine:**
 - Cold reservoir: $T_L = 300 \text{ K}$ (ambient temperature)
 - Hot reservoir: T_H varies (300 K to 2600 K)
- **Analysis Scope:** Combined receiver-engine efficiency optimization

Assumptions:

- **Ideal Carnot cycle:** Reversible processes, no irreversibilities in heat engine
- **Steady-state operation:** Constant solar irradiance and ambient conditions
- **Perfect solar concentration:** All concentrated solar energy reaches receiver surface
- **Black-body radiation:** Receiver emits according to Stefan-Boltzmann law
- **Ambient cold reservoir:** $T_L = 300 \text{ K}$ (environmental heat sink)
- **Negligible convection losses:** Only radiative losses from receiver considered
- **Uniform receiver temperature:** Isothermal hot reservoir at T_H

PART A: Plot Total Efficiency

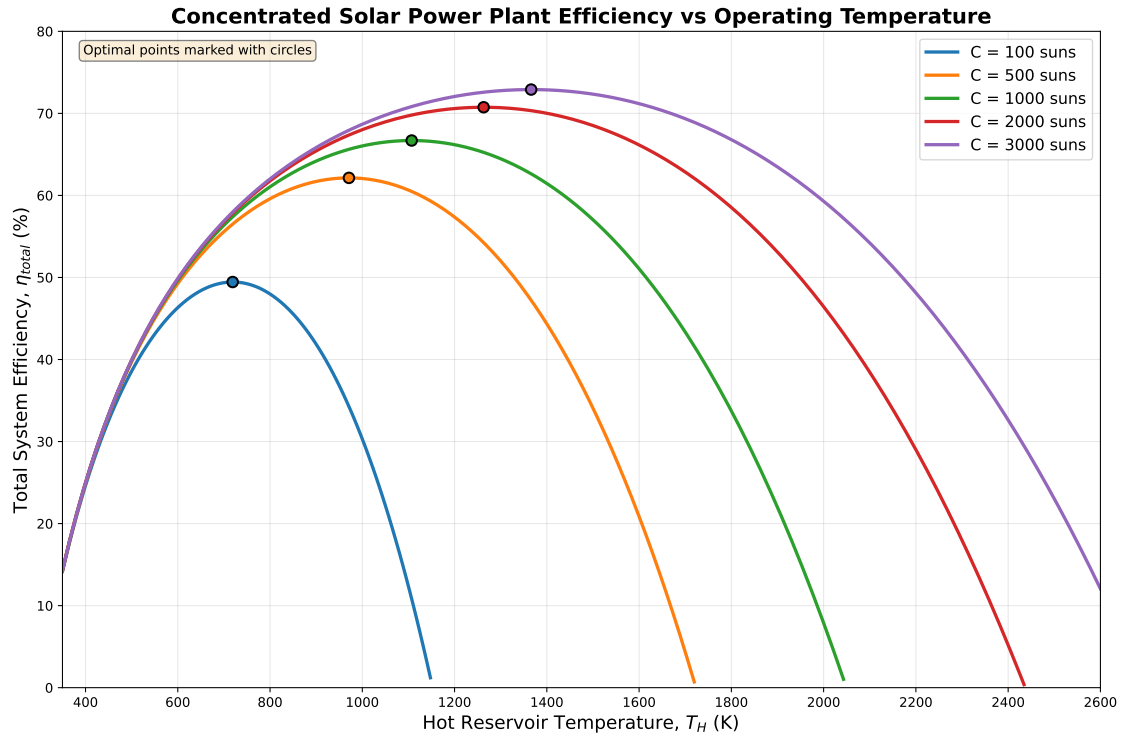


Figure 1: Total system efficiency vs hot reservoir temperature for different concentration ratios. Optimal operating points are marked with circles.

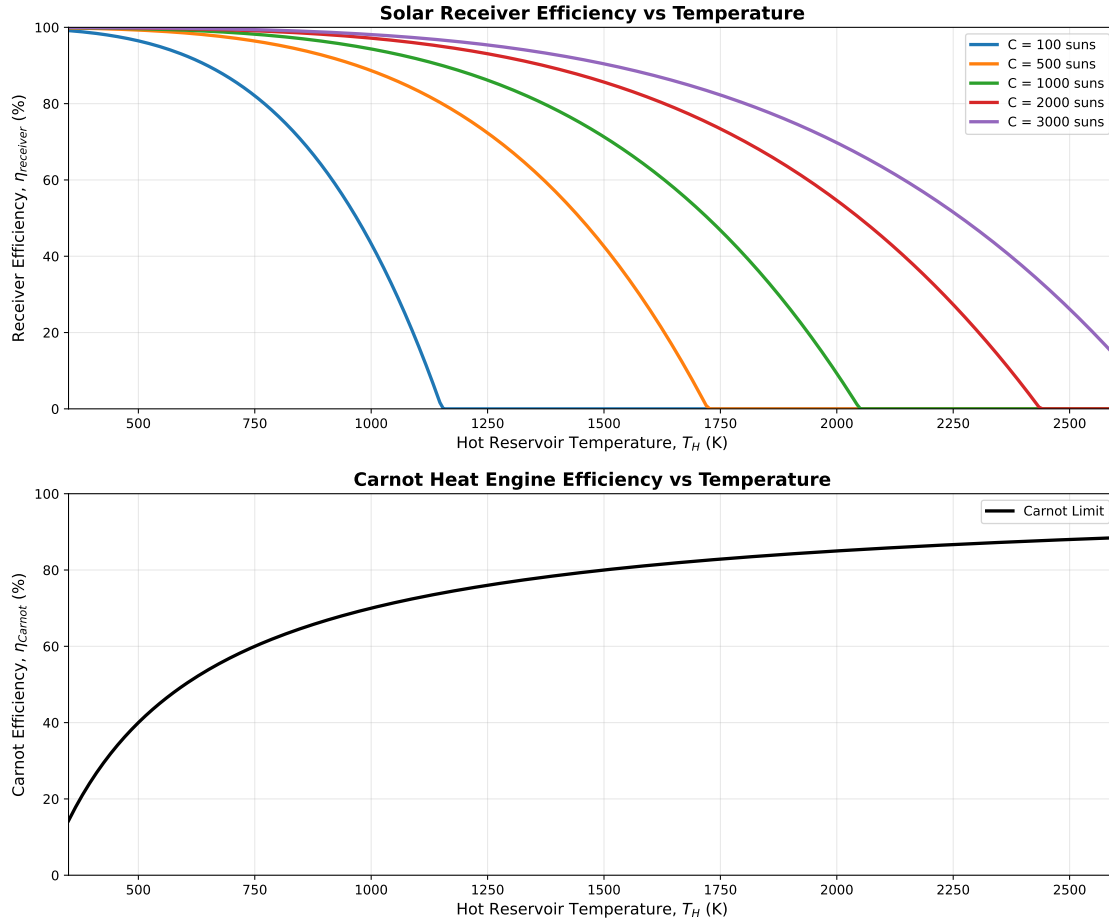


Figure 2: Component efficiencies: (top) Solar receiver efficiency decreases with temperature due to radiation losses, (bottom) Carnot heat engine efficiency increases with temperature.

Relevant Equations

- Receiver Efficiency:

$$\eta_{receiver} = 1 - \frac{\sigma T_H^4}{C \cdot I}$$

- Carnot Heat Engine Efficiency:

$$\eta_{Carnot, HE} = 1 - \frac{T_L}{T_H}$$

- Total System Efficiency:

$$\begin{aligned} \eta_{Total} &= \eta_{receiver} \cdot \eta_{Carnot, HE} \\ &= \left(1 - \frac{5.67 \times 10^{-8} \cdot T_H^4}{C \cdot 1000} \right) \cdot \left(1 - \frac{300}{T_H} \right) \end{aligned}$$

Solution using Python Analysis

- **Python Scripts:** `carnot_simple.py` for optimization, `carnot_plotter.py` for visualization
- **Method:** Temperature optimization scan to find maximum η_{total} for each concentration ratio

Results - Optimal Operating Conditions:

C [suns]	$T_{H,opt}$ [K]	$T_{H,opt}$ [°C]	$\eta_{total,max}$	$\eta_{receiver}$	η_{Carnot}
100	719	446	49.4%	84.8%	58.3%
500	971	698	62.1%	89.9%	69.1%
1000	1107	834	66.7%	91.5%	72.9%
2000	1263	990	70.7%	92.8%	76.2%
3000	1366	1093	72.9%	93.4%	78.0%

Sample Data for Plotting:

T_H [K]	$C = 100$	$C = 500$	$C = 1000$	$C = 2000$	$C = 3000$
500	0.386	0.397	0.399	0.399	0.400
750	0.492	0.578	0.589	0.595	0.596
1000	0.303	0.621	0.660	0.680	0.687
1250	—	0.550	0.655	0.707	0.725
1500	—	0.341	0.570	0.685	0.723
1750	—	—	0.388	0.608	0.682
2000	—	—	0.079	0.464	0.593

PART B: Identify Optimal Receiver Temperature

Key Trends Observed:

1. Optimal Temperature Increases with Concentration:

- $C = 100$ suns: $T_{H,opt} = 719$ K (446°C)
- $C = 3000$ suns: $T_{H,opt} = 1366$ K (1093°C)
- Higher concentration allows operation at higher temperatures

2. Maximum Efficiency Increases with Concentration:

- $C = 100$ suns: $\eta_{max} = 49.4\%$
- $C = 3000$ suns: $\eta_{max} = 72.9\%$
- Improvement of 23.5 percentage points

3. Physical Trade-offs:

- Carnot efficiency increases with T_H : $\eta_{Carnot} = 1 - T_L/T_H$
- Receiver efficiency decreases with T_H : $\eta_{receiver} = 1 - \sigma T_H^4/(C \cdot I)$
- Optimal point balances these competing effects
- Higher concentration shifts optimum to higher temperatures and efficiencies

Python Code Implementation

The analysis was performed using custom Python scripts for accurate calculations:

```
#!/usr/bin/env python3
"""
Problem 14: Carnot Cycle Heat Engines - Analysis
"""

import math
import numpy as np
import matplotlib.pyplot as plt

# Physical constants
SIGMA = 5.67e-8 # Stefan-Boltzmann constant [W/(m^2\cdot K^4)]
I = 1000 # Solar irradiance [W/m^2]
T_L = 300 # Cold reservoir temperature [K]

# Concentration ratios [suns]
concentration_ratios = [100, 500, 1000, 2000, 3000]

def receiver_efficiency(T_H, C):
    """Calculate receiver efficiency:  $\eta_{\text{receiver}} = 1 - \sigma T_H^4 / (C \cdot I)$ """
    return 1 - (SIGMA * T_H**4) / (C * I)

def carnot_efficiency(T_H, T_L):
    """Calculate Carnot heat engine efficiency:  $\eta_{\text{Carnot}} = 1 - T_L / T_H$ """
    return 1 - T_L / T_H

def total_efficiency(T_H, C, T_L):
    """Calculate total system efficiency:  $\eta_{\text{total}} = \eta_{\text{receiver}} \times \eta_{\text{Carnot}}$ """
    eta_recv = receiver_efficiency(T_H, C)
    eta_carnot = carnot_efficiency(T_H, T_L)
    if eta_recv <= 0: # Invalid operating regime
        return 0
    return eta_recv * eta_carnot

def find_optimal_temperature(C, T_L):
    """Find optimal temperature by scanning the valid range"""
    T_range = np.arange(T_L + 50, 2601, 1) # 350K to 2600K

    max_efficiency = -1
    optimal_T_H = T_L + 50

    for T_H in T_range:
        eta = total_efficiency(T_H, C, T_L)
```

```

        if eta > max_efficiency:
            max_efficiency = eta
            optimal_T_H = T_H

    return float(optimal_T_H), max_efficiency

# Find optimal conditions for each concentration ratio
for C in concentration_ratios:
    T_opt, eta_opt = find_optimal_temperature(C, T_L)
    print(f"C = {C:4d} suns: T_opt = {T_opt:6.1f} K, "
          f"eta_max = {eta_opt:5.1%}")

```

Command Line Execution:

```

$ python carnot_simple.py      # Analysis and optimization
$ python carnot_plotter.py    # Generate efficiency plots

```

Key Insights from Analysis:

1. Physical Trade-offs:

- **Carnot efficiency** increases with T_H (better thermodynamic limit)
- **Receiver efficiency** decreases with T_H (radiation losses $\propto T_H^4$)
- Optimum occurs where product is maximized

2. Concentration Effects:

- Higher C shifts optimum to higher temperatures
- Concentrated sunlight ($C \cdot I$) can overcome larger radiation losses (σT_H^4)
- Engineering trade-off: system complexity vs. performance

Physical Interpretation

The total efficiency is the product of two competing effects:

$$\eta_{total}(T_H) = \underbrace{\left(1 - \frac{\sigma T_H^4}{C \cdot I}\right)}_{\text{decreases with } T_H} \times \underbrace{\left(1 - \frac{T_L}{T_H}\right)}_{\text{increases with } T_H}$$

At the optimum: $\frac{d\eta_{total}}{dT_H} = 0$, which yields the balance between thermodynamic benefits and radiation losses.

Python Command Used:

- `python carnot_simple.py` → Complete analysis with optimization

A Appendix