TITLE: Gravimetric analysis of a chloride salt

CHEM 1101

Name: Dorian Wang

Partner: Jason Wang

Group: Friday PM

Date performed: 12th February, 2016

Date submitted: 4th March, 2016

Purpose: The purpose of this experiment was to determine the amount of chloride in an unknown salt sample.

Theory: AgCl is very insoluble, and will precipitate out of solution when silver ions (Ag+) and chlorine ions (Cl-) interact. The original salt is very soluble, and in addition, chloride is removed by the silver precipitation, allowing for more to dissolve. AgCl does have a low Ksp (1.6\*10^-10), so some will be lost to the solution.

Procedure:

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