

Statistical Mechanics
from the context of the course
PHY 410: Statistical Mechanics

Kaedon Cleland-Host

March 9, 2022

Contents

0.1	The SI System	1
0.2	Introduction	2
0.3	Reference	2
1	Microcanonical Ensemble	4
1.1	Multiplicity Functions	4
1.1.2	Boltzmann's Assumption	4
1.1.3	Entropy	4
1.2	Binary System	4
1.2.3	Stirling's Approximation	5
1.3	Thermal Equilibrium	5
1.3.3	Temperature	6
2	The Canonical Ensemble	7
2.1	Partition Function	7
2.2	Reversible Processes	8
2.2.3	Pressure	8
2.3	Helmholtz Free Energy	8
2.4	Ideal Gas	9
2.5	Ideal Diatomic Gas	9
2.6	Thermal Radiation	10
2.7	Phonons in a Solid	11
2.7.3	Einstein Model	11
2.7.5	Debye Model	12
3	Grand Canonical Ensemble	13

0.1 The SI System

In physics it's often important to have precisely defined units for the purposes of making very accurate measurements or simply having a coherent unit system. It's possible to derive all necessary units from five measurements of **length, mass, time, current, and temperature**. The standard SI units for these properties are listed below:

Type	Unit	Definition
Length	Meter(m)	Length of distance light in a vacuum travels in $\frac{1}{299792458}$ seconds
Mass	Kilogram(kg)	Defined by fixing the Planck's constant $h = 6.62607015 \times 10^{-34} kg \cdot m^2 s^{-1}$
Time	Second(s)	Defined by fixing the ground-state hyperfine transition frequency of the caesium-133 atom, to be $9192631770 s^{-1}$
Current	Ampere(A)	Defined by fixing the charge of an electron as $1.602176634 \times 10^{-19} A \cdot s$
Temperature	Kelvin(K)	Defined by fixing the value of the Boltzmann constant k to $1.380649 \times 10^{-23} kg \cdot m^2 s^{-2} K^{-1}$

Common prefixes are listed below:

Prefix	Symbol	Definition
mega	M	10^6
kilo	k	10^3
milli	m	10^{-3}
micro	μ	10^{-6}
nano	n	10^{-9}
pico	p	10^{-12}
femto	f	10^{-15}

Additionally, the following are defined constants:

Symbol	Definition
k_B	$k_B = 1.380649 \times 10^{-23} \text{m}^2 \text{s}^{-2} \text{K}^{-1}$
h	$h = 6.62607015 \times 10^{-34} \text{kg m}^2 \text{s}^{-1}$
\hbar	$\hbar = \frac{h}{2\pi} \approx 1.0546 \times 10^{-34} \text{kg m}^2 \text{s}^{-1}$

0.2 Introduction

Statistical mechanics is the probabilistic study of extremely complex systems. It seeks to describe the equilibrium or averages behaviors of such systems. Very often in nature we deal with systems containing so many possible states that it would be impossible to completely describe and calculate the expected behavior. By focusing on large measurable properties of such systems and the average behavior of the microstates we can derive there behavior. With a combination of statistical arguments and microscopic knowledge, statistical mechanics predicts the macro equilibrium properties, which can be used to derive the laws of thermodynamics.

Thermodynamics is fundamentally connected with statistical mechanics and we will derive the laws of thermodynamics in the book. For reference here are the laws of thermodynamics.

Law 0.2.1. The 0th Law of Thermodynamics - Thermal equilibrium is transitive. If system A and B are in equilibrium and systems B and C are also in equilibrium, then systems A and C are in equilibrium.

Law 0.2.2. The 1st Law of Thermodynamics - Heat is a form of energy. Let ΔE be the change in energy, W be the work done on the system, and Q the heat energy added to the system.

$$\Delta E = W + Q$$

Law 0.2.3. The 2nd Law of Thermodynamics - Total entropy always increases. Let ΔS be the change in entropy and S_f and S_i be the final and initial entropy.

$$\Delta S = S_f - S_i > 0$$

$$\Delta S = \int_i^f \frac{1}{T} \partial Q = S_f - S_i$$

Law 0.2.4. The 3rd Law of Thermodynamics - Absolute Zero. Let S be entropy and T be temperature.

$$\lim_{T \rightarrow 0} S(T) = 0$$

We will seek to derive each of these laws in this book. Additionally, the logarithm function is used very often in statistical mechanics, we will used \log to refer to the natural logarithm (aka \log_e).

0.3 Reference

A reference sheet is include on the next page. [Click here for a pdf file](#)



PHY 410 - Reference Sheet

Stirling's approximation - for very large N the factorial can be very accurately approximated with the following

$$\log N! \approx N \log N - N$$

$$N! \approx \sqrt{2\pi N} N^N e^{-N}$$

Fractional uncertainty of \mathbb{X} is uncertainty of expected value per particle.

$$\frac{\Delta \mathbb{X}}{N} = \frac{\sqrt{\langle \mathbb{X}^2 \rangle - \langle \mathbb{X} \rangle^2}}{N}$$

Boltzmann's constant

$$k_B = 1.380649 \times 10^{-23} \text{ m}^2 \text{ s}^{-2} \text{ K}^{-1}$$

$$\text{Entropy } S = k_B \sigma, \quad \sigma_{TOT} = \sigma_1 + \sigma_2$$

$$\text{Temperature } T = \tau / k_B$$

Microcanonical Ensemble

Multiplicity function

$$g = \# \text{ of microstates}, \quad \mathcal{P}(n) = \frac{1}{g}$$

Expected value of \mathbb{X} is the average across all microstates.

$$\langle \mathbb{X} \rangle = \sum_n \mathbb{X}(n) \mathcal{P}(n) = \frac{1}{g} \sum_n \mathbb{X}(n)$$

Entropy can be written in terms of the multiplicity function.

$$\sigma(N, T, U, V, P) \equiv \log[g(N, T, U, V, P)]$$

Binary System

A **binary system** is a system of N particles where each particles has two possible states. Let N_\uparrow is the number of particle in the up state and N_\downarrow be the number of particles in the down state.

$$g(N, N_\uparrow) = \frac{N!}{N_\uparrow!(N - N_\uparrow)!}, \quad \sum_{N_\uparrow=0}^N g(N, N_\uparrow) = 2^N$$

The binary system can be rewritten in terms of the difference between up states and down states this is the **spin excess**.

$$2S = N_\uparrow - N_\downarrow$$

$$g(N, S) = \frac{N!}{(\frac{N}{2} + S)!(\frac{N}{2} - S)!}$$

$$\sum_{S=-\frac{N}{2}}^{S=\frac{N}{2}} g(N, N_\uparrow) = 2^N$$

Applying Stirling's approximation to the binary model, for large N the multiplicity function and fractional uncertainty are

$$g(N, S) \approx g(N, 0) e^{-2s^2/N}$$

$$g(N, S) \approx \sqrt{\frac{2}{\pi N}} 2^N e^{-2s^2/N}$$

$$\frac{\Delta S}{N} \approx \frac{1}{\sqrt{N}}$$

An example of a binary system is N spin $1/2$ particles in an external **magnetic field** B . The total energy U and magnetization M of the system are

$$U = \sum_{i=1}^N -m_i \cdot \vec{B} = -(N_\uparrow - N_\downarrow) m B = -2SmB$$

$$M = 2Sm = -U/B$$

$$g(N, U) = \frac{N!}{(\frac{N}{2} - \frac{U}{2mB})!(\frac{N}{2} + \frac{U}{2mB})!}$$

$$\sigma(N, S) \approx - \left(\frac{N}{2} + S \right) \log \left(\frac{1}{2} + \frac{S}{N} \right) - \left(\frac{N}{2} - S \right) \log \left(\frac{1}{2} - \frac{S}{N} \right)$$

$$M = Nm \tanh(mB/\tau)$$

Einstein Solid

An **einstein solid** is a system of N atoms where each atom is modeled as a harmonic oscillator the energy of the system is determined by the number of atoms n oscillating at frequency ω .

$$U = n\hbar\omega$$

$$g(N, n) = \frac{(n + N - 1)!}{n!(N - 1)!}$$

$$g(N, n) \approx \left(\frac{n+N}{n} \right)^n \left(\frac{n+N}{N} \right)^N \frac{1}{\sqrt{2\pi n(n+N)/N}}$$

Thermal Equilibrium

Temperature

$$\frac{1}{\tau} = \left(\frac{\partial \sigma}{\partial U} \right)_{N, V}$$

Thermal Equilibrium

$$\left(\frac{\partial \sigma_1}{\partial U_1} \right)_{N_1, V_1} = \left(\frac{\partial \sigma_2}{\partial U_2} \right)_{N_2, V_2}$$

$$\frac{1}{\tau_1} = \frac{1}{\tau_2}$$

2nd law of thermo - Change in entropy ≥ 0 .

Sharpness of Equilibrium For a two binary systems, the number of states in a configuration of deviation δ from equilibrium is

$$g_1 g_2 = (g_1 g_2)_{max} e^{-\left(\frac{2\delta^2}{N_1} - \frac{2\delta^2}{N_2} \right)}$$

Canonical Ensemble

Partition Function - partition by energy levels for a fixed temperature

$$z = \sum_n e^{-\varepsilon_n/\tau}, \quad \mathcal{P}(n) = \frac{1}{z} e^{-\varepsilon_n/\tau}$$

$$z = \sum_\alpha g(\varepsilon_\alpha) e^{-\varepsilon_\alpha/\tau}, \quad \text{for degeneracy } g(\varepsilon_\alpha)$$

Expected Value of \mathbb{X} is the average across all energies (Thermal Average).

$$\langle \mathbb{X} \rangle = \sum_n \mathbb{X}(n) \mathcal{P}(n) = \frac{1}{z} \sum_n \mathbb{X}(n) e^{-\varepsilon_n/\tau}$$

Expected Energy in the canonical ensemble is

$$U = \langle \varepsilon \rangle = \frac{1}{z} \sum_n \varepsilon_n e^{-\varepsilon_n/\tau}$$

$$U = \langle \varepsilon \rangle = \tau^2 \frac{1}{z} \frac{\partial z}{\partial \tau} = \tau^2 \frac{\partial}{\partial \tau} \log z$$

The total partition function and expected energy for N non-interacting particles is simply

$$z_N = z_1^N$$

$$U_N = \langle \varepsilon \rangle_N = N U_1 = N \langle \varepsilon \rangle_1$$

(this also applies for expected value of any \mathbb{X})

Thermodynamic Relations

1st Law of Thermo

$$dU = dQ + dW = \tau d\sigma - PdV$$

$$d\sigma = \frac{1}{\tau} dU + \frac{P}{\tau} dV$$

Temperature $\tau = \left(\frac{\partial U}{\partial \sigma} \right)_V$

Quasi-static Compression Equilibrium

the equilibrium condition for quasi-static compression is

$$\left(\frac{\partial U_1}{\partial V_1} \right)_{\sigma_1} = \left(\frac{\partial U_2}{\partial V_2} \right)_{\sigma_2}$$

Helmholtz Free Energy

$$F = U - \tau \sigma = U - ST = -\tau \log z$$

$$dF = -\sigma d\tau - PdV$$

Entropy $\sigma = - \left(\frac{\partial F}{\partial \tau} \right)_V$, $S = k_B \sigma$

Pressure

$$P = - \left(\frac{\partial U}{\partial V} \right)_\sigma = \tau \left(\frac{\partial \sigma}{\partial V} \right)_U = - \left(\frac{\partial F}{\partial V} \right)_\tau$$

Energy

$$U = -\tau^2 \frac{\partial}{\partial \tau} \left(\frac{F}{\tau} \right)$$

Ideal Gas

DeBroglie Thermal Wavelength is the wavelength of the wave functions of matter at a given temperature.

$$\lambda_T = \sqrt{\frac{2\pi\hbar^2}{m\tau}}$$

Concentration of a system is the inverse of the volume

$$n = \frac{1}{V}$$

Quantum Concentration is the density of quantum state per particle. It is used to define when a system will behave classically ($n \ll n_Q$) and when a system will be dominated by quantum effects ($n \gg n_Q$).

$$n_Q = \frac{1}{\lambda_T^3}$$

Single Particle Ideal Gas is a system in the canonical ensemble consisting of a single particle in a box of side lengths L . The energy levels, partition function and average energy are

$$E_{n_x, n_y, n_z} = \frac{\hbar^2}{2m} \left(\frac{\pi}{L} \right)^2 (n_x^2 + n_y^2 + n_z^2)$$

$$z_1 = \frac{V}{\lambda_T^3}$$

$$U_1 = \frac{3}{2} \tau$$

$$\sigma_1 = \log \left(\frac{V}{\lambda_T^3} \right) + \frac{3}{2}$$

Gibbs Resolution states that for systems in the classical regime the partition function for an ideal gas with N particles is

$$z_N = \frac{1}{N!} (z_1)^N$$

$$U_N = \frac{3}{2} N \tau$$

$$\sigma_N = N \left[\log \left(\frac{V}{N \lambda_T^3} + \frac{5}{2} \right) \right]$$

N-Particle Ideal Gas - by applying Gibbs resolution and properties of expected values we can find the classical ideal gas results

$$PV = N \tau$$

$$U = \frac{3}{2} N \tau$$

$$\sigma = N \left[\log \left(\frac{V}{N \lambda_T^3} \right) + \frac{5}{2} \right]$$

DOG (bork)



Kaedon.net

Chapter 1

Microcanonical Ensemble

Throughout this book we will refer to microscopic and macroscopic systems. Using Quantum mechanics one can predict the behavior of a small number of particles or similar small systems with high levels of symmetry. However, as the number of particles increases, Schrödinger's becomes practically impossible to solve. However the microscopic quantum states of a system determine many measurable macroscopic properties, such as temperature, energy, volume or pressure. Towards making predictions about these macroscopic properties it is useful to count the number of possible microstates for a given macrostate.

Definition 1.0.1. The **microcanonical ensemble** is the system of calculations where the total energy, particle number, and volume are fixed.

1.1 Multiplicity Functions

Definition 1.1.1. A **Multiplicity function** is the number of microstates that produce a given macrostate denoted

$$g(N, T, U, V, P) = \# \text{ of microstates that fit the given macrostate}$$

Assumption 1.1.2. Boltzmann's Assumption - All accessible microstates are equally probable. Let $P(n)$ be the probability that a system is in microstate n .

$$\mathcal{P}(n) = \frac{1}{g(N, T, U, V, P)}$$

Definition 1.1.3. The **entropy** of a system is defined as the natural log of the multiplicity function.

$$\sigma(N, T, U, V, P) \equiv \log[g(N, T, U, V, P)]$$

Definition 1.1.4. The **expected value** of a property \mathbb{X} of a system is determined by the average of that value across all possible microstates.

$$\langle \mathbb{X} \rangle = \sum_n \mathbb{X}(n) \mathcal{P}(n) = \frac{1}{g} \sum_n \mathbb{X}(n)$$

Definition 1.1.5. The **fractional uncertainty** of a property \mathbb{X} is defined in terms of the expected value.

$$\frac{\Delta \mathbb{X}}{N} = \frac{\sqrt{\langle \mathbb{X}^2 \rangle - \langle \mathbb{X} \rangle^2}}{N}$$

1.2 Binary System

Definition 1.2.1. A **binary system** is a system of N particles where each particle has two possible states. Let N_\uparrow be the number of particles in the up state and N_\downarrow be the number of particles in the down state.

$$g(N, N_\uparrow) = \frac{N!}{N_\uparrow!(N - N_\uparrow)!}, \quad \sum_{N_\uparrow=0}^N g(N, N_\uparrow) = 2^N$$

Corollary 1.2.2. The binary system can be rewritten in terms of the **difference between up states and down states**.

$$2S = N_{\uparrow} - N_{\downarrow}$$

$$g(N, S) = \frac{N!}{(\frac{N}{2} + S)!(\frac{N}{2} - S)!}, \quad \sum_{S=-\frac{N}{2}}^{\frac{N}{2}} g(N, N_{\uparrow}) = 2^N$$

Theorem 1.2.3. Stirling's Approximation - for very large N the factorial can be very accurately approximated with the following

$$\log N! \approx N \log N - N$$

$$N! \approx \sqrt{2\pi N} N^N e^{-N}$$

Proposition 1.2.4. Applying Stirling's approximation to the binary model, for large N the multiplicity function and fractional uncertainty are

$$g(N, S) \approx g(N, 0) e^{-2s^2/N}$$

$$g(N, S) \approx \sqrt{\frac{2}{\pi N}} 2^N e^{-2s^2/N}$$

$$\frac{\Delta S}{N} \approx \frac{1}{\sqrt{N}}$$

Example. Consider a system of N spin 1/2 particles in an external magnetic field B . The total energy U and magnetization M of the system are

$$U = \sum_{i=1}^N -\vec{m}_i \cdot \vec{B} = -(N_{\uparrow} - N_{\downarrow})mB = -2SmB$$

$$M = 2Sm = -U/B$$

$$g(N, U) = \frac{N!}{(\frac{N}{2} - \frac{U}{2mB})!(\frac{N}{2} + \frac{U}{2mB})!}$$

$$\sigma(N, S) \approx -\left(\frac{N}{2} + S\right) \log\left(\frac{1}{2} + \frac{S}{N}\right) - \left(\frac{N}{2} - S\right) \log\left(\frac{1}{2} - \frac{S}{N}\right)$$

$$M = Nm \tanh(mB/\tau)$$

Definition 1.2.5. An **einstein solid** is a system of N atoms where each atom is modeled as a harmonic oscillator the total energy of the system is determined by the number of atoms n oscillating at frequency ω .

$$U = n\hbar\omega$$

$$g(N, n) = \frac{(n + N - 1)!}{n!(N - 1)!}$$

$$g(N, n) \approx \frac{\left(\frac{n+N}{n}\right)^n \left(\frac{n+N}{N}\right)^N}{\sqrt{2\pi n(n+N)/N}}$$

1.3 Thermal Equilibrium

Since the uncertainty of our system is extremely small for large N , the most probable state as determined by Boltzmann's assumption will determine the macrostate of the system.

Definition 1.3.1. The **equilibrium** state two systems is the most probable macrostate of the shared system.

$$g_{TOT}(N, V, U) = \sum_{U'_1} g_1(N_1, V_1, U'_1) g_2(N_2, V_2, U - U'_1)$$

$$\mathcal{P}(U'_1) = \frac{g_1(N_1, V_1, U'_1) g_2(N_2, V_2, U - U'_1)}{g_{TOT}(N, V, U)}$$

Example. Consider two isolated binary systems

$$g_1(4, 2) \text{ and } g_2(6, 0)$$

The initial condition of the two states are $N_1 = 4, S_1 = 2$ and $N_2 = 6, S_2 = 0$. Now, allowing contact between the two systems the equilibrium state is the most probable state that conserves energy. To find this we need to maximize the total multiplicity function

$$g(N_1 + N_2, S'_1, S'_2) = g_1(N_1, S'_1)g_2(N_2, S'_2)$$

For this simple system the max is found when $S'_1 = 1$ and $S'_2 = 1$.

Definition 1.3.2. Thermal equilibrium is reached when the energy derivative of the total multiplicity function is zero.

$$\left(\frac{\partial \sigma_1}{\partial U_1}\right)_{N_1, V_1} = \left(\frac{\partial \sigma_2}{\partial U_2}\right)_{N_2, V_2}$$

Definition 1.3.3. The **temperature** of a system is defined with the derivative of entropy σ in terms of energy U . Kelvin temperature T is directly proportional to the fundamental temperature via the Boltzmann constant k_B .

$$\frac{1}{\tau} = \left(\frac{\partial \sigma}{\partial U}\right)_{N, V}$$

$$\tau = k_B T$$

Chapter 2

The Canonical Ensemble

Up until this point we've been working with closed systems where the total amount of energy is fixed. In the real world it is often very difficult to thermally isolate a system. It is much easier to operate at a fixed temperature and a fixed number of particles. We separate each of these cases into ensembles.

Definition 2.0.1. The **canonical ensemble** is the system of calculations where the temperature and particle number are fixed.

2.1 Partition Function

Definition 2.1.1. The **partition function** determines how we calculate probability. For a fixed temperature we partition by energy levels. The probability of a particular energy ε_n is given by

$$\mathcal{P}(\varepsilon_n) = \frac{1}{z} e^{-\varepsilon_n/\tau}$$
$$z = \sum_n e^{-\varepsilon_n/\tau}$$

Corollary 2.1.2. For systems with degeneracy $g(\varepsilon_\alpha)$ for each distinct energy level ε_α , the partition function can be rewritten.

$$z = \sum_\alpha g(\varepsilon_\alpha) e^{-\varepsilon_\alpha/\tau}$$

Definition 2.1.3. The **thermal average** of \mathbb{X} is the average across all energies).

$$\langle \mathbb{X} \rangle = \sum_n \mathbb{X}(n) \mathcal{P}(n) = \frac{1}{z} \sum_n \mathbb{X}(n) e^{-\varepsilon_n/\tau}$$

Definition 2.1.4. The **expected energy** in the canonical ensemble is

$$U = \langle \varepsilon \rangle = \frac{1}{z} \sum_n \varepsilon_n e^{-\varepsilon_n/\tau}$$
$$U = \langle \varepsilon \rangle = \tau^2 \frac{1}{z} \frac{\partial z}{\partial \tau} = \tau^2 \frac{\partial}{\partial \tau} \log z$$

Example. One and many simple harmonic oscillators at temperature τ . Starting with a system of just one harmonic oscillator we find

$$\varepsilon_n = \left(n + \frac{1}{2}\right) \hbar \omega$$
$$z_1 = \sum_n e^{-(n+\frac{1}{2})\frac{\hbar \omega}{\tau}} = \frac{e^{-\hbar \omega/2\tau}}{1 - e^{-\hbar \omega/\tau}}$$
$$U_1 = \langle \varepsilon \rangle = \frac{\hbar \omega}{2} + \frac{\hbar \omega e^{-\hbar \omega/\tau}}{1 - e^{-\hbar \omega/\tau}} = \frac{\hbar \omega}{2} + \frac{\hbar \omega}{e^{-\hbar \omega/\tau} - 1}$$

Now, consider a system of N simple harmonic oscillators at temperature τ .

$$z_N = \sum_{n_1, n_2, \dots, n_N} e^{-\frac{1}{\tau}(\varepsilon_{n_1} + \varepsilon_{n_2} + \dots + \varepsilon_{n_N})} = z_1^N$$
$$U_N = \langle \varepsilon \rangle = \tau^2 \frac{\partial}{\partial \tau} \log z_N = N \tau^2 \frac{\partial}{\partial \tau} \log z = \frac{N \hbar \omega}{2} + \frac{N \hbar \omega}{e^{-\hbar \omega/\tau} - 1} = N U_1$$

2.2 Reversible Processes

Definition 2.2.1. A **quasi-static process** is an idealization where parameters are changed so slowly that the system is indefinitely close to thermal and mechanical equilibrium during the process.

Definition 2.2.2. A **reversible process** is a quasi-static process where the system retraces its steps if you reverse the changes in the parameters.

Definition 2.2.3. The **pressure** of a system is defined as the negative derivative of energy with respect to volume.

$$P_n = - \left(\frac{\partial \varepsilon_n}{\partial V} \right)_\sigma$$

$$P = \langle -\frac{\partial \varepsilon_n}{\partial V} \rangle = \sum_n -\frac{\partial \varepsilon_n}{\partial V} e^{-\varepsilon_n/\tau}$$

Proposition 2.2.4. For quasi-static compression the pressure is simply the derivative of energy with respect to volume.

$$P = -\frac{\partial}{\partial V} \frac{1}{z} \sum_n \varepsilon_n e^{-\varepsilon_n/\tau} = - \left(\frac{\partial U}{\partial V} \right)_\sigma$$

2.3 Helmholtz Free Energy

Definition 2.3.1. The **Helmholtz Free Energy** F of a system is defined in terms of the energy U the entropy σ , S and the temperature τ , T .

$$F = U - \sigma\tau = U - ST$$

Proposition 2.3.2. Equilibrium in the canonical ensemble is achieved when the Helmholtz free energy F is minimized

$$\partial F = 0, \quad \Delta F \leq 0$$

Proposition 2.3.3. The entropy, pressure, and expected energy of a system in the canonical ensemble can be written in terms of the Helmholtz free energy.

$$\sigma = - \left(\frac{\partial F}{\partial \tau} \right)_V, \quad P = - \left(\frac{\partial F}{\partial V} \right)_\tau, \quad U = -\tau^2 \frac{\partial}{\partial \tau} \left(\frac{F}{\tau} \right)$$

Proposition 2.3.4. The Helmholtz free energy is determined by the temperature τ and partition function z .

$$F = -\tau \log z$$

$$z = e^{-F/\tau}$$

Definition 2.3.5. **Maxwell relations** are relations that are derive from the following form

$$df = A dx + B dy$$

$$\frac{\partial}{\partial y} \frac{\partial f}{\partial x} = \frac{\partial}{\partial x} \frac{\partial f}{\partial y}$$

Proposition 2.3.6. Using Maxwell relations we can derive the following

$$dU = \tau d\sigma - P dV$$

$$\left(\frac{\partial \sigma}{\partial V} \right)_\tau = \left(\frac{\partial P}{\partial \tau} \right)_V$$

$$\left(\frac{\partial P}{\partial \sigma} \right)_V = - \left(\frac{\partial \tau}{\partial V} \right)_\sigma$$

2.4 Ideal Gas

Definition 2.4.1. The **concentration** of a system is the inverse of the volume

$$n = \frac{1}{V}$$

Definition 2.4.2. The **quantum concentration** is the density of quantum state per particle. It is used to define when a system will behave classically (when $n \ll n_Q$) and when a system will be dominated by quantum effects (when $n \gg n_Q$).

$$n_Q = \frac{1}{\lambda_T^3}$$

Definition 2.4.3. The **DeBroglie thermal wavelength** is the wavelength of the wave functions of matter at a given temperature.

$$\lambda_T = \sqrt{\frac{2\pi\hbar^2}{m\tau}}$$

Definition 2.4.4. **Single particle ideal gas** is a system in the canonical ensemble consisting of a single particle in a box of side lengths L . The energy levels, partition function and average energy are

$$\begin{aligned} E_{n_x, n_y, n_z} &= \frac{\hbar^2}{2m} \left(\frac{\pi}{L}\right)^2 (n_x^2 + n_y^2 + n_z^2) \\ z_1 &= \sum_{n_x, n_y, n_z} e^{\frac{-E_{n_x, n_y, n_z}}{\tau}} = \left(L\sqrt{\frac{m\tau}{\hbar^2 2\pi}}\right)^3 = \frac{V}{\lambda_T^3} = \frac{n_Q}{n} = \frac{V}{\lambda_T^3} \\ U_1 &= \frac{3}{2}\tau \\ \sigma_1 &= \log\left(\frac{V}{\lambda_T^3}\right) + \frac{3}{2} \end{aligned}$$

Proposition 2.4.5. **Gibbs resolution** states that for systems in the classical regime the partition function for an ideal gas with N particles is

$$\begin{aligned} z_N &= \frac{1}{N!} (z_1)^N \\ U_N &= \frac{3}{2} N\tau \\ \sigma_N &= N \left[\log\left(\frac{V}{N\lambda_T^3} + \frac{5}{2}\right) \right] \end{aligned}$$

Law 2.4.6. **Ideal gas law** states that the pressure and volume are related to temperature and particle number

$$PV = N\tau$$

$$PV = Nk_B T$$

2.5 Ideal Diatomic Gas

Definition 2.5.1. **Single particle diatomic gas** is a system in the canonical ensemble consisting of a single particle in a box with vibration and rotational excitation. The energy levels, partition function and average energy are

$$\begin{aligned} E_{\vec{p}, n, \ell} &= \frac{\rho^2}{2m} + \left(n + \frac{1}{2}\right)\hbar\omega_{vib} + \frac{\hbar^2}{2I}\ell(\ell+1) \\ z_1 &= \sum_{\vec{p}, n, \ell} (2\ell+1) e^{-[E_{KE} + E_{Vib} + E_{Rot}]/\tau} = \frac{V}{\lambda_T^3} + \sum_n e^{-(n+\frac{1}{2})\hbar\omega/\tau} + \sum_\ell (2\ell+1) e^{\frac{\hbar^2 \ell(\ell+1)}{2I}/\tau} \\ U_1 &= \frac{3}{2}N\tau + \delta U_{Vib} + \delta U_{Rot} \\ U_1 &= -\tau \log\left(\frac{V}{\lambda_T^3}\right) + \delta F_{Vib} + \delta F_{Rot} \end{aligned}$$

2.6 Thermal Radiation

Definition 2.6.1. Photons - For light at a given frequency $\omega = 2\pi f$, the energy levels are

$$E_s = s\hbar\omega$$

These energy levels are very similar to the simple harmonic oscillator $E_n = (n + \frac{1}{2})\hbar\omega$. So we can use the example from earlier to derive the partition function.

$$\begin{aligned} z &= \sum_{s=0}^{\infty} e^{-s\hbar\omega/\tau} = \frac{1}{1 - e^{-\hbar\omega/\tau}} \\ \mathcal{P}(s) &= \frac{e^{-s\hbar\omega/\tau}}{z} \\ \langle s \rangle &= \frac{1}{z} \sum_{s=0}^{\infty} s e^{-s\hbar\omega/\tau} = \frac{1}{e^{\hbar\omega/\tau} - 1} \\ \langle \epsilon \rangle &= \frac{\hbar\omega}{e^{\hbar\omega/\tau} - 1} \end{aligned}$$

Definition 2.6.2. A **thermal cavity** or **photon gas** is a system in the microcanonical ensemble that describes the extension of a single photon to account for all the possible frequencies in a chamber. Consider a cubic thermal cavity of side length L . There are two possible polarizations for each of the three directions in the cube.

$$\begin{aligned} \omega_N &= \frac{c\pi}{L} \sqrt{n_x^2 + n_y^2 + n_z^2} = \frac{c\pi}{L} \vec{n} \\ U &= 2 \sum_n \frac{\hbar\omega_n}{e^{\hbar\omega_n/\tau} - 1} = \frac{\hbar V}{\pi^2 c^3} \int_0^{\infty} \frac{\omega^3}{e^{\hbar\omega/\tau} - 1} d\omega = \frac{\pi^2 V}{15(\hbar c)^3} \tau^4 \\ \sigma &= \frac{4\pi^2 V}{45(\hbar c)^3} \tau^3 \\ P &= \frac{1}{3} \frac{U}{V} \end{aligned}$$

Definition 2.6.3. The **Stefan Boltzmann equation** states that the energy per unit volume of a thermal cavity at temperature τ is

$$\frac{U}{V} = \frac{\hbar}{\pi^2 c^3} \int_0^{\infty} \frac{\omega^3}{e^{\hbar\omega/\tau} - 1} d\omega = \frac{\pi^2}{15(\hbar c)^3} \tau^4$$

In classical physics, this energy is defined with an infinite sum that doesn't converge. Therefore, the energy is infinite. This is known as the ultraviolet catastrophe.

Definition 2.6.4. The **spectral density function** of a thermal cavity is the expected energy per unit volume per unit frequency at a given frequency ω

$$\frac{\partial}{\partial \omega} \frac{U}{V} = \frac{\hbar}{\pi^2 c^3} \frac{\omega^3}{e^{\hbar\omega/\tau} - 1}$$

Proposition 2.6.5. The maximal spectral frequency of a radiating object at fixed temperature is found when

$$\begin{aligned} 3 - 3e^{-\hbar\omega_{\max}/\tau} &= \frac{\hbar\omega_{\max}}{\tau} \\ \tau &\approx \frac{\hbar\omega_{\max}}{2.82} \end{aligned}$$

Proposition 2.6.6. From the 1st law of thermodynamics, entropy of thermal radiation is

$$\sigma(\tau, V) = \frac{4\pi^2 V}{45(\hbar c)^3} \tau^3$$

Proposition 2.6.7. Furthermore the pressure of thermal radiation is

$$P = \frac{1}{3} \frac{U}{V} = \frac{\pi^2}{45(\hbar c)^3} \tau^4$$

Proposition 2.6.8. The **Heat capacity** of the photon gas is

$$C_v = \left(\frac{\partial U}{\partial T} \right)_{N,V} = \frac{4V\pi^2}{15(\hbar c)^3} \tau^3$$

Towards deriving some useful insights from this model we will consider a small hole in the side of the cavity. This represents a perfect black body. From the energy density and some geometry we can show that

$$J_\mu = \frac{U}{V} \frac{U}{V} = \frac{\pi^2}{60(\hbar c)^3} \tau^4$$

Definition 2.6.9. The **Stefan Boltzmann Constant** is defined as the relation between the energy emitted per unit time per unit area from a black body.

$$J_\mu = \frac{\pi^2}{60(\hbar c)^3} \tau^4 = \frac{\pi^2 k_B^4}{60(\hbar c)^3} T^4 = \sigma_B T^4$$

Definition 2.6.10. The **emissivity** of an object is a factor between 0 and 1 that determines how much radiation is emitted or reflected. An emissivity of 1 is a perfect black body and 0 is a perfect mirror.

$$J_{real} = e \cdot J_\mu$$

2.7 Phonons in a Solid

The excitations that we have been considering for light can also be generalized to describe vibrations and sound waves in solid materials.

Definition 2.7.1. A **phonon** is a vibrational excitation in a solid. We will consider N atoms in a lattice. The energy levels are

$$E_s = s\hbar\omega$$

The energy levels can be used identically to the photon case to derive a partition function and the average number of phonons at a given frequency and temperature.

$$z = \sum_{s=0}^{\infty} e^{-s\hbar\omega/\tau} = \frac{1}{1 - e^{-\hbar\omega/\tau}}$$

$$\mathcal{P}(s) = \frac{e^{-s\hbar\omega/\tau}}{z}$$

$$\langle s \rangle = \frac{1}{z} \sum_{s=0}^{\infty} s e^{-s\hbar\omega/\tau} = \frac{1}{e^{\hbar\omega/\tau} - 1}$$

$$\langle \epsilon \rangle = \frac{\hbar\omega}{e^{\hbar\omega/\tau} - 1}$$

Definition 2.7.2. A **phonon gas** is a system in the microcanonical ensemble that describes the extension of a single phonon to account for the distribution of frequencies in a solid. The system uses a model of the distribution of frequencies to make predictions

2.7.3 Einstein Model

The einstein model assumes that all the oscillations in a solid are at the same frequency ω .

Proposition 2.7.4. With the Einstein model the average energy and heat capacity are

$$U = \frac{3N\hbar\omega}{e^{\hbar\omega/\tau} - 1}$$

$$C_V = k_B \left(\frac{\partial U}{\partial \tau} \right)_{V,N} = 3N \left(\frac{\hbar\omega}{\tau} \right)^2 \frac{e^{\hbar\omega/\tau}}{(e^{\hbar\omega/\tau} - 1)^2}$$

2.7.5 Debye Model

The Debye model takes into account the possible standing waves in the solid. For a given speed of sound c_s the Debye frequencies are

$$\omega_{n_x, n_y, n_z} = \frac{\pi c_s}{L} \sqrt{n_x^2 + n_y^2 + n_z^2} = \frac{\pi c_s}{L} n$$

For phonons, there are three modes of polarization. Unlike the photon case where the frequency continues to infinity, there is a maximum oscillation frequency that the solid can support. This frequency is determined by the spacing of the lattice and the speed of sound.

Definition 2.7.6. The **Debye cutoff frequency** ω_D is the maximum oscillation frequency that the solid with N atoms, volume V , and speed of sound c_s can support.

$$\omega_D = c_s \left(\frac{6\pi^2 N}{V} \right)^{1/3}$$

Proposition 2.7.7. The average energy of the phonon gas is determined by the following integral.

$$U = \frac{3V}{2\pi^2} \frac{\hbar}{c_s^3} \int_0^{\omega_D} \frac{\omega^3 d\omega}{e^{\hbar\omega/\tau} - 1}$$

In the dimensionless variable $x = \hbar\omega/\tau$ the energy is

$$U = \frac{3}{2} \frac{V\tau^4}{\pi^2(\hbar c_s)^3} \int_0^{x_D} \frac{x^3 dx}{e^x - 1} \quad x_D = \frac{\hbar\omega_D}{\tau}$$

Proposition 2.7.8. In the low τ limit ($\tau \ll \hbar\omega$), this integral can be evaluated exactly. The average energy and heat capacity at this limit are

$$U \approx \frac{3c_s^3 \tau^4 \pi^4}{5\omega_D^3 \pi^2 (\hbar c_s)^3}$$

$$C_V = \left(\frac{\partial U}{\partial T} \right)_{N,V} = \frac{12\pi^4 N k_B}{5(\omega_D \hbar c_s)^3} T^3$$

Chapter 3

Grand Canonical Ensemble

Up until this point we've been working with a system where the total number of particles is fixed. In the Grand Canonical Ensemble the particle number of the system is allowed to exchange with the environment. The temperature, volume, and chemical potential are fixed.

Definition 3.0.1. The **chemical potential** can be used to determine if a system is in diffusive equilibrium.

$$\mu = \left(\frac{\partial F}{\partial N} \right)_{\tau, V}$$

Remark. When several species of particles are present each species has its own μ . In the grand canonical ensemble, each of these chemical potentials is fixed by the environment.

Proposition 3.0.2. Using thermodynamic relations we can derive the following alternate equations for the chemical potential

$$\mu = \tau \log \left(\frac{N \lambda_T^3}{V} \right) = \tau \log \left(\frac{n}{n_Q} \right)$$

$$\mu = -\tau \left(\frac{\partial \sigma}{\partial N} \right)_{U, V}$$

$$\mu = \left(\frac{\partial U}{\partial N} \right)_{\sigma, V}$$

Definition 3.0.3. The **grand partition function** determines the probability of a given energy. For a fixed temperature and chemical potential, the probability of a particular energy ε_n is given by

$$\mathcal{P}(N, \varepsilon_n) = \frac{1}{\mathcal{Z}_\mu}$$

$$\mathcal{Z}_\mu = \sum_N \sum_{n(N)} e^{(\varepsilon_n - \mu N)/\tau}$$

Definition 3.0.4.

$$\lambda = e^{\mu/\tau}$$

$$\mathcal{Z}_\mu = \sum_N \lambda^N \sum_{s(N)} e^{-\varepsilon_s/\tau}$$

$$\langle \mathbb{X} \rangle = \frac{1}{\mathcal{Z}_\mu} \sum_N \sum_s \mathbb{X}(N, s) e^{(-\varepsilon_s - \mu N)/\tau}$$

$$\langle N \rangle = \tau \frac{\partial}{\partial \mu} \log \mathcal{Z}_\mu$$

$$U = \langle \varepsilon \rangle = \frac{1}{\mathcal{Z}_\mu} \sum_N \sum_{n(N)} \varepsilon_n e^{-(\varepsilon_n - \mu N)/\tau} = \tau^2 \frac{\partial}{\partial \tau} (\log \mathcal{Z}_\mu)_\lambda$$

$$\Omega = U - \sigma\tau - \mu N$$

$$\Omega = -\tau \log \mathcal{Z}_\mu$$

$$\sigma = - \left(\frac{\partial \Omega}{\partial \tau} \right)_{V, \mu}$$

$$P = - \left(\frac{\partial \Omega}{\partial V} \right)_{\tau, \mu}$$

$$N = - \left(\frac{\partial \Omega}{\partial \mu} \right)_{\tau, V}$$