# 16 – Redox equilibria

## Redox review

Oxidation - e- #o - agents are Reduced to Oxidize other molecules

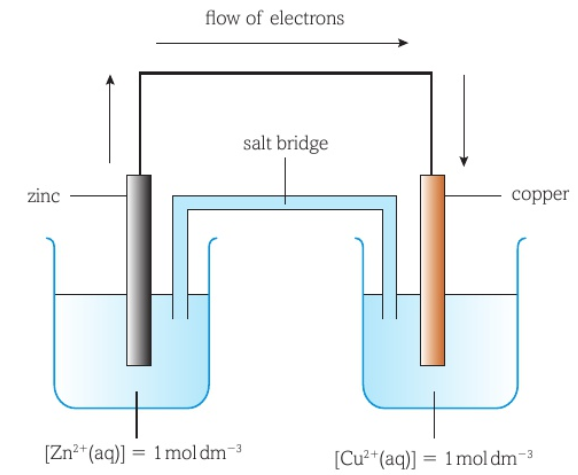
Reduction - e- #o - agents are Oxidized to Reduce other molecules

Disproportionation reaction - Elements of the same compound undergoes both oxidation and reduction

Standard conditions - 100kPa, 298K, 1 mol dm-3

## Electrical chemical cell

Standard electrode potential - Potentials measured of cell against std. H2 electrode at standard conditions



Largerreduced, smaller oxidized

Largerstronger oxidizing / weaker reducing agents

Flipping ECS formulas flips the sign of

The compounds in the cell can be determined from ECS equations of the cell

(Zn oxidized, Cu reduced)

Electrolytes (aq in beaker) must be **soluble**. (Group 1 nitrates used as always soluble)

Salt bridge: NaCl

### Electro-cell diagram

e- flows from smaller larger

Left oxidized, right reduced.

### Shorthand display

Example:

### Measuring

The standard hydrogen electrode is used to measure.

Swap out H2 with anything if different potential is measured.

Half equation:

A reference electrode (H2) is needed as we can only measure a potential difference

Porous Pt is used as to increase surface area

### Feasibility

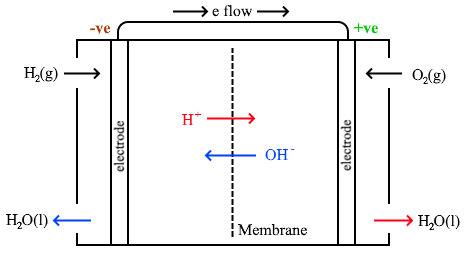
Kinetically stable -

Thermodynamically stable -

reaction is feasible (spontaneous).

This only predicts that reaction is possible, not that it will occur as it depends on factors like

## Hydrogen-Oxygen Fuel cells

* Both sides overall equations are the same: (so same )
* The membrane separates H2 and O2 gases
* Electrodes are coated in Pt as catalyst
* Alkali / Acid is used as electrolytes allowing movement of H+

### Electrolyte equations

Anode:

Cathode:

Acidic: add H+

Alkaline: add OH-

**Acidic**

Anode (-):

Cathode (+):

**Alkaline**

Anode (-):

Cathode (+):

### Advantages & disadvantages

|  |  |
| --- | --- |
| * Environmentally friendly * No harmful product pollutants * Alternative to use of fossil fuels | * H2(g) is a flammable gas * H2(g) is not renewable * Storage of H2(g) has hazards * Storage of H2(g) is costly |

## Titration related

Common color change:

Color change: PurplePale pink

Acidic conditions are used to prevent formation of (brown p.p.)

# 17 – Transitional metals

Transitional metals - Metals that can form > 0 stable ions with partially filled d subshell

## Electronic configuration

If moving 4s e- to 3d, d is d5 / d10, there is more stability for half-filled subshells, so e- would leave 4s.

When forming ions, 4s e- leaves first.

[Ar] corresponds to configuration up to 3d6

## Introduction to complexes

Ligands - Species with lone pair of e- that can form dative bond to t.m. ion

Co-ordinate bonds (c.o. bonds) - Dative covalent bond from ligands

Co-ordination number (c.o. num) - Number of c.o. bonds



### Shapes

Large ligands (Cl-) 4 bonds: Tetrahedral 109.5°

Small ligands (H2O / NH3) forms 6 bonds: Octahedral 90°

*Number of ligands around metal:*

* Ligands repel each other as far as possible
* Giving shape with bond angles of X°

### Naming

H2O aqua | NH3 ammine | X- halo | OH- hydroxo (All negatives end with -o)

### Multidentate ligands

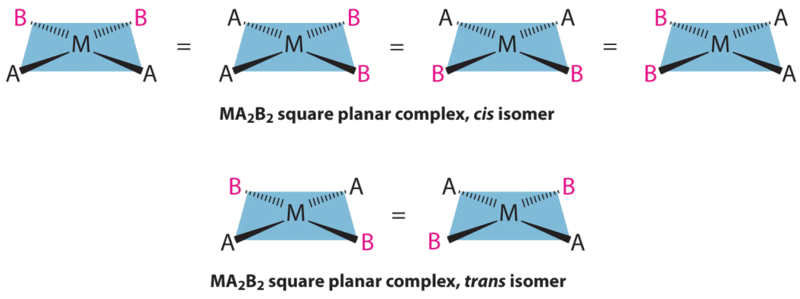
Multidentate ligands - Ligands that has > 1 lone pair of e- that can bond to t.m. ion

Names of multidentate is related to number of c.o. bonds formed.

### Ligand isomerism

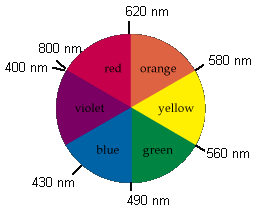
Optical isomerism: only 3 bidentate ligands

Cis-trans geometric isomerism: quad-bi octahedral ligands / bi-bi square planar ligands

Cis - same side | Trans - opposite side

## Complex reactions & color

### Cause of colors – 3d Energy levels

1. Partially filled d-orbital
2. Ligands cause d-d splitting
3. Visible spectrum radiation absorbed
4. Light causes d-d transitions
5. Complementary color observed

If no partially filled d-orbital, there will be no d-d transitions hence no color

### Reactions

1. Deprotonation Ligands ± H+
2. Ligand exchange Ligands replaced
3. Redox Oxidation number
4. C.o. number change Number of ligands in complex

## Summary of reactions

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| **Type** | **Ligand** | **Formula** | **Exception** | **Metals** |
| D | OH- |  |  | All |
| D | OH- (Excess) |  |  | **Cr Zn** |
| D | NH3 |  |  | All |
| L | NH3 (Excess) |  |  | **Cu Co Cr Ni Zn** |
| L | Conc. Cl- |  |  | **Cu Co** |

|  |  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- | --- |
| **T.M.** | **Oxi.** | **Color** | | | **OH- / NH3** *p.p.* | | **Excess OH-** *sol.* | **Excess NH3***sol.* | **Conc. Cl-** |
| Cu | 2 | B | | | B | |  | B | dBr |
| Co | 2 | P | | | B | | Br | B |
| Cr | 2, 3, 6 | B | G | Y | G | | G | G |  |
| Ni | 2 | G | | | G | |  | B |
| Zn | 2 | c | | | W | | c | c |
| Fe | 2, 3 | pG | Y | | dG | Br |  |  |
| Mn | 2 | pP | | | W | |

* Fe2+ turns to Fe3+ upon standing (Dark green to brown)
* Most other reactions can be explained by ECS balancing & comparison

### Chromate reactions

Use ECS to form equations.

## Catalyst properties

Transitional metals can form compounds with catalytic properties.

### Homogeneous

* Phase: Same phase as [R]
* Mechanism: React with [R] to form **intermediates → products** + **re-form** catalyst
* Reformation: Yes

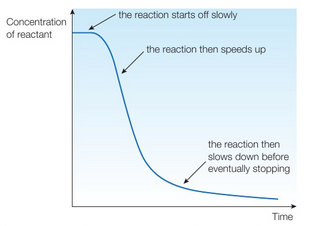
Note that the Ea of forming intermediate + products is lower than the uncatalysted reaction

### Heterogeneous

* Phase: Different phase as [R]
* Mechanism:

1. R A**d**sorb onto surface
2. Weakens bond in R
3. Products then desorb

∴Surface area ↑Rate ↑

**Support mediums** are used to ↑area → ↓cost

* Reformation: Not used up

Advantage over homo: Does **not need separation from products**

### Autocatalysis

**Product of reaction acts as catalyst**: [P] ↑→ rate ↑

* Phase: Same phase as [R]
* Mechanism:

1. Reaction is slow at start
2. Catalyst [P] is formed → reaction speeds up (rate↑)
3. [R] ↓→ reactions slows down (rate↓)

* Reformation: None after reaction

# 18 & 19 – Finally done with Organic Chemistry

## Arene

Phenyl / Benzene prefix

### Disproving Kekule structure of Arene

|  |  |
| --- | --- |
| Phenomenon | Explanation |
| of reaction with H2 is less than expected | Delocalization of more stable than Kekule |
| All C-C bond in Arene has the same length | Delocalization of same bond length instead of diff for C=C |
| Br water is not decolorized | Delocalization of more stable Substitution instead of addition occursstability maintained |

### Reaction speeds

**Alkylarene** alkyl group donates e- to ring, making it a stronger nucleophile

**Phenol**:

1. OH group attached to ring
2. lone pair of e- on O overlaps with cloud & donated to the ring
3. e- density of ring increased

### Arene reactions

Reaction with oxygen:

Halogenation: (ArBr3 Anhydrous)

Nitration: (H2SO4 & HNO3 reactant)

Friedel-crafts alkylation: (AlCl3 catalyst)

Friedel-crafts acylation: (AlCl3 catalyst)

Sulfonation: (Ar warmed with sulfuric acid)

### Phenol reactions

Bromination: (Room temp.) (Orange sol.white p.p.)

## Amines RNH2

### Preparation

Halogenoalkanes: (Excess, hot ethanolic NH3 under pressure) [NS]

Nitriles reduction: (LiAlH4)

Nitriles catalyzed: (Nickel catalyst)

### Basicity

Higher availability of N lone pair e-stronger base

2° > 1° > NH3 > ArNH2

∵ Alkyl groups are e- pushing

∵ Lone pair is pulled into arene ring

### Reactions

*Amine / phenylamine have the same reactions*

With water:

With acids: (revert with OH-)

With acyl chlorides:

Halogenation:

2° to 3°:

3° to 4°:

### Preparing aromatic amines

(Tin [Sn] mixed with HCl, heat under reflux)

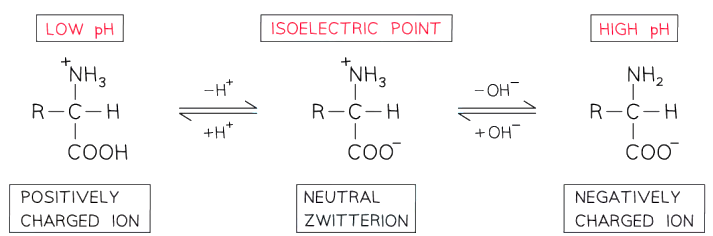
## Amino acids

Compounds that contain NH2 and COOH functional groups

* They are amphoteric, acting as both acid and base. The “R” group decides the basicity
* Amino acids undergoes reaction of amines with acids & carboxylic acid with bases
* Identify different amino acids with Chromatography
* CONH group formed by polymers of amino acids is called a peptide bond

In aqueous solutions, intramolecular reactions forms **zwitterions**

* Charges promote strong intermolecular forces of attraction intermolecularly
* All amino acids are neutral in solution
* Solution in water acts as buffer solutions



## Condensation polymerization

## Azo compounds

Azo / diazonium compounds are organic compounds with the group

### Preparation

1. Formation of nitrous acid: (ice cold NaNO2 + dilute HCl)
2. Formation of diazonium ion: (Ice-water temp, if not ArOH forms)
3. Coupling reaction: ,

## Amides RCON~

### Preparation

RCONH2:

RCONHR’:

RCONR’R”:

Further reaction:

# 20 – Organic synthesis

## Reactions

|  |  |  |  |
| --- | --- | --- | --- |
| R = CnH2n+1 *n = chain name - 1* | | |  |
| Alcohol  Aldehydes  Carboxylic acid  Ketone  Esters  Nitriles  Amines  Amides | R-OH  R-CHO  R-COOH  R-CO-R’  R-COO-R’  R-C≡N  R-NH2  R-ONH-R’ | -ol  -al  -oic acid  -one  R’-yl R-oate  -nitrile  -amine  N-R’-yl R-amide |

## Combustion analysis

## Methods of separation

Distilling Get (l) w/ b.p. << others

Steam distilling Get insoluble (l) from (aq)

Extracting Get wanted organic product

Washing Remove impurities from (s) / (l)

Drying Remove H2O

Filtration Remove (s) from (l)

## Extending carbon chains

### Grignard reactions

Preparing Grignard reagent: (dry ether, heat under reflux)

Reactions (followed by reaction with):

# Essentials

1. Previous O chem test & map
2. NMR, infrared, mass spectrum
3. Polymers
4. Equilibrium constant
5. Chirality, isomerism
6. Redox

# References

## Solubility rules

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| **Soluble** | | | | | **Exceptions** |
| Nitrate | Ammonium | Potassium | Sodium |  | - |
| Chlorine | Iodine |  |  |  | Pb2+ Hg22+ Ag+ (PHAg) |
| Sulphate |  |  |  |  | Pb2+ Hg22+ Ag+ Ca2+ Ba2+ Sr2+ (Castro bear) |
| **Insoluble** | | | | |  |
| Carbonate | Phosphate |  |  |  | Group 1, NH4+ |
| Hydroxide |  |  |  |  | Group 1, NH4+ Ca2+ Ba2+ Sr2+ (Castro bear) |