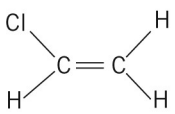
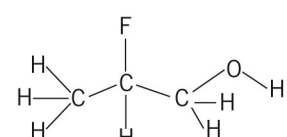
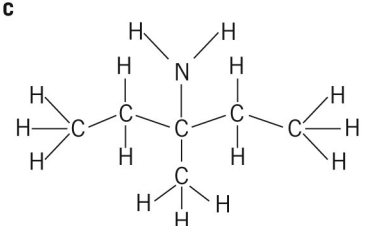
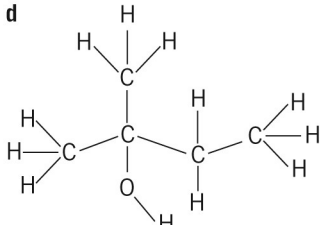
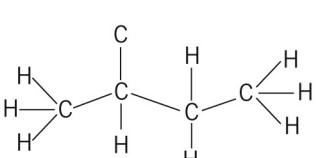
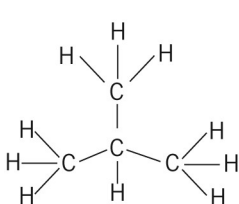


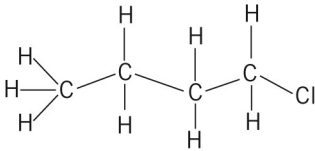
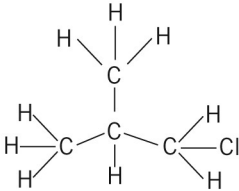
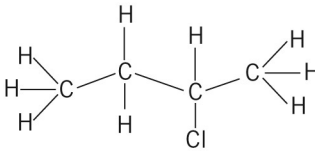
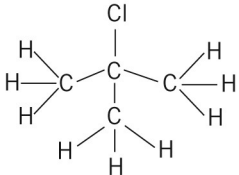
Worksheet 12.1: Solutions

Naming organic compounds

No.	Answer
1	a 2-fluoropropene b propane c 3-methylbutanoic acid d 2-methylpropan-1-ol e trichlorofluoromethane
2	<div style="display: flex; justify-content: space-around; align-items: flex-start;"> <div style="text-align: center;"> a  </div> <div style="text-align: center;"> b  </div> <div style="text-align: center;"> c  </div> </div> <div style="text-align: center; margin-top: 20px;"> d  </div>
3	a C_nH_{2n+2} b C_nH_{2n} c $C_nH_{2n+1}OH$ d $C_{n-1}H_{2n-1}COOH$
4	<div style="display: flex; justify-content: space-around; align-items: center;"> <div style="text-align: center;"> a  butane </div> <div style="text-align: center;">  2-methylpropane </div> </div> <p>b Butane has the higher boiling point, because the more linear molecule means that molecules pack well together, giving greater contact between adjacent molecules and hence stronger dispersion forces. Stronger forces mean a higher boiling point. The branched molecules of 2-methylpropane do not pack together very well.</p>
5	Pent-1-ene (<i>cis</i> and <i>trans</i>) pent-2-ene, 2-methylbut-1-ene, 2-methylbut-2-ene, 3-methylbut-1-ene

Worksheet 12.1: Solutions

Naming organic compounds

No.	Answer
6	<div style="display: flex; flex-wrap: wrap; justify-content: space-around;"> <div style="text-align: center;">  <p>1-chlorobutane</p> </div> <div style="text-align: center;">  <p>1-chloro-2-methylpropane</p> </div> <div style="text-align: center;">  <p>2-chlorobutane</p> </div> <div style="text-align: center;">  <p>2-chloro-2-methylpropane</p> </div> </div>