Trinity College

Chemistry
Organic Chemistry Test

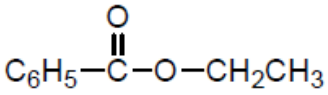
**DO NOT MARK THIS PAPER**

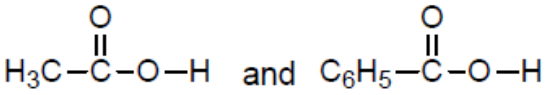
## **Multiple** choice 10 marks

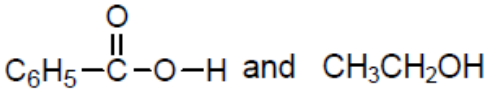
1. Which of the following is **not** a pair of isomers?

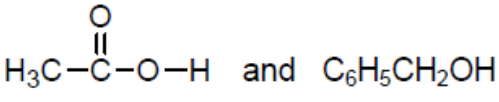
   A. ethyl benzene ($C_6H_5$-$C_2H_5$) and dimethyl benzene, $C_6H_4(CH_3)_2$
   B. 1-propanol ($CH_3CH_2CH_2OH$) and 2-propanol ($CH_3CHOHCH_3$)
   C. ethanol ($C_2H_5OH$) and dimethyl ether ($CH_3OCH_3$)
   D. 2-butanone ($CH_3COCH_2CH_3$) and 1-butanol ($CH_3CH_2CH_2CH_2OH$)

2. Which reagents react to give ethyl benzoate ($C_6H_5COOC_2H_5$) and water? The structure of ethyl benzoate is given below.



**A** 

**B** 

**C** 

**D** $CH_3CH_2OH$ and $C_6H_5OH$

3. In which of the following alternatives are the three compounds listed in order of increasing boiling point?

A. Pentane, butan-1-ol, propanoic acid
B. Propanoic acid, butan-1-ol, pentane
C. Propanoic acid, pentane, butan-1-ol
D. Butan-1-ol, propanoic acid, pentane

4. The raspberry-flavoured food additive, butyl methanoate, can be prepared from $CH_3CH_2CH_2CH_2OH$ using

A. an addition reaction with HCOOH.
B. an addition reaction with $CH_3COOH$.
C. a condensation reaction with HCOOH.
D. a condensation reaction with $CH_3COOH$.

5. Which compound is least soluble in water?

A. $CH_3CH_2CH_2F$
B. $CH_3CH_2CH_2NH_2$
C. $CH_3CH_2CH_2OH$
D. $CH_3CH_2CH_2COOH$

6. How many different aldehydes have the formula $C_5H_{10}O$?

A. 2
B. 3
C. 4
D. 5

7. The boiling points of $CH_3COCH_3$, $CH_3COC_2H_5$, and $CH_3COC_3H_7$ are 56 °C, 80 °C, and 102 °C, respectively. This increase is best attributed to an increase in which of the following?

**I** **dipole-dipole interactions**
**II** **dispersion forces**
**III** **hydrogen bonding**

A. I only
B. II only
C. III only
D. II and III only

8. Which of the following statements would apply to compounds that belong to the same homologous series?

**I** **they have similar physical properties**
**II** **they have similar chemical properties**
**III** **they contain the same functional group**
**IV** **they have the same molecular formula but different structures**

A. III only
B. IV only
C. II and III only
D. I, II, III and IV

9. Pure samples of which of the following exhibit hydrogen bonding?

**I** **$CH_3OH$**
**II** **$CH_3NO_2$**
**III** **$CH_3CN$**

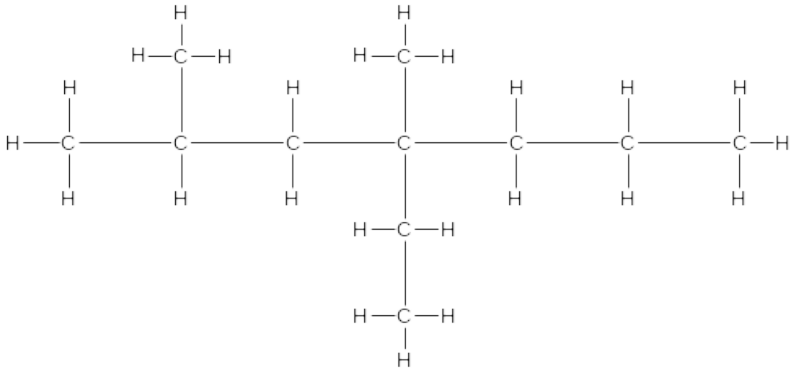
A. I only
B. I and II only
C. II and III only
D. I, II, and III

10. The compound that is **not** an isomer of 2,2,4-trimethylpentane is

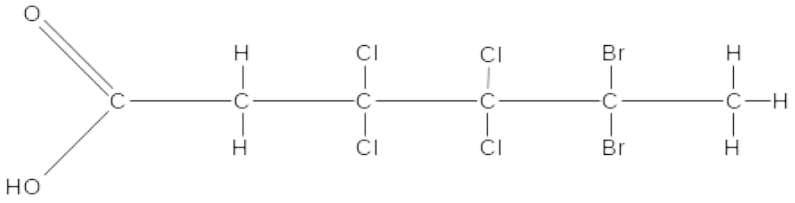
A. octane
B. 3-ethylhexane
C. 2,4-dimethylpentane
D. 2,4-dimethylhexane

1. Give the IUPAC name of the following structures:

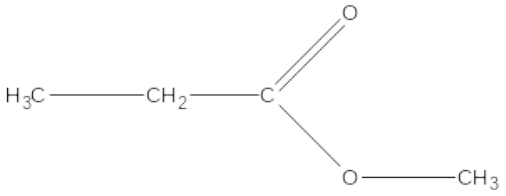
(a)



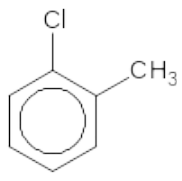
(b)



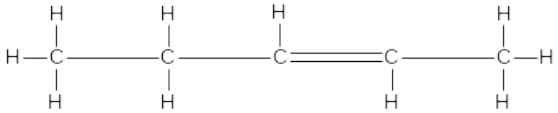
(c)



(d)



(e)



(f)

$CH_3CH_2CH_2COONa$

[6 marks]

2. Give the <u>full structural formula</u> for the following organic chemicals:

(a) 1,2,3 - propantriol

(b) 6-amino-7-bromo-3-heptanone

(c) 2-ethylhexyl ethanoate

(d) *cis*-oct-3-ene

(e) *trans*-3-heptene

(f) 2,4-dimethylpentan-1-amine

[6 marks]

3. For each of the situations described below, determine whether or not a redox reaction would be expected and, if so:

(i) Write a balanced redox reaction showing the changes that take place;

(ii) Give a brief observation that would be expected to accompany the reaction.

(a) Acidified potassium permanganate solution is added to 2-methyl-3-pentanol.

(b) Limited acidified potassium dichromate is added to hexanol.

[8 marks]

4. DRAW and NAME the major organic PRODUCT or PRODUCTS in the following reactions assuming appropriate conditions. NB. No balancing is required.

(a) Ethanol and pentanoic acid with $H_2SO_4$ as a catalyst.

(b) Ethane and chlorine.

(c) Bromine and Benzene in the presence of UV light.

(d) Pentene and hydrogen chloride in the presence of a catalyst

[8 marks]

6. (a) Elementary analysis of a compound indicated that it contained only carbon, hydrogen, nitrogen and oxygen. A 1.279g sample was burned

completely in oxygen such that all the carbon was converted to carbon dioxide and the hydrogen to water. This resulted in 1.600g of carbon dioxide and 0.770g of water. A separate 1.279g sample was shown by analysis to contain 0.1697g of nitrogen. Calculate the empirical formula of the compound.

(b) Given that the molecular mass of the compound was found to be $105 g.mol^{-1}$, determine the molecular formula.

(c) Given that the compound is a primary amine, reacts rapidly with sodium metal yielding an alkanoate and can be neutralized with NaOH, draw a possible structure.

[8 marks]