

Chapter test

Chapter 9 Synthesis reactions

Name: _____

Class: _____

Time permitted: 50 minutes

	Section	Number of questions	Marks available	Marks achieved
A	Multiple choice	15	15	
B	Short answer	5	15	
	Total	20	30	

Grade: _____

Scale:

A+	29–30	A	26–28	B	23–25	C	19–22	D	15–18	E	9–14	UG	0–8
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Comments:

Section A Multiple choice (15 marks)

Section A consists of 15 questions, each worth one mark. Each question has only one correct answer. Circle the correct answer. Attempt all questions. Marks will not be deducted for incorrect answers. You are advised to spend no more than 15 minutes on this section.

- 1 Which of the following is not a synthetic material?
 - A Nylon
 - B Lycra
 - C Teflon
 - D Silk
- 2 Retrosynthetic analysis is:
 - A a process developed back in time (retrospectively).
 - B a process used to determine how to make a compound.
 - C analysis of new materials.
 - D a method of analysis.
- 3 Reaction mechanisms do not include:
 - A multi-step reaction steps.
 - B linear sequences of reactions.
 - C common pathways for different materials.
 - D convergent reaction steps.
- 4 What is Le Chatelier's principle important for?
 - A Maximising yields
 - B Improving purity of product
 - C Choosing reaction pathways
 - D Reducing waste products
- 5 A catalyst is important in producing new materials, because:
 - A it costs little to produce.
 - B is often a useful by-product.
 - C produces the same result by an alternative reaction.
 - D provides an alternative and cheaper pathway for a reaction.
- 6 A micelle is:
 - A produced by soap action on grease.
 - B a name of a detergent molecule.
 - C an important step in the saponification process.
 - D an organic molecule used to make soap.
- 7 Which of the following is not a biofuel made in Australia?
 - A Ethanol

- B** Bioethanol
 - C** Premium petrol
 - D** Biodiesel
- 8** Green chemistry is an initiative designed to:
- A** prevent pollution, treat chemicals to make them safe and dispose of them safely.
 - B** reduce pollution, neutralise chemicals and dispose of them.
 - C** treat pollution, reduce chemical waste, and produce disposal methods.
 - D** prevent pollution, use safe solvents and dispose of them quickly.
- 9** Atom economy means:
- A** the amount of product atoms less the amount used as reactants.
 - B** the percentage waste of atoms used to produce a product.
 - C** the molar fraction of reactants to products used to produce a product.
 - D** the fraction of the mass atoms in reactants to products as a percentage.
- 10** What process is commonly used to produce sulfuric acid?
- A** The Haber process
 - B** The acid rain process
 - C** The acid synthesis process
 - D** The Contact process
- 11** What is the function of limiting reagents?
- A** To limit the effectiveness of catalysts in reactions
 - B** To limit how much product will be made in a reaction
 - C** To limit the amount of steps in a synthesis reaction
 - D** To reduce the yield produced in a reaction
- 12** How does biodiesel differ from crude-based diesel?
- A** One has an oxygen molecule.
 - B** Biodiesel is an ester.
 - C** Diesel has up to 21 carbon atoms in a linear chain.
 - D** Biodiesel is made from fossils.
- 13** The two main processes for producing ethanol exclude:
- A** fermentation.
 - B** hydration of ethene.
 - C** the reaction of water and ethene.
 - D** the Haber process.
- 14** An excess reagent:
- A** is wasted in synthesis reactions.
 - B** is the other reactant that reacts with a limiting reagent.

- C** is usually sent back to the manufacturer to reduce costs.
- D** has limited effectiveness, hence it is left over.

15 A successful product:

- A is determined by the consumer response only.
- B has successful reactants.
- C is determined by different factors and viewpoints.
- D has a low cost and high profit.

Section B Short answer (15 marks)

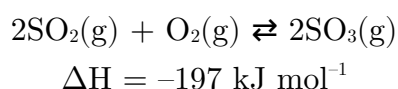
Section B consists of five questions. Write your answers in the spaces provided. You are advised to spend 25 minutes on this section.

1 The synthesis of sulfuric acid involves several steps:

Step 1: The formation of SO_2



Step 2: The formation of SO_3



Step 3: Dissolving SO_3 in water to produce oleum



Step 4: Reaction of oleum and water to produce sulfuric acid



The second step can be affected by a number of factors; comment on the following factors:

- a proportion of reactants
- b temperature
- c pressure.

(1 + 1 + 1 = 3 marks)

2 a Complete the following equation: a zinc chloride solution reacting with a sodium phosphate solution to form insoluble zinc phosphate and a soluble salt.



- b If you had 4.32 g of ZnCl_2 and 222.0 mL of 0.100 M Na_3PO_4 , identify the limiting reactant.

b What mass of precipitate would be recoverable?

(1 + 1 + 1 = 3 marks)

3 Describe and compare the cleaning action of soaps with detergents. What advantages do detergents have in what solutions?

(= 2 marks)

4 a Write out the formula for calculating atom economy.

b Why is this an important consideration in chemical processes?

(1 + 1 = 2 marks)

5 a Calculate the atom economy of the following processes for producing ethanol: hydration of ethene, and fermentation of sugar.

b Compare these two processes and comment on their sustainability.

(3 + 2 = 5 marks)