


STRIVE FOR THE HIGHEST
Western Australian Certificate of Education
Sample Mock Examination
Question/Answer Booklet
NAME
CHEMISTRY
Stage 3

Time allowed for this paper

Reading time before commencing work:

ten minutes

Working time for paper:

three hours

Materials required/recommended for this paper

To be provided by the supervisor

This Question/Answer Booklet
Multiple-choice Answer Sheet
Chemistry Data Sheet
To be provided by the candidate
Standard items: pens, pencils, eraser, correction fluid, ruler, highlighters

Special items: non-programmable calculators satisfying the conditions set out by the Curriculum Council for this course

Important note to candidates

No other items may be taken into the examination room. It is your responsibility to ensure that you do not have any unauthorised notes or other items of a non-personal nature in the examination room. If you have any unauthorised material with you, hand it to the supervisor before reading any further.

OFFICE USE ONLY

	MARK	TOTAL	%
Section 1		50	
Section 2		70	
Section 3		80	
TOTAL		200	

Structure of this paper

Section	Number of questions available	Number of questions to be answered	Suggested working time (minutes)	Marks available	Percentage of exam
Section One: Multiple-choice	25	25	50	25	25
Section Two: Short answer	13	13	70	80	40
Section Three: Extended answer	5	5	60	70	35
					100

Instructions to candidates

1. The rules for the conduct of Western Australian external examinations are detailed in the *Year 12 Information Handbook 2010*. Sitting this examination implies that you agree to abide by these rules.

2. Answer the questions according to the following instructions.

Section One: Answer all questions on the separate Multiple-choice Answer Sheet provided. For each question shade the box to indicate your answer. Use only a blue or black pen to shade the boxes. If you make a mistake, place a cross through that square, do not erase or use correction fluid, and shade your new answer. Marks will not be deducted for incorrect answers. No marks will be given if more than one answer is completed for any question.

Sections Two and Three: Write answers in this Question/Answer Booklet.

3. When calculating numerical answers, show your working or reasoning clearly unless instructed otherwise.
4. You must be careful to confine your responses to the specific questions asked and to follow any instructions that are specific to a particular question.
5. Spare pages are included at the end of this booklet. They can be used for planning your responses and/or as additional space if required to continue an answer.
 - Planning: If you use the spare pages for planning, indicate this clearly at the top of the page.
 - Continuing an answer: If you need to use the space to continue an answer, indicate in the original answer space where the answer is continued, i.e. give the page number. Fill in the number of the question(s) that you are continuing to answer at the top of the page.

SECTION 1: 25 multiple choice questions (25 marks 25 %)

Answer ALL questions in Part 1 on the Separate Multiple Choice Answer Sheet provided, using a 2B pencil. Each question in this part is worth 1 mark.

1. In which of the following lists do all the species have the same electron configuration?

- (a) Na^+ K^+ Rb^+
- (b) Cl^- Ar Ca^{2+}
- (c) Ne Ar Kr
- (d) Cl^- Br^- I^-

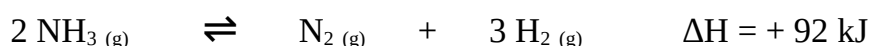
2. A sulfate ion (SO_4^{2-}) contains the isotopes S-33 and O-15.
How many electrons, protons and neutrons does this ion possess?

	electrons	protons	neutrons
(a)	46	48	48
(b)	93	91	96
(c)	50	48	45
(d)	48	45	50

3. Which of the following elemental properties do **NOT** have an increasing trend across Period 3 of the Periodic Table?

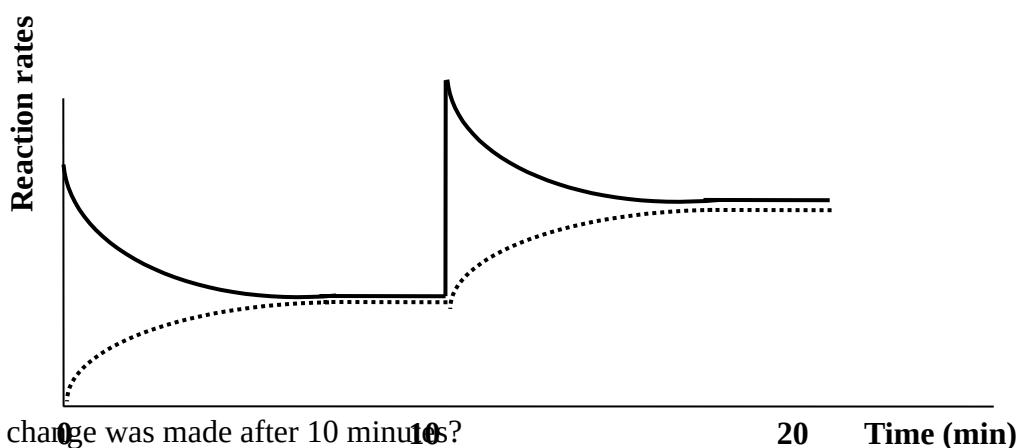
- I. Atomic number
 - II. Atomic size
 - III. Electronegativity
 - IV. Ionization energy
 - V. Melting point
- (a) III and IV
 - (b) I, III and IV
 - (c) II and V
 - (d) I, II, IV and V

4. In which of the following solid substances is there only one type of bonding between the particles?
- Aluminium hydroxide
 - Barium oxide
 - Calcium nitrate
 - Graphite
5. Some ammonia gas is pumped into a sealed container whose volume can be increased or decreased. The reversible decomposition reaction begins immediately.



The forward and reverse reaction rates are measured and after 10 minutes a change is made to the gas system. The reaction rates are measured for a further 10 minutes.

The following graph shows how the forward and reverse reaction rate changed during the 20 minutes.



What change was made after 10 minutes?

- The volume of the gas mixture was decreased.
 - More ammonia was pumped in.
 - The mixture was heated.
 - Some hydrogen was pumped in.
6. Adding a catalyst to an equilibrium system increases
- the proportion of particles that have sufficient energy to react.
 - the proportion of products present.
 - the average kinetic energy of particles.
 - the forward reaction rate more than the reverse reaction rate.

7. Carbon monoxide and chlorine react to form phosgene (COCl_2). The reaction is reversible.



Which of the following conditions will increase the rate of formation of phosgene?

- I. Increasing the temperature
- II. Increasing the pressure
- III. Removal of phosgene
- IV. Decreasing the temperature
- V. Decreasing the pressure

- (a) I and II
- (b) I, II and III
- (c) III, IV and V
- (d) IV and V

8. When carbon dioxide is bubbled through a suspension of calcium carbonate some of the calcium carbonate dissolves as the soluble calcium hydrogencarbonate forms. The following equilibrium is established.



The correct equilibrium expression for this system is

- (a)
$$\frac{[\text{Pb}(\text{HCO}_3)_2]}{[\text{PbCO}_3] [\text{H}_2\text{O}] [\text{CO}_2]}$$
- (b)
$$\frac{[\text{PbCO}_3] [\text{H}_2\text{O}] [\text{CO}_2]}{[\text{Pb}(\text{HCO}_3)_2]}$$
- (c)
$$\frac{[\text{Pb}(\text{HCO}_3)_2]}{[\text{H}_2\text{O}] [\text{CO}_2]}$$
- (d)
$$\frac{[\text{Pb}(\text{HCO}_3)_2]}{[\text{CO}_2]}$$

9. In which of the following is the first species acting as a base?

- (a) $\text{HCO}_3^- + \text{NH}_3 \rightleftharpoons \text{CO}_3^{2-} + \text{NH}_4^+$
- (b) $\text{H}_2\text{PO}_4^- + \text{HCO}_3^- \rightleftharpoons \text{HPO}_4^{2-} + \text{H}_2\text{CO}_3$
- (c) $\text{HPO}_4^{2-} + \text{NH}_4^+ \rightleftharpoons \text{H}_2\text{PO}_4^- + \text{NH}_3$
- (d) $\text{SO}_4^{2-} + \text{Ba}^{2+} \rightleftharpoons \text{BaSO}_4$

10. Which of the following 0.1 mol L⁻¹ solutions has the lowest pH?

- (a) sodium sulfate Na_2SO_4
- (b) ammonium acetate $\text{NH}_4\text{CH}_3\text{COO}$
- (c) ammonium chloride NH_4Cl
- (d) sodium nitrate NaNO_3

11. An acid-base titration is performed to determine the concentration of a sodium hydroxide solution. The flask contains the solution of sodium hydroxide. Standardised ethanoic (acetic) acid is delivered from the burette. A student incorrectly uses the indicator bromophenol blue, which changes colour at about pH 4. Because of this incorrect choice

- (a) too much acid will be delivered and the calculated sodium hydroxide concentration will be too high
- (b) too much acid will be delivered and the calculated sodium hydroxide concentration will be too low
- (c) not enough acid will be delivered and the calculated sodium hydroxide concentration will be too high
- (d) not enough acid will be delivered and the calculated sodium hydroxide concentration will be too low

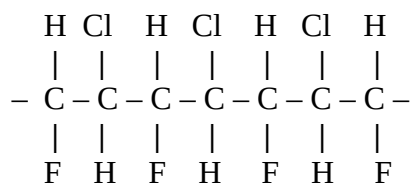
12. Which of the following ions is least likely to act as a base in aqueous solution?

- (a) CO_3^{2-}
- (b) HPO_4^{2-}
- (c) CH_3COO^-
- (d) HSO_4^-

13. A buffer solution is one that does not have its pH changed significantly when a small amount of acid or base is added. Which of the following solutions (where each substance has a concentration of 1 mol L^{-1}) would best act as a buffer?
- (a) A solution of CH_3COOH and HCl
 - (b) A solution of NaH_2PO_4 and Na_2HPO_4
 - (c) A solution of CH_3COOH and H_3PO_4
 - (d) A solution of NaOH and NH_3
14. In which of the following species is the oxidation state of sulfur the lowest?
- (a) SO_3
 - (b) SO_3^{2-}
 - (c) $\text{S}_2\text{O}_4^{2-}$
 - (d) $\text{S}_2\text{O}_6^{2-}$
15. Iodide ion (I^-) can be oxidized by X but not by Y. The identities of X and Y, respectively, could be
- | X | | Y |
|--------------------------------|-----|-------------|
| (a) bromine | and | chlorine |
| (b) gold (III) ions | and | silver ions |
| (c) acidified MnO_4^- | and | nitric acid |
| (d) iron (III) ions | and | nickel ions |
16. A student designs an electrochemical cell. One half cell consists of a nickel rod in a 1 mol L^{-1} nickel (II) sulfate solution. As the cell operates he notices that the green colour of this half cell becomes darker green. Which of the following could correctly describe the other half cell?
- I. Chromium rod in a chromium (III) chloride solution
 - II. Copper rod in a copper (II) sulfate solution
 - III. Lead rod in a lead nitrate solution
 - IV. Zinc rod in a zinc chloride solution
- (a) I and II
 - (b) II and III
 - (c) III and IV
 - (d) I and IV

17. In an electrochemical cell
- (a) there is a flow of electrons.
 - (b) there is a flow of ions.
 - (c) oxidation and reduction occur at the same time.
 - (d) all of the above are occurring.
18. A colourless organic liquid (X) is reacted with an acidified potassium permanganate solution. The product is a liquid (Y). The liquid (X) and liquid (Y) are then reacted to produce a liquid (Z).
- Which of the following general formulae could represent the liquid (Y)?
[R represents the rest of the molecule.]
- (a) RCOOH
 - (b) RCH_2OH
 - (c) RCOR
 - (d) RCOOR
19. Chlorine gas reacts with hydrocarbons by both addition to multiple bonds and substitution of hydrogen atoms. Excess chlorine gas is mixed with 1 mole of each of the following. Which one will react with the most chlorine?
- (a) ethane
 - (b) ethene
 - (c) dichloroethane
 - (d) dichloroethene
20. How many of the following compounds can exhibit geometric (cis-trans) isomerism?
- 1,1 – dibromo propene
 - 1,2 – dibromo propene
 - 2,3 – dibromo propene
 - 3,3 – dibromo propene
- (a) 1
 - (b) 2
 - (c) 3
 - (d) 4

21. Alpha amino acids ($\text{RCHNH}_2\text{COOH}$)
- (a) are strong acids.
 - (b) have an amino group at one end of the molecule and a carboxyl group at the other end of a long hydrocarbon chain.
 - (c) dissolve in water to form ionic solutions.
 - (d) are non-polar molecular substances.
22. One mole of an organic compound was burned in six moles of oxygen, producing four moles of carbon dioxide and five moles of water. Which of the following formulae represents the organic compound?
- (a) $\text{CH}_3\text{CH}_2\text{COCH}_3$
 - (b) $\text{CH}_3\text{CH}_2\text{CH}_2\text{CHO}$
 - (c) $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH}$
 - (d) $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{OH}$
23. Part of a polymer chains is represented by



Which of the following could be the monomer from which the polymer was made?

- I. cis – 1,2 – chlorofluoroethene
 - II. trans – 1,2 – chlorofluoroethene
 - III. 1,2 – chlorofluoroethane
- (a) I only
 - (b) II only
 - (c) I or II
 - (d) I or II or III

24. An industrial cooking oven has a thick layer of grease sticking to the inside walls. Which of the following substances could be used to remove this layer?

I.	Ammonium stearate	$\text{CH}_3(\text{CH}_2)_{16}\text{COONH}_4$
II.	Calcium stearate	$(\text{CH}_3(\text{CH}_2)_{16}\text{COO})_2\text{Ca}$
III.	Sodium hydroxide	NaOH
IV.	Stearic acid	$\text{CH}_3(\text{CH}_2)_{16}\text{COOH}$

- (a) I or II
(b) I or III
(c) II or IV
(d) III or IV
25. Three of the following substances can be polymerized on their own by adding a starter catalyst. Which one of the following will **NOT** polymerize?

- (a) $\text{HOCH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{COOH}$
(b) $\text{H}_2\text{NCH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{COOH}$
(c) $\begin{array}{c} \text{CH}_3\text{CH}_2\text{CHCH}_2\text{COOH} \\ | \\ \text{OH} \end{array}$
(d) $\text{HOOCCH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{COOH}$

SECTION 2 13 questions (80 marks 40 %)

Answer ALL questions in Section 2 in the spaces provided below.

1. Write equations for the reaction that occurs in each of the following procedures.
If no reaction occurs, write 'no reaction'.

In each case describe what you would observe, including any

* colour change

* odour

* precipitate (give the colour)

* Gas evolutions (state the colour or describe as colourless)

If a reaction occurs but the change is not observable, you should state this.

- (a) Oxygen gas is bubbled through an acidified solution of iron (II) sulfate.

Equation _____

Observation _____

_____ (3marks)

- (b) Ethene gas is bubbled through bromine water (aqueous solution of bromine).

Equation _____

Observation _____

_____ (3marks)

2. For each of the following sets of observations:

(i) write a description of any **one** reaction that matches the observations, and

(ii) give an appropriate equation (full or ionic) for **that** reaction.

e.g. A brown solution is added to a colourless solution, producing a brown precipitate.

Reaction *iron (III) nitrate solution is mixed with sodium hydroxide solution.*

Equation $Fe^{3+} + 3 OH^- \rightarrow Fe(OH)_3$

- a) A purple solution is mixed with a colourless solution, producing a colourless solution and a colourless gas

Reaction _____

Equation _____

_____ (3 marks)

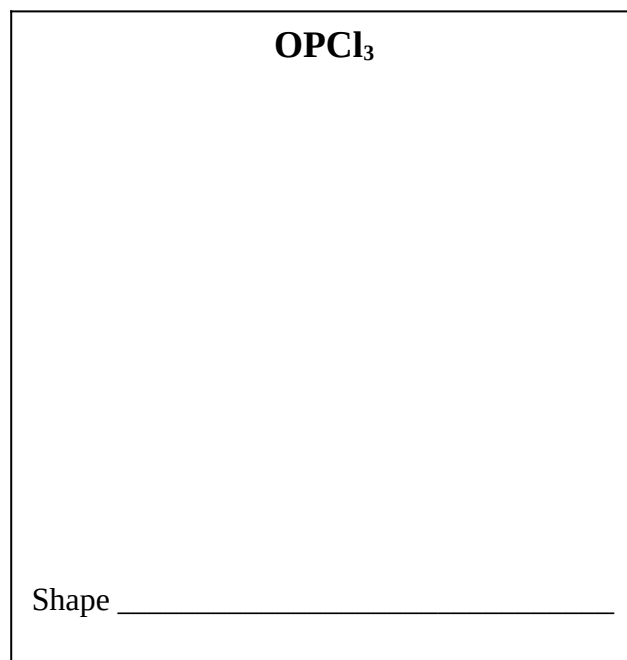
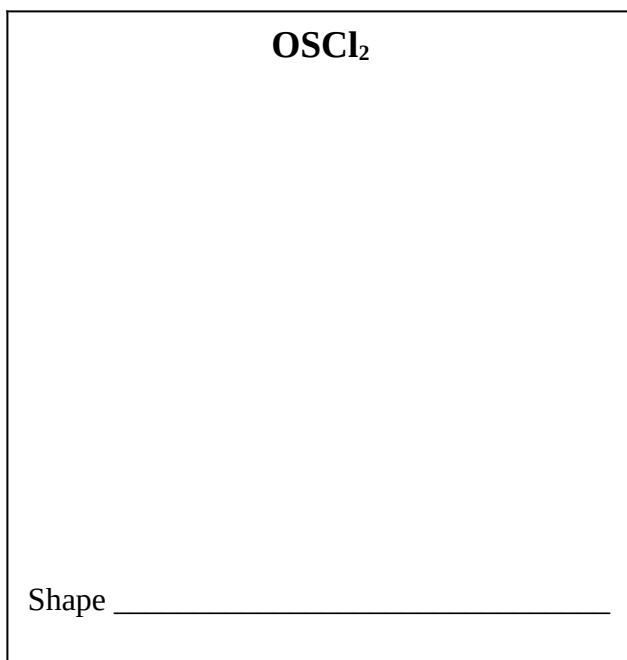
- b) A metal strip is placed in a green solution. Silvery-white crystals form on the strip and the green colour fades.

Reaction _____

Equation _____

(3 marks)

3. Draw electron-dot diagrams showing the arrangement of all valence electrons in the following chemical species.
Describe the shape of each (eg: linear/bent/etc)



(6 marks)

4. Methane reacts with fluorine to form four different fluorinated compounds.
Write the names and formulae of all the fluorinated methane molecules that are **polar**.

Name	Formula
_____	_____
_____	_____
_____	_____
_____	_____
_____	_____

(4 marks)

5. The following table shows the solubilities of two amines in water.

Amine	Methyl amine CH_3NH_2	Dodecyl amine $\text{CH}_3(\text{CH}_2)_{11}\text{NH}_2$
Solubility (g/100 mL)	108	0.05

Explain why their solubilities are so different.
Include a labelled diagram.

(6 marks)

6. Three unlabelled beakers each contain the same volume of 1 mol L⁻¹ solution. The three solutions are:
- sodium hydrogensulfate (NaHSO₄)
 - sulfuric acid (H₂SO₄), and
 - phosphoric acid (H₃PO₄).

The student is asked to identify the solutions. He is also given a bottle of sodium hydroxide (NaOH) solution, a choice of indicators and is allowed to use any other item of laboratory glassware. The student was successful.

How did the student correctly identify the acids?

Include equations to support your answer.

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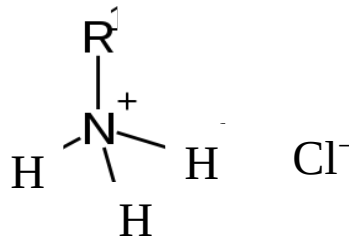
(7 marks)

7. Quaternary ammonium salts can be represented by the following structural formula.

If the alkyl group (**R**) is long then the salt acts like a soap or detergent. If it is short the salt has no cleaning properties.

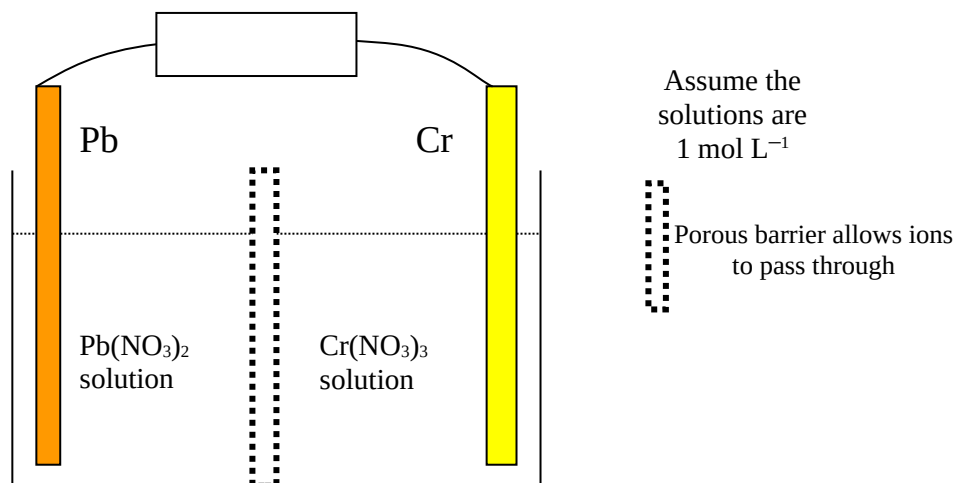
Explain these two differences in properties.

Include a labelled diagram.



(6 marks)

8. An electrochemical cell contains the two half cells separated by a porous membrane, which allows ions to migrate through. Each half cell has a metal rod placed in a solution of its nitrate.



- (a) Write the two half reactions that occur, their standard reduction potentials and state whether each is oxidation, or reduction,

_____ $E^\circ =$ _____
 _____ $E^\circ =$ _____

(4 marks)

- (b) Write the equation for the net redox equation.

(2 mark)

- (c) What is the emf (electromotive force, or voltage) of the cell?

(1 mark)

- (d) Draw an arrow in the top box to show the direction of current (electron flow) in the wire connecting the two electrodes.

(1 mark)

- (e) What change (or changes) will be observed in the cell?

(3 marks)

10. The following table gives information about two substances. Use the information to determine whether each substance is acting as an oxidising agent (oxidant), or reducing agent (reductant) and provide a brief explanation to justify your answer.

Substance	Information	Oxidant, or reductant?
Concentrated sulfuric acid H_2SO_4	Reacts with copper to produce sulfur dioxide.	Expalnation:
Hydrogen peroxide H_2O_2	Reacts with chlorine to produce chloride ion.	Explanation:

(4 marks)

11. A student pours some silver nitrate solution into a bronze (copper-tin alloy) container in order to store it.

Is this wise?

Explain why, or why not. Include an equation.

(3 marks)

12. Vinegar is about 4% by mass acetic acid and is safe to consume in foods. The same strength sulfuric acid is not safe to consume. Explain why. Include equations.

(4 marks)

13. Name, and draw structural diagrams for, the following organic compounds.

Compound	Structural diagram	Name
An isomer of dibromobutane		
An ester containing 4 carbon atoms		
The ketone (alkanone) with the least number of carbon atoms		

(9 marks)

SECTION 3

5 questions (70 marks 35 %)

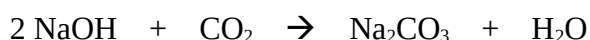
Extended answers

Answer ALL questions in Section 3 in the spaces provided.

1. Treatment of waste by-products in chemical industry

16 marks

In a chemical industries complex one production plant produces a waste caustic soda (NaOH) solution, which it stores in a large pond. Another production plant produces waste carbon dioxide. The chemical engineers decide to combine both wastes to produce the environmentally friendly by-product, sodium carbonate, by bubbling the carbon dioxide through the caustic soda solution.



The caustic soda pond contains 500 kL and has a hydroxide (OH^-) concentration of $1.00 \times 10^{-2} \text{ mol L}^{-1}$.

- (a) What is the pH of the solution?

(2 marks)

- (b) What is the mass of sodium hydroxide in the caustic soda pond?

[illegible]

(3 marks)

- (c) What mass of carbon dioxide is needed to completely react with sodium hydroxide?

If you did not answer Part (b) above, use a mass of 100 kg sodium hydroxide

(4 marks)

- (d) The carbon dioxide is first cooled to 10°C and is pumped at a pressure of 200 kPa, delivering 150 L per minute.

How long does it take to complete the reaction?

(5 marks)

- (e) The pond solution is still found to be alkaline (pH of about 9).

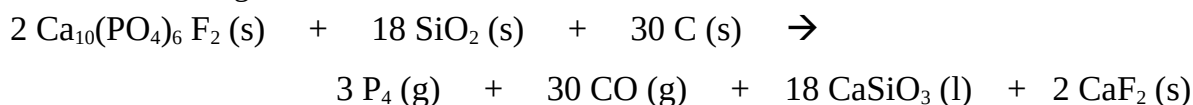
Assuming all the carbon dioxide has reacted suggest a reason why is it still alkaline.

(2 marks)

2. Production of phosphorus from fluoroapatite**16 marks**

The mineral fluoroapatite $[\text{Ca}_{10}(\text{PO}_4)_6\text{F}_2]$ is mixed with sand $[\text{SiO}_2]$ and powdered carbon in a high temperature furnace. The phosphorus is produced as a gas $[\text{P}_4]$, along with carbon monoxide. The reaction actually produces calcium oxide $[\text{CaO}]$, which has a very high melting point. This would make the mixture difficult to control. So, as the calcium oxide is produced it reacts with the sand to form a low melting point slag, calcium silicate $[\text{CaSiO}_3]$. This liquid slag is easily separated from the furnace.

The reaction occurring is:



- (a) Is this reaction exothermic, or endothermic? _____
Give a reason for your choice.

(2 marks)

- (b) The main reaction can be represented by the two half reactions:
- phosphate ion producing phosphorus (P_4) and oxide ions (O^{2-}), and
 - carbon reacting with oxide ion producing carbon monoxide

Which element, phosphorus or carbon, is being oxidised? _____

Justify your answer by referring to oxidation numbers.

- (c) List three elements whose oxidation states are not changing.

(6 marks)

- (d) Some of the oxide ions produced in Part (b) becomes part of the liquid slag by reacting with calcium ions and sand.

Write the equation for the formation of the slag.

(2 marks)

- (d) In a laboratory trial a 155 g sample of fluoroapatite (molar mass = 1008.62) is heated with excess sand and 25.0 g of carbon.
What mass of phosphorus would be produced?

[illegible]

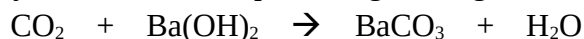
(6 marks)

3. Analysing an organic compound**13 marks**

A certain organic compound is known to contain only carbon, hydrogen and oxygen.

The compound was analysed as follows.

- A 2.149 g sample was burned and the carbon dioxide produced was bubbled through a barium hydroxide solution, producing 11.27 g of barium carbonate (BaCO_3).



- The mass of water produced by burning of the sample was 0.7721 g
- The compound was found to have a molecular weight of 150.1

a) What is the empirical formula of the compound? (10 marks)

b) What is the molecular formula of the compound? (2 marks)

c) The compound is also known to be a carboxylic acid; that is, containing one COOH group.
Write the molecular formula in the form of $\text{C}_x\text{H}_y\text{O}_z\text{COOH}$ (giving values for X, Y and Z).

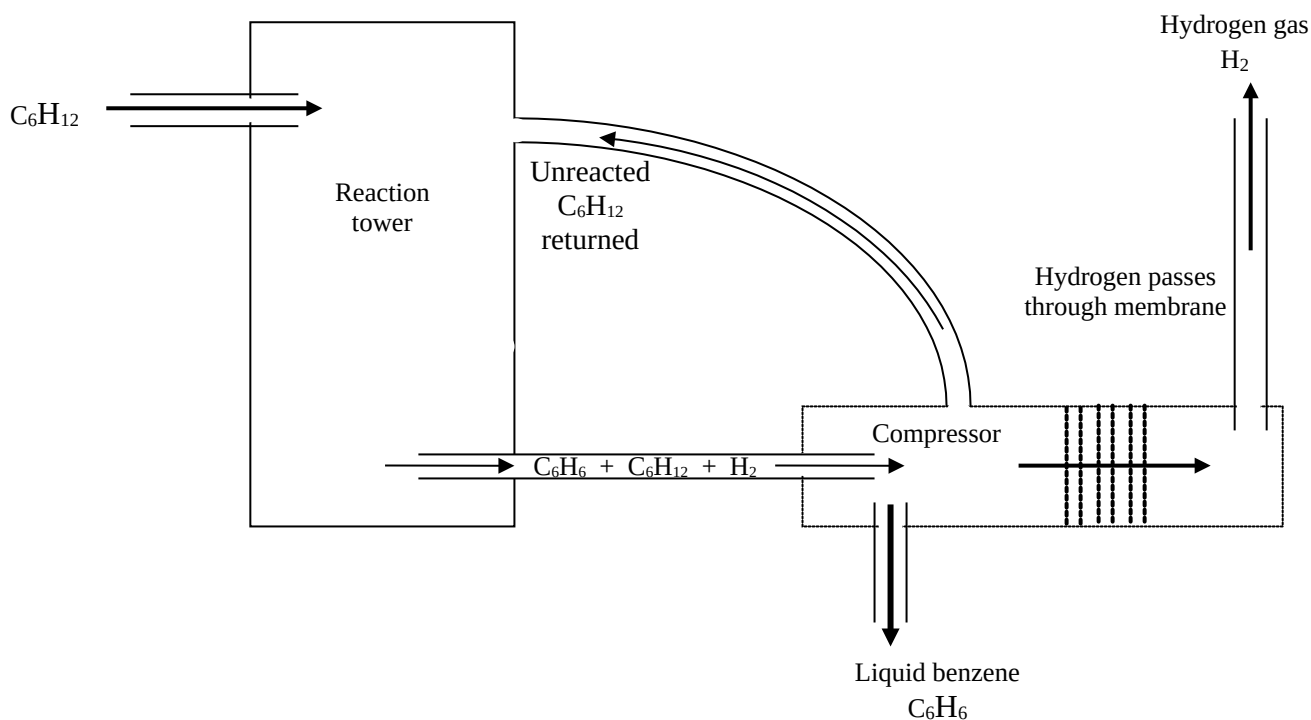
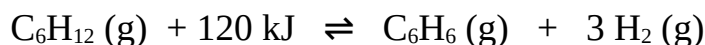
(1 mark)

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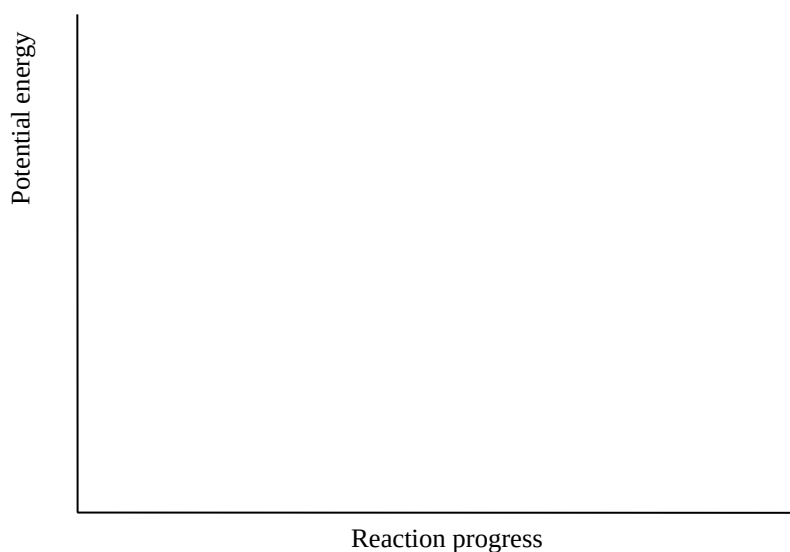
4. Production of benzene

14 marks

Benzene (C_6H_6) can be produced by the dehydrogenation of cyclohexane (C_6H_{12}) gas. The reaction has a high activation energy (880 kJ mol^{-1}), is also endothermic and reversible. The cyclohexane (C_6H_{12}) passes through a special reaction tower where hydrogen is chemically removed. The benzene/cyclohexane/hydrogen mixture then passes through a compressor, where the benzene is liquefied. A special membrane in the compressor allows the small hydrogen molecules to pass through, and out. The unreacted cyclohexane (C_6H_{12}) gas is then returned to the reaction tower.



- a) Draw a labelled energy profile diagram for the reaction.



(3 marks)

- b) Write an equilibrium constant expression for the reaction.

(2 marks)

- c) Under what conditions will the rate of the forward reaction be greatest?

(3 marks)

- d) For a mixture of all three gases at equilibrium in a sealed container, what conditions will produce the maximum yield of benzene?

(2 marks)

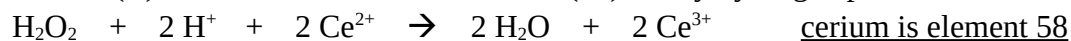
- e) Suggest conditions that would be used for the commercial production of benzene using this process.

Explain why you chose these conditions.

(4 marks)

5. Determining concentration of cerium (II) sulfate solution by titration 10 marks

Cerium (II) ion can be converted to cerium (III) ion by hydrogen peroxide.



A solution of cerium (II) sulfate was analysed by the following steps:

- I. 50.00 mL of the solution was diluted to 500.0 mL in a volumetric flask
- II. 20.00 mL of this diluted solution was pipetted into a conical flask
- III. About 20 mL of dilute sulfuric acid was added to the flask
- IV. Standardised hydrogen peroxide solution of concentration $0.05145 \text{ mol L}^{-1}$ was delivered from a burette
- V. 35.45 mL of the hydrogen peroxide was required for complete reaction

What was the concentration in moles per litre (mol L^{-1}) and in grams per litre (g L^{-1}) of the original undiluted cerium sulfate solution?

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END of EXAM