



Reaction Kinetics

Reaction kinetics is the study of how fast chemical reactions take place, what factors influence the rate of reaction, and what mechanisms are responsible.

A chemical reaction is usually depicted in the form of a chemical equation which describes the transformation of one or more reactants into one or more products.



Equilibrium if forward rate equals backward rate

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Law of Mass-Action

The rate of an elementary reaction (a reaction that proceeds through only one transition state, that is one mechanistic step) is proportional to the product of the concentrations of the participating molecules. (source: Wikipedia)

Biochemical Reaction

Forward Rate

$$[A] \xrightarrow{k_1^+} [B] \qquad \qquad k_{\perp}^+ [A]$$

$$[A] + [B] \xrightarrow{k_2^+} \rightarrow [C] \longrightarrow k_2^+ [A][B]$$

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Law of Mass-Action

$$[L]+[R] \xrightarrow{k_{+}} [LR]$$

$$\frac{d[L]}{dt} = -k_{+}[L][R] + k_{-}[LR]$$

$$\frac{d[R]}{dt} = -k_{+}[L][R] + k_{-}[LR]$$

$$\frac{d[LR]}{dt} = k_+[L][R] - k_-[LR]$$

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Reactions Representation





















