55.

$\longleftrightarrow$

a. 21 σ bonds; 4 π bonds (The electrons in the 3 π bonds in the ring are delocalized.) **b.** angles a, c, and g: $\approx 109.5°$; angles b, d, e, and f: $\approx 120°$; **c.** 6 sp^2 carbons; **d.** 4 sp^3 atoms; **e.** Yes, the π electrons in the ring are delocalized. The atoms in the ring are all sp^2 hybridized. This leaves a p orbital perpendicular to the plane of the ring from each atom. Overlap of all six of these p orbitals results in a π molecular orbital system where the electrons are delocalized above and below the plane of the ring (similar to benzene in Fig. 9.48 of the text). **57.** 267 kJ/mol; this amount of energy must be supplied to break the π bond.
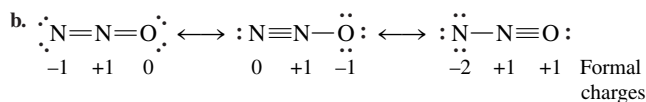
59. a.

Trigonal planar, nonpolar, 120°, sp^2

b. N_2F_2

can also be
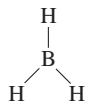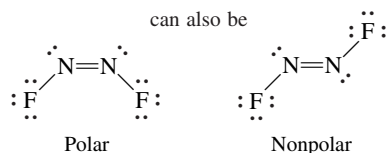
Polar   Nonpolar

V-shaped about both N atoms, $\approx 120°$, sp^2

These are distinctly different molecules.

c.
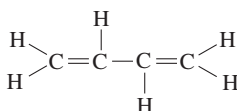
Trigonal planar about all C atoms, nonpolar, 120°, sp^2

d. ICl_3
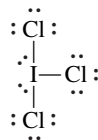
T-shaped, polar, $\approx 90°$, dsp^3

61. a. The NNO structure is correct. From the Lewis structures we would predict both NNO and NON to be linear, but NON would be nonpolar. NNO is polar.

b.

$$: \overset{..}{N} = N = \overset{..}{O} : \longleftrightarrow : N \equiv N - \overset{..}{\underset{..}{O}} : \longleftrightarrow : \overset{..}{\underset{..}{N}} - N \equiv O :$$

$$\begin{array}{ccccccc} -1 & +1 & 0 & & 0 & +1 & -1 & & -2 & +1 & +1 & \text{Formal} \\ & & & & & & & & & & & \text{charges} \end{array}$$

The central N is sp hybridized. We can probably ignore the third resonance structure on the basis of formal charge. **c.** sp hybrid orbitals on the center N overlap with atomic orbitals (or hybrid orbitals) on the other two atoms to form two σ bonds. The remaining p orbitals on the center N overlap with p orbitals on the other N to form two π bonds. **63.** N_2: $(\sigma_{2s})^2(\sigma_{2s}*)^2(\pi_{2p})^4(\sigma_{2p})^2$ in ground state, B.O. = 3, diamagnetic; 1st excited state: $(\sigma_{2s})^2(\sigma_{2s}*)^2(\pi_{2p})^4(\sigma_{2p})^1(\pi_{2p}*)^1$, B.O. = 2, paramagnetic (two unpaired electrons) **65.** F_2 $(\sigma_{2s})^2(\sigma_{2s}*)^2(\sigma_{2p})^2(\pi_{2p})^4(\pi_{2p}*)^4$; F_2 should have a lower ionization energy than F. The electron removed from F_2 is in a $\pi_{2p}*$ antibonding molecular orbital, which is higher in energy than the $2p$ atomic orbitals from which the electron in atomic fluorine is removed. Since the electron removed from F_2 is higher in energy than the electron removed from F, then it should be easier to remove an electron from F_2 than from F. **67.** π molecular orbital **69.** 6 sp^2; 6 sp^3; 0 sp; 25 σ; 4π **71. a.** No, some atoms are attached differently; **b.** Structure 1: All N = sp^3, all C = sp^2; structure 2: All C and N = sp^2; **c.** The first structure with the carbon–oxygen double bonds is slightly more stable.
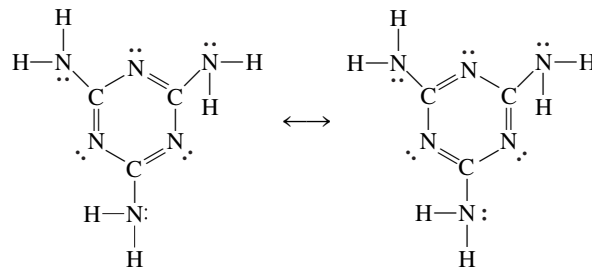
73. a. NCN^{2-}:

H_2NCN:
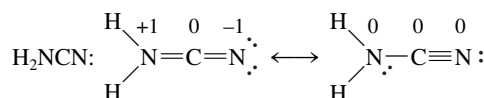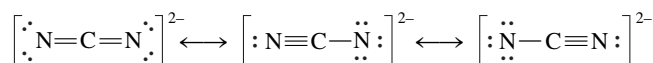
Favored by formal charge

Dicyandiamide:

Favored by formal charge
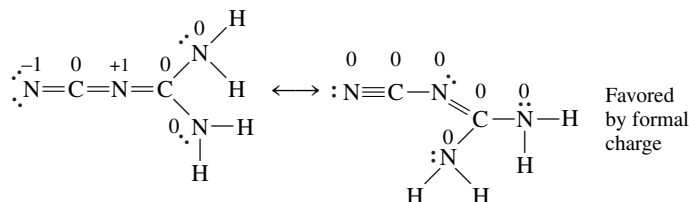
Melamine:

$\longleftrightarrow$

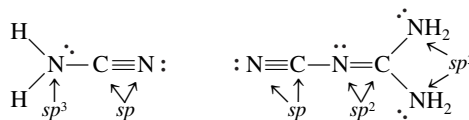**b.** NCN^{2-}: C is sp hybridized. Each resonance structure predicts a different hybridization for the N atoms. For the remaining compounds, we will predict hybrids for the favored resonance structures only.

Melamine: N in NH_2 groups are all sp^3 hybridized. Atoms in ring are all sp^2 hybridized; **c.** NCN^{2-}: 2 σ and 2 π bonds; H_2NCN: 4 σ and 2 π bonds; dicyandiamide: 9 σ and 3 π bonds; melamine: 15 σ and 3 π bonds; **d.** The π system forces the ring to be planar just as the benzene ring is planar.

e. The structure

$$: N \equiv C - \overset{..}{N} \overset{..}{\underset{\underset{H \quad H}{\overset{|}{N}}}{C} - \overset{..}{N} - H}$$

is the most important because it has three different CN bonds. This structure is also favored on the basis of formal charge.   **75. a.** 25-nm light has sufficient energy to ionize N and N_2 and to break the triple bond in N_2. Thus, N_2, N_2^+, N, and N^+ will all be present, assuming excess N_2. **b.** 85.33 nm $< \lambda \leq 127$ nm; **c.** The ionization energy of a substance is the energy it takes to completely remove an electron. N_2: $(\sigma_{2s})^2(\sigma_{2s}^*)^2(\pi_{2p})^4(\sigma_{2p})^2$; the electron removed from N_2 is in the σ_{2p} molecular orbital, which is lower in energy than the $2p$ atomic orbital from which the electron in atomic nitrogen is removed. Since the electron removed from N_2 is lower in energy than the electron removed in N, then the ionization energy of N_2 is greater than the ionization energy of N. **77.** Both reactions apparently involve only the breaking of the N—Cl bond. However, in the reaction ONCl $\rightarrow$ NO + Cl some energy is released in forming the stronger NO bond, lowering the value of ΔH. Therefore, the apparent N—Cl bond energy is artificially low for this reaction. The first reaction involves only the breaking of the N—Cl bond.   **79. a.** The CO bond is polar, with the negative end around the more electronegative oxygen atom. We would expect metal cations to be attracted to and to bond to the oxygen end of CO on the basis of electronegativity. **b.** The formal charge on C is -1, and the formal charge on O is $+1$. From formal charge, we would expect metal cations to bond to the carbon (with the negative formal charge.) **c.** In molecular orbital theory, only orbitals with proper symmetry overlap to form bonding orbitals. The metals that form bonds to CO are usually transition metals, all of which have outer electrons in the d orbitals. The only molecular orbitals of CO that have proper symmetry to overlap with d orbitals are the π_{2p}^* orbitals, whose shape is similar to that of the d orbitals (see Fig. 9.34). Since the antibonding molecular orbitals have more carbon character, one would expect the bond to form through carbon.   **81.** The species with the smallest ionization energy has the highest energy electron. O_2, N_2^{2-}, N_2^-, and O_2^+ all have at least one electron in the high-energy π_{2p}^* orbitals. Because N_2^{2-} has the highest ratio of electrons to protons, the π_{2p}^* electrons are least attracted to the nuclei and easiest to remove, translating into the smallest ionization energy.   **83. a.** Li_2 bond order = 1; B_2 bond order = 1; **b.** 4 electrons must be removed; **c.** 4.5×10^5 kJ   **85.** T-shaped and dsp^3 hybridized

Chapter 10

13. Atoms have an approximately spherical shape. It is impossible to pack spheres together without some empty space between the spheres.   **15.** Evaporation takes place when some molecules at the surface of a liquid have enough energy to break the intermolecular forces holding them in the liquid phase. When a liquid evaporates, the molecules that escape have high kinetic energies. The average kinetic energy of the remaining molecules is lower; thus the temperature of the liquid is lower.   **17.** An alloy is a substance that contains a mixture of elements and has metallic properties. In a substitutional alloy, some of the host metal atoms are replaced by other metal atoms of similar size (e.g., in brass, pewter, plumber's solder). An interstitial alloy is formed when some of the interstices (holes) in the closest packed metal structure are occupied by smaller atoms (e.g., in carbon steels).   **19. a.** As intermolecular forces increase, the rate of evaporation decreases. **b.** increase T: increase rate; **c.** increase surface area: increase rate   **21.** Sublimation will occur, allowing water to escape as $H_2O(g)$.   **23.** The strength of intermolecular forces determines relative boiling points. The types of intermolecular forces for covalent compounds are London dispersion forces, dipole forces, and hydrogen bonding. Because the three compounds are assumed to have similar molar mass and shape, the strength of the London dispersion forces will be about equal between the three compounds. One of the compounds will be nonpolar so it only has London dispersion forces. The other two compounds will be polar so they have additional dipole forces and will boil at a higher temperature than the

nonpolar compound. One of the polar compounds will have an H covalently bonded to either N, O, or F. This gives rise to the strongest type of covalent intermolecular force, hydrogen bonding. This compound exhibiting hydrogen bonding will have the highest boiling point while the polar compound with no hydrogen bonding will boil at an intermediate temperature.   **25. a.** Both CO_2 and H_2O are molecular solids. Both have an ordered array of the individual molecules, with the molecular units occupying the lattice points. A difference within each solid lattice is the strength of the intermolecular forces. CO_2 is nonpolar and only exhibits London dispersion forces. H_2O exhibits the relatively strong hydrogen bonding interactions. The difference in strength is evidenced by the solid phase change that occurs at 1 atm. CO_2 sublimes at a relatively low temperature of $-78°C$. In sublimation, all of the intermolecular forces are broken. However, H_2O doesn't have a solid phase change until 0°C, and in this phase change from ice to water, only a fraction of the intermolecular forces are broken. The higher temperature and the fact that only a portion of the intermolecular forces are broken are attributed to the strength of the intermolecular forces in H_2O as compared to CO_2. Related to the intermolecular forces is the relative densities of the solid and liquid phases for these two compounds. $CO_2(s)$ is denser than $CO_2(l)$ while $H_2O(s)$ is less dense than $H_2O(l)$. For $CO_2(s)$ and for most solids, the molecules pack together as close as possible, which is why solids are usually more dense than the liquid phase. For H_2O, each molecule has two lone pairs and two bonded hydrogen atoms. Because of the equal number of lone pairs and O—H bonds, each H_2O molecule can form two hydrogen bonding interactions to other H_2O molecules. To keep this symmetric arrangement (which maximizes the hydrogen bonding interactions), the $H_2O(s)$ molecules occupy positions that create empty space in the lattice. This translates into smaller density for $H_2O(s)$ (less mass per unit volume).   **b.** Both NaCl and CsCl are ionic compounds with the anions at the lattice points of the unit cell and the cations occupying the empty spaces created by the anions (called holes). In NaCl, the Cl^- anions occupy the lattice points of a face-centered unit cell with the Na^+ cations occupying the octahedral holes. Octahedral holes are the empty spaces created by six Cl^- ions. CsCl has the Cl^- ions at the lattice points of a simple cubic unit cell with the Cs^+ cations occupying the middle of the cube.   **27.** In the ln P_{vap} versus $1/T$ plot, the slope of the straight line is equal to $-\Delta H_{vap}/R$. Because ΔH_{vap} is always positive, the slope of the line will always be negative.   **29. a.** LD (London dispersion); **b.** dipole, LD; **c.** hydrogen bonding, LD; **d.** ionic; **e.** LD; **f.** dipole, LD; **g.** ionic   **31. a.** OCS; **b.** SeO_2; **c.** $H_2NCH_2CH_2NH_2$; **d.** H_2CO; **e.** CH_3OH **33. a.** Neopentane is more compact than n-pentane. There is less surface area contact among neopentane molecules. This leads to weaker London dispersion forces and a lower boiling point. **b.** HF is capable of hydrogen bonding; HCl is not. **c.** LiCl is ionic, and HCl is a molecular solid with only dipole forces and London dispersion forces. Ionic forces are much stronger than the forces for molecular solids. **d.** n-Hexane is a larger molecule, so it has stronger London dispersion forces.   **35. a.** HBr has dipole forces in addition to LD forces; **b.** NaCl, stronger ionic forces; **c.** I_2, larger molecule so stronger LD forces; **d.** N_2, smallest nonpolar compound present, has weakest LD forces; **e.** CH_4, smallest nonpolar compound present, has weakest LD forces; **f.** HF, can form relatively strong hydrogen bonding interactions, unlike the other compounds; **g.** $CH_3CH_2CH_2OH$, unlike others, has relatively strong hydrogen bonding.   **37.** H_2O is attracted to glass while Hg is not.   **39.** The structure of H_2O_2 produces greater hydrogen bonding than water. Long chains of hydrogen bonded H_2O_2 molecules then get tangled together.   **41.** 313 pm   **43.** 0.704 Å   **45.** 1.54 g/cm^3 **47.** 174 pm; 11.6 g/cm^3 **49.** edge, 328 pm; radius, 142 pm   **51.** face-centered cubic unit cell   **53.** For a cubic closest packed structure, 74.06% of the volume of each unit cell is occupied by atoms; in a simple cubic unit cell structure, 52.36% is occupied. The cubic (and hexagonal) closest packed structures provide the most efficient means for packing atoms.   **55.** Doping silicon with phosphorus produces an n-type semiconductor. The phosphorus adds electrons at energies near the conduction band of silicon. Electrons do not need as much energy to move from filled to unfilled energy levels so conduction increases. Doping silicon with gallium produces a p-type semiconductor. Because gallium has fewer valence electrons than silicon, holes (unfilled energy levels) at energies in the previously filled

molecular orbitals are created, which induces greater electron movement (greater conductivity). **57.** p-type **59.** 5.0×10^2 nm **61.** NaCl: $4Na^+$, $4Cl^-$; CsCl: $1Cs^+$, $1Cl^-$; ZnS: $4Zn^{2+}$, $4S^{2-}$; TiO$_2$: $2Ti^{4+}$, $4O^{2-}$ **63.** CoF$_2$ **65.** ZnAl$_2$S$_4$ **67.** MF$_2$ **69.** $r_{O^{2-}} = 1.49 \times 10^{-8}$ cm; $r_{Mg^{2+}} = 6.15 \times 10^{-9}$ cm **71. a.** CO$_2$: molecular; **b.** SiO$_2$: covalent network; **c.** Si: atomic, covalent network; **d.** CH$_4$: molecular; **e.** Ru: atomic, metallic; **f.** I$_2$: molecular; **g.** KBr: ionic; **h.** H$_2$O: molecular; **i.** NaOH: ionic; **j.** U: atomic, metallic; **k.** CaCO$_3$: ionic; **l.** PH$_3$: molecular **73. a.** The unit cell consists of Ni at the cube corners and Ti at the body center or Ti at the cube corners and Ni at the body center. **b.** NiTi; **c.** Both have a coordination number of 8. **75.** CaTiO$_3$; six oxygens around each Ti **77. a.** YBa$_2$Cu$_3$O$_9$; **b.** The structure of this superconductor material is based on the second perovskite structure. The YBa$_2$Cu$_3$O$_9$ structure is three of these cubic perovskite unit cells stacked on top of each other. The oxygens are in the same places, Cu takes the place of Ti, two Ca are replaced by two Ba, and one Ca is replaced by Y. **c.** YBa$_2$Cu$_3$O$_7$ **79.** Li, 158 kJ/mol; Mg, 139 kJ/mol. Bonding is stronger in Li. **81.** 89°C **83.** 77°C
85.
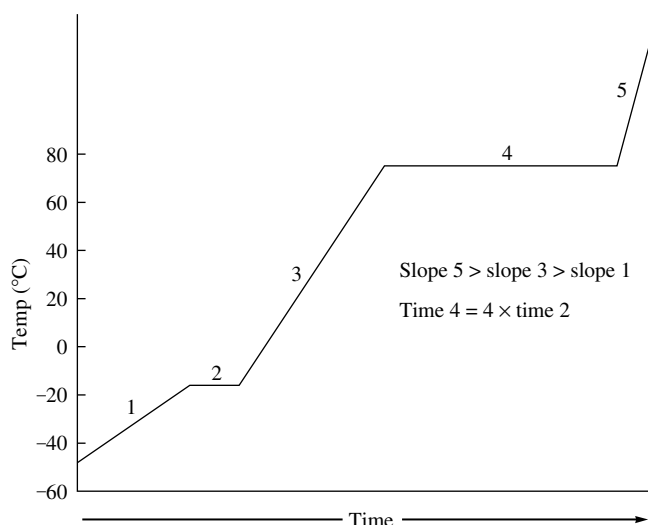
87. 1680 kJ **89.** 1490 g **91.** A: solid; B: liquid; C: vapor; D: solid + vapor; E: solid + liquid + vapor (triple point); F: liquid + vapor; G: liquid + vapor (critical point); H: vapor; the first dashed line (at the lower temperature) is the normal melting point, and the second dashed line is the normal boiling point. The solid phase is denser. **93. a.** two; **b.** higher pressure triple point: graphite, diamond, and liquid; lower pressure triple point: graphite, liquid and vapor; **c.** It is converted to diamond (the more dense solid form); **d.** Diamond is more dense, which is why graphite can be converted to diamond by applying pressure. **95.** Because the density of the liquid phase is greater than the density of the solid phase, the slope of the solid–liquid boundary line is negative (as in H$_2$O). With a negative slope, the melting points increase with a decrease in pressure so the normal melting point of X should be greater than 225°C. **97.** Chalk is composed of the ionic compound calcium carbonate (CaCO$_3$). The electrostatic forces in ionic compounds are much stronger than the intermolecular forces in covalent compounds. Therefore, CaCO$_3$ should have a much higher boiling point than the covalent compounds found in motor oil and in H$_2$O. Motor oil is composed of nonpolar C—C and C—H bonds. The intermolecular forces in motor oil are therefore London dispersion forces. We generally consider these forces to be weak. However, with compounds that have large molar masses, these London dispersion forces add up significantly and can overtake the relatively strong hydrogen-bonding interactions in water. **99.** A: CH$_4$; B: SiH$_4$ C: NH$_3$ **101.** If TiO$_2$ conducts electricity as a liquid, then it would be ionic. **103.** B$_2$H$_6$, molecular; SiO$_2$, network; CsI, ionic; W, metallic **105.** 4.65 kg/h **107.** $\Delta E = 27.86$ kJ/mol;

$\Delta H = 30.79$ kJ/mol **109.** 46.7 kJ/mol; 90% **111.** The solids with high melting points (NaCl, MgCl$_2$, NaF, MgF$_2$, AlF$_3$) are all ionic solids. SiCl$_4$, SiF$_4$, Cl$_2$, F$_2$, PF$_5$, and SF$_6$ are nonpolar covalent molecules with LD forces. PCl$_3$ and SCl$_2$ are polar molecules with LD and dipole forces. In these 8 molecular substances the intermolecular forces are weak and the melting points low. AlCl$_3$ is intermediate. The melting point indicates there are stronger forces present than in the nonmetal halides, but not as strong as for an ionic solid. AlCl$_3$ illustrates a gradual transition from ionic to covalent bonding; from an ionic solid to discrete molecules. **113.** TiO$_{1.182}$ or Ti$_{0.8462}$O; 63.7% Ti^{2+}, 36.3% Ti^{3+} **115.** 6.58 g/cm^3

117.
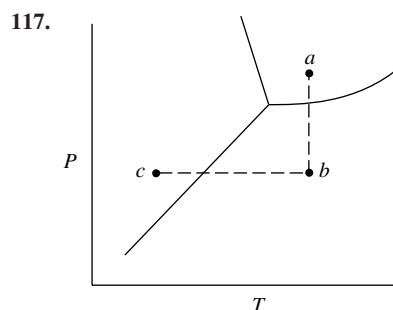
As P is lowered, we go from a to b on the phase diagram. The water boils. The evaporation of the water is endothermic and the water is cooled ($b \rightarrow c$), forming some ice. If the pump is left on, the ice will sublime until none is left. This is the basis of freeze drying.

119. The volume of the hole is $\frac{4}{3}\pi\left[\left(\frac{2\sqrt{3}-2}{2}\right)r\right]^3$ **121.** CdS; n-type

123. 2.53 torr; 6.38×10^{22} atoms

Chapter 11

9. 9.74 M **11.** 4.5 M **13.** As the temperature increases, the gas molecules will have a greater average kinetic energy. A greater fraction of the gas molecules in solution will have kinetic energy greater than the attractive forces between the gas molecules and the solvent molecules. More gas molecules will escape to the vapor phase, and the solubility of the gas will decrease. **15.** The levels of the liquids in each beaker will become constant when the concentration of solute is the same in both beakers. Because the solute is less volatile, the beaker on the right will have a larger volume when the concentrations become equal. Water will initially condense in this beaker in a larger amount than solute is evaporating, while the net change occurring initially in the other beaker is for water to evaporate in a larger amount than solute is condensing. Eventually the rate that solute and H$_2$O leave and return to each beaker will become equal when the concentrations become equal. **17.** No. For an ideal solution, $\Delta H_{soln} = 0$ **19.** Normality is the number of equivalents per liter of solution. For an acid or a base, an equivalent is the mass of acid or base that can furnish 1 mol of protons (if an acid) or accept 1 mol of protons (if a base). A proton is an H$^+$ ion. Molarity is defined as the moles of solute per liter of solution. When the number of equivalents equals the number of moles of solute, then normality = molarity. This is true for acids which only have one acidic proton in them and for bases that accept only one proton per formula unit. Examples of acids where equivalents = moles solute are HCl, HNO$_3$, HF, and HC$_2$H$_3$O$_2$. Examples of bases where equivalents = moles solute are NaOH, KOH, and NH$_3$. When equivalents ≠ moles solute, then normality ≠ molarity. This is true for acids that donate more than one proton (H$_2$SO$_4$, H$_3$PO$_4$, H$_2$CO$_3$, etc.) and for bases that react with more than one proton per formula unit [Ca(OH)$_2$, Ba(OH)$_2$, Sr(OH)$_2$, etc.]. **21.** Only statement b is true. A substance freezes when the vapor pressure of the liquid and solid phases are the same. When a solute is added to water, the vapor pressure of the solution at 0°C is less than the vapor pressure of the solid; the net result is for any ice

present to convert to liquid in order to try to equalize the vapor pressures (which never can occur at 0°C). A lower temperature is needed to equalize the vapor pressure of water and ice, hence the freezing point is depressed. For statement a, the vapor pressure of a solution is directly related to the mole fraction of solvent (not solute) by Raoult's law. For statement c, colligative properties depend on the number of solute particles present and not on the identity of the solute. For statement d, the boiling point of water is increased because the sugar solute decreases the vapor pressure of the water; a higher temperature is required for the vapor pressure of the solution to equal the external pressure so boiling can occur. **23.** Isotonic solutions are those which have identical osmotic pressures. Crenation and hemolysis refer to a phenomena that occurs when red blood cells are bathed in solutions having a mismatch in osmotic pressure between the inside and the outside of the cell. When red blood cells are in a solution having a higher osmotic pressure than that of the cells, the cells shrivel as there is a net transfer of water out of the cells. This is called crenation. Hemolysis occurs when the red blood cells are bathed in a solution having lower osmotic pressure than that inside the cell. Here, the cells rupture as there is a net transfer of water to inside the red blood cells. **25.** 1.06 g/mL; 0.0180 mole fraction H_3PO_4, 0.9820 mole fraction H_2O; 0.981 mol/L; 1.02 mol/kg **27.** HCl: 12 M, 17 m, 0.23; HNO_3: 16 M, 37 m, 0.39; H_2SO_4: 18 M, 200 m, 0.76; $HC_2H_3O_2$: 17 M, 2000 m, 0.96; NH_3: 15 M, 23 m, 0.29 **29.** 35%; 0.39; 7.3 m; 3.1 M **31.** 23.9%; 1.6 m, 0.028, 4.11 N **33.** $NaI(s) \rightarrow Na^+(aq) + I^-(aq)$ $\Delta H_{soln} = -8$ kJ/mol **35.** The attraction of water molecules for Al^{3+} and OH^- cannot overcome the larger lattice energy of $Al(OH)_3$. **37. a.** CCl_4; **b.** H_2O; **c.** H_2O; **d.** CCl_4; **e.** H_2O; **f.** H_2O; **g.** CCl_4; **39.** Ability to form hydrogen bonding interactions, ability to break up into ions, and polarity are some factors affecting solute solubility. **a.** CH_3CH_2OH; **b.** $CHCl_3$; **c.** CH_3CH_2OH **41.** As the length of the hydrocarbon chain increases, the solubility decreases because the nonpolar hydrocarbon chain interacts poorly with the polar water molecules. **43.** 1.04×10^{-3} mol/L · atm; 1.14×10^{-3} mol/L **45.** 50.0 torr **47.** 3.0×10^2 g/mol **49. a.** 290 torr; **b.** 0.69 **51.** $\chi_{methanol} = \chi_{propanol} = 0.500$ **53.** solution c **55.** $P_{ideal} = 188.6$ torr; $\chi_{acetone} = 0.512$, $\chi_{methanol} = 0.488$; the actual vapor pressure of the solution is smaller than the ideal vapor pressure, so this solution exhibits a negative deviation from Raoult's law. This occurs when solute–solvent attractions are stronger than for the pure substances. **57.** 101.5°C **59.** 14.8 g $C_3H_8O_3$ **61.** $T_f = -29.9$°C, $T_b = 108.2$°C **63.** 776 g/mol **65. a.** $\Delta T = 2.0 \times 10^{-5}$°C, $\pi = 0.20$ torr; **b.** Osmotic pressure is better for determining the molar mass of large molecules. A temperature change of 10^{-5}°C is very difficult to measure. A change in height of a column of mercury by 0.2 mm is not as hard to measure precisely. **67.** 0.327 M **69. a.** 0.010 m Na_3PO_4 and 0.020 m KCl; **b.** 0.020 m HF; **c.** 0.020 m $CaBr_2$ **71. a.** $T_f = -0.28$°C; $T_b = 100.077$°C; **b.** $T_f = -0.37$°C; $T_b = 100.10$°C **73.** 2.63 (0.0225 m), 2.60 (0.0910 m), 2.57 (0.278 m); $i_{average} = 2.60$ **75. a.** yes; **b.** no **77. a.** 26.6 kJ/mol; **b.** -657 kJ/mol **79. a.** Water boils when the vapor pressure equals the pressure above the water. In an open pan, $P_{atm} \approx 1.0$ atm. In a pressure cooker, $P_{inside} > 1.0$ atm and water boils at a higher temperature. The higher the cooking temperature, the faster the cooking time. **b.** Salt dissolves in water, forming a solution with a melting point lower than that of pure water ($\Delta T_f = K_f m$). This happens in water on the surface of ice. If it is not too cold, the ice melts. This process won't occur if the ambient temperature is lower than the depressed freezing point of the salt solution. **c.** When water freezes from a solution, if freezes as pure water, leaving behind a more concentrated salt solution. **d.** On the CO_2 phase diagram, the triple point is above 1 atm and $CO_2(g)$ is the stable phase at 1 atm and room temperature. $CO_2(l)$ can't exist at normal atmospheric pressures, which explains why dry ice sublimes rather than boils. In a fire extinguisher, $P > 1$ atm and $CO_2(l)$ can exist. When CO_2 is released from the fire extinguisher, $CO_2(g)$ forms as predicted from the phase diagram. **e.** Adding a solute to a solvent increases the boiling point and decreases the freezing point of the solvent. Thus, the solvent is a liquid over a wider range of temperatures when a solute is dissolved. **81.** 0.600 **83.** $C_2H_4O_3$; 151 g/mol (exp.); 152.10 g/mol (calc.); $C_4H_8O_6$ **85.** 1.97% NaCl **87. a.** 100.77°C; **b.** 23.1 mm Hg; **c.** Assume an ideal solution; assume no ions form (i = 1).

89. 30.% A: $\chi_A = \dfrac{0.30y}{0.70x - 0.30y}$, $\chi_B = 1 - \chi_A$;

50.% A: $\chi_A = \dfrac{y}{x - y}$, $\chi_B = 1 - \dfrac{y}{x - y}$;

80.% A: $\chi_A = \dfrac{0.80y}{0.20x - 0.80y}$, $\chi_B = 1 - \chi_A$;

30.% A: $\chi_A{}^V = \dfrac{0.30x}{0.30x + 0.70y}$, $\chi_B{}^V = 1 - \dfrac{0.30x}{0.30x + 0.70y}$;

50.% A: $\chi_A{}^V = \dfrac{x}{x + y}$, $\chi_B{}^V = 1 - \chi_A{}^V$;

80.% A: $\chi_A{}^V = \dfrac{0.80x}{0.80x + 0.20y}$, $\chi_B{}^V = 1 - \chi_A{}^V$

91. 72.7% sucrose and 27.3% NaCl by mass; 0.2 **93.** 0.050 **95.** 44% naphthalene, 56% anthracene **97.** -0.20°C, 100.056°C **99. a.** 46 L; **b.** No; A reverse osmosis system that applies 8.0 atm can purify only water with solute concentrations less than 0.32 mol/L. Salt water has a solute concentration of 2(0.60 M) = 1.2 M ions. The solute concentration of salt water is much too high for this reverse osmosis unit to work. **101.** i = 3.00; $CdCl_2$

Chapter 12

9. In a unimolecular reaction, a single reactant molecule decomposes to products. In a bimolecular reaction, two molecules collide to give products. The probability of the simultaneous collision of three molecules with enough energy and orientation is very small, making termolecular steps very unlikely. **11.** All of these choices would affect the rate of the reaction, but only b and c affect the rate by affecting the value of the rate constant k. The value of the rate constant is dependent on temperature. It also depends on the activation energy. A catalyst will change the value of k because the activation energy changes. Increasing the concentration (partial pressure) of either H_2 or NO does not affect the value of k, but it does increase the rate of the reaction because both concentrations appear in the rate law. **13.** The average rate decreases with time because the reverse reaction occurs more frequently as the concentration of products increase. Initially, with no products present, the rate of the forward reaction is at its fastest; but as time goes on, the rate gets slower and slower since products are converting back into reactants. The instantaneous rate will also decrease with time. The only rate that is constant is the initial rate. This is the instantaneous rate taken at $t \approx 0$. At this time, the amount of products is insignificant and the rate of the reaction only depends on the rate of the forward reaction. **15.** When the rate doubles as the concentration quadruples, the order is 1/2. For a reactant that has an order of -1, the rate will decrease by a factor of 1/2 when the concentrations are doubled. **17.** Two reasons are: **a.** the collision must involve enough energy to produce the reaction; i.e., the collision energy must equal or exceed the activation energy. **b.** the relative orientation of the reactants must allow formation of any new bonds necessary to produce products. **19.** P_4: 6.0×10^{-4} mol/L · s; H_2: 3.6×10^{-3} mol/L · s **21. a.** average rate of decomposition of $H_2O_2 = 2.31 \times 10^{-5}$ mol/L · s, rate of production of $O_2 = 1.16 \times 10^{-5}$ mol/L · s; **b.** average rate of decomposition of $H_2O_2 = 1.16 \times 10^{-5}$ mol/L · s, rate of production of $O_2 = 5.80 \times 10^{-6}$ mol/L · s **23. a.** mol/L · s; **b.** mol/L · s; **c.** s^{-1}; **d.** L/mol · s; **e.** L^2/mol^2 · s **25. a.** rate = $k[NO]^2[Cl_2]$; **b.** 1.8×10^2 L^2/mol^2 · min **27. a.** rate = $k[NOCl]^2$; **b.** 6.6×10^{-29} cm^3/molecules · s; **c.** 4.0×10^{-8} L/mol · s **29. a.** first order in Hb and first order in CO; **b.** rate = $k[Hb][CO]$; **c.** 0.280 L/μmol · s; **d.** 2.26 μmol/L · s **31.** rate = $k[H_2O_2]$; $\ln[H_2O_2] = -kt + \ln[H_2O_2]_0$; $k = 8.3 \times 10^{-4}$ s^{-1}; 0.037 M **33.** rate = $k[NO_2]^2$; $\dfrac{1}{[NO_2]} = kt + \dfrac{1}{[NO_2]_0}$; $k = 2.08 \times 10^{-4}$ L/mol · s; 0.131 M **35. a.** rate = k; $[C_2H_5OH] = -kt + [C_2H_5OH]_0$; because slope = $-k$, then $k = 4.00 \times 10^{-5}$ mol/L · s; **b.** 156 s; **c.** 313 s **37.** rate = $k[C_4H_6]^2$; $\dfrac{1}{[C_4H_6]} = kt + $

$\dfrac{1}{[C_4H_6]_0}$; $k = 1.4 \times 10^{-2}$ L/mol · s **39.** second order; 0.1 M **41. a.** $[A] = -kt + [A]_0$; **b.** 1.0×10^{-2} s; **c.** 2.5×10^{-4} M **43. a.** 160. s $= t_{1/2}$ for both the first and second half-life; **b.** 532 s **45.** 12.5 s **47. a.** 1.1×10^{-2} M; **b.** 0.025 M **49. a.** rate $= k[CH_3NC]$; **b.** rate $= k[O_3][NO]$; **c.** rate $= k[O_3]$; **d.** rate $= k[O_3][O]$ **51.** Rate $= k[C_4H_9Br]$; $C_4H_9Br + 2H_2O \rightarrow C_4H_9OH + Br^- + H_3O^+$; the intermediates are $C_4H_9^+$ and $C_4H_9OH_2^+$.

53.
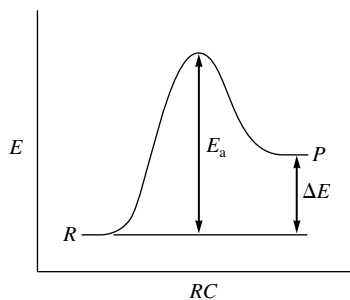
55. 341 kJ/mol **57.** The graph of $\ln k$ versus $1/T$ is linear with slope $= -E_a/R = -1.2 \times 10^4$ K; $E_a = 1.0 \times 10^2$ kJ/mol **59.** 9.5×10^{-5} L/mol · s **61.** 51°C **63.** $H_3O^+(aq) + OH^-(aq) \rightarrow 2H_2O(l)$ should have the faster rate. H_3O^+ and OH^- will be electrostatically attracted to each other; Ce^{4+} and Hg_2^{2+} will repel each other (so E_a is much larger). **65. a.** NO; **b.** NO_2; **c.** 2.3 **67.** CH_2D-CH_2D should be the product. If the mechanism is possible, then the reaction must be $C_2H_4 + D_2 \rightarrow CH_2DCH_2D$. If we got this product, then we would conclude that this is a possible mechanism. If we got some other product, e.g., CH_3CHD_2, then we would conclude that the mechanism is wrong. Even though this mechanism correctly predicts the products of the reaction, we cannot say conclusively that this is the correct mechanism; we might be able to conceive of other mechanisms that would give the same product as our proposed one. **69.** 215°C **71.** 5.68×10^{18} molecules/cm³ · s **73.** 1.0×10^2 kJ/mol **75.** At high [S], the enzyme is completely saturated with substrate. Once the enzyme is completely saturated, the rate of decomposition of ES can no longer increase, and the overall rate remains constant. **77. a.** 115 L³/mol³ · s; **b.** 87.0 s; **c.** $[A] = 1.27 \times 10^{-5}$ M, $[B] = 1.00$ M **79.** rate $= \dfrac{k[I^-][OCl^-]}{[OH^-]}$; $k = 6.0 \times 10^1$ s^{-1} **81. a.** first order with respect to both reactants; **b.** rate $= k[NO][O_3]$; **c.** $k' = 1.8$ s^{-1}; $k'' = 3.6$ s^{-1}; **d.** $k = 1.8 \times 10^{-14}$ cm³/molecules · s **83. a.** 25 kJ/mol; **b.** 12 s;

c.

T	Interval	$54 - 2$(Intervals)
21.0°C	16.3 s	21°C
27.8°C	13.0 s	28°C
30.0°C	12 s	30.°C

This rule of thumb gives excellent agreement to two significant figures.

85. a. $[B] \gg [A]$ so that $[B]$ can be considered constant over the experiments. (This gives us a pseudo-order rate law equation.) **b.** Rate $= k[A]^2[B]$, $k = 0.050$ L²/mol² · s **87.** Rate $= k[A][B]^2$, $k = 1.4 \times 10^{-2}$ L²/mol² · s **89.** 2.20×10^{-5} s^{-1}; 5.99×10^{21} molecules **91.** 1.3×10^{-5} s^{-1}; 112 torr

Chapter 13

9. No, equilibrium is a dynamic process. Both the forward and reverse reactions are occurring at equilibrium, just at equal rates. Thus the forward and reverse reactions will distribute ^{14}C atoms between CO and CO_2. **11.** 4 molecules H_2O, 2 molecules CO, 4 molecules H_2, and 4 molecules CO_2 are present at equilibrium. **13.** K and K_p are equilibrium constants as determined by the law of mass action. For K, the units used for concentrations are mol/L, for K_p, partial pressures in units of atm are used (generally). Q is called the reaction quotient. Q has the exact same form as K or K_p, but instead of equilibrium concentrations, initial concentrations are used to calculate the Q value. Q is of use when it is compared to the K value. When $Q = K$ (or when $Q_p = K_p$), the

reaction is at equilibrium. When $Q \neq K$, the reaction is not at equilibrium and one can determine what has to be the net charge for the system to get to equilibrium. **15.** We always try to make good assumptions that simplify the math. In some problems, we can set up the problem so that the net change, x, that must occur to reach equilibrium is a small number. This comes in handy when you have expressions like $0.12 - x$ or $0.727 + 2x$. Since x is small, we assume that it makes little difference when subtracted from or added to some relatively big number. When this is true, $0.12 - x \approx 0.12$ and $0.727 + 2x \approx 0.727$. If the assumption holds by the 5% rule, then the assumption is assumed valid. The 5% rule refers to x (or $2x$ or $3x$, etc.) that was assumed small compared to some number. If x (or $2x$ or $3x$, etc.) is less than 5% of the number the assumption was made against, then the assumption will be assumed valid. If the 5% rule fails to work, one can generally use a math procedure called the method of successive approximations to solve the quadratic or cubic equation.

17. a. $K = \dfrac{[NO]^2}{[N_2][O_2]}$; **b.** $K = \dfrac{[NO_2]^2}{[N_2O_4]}$; **c.** $K = \dfrac{[SiCl_4][H_2]^2}{[SiH_4][Cl_2]^2}$;

d. $K = \dfrac{[PCl_3]^2[Br_2]^3}{[PBr_3]^2[Cl_2]^3}$ **19. a.** 0.11; **b.** 77; **c.** 8.8; **d.** 4.6×10^{-4}

21. 4.0×10^6 **23.** 1.7×10^{-5} **25.** 6.3×10^{-13} **27.** 1.1×10^3

29. a. $K = \dfrac{[H_2O]}{[NH_3]^2[CO_2]}$, $K_p = \dfrac{P_{H_2O}}{P_{NH_3}^2 \times P_{CO_2}}$;

b. $K = [N_2][Br_2]^3$, $K_p = P_{N_2} \times P_{Br_2}^3$; **c.** $K = [O_2]^3$, $K_p = P_{O_2}^3$;

d. $K = \dfrac{[H_2O]}{[H_2]}$, $K_p = \dfrac{P_{H_2O}}{P_{H_2}}$ **31.** 8.0×10^9 **33. a.** $Q > K$, so reaction shifts

left to reach equilibrium. **b.** $Q = K$, so reaction is at equilibrium. **c.** $Q < K$, so reaction shifts right to reach equilibrium. **35. a.** decrease; **b.** no change; **c.** no change; **d.** increase **37.** 8.0×10^{-2} M **39.** 3.4 **41.** 0.056 **43.** $[N_2]_0 = 10.0$ M, $[H_2]_0 = 11.0$ M **45.** $[SO_3] = [NO] = 1.06$ M; $[SO_2] = [NO_2] = 0.54$ M **47.** 7.8×10^{-2} atm **49.** $P_{SO_2} = 0.38$ atm; $P_{O_2} = 0.44$ atm; $P_{SO_3} = 0.12$ atm **51. a.** $[NO] = 0.032$ M, $[Cl_2] = 0.016$ M, $[NOCl] = 1.0$ M; **b.** $[NO] = [NOCl] = 1.0$ M, $[Cl_2] = 1.6 \times 10^{-5}$ M; **c.** $[NO] = 8.0 \times 10^{-3}$ M, $[Cl_2] = 1.0$ M, $[NOCl] = 2.0$ M **53.** $[CO_2] = 0.39$ M, $[CO] = 8.6 \times 10^{-3}$ M, $[O_2] = 4.3 \times 10^{-3}$ M **55.** 0.27 atm **57. a.** no effect; **b.** shifts left; **c.** shifts right **59. a.** right **b.** right; **c.** no effect; **d.** left; **e.** no effect **61. a.** left; **b.** right; **c.** left; **d.** no effect; **e.** no effect; **f.** right **63.** increase **65.** 2.6×10^{81} **67. a.** 0.379 atm; **b.** 0.786 **69. a.** 1.16 atm; **b.** 0.10 atm; **c.** 2.22 atm; **d.** 91.4% **71.** $[H_2] = [F_2] = 0.0251$ M; $[HF] = 0.450$ M **73.** Added OH^- reacts with H^+ to produce H_2O. As H^+ is removed, the reaction shifts right to produce more H^+ and CrO_4^{2-}. Because more CrO_4^{2-} is produced, the solution turns yellow. **75.** 9.0×10^{-3} M **77.** $P_{PCl_3} = P_{Cl_2} = 0.2230$ atm, $P_{PCl_5} = 0.0259$ atm; $K_p = 1.92$ **79.** $[NOCl] = 2.0$ M, $[NO] = 0.050$ M, $[Cl_2] = 0.025$ M **81.** 2.1×10^{-3} atm **83.** $P_{NO_2} = 0.704$ atm, $P_{N_2O_4} = 0.12$ atm **85.** 0.63 **87.** 0.240 atm **89.** 9.17×10^{-3} **91.** 192 g; 0.25 atm **93.** $C_{10}H_8$; 0.0919%

Chapter 14

17. 10.78 (4 significant figures); 6.78 (3 significant figures); 0.78 (2 significant figures); A pH value is a logarithm. The numbers to the left of the decimal place identify the power of 10 to which $[H^+]$ is expressed in scientific notation—for example, 10^{-11}, 10^{-7}, 10^{-1}. The number of decimal places in a pH value identifies the number of significant figures in $[H^+]$. In all three pH values, the $[H^+]$ should be expressed only to two significant figures since these pH values have only two decimal places. **19. a.** These would be 0.10 M solutions of strong acids like HCl, HBr, HI, HNO_3, H_2SO_4 or $HClO_4$. **b.** These are salts of the conjugate acids of the bases in Table 14.3. These conjugate acids are all weak acids. Three examples would be 0.10 M solutions of NH_4Cl, $CH_3NH_3NO_3$, and $C_2H_5NH_3Br$. Note that the anions used to form these salts are conjugate bases of strong acids; this is because they have no acidic or basic properties in water (with the exception of HSO_4^-, which has weak acid properties). **c.** These would be 0.10 M solutions of strong bases like LiOH,

NaOH, KOH, RbOH, CsOH, Ca(OH)$_2$, Sr(OH)$_2$ and Ba(OH)$_2$. **d.** These are salts of the conjugate bases of the neutrally charged weak acids in Table 14.2. The conjugate bases of weak acids are weak bases themselves. Three examples would be 0.10 M solutions of NaClO$_2$, KC$_2$H$_3$O$_2$, and CaF$_2$. The cations used to form these salts are Li$^+$, Na$^+$, K$^+$, Rb$^+$, Cs$^+$, Ca^{2+}, Sr^{2+}, and Ba^{2+} since these cations have no acidic or basic properties in water. Notice that these are the cations in the list of the strong bases listed in part c that you should memorize. **e.** There are two ways to make a neutral salt. The easiest way is to combine a conjugate base of a strong acid (except for HSO$_4^-$) with one of the cations from the strong bases. These ions have no acidic/basic properties in water so salts of these ions are neutral. Three examples would be 0.10 M solutions of NaCl, KNO$_3$, and SrI$_2$. Another type of strong electrolyte that can produce neutral solutions are salts that contain an ion with weak acid properties combined with an ion of opposite charge having weak base properties. If the K_a for the weak acid ion is equal to the K_b for the weak base ion, then the salt will produce a neutral solution. The most common example of this type of salt is ammonium acetate, NH$_4$C$_2$H$_3$O$_2$. For this salt, K_a for NH$_4^+$ = K_b for C$_2$H$_3$O$_2^-$ = 5.6×10^{-10}. This salt, at any concentration, produces a neutral solution.
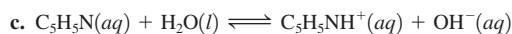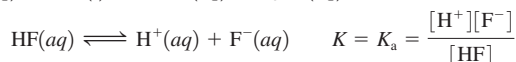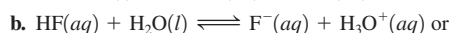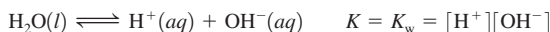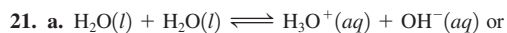
21. a. H$_2$O(l) + H$_2$O(l) $\rightleftharpoons$ H$_3$O$^+$(aq) + OH$^-$(aq) or
H$_2$O(l) $\rightleftharpoons$ H$^+$(aq) + OH$^-$(aq) $K = K_w = [\text{H}^+][\text{OH}^-]$
b. HF(aq) + H$_2$O(l) $\rightleftharpoons$ F$^-$(aq) + H$_3$O$^+$(aq) or
HF(aq) $\rightleftharpoons$ H$^+$(aq) + F$^-$(aq) $K = K_a = \dfrac{[\text{H}^+][\text{F}^-]}{[\text{HF}]}$
c. C$_5$H$_5$N(aq) + H$_2$O(l) $\rightleftharpoons$ C$_5$H$_5$NH$^+$(aq) + OH$^-$(aq)
$K = K_b = \dfrac{[\text{C}_5\text{H}_5\text{NH}^+][\text{OH}^-]}{[\text{C}_5\text{H}_5\text{N}]}$

23. a. This expression holds true for solutions of strong acids having a concentration greater than 1.0×10^{-6} M. For example, 0.10 M HCl, 7.8 M HNO$_3$, and 3.6×10^{-4} M HClO$_4$ are solutions where this expression holds true. **b.** This expression holds true for solutions of weak acids where the two normal assumptions hold. The two assumptions are that the contribution of H$^+$ from water is negligible and that the acid is less than 5% dissociated in water (from the assumption that x is small compared to some number). This expression will generally hold true for solutions of weak acids having a K_a value less than 1×10^{-4}, as long as there is a significant amount of weak acid present. Three example solutions are 1.5 M HC$_2$H$_3$O$_2$, 0.10 M HOCl, and 0.72 M NH$_4$NO$_3$.   **c.** This expression holds true for strong bases that donate 2 OH$^-$ ions per formula unit. As long as the concentration of the base is above 5×10^{-7} M, this expression will hold true. Three examples are 5.0×10^{-3} M Ca(OH)$_2$, 2.1×10^{-4} M Sr(OH)$_2$, and 9.1×10^{-5} M Ba(OH)$_2$. **d.** This expression holds true for solutions of weak bases where the two normal assumptions hold. The assumptions are that the OH$^-$ contribution from water is negligible and that the base is less than 5% ionized in water. For the 5% rule to hold, you generally need bases with $K_b < 1 \times 10^{-4}$ and concentrations of weak base greater than 0.10 M. Three examples are 0.10 M NH$_3$, 0.54 M C$_6$H$_5$NH$_2$, and 1.1 M C$_5$H$_5$N.   **25.** One reason HF is a weak acid is that the H—F bond is unusually strong and thus, is difficult to break. This contributes to the reluctance of the HF molecules to dissociate in water.   **27. a.** HClO$_4$(aq) + H$_2$O(l) $\rightarrow$ H$_3$O$^+$(aq) + ClO$_4^-$(aq) or HClO$_4$(aq) $\rightarrow$ H$^+$(aq)+ClO$_4^-$(aq); water is commonly omitted from K_a reactions. **b.** CH$_3$CH$_2$CO$_2$H(aq) $\rightleftharpoons$ H$^+$(aq) + CH$_3$CH$_2$CO$_2^-$(aq); **c.** NH$_4^+$(aq) $\rightleftharpoons$ H$^+$(aq) + NH$_3$(aq)   **29. a.** H$_2$O, base; H$_2$CO$_3$, acid; H$_3$O$^+$, conjugate acid; HCO$_3^-$, conjugate base; **b.** C$_5$H$_5$NH$^+$, acid; H$_2$O, base; C$_5$H$_5$N, conjugate base; H$_3$O$^+$, conjugate acid; **c.** HCO$_3^-$, base; C$_5$H$_5$NH$^+$, acid; H$_2$CO$_3$, conjugate acid; C$_5$H$_5$N, conjugate base   **31. a.** HClO$_4$, strong acid; **b.** HOCl, weak acid; **c.** H$_2$SO$_4$, strong acid; **d.** H$_2$SO$_3$, weak acid   **33.** HClO$_4$ > HClO$_2$ >NH$_4^+$ > H$_2$O   **35. a.** HCl; **b.** HNO$_2$; **c.** HCN since it has a larger K_a value.   **37. a.** 1.0×10^{-7} M, neutral; **b.** 12 M, basic; **c.** 8.3×10^{-16} M, acidic; **d.** 1.9×10^{-10} M, acidic   **39. a.** endothermic; **b.** [H$^+$] = [OH$^-$] = 2.34×10^{-7} M **41.** [37] **a.** pH = pOH = 7.00; **b.** pH = 15.08, pOH = −1.08; **c.** pH = −1.08, pOH = 15.08; **d.** pH = 4.27, pOH = 9.73 [38] **a.** pH = 14.18, pOH =

−0.18; **b.** pH = −0.44, pOH = 14.44; **c.** pH = pOH = 7.00; **d.** pH = 10.85, pOH = 3.14   **43. a.** pH = 6.88, pOH = 7.12, [H$^+$] = 1.3×10^{-7} M, [OH$^-$] = 7.6×10^{-8} M, acidic; **b.** pH = 0.92, pOH = 13.08, [H$^+$] = 0.12 M,   [OH$^-$] = 8.4×10^{-14} M,   acidic; **c.** pH = 10.89, pOH = 3.11, [H$^+$] = 1.3×10^{-11} M, [OH$^-$] = 7.8×10^{-4} M, basic; **d.** pH = pOH = 7.00, [H$^+$] = [OH$^-$] = 1.0×10^{-7} M, neutral   **45.** pOH = 11.9, [H$^+$] = 8×10^{-3} M, [OH$^-$] = 1×10^{-12} M, acidic   **47. a.** H$^+$, ClO$_4^-$, H$_2$O; 0.602; **b.** H$^+$, NO$_3^-$, H$_2$O; 0.602   **49.** [H$^+$] = 0.088 M, [OH$^-$] = 1.1×10^{-13} M, [Cl$^-$] = 0.013 M, [NO$_3^-$] = 0.075 M **51.** Add 4.2 mL of 12 M HCl to water with mixing; add enough water to bring the solution volume to 1600 mL.   **53. a.** HNO$_2$ and H$_2$O, 2.00; **b.** HC$_2$H$_3$O$_2$ and H$_2$O, 2.68   **55.** [H$^+$] = [CH$_3$COO$^-$] = 5.8×10^{-4} M,   [CH$_3$COOH] = 0.0181 M, pH = 3.24   **57.** [H$^+$] = [F$^-$] = 3.5×10^{-3} M, [OH$^-$] = 2.9×10^{-12} M, [HF] = 0.017 M, 2.46   **59.** 1.96   **61. a.** 1.00; **b.** 1.30   **63. a.** 0.60%; **b.** 1.9%; **c.** 5.8%; **d.** Dilution shifts equilibrium to the side with the greater number of particles (% dissociation increases). **e.** [H$^+$] also depends on initial concentration of weak acid.   **65.** 1.4×10^{-4} **67.** 3.5×10^{-4} **69.** 0.024 M
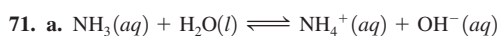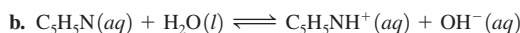
71. a. NH$_3$(aq) + H$_2$O(l) $\rightleftharpoons$ NH$_4^+$(aq) + OH$^-$(aq)
$K_b = \dfrac{[\text{NH}_4^+][\text{OH}^-]}{[\text{NH}_3]}$;
b. C$_5$H$_5$N(aq) + H$_2$O(l) $\rightleftharpoons$ C$_5$H$_5$NH$^+$(aq) + OH$^-$(aq)
$K_b = \dfrac{[\text{C}_5\text{H}_5\text{NH}^+][\text{OH}^-]}{[\text{C}_5\text{H}_5\text{N}]}$

73. NH$_3$ > C$_5$H$_5$N > H$_2$O > NO$_3^-$ **75. a.** C$_6$H$_5$NH$_2$; **b.** C$_6$H$_5$NH$_2$; **c.** OH$^-$; **d.** CH$_3$NH$_2$ **77. a.** 13.00; **b.** 7.00; **c.** 14.30   **79. a.** K$^+$, OH$^-$, and H$_2$O, 0.015 M, 12.18; **b.** Ba^{2+}, OH$^-$, and H$_2$O, 0.030 M, 12.48   **81.** 0.16 g   **83.** NH$_3$ and H$_2$O, 1.6×10^{-3} M, 11.20   **85. a.** [OH$^-$] = 8.9×10^{-3} M, [H$^+$] = 1.1×10^{-12} M, 11.96; **b.** [OH$^-$] = 4.7×10^{-5} M, [H$^+$] = 2.1×10^{-10} M, 9.68   **87.** 12.00   **89. a.** 1.3%; **b.** 4.2%   **91.** 9.2×10^{-7}

93. H$_2$SO$_3$(aq) $\rightleftharpoons$ HSO$_3^-$(aq) + H$^+$(aq) K_{a_1} reaction
HSO$_3^-$(aq) $\rightleftharpoons$ SO$_3^{2-}$(aq) + H$^+$(aq) K_{a_2} reaction
95. a. 1.62; **b.** 3.68   **97.** −0.30   **99.** HCl > NH$_4$Cl > KCl > KCN > KOH   **101.** OCl$^-$ **103.** [HN$_3$] = [OH$^-$] = 2.3×10^{-6} M, [Na$^+$] = 0.010 M, [N$_3^-$] = 0.010 M, [H$^+$] = 4.3×10^{-9} M **105. a.** 5.82; **b.** 10.95   **107.** NaF   **109.** 3.08   **111. a.** neutral; **b.** basic; NO$_2^-$ + H$_2$O $\rightleftharpoons$ HNO$_2$ + OH$^-$; **c.** acidic; C$_5$H$_5$NH$^+$ $\rightleftharpoons$ C$_5$H$_5$N + H$^+$; **d.** acidic because NH$_4^+$ is a stronger acid than NO$_2^-$ is a base; NH$_4^+$ $\rightleftharpoons$ NH$_3$ + H$^+$; NO$_2^-$ + H$_2$O $\rightleftharpoons$ HNO$_2$ + OH$^-$; **e.** basic; OCl$^-$ + H$_2$O $\rightleftharpoons$ HOCl + OH$^-$; **f.** basic because OCl$^-$ is a stronger base than NH$_4^+$ is an acid; OCl$^-$ + H$_2$O $\rightleftharpoons$ HOCl + OH$^-$, NH$_4^+$ $\rightleftharpoons$ NH$_3$ + H$^+$ **113. a.** HIO$_3$ < HBrO$_3$; as the electronegativity of the central atom increases, acid strength increases. **b.** HNO$_2$ < HNO$_3$; as the number of oxygen atoms attached to the central atom increases, acid strength increases. **c.** HOI < HOCl; same reasoning as in part a. **d.** H$_3$PO$_3$ < H$_3$PO$_4$; same reasoning as in part b.   **115. a.** H$_2$O < H$_2$S < H$_2$Se; acid strength increases as bond energy decreases. **b.** CH$_3$CO$_2$H < FCH$_2$CO$_2$H < F$_2$CHCO$_2$H < F$_3$CCO$_2$H; as the electronegativity of the neighboring atoms increases, acid strength increases. **c.** NH$_4^+$ < HONH$_3^+$; same reasoning as in part b. **d.** NH$_4^+$ < PH$_4^+$; same reasoning as in part a.   **117. a.** basic; CaO(s) + H$_2$O(l) $\rightarrow$ Ca(OH)$_2$(aq); **b.** acidic; SO$_2$(g) + H$_2$O(l) $\rightarrow$ H$_2$SO$_3$(aq); **c.** acidic; Cl$_2$O(g) + H$_2$O(l) $\rightarrow$ 2HOCl(aq)   **119. a.** B(OH)$_3$, acid; H$_2$O, base; **b.** Ag$^+$, acid; NH$_3$, base; **c.** BF$_3$, acid; F$^-$, base   **121.** Al(OH)$_3$(s) + 3H$^+$(aq) $\rightarrow$ Al^{3+}(aq) + 3H$_2$O(l); Al(OH)$_3$(s) + OH$^-$(aq) $\rightarrow$ Al(OH)$_4^-$(aq)   **123.** Fe^{3+}; because it is smaller with a greater positive charge, Fe^{3+} will be more strongly attracted to a lone pair of electrons from a Lewis base.   **125.** 990 mL H$_2$O   **127. a.** 2.80; **b.** 1.1×10^{-3} M **129.** NH$_4$Cl   **131.** 4.2×10^{-2} M **133.** 3.00   **135. a.** 2.62; **b.** 2.4%; **c.** 8.48   **137. a.** 1.66; **b.** Fe^{2+} ions will produce a less acidic solution (higher pH) due to the lower charge on Fe^{2+} as compared with Fe^{3+}. As the charge on a metal ion increases, acid strength of the hydrated ion increases.   **139.** acidic; HSO$_4^-$ $\rightleftharpoons$ SO$_4^{2-}$ + H$^+$; 1.54   **141. a.** Hb(O$_2$)$_4$ in lungs, HbH$_4^{4+}$ in cells; **b.** Decreasing [CO$_2$] will decrease [H$^+$], favoring Hb(O$_2$)$_4$ formation. Breathing into a bag raises [CO$_2$]. **c.** NaHCO$_3$ lowers the acidity from accumulated CO$_2$.

143. a. H_2SO_3; **b.** $HClO_3$; **c.** H_3PO_3; NaOH and KOH are ionic compounds composed of either Na^+ or K^+ cations and OH^- anions. When soluble ionic compounds dissolve in water, they form the ions from which they are formed. The acids in this problem are all covalent compounds. When these acids dissolve in water, the covalent bond between oxygen and hydrogen breaks to form H^+ ions. **145.** 7.20. **147.** 4540 mL **149.** 4.17 **151.** 0.022 M **153.** 2.5×10^{-3} **155.** PO_4^{3-}, $K_b = 0.021$; HPO_4^{2-}, $K_b = 1.6 \times 10^{-7}$; $H_2PO_4^-$, $K_b = 1.3 \times 10^{-12}$; from the K_b values, PO_4^{3-} is the strongest base. **157. a.** basic; **b.** acidic; **c.** basic; **d.** acidic; **e.** acidic **159.** 1.0×10^{-3} **161.** 5.4×10^{-4} **163.** 3.36

Chapter 15

13. When an acid dissociates or when a salt dissolves, ions are produced. A common ion is when one of the product ions in a particular equilibrium is added from an outside source. For a weak acid dissociating to its conjugate base and H^+, the common ion would be the conjugate base; this would be added by dissolving a soluble salt of the conjugate base into the acid solution. The presence of the conjugate base from an outside source shifts the equilibrium to the left so less acid dissociates. For the K_{sp} reaction of a salt dissolving into its respective ions, the common ion would be if one of the ions in the salt was added from an outside source. When a common ion is present, the K_{sp} equilibrium shifts to the left resulting in less of the salt dissolving into its ions. **15.** The more weak acid and conjugate base present, the more H^+ and/or OH^- that can be absorbed by the buffer without significant pH change. When the concentrations of weak acid and conjugate base are equal (so that $pH = pK_a$), the buffer system is equally efficient at absorbing either H^+ or OH^-. If the buffer is overloaded with weak acid or with conjugate base, then the buffer is not equally efficient at absorbing either H^+ or OH^-. **17.** The three key points to emphasize in your sketch are the initial pH, pH at the halfway point to equivalence, and the pH at the equivalence point. For the two weak bases titrated, $pH = pK_a$ at the halfway point to equivalence (50.0 mL HCl added) because [weak base] = [conjugate acid] at this point. For the initial pH, the strong base has the highest pH (most basic), while the weakest base has the lowest pH (least basic). At the equivalence point, the strong base titration has $pH = 7.0$. The weak bases titrated have acidic pHs at the equivalence point because the conjugate acids of the weak bases titrated are the major species present. The weakest base has the strongest conjugate acid so its pH will be lowest (most acidic) at the equivalence point.
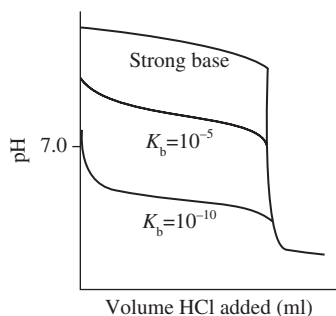
19. i. This is the result when you have a salt that breaks up into two ions. Examples of these salts (but not all) would be AgCl, $SrSO_4$, $BaCrO_4$, and $ZnCO_3$ **ii.** This is the result when you have a salt that breaks up into three ions, either two cations and one anion or one cation and two anions. Some examples are SrF_2, Hg_2I_2, and Ag_2SO_4. **iii.** This is the result when you have a salt that breaks up into four ions, either three cations and one anion (Ag_3PO_4) or one cation and three anions (ignoring the hydroxides, there are no examples of this type of salt in Table 15.4). **iv.** This is the result when you have a salt that breaks up into five ions, either three cations and two anions $[Sr_3(PO_4)_2]$ or two cations and three anions (no examples of this type of salt are in Table 15.4). **21.** When strong acid or strong base is added to a sodium bicarbonate/sodium carbonate buffer mixture, the strong acid/base is neutralized. The reaction goes to completion resulting in the strong acid/base being replaced with a weak acid/base. This results in a new buffer solution. The reactions are $H^+(aq) + CO_3^{2-}(aq) \rightarrow$ $HCO_3^-(aq)$; $OH^-(aq) + HCO_3^-(aq) \rightarrow CO_3^{2-}(aq) + H_2O(l)$ **23. a.** 2.96; **b.** 8.94; **c.** 7.00; **d.** 4.89 **25.** 1.1% vs. 1.3×10^{-2}% dissociated; the presence of $C_3H_5O_2^-$ in solution 23d greatly inhibits the dissociation of $HC_3H_5O_2$. This is called the *common ion effect*. **27. a.** 1.70; **b.** 5.49; **c.** 1.70; **d.** 4.71 **29. a.** 4.29; **b.** 12.30; **c.** 12.30; **d.** 5.07 **31.** solution d; solution d is a buffer solution that resists pH changes. **33.** 3.40 **35.** 3.48; 3.22 **37.** 4.36 **39. a.** 7.97; **b.** 8.73; both solutions have an initial pH = 8.77. The two solutions differ in their buffer capacity. Solution b with the larger concentrations has the greater capacity to resist pH change. **41.** 15 g **43. a.** 0.19; **b.** 0.59; **c.** 1.0; **d.** 1.9 **45.** HOCl; there are many possibilities. One possibility is a solution with [HOCl] = 1.0 M and [NaOCl] = 0.35 M. **47.** solution d **49. a.** 1.0 mol; **b.** 0.30 mol; **c.** 1.3 mol **51. a.** ~22 mL base added; **b.** buffer region is from ~1 mL to ~21 mL base added. The maximum buffering region would be from ~5 mL to ~17 mL of base added with the halfway point to equivalence (~11 mL) as the best buffer point. **c.** ~11 mL base added; **d.** 0 mL base added; **e.** ~22 mL base added (the stoichiometric point); **f.** any point after the stoichiometric point (volume base added > ~22 mL) **53. a.** 0.699; **b.** 0.854; **c.** 1.301; **d.** 7.00; **e.** 12.15 **55. a.** 2.72; **b.** 4.26; **c.** 4.74; **d.** 5.22; **e.** 8.79; **f.** 12.15

57.

Volume (mL)	pH
0.0	2.43
4.0	3.14
8.0	3.53
12.5	3.86
20.0	4.46
24.0	5.24
24.5	5.6
24.9	6.3
25.0	8.28
25.1	10.3
26.0	11.29
28.0	11.75
30.0	11.96

See *Solutions Guide* for pH plot.

59.

Volume (mL)	pH
0.0	11.11
4.0	9.97
8.0	9.58
12.5	9.25
20.0	8.65
24.0	7.87
24.5	7.6
24.9	6.9
25.0	5.28
25.1	3.7
26.0	2.71
28.0	2.24
30.0	2.04

See *Solutions Guide* for pH plot.

61. a. 4.19, 8.45; **b.** 10.74; 5.96; **c.** 0.89, 7.00 **63.** 2.1×10^{-6} **65. a.** yellow; **b.** 8.0; **c.** blue **67.** phenolphthalein **69.** Phenol red is one possible indicator for the titration in Exercise 53. Phenolphthalein is one possible indicator for the titration in Exercise 55. **71.** Phenolphthalein is one possible indicator for Exercise 57. Bromcresol green is one possible indicator for Exercise 59. **73.** The pH is between 5 and 8. **75. a.** $AgC_2H_3O_2(s) \rightleftharpoons$ $Ag^+(aq) + C_2H_3O_2^-(aq)$; $K_{sp} = [Ag^+][C_2H_3O_2^-]$; **b.** $Al(OH)_3(s) \rightleftharpoons$ $Al^{3+}(aq) + 3OH^-(aq)$; $K_{sp} = [Al^{3+}][OH^-]^3$; **c.** $Ca_3(PO_4)_2(s) \rightleftharpoons$ $3Ca^{2+}(aq) + 2PO_4^{3-}(aq)$; $K_{sp} = [Ca^{2+}]^3[PO_4^{3-}]^2$ **77. a.** 2.3×10^{-9}; **b.** 8.20×10^{-19} **79.** 3.92×10^{-5} **81. a.** 1.6×10^{-5} mol/L; **b.** 9.3×10^{-5} mol/L; **c.** 6.5×10^{-7} mol/L **83.** 3.30×10^{-43} **85.** 2.5×10^{-22} mol/L **87. a.** CaF_2; **b.** $FePO_4$ **89. a.** 4×10^{-17} mol/L; **b.** 4×10^{-11} mol/L; **c.** 4×10^{-29} mol/L **91.** 2.3×10^{-11} mol/L **93.** 1.5×10^{-19} g **95.** If the anion in the salt can act as a base in water, then the solubility of the salt will increase

as the solution becomes more acidic. Added H^+ will react with the base, forming the conjugate acid. As the basic anion is removed, more of the salt will dissolve to replenish the basic anion. The salts with basic anions are Ag_3PO_4, $CaCO_3$, $CdCO_3$, and $Sr_3(PO_4)_2$. Hg_2Cl_2 and PbI_2 do not have any pH dependence because Cl^- and I^- are terrible bases (the conjugate bases of strong acids).

$$Ag_3PO_4(s) + H^+(aq) \longrightarrow 3Ag^+(aq) + HPO_4^{2-}(aq) \xrightarrow{\text{excess } H^+}$$
$$3Ag^+(aq) + H_3PO_4(aq)$$

$$CaCO_3(s) + H^+(aq) \longrightarrow Ca^{2+}(aq) + HCO_3^-(aq) \xrightarrow{\text{excess } H^+}$$
$$Ca^{2+}(aq) + H_2CO_3(aq) \, [H_2O(l) + CO_2(g)]$$

$$CdCO_3(s) + H^+(aq) \longrightarrow Cd^{2+}(aq) + HCO_3^-(aq) \xrightarrow{\text{excess } H^+}$$
$$Cd^{2+}(aq) + H_2CO_3(aq) \, [H_2O(l) + CO_2(g)]$$

$$Sr_3(PO_4)_2(s) + 2H^+(aq) \longrightarrow 3Sr^{2+}(aq) + 2HPO_4^{2-}(aq) \xrightarrow{\text{excess } H^+}$$
$$3Sr^{2+}(aq) + 2H_3PO_4(aq)$$
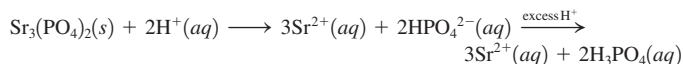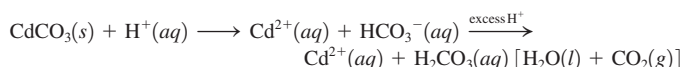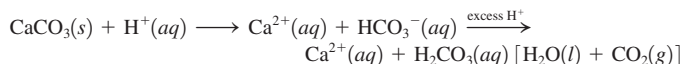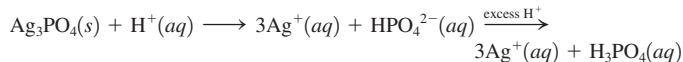
97. yes; $Q = 1.9 \times 10^{-4} > K_{sp}$ **99.** $[K^+] = 0.160 \, M$, $[C_2O_4^{2-}] = 3.3 \times 10^{-7} \, M$, $[Ba^{2+}] = 0.0700 \, M$, $[Br^-] = 0.300 \, M$ **101.** $[Ag^+] > 5.6 \times 10^{-5} \, M$
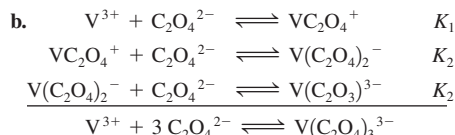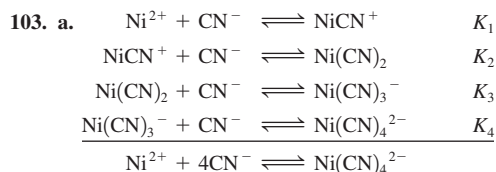
103. a.

$$Ni^{2+} + CN^- \rightleftharpoons NiCN^+ \qquad K_1$$
$$NiCN^+ + CN^- \rightleftharpoons Ni(CN)_2 \qquad K_2$$
$$Ni(CN)_2 + CN^- \rightleftharpoons Ni(CN)_3^- \qquad K_3$$
$$\underline{Ni(CN)_3^- + CN^- \rightleftharpoons Ni(CN)_4^{2-} \qquad K_4}$$
$$Ni^{2+} + 4CN^- \rightleftharpoons Ni(CN)_4^{2-}$$

b.

$$V^{3+} + C_2O_4^{2-} \rightleftharpoons VC_2O_4^+ \qquad K_1$$
$$VC_2O_4^+ + C_2O_4^{2-} \rightleftharpoons V(C_2O_4)_2^- \qquad K_2$$
$$\underline{V(C_2O_4)_2^- + C_2O_4^{2-} \rightleftharpoons V(C_2O_3)^{3-} \qquad K_2}$$
$$V^{3+} + 3\,C_2O_4^{2-} \rightleftharpoons V(C_2O_4)_3^{3-}$$

105. 6.2×10^5 **107.** $Hg^{2+}(aq) + 2I^-(aq) \rightarrow HgI_2(s)$ (orange precipitate); $HgI_2(s) + 2I^-(aq) \rightarrow HgI_4^{2-}(aq)$ (soluble complex ion)   **109.** $3.3 \times 10^{-32} \, M$ **111. a.** 1.2×10^{-8} mol/L; **b.** 1.5×10^{-4} mol/L; **c.** The presence of NH_3 increases the solubility of AgI. Added NH_3 removes Ag^+ from solution by forming the complex ion $Ag(NH_3)_2^+$. As Ag^+ is removed, more AgI will dissolve to replenish the Ag^+ concentration.   **113.** 4.7×10^{-2} mol/L   **115.** Test tube 1: added Cl^- reacts with Ag^+ to form the silver chloride precipitate. The net ionic equation is $Ag^+(aq) + Cl^-(aq) \rightarrow AgCl(s)$. Test tube 2: added NH_3 reacts with Ag^+ ions to form the soluble complex ion $Ag(NH_3)_2^+$. As this complex ion forms, Ag^+ is removed from solution, which causes AgCl(s) to dissolve. When enough NH_3 is added, then all of the silver chloride precipitate will dissolve. The equation is $AgCl(s) + 2NH_3(aq) \rightarrow Ag(NH_3)_2^+(aq) + Cl^-(aq)$. Test tube 3: added H^+ reacts with the weak base NH_3 to form NH_4^+. As NH_3 is removed, Ag^+ ions are released to solution, which can then react with Cl^- to reform AgCl(s). The equations are $Ag(NH_3)_2^+(aq) + 2H^+(aq) \rightarrow Ag^+(aq) + 2NH_4^+(aq)$ and $Ag^+(aq) + Cl^-(aq) \rightarrow AgCl(s)$.   **117.** $pOH = pK_b + \log\dfrac{[\text{acid}]}{[\text{base}]}$ **119. a.** 10.44; **b.** 10.74   **121. a.** potassium fluoride + HCl; **b.** benzoic acid + NaOH; **c.** acetic acid + sodium acetate; **d.** HOCl + NaOH; **e.** ammonium chloride + NaOH   **123. a.** 1.8×10^9; **b.** 5.6×10^4; **c.** 1.0×10^{14} **125.** 4.4 L   **127.** 180. g/mol; 3.3×10^{-4}; assume acetylsalicylic acid is a weak monoprotic acid.   **129.** 65 mL   **131.** 0.210 M **133. a.** 1.6×10^{-6}; **b.** 0.056 mol/L   **135.** 2.7×10^{-5} mol/L; the solubility of hydroxyapatite will increase as a solution gets more acidic, since both phosphate and hydroxide can react with H^+. 6×10^{-8} mol/L; the hydroxyapatite in the tooth enamel is converted to the less soluble fluorapatite by fluoride-treated water. The less soluble fluorapatite will then be more difficult to dissolve, making teeth less susceptible to decay. See Chemical Impact on "The Chemistry of Teeth."   **137. a.** 6.7×10^{-6} mol/L; **b.** 1.2×10^{-13} mol/L; **c.** $Pb(OH)_2(s)$ will not form since $Q < K_{sp}$ **139.** 49 mL   **141.** 3.9 L   **143. a.** 200.0 mL; **b.** i. H_2A, H_2O; ii. H_2A, HA^-, H_2O, Na^+; iii. HA^-, H_2O, Na^+; iv. HA^-, A^{2-}, H_2O, Na^+; v. A^{2-}, H_2O, Na^+; vi. A^{2-}, H_2O, Na^+, OH^-;

c. $K_{a_1} = 1 \times 10^{-4}$; $K_{a_2} = 1 \times 10^{-8}$ **145.** $pH \approx 5.0$; $K_a \approx 1 \times 10^{-10}$ **147.** 3 M **149. a.** 5.8×10^{-4} mol/L; **b.** Greater; F^- is a weak base ($K_b = 1.4 \times 10^{-11}$), so some of the F^- is removed by reaction with water. As F^- is removed, more SrF_2 will dissolve; **c.** 3.5×10^{-3} mol/L   **151.** 3.00   **153.** 2.78

Chapter 16

7. Living organisms need an external energy source to produce the necessary "ordering."   **9.** ΔS_{univ} **11.** As any process occurs, ΔS_{univ} will increase; ΔS_{univ} cannot decrease. Time also goes in one direction, just as ΔS_{univ} goes in one direction.   **13.** Possible arrangements for one molecule:

1 way   1 way
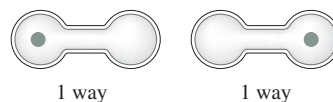
Both are equally probable.
Possible arrangements for two molecules:

1 way   2 ways   1 way
Most probable
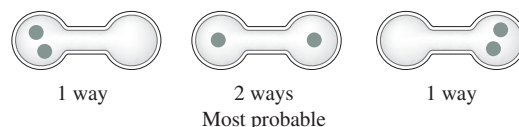
Possible arrangement for three molecules:

1 way   3 ways   3 ways   1 way

Equally most probable
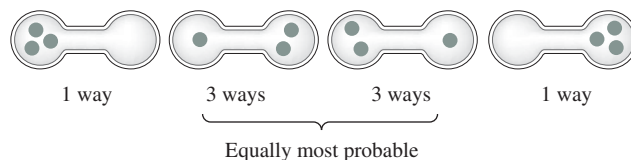
15. Note that these substances are not in the solid state, but are in the aqueous state; water molecules are also present. There is an apparent increase in ordering when these ions are placed in water. The hydrating water molecules must be in a highly ordered state when surrounding these anions.   **17.** One can determine $\Delta S°$ and $\Delta H°$ for the reaction using the standard entropies and standard enthalpies of formation in Appendix 4, then use the equation $\Delta G° = \Delta H° - T\Delta S°$. One can also use the standard free energies of formation in Appendix 4. And finally, one can use Hess's law to calculate $\Delta G°$. Here, reactions having known $\Delta G°$ values are manipulated to determine $\Delta G°$ for a different reaction. For temperatures other than 25(C, $\Delta G°$ is estimated using the $\Delta G° = \Delta H° - T\Delta S$ equation. The assumptions made are that the $\Delta H°$ and $\Delta S°$ values determined from Appendix 4 data are temperature independent. We use the same $\Delta H°$ and $\Delta S°$ values as determined when $T = 25°C$, then plug in the new temperature in Kelvin into the equation to estimate $\Delta G°$ at the new temperature.   **19.** a, b, c   **21.** We draw all of the possible arrangements of the two particles in the three levels.

2 kJ	—	—	X	—	X	XX
1 kJ	—	X	—	XX	X	—
0 kJ	XX	X	X	—	—	—
Total E =	0 kJ	1 kJ	2 kJ	2 kJ	3 kJ	4 kJ

The most likely total energy is 2 kJ.   **23. a.** H_2 at 100°C and 0.5 atm; **b.** N_2 at STP; **c.** $H_2O(l)$ **25. a.** negative; **b.** positive   **27.** $\Delta G < 0$ for b, c, d   **29.** 89.3 J/K · mol   **31. a.** yes ($\Delta G < 0$); **b.** 196 K   **33. a.** negative; **b.** positive; **c.** negative; **d.** positive   **35. a.** $C_{graphite}(s)$; **b.** $C_2H_5OH(g)$; **c.** $CO_2(g)$ **37. a.** negative, -186 J/K; **b.** positive, 187 J/K; **c.** hard to predict since $\Delta n = 0$.; 138 J/K   **39.** 262 J/K · mol   **41. a.** ΔH and ΔS are both positive; **b.** $S_{rhombic}$ **43. a.** ΔH and ΔS are both negative; **b.** low

temperatures **45. a.** $\Delta H° = -803$ kJ, $\Delta S° = -4$ J/K, $\Delta G° = -802$ kJ; **b.** $\Delta H° = 2802$ kJ, $\Delta S° = -262$ J/K, $\Delta G° = 2880.$ kJ; **c.** $\Delta H° = -416$ kJ, $\Delta S° = -209$ J/K, $\Delta G° = -354$ kJ; **d.** $\Delta H° = -176$ kJ, $\Delta S° = -284$ J/K, $\Delta G° = -91$ kJ **47.** -5.40 kJ; 328.6 K; $\Delta G°$ is negative below 328.6 K. **49.** $CH_4(g) + CO_2(g) \rightarrow CH_3CO_2H(l)$, $\Delta H° = -16$ kJ, $\Delta S° = -240.$ J/K, $\Delta G° = 56$ KJ; $CH_3OH(g) + CO(g) \rightarrow CH_3CO_2H(l)$, $\Delta H° = -173$ kJ, $\Delta S° = -278$ J/K, $\Delta G° = -90.$ kJ; the second reaction is preferred. It should be run at temperatures below 622 K. **51.** -817 kJ **53.** -731 kJ/mol **55.** yes **57.** -188 kJ **59. a.** shifts right; **b.** no shift since the reaction is at equilibrium; **c.** shifts left **61.** 8.72; 0.0789 **63.** 140 kJ **65.** -71 kJ/mol **67.** $\Delta H° = 1.1 \times 10^5$ J/mol; $\Delta S° = 330$ J/K · mol; The major difference in the plot is the slope of the line. An endothermic process has a negative slope for the $\ln(K)$ versus $1/T$ plot, whereas an exothermic process has a positive slope. **69.** -447 J/K · mol **71.** decreases; ΔS will be negative since 2 mol of gaseous reactants form 1 mol of gaseous product. For ΔG to be negative, ΔH must be negative (exothermic). For exothermic reactions, K decreases as T increases, so the ratio of the partial pressure of PCl_5 to the partial pressure of PCl_3 will decrease. **73.** 43.7 K **75.** 60 **77. a.** 1.8×10^4 J/mol; shifts left; **b.** 0; no shift since at equilibrium; **c.** -1.1×10^4 J/mol; shifts right; **d.** 0; no shift since at equilibrium; **e.** 2×10^3 J/mol; shifts left **79. a.** 2.22 $\times 10^5$; **b.** 94.3; **c.** 0.29 mol ATP **81.** ΔS is more favorable for reaction 2 than for reaction 1, resulting in $K_2 > K_1$. In reaction 1, seven particles in solution form one particle. In reaction 2, four particles form one, which results in a smaller decrease in disorder than for reaction 1. **83.** 725 K **85.** $\Delta H° = 286$ kJ; $\Delta G° = 326$ kJ; $K = 7.22 \times 10^{-58}$; $P_{O_3} = 3.3 \times 10^{-41}$ atm; This partial pressure represents one molecule of ozone per 9.5×10^{17} L of air. Equilibrium is probably not maintained under the conditions because the concentration of ozone is not large enough to maintain equilibrium **87. a.** Because

$$k_f = A \exp\left(\frac{-E_a}{RT}\right) \text{ and } k_r = A \exp\left(\frac{-(E_a - \Delta G°)}{RT}\right),$$

Because $\dfrac{k_f}{k_r} = \exp\left(\dfrac{-E_a}{RT} + \dfrac{(E_a - \Delta G°)}{RT}\right) = \exp\left(\dfrac{-\Delta G°}{RT}\right)$. Because $K = \exp\left(\dfrac{-\Delta G°}{RT}\right)$, then $K = \dfrac{k_f}{k_r}$. **b.** A catalyst increases the value of the rate constant (increases rate) by lowering the activation energy. For the equilibrium constant K to remain constant, both k_f and k_r must increase by the same factor. Therefore, a catalyst must increase the rate of both the forward and the reverse reactions. **89. a.** 0.333; **b.** $P_A = 1.50$ atm; $P_B = 0.50$ atm; **c.** $\Delta G = \Delta G° + RT \ln(P_B/P_A) = 2722$ J $- 2722$ J $= 0$ **91.** at least 7.5 torr **93.** 16 g **95.** 61 kJ **97.** -4.1 kJ/mol

Chapter 17

13. Oxidation: increase in oxidation number, loss of electrons; reduction: decrease in oxidation number, gain of electrons **15.** Reactions a, b, and c are oxidation–reduction reactions.

Oxidizing Agent	Reducing Agent	Substance Oxidized	Substance Reduced
a. H_2O	CH_4	$CH_4(C)$	$H_2O(H)$
b. $AgNO_3$	Cu	Cu	$AgNO_3(Ag)$
c. HCl	Zn	Zn	HCl(H)

17. Magnesium is an alkaline earth metal; Mg will oxidize to Mg^{2+}. The oxidation state of hydrogen in HCl is $+1$. To be reduced, the oxidation state of H must decrease. The obvious choice for the hydrogen product is $H_2(g)$ where hydrogen has a zero oxidation state. The balanced reaction is: $Mg(s) + 2HCl(aq) \rightarrow MgCl_2(aq) + H_2(g)$. Mg goes from the 0 to the $+2$ oxidation state by losing two electrons. Each H atom goes from the 0 to the -1 oxidation state by gaining one electron. Since there are two H atoms in the balanced equation, then a total of two electrons are gained by the H atoms. Hence, two electrons are transferred in the balanced reaction. When the electrons are transferred directly from Mg to H^+, no work is obtained. In order to harness this reaction to do useful work, we must control the flow of

electrons through a wire. This is accomplished by making a galvanic cell which separates the reduction reaction from the oxidation reaction in order to control the flow of electrons through a wire to produce a voltage. **19.** An extensive property is one that depends on the amount of substance. The free energy change for a reaction depends on whether 1 mol of product is produced or 2 mol of product is produced or 1 million mol of product is produced. This is not the case for cell potentials which do not depend on the amount of substance. The equation that relates ΔG to E is $\Delta G = -nFE$. It is the n term that converts the intensive property E into the extensive property ΔG. The n is the number of mol of electrons transferred in the balanced reaction that ΔG is associated with. **21.** A potential hazard when jump-starting a car is that the electrolysis of $H_2O(l)$ can occur. When $H_2O(l)$ is electrolyzed, the products are the explosive gas mixture of $H_2(g)$ and $O_2(g)$. A spark produced during jump starting a car could ignite any $H_2(g)$ and $O_2(g)$ produced. Grounding the jumper cable far from the battery minimizes the risk of a spark nearby the battery where $H_2(g)$ and $O_2(g)$ could be collecting. **23.** You need to know the identity of the metal so you know which molar mass to use. You need to know the oxidation state of metal ion in the salt so the mol of electrons transferred can be determined. And finally, you need to know the amount of current and the time the current was passed through the electrolytic cell. If you know these four quantities, then the mass of metal plated out can be calculated. **25.** See Figure 17.3 of the text for a typical galvanic cell. The anode compartment contains the oxidation half-reaction compounds/ions, and the cathode compartment contains the reduction half-reaction compounds/ions. The electrons flow from the anode to the cathode. For each of the following answers, all solutes are 1.0 M and all gases are at 1.0 atm. **a.** $7H_2O(l) + 2Cr^{3+}(aq) + 3Cl_2(g) \rightarrow Cr_2O_7^{2-}(aq) + 6Cl^-(aq) + 14H^+(aq)$; cathode: Pt electrode; Cl_2 bubbled into solution, Cl^- in solution; anode: Pt electrode; Cr^{3+}, H^+, and $Cr_2O_7^{2-}$ in solution; **b.** $Cu^{2+}(aq) + Mg(s) \rightarrow Cu(s) + Mg^{2+}(aq)$; cathode: Cu electrode; Cu^{2+} in solution; anode: Mg electrode; Mg^{2+} in solution **27. a.** 0.03 V; **b.** 2.71 V **29.** See Exercise 25 for a description of a galvanic cell. For each of the following answers, all solutes are 1.0 M and all gases are at 1.0 atm. In the salt bridge, cations flow to the cathode and anions flow to the anode. **a.** $Cl_2(g) + 2Br^-(aq) \rightarrow Br_2(aq) + 2Cl^-(aq)$, $\mathscr{E}° = 0.27$ V; cathode: Pt electrode; $Cl_2(g)$ bubbled in, Cl^- in solution; anode: Pt electrode; Br_2 and Br^- in solution; **b.** $3H_2O(l) + 5IO_4^-(aq) + 2Mn^{2+}(aq) \rightarrow 5IO_3^-(aq) + 2MnO_4^-(aq) + 6H^+(aq)$, $\mathscr{E}° = 0.09$ V; cathode: Pt electrode; IO_4^-, IO_3^-, and H_2SO_4 (as a source of H^+) in solution; anode: Pt electrode; Mn^{2+}, MnO_4^-, and H_2SO_4 in solution **31.** 25a. $Pt|Cr^{3+}$ (1.0 M), H^+ (1.0 M), $Cr_2O_7^{2-}$ (1.0 M) $\|$ Cl_2 (1.0 atm)$|Cl^-$ (1.0 M) | Pt; 25b. Mg | Mg^{2+} (1.0 M) $\|$ Cu^{2+} (1.0 M) | Cu; 29a. Pt | Br^- (1.0 M), Br_2 (1.0 M) $\|$ Cl_2 (1.0 atm) | Cl^- (1.0 M) | Pt; 29b. Pt | Mn^{2+} (1.0 M), MnO_4^- (1.0 M), H^+ (1.0 M) $\|$ IO_4^- (1.0 M), H^+(1.0 M) IO_3^- (1.0 M), | Pt **33. a.** $Au^{3+}(aq) + 3Cu^+(aq) \rightarrow 3Cu^{2+}(aq) + Au(s)$, $\mathscr{E}° = 1.34$ V; **b.** $2VO_2^+(aq) + 4H^+(aq) + Cd(s) \rightarrow Cd^{2+}(aq) + 2VO^{2+}(aq) + 2H_2O(l)$, $\mathscr{E}° = 1.40$ V **35. a.** $16H^+ + 2MnO_4^- + 10I^- \rightarrow 5I_2 + 2Mn^{2+} + 8H_2O$, $\mathscr{E}°_{cell} = 0.97$ V, spontaneous; **b.** $16H^+ + 2MnO_4^- + 10F^- \rightarrow 5F_2 + 2Mn^{2+} + 8H_2O$, $\mathscr{E}°_{cell} = -1.36$ V, not spontaneous **37.** $\mathscr{E}° = 0.41$ V, $\Delta G° = -79$ kJ **39.** 33a. -388 kJ; 33b. $-270.$ kJ **41.** 1.21 V **43.** $K^+ < H_2O < Cd^{2+} < I_2 < AuCl_4^- < IO_3^-$ **45. a.** no; **b.** yes; **c.** yes; **d.** no **47. a.** $Cr_2O_7^{2-}$, O_2, MnO_2, IO_3^-; **b.** $PbSO_4$, Cd^{2+}, Fe^{2+}, Cr^{3+}, Zn^{2+}, H_2O **49.** $ClO^-(aq) + 2NH_3(aq) \rightarrow Cl^-(aq) + N_2H_4(aq) + H_2O(l)$, $\mathscr{E}°_{cell} = 1.00$ V; Because $\mathscr{E}°_{cell}$ is positive for this reaction, at standard conditions ClO^- can spontaneously oxidize NH_3 to the somewhat toxic N_2H_4. **51. a.** larger; **b.** smaller **53.** Electron flow is always from the anode to the cathode. For the cells with a nonzero cell potential, we will identify the cathode, which means the other compartment is the anode. **a.** 0; **b.** 0.018 V; compartment with $[Ag^+] = 2.0$ M is cathode; **c.** 0.059 V; compartment with $[Ag^+] = 1.0$ M is cathode; **d.** 0.26 V; compartment with $[Ag^+] = 1.0$ M is cathode; **e.** 0 **55.** 2.12 V **57.** 1.09 V **59. a.** 0.23 V; **b.** 1.2×10^{-5} M **61.** 0.16 V, copper is oxidized. **63.** 1.7×10^{-30} **65.** [25] **a.** $\Delta G° = -20$ kJ; 1×10^3; **b.** $\Delta G° = -523$ kJ; 5.12×10^{91}; [29] **a.** $\Delta G° = -52$ kJ; 1.4×10^9; **b.** $\Delta G° = -90$ kJ; 2×10^{15} **67.** 2.5×10^{26} **69. a.** no reaction; **b.** $Cl_2(g) + 2I^-(aq) \rightarrow I_2(s) + 2Cl^-(aq)$, $\mathscr{E}°_{cell} = 0.82$ V; $\Delta G° = -160$ kJ; $K = 5.6 \times 10^{27}$; **c.** no reaction; **d.** $4Fe^{2+}(aq) + 4H^+(aq) + O_2(g) \rightarrow$

$4Fe^{3+}(aq) + 2H_2O(l)$, $\mathscr{E}°_{cell} = 0.46$ V; $\Delta G° = -180$ kJ; $K = 1.3 \times 10^{31}$; **71. a.** $Au^{3+}(aq) + 3Tl(s) \rightarrow Au(s) + 3Tl^+(aq)$; $\mathscr{E}°_{cell} = 1.84$ V; **b.** $\Delta G° = -533$ kJ; $K = 2.52 \times 10^{93}$; **c.** 2.04 V   **73.** 5.1×10^{-20} **75.** 6.19×10^{52} **77. a.** 30. hours; **b.** 33 s; **c.** 1.3 hours   **79. a.** 16 g; **b.** 25 g; **c.** 71 g; **d.** 4.9 g   **81.** Bi   **83.** 9.12 L F_2 (anode), 29.2 g K (cathode)   **85.** 7.44×10^4 A   **87.** 1.14×10^{-2} M **89.** Au followed by Ag followed by Ni followed by Cd   **91. a.** cathode: $Ni^{2+} + 2e^- \rightarrow Ni$; anode: $2Br^- \rightarrow Br_2 + 2e^-$; **b.** cathode: $Al^{3+} + 3e^- \rightarrow Al$; anode: $2F^- \rightarrow F_2 + 2e^-$; **c.** cathode: $Mn^{2+} + 2e^- \rightarrow Mn$; anode: $2I^- \rightarrow I_2 + 2e^-$ **93. a.** 0.10 V, SCE is anode; **b.** 0.53 V, SCE is anode; **c.** 0.02 V, SCE is cathode; **d.** 1.90 V, SCE is cathode; **e.** 0.47 V, SCE is cathode   **95. a.** decrease; **b.** increase; **c.** decrease; **d.** decrease; **e.** same   **97. a.** $\Delta G° = -582$ kJ; $K = 3.45 \times 10^{102}$; $\mathscr{E}° = 1.01$ V; **b.** -0.65 V;   **99.** Aluminum has the ability to form a durable oxide coating over its surface. Once the HCl dissolves this oxide coating, Al is exposed to H^+ and is easily oxidized to Al^{3+}. Thus, the Al foil disappears after the oxide coating is dissolved.   **101.** The claim is impossible. The strongest oxidizing agent and reducing agent when combined give $\mathscr{E}°$ of only about 6 V.   **103.** $w_{max} = -13,200$ kJ; the work done can be no larger than the free energy change. If the process were reversible all of the free energy released would go into work, but this does not occur in any real process. Fuel cells are more efficient in converting chemical energy to electrical energy; they are also less massive. Major disadvantage: They are expensive.   **105.** 0.98 V

107. 0.250 mol   **109.** +3   **111.** $\mathscr{E}° = \dfrac{T\Delta S°}{nF} - \dfrac{\Delta H°}{nF}$; if we graph $\mathscr{E}°$ versus T, we should get a straight line ($y = mx + b$). The slope of the line is equal to $\Delta S°/nF$ and the y-intercept is equal to $-\Delta H°/nF$. $\mathscr{E}°$ will have little temperature dependence for cell reactions with $\Delta S°$ close to zero.   **113.** 9.8×10^{-6} **115.** 2.39×10^{-7} **117. a.** ± 0.02 pH units; $\pm 6 \times 10^{-6}$ M H^+; **b.** ± 0.001 V   **119. a.** 0.16 V; **b.** 8.6 mol   **121.** $[Ag^+] = 4.6 \times 10^{-18}$ M; $[Ni^{2+}] = 1.5$ M **123. a.** 0.12 V; **b.** 0.54 V   **125. a.** 5.77×10^{10}; **b.** -12.2 kJ/mol   **127.** Osmium(IV) nitrate; $[Ar]4s^13d^{10}$

Chapter 18

1. The characteristic frequencies of energies emitted in a nuclear reaction suggest that discrete energy levels exist in the nucleus. The extra stability of certain numbers of nucleons and the predominance of nuclei with even numbers of nucleons suggest that the nuclear structure might be described by using quantum numbers.   **3.** β-particle production has the net effect of turning a neutron into a proton. Radioactive nuclei having too many neutrons typically undergo β-particle decay. Positron production has the net effect of turning a proton into a neutron. Nuclei having too many protons typically undergo positron decay.   **5.** The transuranium elements are the elements having more protons than uranium. They are synthesized by bombarding heavier nuclei with neutrons and positive ions in a particle accelerator.   **7.** $\Delta E = \Delta mc^2$; The key difference is the mass change when going from reactants to products. In chemical reactions, the mass change is indiscernible. In nuclear processes, the mass change is discernable. It is the conversion of this discernable mass change into energy that results in the huge energies associated with nuclear processes. **9.** Sr-90 is an alkaline earth metal having chemical properties similar to calcium. Sr-90 can collect in bones replacing some of the calcium. Once imbedded inside the human body, β particles can do significant damage. Rn-222 is a noble gas so one would expect Rn to be unreactive and pass through the body quickly; it does. The problem with Rn-222 is the rate at which it produces alpha particles. With a short half-life, the few moments that Rn-222 is in the lungs, a significant number of decay events can occur; each decay event produces an alpha particle which is very effective at causing ionization and can produce a dense trail of damage.   **11. a.** $^{51}_{24}Cr + ^{0}_{-1}e \rightarrow ^{51}_{23}V$; **b.** $^{131}_{53}I \rightarrow ^{0}_{-1}e + ^{131}_{54}Xe$ **13. a.** $^{68}_{31}Ga + ^{0}_{-1}e \rightarrow ^{68}_{30}Zn$; **b.** $^{62}_{29}Cu \rightarrow ^{0}_{+1}e + ^{62}_{28}Ni$; **c.** $^{212}_{87}Fr \rightarrow ^{4}_{2}He + ^{208}_{85}At$; **d.** $^{129}_{51}Sb \rightarrow ^{0}_{-1}e + ^{129}_{52}Te$ **15.** 10 α particles; 5 β particles   **17.** $^{53}_{26}Fe$ has too many protons. It will undergo positron production, electron capture, and/or alpha-particle production. $^{59}_{26}Fe$ has too many neutrons and will undergo beta-particle production. (See Table 18.2 of the text.)

19. a. $^{249}_{98}Cf + ^{18}_{8}O \rightarrow ^{263}_{106}Sg + 4^{1}_{0}n$; **b.** $^{259}_{104}Rf$ **21.** 690 hours   **23.** ^{81}Kr is most stable since it has the longest half-life. ^{73}Kr is "hottest" since it decays very rapidly due to its very short half-life. ^{73}Kr, 81s; ^{74}Kr, 34.5 min; ^{76}Kr, 44.4 h; ^{81}Kr, 6.3×10^5 yr   **25.** 6.22 mg ^{32}P remains   **27.** 0.230   **29.** 26 g   **31.** 2.3 counts per minute per gram of C. No; for a 10.-mg C sample, it would take roughly 40 min to see a single disintegration. This is too long to wait, and the background radiation would probably be much greater than the ^{14}C activity. Thus ^{14}C dating is not practical for very small samples.   **33.** 3.8×10^9 yr   **35.** 4.3×10^6 kg/s   **37.** ^{232}Pu, -1.715×10^{14} J/mol; ^{231}Pa, -1.714×10^{14} J/mol   **39.** ^{12}C: 1.23×10^{-12} J/nucleon; ^{235}U: 1.2154×10^{-12} J/nucleon; since ^{56}Fe is the most stable known nucleus, then the binding energy per nucleon for ^{56}Fe would be larger than that of ^{12}C or ^{235}U. (See Fig. 18.9 of the text.)   **41.** 6.01513 amu   **43.** -2.0×10^{10} J/g of hydrogen nuclei   **45.** The Geiger–Müller tube has a certain response time. After the gas in the tube ionizes to produce a "count," some time must elapse for the gas to return to an electrically neutral state. The response of the tube levels off because, at high activities, radioactive particles are entering the tube faster than the tube can respond to them.   **47.** All evolved $O_2(g)$ comes from water.   **49.** 2 neutrons; 4 β particles   **51.** Strontium. Xe is chemically unreactive and not readily incorporated into the body. Sr can be easily oxidized to Sr^{2+}. Strontium is in the same family as calcium and could be absorbed and concentrated in the body in a fashion similar to Ca^{2+}. The chemical properties determine where radioactive material may be concentrated in the body or how easily it may be excreted.   **53. a.** unstable; beta production; **b.** stable; **c.** unstable; positron production or electron capture; **d.** unstable, positron production, electron capture, or alpha production.   **55.** 49.7 yr   **57.** 1975   **59.** 900 g ^{235}U **61.** 7×10^5 m/s; 8×10^{-16} J/nuclei;   **63.** Assuming that (1) the radionuclide is long lived enough that no significant decay occurs during the time of the experiment, and (2) the total activity is uniformly distributed only in the rat's blood; $V = 10.$ mL;   **65. a.** $^{12}_{6}C$; **b.** ^{13}N, ^{13}C, ^{14}N, ^{15}O, and ^{15}N; **c.** -5.950×10^{11} J/mol ^1H **67.** 4.3×10^{-29} **69.** $^{249}_{97}Bk + ^{22}_{10}Ne \rightarrow ^{267}_{107}Bh + 4^{1}_{0}n$; 62.7s; $[Rn]7s^25f^{14}6d^5$

Chapter 19

1. The gravity of the earth cannot keep H_2 in the atmosphere.   **3.** The acidity decreases. Solutions of Be^{2+} are acidic, while solutions of the other M^{2+} ions are neutral.   **5.** The planes of carbon atoms slide easily. Graphite is not volatile so the lubricant will not be lost when used in a high vacuum environment.   **7.** p-type semiconductor   **9.** For groups 1A–3A, the small size of H (as compared to Li), Be (as compared to Mg), and B (as compared to Al) seems to be the reason why these elements have nonmetallic properties, while others in the groups 1A–3A are strictly metallic. The small size of H, Be, and B also causes these species to polarize the electron cloud in nonmetals, thus forcing a sharing of electrons when bonding occurs. For groups 4A–6A, a major difference between the first and second members of a group is the ability to form π bonds. The smaller elements form stable π bonds, while the larger elements do not exhibit good overlap between parallel p orbitals and, in turn, do not form strong π bonds. For group 7A, the small size of F as compared to Cl is used to explain the low electron affinity of F and the weakness of the F—F bond.   **11.** In order to maximize hydrogen bonding interactions in the solid phase, ice is forced into an open structure. This open structure is why $H_2O(s)$ is less dense than $H_2O(l)$.   **13. a.** $\Delta H° = 207$ kJ, $\Delta S° = 216$ J/K; **b.** $T > 958$ K   **15. a.** lithium oxide; **b.** potassium superoxide; **c.** sodium peroxide   **17. a.** $Li_2O(s) + H_2O(l) \rightarrow 2LiOH(aq)$; **b.** $Na_2O_2(s) + 2H_2O(l) \rightarrow 2NaOH(aq) + H_2O_2(aq)$; **c.** $LiH(s) + H_2O(l) \rightarrow H_2(g) + LiOH(aq)$; **d.** $2KO_2(s) + 2H_2O(l) \rightarrow 2KOH(aq) + O_2(g) + H_2O_2(aq)$ **19.** $2Li(s) + 2C_2H_2(g) \rightarrow 2LiC_2H(s) + H_2(g)$; oxidation–reduction   **21. a.** magnesium carbonate; **b.** barium sulfate; **c.** strontium hydroxide   **23.** $CaCO_3(s) + H_2SO_4(aq) \rightarrow CaSO_4(aq) + H_2O(l) + CO_2(g)$ **25.** In the gas phase, linear molecules would exist:

$$: \ddot{F} - Be - \ddot{F} :$$
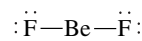
In the solid state, BeF_2 has the following extended structure:
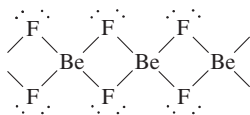
27. 2×10^{-2} *M* **29.** 3.84×10^6 g Ba **31. a.** AlN; **b.** GaF_3; **c.** Ga_2S_3
33. $B_2H_6(g) + 3O_2(g) \rightarrow 2B(OH)_3(s)$ **35.** $In_2O_3(s) + 6H^+(aq) \rightarrow 2In^{3+}(aq) +$
$3H_2O(l)$; $In_2O_3(s) + OH^-(aq) \rightarrow$ no reaction; $Ga_2O_3(s) + 6H^+(aq) \rightarrow$
$2Ga^{3+}(aq) + 3H_2O(l)$; $Ga_2O_3(s) + 2OH^-(aq) + 3H_2O(l) \rightarrow 2Ga(OH)_4^-(aq)$
37. $2Ga(s) + 3F_2(g) \rightarrow 2GaF_3(s)$; $4Ga(s) + 3O_2(g) \rightarrow 2Ga_2O_3(s)$; $16Ga(s) +$
$3S_8(s) \rightarrow 8Ga_2S_3(s)$; $2Ga(s) + N_2(g) \rightarrow 2GaN(s)$; $2Ga(s) + 6HCl(aq) \rightarrow$
$2GaCl_3(aq) + 3H_2(g)$

39.

Tetrahedral;	Tetrahedral;	Octahedral;
$109.5°$; sp^3	$109.5°$; sp^3	$90°$; d^2sp^3
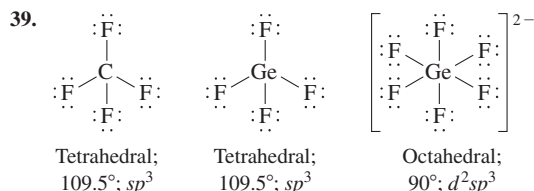
To form CF_6^{2-}, carbon would have to expand its octet of electrons. Carbon compounds do not expand their octet because of the small atomic size of carbon and because there are no low energy *d* orbitals on carbon to accommodate the extra electrons. **41. a.** $SiO_2(s) + 2C(s) \rightarrow Si(s) + 2CO(g)$; **b.** $SiCl_4(l) +$
$2Mg(s) \rightarrow Si(s) + 2MgCl_2(s)$; **c.** $Na_2SiF_6(s) + 4Na(s) \rightarrow Si(s) + 6NaF(s)$
43. Lead is very toxic. As the temperature of the water increases, the solubility of Pb increases. Drinking hot tap water from pipes containing lead solder could result in higher Pb concentrations in the body. **45.** $C_6H_{12}O_6(aq) \rightarrow$
$2C_2H_5OH(aq) + 2CO_2(g)$ **47.** The π electrons are free to move in graphite, thus giving it a greater conductivity (lower resistance). The electrons have the greatest mobility within the sheets of carbon atoms. Electrons in diamond are not mobile (high resistance). The structure of diamond is uniform in all directions; thus there is no directional dependence of the resistivity. **49.** Only some of the ice will melt; 0°C **51.** It is feasible to recycle Al by melting the metal because, in theory, it takes less than 1% of the energy required to produce the same amount of Al by the Hall–Heroult process. **53.** 60 **55.** 2.12 V
57. Strontium and calcium are both alkaline earth metals, so both have similar chemical properties. Since milk is a good source of calcium, strontium could replace some calcium in milk without much difficulty. **59.** The inert-pair effect refers to the difficulty of removing the pair of valence *s* electrons from some of the elements in the fifth and sixth periods of the periodic table. As a result, multiple oxidation states are exhibited for the heavier elements of Groups 3A (and 4A). In^+, In^{3+}, Tl^+, and Tl^{3+} oxidation states are all important to the chemistry of In and Tl. **61.** 3.08 **63.** If the compound contained Ga(II) it would be paramagnetic and if the compound contained Ga(I) and Ga(III), it would be diamagnetic. Paramagnetic compounds have an apparent greater mass in a magnetic field. **65.** 59 atm in the gas phase; 1.8 mol CO_2/L in the wine
67. $Pb(NO_3)_2(aq) + H_3AsO_4(aq) \rightarrow PbHAsO_4(s) + 2HNO_3(aq)$ **69.** Ca;
12.698 **71.** $3.34, 5.6 \times 10^{-11}$ *M* **73.** 2.0×10^{-37} *M* **75.** Carbon is much smaller than Si and cannot form a fifth bond in the transition state. **77.** I
79. a. 7.1 g; **b.** 979 nm; This electromagnetic radiation is not visible to humans; it is in the infrared region of the electromagnetic radiation spectrum, **c.** n-type

Chapter 20

1. This is due to nitrogen's ability to form strong π bonds whereas heavier group 5A elements do not form strong π bonds. Therefore, P_2, As_2, and Sb_2 do not form since two π bonds are required to form these diatomic substances.
3. There are medical studies that have shown an inverse relationship between the incidence of cancer and the selenium levels in soil. The foods grown in these soils and eventually digested are assumed to somehow furnish protection from cancer. Selenium is also involved in the activity of vitamin E and certain enzymes in the human body. In addition, selenium deficiency has been linked

to the occurrence of congestive heart failure. **5.** +6 oxidation state: SO_4^{2-}, SO_3, SF_6; +4 oxidation state: SO_3^{2-}, SO_2, SF_4; +2 oxidation state: SCl_2; 0 oxidation state: S_8 and all other elemental forms of sulfur; −2 oxidation state: H_2S, Na_2S **7. a.** $H_2(g) + Cl_2(g) \rightarrow 2HCl(g)$; this reaction produces a lot of energy which can be used in a cannon apparatus to send a stopper flying across the room. To initiate this extremely slow reaction, light of specific wavelengths is needed. This is the purpose of lighting the magnesium strip. When magnesium is oxidized to MgO, an intense white light is produced. Some of the wavelengths of this light can break Cl—Cl bonds and get the reaction started. **b.** Br_2 is brown. The disappearance of the brown color indicates that all of the Br_2 has reacted with the alkene (no free Br_2 remains). **c.** $2Al(s) + 3I_2(s) \rightarrow 2AlI_3(s)$; This is a highly exothermic reaction, hence the sparks that accompany this reaction. The purple smoke is excess $I_2(s)$ being vaporized [the purple smoke is $I_2(g)$]. **9.** Nitrogen's small size does not provide room for all four oxygen atoms, making NO_4^{3-} unstable. Phosphorus is larger so PO_4^{3-} is more stable. To form NO_3^-, a pi bond must form. Phosphorus doesn't form strong pi bonds as readily as nitrogen. **11.** $NO_2 = N_2O_4 < NO < N_2O$ **13. a.** $NH_4NO_3(s) \xrightarrow{Heat}$
$N_2O(g) + 2H_2O(g)$; **b.** $2N_2O_5(g) \rightarrow 4NO_2(g) + O_2(g)$; **c.** $2K_3P(s) +$
$6H_2O(l) \rightarrow 2PH_3(g) + 6KOH(aq)$; **d.** $PBr_3(l) + 3H_2O(l) \rightarrow H_3PO_3(aq) +$
$3HBr(aq)$; **e.** $2NH_3(aq) + NaOCl(aq) \rightarrow N_2H_4(aq) + NaCl(aq) + H_2O(l)$
15. $CaF_2 \cdot 3Ca_3(PO_4)_2(s) + 10H_2SO_4(aq) + 20H_2O(l) \rightarrow 6H_3PO_4(aq) +$
$2HF(aq) + 10CaSO_4 \cdot 2H_2O(s)$ **17.** 2.08 mol **19.** $N_2H_4(l) + O_2(g) \rightarrow$
$N_2(g) + 2H_2O(g)$; $\Delta H = -590$. kJ **21.** $\frac{1}{2}N_2(g) + \frac{1}{2}O_2(g) \rightarrow NO(g)$
$\Delta G° = \Delta G°_{f(NO)}$; NO (and some other oxides of nitrogen) have weaker bonds as compared with the triple bond of N_2 and the double bond of O_2. Because of this, NO (and some other oxides of nitrogen) has a higher (positive) standard free energy of formation as compared to the relatively stable N_2 and O_2 molecules.
23.

M.O.		Bond order	# unpaired e$^-$
	NO	2.5	1
	NO$^+$	3	0
	NO$^-$	2	2

Lewis

Lewis structures are not adequate for NO and NO^-. The M.O. model gives correct results for all three species. For NO, Lewis structures fail for odd-electron species. For NO^-, Lewis structures fail to predict that NO^- is paramagnetic. **25. a.** $H_3PO_4 > H_3PO_3$; **b.** $H_3PO_4 > H_2PO_4^- > HPO_4^{2-}$
27. The acidic protons are attached to oxygen.

$H_4P_2O_6$ $H_4P_2O_5$
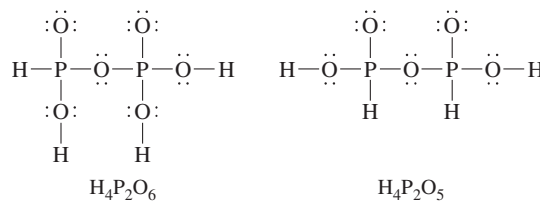
29. 821 nm **31. a.** $2SO_2(g) + O_2(g) \rightarrow 2SO_3(g)$; **b.** $SO_3(g) + H_2O(l) \rightarrow$
$H_2SO_4(aq)$; **c.** $2Na_2S_2O_3(aq) + I_2(aq) \rightarrow Na_2S_4O_6(aq) + 2NaI(aq)$; **d.** $Cu(s) +$
$2H_2SO_4(aq) \rightarrow CuSO_4(aq) + 2H_2O(l) + SO_2(aq)$
33. a.

Trigonal pyramid; $\approx 109.5°$; sp^3

b.

V-shaped; $\approx 120°$; sp^2

c.

V-shaped; $\approx 109.5°$; sp^3

d.

See-saw; $a \approx 120°$, $b \approx 90°$; dsp^3

e.

Octahedral; $90°$; d^2sp^3
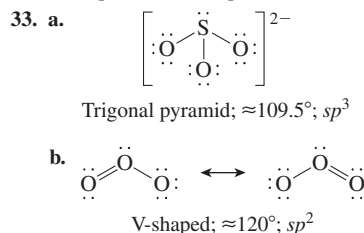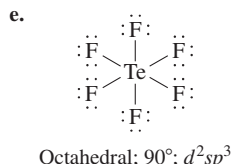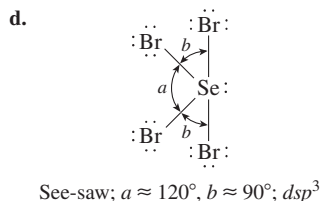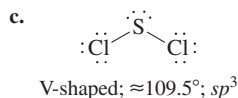
35. 0.301 g H_2O_2; 3.6×10^{-2} g excess HCl **37.** From the following Lewis structure, each oxygen atom has a tetrahedral arrangement of electron pairs. Therefore, bond angles $\approx 109.5°$, and each O is sp^3 hybridized.
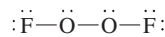
Formal charge: 0   0   0   0
Oxidation state: −1 +1 +1 −1

Oxidation states are more useful. We are forced to assign +1 as the oxidation state for this element. **39. a.** $BaCl_2(s) + H_2SO_4(aq) \rightarrow BaSO_4(s) +$ $2HCl(g)$; **b.** $BrF(s) + H_2O(l) \rightarrow HF(aq) + HOBr(aq)$; **c.** $SiO_2(s) +$ $4HF(aq) \rightarrow SiF_4(g) + 2H_2O(l)$ **41.** ClO^- can oxide NH_3 to the somewhat toxic N_2H_4. **43. a.** IO_4^-; **b.** IO_3^-; **c.** IF_2^-; **d.** IF_4^-; **e.** IF_6^- **45.** XeF_2 can react with oxygen and water to produce explosive xenon oxides and oxy-fluorides. **47.** Release of Sr is probably more harmful. Xe is chemically un-reactive. Strontium is in the same family as calcium and could be absorbed and concentrated in the body in a fashion similar to Ca. This puts the radioactive Sr in the bones, and red blood cells are produced in bone marrow. Xe would not be readily incorporated into the body. The chemical properties determine where a radioactive material may concentrate in the body or how easily it may be excreted. The length of time of exposure and the body part exposed to ra-diation significantly affects the health hazard. **49.** As the halogen atoms get larger, it becomes more difficult to fit three halogens around the small N.

51.

Formal charge: 0   0 −1   −1 0   0   +1 0 −2

Formal charge: −2 +1 0   −1 +1 −1   −3 +1 +1
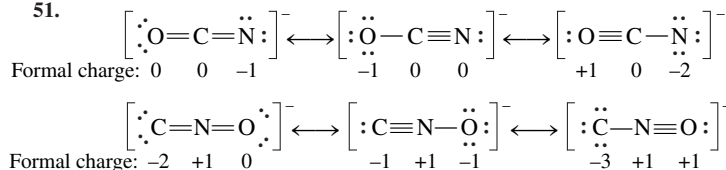
All the resonance structures for fulminate involve greater formal charges than in cyanate, making fulminate more reactive (less stable). **53.** 32 kg bacterial tissue

55.
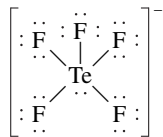
The lone pair of electrons around Te exerts a stronger repulsion than the bond-ing pairs, pushing the four square planar F's away from the lone pair. **57.** exothermic **59.** $MgSO_4(s) \rightarrow Mg^{2+}(aq) + SO_4^{2-}(aq)$; $NH_4NO_3(s) \rightarrow$ $NH_4^+(aq) + NO_3^-(aq)$ **61.** F, Cl, Br, or I; trigonal pyramid; $\approx 109.5°$ **63.** XeF_2

65. For the reaction
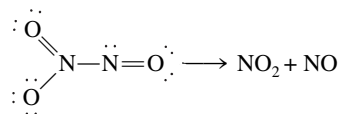
the activation energy must in some way involve the breaking of a nitrogen–nitrogen single bond. For the reaction
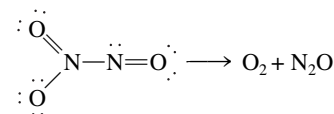
at some point nitrogen–oxygen bonds must be broken. N—N single bonds (160 kJ/mol) are weaker than N—O single bonds (201 kJ/mol). In addition, reso-nance structures indicate that there is more double-bond character in the N—O bonds than in the N—N bond. Thus NO_2 and NO are preferred by kinetics be-cause of the lower activation energy. **67. a.** NO; **b.** NO_2; **c.** $k_{cat}/k_{un} = 2.3$; **d.** $ClO(g) + O(g) \rightarrow O_2(g) + Cl(g)$; $O_3(g) + O(g) \rightarrow 2O_2(g)$; **e.** The Cl-catalyzed reaction is roughly 52 times faster (more efficient) than the NO-catalyzed reaction. **69.** 5.89 **71.** 20. g **73. a.** 287 kJ/mol; **b.** $[IF_2]^+$: V-shaped; sp^3; $[BF_4]^-$: tetrahedral; sp^3

Chapter 21

5. $Fe_2O_3(s) + 6H_2C_2O_4(aq) \rightarrow 2Fe(C_2O_4)_3^{3-}(aq) + 3H_2O(l) + 6H^+(aq)$; The oxalate anion forms a soluble complex ion with iron in rust (Fe_2O_3), which allows rust stains to be removed. **7.** No; both the *trans* or the *cis* forms of $Co(NH_3)_4Cl_2^+$ have mirror images that are superimposable. For the *cis* form, the mirror image only needs a $90°$ rotation to produce the original structure. Hence, neither the *trans* nor *cis* forms are optically active.

9. a.

small Δ

$CoCl_4^{2-}$ is an example of a weak-field case having three unpaired electrons.

b.

large Δ

CN^- is a strong field ligand so $Co(CN)_6^{3-}$ will be a low-spin case having zero unpaired electrons.

11. At high altitudes, the oxygen content of air is lower, so less oxyhemoglo-bin is formed which diminishes the transport of oxygen in the blood. A seri-ous illness called high-altitude sickness can result from the decrease of O_2 in the blood. High altitude acclimatization is the phenomena that occurs in the human body in response to the lower amounts of oxyhemoglobin in the blood. This response is to produce more hemoglobin, and, hence, increase the oxyhemoglobin in the blood. High-altitude acclimatization takes several weeks to take hold for people moving from lower altitudes to higher altitudes. **13. a.** Ni: $[Ar]4s^23d^8$; **b.** Cd: $[Kr]5s^24d^{10}$; **c.** Zr: $[Kr]5s^24d^2$; **d.** Os: $[Xe]6s^24f^{14}5d^6$ **15. a.** Ti: $[Ar]4s^23d^2$; Ti^{2+}: $[Ar]3d^2$; Ti^{4+}: $[Ne]3s^23p^6$ or $[Ar]$; **b.** Re: $[Xe]6s^24f^{14}5d^5$; Re^{2+}: $[Xe]4f^{14}5d^5$; Re^{3+}: $[Xe]4f^{14}5d^4$; **c.** Ir: $[Xe]6s^24f^{14}5d^7$; Ir^{2+}: $[Xe]4f^{14}5d^7$; Ir^{3+}: $[Xe]4f^{14}5d^6$ **17. a.** Fe^{3+}: $[Ar]3d^5$; **b.** Ag^+: $[Kr]4d^{10}$; **c.** Ni^{2+}: $[Ar]3d^8$; **d.** Cr^{3+}: $[Ar]3d^3$ **19. a.** molybdenum(IV) sulfide, molybdenum(VI) oxide; **b.** MoS_2, +4; MoO_3, +6; $(NH_4)_2 Mo_2O_7$, +6; $(NH_4)_6Mo_7O_{24} \cdot 4H_2O$, +6 **21.** The lan-thanide elements are located just before the 5d transition metals. The lanthanide contraction is the steady decrease in the atomic radii of the lanthanide elements when going from left to right across the periodic table. As a result of the lan-thanide contraction, the sizes of the 4d and 5d elements are very similar. This leads to a greater similarity in the chemistry of the 4d and 5d elements in a given vertical group. **23.** If rain is imminent, the large amount of water vapor in the air would cause the reaction to shift to the right. The indicator would take

on the color of the pink $CoCl_2 \cdot 6H_2O$. **25.** Test tube 1: added Cl^- reacts with Ag^+ to form the silver chloride precipitate. The net ionic equation is $Ag^+(aq) + Cl^-(aq) \rightarrow AgCl(s)$. Test tube 2: added NH_3 reacts with Ag^+ ions to form the soluble complex ion $Ag(NH_3)_2^+$. As this complex ion forms, Ag^+ is removed from the solution, which causes the $AgCl(s)$ to dissolve. When enough NH_3 is added, then all of the silver chloride precipitate will dissolve. The equation is $AgCl(s) + 2NH_3(aq) \rightarrow Ag(NH_3)_2^+(aq) + Cl^-(aq)$. Test tube 3: added H^+ reacts with the weak base NH_3 to form NH_4^+. As NH_3 is removed, Ag^+ ions are released to solution, which can then react with Cl^- to reform $AgCl(s)$. The equations are $Ag(NH_3)_2^+(aq) + 2H^+(aq) \rightarrow Ag^+(aq) + 2NH_4^+(aq)$ and $Ag^+(aq) + Cl^-(aq) \rightarrow AgCl(s)$. **27.** $[Co(NH_3)_6]I_3$: 3 mol AgI; $[Pt(NH_3)_4I_2]I_2$: 2 mol AgI; $Na_2[PtI_6]$: 0 mol AgI; $[Cr(NH_3)_4I_2]I$: 1 mol AgI. **29. a.** pentaamminechlororuthenium(III) ion; **b.** hexacyanoferrate(II) ion; **c.** tris(ethylenediamine)manganese(II) ion; **d.** pentaamminenitrocobalt(III) ion **31. a.** hexaamminecobalt(II) chloride; **b.** hexaaquacobalt(III) iodide; **c.** potassiumtetrachloro-platinate(II); **d.** potassium hexachloroplatinate(II); **e.** pentaamminechlorocobalt(III) chloride; **f.** triamminetrinitrocobalt(III) **33. a.** $K_2[CoCl_4]$; **b.** $[Pt(H_2O)(CO)_3]Br_2$; **c.** $Na_3[Fe(CN)_2(C_2O_4)_2]$; **d.** $[Cr(NH_3)_3Cl(H_2NCH_2CH_2NH_2)]I_2$

35. a.

cis   trans

b.

cis   trans

c.

cis   trans

d.
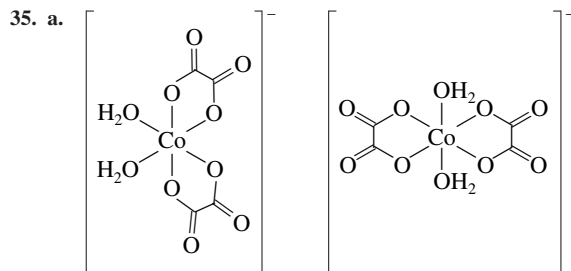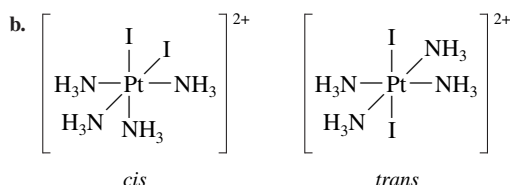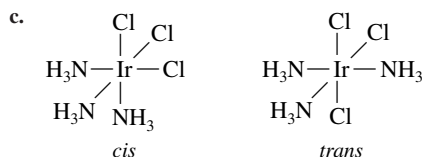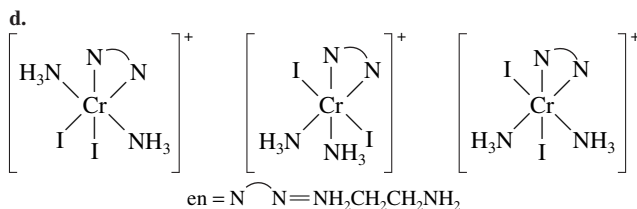
$en = N\frown N = NH_2CH_2CH_2NH_2$

37.
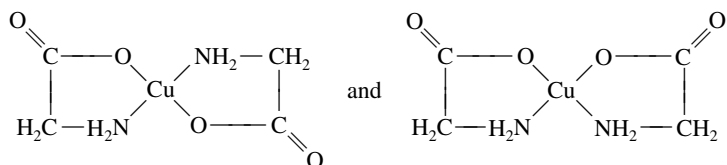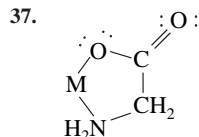
and

39. SCN^-, NO_2^-, and OCN^- can form linkage isomers; all are able to bond to the metal ion in two different ways. **41.** $Cr(acac)_3$ and cis-$Cr(acac)_2(H_2O)_2$ are optically active.

43. a. Fe^{2+}

High spin   Low spin

b. Fe^{3+} **c.** Ni^{2+}

High spin
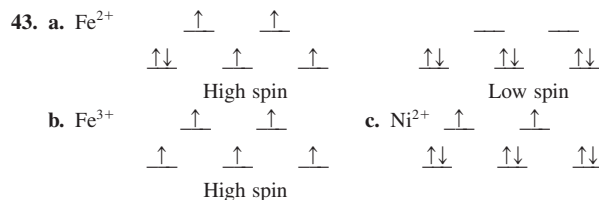
45. weak field **47. a.** 0; **b.** 2; **c.** 2 **49.** $[Co(CN)_6]^{3-} < [Co(en)_3]^{3+} < [Co(H_2O)_6]^{3+} < [CoI_6]^{3-}$ **51.** The violet complex ion absorbs yellow-green light ($\lambda \approx 570$ nm), the yellow complex ion absorbs blue light ($\lambda \approx 450$ nm), and the green complex ion absorbs red light ($\lambda \approx 650$ nm). The violet complex ion is $Cr(H_2O)_6^{3+}$, the yellow complex ion is $Cr(NH_3)_6^{3+}$, and the green complex ion is $Cr(H_2O)_4Cl_2^+$. **53.** $CoBr_4^{2-}$ is a tetrahedral complex ion, while $CoBr_6^{4-}$ is an octahedral complex ion. Since tetrahedral d-orbital splitting is less than one-half the octahedral d-orbital splitting, the octahedral complex ion ($CoBr_6^{4-}$) will absorb higher-energy light, which will have a shorter wavelength than 3.4×10^{-6} m ($E = hc/\lambda$). **55.** 5 **57. a.** -11 kJ; **b.** $\Delta H° = 172.5$ kJ; $\Delta S° = 176$ J/K; $T > 980.$ K **59.** $8CN^-(aq) + 4Ag(s) + O_2(g) + 2H_2O(l) \rightarrow 4Ag(CN)_2^-(aq) + 4OH^-(aq)$ **61.** $[Cr(NH_3)_5I]I_2$; octahedral **63.** $Hg^{2+}(aq) + 2I^-(aq) \rightarrow HgI_2(s)$, orange precipitate; $HgI_2(s) + 2I^-(aq) \rightarrow HgI_4^{2-}(aq)$, soluble complex ion **65. a.** 2; **b.** 3; **c.** 4; **d.** 4 **67. a.** optical isomerism

b.
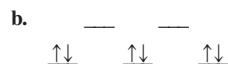
69. Octahedral Cr^{2+} complexes should be used. Cr^{2+}: $[Ar]3d^4$; High-spin (weak-field) Cr^{2+} complexes have four unpaired electrons and low-spin (strong-field) Cr^{2+} complexes have two unpaired electrons. Ni^{2+}: $[Ar]3d^8$; Octahedral Ni^{2+} complexes will always have two unpaired electrons, whether high or low spin. Therefore, Ni^{2+} complexes cannot be used to distinguish weak- from strong-field ligands by examining magnetic properties. Alternatively, the ligand field strengths can be measured using visible spectra. Either Cr^{2+} or Ni^{2+} complexes can be used for this method. **71.** $Pb(OH)_2$ will not form since Q is less than K_{sp}. **73.** 60 **75.** $Ni^{2+} = d^8$; If A and B produced very similar crystal fields, the complex $trans$-$[NiA_2B_4]^{2+}$ would give an octahedral crystal field diagram:

This is paramagnetic. Because the complex ion is diamagnetic, ligands A and B must produce different crystal fields resulting in a unique d-orbital splitting diagram. **77. a.** -0.26 V; **b.** From standard reduction potentials, Co^{3+} ($\mathscr{E}° = 1.82$ V) is a much stronger oxidizing agent than $Co(en)_3^{3+}$ ($\mathscr{E}° = -0.26$ V); **c.** In aqueous solution, Co^{3+} forms the hydrated transition metal complex, $Co(H_2O)_6^{3+}$. In both complexes, $Co(H_2O)_6^{3+}$ and $Co(en)_3^{3+}$, cobalt exists as Co^{3+}, which has 6 d electrons. If we assume a strong-field case for each complex ion, then the d-orbital splitting diagram for each has the six electrons paired in the lower-energy t_{2g} orbitals. When each complex ion gains an electron, the electron enters the higher-energy e_g orbitals. Since en is a stronger-field ligand than H_2O, then the d-orbital splitting is larger for $Co(en)_3^{3+}$, and it takes more energy to add an electron to $Co(en)_3^{3+}$ than to $Co(H_2O)_6^{3+}$. Therefore, it is more favorable for $Co(H_2O)_6^{3+}$ to gain an electron than for $Co(en)_3^{3+}$ to gain an electron. **79.** No, since in all three cases six bonds are formed between Ni^{2+} and nitrogen. So ΔH values should be similar. $\Delta S°$ for formation of the complex ion is most negative for $6NH_3$ molecules reacting with a metal ion (7 independent species become 1). For penten reacting with a metal ion, 2 independent species become 1, so $\Delta S°$ is the least negative. Thus the chelate effect occurs because the more bonds a chelating agent can form to the metal,

5342

the more favorable $\Delta S°$ is for the formation of the complex ion, and the larger the formation constant.

81.
$$\underline{\quad}\qquad d_z{}^2$$
$$\underline{\quad}\ \underline{\quad}\qquad d_{x^2-y^2},\ d_{xy}$$
$$\underline{\quad}\ \underline{\quad}\qquad d_{xz},\ d_{yz}$$
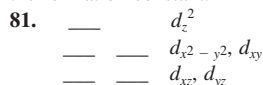
83. Place the *trans* NH_3 ligands on the z axis with the CN^- ligands oriented as indicated in the following diagram. Since CN^- produces a much stronger crystal field, the diagram will most resemble that of a square planar complex:

$$
\begin{array}{ll}
d_{x^2-y^2} & \underline{\quad} \\
d_{xy} & \text{⥮} \\
d_{z^2} & \text{⥮} \\
d_{xz} & \text{⥮}\ \text{⥮}\quad d_{yz}
\end{array}
$$
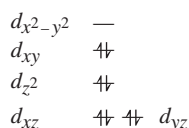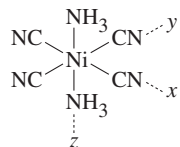
85. a. $+6$; $[Ar]3d^2$; **b.** $0.0496\ L\ N_2$

Chapter 22

1. a. 1-sec-butylpropane

$$CH_3\underset{\underset{CH_2CH_3}{|}}{CH}CH_2CH_3$$
(CH_2CH_2CH_3)
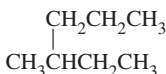
3-methylhexane is correct.

c. 2-ethylpentane

$$CH_3\underset{\underset{CH_2CH_3}{|}}{CH}CH_2CH_2CH_3$$
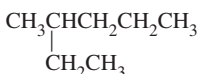
3-methylhexane is correct.

e. 3-methylhexane

$$CH_3CH_2\underset{\underset{CH_3}{|}}{CH}CH_2CH_2CH_3$$

b. 4-methylhexane

$$CH_3CH_2CH_2\underset{\underset{CH_3}{|}}{CH}CH_2CH_3$$

3-methylhexane is correct.

d. 1-ethyl-1-methylbutane

$$\underset{\underset{CH_3}{|}}{\underset{|}{CH}}CH_2CH_2CH_3$$ (CH_2CH_3)
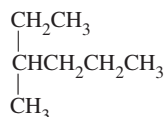
3-methylhexane is correct.

f. 4-ethylpentane

$$CH_3CH_2CH_2\underset{\underset{CH_2CH_3}{|}}{CH}CH_3$$

3-methylhexane is correct.

All six of these are the same compound. They only differ from each other by rotations about one or more carbon-carbon single bonds. Only one isomer of C_7H_{16} is present in all of these names, 3-methylhexane.

3. a.

$$CH_3\underset{\underset{CH_2CH_3}{|}}{CH}CH_3$$

The longest chain is 4 carbons long. The correct name is 2-methylbutane.

b.

$$CH_3CH_2CH_2CH_2\underset{\underset{CH_3}{|}}{\overset{\overset{I}{|}}{C}}CH_2$$ (CH_3)
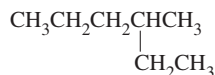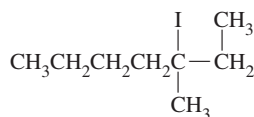
The longest chain is 7 carbons long and we would start the numbering system at the other end for lowest possible numbers. The correct name is 3-iodo-3-methylheptane.

c.

$$CH_3CH_2CH=\underset{\underset{CH_3}{|}}{C}—CH_3$$
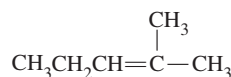
This compound cannot exhibit *cis– trans* isomerism since one of the double bonded carbons has the same two groups (CH_3) attached. The numbering system should also start at the other end to give the double bond the lowest possible number. 2-methyl-2-pentene is correct.

d.

$$CH_3\underset{\underset{Br}{|}}{CH}\underset{\underset{OH}{|}}{CH}CH_3$$

The OH functional group gets the lowest number. 3-bromo-2-butanol is correct.

5. Hydrocarbons are nonpolar substances exhibiting only London dispersion forces. Size and shape are the two most important structural features relating to the strength of London dispersion forces. For size, the bigger the molecule (the larger the molar mass), the stronger the London dispersion forces and the higher the boiling point. For shape, the more branching present in a compound, the weaker the London dispersion forces and the lower the boiling point.

7. The correct order of strength is polyhydrocarbon $<$ polyester $<$ polyamide. The difference in strength is related to the types of intermolecular forces present. All of these polymers have London dispersion forces. However, polyhydrocarbons only have London dispersion forces. The polar ester group in polyesters and the polar amide group in polyamides give rise to additional dipole forces. The polyamide has the ability to form relatively strong hydrogen bonding interactions, hence why it would form the strongest fibers.

9. a.

$$CH_2=CH_2 + H_2O \xrightarrow{H^+} CH_2—CH_2 \quad 1° \text{ alcohol}$$ (OH, H)

b.

$$CH_2CH=CH_2 + H_2O \xrightarrow{H^+} CH_2CH—CH_2 \quad 2° \text{ alcohol}$$
major product

c.

$$CH_3\underset{\underset{CH_3}{|}}{C}=CH_2 + H_2O \xrightarrow{H^+} CH_3\underset{\underset{CH_3}{|}}{C}—CH_2 \quad 3° \text{ alcohol}$$ (OH, H)
major product

d.

$$CH_3CH_2OH \xrightarrow{oxidation} CH_3\overset{\overset{O}{||}}{C}H \quad \text{aldehyde}$$

e.

$$CH_3\underset{\underset{OH}{|}}{CH}CH_3 \xrightarrow{oxidation} CH_3\overset{\overset{O}{||}}{C}CH_3 \quad \text{ketone}$$

f.

$$CH_3CH_2CH_2OH \xrightarrow{oxidation} CH_3CH_2\overset{\overset{O}{||}}{C}—OH \quad \text{carboxylic acid}$$
or
$$CH_3CH_2\overset{\overset{O}{||}}{C}H \xrightarrow{oxidation} CH_3CH_2\overset{\overset{O}{||}}{C}—OH$$

g.

$$CH_3OH + HO\overset{\overset{O}{||}}{C}CH_3 \xrightarrow{H^+} CH_3—O—\overset{\overset{O}{||}}{C}CH_3 + H_2O \quad \text{ester}$$
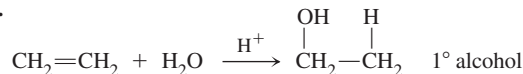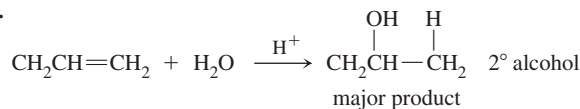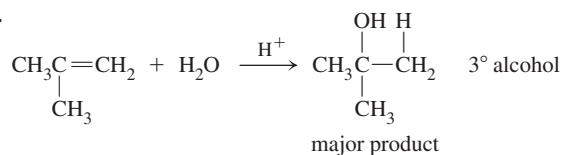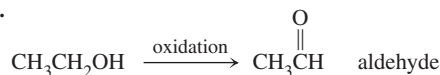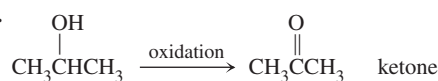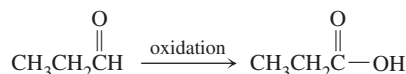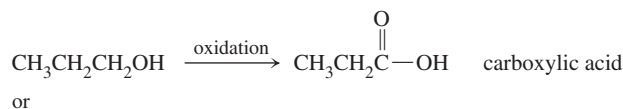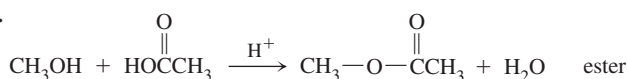
11. a. A polyester forms when an alcohol functional group reacts with a carboxylic acid functional group. The monomer for a homopolymer polyester must have an alcohol functional group and a carboxylic acid functional group present

in the structure.   **b.** A polyamide forms when an amine functional group re-acts with a carboxylic acid functional group. For a copolymer polyamide, one monomer would have at least two amine functional groups present and the other monomer would have at least two carboxylic acid functional groups present. For polymerization to occur, each monomer must have two reactive functional groups present.   **c.** To form an addition polymer, a carbon-carbon double bond must be present. To form a polyester, the monomer would need the alcohol and carboxylic acid functional groups present. To form a polyamide, the monomer would need the amine and carboxylic acid functional groups present. The two possibilities are for the monomer to have a carbon-carbon double bond, an alcohol functional group, and a carboxylic acid functional group all present, or to have a carbon-carbon double bond, an amine functional group, and a carboxylic acid functional group present.

13. $CH_3-CH_2-CH_2-CH_2-CH_2-CH_3$

15. a.

2-methylheptane

3-methylheptane

4-methylheptane

b.

2,2,3,3-tetramethylbutane

17. a.

b.

c.

d.

19. a. $CH_3CH_2-CH-CH_2CH_2CH_3$

b.

c.

d. 4-ethyl-2-methylheptane; 2,2,3-trimethylhexane
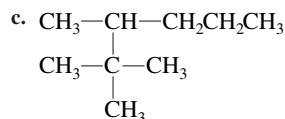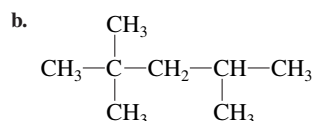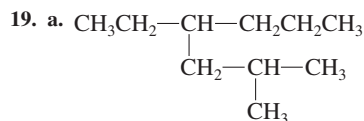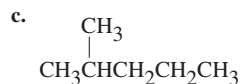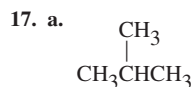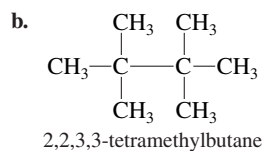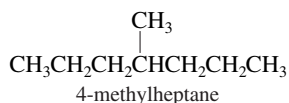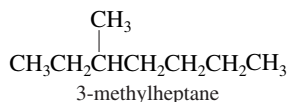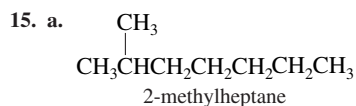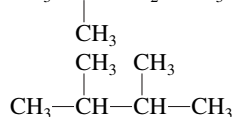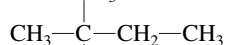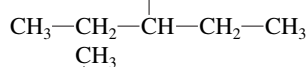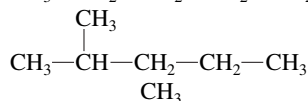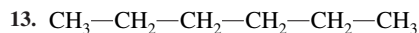
21. a. 2,2,4-trimethylhexane: **b.** 5-methylnonane; **c,** 2,2,4,4-tetra-methylpentane; **d.** 3-ethyl-3-methyloctane
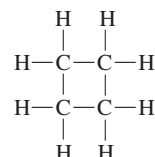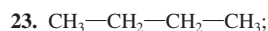
23. $CH_3-CH_2-CH_2-CH_3$;

Each carbon is bonded to four other atoms.

25. a. 1-butene; **b.** 4-methyl-2-hexene; **c.** 2,5-dimethyl-3-heptene

27. a. $CH_3CH_2CH=CHCH_2CH_3$; **b.** $CH_3CH=CHCH=CHCH_2CH_3$;

c. $CH_3CHCH=CHCH_2CH_2CH_2CH_3$

29. a.

b.

c.

d. $CH_2CH=CHCH_3$
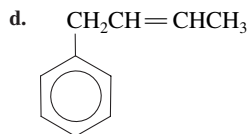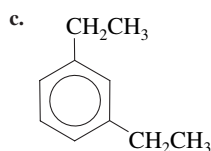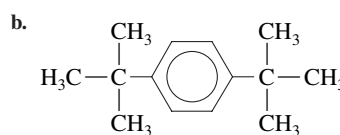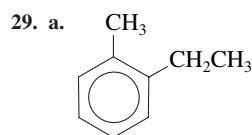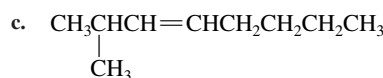
31. a. 1,3-dichlorobutane; **b.** 1,1,1-trichlorobutane; **c.** 2,3-dichloro-2,4-dimethylhexane; **d.** 1,2-difluoroethane;   **33.** CH_2Cl-CH_2Cl,1-2-dichloroethane: There is free rotation about the C—C single bond that doesn't lead to different compounds. CHCl=CHCl, 1-2-dichloroethene: There is no rotation about the C=C double bond. This creates the *cis* and *trans* isomers, which are different compounds.   **35.** [25], compounds b and c; [27], all compounds

37. $CH_2=CHCH_2CH_2CH_3$ $CH_3CH=CHCH_2CH_3$

39.

41.

43. a. (structure: CH$_3$/H on C=C with CH$_2$CH$_2$CH$_3$/H) **b.** (structure: CH$_3$/H on C=C with H/CH$_3$)

c. (structure: CH$_3$/Cl on C=C with CH$_2$CH$_3$/Cl)
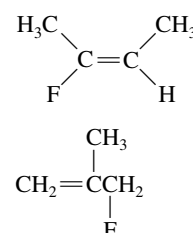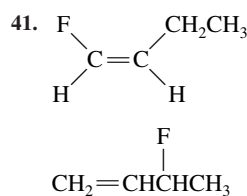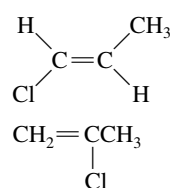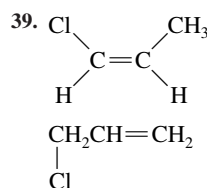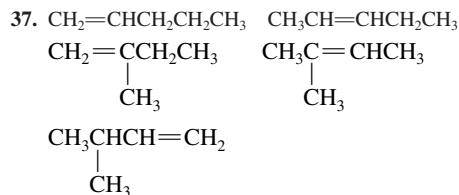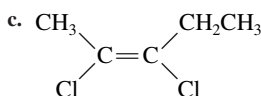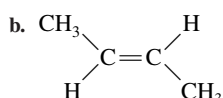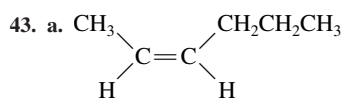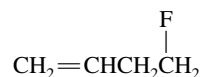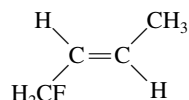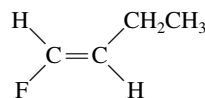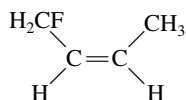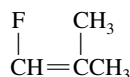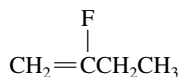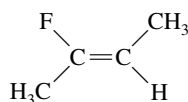
45. a. 3 monochloro isomers of *n*-pentane; **b.** 4 monochloro isomers of 2-methylbutane; **c.** 3 monochloro isomers of 2,4-dimethylpentane; **d.** 4 monochloro isomers of methylcyclobutane **47. a.** ketone; **b.** aldehyde; **c.** carboxylic acid; **d.** amine

49. a. (structure with labels: amine, ketone, alcohol, carboxylic acid, amine)
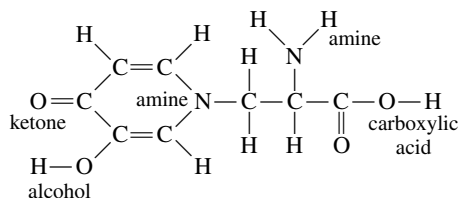
b. 5 carbons in ring and the carbon in —CO$_2$H: sp^2; the other two carbons: sp^3; **c.** 24 sigma bonds, 4 pi bonds **51. a.** 3-chloro-1-butanol, primary: **b.** 3-methyl-3-hexanol,   tertiary; **c.** 2-methylcyclopentanol,   secondary **53.** 1-pentanol; 2-pentanol; 3-pentanol; 2-methyl-1-butanol; 2-methyl-2-butanol; 3-methyl-2-butanol; 3-methyl-1-butanol; 2,2-dimethyl-1-propanol; 6   ethers **55. a.** 4,5-dichloro-3-hexanone; **b.** 2,3-dimethylpentanal; **c.** 3-methylbenzaldehyde or *m*-methylbenzaldehyde **57. a.** 4-chlorobenzoic acid or *p*-chlorobenzoic acid; **b.** 3-ethyl-2-methylhexanoic acid; **c.** methanoic acid (common name = formic acid) **59.** Only statement d is false.

2-butenal: $\overset{\overset{\displaystyle O}{\|}}{H C} CH{=}CHCH_3$.
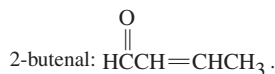
The formula of 2-butenal is C$_4$H$_6$O, while the ether has a formula of C$_4$H$_8$O.

61. a. CH$_3$CH$_2$CH$_2$CH$_3$

b. (structure: CH$_2$CHCHCHCH with Cl Cl / Cl Cl and CH$_3$ CH$_3$) **c.** (chlorobenzene structure) + HCl
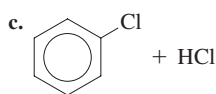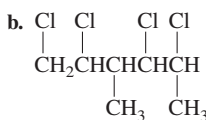
d. C$_4$H$_8$(g) + 6O$_2$(g) $\longrightarrow$ 4CO$_2$(g) + 4H$_2$O(g)

63. For the iron-catalyzed reaction, one of the *ortho* or *para* hydrogens in benzene is replaced by chlorine. When an iron catalyst is not present, then the benzene hydrogens are unreactive, which is seen for the light-catalyzed reaction where one of the methyl hydrogens is replaced by chlorine.

65. a. $H{-}\overset{\overset{\displaystyle O}{\|}}{C}{-}CH_2CHCH_3$ + $HO{-}\overset{\overset{\displaystyle O}{\|}}{C}{-}CH_2CHCH_3$ (both with CH$_3$ branch)

b. $CH_3{-}\overset{\overset{\displaystyle O}{\|}}{C}{-}CHCH_3$ (with CH$_3$ branch)

c. No reaction

d. (benzaldehyde structure) + (benzoic acid structure)

e. (2-methylcyclohexanone structure)

f. (two structures with OH, CH$_3$, and C—H / C—OH groups)
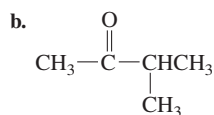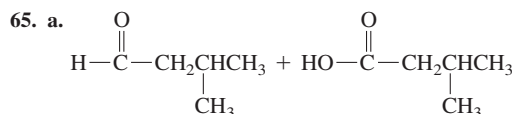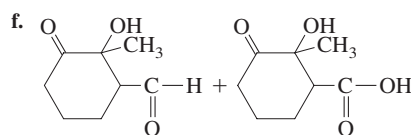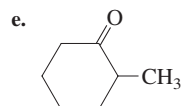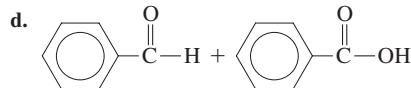
67. a. CH$_3$CH=CH$_2$ + Br$_2$ → CH$_3$CHBrCH$_2$Br;

b. $CH_3{-}\overset{\overset{\displaystyle OH}{|}}{CH}{-}CH_3$ $\xrightarrow{\text{oxidation}}$ $CH_3{-}\overset{\overset{\displaystyle O}{\|}}{C}{-}CH_3$;

c. $CH_2{=}\overset{\overset{\displaystyle CH_3}{|}}{C}{-}CH_3$ + H$_2$O $\xrightarrow{H^+}$ $CH_2{-}\overset{\overset{\displaystyle CH_3}{|}}{\underset{\underset{\displaystyle OH}{|}}{C}}{-}CH_3$ (with H);

d. CH$_3$CH$_2$CH$_2$OH $\xrightarrow{KMnO_4}$ $CH_3CH_2\overset{\overset{\displaystyle O}{\|}}{C}{-}OH$

69. a. $CH_3\overset{\overset{\displaystyle O}{\|}}{C}{-}OH$ + HOCH$_2$(CH$_2$)$_6$CH$_3$ $\longrightarrow$

$CH_3\overset{\overset{\displaystyle O}{\|}}{C}{-}O{-}CH_2(CH_2)_6CH_3$ + H$_2$O;

b. $CH_3CH_2\overset{\overset{\displaystyle O}{\|}}{C}{-}OH$ + HOCH$_2$(CH$_2$)$_4$CH$_3$ $\longrightarrow$

$CH_3CH_2\overset{\overset{\displaystyle O}{\|}}{C}{-}O{-}CH_2(CH_2)_4CH_3$ + H$_2$O

71. CFCl=CF$_2$

73. (polymer structure with CN, C—OCH$_3$, O groups)

75. (polymer structure with HN, NHC, O groups)

77.

H₂N—⟨benzene⟩—NH₂   and   HO₂C—⟨benzene⟩(CO₂H, HO₂C, CO₂H)
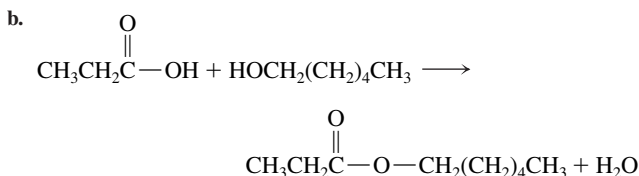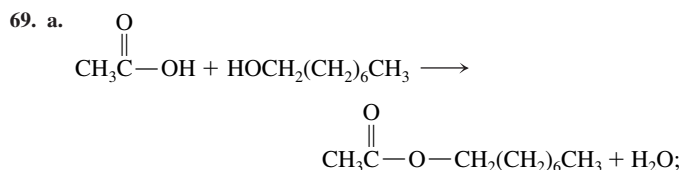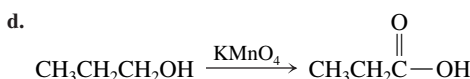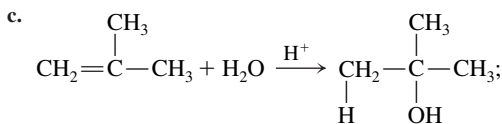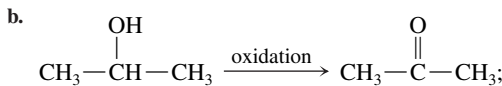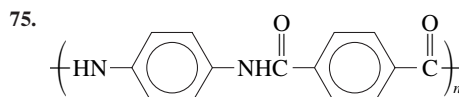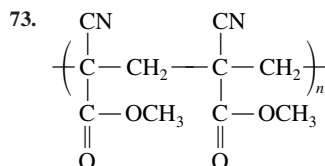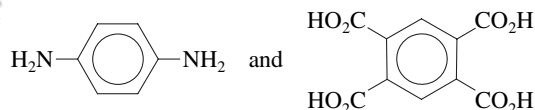
79. Divinylbenzene inserts itself into two adjacent polymer chains and bonds them together. The chains cannot move past each other because of the crosslinks making the polymer more rigid.   **81. a.** The polymer from 1,2-diaminoethane and terephthalic acid is stronger because of the possibility of hydrogen bonding between chains.   **b.** The polymer of

HO—⟨benzene⟩—CO₂H
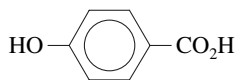
is more rigid because the chains are stiffer due to the rigid benzene rings in the chains.   **c.** Polyacetylene is $nHC\equiv CH \rightarrow -(CH=CH)_n$. Polyacetylene is more rigid because the double bonds in the chains make the chains stiffer.   **83. a.** serine; tyrosine; threonine;   **b.** aspartic acid; glutamic acid;   **c.** histidine; lysine; arginine; tryptophan;   **d.** glutamine; asparagine   **85. a.** aspartic acid and phenylalanine.   **b.** Aspartame contains the methyl ester of phenylalanine. This ester can hydrolyze to form methanol, $R-CO_2CH_3 + H_2O \rightleftharpoons RCO_2H + CH_3OH$.

87.

ser–ala   ala–ser
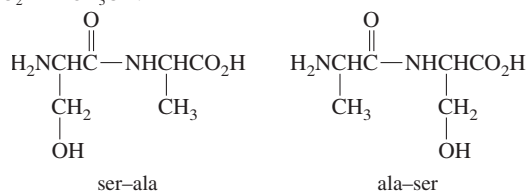
89. a. Six tetrapeptides are possible. From NH₂ to CO₂H end: phe–phe–gly–gly, gly–gly–phe–phe, gly–phe–phe–gly, phe–gly–gly–phe, phe–gly–phe–gly, gly–phe–gly–phe   **b.** Twelve tetrapeptides are possible. From NH₂ to CO₂H end: phe–phe–gly–ala, phe–phe–ala–gly, phe–gly–phe–ala, phe–gly–ala–phe, phe–ala–phe–gly, phe–ala–gly–phe, gly–phe–phe–ala, gly–phe–ala–phe, gly–ala–phe–phe, ala–phe–phe–gly, ala–phe–gly–phe, ala–gly–phe–phe   **91.** Ionic: his, lys, or arg with asp or glu; hydrogen bonding: ser, glu, tyr, his, arg, asn, thr, asp, gln, or lys with any amino acid; covalent: cys with cys; London dispersion: all amino acids with nonpolar R groups (gly, ala, pro, phe, ile, trp, met, leu, val); dipole–dipole: tyr, thr, and ser with each other   **93.** Glutamic acid has a polar R group and valine has a nonpolar R group. The change in polarity of the R groups could affect the tertiary structure of hemoglobin and affect the ability of hemoglobin to bond to oxygen.

95.
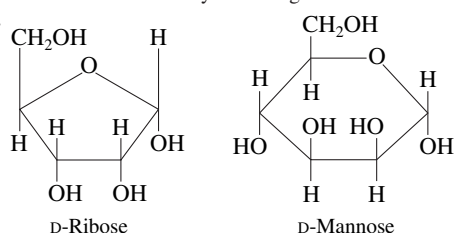
D-Ribose   D-Mannose

97. aldohexose: glucose, mannose, galactose; aldopentose: ribose, arabinose; ketohexose: fructose; ketopentose: ribulose   **99.** They differ in the orientation of a hydroxy group on a particular carbon. Starch is composed from α-D-glucose, and cellulose is composed from β-D-glucose.

101. The chiral carbons are marked with asterisks.

Isoleucine   Threonine

103.

Br—C*—CH=CH₂ (Cl, H)
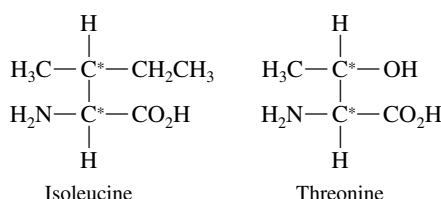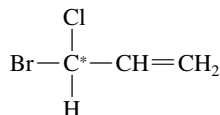
is optically active. The chiral carbon is marked with an asterisk.   **105.** C–C–A–G–A–T–A–T–G   **107.** Uracil will H-bond to adenine.

Uracil   Adenine
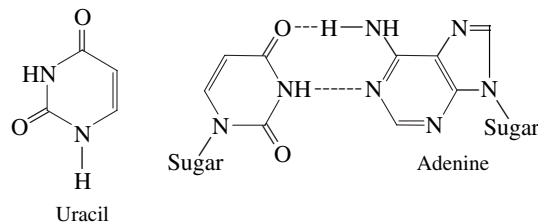
109. a. glu: CTT, CTC; val: CAA, CAG, CAT, CAC; met: TAC; trp: ACC; phe: AAA, AAG; asp: CTA, CTG;   **b.** ACC–CTT–AAA–TAC or ACC–CTC–AAA–TAC or ACC–CTT–AAG–TAC or ACC–CTC–AAG– TAC;   **c.** four (see answer in part b);   **d.** met–asp–phe   **e.** TAC–CTA–AAG; TAC–CTA–AAA; TAC–CTG–AAA   **111. a.** 2,3,5,6-tetramethyloctane;   **b.** 2,2,3,5-tetramethylheptane;   **c.** 2,3,4-trimethylhexane;   **d.** 3-methyl-1-pentyne

113.
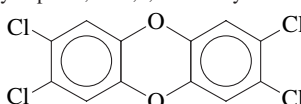
There are many possibilities for isomers. Any structure with four chlorines replacing four hydrogens in any of the numbered positions would be an isomer; i.e., 1,2,3,4-tetrachloro-dibenzo-*p*-dioxin is a possible isomer.   **115.** $-23°C$: CH_3-O-CH_3; 78.5°C: CH_3-CH_2-OH **117.** Alcohols consist of two parts, the polar OH group and the nonpolar hydrocarbon chain attached to the OH group. As the length of the nonpolar hydrocarbon chain increases, the solubility of the alcohol decreases. In methyl alcohol (methanol), the polar OH group can override the effect of the nonpolar CH_3 group, and methyl alcohol is soluble in water. In stearyl alcohol, the molecule consists mostly of the long nonpolar hydrocarbon chain, so it is insoluble in water.   **119.** *n*-hexane, 69°C; pentanal, 103°C; 1-pentanol, 137°C; butanoic acid, 164°C.   **121.** 1-butene   **123.** ethanoic acid   **125.** In nylon, hydrogen-bonding interactions occur due to the presence of N—H bonds in the polymer. For a given polymer chain length, there are more N—H groups in nylon-46 as compared to nylon-6. Hence, nylon-46 forms a stronger polymer compared to nylon-6 due to the increased hydrogen-bonding interactions.

127. a.

H₂N—⟨benzene⟩—NH₂   and   HO₂C—⟨benzene⟩—CO₂H
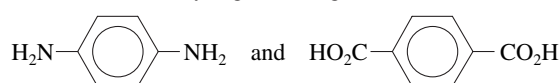
b. Repeating unit:
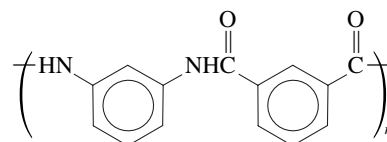
The two polymers differ in the substitution pattern on the benzene rings. The Kevlar chain is straighter, and there is more efficient hydrogen bonding between Kevlar chains than between Nomex chains.   **129. a.** The bond angles in the ring are about 60°. VSEPR predicts bond angles close to 109°. The bonding electrons are much closer together than they prefer, resulting in strong electron–electron repulsions. Thus ethylene oxide is unstable (reactive).   **b.** The ring opens up during polymerization and the monomers link together through the formation of O—C bonds.

$-(O-CH_2CH_2-O-CH_2CH_2-O-CH_2CH_2)_n$

131.

H₂N—CH—CO₂H   or   H₂N—CH—CO₂⁻Na⁺
         | |
   CH₂CH₂CO₂⁻Na⁺   CH₂CH₂CO₂H
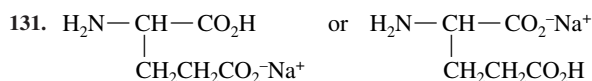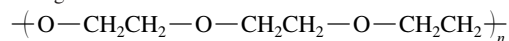
The first structure is MSG, which is impossible for you to predict.   **133.** In the reaction, P—O and O—H bonds are broken and P—O and O—H bonds are formed. Thus $\Delta H \approx 0$. $\Delta S < 0$, since two molecules are going to form one molecule. Thus $\Delta G > 0$, not spontaneous.

135.

$$1.0\ M\ \text{H}^+:\quad \overset{+}{\text{H}_3\text{N}}-\underset{\underset{\text{CH}_3}{|}}{\text{CH}}-\overset{\overset{\text{O}}{\|}}{\text{C}}-\text{OH};$$

$$1.0\ M\ \text{OH}^-:\quad \text{H}_2\text{N}-\underset{\underset{\text{CH}_3}{|}}{\text{CH}}-\overset{\overset{\text{O}}{\|}}{\text{C}}-\text{O}^-$$
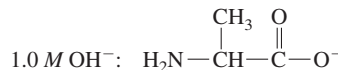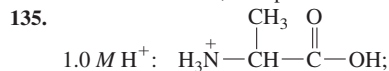
137. Both ΔH and ΔS are positive values.   **139.** 6.07   **141. a.** No; the mirror image is superimposable.

b.

143.

$$\text{HC}\equiv\text{C}-\text{C}\equiv\text{C}-\text{HC}=\text{C}=\text{CH}-\text{HC}=\text{CH}-\text{HC}=\text{CH}-\text{CH}_2-\overset{\overset{\text{O}}{\|}}{\text{C}}-\text{OH}$$

13 12 11 10   9   8   7   6   5   4   3   2   1

145. a.
$$\underset{\underset{\text{CH}_3}{|}}{\text{CH}_3\text{CHCH}_2\text{CH}_3}$$

b.
$$\underset{\underset{\text{CH}_3}{|}}{\text{CH}_2{=}\text{CCH}_3}$$

c.

d. $\text{CH}_2{=}\text{CHCH}_3$

e. $\underset{\underset{\text{OH}}{|}}{\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3}$ $\underset{\underset{\text{OH}}{|}}{\overset{\overset{\text{CH}_3}{|}}{\text{CH}_2\text{CHCH}_2\text{CH}_3}}$
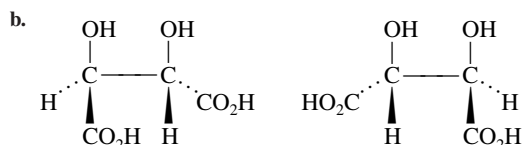
$\underset{\underset{\text{OH}}{|}}{\overset{\overset{\text{CH}_3}{|}}{\text{CH}_3\text{CHCH}_2\text{CH}_2}}$ $\underset{\underset{\text{OH}\ \ \ \text{CH}_3}{|\ \ \ \ \ |}}{\overset{\overset{\text{CH}_3}{|}}{\text{CH}_2{-}\text{C}{-}\text{CH}_3}}$
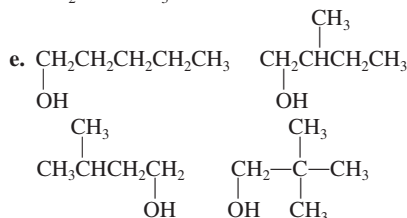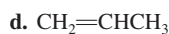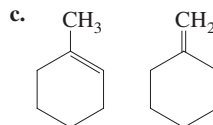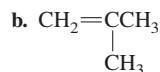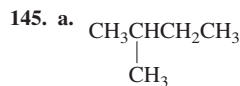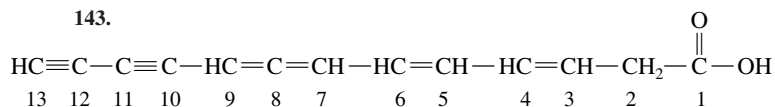
147.
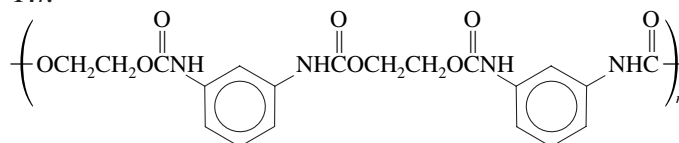
149. a. The temperature of the rubber band increases when it is stretched; **b.** exothermic (heat is released); **c.** As the chains are stretched, they line up more closely resulting in stronger dispersion forces between the chains. Heat is released as the strength of the intermolecular forces increases. **d.** ΔG is positive and ΔS is negative; **e.** The structure of the stretched polymer chains is more ordered than in unstretched rubber. Disorder decreases as the rubber band is stretched.   **151.** 0.11%   **153.** 5   **155. a.** 37.50%; **b.** The hybridization changes from sp^2 to sp^3; **c.** 3,4-dimethyl-3-hexanol

Index

FÈUE WHD

Periodic Table of the Elements

Noble gases

Alkaline earth metals

Halogens 18
8A

1 1A	2 2A											13 3A	14 4A	15 5A	16 6A	17 7A	18 8A
1 **H** 1.008																	2 **He** 4.003
3 **Li** 6.941	4 **Be** 9.012											5 **B** 10.81	6 **C** 12.01	7 **N** 14.01	8 **O** 16.00	9 **F** 19.00	10 **Ne** 20.18
11 **Na** 22.99	12 **Mg** 24.31	3	4	5	6	7	8	9	10	11	12	13 **Al** 26.98	14 **Si** 28.09	15 **P** 30.97	16 **S** 32.07	17 **Cl** 35.45	18 **Ar** 39.95
19 **K** 39.10	20 **Ca** 40.08	21 **Sc** 44.96	22 **Ti** 47.88	23 **V** 50.94	24 **Cr** 52.00	25 **Mn** 54.94	26 **Fe** 55.85	27 **Co** 58.93	28 **Ni** 58.69	29 **Cu** 63.55	30 **Zn** 65.38	31 **Ga** 69.72	32 **Ge** 72.59	33 **As** 74.92	34 **Se** 78.96	35 **Br** 79.90	36 **Kr** 83.80
37 **Rb** 85.47	38 **Sr** 87.62	39 **Y** 88.91	40 **Zr** 91.22	41 **Nb** 92.91	42 **Mo** 95.94	43 **Tc** (98)	44 **Ru** 101.1	45 **Rh** 102.9	46 **Pd** 106.4	47 **Ag** 107.9	48 **Cd** 112.4	49 **In** 114.8	50 **Sn** 118.7	51 **Sb** 121.8	52 **Te** 127.6	53 **I** 126.9	54 **Xe** 131.3
55 **Cs** 132.9	56 **Ba** 137.3	57 **La*** 138.9	72 **Hf** 178.5	73 **Ta** 180.9	74 **W** 183.9	75 **Re** 186.2	76 **Os** 190.2	77 **Ir** 192.2	78 **Pt** 195.1	79 **Au** 197.0	80 **Hg** 200.6	81 **Tl** 204.4	82 **Pb** 207.2	83 **Bi** 209.0	84 **Po** (209)	85 **At** (210)	86 **Rn** (222)
87 **Fr** (223)	88 **Ra** 226	89 **Ac**† (227)	104 **Rf** (261)	105 **Db** (262)	106 **Sg** (263)	107 **Bh** (264)	108 **Hs** (265)	109 **Mt** (268)	110 **Ds** (271)	111 **Rg** (272)	112 **Uub**	113 **Uut**	114 **Uuq**	115 **Uup**			

Transition metals

Alkali metals

metals ← → nonmetals

	58 **Ce** 140.1	59 **Pr** 140.9	60 **Nd** 144.2	61 **Pm** (145)	62 **Sm** 150.4	63 **Eu** 152.0	64 **Gd** 157.3	65 **Tb** 158.9	66 **Dy** 162.5	67 **Ho** 164.9	68 **Er** 167.3	69 **Tm** 168.9	70 **Yb** 173.0	71 **Lu** 175.0
*Lanthanides														
†Actinides	90 **Th** 232.0	91 **Pa** (231)	92 **U** 238.0	93 **Np** (237)	94 **Pu** (244)	95 **Am** (243)	96 **Cm** (247)	97 **Bk** (247)	98 **Cf** (251)	99 **Es** (252)	100 **Fm** (257)	101 **Md** (258)	102 **No** (259)	103 **Lr** (260)

Group numbers 1–18 represent the system recommended by the International Union of Pure and Applied Chemistry.

Table of Atomic Masses*

Element	Symbol	Atomic Number	Atomic Mass	Element	Symbol	Atomic Number	Atomic Mass	Element	Symbol	Atomic Number	Atomic Mass
Actinium	Ac	89	[227]§	Gold	Au	79	197.0	Praseodymium	Pr	59	140.9
Aluminum	Al	13	26.98	Hafnium	Hf	72	178.5	Promethium	Pm	61	[145]
Americium	Am	95	[243]	Hassium	Hs	108	[265]	Protactinium	Pa	91	[231]
Antimony	Sb	51	121.8	Helium	He	2	4.003	Radium	Ra	88	226
Argon	Ar	18	39.95	Holmium	Ho	67	164.9	Radon	Rn	86	[222]
Arsenic	As	33	74.92	Hydrogen	H	1	1.008	Rhenium	Re	75	186.2
Astatine	At	85	[210]	Indium	In	49	114.8	Rhodium	Rh	45	102.9
Barium	Ba	56	137.3	Iodine	I	53	126.9	Roentgenium	Rg	111	[272]
Berkelium	Bk	97	[247]	Iridium	Ir	77	192.2	Rubidium	Rb	37	85.47
Beryllium	Be	4	9.012	Iron	Fe	26	55.85	Ruthenium	Ru	44	101.1
Bismuth	Bi	83	209.0	Krypton	Kr	36	83.80	Rutherfordium	Rf	104	[261]
Bohrium	Bh	107	[264]	Lanthanum	La	57	138.9	Samarium	Sm	62	150.4
Boron	B	5	10.81	Lawrencium	Lr	103	[260]	Scandium	Sc	21	44.96
Bromine	Br	35	79.90	Lead	Pb	82	207.2	Seaborgium	Sg	106	[263]
Cadmium	Cd	48	112.4	Lithium	Li	3	6.9419	Selenium	Se	34	78.96
Calcium	Ca	20	40.08	Lutetium	Lu	71	175.0	Silicon	Si	14	28.09
Californium	Cf	98	[251]	Magnesium	Mg	12	24.31	Silver	Ag	47	107.9
Carbon	C	6	12.01	Manganese	Mn	25	54.94	Sodium	Na	11	22.99
Cerium	Ce	58	140.1	Meitnerium	Mt	109	[268]	Strontium	Sr	38	87.62
Cesium	Cs	55	132.90	Mendelevium	Md	101	[258]	Sulfur	S	16	32.07
Chlorine	Cl	17	35.45	Mercury	Hg	80	200.6	Tantalum	Ta	73	180.9
Chromium	Cr	24	52.00	Molybdenum	Mo	42	95.94	Technetium	Tc	43	[98]
Cobalt	Co	27	58.93	Neodymium	Nd	60	144.2	Tellurium	Te	52	127.6
Copper	Cu	29	63.55	Neon	Ne	10	20.18	Terbium	Tb	65	158.9
Curium	Cm	96	[247]	Neptunium	Np	93	[237]	Thallium	Tl	81	204.4
Darmstadtium	Ds	110	[271]	Nickel	Ni	28	58.69	Thorium	Th	90	232.0
Dubnium	Db	105	[262]	Niobium	Nb	41	92.91	Thulium	Tm	69	168.9
Dysprosium	Dy	66	162.5	Nitrogen	N	7	14.01	Tin	Sn	50	118.7
Einsteinium	Es	99	[252]	Nobelium	No	102	[259]	Titanium	Ti	22	47.88
Erbium	Er	68	167.3	Osmium	Os	76	190.2	Tungsten	W	74	183.9
Europium	Eu	63	152.0	Oxygen	O	8	16.00	Uranium	U	92	238.0
Fermium	Fm	100	[257]	Palladium	Pd	46	106.4	Vanadium	V	23	50.94
Fluorine	F	9	19.00	Phosphorus	P	15	30.97	Xenon	Xe	54	131.3
Francium	Fr	87	[223]	Platinum	Pt	78	195.1	Ytterbium	Yb	70	173.0
Gadolinium	Gd	64	157.3	Plutonium	Pu	94	[244]	Yttrium	Y	39	88.91
Gallium	Ga	31	69.72	Polonium	Po	84	[209]	Zinc	Zn	30	65.38
Germanium	Ge	32	72.59	Potassium	K	19	39.10	Zirconium	Zr	40	91.22

*The values given here are to four significant figures where possible.   §A value given in parentheses denotes the mass of the longest-lived isotope.

Page Numbers of Some Important Tables

Physical Constants

Constant	Symbol	Value
Atomic mass unit	amu	1.66054×10^{-27} kg
Avogadro's number	N	6.02214×10^{23} mol^{-1}
Bohr radius	a_0	5.292×10^{-11} m
Boltzmann constant	k	1.38066×10^{-23} J/K
Charge of an electron	e	1.60218×10^{-19} C
Faraday constant	F	96,485 C/mol
Gas constant	R	8.31451 J/K $\cdot$ mol
		0.08206 L $\cdot$ atm/K $\cdot$ mol
Mass of an electron	m_e	9.10939×10^{-31} kg
		5.48580×10^{-4} amu
Mass of a neutron	m_n	1.67493×10^{-27} kg
		1.00866 amu
Mass of a proton	m_p	1.67262×10^{-27} kg
		1.00728 amu
Planck's constant	h	6.62608×10^{-34} J $\cdot$ s
Speed of light	c	2.99792458×10^8 m/s

General Biology

Wikibooks.org

March 15, 2013

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1 Getting Started

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2 Biology - The Life Science

The word ***biology*** means, "the science of life", from the Greek **bios,** *life*, and **logos,** *word* or *knowledge.* Therefore, Biology is the science of Living Things. That is why Biology is sometimes known as Life Science.

The science has been divided into many subdisciplines, such as botany[1], bacteriology, anatomy[2], zoology, histology, mycology, embryology, parasitology, genetics[3], molecular biology[4], systematics, immunology, microbiology[5], physiology, cell biology[6], cytology, ecology[7], and virology. Other branches of science include or are comprised in part of biology studies, including paleontology[8], taxonomy, evolution, phycology, helimentology, protozoology, entomology, biochemistry, biophysics, biomathematics, bio engineering, bio climatology and anthropology.

2.1 Characteristics of life

Not all scientists agree on the definition of just what makes up life. Various characteristics describe most living things. However, with most of the characteristics listed below we can think of one or more examples that would seem to break the rule, with something nonliving being classified as living or something living classified as nonliving. Therefore we are careful not to be too dogmatic in our attempt to explain which things are living or nonliving.

- Living things are composed of **matter structured in an orderly way** where simple molecules are ordered together into much larger macromolecules.

An easy way to remember this is GRIMNERD C All organisms; - **G**row, **R**espire, **I**nteract, **M**ove, Need **N**utrients, **E**xcrete (Waste), **R**eproduce, **D**eath, **C**ells (Made of)

- Living things are **sensitive,** meaning they are able to **respond to stimuli.**

- Living things are able to **grow**, **develop**, and **reproduce**.

- Living things are able to **adapt** over time by the process of **natural selection**.

- All known living things use the **hereditary molecule, DNA**[9].

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1   http://en.wikibooks.org/wiki/botany
2   http://en.wikibooks.org/wiki/anatomy
3   http://en.wikibooks.org/wiki/genetics
4   http://en.wikibooks.org/wiki/Molecular%20Biology
5   http://en.wikibooks.org/wiki/microbiology
6   http://en.wikibooks.org/wiki/Cell%20Biology
7   http://en.wikibooks.org/wiki/ecology
8   http://en.wikibooks.org/wiki/paleontology
9   http://en.wikipedia.org/wiki/DNA

- Internal functions are coordinated and **regulated** so that the internal environment of a living thing is relatively constant, referred to as **homeostasis[10].**

Living things are organized in the microscopic level from atoms up to cells[11]. Atoms are arranged into molecules, then into macromolecules[12], which make up organelles[13], which work together to form cells. Beyond this, cells are organized in higher levels to form entire multicellular organisms. Cells together form tissues[14], which make up organs, which are part of organ systems, which work together to form an entire organism. Of course, beyond this, organisms form populations which make up parts of an ecosystem. All of the Earth's ecosystems together form the diverse environment that is the earth.

Example:-

sub atoms, atoms, molecules, cells, tissues, organs, organ systems, organisms, population, community, eco systems

2.2 Nature of science

Science is a **methodology** for **learning about the world**. It involves the **application of knowledge**.

The scientific method deals with **systematic investigation**, **reproducible results**, the formation and testing of **hypotheses**, and **reasoning.**

Reasoning can be broken down into two categories, **induction** (specific data is used to develop a generalized observation or conclusion) and **deduction** (general information leads to specific conclusion). Most reasoning in science is done through induction.

Science as we now know it arose as a discipline in the 17th century.

2.3 Scientific method

The scientific method is not a step by step, linear process. It is an intuitive process, a methodology for learning about the world through the application of knowledge. Scientists must be able to have an "imaginative preconception" of what the truth is. Scientists will often observe and then hypothesize the reason why a phenomenon occurred. They use all of their knowledge and a bit of imagination, all in an attempt to uncover something that might be true. A typical scientific investigation might go like so:

You *observe* that a room appears dark, and you ponder *why* the room is dark. In an attempt to find explanations to this curiosity, your mind unravels several different *hypotheses*. One hypothesis might state that the lights are turned off. Another hunch might be that the room's lightbulb has burnt out. Worst yet, you could be going blind. To discover the truth,

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10   http://en.wikipedia.org/wiki/homeostasis
11   http://en.wikipedia.org/wiki/cell
12   http://en.wikipedia.org/wiki/macromolecule
13   http://en.wikipedia.org/wiki/organelle
14   http://en.wikibooks.org/wiki/General%20Biology%2FTissues

6

you *experiment*. You feel your way around the room and find a light switch and turn it on. No light. You *repeat* the experiment, flicking the switch back and forth. Still nothing. That means your initial hypothesis, the room is dark because the lights are off, has been *rejected*. You devise more experiments to test your hypotheses, utilizing a flashlight to prove that you are indeed not blind. In order to *accept* your last remaining hypothesis as the truth, you could *predict* that changing the light bulb will fix the problem. If all your predictions succeed, the original hypothesis is valid and is accepted. In some cases, however, your predictions will not occur, in which you'll have to start over. Perhaps the power is off.
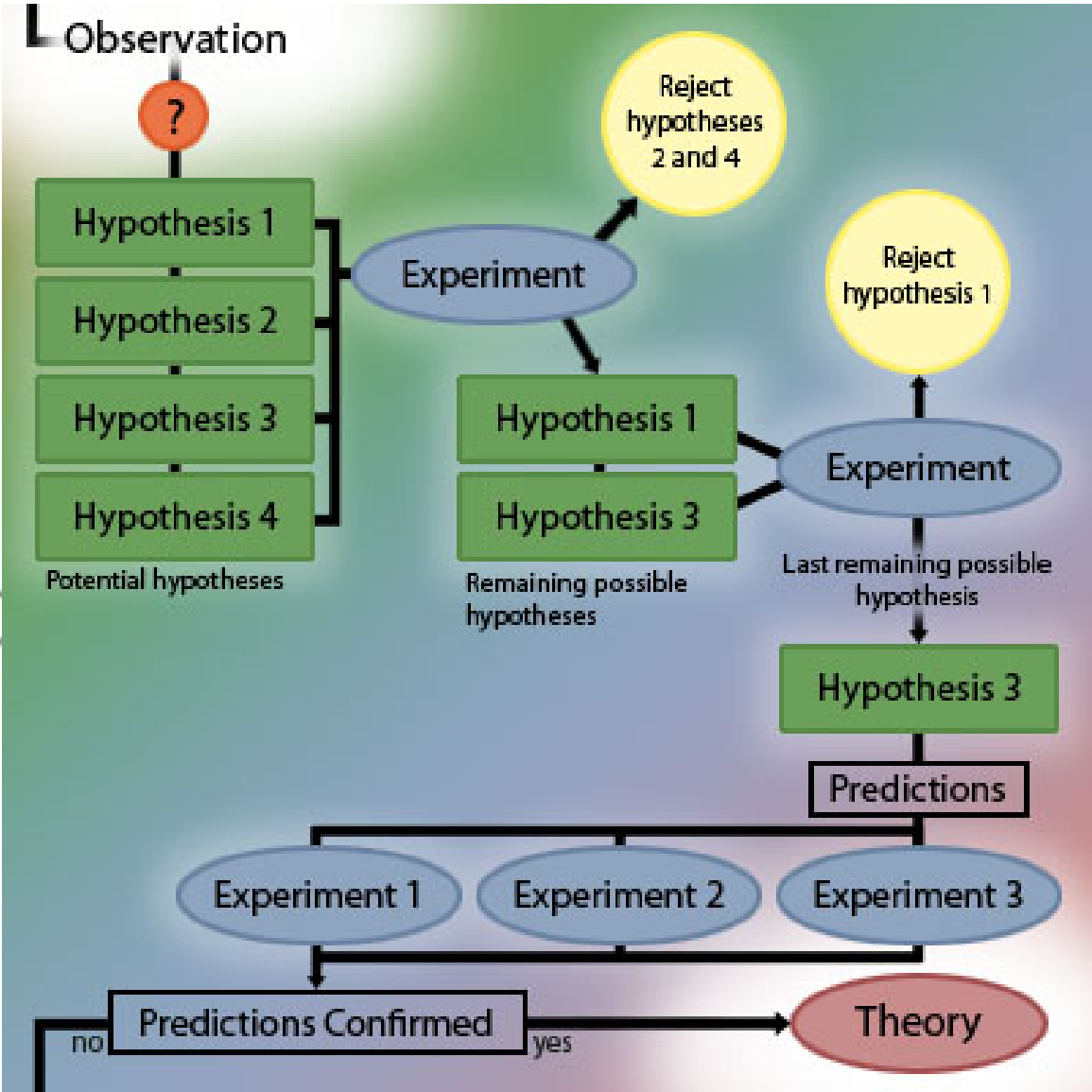
Figure 1   How Science is Done
A diagram that illustrates scientific investigation

Scientists first make observations that raise a particular question. In order to explain the observed phenomenon, they develop a number of possible explanations, or hypotheses. This is the inductive part of science, observing and constructing plausible arguments for why

an event occurred. Experiments are then used to eliminate one of more of the possible hypotheses until one hypothesis remains. Using deduction, scientists use the principles of their hypothesis to make predictions, and then test to make sure that their predictions are confirmed. After many trials (repeatability) and all predictions have been confirmed, the hypothesis then may become a theory.

Quick Definitions

Observation - Quantitative and qualitative measurements of the world.

Inference - Deriving new knowledge based upon old knowledge.

Hypotheses - A suggested explanation.

Rejected Hypothesis - An explanation that has been ruled out through experimentation.

Accepted Hypothesis - An explanation that has not been ruled out through excessive experimentation and makes verifiable predictions that are true.

Experiment - A test that is used to rule out a hypothesis or validate something already known.

Scientific Method - The process of scientific investigation.

Theory - A widely accepted hypothesis that stands the test of time. Often tested, and usually never rejected.

The scientific method is based primarily on the testing of hypotheses by experimentation. This involves a **control**, or subject that does not undergo the process in question. A scientist will also seek to limit variables to one or another very small number, single or minimum number of variables. The procedure is to form a hypothesis or prediction about what you believe or expect to see and then do everything you can to violate that, or falsify the hypotheses. Although this may seem unintuitive, the process serves to establish more firmly what is and what is not true.

A founding principle in science is a lack of absolute truth: the accepted explanation is the most likely and is the basis for further hypotheses as well as for falsification. All knowledge has its relative uncertainty.

Theories are hypotheses which have withstood repeated attempts at falsification. Common theories include evolution by natural selection and the idea that all organisms consist of cells. The scientific community asserts that much more evidence supports these two ideas than contradicts them.

2.4 Charles Darwin
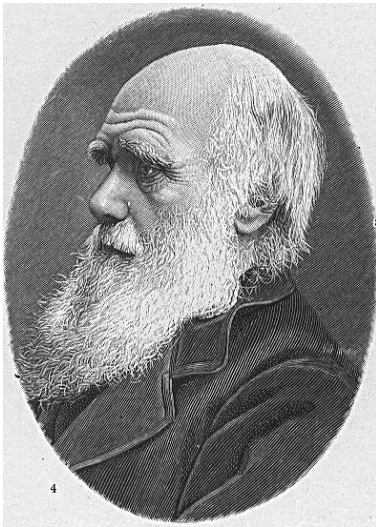
Figure 2

Charles Darwin is most remembered today for his contribution of the theory of **evolution through natural selection**.

The seeds of this theory were planted in Darwin's mind through observations made on a five-year voyage through the New World on a ship called the Beagle. There, he studied fossils and the geological record, geographic distribution of organisms, the uniqueness and relatedness of island life forms, and the affinity of island forms to mainland forms.

Upon his return to England, Darwin pondered over his observations and concluded that evolution must occur through natural selection. He declined, however, to publish his work because of its controversial nature. However, when another scientist, Wallace, reached similar conclusions, Darwin was convinced to publish his observations in 1859. His hypothesis revolutionized biology and has yet to be falsified by empirical data collected by mainstream scientists.

2.5 After Darwin

Since Darwin's day, scientists have amassed a more complete **fossil record**, including **microorganisms** and **chemical fossils**. These fossils have supported and added subtleties to Darwin's theories. However, the age of the Earth is now held to be much older than Darwin thought. Researchers have also uncovered some of the preliminary mysteries of the mechanism of heredity as carried out through **genetics** and **DNA**, areas unknown to Darwin. Another growing area is **comparative anatomy** including homology and analogy.

Today we can see a bit of evolutionary history in the **development** of embryos, as certain (although not all) aspects of development recapitulate evolutionary history.

The molecular biology[15] study of slowly mutating genes reveal considerable evolutionary history consistent with fossil and anatomical record.

2.6 Challenges to Darwin

Figure 3

Darwin and his theories have been challenged many times in the last 150 years. The challenges have been primarily religious based on a perceived conflict with the preconceived notion of creationism. Many of those who challenge Darwin have been adherents to the young earth hypothesis that says that the Earth is only some 6000 years old and that all species were individually created by a god. Some of the proponents of these theories have suggested that chemical and physical laws that exist today were different or nonexistent in earlier ages. However, for the most part, these theories are either not scientifically testable and fall outside the area of attention of the field of biology, or have been disproved by one or more fields of science.

This text is based on notes very generously donated by Dr. Paul Doerder, Ph.D., of Cleveland State University.

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15   http://en.wikibooks.org/wiki/molecular%20biology

3 The Nature of Molecules

3.1 Matter

Matter is defined as anything that has mass[1] (an amount of matter in an object) and occupies space[2] (which is measured as volume[3]).

- Particles, from smallest to largest
    1. Subatomic particles
        - Electrons[4]
        - Protons[5]
        - Neutrons[6]
    2. Atoms
    3. Molecules
    4. Macromolecules
- Origin of matter
    1. Big Bang[7], about 13.7 billion years ago
    2. Hydrogen[8], helium[9]
    3. Heavier elements formed in suns, super nova
        - Earth's matter predates formation of sun, 4.5 billion years ago
- All matter consists of atoms, which are composed of : electrons, protons, neutrons

3.2 The atom

- Example: Hydrogen
    - The simplest element
    - One proton (+)
    - One electron in orbit (-)
- Built by adding one proton (and one electron) at a time
- Number of protons determines atomic number and number of electrons
- Neutrons
    - Neutral charge

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1   http://en.wikipedia.org/wiki/mass
2   http://en.wikipedia.org/wiki/space%23Physics
3   http://en.wikipedia.org/wiki/volume
4   http://en.wikipedia.org/wiki/Electrons
5   http://en.wikipedia.org/wiki/Protons
6   http://en.wikipedia.org/wiki/Neutrons
7   http://en.wikipedia.org/wiki/Big%20Bang
8   http://en.wikipedia.org/wiki/Hydrogen
9   http://en.wikipedia.org/wiki/Helium

- Contribute mass
- May decay
- Oxygen[10]
  - 8 protons (mass)
  - 8 electrons
  - 8 neutrons (mass)

3.3 Mass and isotopes

- Atomic mass
  - Sum of masses of protons and neutrons
  - Measured in daltons or AMU (Atomic Mass Unit)
  - An AMU is 1/12 the mass of Carbon-12
  - proton ~1 AMU or dalton
  - 6.024 x 10^{23} daltons/gram
- Atoms with same atomic number belong to same element
- Isotopes
  - Same atomic number but different atomic mass
  - Some are radioactive
  - Uses of isotopes
    - Radioactive: 3H, 14C, 32P, 35S
      - Tracers in biochemical reactions
      - Detection of molecules in recombinant DNA technology (genetic engineering)
      - Half-life: dating of rocks, fossils
    - Non-radioactive (N, C, O)
      - Diet of organisms (including fossils)
      - Biochemical tracers

3.4 Electrons

- Negative charge
- Held in orbit about nucleus by attraction to positively charged nucleus
- Atom may gain or lose electron, altering charge
  - Cation: loses electron, positive charge
    - Na^+
  - Anion: gains electron, negative charge
    - Cl^-
- Determine chemical properties of atoms
  - Number
  - Energy level

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10   http://en.wikipedia.org/wiki/Oxygen

3.5 Chemical bonds

- Form molecules
- Enzymes: make, break, rearrange chemical bonds in living systems
- Ionic
- Covalent
    - Sharing of one or more pairs of electrons
        - Called single, double, or triple
    - No net charge (as in ionic bonds)
    - No free electrons
    - Give rise to discrete molecules
    - Hydrogen

3.6 Chemical reactions

- Formation and breaking of chemical bonds
- Shifting arrangement of atoms
- Reactants -> products
- Reactions are influenced by:
    - Temperature
    - Concentration of reactants, products
    - Presence of catalysts (enzymes)
- Oxidation:reduction

3.7 Water

- Essential for life
- ~75% earth's surface is water
- Life evolved in water
- Solvent for many types of solutes
- High specific heat
- High polarity
    - Creates a slightly negative Oxygen and a Slightly positive hydrogen
    - allows formation of Hydrogen Bonds

3.7.1 Hydrogen bonding

- A type of polar interaction
- Critical for:
    - Protein structure
    - Enzymatic reactions
    - Movement of water in plant stems
- Weak and transient
- Powerful cumulative effect

- Solubility of many compounds
- Cohesion (capillary action)
- Lower density of ice
- Formed between molecules other than water
  - Protein structure
  - DNA[11], RNA[12] structure

Water organizes nonpolar molecules

- Nonpolar molecules: no polarity (+/-) charges
- Hydrophobic: exclude water because they don't form hydrogen bonds with it
- Consequences:
  - Membranes
  - Protein structure
- Hydrophilic: polar substances associate with water

Ionization of water: $H_2O -> H^+ + OH^-$

- Forms a Hydrogen ion (H^+), hydroxide ion (OH^-)
- Due to spontaneous breakage of covalent bond
- At 25°C, 1 liter of water contains 10-7 moles of H^+ ions: 10-7 moles/liter

pH

- A convenient way of indicating H^+ concentration
- pH^{13} = -log[H^+]
- For water, pH = -log[10-7] = 7
- Since for each H^+ in pure water, there is one OH^-, pH of 7 indicates neutrality
- Logarithmic scale

Buffer

- Reservoir for H^+
- Maintains relatively constant pH over buffering range

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11   http://en.wikipedia.org/wiki/DNA
12   http://en.wikipedia.org/wiki/RNA
13   http://en.wikipedia.org/wiki/pH

14

4 The Chemical Building Blocks of Life

Building blocks of life

- Carbon based: organic molecules
- Carbohydrates: CHO
- Lipids: CHO, water insoluble
- Proteins: CHONS, structure/function in cells
- Nucleic acids: CHONP, hereditary (genetic) information

4.1 Carbon

- Can make 4 covalent bonds[1]
  - Chains
    - Straight
    - Branched
    - Ring
  - Hydrocarbons[2] (C, H): store energy
  - Functional groups
    - Attach to carbon
    - Alter chemical properties
    - Form macromolecules
    - Sapoteton

4.2 Carbohydrates

- Principally CHO (rare N, S and P)
  - 1C:2H:1O ratio
  - Energy rich (many C-H bonds)
- Monosaccharides (principal: glucose[3])
  - Simple sugars
  - Principle formula: $C_6H_{12}O_6$
  - Form rings in water solution
- Disaccharides (sucrose, lactose)
- Polysaccharides (starch, glycogen, cellulose, chitin)

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1   http://en.wikipedia.org/wiki/covalent%20bonds
2   http://en.wikipedia.org/wiki/Hydrocarbons
3   http://en.wikipedia.org/wiki/glucose

4.3 Stereoisomers

- Bond angles of carbon point to corners of a tetrahedron
- When 4 different groups are attached to a carbon, it is asymmetric, leading to various types of isomerism
  - Stereoisomers: (D, L)
- Same chemical properties
- Different biological properties
- D sugars, L amino acids

4.4 Lipids

- C-H bonds (nonpolar) instead of C-OH bonds as in carbohydrates
  - High energy
  - Hydrophobic (insoluble in water)
- Categories
  - Fats: glycerol and three fatty acids
  - Phospholipids: primary component of membranes
  - Prostaglandins: chemical messengers (hormones)
  - Steroids: membrane component; hormones
  - Terpenes: pigments; structure

4.4.1 Fatty acids

- Hydrocarbon chain
  - Even number of C, 14->20
  - Terminates in carboxyl group
- Saturated: contain maximum number of hydrogens (all single bonds); maximum energy
- Unsaturated: one or more double bonds
  - Usually higher melting point
  - Many common oils are polyunsaturated

4.5 Proteins

- Polymer of amino acids
  - 21 different amino acids found in proteins
  - Sequence of amino acids determined by gene
- Amino acid sequence determines shape of molecule
  - Linked by peptide bond (covalent)
- Functions
  - regulate chemical reactions and cell processes [enzymes]
  - form bone and muscle; various other tissues
  - facilitate transport across cell membrane [carrier proteins]

- fight disease [antibodies]
- Motifs: folding patterns of secondary structure
- Domains: structural, functional part of protein often independent of another part; often encoded by different exons
- Shape determines protein's function

4.5.1 Amino acids

- 21 commonly found in proteins
  - 21st is selenocysteine, not mentioned in text
- Common structure
  - Amino group: NH_2
  - Carboxyl group: COOH
  - R group- 4 different kinds of R groups
    - acidic
    - basic
    - hydrophilic (polar)
    - hydrophobic (nonpolar)
- Confer individual properties on amino acids
- List of amino acids[4]

4.5.2 Structure

- Primary structure: the amino acid sequence
  - Determines higher orders of structure
  - Critical for structure and function of protein
- Secondary: stabilized by intramolecular hydrogen bonding
  - helix
  - sheet
- Tertiary: folding, stabilized by ionic bonds (between R groups), hydrogen bonding, van der Waal's forces, hydrophobic interactions
- Quaternary: _2 polypeptides

4.5.3 Function

- Requires proper folding, cofactors, pH, temperature, etc.
- Proteins are often modified after synthesis
  - Chemical modification
  - Addition of heme groups (hemoglobin, cytochrome)
- Denatured proteins can not function properly
- Proteins are degraded by proteosome as part of constant turnover of cell components

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4   `http://en.wikipedia.org/wiki/Amino%20acid%23List%20of%20amino%20acids`

4.6 Hereditary (Genetic) information

- Nucleic acids
  - DNA: deoxyribonucleic acid
- Hereditary information of all cells
- Hereditary information for many viruses
  - RNA: ribonucleic acid
- Hereditary information of certain viruses (HIV[5])
- Intermediate in gene expression
- Composed of nucleotides
  - Ribonucleotides
  - Deoxyribonucleotides

4.6.1 RNA DNA origin

- Which came first?
- Paradox: DNA encodes protein necessary for its own replication
- Discovery of catalytic RNA by Cech and Altman suggested that RNA might have been first self-replicating molecule
- DNA evolved as more stable type of storage molecule

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Proteins: Their building block is amino acids. The bond connecting 2 of the amino acids together are called peptide bonds. One of these bonds makes a monopeptide, two a dipeptide, and any more than that makes a polypeptide.

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5   http://en.wikipedia.org/wiki/HIV

5 Life: History and Origin

5.1 Properties of life

1. Organization: Being structurally composed of one or more cells, which are the basic units of life.
   - prokaryote: no nucleus
   - eukaryote: membrane bound nucleus.

2. Sensitivity: respond to stimuli.

3. Energy Processing

4. Growth and Development

5. Reproduction

   - hereditary mechanisms to make more of self; DNA based.
6. Regulation, including homeostasis.

7. Evolution.

5.2 Origin of life: 3 hypotheses

- Extraterrestrial origin (panspermia): meteor, comet borne from elsewhere in universe
  - evidence of amino acids and other organic material in space (but often both D & L forms)
  - questionable bacterial fossils in Martian rock

-However, this would imply that some other origin of life was likely because it would have had to happen elsewhere before it could be transported here, and the only difference would be that life did not originate on Earth.

- Spontaneous origin on earth: primitive self-replicating macromolecules acted upon by natural selection ((macro)Evolution is one example of this)

-This is often attacked for the seeming impossibility for life to have been produced by a chemical reaction triggered by lightning and the ability of any produced DNA to actually be in a sequence that could produce a working model of life if replicated. It is also attacked for religious reasons, as it bypasses things like the idea of a supreme being directly creating humans. It also seems unlikely to some that such huge changes are possible in evolution

without evidence of an "in-between stage" that is credible. Many of the stages of man are disputed due to their somewhat shakey grounds. For example, bones from other animals have been taken accidentally in some cases to be part of a humanoid, and complete skeletons have been sketched out from a limited number of bones.

- Special creation: religious explanations (Intelligent Design is one popular example of this.) These explanations contend that life was created by God (or perhaps some other Intelligent Designer).
  - Proponents of Intelligent design suggest that the vast complexity of life could only have been intentionally designed while other creationists cite biblical support.

-This is often attacked for many of the same reasons that religion is attacked, and is often regarded as superstitious and/or unscientific.

- It is debated as to whether schools should teach one hypothesis or the other when talking about the origin of life. However, since they are all currently known major hypotheses (and sometimes hypotheses proven wrong are shown for educational purposes), this wikibook includes what it can without discriminating unfairly against one hypothesis or the other.

5.3 The early earth

It is believed that the Earth was formed about 4.5 billion years ago.

- Heavy bombardment by rubble ceased about 3.8 billion years ago.
- Reducing atmosphere: much free H
  - also H_2O, NH_3, CH_4
  - little, if any, free O_2
  - with numerous H electrons, require little energy to form organic compounds with C
- Warm oceans, estimated at 49-88°C
- Lack of O_2 and consequent ozone (O_3) meant considerable UV energy

Chemical reactions on early earth

- UV and other energy sources would promote chemical reactions and formation of organic molecules
- Testable hypothesis: Miller-Urey experiment
  - simulated early atmospheric conditions
  - found amino acids, sugars, etc., building blocks of life
  - won Nobel prize for work
  - experiment showed prebiotic synthesis of biological molecules was possible

Issues

- Miller later conceded that the conditions in his experiments were not representative of what is currently thought to be those of early earth
- He also conceded that science has no answer for how amino acids could self-organize into replicating molecules and cells
- In the 50 years since Miller-Urey, significant issues and problems for biogenesis have been identified. This is a weak hypothesis at this time.

- Conclusion: Life exists, we don't know why.

5.4 Origin of cells

Cells are very small and decompose quickly after death. As such, fossils of the earliest cells do not exist. Scientists have had to form a variety of theories on how cells (and hence life) was created on Earth.

- Bubble hypothesis
  - A. Oparin, J.B.S. Haldane, 1930's
- Primary abiogenesis: life as consequence of geochemical processes
- Protobionts: isolated collections of organic material enclosed in hydrophobic bubbles
  - Numerous variants: microspheres, protocells, protobionts, micelles, liposomes, coacervates
- Other surfaces for evolution of life
  - deep sea thermal vents
  - ice crystals
  - clay surfaces
  - tidal pools

5.5 The RNA world?

- DNA → RNA → polypeptide (protein)
- Catalytic RNA: ribozyme
  - discovered independently by Tom Cech and Sid Altman (Nobel prize)
  - catalytic properties: hydrolysis, polymerization, peptide bond formation, etc.
- Self-replicating RNA molecule may have given rise to life
  - consistent with numerous roles for RNA in cells as well as roles for ribonucleotides (ATP)
  - relationship to bubble-like structures is uncertain

5.6 The earliest cells

- Microfossils
  - ˜3.5 by
  - resemble bacteria: prokaryotes
  - biochemical residues
  - stromatolites
- Archaebacteria (more properly Archaea)
  - extremophiles: salt, acid, alkali, heat, methanogens
  - may not represent most ancient life
- Eubacteria
  - cyanobacteria: photosynthesis
- atmospheric O_2; limestone deposits

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- chloroplasts of eukaryotes

Cyanobacteria

5.7 Major steps in evolution of life

- Prebiotic synthesis of macromolecules
- Self replication
  - RNA? (primitive metabolism)
- DNA as hereditary material
- 1st cells
- Photosynthesis
- Aerobic respiration
- Multicellularity (more than once)

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6 Cells

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7 Cell structure

7.1 What is a cell?

The word cell comes from the Latin word "cella", meaning "small room", and it was first coined by a microscopist observing the structure of cork. The cell is the basic unit of all living things, and all organisms are composed of one or more cells. Cells are so basic and critical to the study of life, in fact, that they are often referred to as "the building blocks of life". Organisms - bacteria, amoebae and yeasts, for example - may consist of as few as one cell, while a typical human body contains about a trillion cells.

According to *Cell Theory,* first proposed by Schleiden and Schwann in 1839, all life consists of cells. The theory also states that all cells come from previously living cells, all vital functions (chemical reactions) of organisms are carried out inside of cells, and that cells contain necessary hereditary information to carry out necessary functions and replicate themselves.

All cells contain:

- Lipid bilayer boundary (plasma membrane[1])
- Cytoplasm[2]
- DNA[3] (hereditary information)
- Ribosomes[4] for protein synthesis

Eukaryotic cells also contain:

- At least one nucleus[5]
- Mitochondria[6] for cell respiration and energy

Cells may also contain:

- Lysosomes[7]
- Peroxisomes[8]
- Vacuoles[9]
- Cell walls[10]

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1   http://en.wikipedia.org/wiki/plasma%20membrane
2   http://en.wikipedia.org/wiki/Cytoplasm
3   http://en.wikipedia.org/wiki/DNA
4   http://en.wikipedia.org/wiki/Ribosome
5   http://en.wikipedia.org/wiki/Cell%20Nucleus
6   http://en.wikipedia.org/wiki/Mitochondrion
7   http://en.wikipedia.org/wiki/Lysosome
8   http://en.wikipedia.org/wiki/Peroxisome
9   http://en.wikipedia.org/wiki/Vacuole
10   http://en.wikipedia.org/wiki/Cell%20Wall

7.1.1 Concepts

Plasma Membrane

Phospholipid bilayer, which contains great amount of proteins, the most important functions are the following:

1. It selectively isolates the content of the cell of the external atmosphere.
2. It regulates the interchange of substances between the cytoplasm and the environment.
3. Communicates with other cells.

Model of the fluid mosaic

Describes the structure of the plasma membrane, this model was developed in 1972 by cellular biologists J. Singer and L. Nicholson.

Phospholipid bilayer

Is in the plasma membrane and produces the fluid part of membranes.

Proteins

Long chains of amino acids.

Glucose proteins

Proteins together with carbohydrates in the plasma membrane, mostly in the outer parts of the cell.

Functions of proteins

Transport oxygen, they are components of hair and nails, and allow the cell interact with its environment.

Transport Proteins

Regulate the movement of soluble water molecules, through the plasma membrane. Some transport proteins called **channel proteins** form pores or channels in the membrane so that water soluble molecules pass.

Carrying proteins

Have union sites that can hold specific molecules.

Reception proteins

They activate **cellular responses** when specific molecules join.

Proteins of recognition

They work as identifiers and as place of union to the cellular surface.

Fluid

It is any substance that can move or change of form.

Concentration

Number of molecules in a determined unit of volume.

Gradient

Physical difference between two regions of space, in such a way that the molecules tend to move in response to the gradients.

Diffusion

Movement of the molecules in a fluid, from the regions of high concentration to those of low concentration.

Passive transport

Movement of substances in a membrane that doesn't need to use energy.

Simple diffusion

Diffusion of water, gases or molecules across the membrane.

Facilitated diffusion

Diffusion of molecules across the membranes with the participation of proteins.

Osmosis

Diffusion of the water across a membrane with differential permeability.

Transport that needs energy

Movement of substances across a membrane generally in opposition to a gradient of concentration with the requirement of energy.

Active transport

Movement of small molecules using energy (ATP).

Endocytosis

Movement of big particles towards the interior of the cell using energy. The cells enclose particles or liquids.

Pinocytosis

(Literally cell drinking) Form in which the cell introduces liquids.

Phagocytosis

Way of eating of the cells. It feeds in this case of big particles or entire microorganisms.

Pseudopods

False feet (the amoeba).

Exocitosis

Movement of materials out of the cell with the use of energy. It throws waste material.

Isotonic

The cytoplasm fluid of the interior of the cells is the same that the outer.

Hypertonic solution

The solutions that have a higher concentration of dissolved particles than the cellular cytoplasm and that therefore water of the cells goes out with osmosis.

Hypotonic

The solutions with a concentration of dissolved particles lower than the cytoplasm of a cell and that therefore do that water enters the cell with osmosis.

Swelling

Pressure of the water inside the vacuole.

Endoplasmic Reticulum

It is the place of the synthesis of the cellular membrane.

7.1.2 Structure and function of the cell

Rudolf Virchow

Zoologist, who proposed the postulates of the cellular theory, observes that the living cells could grow and be in two places at the same time, he proposed that all the cells come from other equal cells and proposed 3 postulates:

1. Every living organism is formed from one or more cells
2. The smallest organisms are unicellular and these in turn are the functional units of the multicellular organisms.
3. All the cells come from preexisting cells.

7.1.3 Common characteristics of all the cells

Molecular components

Proteins, amino acids, lipids, sweeten, DNA, RNA.

Structural components

Plasmatic membrane, citoplasm, ribosomes.

Robert Hooke

He postuled for the first time the term *cell*

Prokaryotes

Their genetic material is not enclosed in a membrane ex. Bacterias

Eukaryotes

Their genetic material is contained inside a nucleus closed by a membrane

7.2 History of cell knowledge

The optical microscope was first invented in 17th century. Shortly thereafter scientists began to examine living and dead biological tissues in order to better understand the science of life. Some of the most relevant discovery milestones of the time period include:

- The invention of the microscope[11], which allowed scientists for the first time to see biological cells
- Robert Hooke[12] in 1665 looked at cork under a microscope and described what he called cork "cells"
- Anton van Leeuwenhoek[13] called the single-celled organisms that he saw under the microscope "animalcules"
- Matthias Jakob Schleiden[14], a botanist, in 1838 determined that all *plants* consist of cells
- Theodor Schwann[15], a zoologist, in 1839 determined that all *animals* consist of cells
- Rudolf Virchow[16] proposed the theory that all cells arise from previously existing cells

In 1838, the botanist Matthias Jakob Schleiden and the physiologist Theodor Schwann discovered that both plant cells and animal cells had nuclei. Based on their observations, the two scientists conceived of the hypothesis that all living things were composed of cells. In 1839, Schwann published 'Microscopic Investigations on the Accordance in the Structure and Growth of Plants and Animals', which contained the first statement of their joint cell theory.

7.2.1 Cell Theory

Schleiden and Schwann proposed spontaneous generation as the method for cell origination, but spontaneous generation (also called abiogenesis[17]) was later disproven. Rudolf Virchow famously stated "Omnis cellula e cellula"... "All cells only arise from pre-existing cells." The parts of the theory that did not have to do with the origin of cells, however, held up to scientific scrutiny and are widely agreed upon by the scientific community today.

The generally accepted portions of the modern Cell Theory are as follows: (1) The cell is the fundamental unit of structure and function in living things. (2) All organisms are made up of one or more cells. (3) Cells arise from other cells through cellular division. (4) Cells carry genetic material passed to daughter cells during cellular division. (5) All cells are essentially the same in chemical composition. (6) Energy flow (metabolism and biochemistry) occurs within cells.

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11   http://en.wikipedia.org/wiki/microscope
12   http://en.wikipedia.org/wiki/Robert%20Hooke
13   http://en.wikipedia.org/wiki/Anton%20van%20Leeuwenhoek
14   http://en.wikipedia.org/wiki/Matthias%20Jakob%20Schleiden
15   http://en.wikipedia.org/wiki/Theodor%20Schwann
16   http://en.wikipedia.org/wiki/Rudolf%20Virchow
17   http://en.wikipedia.org/wiki/Abiogenesis

7.3 Microscopes

- Allow greater resolution, can see finer detail
- Eye: resolution of ˜ 100 μm
- Light microscope[18]: resolution of ˜ 200 nm
- Limited to cells are larger organelles within cells
- Confocal microscopy[19]: 2 dimension view
- Electron microscope[20]: resolution of ˜0.2 nm
- Laser tweezers: move cell contents

7.4 Cell size

One may wonder why all cells are so small. If being able to store nutrients is beneficial to the cell, how come there are no animals existing in nature with huge cells? Physical limitations prevent this from occurring. A cell must be able to diffuse gases and nutrients in and out of the cell. A cell's surface area does not increase as quickly as its volume, and as a result a large cell may require more input of a substance or output of a substance than it is reasonably able to perform. Worse, the distance between two points within the cell can be large enough that regions of the cell would have trouble communicating, and it takes a relatively long time for substances to travel across the cell.

That is not to say large cells don't exist. They are, once again, less efficient at exchanging materials within themselves and with their environment, but they are still functional. These cells typically have more than one copy of their genetic information, so they can manufacture proteins locally within different parts of the cell.

Key concepts: Cell size:

- Is limited by need for regions of cell to communicate
- Diffuse oxygen and other gases
- Transport of mRNA[21] and protein[22]s
  - Surface area to volume ratio limited
- Larger cells typically:
  - Have extra copies of genetic information
  - Have slower communication between parts of cell

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18   http://en.wikipedia.org/wiki/Light%20microscope
19   http://en.wikipedia.org/wiki/microscopy
20   http://en.wikipedia.org/wiki/Electron%20microscope
21   http://en.wikipedia.org/wiki/RNA
22   http://en.wikipedia.org/wiki/protein

8 Structure of Eukaryotic cells

Eukaryotic[1] cells feature membrane delimited nucleii containing two or more linear chromosome[2]s; numerous membrane-bound cytoplasmic organelles: mitochondria, RER[3], SER[4], lysosomes, vacuole[5]s, chloroplast[6]s; ribosomes and a cytoskeleton[7]. Also, plants, fungi, and some protists have a cell wall.

8.1 Structure of the nucleus

The nucleus is the round object in the cell that holds the genetic information (DNA) of the cell. It is surrounded by a nuclear envelope and has a nucleolus inside.

8.1.1 Nuclear envelope

The nuclear envelope is a double-layered plasma membrane[8] like the cell membrane, although without membrane proteins. To allow some chemicals to enter the nucleus, the nuclear envelope has structures called Nuclear pore[9]s. The nuclear envelope is continuous with the endoplasmic reticulum.

8.1.2 Nucleolus

The nucleolus appears in a microscope as a small dark area within the nucleus. The nucleolus is the area where there is a high amount of DNA transcription[10] taking place.

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1   http://en.wikipedia.org/wiki/Eukaryote
2   http://en.wikipedia.org/wiki/chromosome
3   http://en.wikipedia.org/wiki/RER
4   http://en.wikipedia.org/wiki/SER
5   http://en.wikipedia.org/wiki/vacuole
6   http://en.wikipedia.org/wiki/chloroplast
7   http://en.wikipedia.org/wiki/cytoskeleton
8   http://en.wikibooks.org/wiki/Plasma%20membrane
9   http://en.wikibooks.org/wiki/Nuclear%20pore
10   http://en.wikibooks.org/wiki/DNA%20transcription

8.2 Chromatin

Chromosomes consist of chromatin[11]. This is made up of strings of DNA, which typically measure centimeters in length if stretched out. This DNA is wound around a histone[12] core and organized into nucleosome[13]s.

The chromatin[14] must be uncoiled for gene expression[15] and replication[16]. Chromosome micrograph

8.3 Endoplasmic reticulum

The endoplasmic reticulum[17] is a cellular organelle[18] made up of a series of extended folded intracellular membranes. It is continuous with the nuclear membane.

There are two main types of endoplasmic reticulum:

- RER: rough endoplasmic reticulum (site of protein synthesis[19]) associated with ribosomes
- SER: smooth endoplasmic reticulum (site of lipid synthesis[20])

8.3.1 Rough Endoplasmic Reticulum

Proteins are directed to the RER by a **signal sequence** of a growing polypeptide[21]s on the ribosome. This is recognised by a signal recognition particle which brings the ribosome/polypeptide complex to a channel on the RER called a translocon. At the translocon, the signal sequence and ribosome/polypeptide complex interact with the translocon to open it. The signal sequence becomes attached to the translocon. The ribosome can continue to translate the polypeptide into the lumen of the RER. As synthesis continues, 2 processes can happen.

1. If the protein is destined to become a membrane bound protein then the protein synthesis will continue until termination. The ribosome can then dissociate, allowing protein folding within the RER lumen to occur and continuation to the golgi apparatus for processing of the polypeptide.
2. If the protein is destined for storage for later secretion after stimulation or for continuous secretion then a protease-enzyme which cuts proteins at the peptide bond-can cut the signal sequence from the growing polypeptide. Continuation to the golgi etc. can then occur.

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11 http://en.wikipedia.org/wiki/chromatin
12 http://en.wikipedia.org/wiki/histone
13 http://en.wikipedia.org/wiki/nucleosome
14 http://en.wikipedia.org/wiki/chromatin
15 http://en.wikipedia.org/wiki/gene%20expression
16 http://en.wikipedia.org/wiki/replication
17 http://en.wikipedia.org/wiki/endoplasmic%20reticulum
18 http://en.wikipedia.org/wiki/organelle
19 http://en.wikipedia.org/wiki/protein%20synthesis
20 http://en.wikipedia.org/wiki/lipid%20synthesis
21 http://en.wikipedia.org/wiki/polypeptide

When produced, proteins are then exported to one of several locations. The proteins are either modified for extracellular membrane insertion or secretion. Note, this is in contrast with ribosomes which do not associate with the RER and produce proteins which will become cytosolic enzymes for example.

8.3.2 Smooth Endoplasmic Reticulum

Smooth endoplasmic reticulum produces enzyme[22]s for lipid and carbohydrate biosynthesis and detoxification RER

8.3.3 Sarcoplasmic Reticulum

This is a specialised form of endoplasmic reticulum found in some muscle cell types-particularly striated, skeletal muscle. Its main function is different from the other 2 types in that is mainly acts as a storage of calcium. This reticulum has voltage gated channels which respond to signals from 'motor neurones' to open and release calcium into the cytoplasm. This can then bring about the next part in muscle contraction.
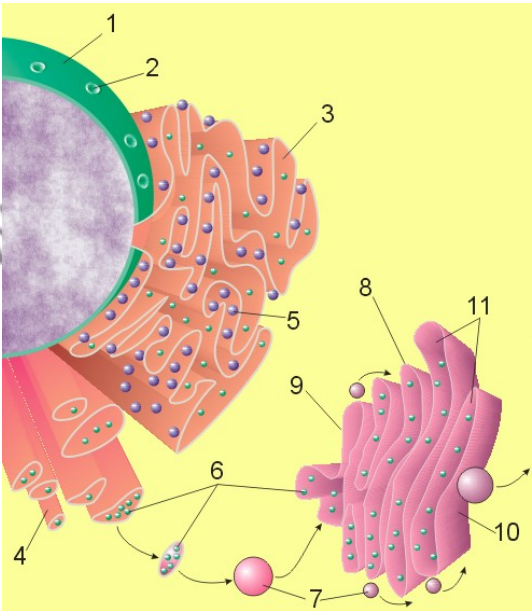
Figure 4

Figure 1 :Image of nucleus[23], endoplasmic reticulum and Golgi apparatus.

1. Nucleus.
2. Nuclear pore.
3. Rough endoplasmic reticulum (RER).
4. Smooth endoplasmic reticulum (SER).
5. Ribosome on the rough ER.
6. Proteins that are transported.
7. Transport vesicle.
8. Golgi apparatus.
9. Cis face of the Golgi apparatus.
10. Trans face of the Golgi apparatus.
11. Cisternae of the Golgi apparatus.
1.

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22   http://en.wikipedia.org/wiki/enzyme
23   http://en.wikipedia.org/wiki/cell%20nucleus

8.4 The Golgi apparatus

The golgi apparatus[24] is made up of multiple stacks of bilipid membranes.

- Proteins made on the RER are modified and then sorted
- Formation of secretory vesicles
- Formation of lysosomes (intracellular digestion)

Other membrane-bound cytoplasmic organelles include:

- Microbodies[25] (generic term)
- Glyoxysome (transforms fat into carbohydrate in plants)
- Peroxisome[26] (uses oxidative metabolism to form hydrogen peroxide and is destroyed by catalase[27])

8.5 Ribosomes

Ribosomes are the site of protein synthesis. Ribosomes themselves are synthesized in the cell nucleoli[28] and are structured as two subunits, the large and the small. These parts are composed of RNA and protein.

Prokaryotic and eukaryotic ribosomes are different, the eukaryotic ones being larger and more complicated.

8.6 DNA-containing organelles

Mitochondria

- Double membrane
- Aerobic metabolism, internal membrane
- DNA, ribosomes
- Give rise to new mitochondria

Chloroplast[29]

- Double membrane
- Photosynthesis, internal membrane
- DNA, ribosomes
- Give rise to new chloroplasts

Centriole[30]s

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24   http://en.wikipedia.org/wiki/Golgi%20apparatus
25   http://en.wikipedia.org/wiki/Microbody
26   http://en.wikipedia.org/wiki/Peroxisome
27   http://en.wikipedia.org/wiki/catalase
28   http://en.wikipedia.org/wiki/nucleoli
29   http://en.wikipedia.org/wiki/Chloroplast
30   http://en.wikipedia.org/wiki/Centriole

- Microtubule organizing centers
  - Animal cells and many protists
  - Pair constitutes the centrosome
  - Give rise to flagellum during spermatogenesis
- Consist of 9 triplet microtubules
- Mitosis[31], meiosis[32]

8.7 Cytoskeleton

Cytoskeleton is a collective term for different filaments of proteins that can give physical shape within the cell and are responsible for the 'roads' which organelles can be carried along.

- Gives the cell shape
- Anchors other organelles
- Vital to intracellular transport of large molecules

The cytoskeleton is composed of 3 main types of filaments:

- Actin[33] filaments (7 nm)
- Microtubule[34]s: (25 nm) polymer of tubulin; 13/ring.
- Intermediate Filament[35]s

Both actin and microtubules can have associated motor proteins.

8.7.1 Intermediate Filaments

These are rope like filaments, 8-10nm in diameter and tend to give the structural stability to cells. Examples inculude Vimentin, neurofilaments and keratin. It is keratin which priniciply makes up hair, nails and horns.

8.7.2 Actin Filaments

Growth
These filaments are 2-stranded and composed of dimeric subunits called G-Actin. They contain a GTP molecule in order to bind (polymerise). As GTP is hydrolysed then the structure becomes unstable and depolymerisation occurs. The growth of actin filaments is concentration dependant-that is, the higher the concentration of free G-actin, the greater the polymerisation. The are also polar, having a + and a - end (not related to charge) and polymerisation tends to happen faster at the + end.

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31   http://en.wikipedia.org/wiki/mitosis
32   http://en.wikipedia.org/wiki/meiosis
33   http://en.wikipedia.org/wiki/Actin
34   http://en.wikipedia.org/wiki/Microtubule
35   http://en.wikipedia.org/wiki/Intermediate%20Filament

Structure of Eukaryotic cells

Cilia and flagella are threads of microtubules that extend from the exterior of cells and used to move single celled organisms as well as move substances away from the surface of the cell. motor proteins-move, wave motion

9 Membranes

9.1 Biological membranes
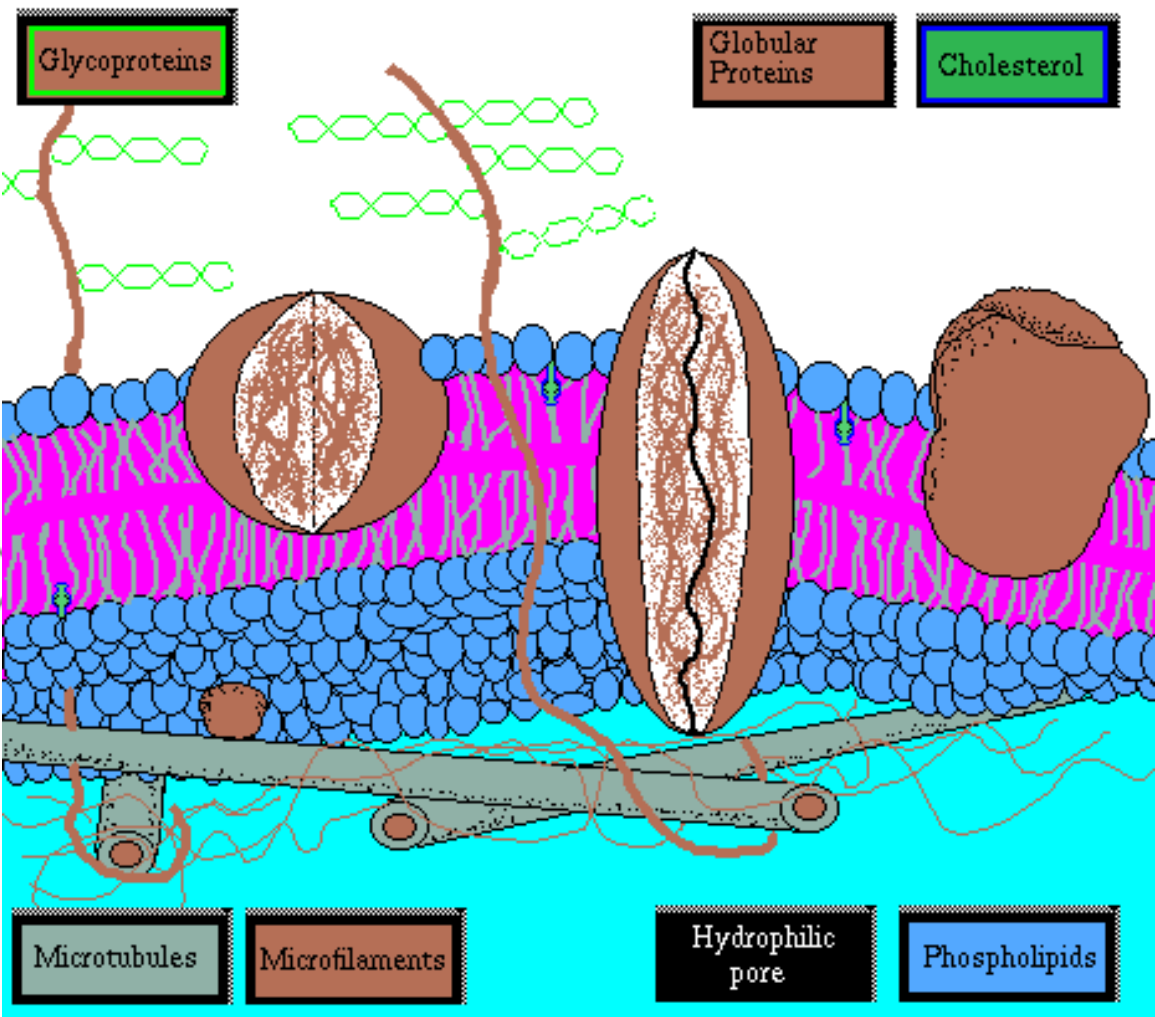
Figure 5   Plasma membrane bilayer

Membranes

Biological membranes surround cells and serve to keep the insides separated from the outsides. They are formed of phospholipid bilayer[1]s, which by definition are a double layer of fatty acid[2] molecules (mostly phospholipid[3]s, lipids containing lots of phosphorus).

Proteins[4] serve very important functions in cellular membranes. They are **active transports** in and out of the cell, acting as gatekeepers. They relay **signals** in and out of the cell. Proteins are the **site of many enzymatic reactions** in the cell, and play a role in **regulation of cellular processes.**

9.2 Phospholipid

Phospholipid bilayer

- basis of biological membranes and cellular organisms
- contains a charged, hydrophilic (attracted to water) head and two hydrophobic (repelled by water) hydrocarbon tails
- In presence of water, phospholipids form bilayer
  - maximize hydrogen bonds between water
  - creates barrier to passage of materials
  - fluid mosaic model shows horizontal (common) and "flip-flop" (rare) movement of phospholipids

9.3 Fluid mosaic model

- Current model of membrane
- Phospholipid bilayer
  - Phospholipids
- Move freely in lipid layer, but rarely switch layers
- Different phospholipids in each layer in different organelles
  - Glycolipids
  - Sterols (cholesterol in animals)
- Transmembrane proteins "float" in fluid lipid bilayer
  - also called intrinsic, integral proteins
- Exterior (extrinsic, peripheral) proteins

9.4 Membrane proteins

- Transport channels
- Enzymes
- Cell surface receptors

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1   http://en.wikibooks.org/wiki/lipid%20bilayer
2   http://en.wikibooks.org/wiki/Fatty%20acid
3   http://en.wikibooks.org/wiki/phospholipid
4   http://en.wikibooks.org/wiki/Proteins

- Cell surface identity markers
- Cell adhesion proteins
- Attachments to cytoskeleton

Integral membrane proteins

- Anchoring to membrane
  - Protein has attached phosphatidylinositol (GPI) linkage, anchors protein in outer layer (no picture)
  - Protein has one or more hydrophobic transmembrane domains
- -helix
- -sheet

Channel protein Transport across membranes *** Diffusion**

- - From higher concentration to lower concentration
  - Membranes are selectively permeable
- Ions diffuse through membrane channels
  - Selective
  - Movement determined by diffusion and voltage differences
- **Facilitated diffusion**
  - Carrier protein, physically binds transported molecule
- **Osmosis**
- Diffusion of water down concentration gradient
- In cell: various solutes (amino acids, ions, sugars, etc.)
  - interact with water, e.g., hydration shells
- Water moves through aquaporin channels into cell
- Depends upon the concentration of all solutes in solution
  - Hyperosmotic solution: higher concentration of solutes
  - Hypoosmotic solution: lower concentration of solutes
  - Isoosmotic solution: solute concentrations equal
- Water moves from hypoosmotic solution to hyperosmotic solution

Osmotic pressure **Bulk transport**

- Endocytosis: energy requiring
  - Phagocytosis
- Solid material, typically food
  - Pinocytosis
- Primarily liquid

** Receptor-mediated endocytosis

- Pits on cell surface coated with clathrin and receptors
- Bind specific proteins
- Exocytosis
  - Discharge of materials from vesicle at cell surface

9.5 Receptor-mediated endocytosis

Active transport

- Energy required (usually ATP)
- Highly selective
- Works against concentration gradient
- Many examples, e.g., Na^+/K^+ pump

Cotransport (coupled transport)

- Does not use ATP directly
- Molecule is transported in connection with another molecule that is moving down a concentration gradient
  - Example: Na^+ gradient is established by a Na^+ pump, with higher concentration on outside of cell. Cotransport channel carries Na+ and another molecule (e.g. glucose) into cell
- May involve proton (H^+) pumps (chemiosmosis - ATP production)

This text is based on notes very generously donated by Dr. Paul Doerder, Ph.D., of the Cleveland State University.

10 Cell-cell interactions

with the environment with each other

10.1 Cell signaling

- Signaling requires
  - Signal
  - Cell receptor (usually on surface)
- Signaling is important in:
  - Response to environmental stimuli
  - Sex
  - Development
- Major area of research in biology today

10.1.1 Types of signaling

- Direct contact (e.g., gap junctions between cells)
- Paracrine: Diffusion of signal molecules in extracellular fluid; highly local
- Endocrine: Signal (hormone) molecule travels through circulatory system
- Synaptic: neurotransmitters

Types of signal molecules

- Hormones: chemically diverse
  - Steroid
  - Polypeptide
  - Vitamin/amino acid derived
- Cell surface proteins/glycoproteins
- Ca^{2+}, NO
- Neurotransmitter
  - Several hundred types
  - Some are also hormones e.g. Estrogen, progesterone

Receptor molecules

- Intracellular
  - Protein that binds signal molecule in cytoplasm

- Bound receptor may act as:
- Gene regulator
- Enzyme
- Cell surface
  - Gated ion channels (neurotransmitter receptor)
  - Enzymic receptors
  - G protein-linked receptors

Cell surface protein

- Tissue identity
  - glycolipids
  - MHC proteins
- Immune systems
  - distinguish self from not-self
- Intercellular adhesion
  - permanent contact
  - help form sheets of cells, tissues
  - may permit signaling

Example: G proteins

- Transmembrane surface receptor binds signal molecule
- Conformational change allows binding of G protein on cytoplasmic side
- G protein binds GTP, becomes activated
- G protein activates intracellular signal cascade
  - Change in gene expression
  - Secrection
  - Many other possible consequences

10.2 Communicating junctions

- Gap junctions
  - animals
  - small molecules and ions may pass
- Plasmodesmata
  - plants
  - lined with plasma membrane
  - permit passage of water, sugars, etc.

10.2.1 Gap junctions

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11 Energy and Metabolism

11.1 Energy

- The capacity to do work.
  - Kinetic energy: energy of motion (ex. jogging).
  - Potential energy: stored energy (ex. a lion that is about to leap on its prey).
- Many forms of energy: e.g.,
  - Heat
  - Sound
  - Electric current
  - Light
  - All convertible to heat
- Most energy for biological world is from sun
- Heat (energy of random molecular motion, thermal energy)
  - Convenient in biology
  - All other energy forms can be converted to heat
  - Thermodynamics: study of thermal energy
- Heat typically measured in kilocalories
  - Kcal: 1000 calories
  - 1 calorie: amount of heat required to raise the temperature of one gram of water one degree Celsius (°C)
- Heat plays major role in biological systems
  - Ecological importance
  - Biochemical reactions

11.2 Oxidation–Reduction

- Energy flows into biological world from sun
- Light energy is captured by photosynthesis
  - Light energy raises electrons to higher energy levels
  - Stored as potential energy in covalent C-H bonds of sugars
- Strength of covalent bond is measured by amount of energy required to break it
  - 98.8 kcal/mole of C-H bonds
- In chemical reaction, energy stored in covalent bonds may transfer to new bonds. When this involves transfer of electrons, it is oxidation–reduction reaction
- Always take place together
  - Electron lost by atom or molecule through oxidation is gained by another atom or molecule through reduction
  - Potential energy is transferred from one molecule to another (but never 100%)

- Often called redox reactions
  - Photosynthesis
  - Cellular Respiration
  - Chemiosynthesis
  - Autotrophs
  - Heterotrophs

11.3 NAD$^+$

- Common electron acceptor/donor in redox reactions
- Energetic electrons often paired with H$^+$

11.4 Free energy

- Energy required to break and subsequently form other chemical bonds
  - Chemical bonds: sharing of electrons, tend to hold atoms of molecule together
  - Heat, by increasing atomic motion, makes it easier to break bonds (entropy)
- Energy available to do work in a system
- In cells, G = H - TS
  - G = Gibbs' free energy
  - H = H (enthalpy) energy in molecule's chemical bonds
  - TS (T, temperature in °K; S, entropy)
- Chemical reactions break and make bonds, producing changes in energy
- Under constant conditions of temperature, pressure and volume, $\Delta G = \Delta H - T\Delta S$
- ΔG, change in free energy
  - If positive (+), H is higher, S is lower, so there is more free energy; endergonic reaction, does not proceed spontaneously; require input of energy (e.g., heat)
  - If negative (–), H is lower, S is higher. Product has less free energy; exergonic; spontaneous

===Activation energy = ==

- Reactions with –ΔG often require activation energy
  - e.g., burning of glucose
  - Must break existing bonds to get reaction started
- Catalysts lower activation energy

11.5 Enzymes

- Biological catalysts
  - Protein
  - RNA (ribozyme)
- Stabilizes temporary association between reactants (substrates) to facilitate reaction
  - Correct orientation
  - Stressing bonds of substrate

- Lower activation energy
- Not consumed (destroyed) in reaction

11.5.1 Carbonic anhydrase

- Important enzyme of red blood cells
- $CO_2 + H_2O \rightarrow H_2CO_3$ -> $HCO_3 + H^+$
- Carbonic anhydrase catalyzes 1st reaction
  - Converts water to hydroxyl
  - Orients the hydroxyl and CO_2

11.5.2 Enzyme mechanism

- One or more active sites which bind substrates (reactants)
  - Highly specific
- Binding may alter enzyme conformation, inducing better fit

11.5.3 Factors affecting enzyme activity

- Substrate concentration
- Product concentration
- Cofactor concentration
- Temperature
- pH
- Inhibitors
  - Competitive: bind to active site
  - Noncompetitive: bind to 2nd site, called allosteric site; changes enzyme conformation
- Activators
  - Bind to allosteric sites, increase enzyme activity

Cofactors

- Required by some enzymes
- Positively charged metal ions
  - e.g., ions of Zn[1], Mo, Mg, Mn
  - Draw electrons away from substrate (stress chemical bonds)
- Non-protein organic molecules (coenzymes)
  - E.g., NAD^+, $NADP^+$, etc.
  - Major role in oxidation/reduction reactions by donating or accepting electrons

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1   `http://en.wikibooks.org/wiki/Zinc`

11.6 ATP

- Adenosine triphosphate
- Major energy currency of cells, power endergonic reactions
- Stores energy in phosphate bonds
  - Highly negative charges, repel each other
- Makes these covalent bonds unstable
  - Low activation energy
- When bonds break, energy is transferred
- ATP $\rightarrow$ ADP + Pi + 7.3 kcal/mole

11.7 Biochemical pathways

- Metabolism: sum of chemical reactions in cell/organism
- Many anabolic and catabolic reactions occur in sequences (biochemical pathways)
- Often highly regulated

Evolution of biochemical pathways

- Protobionts or 1st cells likely used energy rich substrates from environment
- Upon depletion of a substrate, selection would favor catalyst which converts another molecule into the depleted molecule
- By iteration, pathway evolved backward

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12 Respiration: harvesting of energy

Glucose + O_2 → CO_2 + H_2O + ATP

12.1 Energy

- Energy is primarily in C-H bonds (C-O too)
- Chemical energy drives metabolism
  - Autotrophs: harvest energy through photosynthesis or related process (plants, algae, some bacteria)
  - Heterotrophs: live on energy produced by autotrophs (most bacteria and protists, fungi, animals)
- Digestion: enzymatic breakdown of polymers into monomers
- Catabolism: enzymatic harvesting of energy
- Respiration: harvesting of high energy electrons from glucose

12.2 Respiration

- Transfer of energy from high energy electrons of glucose to ATP
- Energy depleted electron (with associated H^+) is donated to acceptor molecule
  - Aerobic respiration: oxygen accepts electrons, forms water
  - Anaerobic respiration: inorganic molecule accepts hydrogen/electron
  - Fermentation: organic molecule accepts hydrogen/electron

12.3 Respiration of glucose

- $C_6H_{12}O_6$ + 6 O_2 → 6 CO_2 + 6 H_2O + energy
- ΔG = -720 kcal/mole under cellular conditions
- Largely from the 6 C-H bonds
- Same energy whether burned or catabolized
- In cells, some energy produces heat, most is transferred to ATP

12.4 Alternative anaerobic respiration

- Methanogens (Archaebacteria).
  - CO_2 is electron acceptor, forming CH_4

- Sulfur bacteria
  - SO_4 reduced to H_2S
  - Formation of H_2S set stage for evolution of photosynthesis (H_2S as electron donor before H_2O)
  - About 2.7 by, based on ratio of 32S/34S, where only biological processes produce 32S enrichment

12.5 Glycolysis overview

Glycolysis accounting

- Oxidation
  - Two electrons (one proton) are transferred from each G3P to NAD^+ forming NADH

2NADH

- Substrate level phosphorylation
  - G3P to pyruvate forms 2 ATP molecules

```
4 ATP (from 2 G3P)
```

–2 ATP (priming)

```
2 ATP (net gain)
```

Summary: The net input of glycolysis is 2 ATP molecules which are used to split one glucose molecule. The net yield of this step is 2 ATP and 2 pyruvate.

12.6 Regeneration of NAD^+

- Reduction of NAD^+ to NADH can deplete NAD^+ supply; it must be regenerated
- Two pathways, coupled to fate of pyruvate
  - With oxygen: enter electron transport chain, forming water (and ATP)
  - Without oxygen: fermentation
- lactate
- ethanol

12.7 Alcohol fermentation

12.8 Lactate formation

Either lactic acid[1] or alcohol can be formed as a result of anaerobic respiration in cells.

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1   http://en.wikipedia.org/wiki/Lactate

12.9 Krebs cycle: overview

- Matrix of mitochondrion
- Priming steps
  - Joining of acetyl-CoA to oxaloacetate
  - Isomerization reactions
- Energy extraction steps in Krebs cycle
  - Per glucose
- 6 NADH
- 2 FADH$_2$
- 2 ATP (from GTP)
- 4 CO$_2$

12.10 ATP production

- Chemiosmosis (Mitchell)
- H$^+$ (from NADH and FADH$_2$) is pumped against a gradient into the intermembranal space of the mitochondrion (creates voltage potential)
- Diffusion back into matrix through ATP synthase channels drives synthesis of ATP (ADP + Pi $\rightarrow$ ATP)
- ATP exits mitochondrion by facilitated transport

12.11 Evolution of aerobic respiration

- Preceded by evolution of photosynthesis (O2 needed; also, prior evolution of electron transport and chemiosmosis)
- High efficiency of ATP production compared to glycolysis
  - Fostered evolution of heterotrophs
  - Fostered evolution of mitochondria by endosymbiosis in eukaryotes

13 Photosynthesis

$$6 \ CO_2 + 6 \ H_2O \rightarrow C_6H_{12}O_6 + 6 \ O_2$$
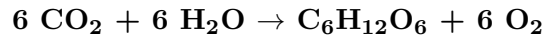
- One of most important reactions in history of life:
  - source of atmospheric O_2
  - ultimately led to aerobic respiration and eukaryotes
- Responsible for bulk of glucose production
- Early experiments showed that mass of plant must be derived from substances in the air, not the soil
- Experiments with isotopes showed that liberated oxygen comes from water
- Experiments also showed that light is essential but that some reactions (e.g., reduction of CO_2) continue in the dark
- Plants do two big, important things during photosynthesis: gain energy (absorb light) and build sugar (glucose).
- Photosynthesis can be divided into two series of chemical reactions: the light (light-dependent) reactions and the dark (light-independent) reactions. In light reactions, light is absorbed; in dark reactions, sugar is built.
- Occurs when plants, algae, and autotrophic bacteria absorb light energy and build glucose.

13.1 Light Reactions

- Part of the electromagnetic spectrum
- Consists of units of energy called photons
- Photons at UV end of spectrum have more energy than those at the red end
- Occur on the surface of thylakoid disks
- Chlorophyll and other plant pigments differentially absorb photons
  - Chlorophyll a: light to chemical energy
  - Chlorophyll b: accessory chlorophyll
  - Chlorophylls absorb primarily blue and red (green reflected back, hence the green color of plants)

13.1.1 Accessory pigments

- Chlorophyll is a major light gathering pigment
  - Absorbs light with considerable efficiency (i.e., retaining energy)
- Accessory pigments
  - Chlorophyll b
  - Carotenoinds
- capture light of wavelengths not captured by chlorophylls

- Confer other colors to plant leaves (autumn colors too)

Photosynthetic steps

- Primary photoevent: light photon captured by photosystem and energy transferred to electron donated by water
- Electron transport: excited electron is shuttled along imbedded series of electron carriers to proton pump and electron is transferred to acceptor
- Chemiosmosis: transport of protons back into chloroplast drives synthesis of ATP

13.1.2 The Even More Detailed Light Reactions

What the Light Reactions Do:

The light reactions of photosynthesis occur in chloroplasts in and on the thylakoid disks. During the light reactions, light energy charges up ATP molecules. More specifically, light turns the chloroplast into an acid battery, and this battery charges up ATP.

How the "Chloroplast-Battery" Charges ATP:

The **stroma** is the fluid inside of the chloroplasts, and it carries a negative charge. This means that it contains about a "gazillion" extra electrons. The solvent of stroma is water.

The fluid inside the thylakoid disks is positively charged because it contains a lot of hydrogen (H+) ions. The pH here is low, making the fluid very acidic. The solvent of thylakoid disk fluid is water.

A chloroplast acts like a battery, because it has separated a strong positive charge and a strong negative charge in two different compartments. Energy is released when H+ ions (free protons) flow from the inside of a thylakoid disk to the stroma. This is electrical energy, since it is a flow of charged particles.

The protons pass through special channels (made of protein) in the thylakoid membrane; this reaction is 'exothermic.' The energy that is given off is used to fuel this reaction (Pi is the phosphate ion):

```
ADP + Pi --> ATP
```
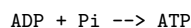
The proton can go to the negative stroma, but only if it uses its energy to charge up ATP. Since one reaction wants to go, and the other one doesn't, and since the first reaction releases energy and the second one absorbs energy, the two reactions are known to be 'coupled' together so that the first fuels the second. Of course, a special enzyme must be involved for this to happen.

Chlorophyll Molecules on a Thylakoid Disk:

Hundreds of chlorophyll molecules cover the surface of a thylakoid disk, making the disk green. The nonpolar "tails" of the chlorophyll molecule are embedded in the membrane of the thylakoid.

13.2 "Dark" reactions

- ATP drives endergonic reactions
- NADPH provides hydrogens for reduction of CO_2 to carbohydrate (C-H bonds)
- Occur in the stroma
- First step in carbon fixation

13.2.1 The Detailed Dark Reactions

What the Dark Reactions Do:

The dark reactions build sugar from carbon dioxide gas (CO2), water (H2O), and energy from ATP molecules that were charged up during the light reactions. The dark reactions occur in the stroma of a chloroplast. Dark reactions usually occur in the light, but they don't have to. They'll occur in the dark until the chloroplast's supply of ATP runs out (usually about 30 seconds).

The Calvin Cycle:

The Calvin Cycle is the fancy name for the metabolic pathway that builds sugar. This means that it involves a whole lot of chemical reactions, and it uses a lot of different enzymes to catalyze the reactions.

Carbon dioxide gas is stable, therefore the bonds that hold the carbon and oxygen atoms are strong. Therefore it takes a lot of energy to break the bonds and separate the carbon atoms from the oxygen atoms. The energy needed to do this comes from ATP molecules.

When inorganic carbon (like from CO2) is being added to an organic molecule (such as sugar), this is called carbon fixation.

It takes 2 complete turns of the Calvin Cycle to make a glucose molecule.

Some portions of this text is based on notes very generously donated by Paul Doerder, Ph.D., of the Cleveland State University. The detailed portions are not provided by Dr. Doerder.

How cells divide

13.3 Prokaryote cell division

- Binary fission
  - Doubling of cell contents, including DNA
  - Fission to divide contents
- Segregation of replicated genomes by growth of membrane between attachment points
- Partitioning of cytoplasmic components
- Escherichia coli
  - Capable of cell division every 20 minutes under optimal conditions (DNA in continuous state of replication)
  - Model organism of bacterial cell division

13.4 Bacterial DNA replication

- Replication follows rules of base pairing, with each polynucleotide chain serving as template for synthesis of its complement.
- Genetic evidence showed that the bacterial chromosome is circular long before there was corroborating physical evidence.

Eukaryotic chromosomes

- Discovered by Walther Fleming in 1882 in dividing cells of salamander larvae, following improvements in microscopes and staining technology
  - He called division mitosis (mitos = "thread")
- Chromosome number is constant in a species
  - Ranges from 2 to >500 (46 in human somatic cells)
  - Homologous pairs, one contributed by each parent
  - Change in number is cause and consequence of speciation
- Chromosome constancy and their precise division in mitosis and meiosis led biologists to postulate that they were carriers of hereditary information

13.5 Chromosome number

- 1N = number of chromosomes in gamete
- 1N = haploid chromosome number
- 2N = number of chromosomes in somatic cells (cells that are not egg or sperm)
- 2N = diploid
- Deviations from N or 2N are usually lethal in animals

Chromosome numbers

13.6 Eukaryotic chromosomes

- Consist of chromatin
  - DNA and associated proteins, mainly histones
  - Nucleosomal organization
  - Euchromatin: unwound chromatin, in basic nucleosomal configuration; genes available for expression
  - Heterochromatin: highly condensed except during replication
- Karyotype: array of chromosomes an individual possesses
  - Clinical importance (Down syndrome; cancer)
  - Evolutionary importance (speciation)

13.7 Chromosome organization

13.8 Human karyotype stained by chromosome painting

13.9 Chromosomes

- Homologous pairs
  - Inherited one from each parent
  - Identical in length and position of centromere
  - Contain identical or similar genes
  - Homologous pair = homologs
- Morphology
  - After replication, consist of two sister chromatids attached to a centromere

13.10 Human chromosomes

- Diploid number = 2n = 46 = 23 pairs of homologs
- Haploid number = 23 (gametes)
- Each replicated chromosome contains 2 sister chromatids = 92 chromatids

Cell cycle

- Growth and division cycle of cells
- Precisely controlled by biochemical and gene activity, except in cancer
- Phases
  - G1: primary growth phase
  - S: DNA replication; chromosome replication
  - G2: second growth phase; preparation for mitosis
  - M: mitosis; nuclear division
  - C: cytoplasmic division

13.11 Mitotic cell cycle

- Cells exiting the cell cycle are said to be in G0
- Cell cycle time varies with stages of life cycle and development, with G1 the most variable
- DNA replication occurs during S phase of the cell cycle following G1.

```
 - at this point the chromosomes are composed of two sister
chromotids connected by a common centromere.
```

13.12 Replicated human chromosomes

13.13 Mitosis

- Nuclear division
  - equational division of replicated chromosomes
  - chromatids move to opposite poles
- Continuous process
  - prophase
  - metaphase
  - anaphase
  - telophase
- Driven by motors and microtubules
- No change in chromosome number
  - N $\rightarrow$ N by mitosis
  - 2N $\rightarrow$ 2N by mitosis
- May be accompanied by cytokinesis

Kinetochore Microtubules attach to kinetochores. Metaphase

- Momentary alignment of chromosomes in center of cell

Anaphase

13.14 Plant mitosis

- Similar to animal mitosis
- New cell wall formed between cells from membrane partition

Cell cycle control

- Cell cycle events are regulated by protein complexes and checkpoints
- Discovered by microinjection of proteins in to eggs, by mutational analysis and by techniques of molecular biology

Molecular control of cell cycle: Cdk and cyclin

- Cyclin dependent protein kinase (Cdk)
  - Phosphorylate serine/threonine of target regulatory proteins
  - Function only when bound to cyclin
- Cyclin: short-lived proteins that bind to cdks

13.15 Controlling the cell cycle

- External signals initiate cell division in multicellular organisms
- Growth factors: extracellular regulatory signals
  - Usually soluble; bind to cell surface receptor

- Sometimes membrane bound, requiring cell-cell contact with receptor
- E.g., upon wound, platelets release PDGF which stimulates fibroblasts to enter cell cycle (exit G0), to heal wound

13.16 Cancer

- Unregulated cell proliferation
- Cancer cells have numerous abnormalities
  - >46 chromosomes
  - Mutations in proto-oncogenes
- Encode proteins stimulating the cell cycle
- May be regulated by phosphorylation
- Often over expressed in cancer cells
  - Mutations in tumor-suppressor genes
- Encode proteins inhibiting the cell cycle
  - Often bind to products of proto-oncogenes
- May be regulated by phosphorylation

13.17 Mutations and cancer

This text is based on notes very generously donated by Paul Doerder, Ph.D., of the Cleveland State University.

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14 Sexual reproduction

14.1 Sexual

- Exclusively eukaryotes
- Fusion of two haploid genomes
  - Fertilization (= syngamy)
  - Forms new individuals in multicellular organisms as result of fusion of egg and sperm
- Plants
- Animals
- Meiosis yields haploid genomes at some point in life cycle

14.2 Sexual life cycle

Typical animal life cycle

- Meiosis occurs in germ line cells in gonads producing haploid gametes
- All other cells are somatic cells
- Alternation of generations
- Sexual intercourse

14.3 Meiosis

- Gives rise to genetic variation
- Reduction division: 2n to n
- Preceded by one round of DNA (chromosome) replication
- Two rounds of nuclear (& usually cell) division
  - Meiosis I
- Synapsis of homologs
- Segregation of homologs
- Reduction division, 2n to n
  - Meiosis II
- No chromosome replication
- Segregation of sister chromatids
- Formation of 4 haploid (n) cells
- Two nuclear divisions, usually 2 cell divisions, only one round of replication
- Meiosis I
  - Prophase: synapsis and crossing over
  - Metaphase

- Anaphase: chromosome segregation
- Telophase
- Meiosis II (mitosis-like)
  - Prophase
  - Metaphase
  - Anaphase: sister chromosome segregation
  - Telophase

14.4 Prophase I: synapsis

- Complete alignment of replicated homologs
- Synapsis occurs throughout the entire length of a pair of homologs
- Key to chromosome segregation
- Synapsis, crossing over
- Subdivided into 5 continuous stages

14.5 Crossing over

- Reciprocal, physical exchange between nonsister chromatids
- Type of recombination; mixes maternal and paternal genes
- Visual evidence: chiasmata

14.6 Microtubules and anaphase I

- During prophase microtubules attach to kinetochores on one side of centromere
- The metaphase checkpoint insures proper attachment
- A phosphorylation event initiates motor activity and anaphase

14.7 Meiosis II

- Cytologically similar to mitosis
  - No preceding DNA replication
  - Chromatids segregate and move to opposite poles as chromosomes
  - 4 haploid cells produced
- In animals, these cells differentiate into gametes
- In plants and many other organisms, these cells divide by mitosis, followed some time later by gamete formation

14.8 Evolution of sex

- Asexual reproduction: all offspring genetically identical to parent

- Sex: recombination destroys advantageous combinations
- So why sex?
  - Many hypotheses
  - Effect repair of genetic damage?
- Much pachytene repair as well as gene conversion
- Some protists form diploid cells in response to stress
  - Recombination breaks up combinations of genes favoring parasites, thus reducing parasitism?

14.9 Consequences of sex

- Recombination: generates genetic diversity
  - Crossing over
  - Independent assortment
- Random fertilization
  - Qualities of gamete usually do NOT reflect qualities of genes enclosed in gamete

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15 Genetics

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16 Gregor Mendel and biological inheritance

Charles Darwin[1], for all he contributed to the science of biology, never knew about the mechanism by which living things inherit traits from previous generations, or how new traits arise.

As any schoolchild can tell you, this mechanism of interitance has since been found to be **DNA[2]**, or deoxyribonucleic acid. DNA allows for stable inheritance of traits: the code in each strand of DNA is replicated precisely through the pairing of basic units along each strand. The error rate in this replication is amazingly low; not even one base pair in a million matches out of sequence.

However, when even one base pair is added to a new strain of DNA in an order differently than in the parent chain, it can be the basis of a **mutation**. These changes in DNA sequences are the microscopic origin of changes in traits of all studied living things. Even the smallest difference in a strand of DNA can result in a change in traits that can cost the life of the organism. Mutations can produce proteins with a new or altered function. In humans, the example of Sickle cell anemia[3] is commonly given as its origin is a difference of only one base pair in a section of DNA that encodes red blood cells.

Individual sequences of DNA that encode for specific proteins are called **genes** and are the units of heredity. Each one has a set nucleotide[4], and together all of the genes (and some sequence of DNA that apparently do not code for any biologically important functions) together make up the entire **chromosome[5]**

16.1 Mendel

- Discovered principle of genetic segregation via numerous experiments utilizing pea plants
- Inferred the existence of genes through segregation of phenotypes
- Used quantitative methods: counted; ratios
- Work is model of scientific method
- In particular, observed the F2 progeny, which lead to the discovery of dominant and recessive traits
- Published work in 1866, went unnoticed
- In 1900 his scientific paper was "rediscovered"

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1   http://en.wikibooks.org/wiki/Charles%20Darwin
2   http://en.wikipedia.org/wiki/DNA
3   http://en.wikipedia.org/wiki/Sickle%20cell%20anaemia
4   http://en.wikipedia.org/wiki/nucleotide
5   http://en.wikipedia.org/wiki/chromosome

- Mendel is acknowledged as founder of Genetics
  - still used alphabet letters to designate genes
  - still refer to dominant and recessive genes
  - still refer to segregation of alleles in meiosis
  - principle of segregation applies to all sexually reproducing organisms; Mendel's results were immediately applied to humans in 1900

16.2 Mendel's experiments

- 1856, began experiments with the garden pea, Pisum sativum
- 1865, presented results to the Bruno natural history society, which he helped found
- 1866, published his results in proceedings of the society
- Naegeli encouraged Mendel to reproduce results in another species, which failed because the species did not undergo true fertilization
- discrete traits in Pisum sativum
  - pure-breeding lines
  - dominant/recessive alleles
- alleles are two alternate versions of a gene
  - gametes contain hybridized chromosomes that are formed during meiosis
  - homozygous has two of the same allele
  - heterozygous has two different alleles
- reciprocal F1 crosses (all exhibiting dominant phenotypes); F2; F3
- counted offspring, noted ratios
- inferred genotypes from phenotypes
- tested hypotheses with testcrosses
- attempted to repeat with another species

16.3 Mendel's seven pairs of traits

1. Seed form (round or wrinkled)
2. Cotyledon color (green or yellow)
3. Seed coat color (white or colored)
4. Pod form (inflated or constricted)
5. Pod color (green or yellow)
6. Flower position (axial or terminal)
7. Plant heights (tall or short)

16.4 Locus

- The location of a specific gene within a chromosome

16.5 Modern Y chromosome

Y-chromosome is the most evolved chromosome. Generally it is thought that if Y- chromosome is present in an individual then he will be male. But if mutation occurs at sex determining region or zinc factor then it will not code for testis determining factor, and results in normal female. This type of female's frequency is 1/250000.

16.6 Chromosome phenomena

- X-chromosome inactivation
  - Barr bodies
- Nondisjunction: failure of chromosome segregation at meiosis or mitosis
  - Results in $2N \pm 1$ chromosome number
- Trisomy $2N + 1$
  - Usually lethal. Trisomy 21 (Down) exception
- Monosomy $2N ** 1$
  - Lethal except XO
  - Usually maternal origin in humans

16.7 X-chromosome inactivation

In females, one X-chromosone is randomly switched off forming a Barr body.

16.8 Barr body

Dense region in the nucleus formed by the inactive X-chromosome.

16.9 Human genetic disorders

Down's Syndrome(Mongolism)

Down's Syndrome is usually produced by the nondisjunction of chromosome 21 during oogenesis and sometimes during spermatogenesis. The individual suffering from this type of syndrome has 47 chromosomes instead of the normal 46. The extra chromosome is not a sex chromosome but an autosome.

Most cases of mongolism were found to occur in children born by women in their forties. The affected children, called mongoloids, show mental retardation and have a shorter life expectancy. Their most prominent feature is the Mongolian folds in their eyes; hence, the term mongolism.

Klinefelter's Syndrome

When an XY-bearing sperm unites with an X-bearing egg, the resulting condition is called Klinefelter's Syndrome, or sexually undeveloped male. Individuals having the syndrome show the following characteristics:

- testes are small
- sperms are never produced
- breasts are enlarged
- body hair is sparse
- individuals are mentally defective

The same abnormal meiotic division may occur in females. They produce eggs with XX or no sex chromosomes. Such egg, when fertilized by a Y-bearing sperm, will not develop (YO). This is because YO is lethal--it wil cause death to the offspring.

17 DNA: The Genetic Material

17.1 DNA

DNA stands for *Deoxyribose Nucleic Acid.* That is, a nucleic acid with two sugars. DNA is the hereditary material of cells and is considered the blueprint of life. DNA is found in all kingdoms of life. Even most viruses have DNA. A molecule of DNA is chemically stable (it does not have a 2-prime alcohol group.)

When someone says DNA, they may be referring to one's genetic material on multiple levels: They may be speaking about a single deoxyribose nucleic acid molecule, a section of a double helix, a section of a chromosome, or one's entire hereditary composition.

- antiparallel
- Double helix
  - Semiconservative replication
  - Sequence of nucleotides encodes functional RNA or polypeptide

17.2 Historical perspective

- Mitosis and meiosis
  - Regular distribution of chromosomes suggested that they contain hereditary information
  - Bridges/Morgan, using Drosophila melanogaster showed that genes are on chromosomes (1910s)
- Hammerling: nucleus contains hereditary information (1930s)
- Griffith: transformation of bacteria (1928)
- Avery, MacLeod, McCarty: transforming substance is DNA (1944)
- Hershey, Chase: DNA is hereditary material of viruses (1952)
- Rosalind Franklin
- Watson and Crick: structure of DNA (1953)

17.3 Hershey-Chase Experiment

The Hershey and Chase experiment was one of the leading suggestions that DNA was a genetic material. Hershey and Chase used phages, or viruses, to implant their own DNA into a bacterium. They did two experiments marking either the DNA in the phage with a radioactive phophorus or the protein of the phage with radioactive sulfur. With the bacteria that was infected by the phages with radioactive DNA the DNA in the bacteria

was radioactive. In the bacteria that was infected with the radioactive protein the bacteria was radioactive, not the DNA. This proves that DNA is a genetic material and it is passed on in viruses.

17.4 DNA/RNA components

- Miescher: discovered DNA, 1869

17.4.1 Structure of DNA

DNA is in a double helix structure made up of nucleotides. The "backbone" of the double helix is composed of phosphates connected to a five carbon sugar called *deoxyribose*, . The "rungs" are composed of nitrogenous bases, Purines and Pyrimidines. Purines contain Adenine(A) and Guanine(G) and have two rings in their structures. Pyrimidines contain Cytosine(C) and Thymine (T) and have one ring in their structures.

17.5 Chemical structure of DNA

- Polynucleotide
  - Phosphodiester bonds between nucleotides
  - 5'-pGpTpCpGpTpApApTp-OH 3'
- Chargaff's rules, in DNA: equimolar amounts
  - A = T
  - G = C

17.6 3D structure of DNA

- James Watson[1] and Francis Crick[2] (1953[3])
  - Nucleotide
  - Keto and amino forms of bases
  - Chargaff's rules
  - X-ray crystallographic data (Rosalind Franklin[4])

17.7 Franklin

- X-ray diffraction of DNA crystals
- revealed regular pattern explained by antiparallel double helix

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1   http://en.wikipedia.org/wiki/James_D._Watson
2   http://en.wikipedia.org/wiki/Francis_Crick
3   http://en.wikipedia.org/wiki/1953
4   http://en.wikipedia.org/wiki/Rosalind_Franklin

DNA model

- Double helix of polynucleotides
  - antiparallel
  - 3'-5' phosphodiester bonds
- Base pairs held by hydrogen bonds
  - AT
  - GC
- There are about 10 base pairs per turn of helix
- model has predictive power
  - mode of DNA replication
  - encoding of genetic information

17.8 DNA replication

- Conservative model
  - One double helix of both old strands
  - One double helix of two new strands
- Dispersed
  - Each strand mixture of old new
- Semiconservative
  - Meselson-Stahl experiment confirmed its viability over the previous two
    - grew *E. coli* bacterium in a culture containing 15N (a heavy isotope of nitrogen)
    - bacterium assimilated the 15N into their DNA
    - a similar process was then done using 14N, a lighter isotope
    - following centrifugation, the densities were observed to be that of combined in the middle, and 14N on top, thereby confirming the semiconservative model

17.9 DNA replication

- Semiconservative
- New nucleotides added to 3' –OH
- Replication fork
  - Replication complex
- DNA polymerase
- Associated enzymes/proteins
  - Energy from phosphate bonds of triphosphate nucleotide substrates (dNTP)

17.10 DNA polymerases

- Prokaryotes, E. coli
  - 3 DNA polymerases
  - III is main enzyme for DNA replication
  - ~1000 nt/sec

- Eukaryotes
  - 6 DNA polymerases
- Add nucleotide to 3' –OH end
- All require primer, i.e., free 3' –OH

17.11 DNA replication complex

- Helicase "unzips" the DNA double helix
- Primase: synthesize RNA primer
- Single-strand binding proteins
- DNA gyrase (topoisomerase)

- DNA polymerase III
- DNA polymerase I (remove primer, fill gaps)

17.12 DNA replication

- $5' \rightarrow 3'$ replication
  - Nucleotide addition at 3' –OH
  - No exceptions
- New strands are oriented in opposite direction due to $5' \rightarrow 3'$ constraint
  - Leading strand: continuous replication
  - Lagging strand: discontinuous replication
    - contains multiple Okazaki fragments
- Joined by DNA ligase

17.13 DNA replication fork

- primer required by all DNA polymerases

17.14 Replication units

17.15 Replicon

A region of DNA that is replicated from a single origin.

17.16 What is gene?

- Garrod
  - "inborn errors of metabolism"
  - Alkaptonuria: enzyme deficiency

- Beadle and Tatum
  - One gene one enzyme
  - Genetic and biochemical analysis in Neurospora
- Today: gene is sequence of nucleotides encoding functional RNA molecule or the amino acid sequence of a polypeptide

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18 Gene expression

Flow of genetic information

- DNA → mRNA → polypeptide
- Transcription: DNA → mRNA
  - RNA polymerase
  - Nucleus in eukaryotes
  - Transcription also makes rRNA and tRNA
- Translation: mRNA → polypeptide
  - Ribosomes: protein and rRNA
  - Genetic code and tRNA

18.1 "Central Dogma"

18.2 The Genetic Code

- Triplet codon
  - 64 triplet codons (4^3)
  - Experimentally deciphered in 1961
- Nearly universal
  - Implies common ancestor to all living things
  - Minor exceptions: certain ciliates, mitochondria, chloroplasts
- Still evolving

18.3 Transcription

- RNA polymerase
  - NTP substrates
  - Synthesizes single stranded RNA complementary to template strand of DNA
  - New nucleotides to 3' end
- Begins at promoter site
  - no primer necessary
- Ends at terminator site
- Much posttranscriptional modification in eukaryotes

18.4 Transcription bubble

Promoter site

- Prokaryotes
  - -10 nt, TATA box
  - -35 nt, additional signal
- Eukayotes
  - -25, TATAAA box
  - Additional signals upstream
- Promoters may be strong or weak
- In eukaryotes, access to promoter depends upon state of chromatin coiling

18.5 Eukaryote mRNA

- Synthesized as pre-mRNA, processed in nucleus
- 5' end: GTP cap placed in inverted position
  - Essential for ribosome recognition
- 3' end: poly-A tail; non-templated addition of ~50-250 A nucleotides; stability
- Introns: intervening sequences removed

18.6 Translation

- Requires:
  - mRNA
  - tRNA
  - ribosomes
  - translation factors (various proteins)
- In prokaryotes, takes place on growing mRNA
- In eukaryotes, in cytoplasm on free ribosomes and RER
- AUG start codon to stop codon

18.7 Translation in bacteria

tRNA

- Transfer RNA
- Two important parts
  - Anticodon
- Hydrogen bonds with mRNA codon
  - 3' end
- Accepts amino acid (using energy of ATP)
- Aminoacyl-tRNA synthetase

18.8 Aminoacyl tRNA synthase

- Enzyme used to bind amino acid from the cytoplasm to tRNA, which then transfers that amino acid to the ribosome for polypeptide formation

18.9 Ribosome structure

18.10 Large ribosome subunit

18.11 Translation

- Initiation complex
  - Small ribosomal subunit
  - mRNA
  - fMet-tRNA (prokaryotes only; met-tRNA in eukaryotes)
  - Initiation factors
- Elongation
  - Ribosome
  - mRNA
  - tRNAs
  - Elongation factors

18.12 Initiation complex

18.13 Elongation, translocation

- incoming tRNA enters the A site
- rRNA catalyzes peptide bond formation. Note that growing peptide is attached to what was incoming tRNA at P site after translocation.
- empty tRNA leaves via E site; recycled
- A site ready for next charged tRNA

18.14 Introns/exons

- In eukaryotes, coding regions of gene may be interrupted by introns, noncoding regions of DNA (RNA)
- Introns
  - 22- >10,000 nt in length
  - 5' GU . . . 3' AG removal sequence
  - Not essential to genes
  - May constitute >90% of gene

- removed from pre-mRNA to form mRNA
- Exon: often codes for functional domain of protein
  - translatable mRNA

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19 Gene regulation

- Not all genes are expressed in a cell
- Gene expression can be turned on and off
- Multiple levels of regulation gene function
  - Transcription initiation
- State of chromatin
- Transcription factors
  - Post-transcriptional
- mRNA processing
- mRNA half-life
- Translational
- Post-translational
  - Protein modification

19.1 Transcriptional control

- State of chromatin
  - Euchromatin: transcriptionally active
  - Heterochromatin: transcriptionally inactive
  - Chemical modification of histones
  - Methylation of bases
- Transcription factors
  - Bind to DNA at promoter or other regulatory sites (enhancers)
- Recognize base sequence through major and minor grooves
  - Recruit RNA polymerase

19.2 DNA grooves

Categories of transcription factors in eukaryotes

- Helix-turn-helix
  - Two small "-helices
  - Fit into DNA groove
- Homeodomain
  - Highly conserved helical domains
  - ~60 amino acids
- Zinc finger motif
  - Zn atom bound
- Leucine zipper

- dimer

19.3 Regulatory proteins

- Activity may depend upon allosteric binding of small molecules
  - cAMP
  - Co-repressors
  - Inhibitors
- Binding to promoter region may "bend" DNA, making it accessible to other regulatory proteins

19.4 Lac operon of E. coli

- Single promoter region for cluster of genes
- Regulated and transcribed as a single unit
- Operons typical in prokaryotes
- Repressor: turns OFF gene expression

lac repressor

- Turns off transcription by blocking access by RNA polymerase
- repressor in activated by allosteric binding of lactose

Regulation in eukaryotes

- Both proximal (promoter) and distal (enhancer) to gene
- Typically transcription unit encodes a single polypeptide
- Promoter
  - TATA box
  - Other elements (regulatory sequences) may be present
- Enhancers
  - Work upstream, downstream, close, far from gene
  - Bend DNA

19.5 Alternative splicing

- Single transcript gives rise to 2 or more mature mRNAs
  - encode different polypeptides with shared domains
  - tissue and developmentally specific

This text is based on notes very generously donated by Paul Doerder, Ph.D., of the Cleveland State University.

(This Page was Last Edited December 2005)

20 Mutation

A mutation is a permanent change to an organism's genetic material (DNA[1] or RNA). Mutations are a rare but significant biological process, since they provide the variation on which evolution[2] acts and are also the source of cancer.

An organism's genetic material is made up of polymers (chains) of four different nucleotides[3], like a recipe book written in a language of only four letters. A mutation event is when the order of the nucleotides in DNA change, usually when the DNA is being copied.

Mutations come in a number of forms:

20.1 Point Mutations

Point mutations are all mutations which involve a single nucleotide. These come in the form of substitutions, insertions and deletions:

20.2 Substitution

Substitution Mutations: In substitution mutations, a nitrogenous base of a triplet codon of DNA is replaced by another nitrogen base or some derivative of the nitrogen base, changing the codon. The altered codon codes for a different amino acid substitution.The substitution mutations are of two types:

1.Transitions: It is the replacement of one purine in a polynucleotide chain by another purine(A by G or C by A) or one pyrimidine by another pyrimidine(T by C or C by T)

2.Transversions:A base pair substitution involving the substitution of a purine by pyrimidine or pyrimidine by a purine is called transversion.

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1   http://en.wikipedia.org/wiki/DNA
2   http://en.wikibooks.org/wiki/General%20Biology%2FEvolution
3   http://en.wikipedia.org/wiki/nucleotide

20.2.1 Insertion

20.2.2 Deletion

20.3 Larger mutations

Larger mutations which involve more than one nucleotide also include insertions and deletions, but can also include inversions, rearrangement of nucleotides and duplication of entire genes:

20.3.1 Inversion

20.3.2 Rearrangement

20.3.3 Gene/Exon Duplications

Transposition

Retrotransposition

20.4 Chromosomal mutations

Chromosomal mutations involve changes to entire chromosomes. These mutations are particularly rare:

20.4.1 Translocation

20.4.2 Fusion

20.4.3 Fission

20.4.4 Segmental Duplication

20.4.5 Chromosomal Duplication

20.4.6 Genome Duplication

20.5 Causes of mutations

20.6 Effects of mutations

Mutations can have a variety of different effects depending on the type of mutation, the significance of the piece of genetic material affected and whether the cells affected are germ-

line cells. Only mutations in germ-line cells can be passed on to children, while mutations elsewhere can cause cell-death or cancer.

Mutations can be classified by their effects:

20.6.1 Silent Mutation

Silent Mutations are DNA mutations that do not result in a change to the amino acid sequence or a protein. They may occur in a non-coding region (outside of a gene or within an intron), or they may occur within an exon in a manner that does not alter the final amino acid chain.

20.6.2 Frameshift

20.6.3 Missense Mutation

Missense mutations are types of point mutations where a single nucleotide is changed to cause substitution of a different amino acid. This in turn can render the resulting protein nonfunctional. Such mutations are responsible for diseases such as Epidermolysis bullosa.

20.6.4 Nonsense Mutation

20.7 Further reading

20.7.1 Books

- Jones, S. 1993. *The Language of the Genes.* Harper Collins ISBN 0006552439.

20.7.2 Websites

- Wikipedia: Mutation[4]
- http://www.evowiki.org/Mutation

20.8 Original notes

- "Rare" change in nucleotide sequence
  - Somatic vs germline
- only those in germline are heritable
  - Point mutation
- Single nucleotide change
  - Change in gene position

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4   http://en.wikipedia.org/wiki/Mutation

- Transposition
- Chromosomal rearrangement
- Mutagenic agents
- Raw material for evolutionary change

20.9 Point mutation

- Ionizing radiation
  - UV light induces thymine dimers
- Reparable
- Error during DNA synthesis
- Movement of transposons
  - McClintock
- Chemical mutagens
- May alter
  - Properties of promoter, enhancer
  - Amino acid sequence of polypeptide

20.10 Acquisition of genetic variability

- Mutation
- Sex (fusion of genomes)
- Recombination
  - Crossing over
- Reciprocal (may result in gene conversion)
- Unequal (gives rise to gene families)
  - Independent segregation
- Transposition by transposons
- Conjugation in bacteria
  - One way transfer from donor to recipient

20.11 Eukaryote genome

- Thousands of transposons
- Millions of transposon derived elements
  - LINES, SINES
- Above may constitute largest portion of genome
- Pseudogenes
- Tandem clusters (rRNA genes; nucleolus)
- Multigene families
- Single-copy genes (one copy per 1n)

20.12 Barbara McClintock

- Discovered transposons in perhaps greatest and ultimately most important intellectual endeavors in genetics
  - Maize
  - Worked alone
- Transposons: likely responsible for considerable evolution in eukaryotic genomes
- Likely origin of viruses

This text is based on notes very generously donated by Paul Doerder, Ph.D., of the Cleveland State University.

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21 Recombinant DNA technology

- Revolutionized modern biology
  - Ability to manipulate genes in vitro
- Hybrid genes, including combining genes of different species
- Detailed study of gene function
  - Determine nucleotide sequences of genes and their regulators (deduce amino acid sequences of proteins)
- Genome projects: complete nucleotide sequence of >40 genomes, including human
- Made possible by convergence of:
  - discovery of restriction enzymes
  - genetics of bacteria and their plasmids

21.1 Recombinant DNA technology

- Uses
  - Detailed study of gene function
- Homeostasis, response to stress
- Development (birth defects)
  - Evolution of genes informs on evolution of life
- Human betterment
  - Medicine
- Identification, treatment of genetic disorders
- Molecular medicine: from deduced amino acid sequences, design better drugs
  - Foods
- Improve crop yield, resistance to disease
- Improve nutritional value
  - Forensics
- DNA fingerprinting: guilt or innocence

21.2 Restriction endonucleases

Originally found in bacteria to prevent invasion of viral DNA, cuts double stranded DNA that is unmethylated, will not cut newly synthesized DNA since hemi-methylated, a product of semi-conservative replication of DNA

- sever phosphodiester bonds of both polynucleotide strands in order to combine foreign DNA
  - create restriction fragments (restriction digestion)
  - 5' phosphate and 3' –OH at ends

- usually nucleotide specific target sequence
  - 4-6 bp most common, the more bases, then the more specific for recombination
  - cuts in or near sequence
  - ends
- sticky=overhanging ends, 5' or 3'
- blunt ends - straight cut, will anneal with any other blunt end in the presence of high ligase
- Hundreds of know restriction endonucleases, usually named after the bacteria that it was found in
  - e.g. EcoR1, Alu1, BAM, HIND3

21.3 Restriction endonucleases

Gene cloning

- Cloning:
  - Restriction digestion of DNA
  - insertion of restriction fragment into cloning vector
- Bacterial plasmid
- Bacterial virus
- Yeast artificial chromosomes
- Transformation of bacteria with recombinant plasmid, virus
- Screening for clone of interest by using reporter genes or resistance upon exposure to anti-biotic

21.4 Uses of cloned gene

- Determine nucleotide sequence and deduce amino acid sequence from genetic code
  - Submit to GenBank (available on WWW)
- Manipulate gene to study function
  - In vitro
  - In vivo
- Transgenic (recombinant) organisms
- Knockout organisms
- Medical and commercial uses

21.5 Other molecular procedures

- Polymerase chain reaction (Mullis)
  - Amplifies target DNA without cloning
  - Target amount can be single molecule
  - Amplified DNA can be sequenced, cloned, etc.
- Southern blotting
  - Used to identify restriction fragments carrying particular gene

- Also used for DNA fingerprinting and RFLP analysis
- cDNA construction
  - Reverse transcription from mRNA template

21.6 RFLP(restriction fragment length polymorphism) analysis

- Basis of DNA fingerprinting using SNP - single nucleotide polymorphisms and repeats of DNA sequence
- Many uses
  - Criminal cases using multiple probes
  - Parentage
  - Species identification
  - Gene evolution
  - Species evolution

21.7 Sanger DNA sequencing

- Uses dideoxynucleotides (ddNTP), a template strand, DNA polymerase 1 (Also known as Kornberg enzymes) and dNTPs
  - Missing 3'-OH for nulceopjilic attack for elongation
  - DNA synthesis stops after one is incorporated into DNA fragment
  - ratio of ddNTP to dNTP determines likelihood of termination
- Manual method with 32P-labeled ddATP and 4 test tubes - ddATP, ddCTP, ddGTP, ddTTP
- Automated method using ddNTPs labeled with fluorescent dyes in capillary tube
  - Often done commercially

21.8 Automated sequencing

Typical machine

- • 2 hour sequencing run
  - 600-1000 bases per sample
  - multiple samples
- Up to 500,000 bases per day (12 hr)
- Data processed by computer
- In big labs, sequencing reactions also are automated

21.9 Genome projects

- Determine entire nucleotide sequence of genome

- >40 genomes sequenced
  - Helicobacter pylori
  - Escherichia coli
  - Saccharomyces cerevisiae
  - Caenorhabditis elegans
  - Drosophila melanogaster
  - Homo sapiens (first rough draft)
- Computer identifies all genes, based on properties of genes (e.g., start/stop codons, introns, etc.).

21.10 Biochips

- Microarray of DNA fragments, size of postage stamp; can be expensive, but has decreased in cost

Microarray chips contain wells of DNA that code for specific genes that uses the concept of hybridization with the gene of interest to see if a gene is expressed or is present.

- Designed to detect:
  - mutated genes (SNPs)
  - expressed genes
- Instant DNA profile ("GATTACA")

21.11 DNA chip controversies

- Medicine
  - Risks and informed consent for gene replacement therapy
  - Alteration of human gene pool
  - Parental choice
  - Privacy
- Genetically modified foods
  - Safety
  - Labeling
- Forensics
  - Mandatory tests
  - Reliability standards

21.12 Gene patenting

- Techniques to study and manipulate genes are patented (e.g., cloning and PCR)
- Should genes be patented?
  - Are they the intellectual property of the discoverer?
  - Don't they belong to all of us?
  - Should indigenous peoples be compensated for useful genes extracted from their local plants and fungi?

21.13 Stem cells

- Totipotent cells from early embryo
  - grow into any tissue or cell type
- Recombinant genes can be introduced
- Considerable use in analyzing gene expression in mice
- Possible therapeutic use in humans
- Very controversial

This text is based on notes very generously donated by Paul Doerder, Ph.D., of the Cleveland State University.

22 Classification of Living Things

22.0.1 Classification of Living Things & Naming

With so many flora and fauna on planet Earth, there must be a method to classify each organism to distinguish it from others so it can be correctly identified. Classification does not only apply to biology. For example, supermarkets and grocery stores organise their products by classifying them. Beverages may occupy one aisle, while cleaning supplies may occupy another. In science, the practice of classifying organisms is called **taxonomy** (Taxis means arrangement and nomos means law). The modern taxonomic system was developed by the Swedish botanist Carolus (Carl) Linneaeus (1707-1788). He used simple physical characteristics of organisms to identify and differentiate between different species.

Linneaeus developed a hierarchy of groups for taxonomy. To distinguish different levels of similarity, each classifying group, called **taxon** (pl. taxa) is subdivided into other groups. To remember the order, it is helpful to use a mnemonic device. The taxa in hierarchical order:

- Domain - Archea, Eubacteria, Eukaryote
- Kingdom - Plants, Animals, Fungi, Protists, Eubacteria (Monera), Archaebacteria
- Phylum
- Class
- Order
- Family
- Genus
- Species

The domain is the broadest category, while species is the most specific category available. The taxon Domain was only introduced in 1990 by Carl Woese, as scientists reorganise things based on new discoveries and information. For example, the European Hare would be classified as follows:

Eukaryote --> Animal --> Chordata --> Mammalia --> Lagomorpha --> Leporidae --> Lepus --> *Lepus europaeus.*

Binomial nomenclature is used to name an organism, where the first word beginning with a capital is the genus of the organism and the second word beginning with lower-case letter is the species of the organism. The name must be in italics and in Latin, which was the major language of arts and sciences in the 18th century. The scientific name can be also abbreviated, where the genus is shortened to only its first letter followed by a period. In our example, *Lepus europaeus* would become *L. europaeus'.*

Taxonomy and binomial nomenclature are both specific methods of classifying an organism. They help to eliminate problems, such as mistaken identity and false assumptions, caused by common names. An example of the former is the fact that a North American robin is

quite different from the English robin. An example of the latter is the comparison between crayfish and catfish, where one might believe that they both are fish when in fact, they are quite different.

Nomenclature is concerned with the assignment of names to taxonomic groups in agreement with published rules.

22.0.2 Eukaryotes & Prokaryotes

Recall that there are two basic types of cells: **eukaryotes** and **prokaryotes**.

Eukaryotes are more complex in structure, with nuclei and membrane-bound organelles. Some characteristics of eukaryotes are:

- Large (100 - 1000 μm)
- DNA in nucleus, bounded by membrane
- Genome consists of several chromosomes.
- Sexual reproduction common, by mitosis and meiosis
- Mitochondria and other organelles present
- Most forms are multicellular
- Aerobic

Prokaryotes refer to the smallest and simplest type of cells, without a true nucleus and no membrane-bound organelles. Bacteria fall under this category. Some characteristics:

- Small (1-10 μm)
- DNA circular, unbounded
- Genome consists of single chromosome.
- Asexual reproduction common, not by mitosis or meiosis.
- No general organelles
- Most forms are singular
- Anaerobic

22.0.3 The Three Domains

The three domains are organised based on the difference between eukaryotes and prokaryotes. Today's living prokaryotes are extremely diverse and different from eukaryotes. This fact has been proven by molecular biological studies (e.g. of RNA structure) with modern technology. The three domains are as follows:

Archea (Archeabacteria) consists of archeabacteria, bacteria which live in extreme environments. The kingdom Archaea belongs to this domain.

Eubacteria consists of more typical bacteria found in everyday life. The kingdom Eubacteria belongs to this domain.

Eukaryote encompasses most of the world's visible living things. The kingdoms Protista, Fungi, Plantae, and Animalia fall under this category.

22.0.4 The Six Kingdoms

Under the three domains are six kingdoms in taxonomy. The first two, **Plants** and **Animals**, are commonly understood and will not be expounded here.

Protista, the third kingdom, was introduced by the German biologist Ernst Haeckel in 1866 to classify micro-organisms which are neither animals nor plants. Since protists are quite irregular, this kingdom is the least understood and the genetic similarities between organisms in this kingdom are largely unknown. For example, some protists can exhibit properties of both animals and plants.

Fungi are organisms which obtain food by absorbing materials in their bodies. Mushrooms and moulds belong in this kingdom. Originally, they were part of the plant kingdom but were recategorised when they were discovered not to photosynthesise.

Eubacteria are bacteria, made up of small cells, which differ in appearance from the organisms in the above kingdoms. They lack a nucleus and cell organelles. They have cell walls made of peptidoglycan.

Archae (or Archaebacteria) are bacteria which live in extreme environments, such as salt lakes or hot, acidic springs. These bacteria are in their own category as detailed studies have shown that they have unique properties and features (ex. unusual lipids that are not found in any other organism)which differ them from other bacteria and which allow them to live where they live. Their cell walls lack peptidoglycan.

22.0.5 Origins of Diversity

The diversity in our planet is attributed to diversity within a species. As the world changed in climate and in geography as time passed, the characteristics of species diverged so much that new species were formed. This process, by which new species evolve, was first described by British naturalist Charles Darwin as **natural selection**.

For an organism to change, genetic mutations must occur. At times, genetic mutations are accidental, as in the case of prokaryotes when they undergo asexual reproduction. For most eukaryotes, genetic mutations occur through sexual reproduction, where meiosis produces haploid gametes from the original parent cells. The fusion of these haploid gametes into a diploid zygote results in genetic variation in each generation. Over time, with enough arrangement of genes and traits, new species are produced. Sexual reproduction creates an immense potential of genetic variety.

One goal of taxonomy is to determine the evolutionary history of organisms. This can be achieved by comparing species living today with species in the past. The comparison in anatomy and structure is based on data from development, physical anatomy, biochemistry, DNA, behaviour, and ecological preferences. The following are examples of how such data is used:

- Anatomy:

Although a horse and a human may look different, there is evidence that their arm structures are quite similar. Their arms' sizes and proportions may be different, but the anatomical structures are quite similar. Such evidence reveals that animals in different taxa may not

be that different. Biological features from a common evolutionary origin are known as **homologous**.

- Development

- Biochemistry:

Biochemical analysis of animals similar in appearance have yielded surprising results. For example, although guinea pigs were once considered to be rodents, like mice, biochemistry led them to be in their taxon of their own.

22.0.6 Phylogeny, Cladistics & Cladogram

Modern taxonomy is based on many hypotheses' of the evolutionary history of organisms, known as **phylogeny**. As with the Scientific Method, scientists develop a hypothesis on the history of an animal and utilise modern science and technology to prove the phylogeny.

Cladistics is a classification system which is based on phylogeny. Expanding on phylogeny, cladistics is based on the assumption that each group of related species has one common ancestor and would therefore retain some ancestral characteristics. Moreover, as these related species evolve and diverge from their common ancestor, they would develop unique characteristics. Such characteristics are known as **derived characteristics**

The principles of phylogeny and cladistics can be expressed visually as a **cladogram**, a branching diagram which acts as a family (phylogenetic) tree for similar species. A cladogram can also be used to test alternative hypotheses for an animal's phylogeny. In order to determine the most likely cladogram, the derived characteristics of similar species are matched and analysed.

22.0.7 Classification of Living Things Practice Questions

1. If taxonomists had to select an existing kingdom to reclassify, which of the six would most likely be chosen? Why?

2. Complete the following without consulting external sources:

a) The species *caudatum* is in the family *Paramecidae*. What would be the binomial name of this organism?

b) Give the abbreviation of the binomial name.

3.

a) Irish moss belongs to the genus *Chondrus*. The name for this species is *crispus*. Give the binomial name.

b) Give the abbreviation of the binomial name.

4. Humans and chimpanzees are alike. Which of the following data would most accurately prove this correct?

a) biochemistry

b) DNA

c) appearance

d) development

e) A, B, C

5. Which of the following is out of order?

a) Kingdom --> Phyllum --> Class

b) Class --> Family --> Order

c) Family --> Order --> Genus

d) Genus --> Species

e) A, C

f) A, B, D

g) B, C

6. A taxonomist discovers Organism A and Organism B and wishes to classify them. Which of the following choices is the most informative?

a) Both organisms are brown.

b) Both organisms have a tail.

c) Both organisms have ears.

d) Both organisms are nocturnal.

7. DNA analysis is usually done using DNA found in a cell's mitochondria, and not in a cell's nucleus. From your knowledge of mitosis, explain why this is so.

1. Arachbacteria 3.a) Chondrus crispus b) C. cripus 4. B 5. G 6. B

22.1 Introduction

Viruses are the smallest biological particle (the tiniest are only 20 nm in diameter). However, they are not biological organisms so they are not classified in any kingdom of living things. They do not have any organelles and cannot respire or perform metabolic functions. Viruses are merely strands of DNA or RNA surrounded by a protective protein coat called a **capsid**. Viruses only come to life when they have invaded a cell. Outside of a host cell, viruses are completely inert.

Since first being identified in 1935, viruses have been classified into more than 160 major groups. Viruses are classified based on their shape, replication properties, and the diseases that they cause. Furthermore, the shape of a virus is determined by the type and arrangement of proteins in its capsid. Viruses pathogenic to humans are currently classified into 21 groups.

Viruses can also attack bacteria and infect bacterial cells. Such viruses are called **bacteriophages**.

22.2 Viral Replication

As previously stated, viruses are not a biological life form so they cannot reproduce by themselves. They need to take over a functioning eukaryotic or prokaryotic cell to replicate its DNA or RNA and to make protein coat for new virus particles.

In order to enter a cell, a virus must attach to a specific receptor site on the plasma membrane of the host cell. The proteins on the surface of the virus act as keys which fit exactly into a matching glycoprotein on the host cell membrane. In some viruses, the attachment protein is not on the surface of the virus but is in the capsid or in the envelope.

There are two forms of viral replication: the **lytic cycle** and the **lysogenic cycle**.

22.2.1 Lytic Cycle

1. Attachment: The virus binds to specific receptors on the host cell.
2. Entry: There are two ways in which a virus can enter cells. Firstly, the virus can inject its nucleic acid into the host cell. Secondly, if a virus is contained in an envelope, the host cell can phagocytosise the entire virus particle into a vacuole. When the virus breaks out of the vacuole, it then releases its nucleic acid into the cell.
3. Replication: The virus's nucleic acid instructs the host cell to replicate the virus's DNA or RNA.
4. Assembly: New virus particles are assembled.
5. Lysis and Release: The virus directs the production of an enzyme which damages the host cell wall, causing the host cell to swell and burst. The newly formed virus particles are now released.

22.2.2 Lysogenic Cycle

1. Attachment: Similar to Lytic Cycle
2. Entry: Similar to Lytic Cycle
3. Incorporation: The viral nucleic acids is not replicated, but instead integrated by genetic combination (crossing over) into the host cell's chromosome. When integrated in a host cell this way, the viral nucleic acid as part of the host cell's chromosome is known as a **prophage**.
4. Host Cell Reproduction: The host cell reproduces normally. Subsequent cell divisions, daughter cells, contain original father cell's chromosome embedded with a prophage.
5. Cycle Induction: Certain factors now determine whether the daughter cell undergoes the lytic or lysogenic cycle. At any time, a cell undergoing the lysogenic cycle can switch to the lytic cycle.

The reproduction cycle of viruses with RNA and no DNA is slightly different. A notable example of a RNA-based virus is HIV, a retrovirus.

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Retrovirus reproductive cycle

1. The retrovirus force RNA into cell, by either one of the two methods of entry (See above).
2. In the retrovirus are reverse transcriptase enzymes, which catalyses the synthesis of a DNA strand complementary to the viral RNA.
3. Reverse transcriptase catalyses a second DNA strand complementary to the first. With these two strands, the double-stranded DNA can be created.
4. DNA is then incorporated into the host cell's chromosomes. Similar to the concept of a prophage, this incorporated DNA is called a **provirus**. However, the provirus never leaves the host cell, unlike a prophage.
5. The infected host cell undergoes the lytic or lysogenic cycle.

22.3 Viral Genome

The genome of a virus consists of DNA or RNA, whose size and configuration vary. The entire genome can exist as a single nucleic acid molecule or several nucleic acid segments. Also, the DNA or RNA may be single-stranded or double-stranded, and either linear or circular.

Not all viruses can reproduce in a host cell by themselves. Since viruses are so small, the size of their genome is limiting. For example, some viruses have coded instructions for only making a few different proteins for the viruses' capsid. On the other hand, the human genome codes for over 30,000 different proteins. Therefore, the lack of coded instructions cause some viruses to need the presence of other viruses to help them reproduce themselves. Such viruses are called **replication defective**.

Lastly, it is worthy to note that 70% of all viruses are RNA viruses. As the process of RNA replication (with enzymes and other organelles of the host cell) is more prone to errors, RNA viruses have much higher mutation rates than do DNA viruses.

22.4 Viruses Practice Questions

1. As the name implies, the Tomato Spotted Wilt Virus targets tomatoes. Would it be possible for this virus to target other fruits as well? Explain.
2. If a DNA and a RNA virus both infected somatic cells, which virus would be more difficult to detect?
3. Many people have had cold sores, which are caused by infection with the herpes simplex virus. One characteristic of cold sores is that after a period of inactivity, they will reappear many times during the course of a person's life. Which cycle would the herpes simplex virus undergo?
4. Chicken pox is a common, non-fatal disease usually acquired in adolescence and caused by the varicella zoster virus. In adulthood, many people suffer from shingles, an altered form of the varicella zoster virus. Which cycle would the varicella zoster virus have undergone?
5. Would an antibiotic work for a person suffering from a cold of flu? Explain.

Answers to Viruses Practice Questions[1]

For Eubacteria, please visit General Biology/Classification of Living Things/Eubacteria[2].

22.5 Archaea

- Proposed as separate group from (eu)bacteria by Carl Woese
  - based on structure and metabolic pathways
  - inhabit extreme environments
  - unique branched lipids in membrane
- Share traits with both eukaryotes and eubacteria, e.g., RNA polymerase, introns
- Biochemically diverse
- Economically important
  - Taq polymerase used in PCR

22.5.1 Types

- Methanogens
- Halophiles
- Thermophiles

Underground bacteria

- Metabolism
  - built around inorganic energy sources
- e.g., basalt reacts with H_2O to release hydrogen which is catalytically combined with CO_2 to form carbohydrate (akin to photosynthesis)
- may result in deposit of minerals
- Unresolved problems
  - Did bacteria move downward from surface or did they first evolve there, protected from harsh surface conditions?
  - Could bacteria be ejected into space in rocks?

22.6 Prokaryote evolution

- Tentative, subject to change
- Derived largely from molecular systematics (rRNA sequences)
- Note: most bacteria can't be cultured, thus hard to study! (Studied by PCR of water/soil samples)

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1   http://en.wikibooks.org/wiki/%2FAnswers

2   http://en.wikibooks.org/wiki/General%20Biology%2FClassification%20of%20Living%20Things%2FEubacteria

22.7 Domains of life: characteristics

This text is based on notes very generously donated by Paul Doerder, Ph.D., of the Cleveland State University.

22.8 Introduction

Out of the six kingdoms, Protista is the most diverse. This is the kingdom of organisms with strange, atypical characteristics. In essence, this kingdom is designated for organisms which do not belong in any other kingdom. The majority of protists are microscopic.

22.9 Classification of Protists

There are three phyla of protists, based on their type of nutrition.

1. **Protozoa** (animal-like protists) are heterotrophs that ingest or absorb their food.

2. **Algae** (plant-like protists) are autotrophs they get nutrition from photosythesis.

3. **Slime moulds** and **water moulds** (fungus-like protists) are also heterotrophs, like protozoa.

22.10 Protozoa

As heterotrophs, protozoa scavenge materials from their surroundings. Others are predators which actively hunt or ambush small organisms such as bacteria and other protozoa for a source of nutrition. Protozoa can be parasitic as well; they may live inside larger organisms, like humans. Most protozoa live as single cells, although a few form colonies.

Protozoa are generally difficult to identify due to their varied shape. They may appear as jelly-like blobs, spherical sunbursts, or a flattened leaf. Tiny blood parasites may be only 2 μm long. On the other hand, shell-covered marine may be 5 cm or more in diameter.

Furthermore, different protozoans have their own complex life cycles. The complexity has led certain organisms to be mistakenly classified for other species.

Nevertheless, protozoa can move, and so, they are classified based on their methods of locomotion.

Characteristics of Protozoa :

- About 30,000 species known
- About 10,000 species are pathogenic, including some of the worst human diseases
- heterotrophic
- highly variable in form and life cycle

- mostly unicellular

- range in size from 0.005 mm to 50 mm

- lack cell walls

they love environment and each other..............

22.11 Algae

Algae are much simpler than protozoa. They are aquatic and contain chlorophyll. Algae can exist as a single cell or as giant seaweeds 60 m in length. Formerly, algae were classified as plants but this was incorrect as algae lack parts of true plants: leaves, stems, roots, xylem, and phloem. Since algae belong in the kingdom Protista, algae is a broad term used to denote all aquatic eukaryotes which photosynthesise; algae can differ in size and shape as well.

There are six phyla of algae:**chlorophytes (green algae)**, **phaeophytes (brown algae)**, **rhodophytes (red algae)**, **chrysophytes (diatoms)**, **pyrrophytes (dinoflagellates)**, and **euglenophytes (euglenoids)**.

22.11.1 Chlorophytes

Chlorophytes resemble plants the most. Like plants, their cell walls contain cellulose and they store food in reserve as starch. Chlorophytes can be unicellular or multicellular. Most chlorophytes use flagellae for some locomotion.

22.11.2 Phaeophytes

Phaeophytes are nearly all multicellular marine organisms, which are known to us as seaweeds. They have cell walls composed of cellulose and alginic acid (a substance similar to pectin). The cellulose and alignic acid help to retain water and prevent seawood from drying out when exposed to air at low tide.

Since phaeophytes live in a tidal environment, they have large, flat fronds (a large leaf) which can withstand pounding by waves. Their bases strongly anchor the algae to the rocky seabed and prevent them from being washed out to sea. Phaeophytes are usually found in areas of cold water.

22.11.3 Rhodophytes

Rhodophytes are typically found in warmer seawater, and are more delicate and smaller than brown algae (phaeophytes). Rhodophytes are also able to grow at deeper depths in the ocean, since red algae absorb green, violet, and blue light, the wavelengths of which penetrate the deepest below the water surface. They also have mucilaginous material to resist drying.

22.11.4 Chryosophytes

Chryosophytes are the most abundant unicellular algae in the oceans. They are also one of the biggest components of plankton, a free-floating collection of microorganisms, eggs, and larvae. As photosynthetic organisms, they produce a significant amount of atmospheric oxygen.

The reproduction cycle of chryosophytes is particularly interesting. Note that diatoms reproduce both asexually and sexually. Since diatoms have a rigid cell wall with an outer layer of silica (found in sand and glass), the daughter cells produced by mitosis must fit inside the original cell wall. Therefore, each generation of diatoms is smaller than the one before. The reduction in size continues until the diatoms produce sexually, producing a zygote which eventually grows to the original size as it matures.

22.11.5 Pyrrophytes

Pyrrophytes are unicellular, photosynthetic, and mostly aquatic. They have protective coats composed of stiff cellulose. They are more easily identifiable, due to the presence of two flagellae. The longer flagellae propels the dinoflagellate, while the second shorter, flatter flagellae functions as a rudder.

Some species of pyrrophytes are **zooxanthellae**. Since they lack cellulose plates, they make their home in coral reefs and animals, such as sea anemones, and molluscs. In returning the favour of sheltering them, dinoflagellates provide carbohydrates to their host through photosynthesis. This is why there are nutrient-rich coral reefs in malnutritions water.

A negative aspect of pyrrophytes is that under certain conditions, species of dinoflagellates reproduce rapidly to form a **harmful algal bloom (HAB)**, known as a red tide if dinoflagellates are the cause. Such pyrrophytes can produce toxins which may injure or kill wildlife, and additionally any consumers of contaminated wildlife.

22.11.6 Euglenophytes

Like pyrrophytes, euglenophytes are small unicellular freshwater organisms with two flagella. They are mainly autotrophic or heterotrophic, depending if they have a red, light-sensitive structure called an **eyespot**.

22.12 Slime molds & Water molds

There are two phyla of slime moulds and one phylum of water moulds.

22.12.1 Oomycotes (Water moulds)

Oomycotes are filamentous organisms which resemble fungi, in that they live as saprotrophs. Oomycotes differ from other moulds with the presence of spores and their sexual life cycle.

22.12.2 Myxomycotes (Plasmodial slime moulds)

Myxomycoties are visible to the naked eye as tiny slug-like organisms which creep over decayed and dead matter. This streaming blob containing many nuclei is called a **plasmodium**.

22.12.3 Acrasiomycotes (Cellular slime moulds) and its reproductive cycle

Acrasiomycotes exist as individual amoeboid cells with one nucleus each. When in unfavourable conditions, each acrasiomycete cell gathers together to form a **pseudoplasmodium**.

Reproductive Cycle:

1. One acrasiomycete cell joins with others to form a pseudoplasmodium.

2. The pseudoplasmodium shrinks and forms a smaller plasmodium.

3. The plasmodium migrates to a suitable environment.

4. The plasmodium develops a sporangia, where original parental nuclei has divided by meiosis into haploid spores to be germinated.

5. When favourable conditions arise, the spores germinate and are carried away by animals or the wind.

6. Cycle repeats.

22.13 Protists Practice Questions

1. Which of the following adjectives describe the major food source of protozoa?

 a) chemoautotrophic

 b) photoheterotrophic

 c) chemoheterotrophic

 d) heterotrophic

 e) A, C, D

 f) C, D

2. The protozoan *Giardia lamblia* can inhabit a human body's intestinal tract and cause gastroenteritis.

a) Give the abbreviated binomial name of this protozoan.

b) Would the relationship between this protozoan and human being be mutualistic, commensalistic, or parasitic?

3. Found in many products, such as Petri dishes, agar is made from mucilagnious material in seaweed. Of the six phyla of algae, which phyllum/phyla would agar be made from?

4. Which of the following adjectives describe the major food source of Euglenophytes without an eyespot?

a) photoautotrophic

b) photoheterotrophic

c) chemoautotrophic

d) chemoheterotrophic

e) B or C

f) C or D

5. Can coral reefs exist in nutrient-poor areas? Explain.

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23 Multicellular Photosynthetic Autotrophs

23.1 Plants

- Multicellular
- Cellulose cell walls
- Chlorophylls a and b
- Develop from embryophyte
- Alternation of generations
- Major food source for terrestrial life
- Atmospheric O2 and CO2 balance
- Coal deposits
- Intimate association with mycorrhizal fungi
- >250,000 species (˜500,000?)
- Taxonomy
  - State of flux
- DNA sequencing
- Developmental studies
  - Division (old literature) = phylum (new literature)
  - ˜12 phyla, 9 of which are vascular plants

23.2 Plant phyla

Phyla are 12 groupings

23.3 Plant evolution

- Evolved from green algae, likely related to charophytes
- Evidence
  - DNA sequences
  - homologous chloroplasts: chlorophyll b and beta-carotene; thylakoids in grana;
  - Cellulose in both groups; also peroxisomes
  - Mitosis and cytokinesis similar
  - Sperm ultrastructure

23.3.1 Terrestrial adaptations

- Stomata: pores in leaves for exchange of gases; prevent desiccation
- Secondary metabolites:
  - cuticle: waxy coating to prevent H2O loss
  - lignin: hardens wood
  - sporopollenin: resistant polymer; coats pollen
  - predator defenses
- Embryonic development
  - gametangia in early plants
  - spores; seeds
- Mycorrhizae
- Water/food conducting systems

23.4 Plant phylogeny

23.5 Plant life cycles

- Alternation of generations
- Sporophyte
  - diploid
  - produces spores in sporangia
- Gametophyte
  - develops from spore
  - haploid
  - produces gametes in gametangia
- Haplodiplontic life cycle

23.6 Moss life cycle

23.7 Vascular plants

- Most have roots
- Aerial shoot systems
- Vascular tissue
  - xylem: water, mineral transport
  - phloem: food transport
- Lignin
- Branched sporophyte is dominant stage
  - amplified production of spores
  - evolution of complex plant bodies
- Dominated Carboniferous (360 my)

23.8 Vascular plant life cycles

- Homosporous (single type of spore)
- Heterosporous (two types of spore)

23.9 Pterophyta (ferns)

- Non-seed plant
- Sporophyte conspicuous (vascular tissue)
- Rhizome: ground stem, roots
- Fronds: leaves
- Sori: clusters of sporangia
- Motile sperm require external water for fertilization
- Originated in Devonian, 350 my

Tree fern Fern life cycle

23.10 Non-seed plants, continued

- Lycophyta: club mosses
  - E.g., Lycopodium ("ground pine")
  - Many species became extinct 270 my, once dominant (coal formations)
  - Gametophyte non-photosynthetic, nourished by fungi
- Arthrophyta: horsetails
  - Equisitum
  - Some fossil forms (300 my) were tree-size (coal)
  - Photosynthetic stems, no leaves
  - Silica deposits in epidermal cells

23.11 Seed plants

- 1st appeared in Devonian, 360 my
- Seed develops from ovule, protects embryo
  - withstands drought
  - dispersal is enhanced
  - no immediate need for water for germination
- Heterosporous
  - male gametophyte: arise from microspores
  - female gametophyte: arise from megaspores in ovule in ovary
- Two groups
  - gymnosperms
  - angiosperms

Multicellular Photosynthetic Autotrophs

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plant
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23.12 Sporophyte/gametophyte

23.13 Megasporangium (nucellus)

- Key to seed development
- Nucellus: solid, fleshy, surrounded by integuments derived from sporophyte (seed coat)
- Entire structure called ovule
- Flower may have many ovules

23.14 Pollen

- Develop from microspores, become male gametophyte
- Protected by sporopellenin
- In most plants, sperm lack flagella (loss)
- Many mechanisms to transport pollen
  - wind
  - insects, birds,

23.15 Gymnosperms

- "naked seed"
  - ovule not fully enclosed by sporophyte at time of pollination
- Conifers, cycads, gnetophytes, Ginkgo
- Small, inconspicuous plants to giants like sequoia
- Conifers: to carry cones fv
  - male cones, Female conesvv
  - evergreen

23.16 Pine life cycle

23.17 Other Coniferophyta

- Cycadophyta: cycads
  - tropical, subtropical
  - flagellated sperm
- Gnetophyta
  - e.g., Ephedra, Mormon Tea
- Ginkgophyta: Ginkgo

- only one surviving species
- diocious (separate % and &trees)

23.18 Other gymnosperms

23.19 Angiosperms

- Flowering plants, Anthophyta
  - monocots- single seed leaf (grasses, lilies, etc.)
  - dicots- two seed leaves (roses, pulses, maples)
- More specialized xylem (water transport)
  - vessel elements
  - fiber cells
- Fossils date to 130 my
- Animal (e.g., insect) coevolution

23.19.1 Monocots vs dicots

23.20 Earliest angiosperm

- What is earliest angiosperm?
- Recent analysis of nucleotide and amino acid sequences suggests that Amborella, a tropical plant found only on the island of New Caledonia, is closest relative to flowering plants

23.21 Angiosperm flower

Insert non-formatted text hereInsert non-formatted text here

23.22 Angiosperm life cycle

This text is based on notes very generously donated by Paul Doerder, Ph.D., of the Cleveland State University.

23.23 Introduction

Although you may not recognise fungi, they are just as prevalent as plants and animals. Their spores are in the air which we breathe, fungi allow us to make bread, and mushrooms (a type of fungi) are eaten by us. A few types of fungi are unicellular. For example, yeasts live as individual oval or cylindrical cells. However, the majority of fungi live are multicellular. Their bodies are composed of **hyphae**, a network of fine filaments. In a mushroom, the

hyphae are densely packed so it is difficult to see the individual structures when a mushroom is eaten. However, a mushroom is only a specialised reproductive part of the whole fungus. The main part of the fungi is underground in a whole web of hyphae, called a **mycelium**.

In the mycelium, each fungal cell is separated from each other by a septum. Each fungal cell may have one or more nuclei and remains connected to the mycelium because the septa are porous, allowing cytoplasm to flow through the hyphae and fungal cell walls, made of a hard material called **chitin**. Some fungi do not have septa, and they appear to be large, branching, multinucleate cells.

23.24 Nutrition

Fungi are **saprophytes**. When they find a source of food (e.g. dead wood, orange peel) , they decompose it and digest it. The enzymes break down larger organic molecules in the substrate into smaller molecules. These smaller molecules diffuse into the fungus, where they are used to allow growth and repair.

Fungi which feed on living cells are parasitic. For example, athlete's foot grows on the human foot. These kinds of fungi produce hyphae called **haustoria**, which can penetrate host cells without immediately killing them.

However, they are friendlier species of fungi. Many fungi live symbiotically with plants or animals. For example, most trees have fungi living in close contact with their roots. In this relationship, known as a **mycorrhiza**, there are many benefits:

- Growing around the plant roots and often entering plant cells, the hyphae absorb minerals from the soil and release them in the roots. The fungi gets its source of food (organic nutrients) while delivering food to the plant.

- The mycelium here would increase the surface area, thus the absorptive surface, of the plant roots.

- The fungal cells help to maintain air and water flow in the soil around the roots.

- The fungi may prevent other potentially pathogenic fungi to attack the tree.

23.25 Fungal Reproduction

Fungi can reproduce in two ways. Firstly, they make asexually produce through **fragmentation**. This occurs when pieces of hyphae are broken off, which then grow into new mycelia.

The second method is by spores. Spores are lightweight structures and windblown designed to be transported over long distances and by many mediums, such as on the bodies of insects and birds. They are additionally light enough to be blown away for hundreds of kilometers. Spores may be asexual and sexual. Their sexual properties can be analysed to classify the four phylla of fungi.

23.26 Types of Fungi

23.26.1 Zygospore Fungi (Zygomycetes)

This phyllum includes bread moulds and other saprotrophs. Comparable to bacteria, this phyllum prefers asexual reproduction over sexual reproduction.

1. Two haploid hyphae of opposite types, also known as **mating strain +** and **mating strain -**, combine and fuse together.

2. **Plasmogamy**, the union of the two parent hyphae, occurs and results in the creation of a heterokaryotic (n + n) **zygosporangium or zygospore**. Note that the zygospore is NOT diploid yet; the haploid nuclei are simply clumped together.

3. Immediately, a thick wall develops around the zygospore to protect it from drying and other hazards. The zygospore becomes dormant.

4. When conditions are favourable, the zygospore absorbs water and undergoes **karyogamy** (n + n = 2n), where the haploid nuclei contributed by the two parents fuse to produce diploid zygosporangia.

5. The now diploid zygosporangium then undergoes meiosis to form haploid sporangia.

6. Through asexual reproduction of fungi (See above for more information), the spores from the sporangia germinate and grow into new mycelia.

7. Back to step #1.

23.26.2 Club Fungi (Basidiomycetes)

This phyllum increases mushrooms and shelf fungi. In many ways, the reproduction stages of this phyllum is similar to that of zygomycetes.

1. Two haploid hyphae of opposite types, also known as **mating strain +** and **mating strain -**, combine and fuse together.

2. Plasmogamy takes place, and a dikaryotic mycelium forms. The dikaryotic mycelium grows faster then the haploid parental mycelia.

3. Environmental factors cause the dikaryotic mycelium to form compact masses which develop into **basidiocarps**, short-lived reproductive structures. An example is the mushroom.

4. The basidiocarp gills are lined with terminal dikaryotic cells called **basidia**, which then undergo karyogamy.

5. The basidia are now diploid. They undergo meiosis to develop haploid **basidiospores**, a term referring to a basidiomycete's spores.

6. Still remaining on the basidiocarp, the haploid basidiospores eject, fall from the basidiocarp, and are dispersed by the wind when mature.

7. In a favourable environment, the basidiospores germinate and grow into short-lived haploid mycelia.

8. Back to Step #1.

23.27 Key Terms

synapomorphy

23.28 Introduction

What makes an animal an animal?

If animals are a monophyletic taxon, then animals should be able to be defined by synapomorphies, (shared, derived characteristics). Ideally, we would NOT define this or any taxon using symplesiomorphies (shared ancestral or primitive characteristics) or homoplastic characters (the independent evolution of similarity, or "convergent evolution"). See pages 654 - 656 and Fig. 32.6 in your text to review these concepts. As you consider the characteristics listed below, ask yourself whether or not each is a synapomorphy.

23.29 Characteristics of an Animal

- There is no one universally accepted definition of an animal. The following treatment follows your text, beginning on page 876.
  - Animals:
    - Are multicellular, heterotrophic eukaryotes ...
    - Lack the distinctive cell walls of plants & fungi
    - Share unique characteristics ...
    - Share certain reproductive characteristics ...
    - Other commonly used definitions ...

Animals are multicellular heterotrophic eukaryotes

- Unfortunately, none of these traits is exclusive to animals:
  - Plants, fungi, and some algae are multicellular.
  - Many bacteria, protists, and all fungi are heterotrophic.
  - Everything other than bacteria and archaea are eukaryotic.
- Moreover, all three of these characteristics also apply to fungi.
  - However, there is a difference here between animals and fungi. Animals generally take in their food through ingestion, or eating and swallowing something. Fungi are absorptive heterotrophs; they secrete their digestive enzymes onto their food, and then absorb the resulting nutrients.

Animals share unique characteristics

- Only animals have muscle tissue and nervous tissue.
- Only animals have collagen, a structural protein
- Only animals have the following types of intercellular junctions: (See pages 135 - 139, Figure 7.15 in your text for more information on these junctions.)

- Tight junctions (sealing function)
- Desmosomes (anchoring function)
- Gap junctions (communication function)

Animals share certain reproductive characteristics

- Most animals reproduce sexually, with the diploid stage dominating.
- In most animals, a small, haploid, flagellated, motile sperm fertilizes a larger, haploid, nonmotile egg to form a diploid zygote.
- Mitotic division of the zygote yields a blastula stage, followed by a gastrula stage. A synapomorphy? This feature could be another "unique characteristic" shared by animals.
- Development may be direct to adult form, or there may be a sexually immature stage (or stages) that are morphologically & ecologically distinct from the adult called a larva (plural: larvae).

Other commonly used definitions or characterizations

- It is surprisingly difficult to find two texts that agree on a precise definition of an animal. Here are a few perspectives from some other texts.
  - Animals are multicellular eukaryotes; they are chemosynthetic heterotrophs that ingest their food.
  - Animals are motile, though many are secondarily sessile. Gametes usually are produced in multicellular sex organs, and the zygote passes through embryonic stages that include a blastula.
  - Animals are organisms that are multicellular, with more than one type of cell. They are heterotrophic. They reproduce sexually (at least sometimes), with a zygote formed from two different haploid gametes. They go through a developmental stage called a blastula.
  - Animals are not photosynthetic, have no cell wall, and no hyphae or mycelia. (What would a cladist think of this definition of the taxon Animalia?)

What kinds of animals are there?

- Kingdom Animalia generally is recognized to have approximately 30 phyla ...
  - There is relatively little dispute over the number of phyla recognized; however, the phylogenetic relationships among the phyla are hotly debated.
  - Molecular techniques for assess similarity based on nucleotide sequences in nucleic acids are providing valuable new perspectives on this question.
- Remember that two animals in different phyla generally are considered to be more different from each other than are animals within one phylum (e.g., nematodes are more different from annelids than humans are from sharks).

This text is based on notes very generously donated by Ralph Gibson, Ph.D.[1] of the Cleveland State University[2].

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[1]   http://bgesweb.artscipub.csuohio.edu/faculty/gibson.htm
[2]   http://www.csuohio.edu

23.30 Introduction to animal phyla

There currently are almost 40 recognized phyla.

Phylum — Number of Species — Common Name

- Placozoa[3] — 1
- Monoblastozoa[4] — 1
- Rhombozoa[5] — 50
- Orthonectida[6] — 50
- **_Porifera_**[7] — 9,000 — sponges (figures)
- **_Cnidaria_**[8] — 9,000 — corals (figures)
- Ctenophora[9] — 100 — comb jellies
- **_Platyhelminthes_**[10] — 20,000 — flatworms (figures)
- Nemertea[11] — 900 — ribbon worms (figures)
- **_Rotifera_**[12] — 1,800 — rotifers (figures)
- Gastrotricha[13] — 450 — gastrotrichs
- Kinorhyncha — 150 — kinorhynchids
- **_Nematoda_** — 12,000 — roundworms (figures)
- Nematomorpha — 230 — horsehair worms
- Priapula — 15
- Acanthocephala — 700 — (figures)
- Entoprocta — 150
- Gnathostomulida — 80
- Loricifera — 35
- **_Annelida_** — 15,000 — segmented worms (figures)
- Sipuncula — 250 — peanut worms (figures)
- Echiura — 135
- Pogonophora — 145 — beard worms
- Vestimentifera — 8 — beard worms
- **_Arthropoda_** — 957,000 — arthropods (figures)
- Onychophora — 80
- Tardigrada — 400 — water bears
- Pentastomida — 95 — tongue worms
- **_Mollusca_** — 100,000 — molluscs (figures)
- Phoronida — 15
- Ectoprocta — 4,500 — sessile zooids
- Brachiopoda — 335 — lampshells

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3   http://en.wikibooks.org/wiki/Placozoa
4   http://en.wikibooks.org/wiki/Monoblastozoa
5   http://en.wikibooks.org/wiki/Rhombozoa
6   http://en.wikibooks.org/wiki/Orthonectida
7   http://en.wikibooks.org/wiki/Porifera
8   http://en.wikibooks.org/wiki/Cnidaria
9   http://en.wikibooks.org/wiki/Ctenophora
10   http://en.wikibooks.org/wiki/Platyhelminthes
11   http://en.wikibooks.org/wiki/Nemertea
12   http://en.wikibooks.org/wiki/Rotifera
13   http://en.wikibooks.org/wiki/Gastrotricha

- *Echinodermata* — 7000 — echinoderms (figures)
- Chaetognatha — 100 — arrow worms (figures)
- Hemichordata — 85 — acorn worms
- *Chordata* — 50,000 — chordates (figures)

23.31 Phylum Porifera

Figure 6 Sponges

Name means "pore-bearing".

This phylum consists of the sponges. The number of species is estimated to be between 5,000 and 10,000. All are aquatic and almost all are marine.

Animals in this phyla have no true tissues, which means, for example, that they have no nervous system or sense organs. Although sponges are multicellular, they are described as being essentially at a cellular level of organization. They are sessile as adults, but have a free swimming larva.

Their bodies are porous. They are filter feeders; water flows in through many small openings, and out through fewer, large openings. They have inner and outer cell layers, and a variable middle layer. The middle layer often is gelatinous with spiny skeletal elements (called spicules) of silica or calcium carbonate, and fibres made of spongin (a form of collagen).

Choanocytes are flagellated cells lining the inside of the body that generate a current, and trap and phagocytize food particles.

Their cells remain totipotent, or developmentally flexible: they can become any type of cell at any point in the sponge's development. This allows for the great regenerative power sponges have.

Sponges are an ancient group, with fossils from the early Cambrian (ca. 540 mya) and possibly from the Precambrian. Sponges often are abundant in reef ecosystems. They somehow are protected from predators (spicules? bad taste?).

Many organisms are commensals of sponges, living inside them. Some sponges harbor endosymbiotic cyanobacteria or algae (dinoflagellates, a.k.a. "zooxanthellae").

23.32 Phylum Cnidaria

See text pages 886 - 889.

Name comes from the Greek knide- meaning "nettle".

This phylum They have one opening, which serves as both mouth and anus. The body wall has an outer ectoderm, an inner endoderm, and a variable undifferentiated middle layer called mesoglea or mesenchyme that may be jelly-like. The mesoglea is **NOT** considered to be true mesoderm and so the Cnidaria are described as diploblastic. Tentacles usually extend from the body wall around the mouth/anus.
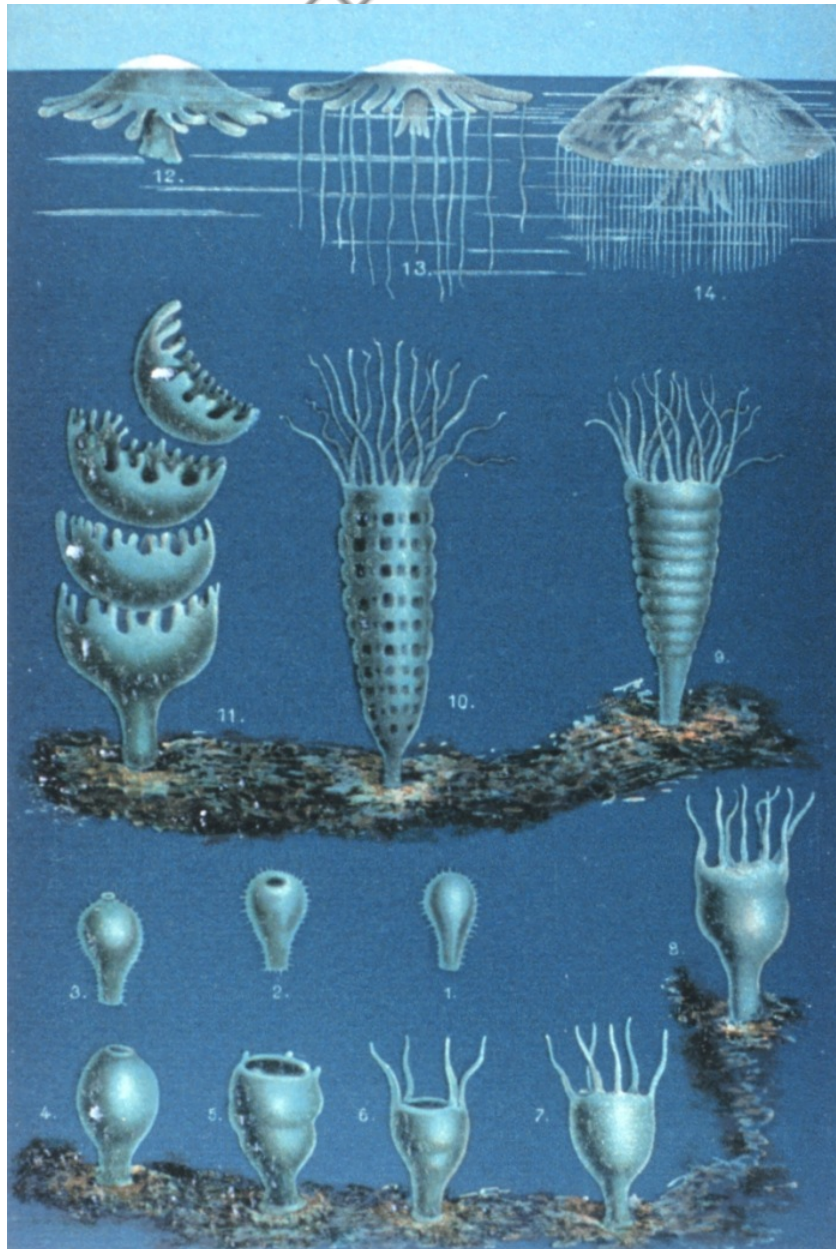
Figure 7   Jellyfish Development

There are two basic body plans: the polyp[14] and the medusa[15]. The polyp is sessile and attaches to substrate by the aboral end (i.e., the end away from the mouth). The medusa ("jellyfish") is a floating form, and looks like an upside-down version of the polyp. Some cnidarians only have the polyp stage, some have only the medusa stage, and others have both.

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14   http://en.wikipedia.org/wiki/polyp
15   http://en.wikipedia.org/wiki/Medusa%20%28biology%29

The typical life cycle of a cnidarian involves what is called "alternation of generations": an alternation between an asexual polyp stage and a sexual medusa stage.

The tentacles are armed with cnidae (or nematocysts), small intracellular "harpoons" that function in defense and prey capture. When fired, the cnidae deliver a powerful toxin that in some cases is dangerous to humans. The phylum is named after the cnidae.

Cnidarians have no head, no centralized nervous system, and no specialized organs for gas exchange, excretion, or circulation. They do have a "nerve net."

Many cnidarians have intracellular algae living within them in a mutualistic symbiotic relationship (Dinoflagellates = zooxanthellae). This combination is responsible for much of the primary productivity of coral reefs.

There are three main classes in the phylum

- Class Hydrozoa[16] (hydras and Portugese man-of-war are well-known but atypical examples of this Class)
- Class Scyphozoa[17] (jellyfish)
  - The medusa stage is dominant and the polyp stage often is reduced.
- Class Anthozoa[18] (sea anemones, most corals)
  - No medusa (jellyfish) stage, so sexual reproduction occurs in the polyp stage in this group. The polyps also can reproduce asexually, which is how individual "corals" grow.

23.33 Phylum Platyhelminthes

See text pages 890 - 893.

Name means "flat worm"

Most members of this phylum are parasitic (flukes and tapeworms), but some are free living (e.g., planaria). There are about 20,000 species.

They are dorsoventrally compressed (i.e., "flat").

Animals in this phylum are acoelomate, triploblastic, bilaterally symmetrical, and unsegmented. Platyhelminths have a simple anterior "brain" and a simple ladder-like nervous system. Their gut has only one opening. Flatworms have NO circulatory or gas exchange systems. They do have simple excretory/osmoregulatory structures (protonephridia or "flame cells").

Platyhelminths are hermaphroditic, and the parasitic species often have VERY complex reproductive (life) cycles.

There are four main classes of platyhelminths:

- Class Turbellaria[19] (mostly free living flatworms, e.g., planaria)

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16   http://en.wikipedia.org/wiki/Hydrozoa
17   http://en.wikipedia.org/wiki/Scyphozoa
18   http://en.wikipedia.org/wiki/Anthozoa
19   http://en.wikipedia.org/wiki/Turbellaria

- Class Monogenea[20] (parasitic flukes)
- Class Trematoda[21] (parasitic flukes, e.g., liver fluke and the human blood fluke, Schistosoma)
- Class Cestoda[22] (tapeworms)
  - Cestodes are endoparasitic in the gut of vertebrates. They do not have a mouth or digestive system.

23.34 Phylum Rotifera

See text page 900

The Rotifers. The name means "wheel bearing," a reference to the corona, a feeding structure (see below).

They are triploblastic, bilaterally symmetrical, and unsegmented. They are considered pseudocoelomates.

Most less than 2 mm, some as large as 2 - 3 mm.

Rotifers have a three part body: head, trunk foot. The head has a ciliary organ called the corona that, when beating, looks like wheels turning, hence the name of the phylum. The corona is a feeding structure that surrounds the animal's jaws. The gut is complete (i.e., mouth & anus), and regionally specialized. They have protonephridia but no specialized circulatory or gas-exchange structures.

Most live in fresh water, a very few are marine or live in damp terrestrial habitats. They typically are very abundant. There are about 2,000 species.

Parthenogenesis, where females produce more females from unfertilized but diploid eggs, is common. Males may be absent (as in bdelloid rotifers) or reduced. When males are present, sexual and asexual life cycles alternate. Males develop from unfertilized haploid eggs and are haploid. Males produce sperm by mitosis which can fertilize haploid eggs, yielding a diploid zygote that develops into a diploid female. Sexual reproduction occurs primarily when living conditions are unfavorable.

Most structures in rotifers are syncytial ("a mulitnucleate mass of protoplasm not divided into separate cells," or "a multinucleated cell") and show eutely (here, "constant or near-constant number of nuclei").

23.35 Phylum Nematoda

See text pages 894 - 895.

Name from the Greek for "thread".

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20   http://en.wikipedia.org/wiki/Monogenea
21   http://en.wikipedia.org/wiki/Trematoda
22   http://en.wikipedia.org/wiki/Cestoda

This phylum consists of the round worms. There are about 12,000 named species but the true number probably is 10 - 100 times this!

These animals are triploblastic, bilaterally symmetrical, unsegmented pseudocoelomates. They are vermiform, or wormlike.

In cross-section, they are round, and covered by a layered cuticle (remember this cuticle !!). Probably due to this cuticle, juveniles in this phylum grow by molting. The gut is complete. They have a unique excretory system but they lack special circulatory or gas-exchange structures. The body has only longitudinal muscle fibers. The sexes are separate.

Nematodes can be incredibly common, widespread, and of great medical and economic importance. They are parasites of humans and our crops. They can live pretty much anywhere.

In one rotting apple, there can be up to 90,000 nematodes, and in one tablespoon of coastal mud, there can be 236 species of nematodes!

Nematodes can be free living or important parasites of our crops, or of humans and other animals. They have become very important in development studies, especially the species Caenorhabditis elegans, presumably due to its small size and constancy of cell number (eutely - 959 cells in C. elegans).

23.36 Phylum Annelida

See text pages 906 - 909.

Name means "ringed", from the Greek annulatus.

This phylum consists of earthworms, leeches, and various marine worms given many different names (e.g., sand worms, tube worms). There are about 12,000 - 15,000 species.

Animals in this phylum are triploblastic, bilaterally symmetrical, segmented coelomates. Development is typically protostomous. They have a complete circulatory system, and a well-developed nervous system. Typically, each segment has paired epidermal "bristles" (setae or chaetae).

Most are marine but they are successful occupants of almost anywhere sufficient water is available. They can be free living, parasitic, mutualistic, or commensalistic.

Major advances of this phylum include the true coelom, segmentation, both longitundinal and circular muscles, a closed circulatory system and, for most, a more advanced excretory system (metanephridia).

There are three main classes of Annelids

- Class Oligochaeta (earthworms)
- Class Polychaeta (marine worms)
- Class Hirudinea (leeches)

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23.37 Phylum Arthropoda
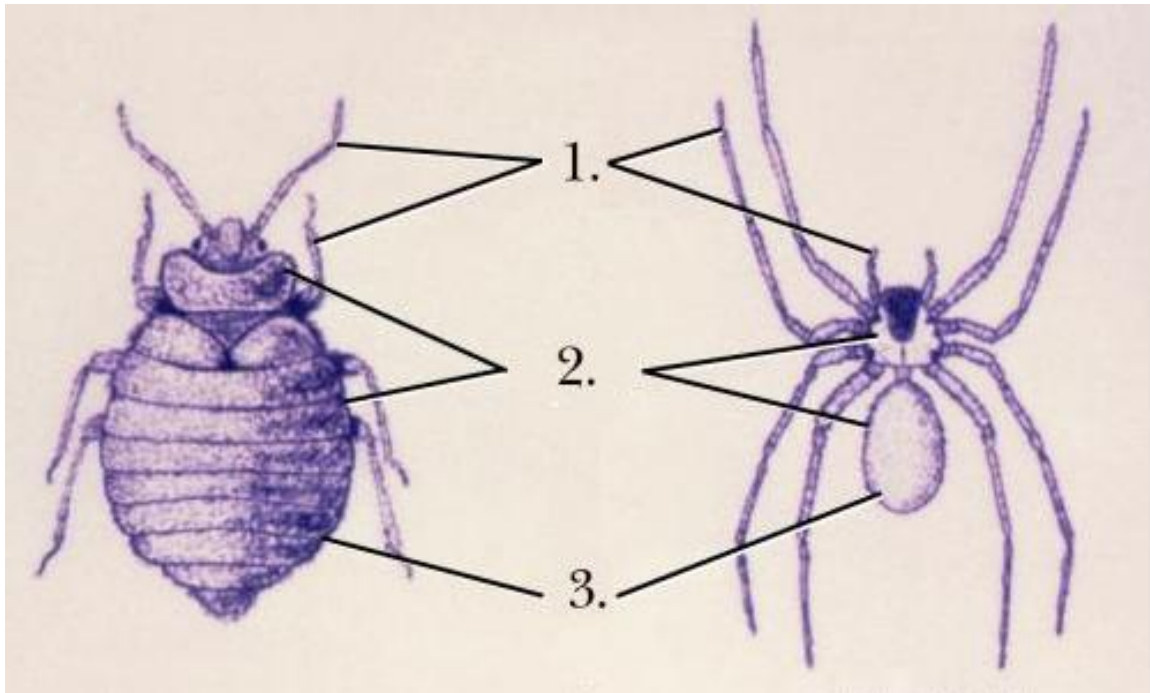
Figure 8   Arthropods

Name means "jointed feet".

This phylum consists of spiders, ticks, mites, insects, lobsters, crabs, and shrimp, and is the largest of all the phyla. So far, over 1 million species have been named, and it is likely that the true number out there is 10 - 100 times greater.

This phylum also includes the extinct trilobites, which were prevalent in the Paleozoic era. Because of their exoskeletons, these animals fossilized well and over 4000 species have been named.

These animals are triploblastic, bilaterally symmetrical, segmented, protostome coelomates. The coelom is generally reduced to portions of the reproductive and excretory systems. They have an open circulatory system.

The most notable advancement of this phylum is a rigid exoskeleton. It has major implications in these organisms' locomotion, flexibility, circulatory systems, gas exchange systems, and growth. It also was partially responsible for the ability of the arthropods to move on to land.

There are several major groupings of arthropods:

Major subgroups include:

- The chelicerates (eurytperids, horseshoe crabs, scorpions, spiders, ticks) have clawlike feeding appendages. They lack antennae and usually have simple eyes.
- The Trilobites...they get their own grouping

- The uniramians (centipedes, millipedes, insects) have one pair of antennae and unbranched (uniramous) appendages.
- The crustaceans (crabs, shrimp, lobsters, barnacles and many others) have two pairs of antennae and branched (biramous) appendages.

Major Classes Include

- Class Arachnida (mites, scorpions, spiders, ticks)
- Class Diplopoda (millipedes)
- Class Chilopoda (centipedes)
- Class Insecta (insects)
- Class Crustacea (crabs, crayfish, lobsters, shrimp)

23.38 Phylum Mollusca

See text pages 900 - 905.

Name means "soft".

This phylum consists of snails, slugs, bivalves, chitons, squids, octopus, and many others. About 110,000 species

All molluscs have a similar body plan:

- A muscular foot, usually used for movement.
- A visceral mass, containing most of the internal organs.
- A mantle, a fold of tissue that drapes over the visceral mass and secretes the shell, if present.
- Most have a radula, or a rasping organ to scrape food.

Molluscs are bilaterally symmetrical, or secondarily asymmetrical. They are coelomates, but the coelom generally has been greatly reduced; the main body cavity is a hemocoel. Development is typically protostomous. The gut is complete with marked regional specialization. Large, complex, metanephridia (excretion).

Many molluscan life cycles include a trochophore larva. This stage also is characteristic of annelids.

There are several major classes of molluscs:

- Class Polyplacophora (chitons)
- Class Gastropoda (snails, slugs, nudibranchs)
- Class Bivalvia (clams, mussels, scallops, oysters)
- Class Cephalopoda (squids, octopuses, chambered nautiluses)

23.39 Phylum Echinodermata

Name means "spiny skin"

This phylum consists of sea stars, brittle stars, sea urchins, and sea cucumbers.

Echinoderms are mostly sessile or very slow moving animals.

As adults, they are radially symmetrical, but in the larval stage, they are bilaterally symmetrical. They are considered deuterostomes.

Echinoderms are unique in that they have a water vascular system composed of a system of fluid-filled canals. These canals branch into the tube feet, which function in feeding, locomotion, and gas exchange.

There are six major classes of echinoderms:

- Class Asteroidea (sea stars)
- Class Ophiuroidea (brittle stars)
- Class Echinoidea (sea urchins, sand dollars)
- Class Crinoidea (sea lilies)
- Class Holothuroidea (sea cucumbers)

23.40 Phylum Chordata

Name means "the chordates", i.e., these animals have a notochord at some stage in their lifecycle.

This phylum consists of tunicates, lancelets, and the vertebrates.

There are four major features that characterize the phylum Chordata.

- A notochord, or a longitudinal, flexible rod between the digestive tube and the nerve cord. In most vertebrates, it is replaced developmentally by the vertebral column. This is the structure for which the phylum is named.
- A dorsal hollow nerve cord which develops from a plate of ectoderm that rolls into a tube located dorsal to the notochord. Other animal phyla have solid nerve cords ventrally located. A chordate nerve cord splits into the central nervous system: the brain and spinal cord.
- Pharyngeal slits, which allow water that enters through the mouth to exit without continuing through the entire digestive tract. In many of the invertebrate chordates, these function as suspension feeding devices; in vertebrates, they have been modified for gas exchange, jaw support, hearing, and other functions.
- A muscular, postanal tail which extends posterior to the anus. The digestive tract of most nonchordates extends the length of the body. In chordates, the tail has skeletal elements and musculature, and can provide most of the propulsion in aquatic species.

Chordates have a segmented body plan, at least in development. This segmentation evolved independently from the segmentation of annelids.

Three subphyla make up the phylum Chordata:

- Subphylum Urochordata (tunicates): the adults are enclosed in a tunic made of a carbohydrate much like cellulose. They squirt water out of an excurrent siphon. Urochordates are characterized by errant (mobile and active) larvae and sessile adults. All are filter feeders. The only "chordate" characteristics retained in adult life are the pharyngeal slits.

Larval urochordates look more like adult cephlochordates & adult vertebrates than adult urochordates.

- Subphylum Cephalochordata: Cephalochordates are known as lancelets because of their blade-like shape; they are also known as amphioxus. They are marine animals and usually live on the bottom, but can swim.
- Subphylum Vertebrata (vertebrates) ...

Formally, the phyla Urochordata and Cephalochordata are considered invertebrates.

23.40.1 Subphylum Vertebrata

Vertebrata refers to the presence of vertebrae and a vertebral column.

This subphylum includes most of the animals with which most people are familiar.

Vertebrates show extreme cephalization.

The notochord generally is replaced by the cranium & vertebral column in adults.

Neural Crest Cells

Later in development, these give rise to many cells of the body, including some cartilage cells, pigment cells, neurons & glial cells of the peripheral nervous systems, much of the cranium, and some of the cells of the endocrine system.

Some scientists would like to classify the neural crest as the fourth germ layer.

Neural crest cells come from the dorsal edge of the neural plate, thus ectoderm.

24 Chordates

The phylum Chordata includes three subphyla. These include vertebrates and invertebrate chordates.

24.1 Characteristics

Notochord: the rod-shaped supporting axis found in the dorsal part of the embryos of all chordates, including vertebrates

Flexible, non-collapsible rod dorsal to the gut/coelom and below the nervous system, hydrostatic, fluid wrapped in tough connective tissue. As bone does not compact, muscles tensed on one side result in movement instead of shortening the animal. This allows much better locomotion than do cilia for larger animals in water, a crucial victory for later success.

Pharyngeal slits: Slits in the pharynx originally used to gather food, water enters the mouth, passes through pharynx and out gill-like slits, passing through a cavity called an antrium and then outside. In humans, present only in embryo.

Dorsal nerve cord: A neural tube dorsal to the notochord

Postanal tail: Elongation of the body and notochord, nerve cord and muscles past anus into tail, early locomotive function led to success.

Non-synapomorphic characteristics (not limited to chordates):

- bisymmetrical (bilateral symmetry)
- segmented muscles and bones

24.2 Subphylum Urochordata

The tunicates are located in this subphylum. Along with the subphylum Cephalochordata, these two subphyla make up the invertebrate chordates. Only the tunicate larvae have notochords, nerve cords, and postanal tails. Most adult tunicates are sessile, filter-feeders which retain their pharyngeal slits. Adult tunicates also develop a sac, called a tunic, which gives tunicates their name. Cilia beating within the turnicate cause water to enter the incurrent siphon. The water enters the body, passes through the pharyngeal slits, and leaves the body through the excurrent siphon. Undigested food is removed through the anus. Tunicates are hemaphrodites and can reproduce asexually through budding.In urochordates notochord is confined to larval tail.These lack cranium. These have an open type of circulatory system.Excretion is by neural gland,nephrocytes.

24.3 Subphylum Cephalochordata

The lancelets are located in this subphylum. Along with the subphylum Urochordata, these two subphyla make up the invertebrate chordates. Lancelets receive their name from their bladelike shape. They resemble fish but they are actually scaleless chordates only a few centimeters long. They spend most of their time buried in the sand with their mouths protruding. Fossils of lancelets have been found to be over 550 million years old.

Dropped out sessile stage, what was the larval stage is now sexually reproductive. Includes Branchiostoma ("amphioxus").

24.4 Subphylum Vertebrata

(Vertebra from Latin vertere, to turn). Characterized by separate bones or cartilage blocks firmly joined as a backbone. The backbone supports and protects a dorsal nerve cord. Vertebrates have tissues which are organized into organs which in turn are organized into organ systems.

All vertebrates share the following characteristics: - segmentation - a true coelom - bilateral symmetry - cephalization - a backbone - a bony skull - a closed circulatory system - chambered heart - two pairs of jointed appendages - tissues organized into organs

Vertebrate Organ Systems: - Nervous System - Circulatory System - Digestive System - Respiratory System - Reproductive System - Excretory System

- Vertebral column: Not present in higher vertebrate adults. (In humans, the gel-like, spongy core of the vertebral column is the only remainder. Ruptured or herniated disc is an injury to this.)
- Cranium: Composite structure of bone/cartilage. Two functions: 1. Supports sensory organs of head and 2. Encloses or partially encloses the brain.

What evolutionary relationship could we imagine between sessile echinoderms and the higher chordate animals?

Paedomorphic (child-form) hypothesis: basically, evolution of sexual reproduction in what had previously been a larval life stage, or the retention of at least one juvenile characteristic into the adult (adult = sexually reproducing) stage. Some scientists believe that this occurred in a proto-chordate animal lineage. Maybe chordates (and vertebrates) arose from sessile (attached) ancestors. Selection in these proto-chordates maybe began to favor more time in the larval stage, as feeding was more successful or mortality lower in this stage. As larvae got bigger physics shows that the cilia become less efficient for locomotion, favoring the undulating motion allowed by a notochord.

Is this hypothesis crazy? A similar example of this today is Epemeroptera, the mayfly, which has almost abandoned its adult stage. Its one-year lifespan is mostly larval with just a brief day of reproduce-and-die as an adult, which doesn't even have usable mouthparts.

Tunicate (sea squirt) larva has all four chordate characteristics, although adult sessile ("attached").

24.4.1 Class Agnatha

"jawless fish"

- Ostracoderms: extinct Agnathans which had primitive fins and massive plates of bony tissue on their body.
- Cyclostomes: "circle mouth" - a group of Agnathans which is still alive in the form of lampreys and hagfish.

Appeared approximately 500 million years ago and dominated the oceans for about 100 million years. The first group of fish to appear. They had neither jaws, paired fins, nor scales, but they were the first organisms with a backbone.

Class Acanthodia

"spiny fish" Appeared about 430 million years ago. An extinct class of fish that developed jaws with bony edges. They had internal skeletons made of cartilage and some bone.

Class Placodermi

Appeared about 410 million years ago, dominated the sea for about 50 million years. An extinct class of fish with massive heads.

24.4.2 Class Chondrichthyes

"cartilaginous fish" Appeared about 400 million years ago with bony fish. Includes sharks, skates and rays, and chimaeras. Their skeletons are made of cartilage strengthened by the mineral calcium carbonate.

The main characteristics and distinguishing features of this class: - gills - single-loop blood circulation - vertebral column - presence of placoid scales on their bodies - internal skeleton of cartilage - paired, fleshy pectoral and pelvic fins - asymmetrical tail fin prevents sinking - fatty liver provides neutral buoyancy - visceral clefts present as separate and distinct gills - no external ear - oviparous - internal fertilization - ectoderms - cold blooded

24.4.3 Class Osteichthyes

"bony fish" Appeared about 400 million years ago with cartilaginous fish. Includes about 95% of today's fish species.

Subclass Sarcopterygii
 fleshy-finned fishes. Fins have bones and muscles, homologous to our limbs.

Order Dipnoi
 lung fishes, two groups isolated when continents separated

Order Crossopterygii

includes coelacanths and rhipodistians, gave rise to amphibians, had lungs which evolved into a swim bladder in bony fishes, and labyrinthodont teeth, characterized by complex folding of enamel.

- Skeleton made of bone, jaws, fins, most with scales, two-chambered heart.

24.4.4 Class Amphibia

means "both lives", aquatic larvae, terrestrial adult Amphibians: - Legs - Lungs - Double-Loop Circulation - Partially Divided Heart - Cutaneous Respiration (Breathes through Skin)

Order Salientia

frogs (jumping) (aka Anura)

Order Urodela

salamanders (tailed)

Labyrinthodont amphibians: oldest known amphibians, inherited characteristic teeth from crossopterygii ancestor, had stocky, aquatic larvae.

Amphibians have limbs instead of fins. Girdles and vertebral column now more substantial and connected, support body on legs.

Lisamphybia: no scales, "smooth", eggs with no shell, laid in water (water-reliant).

Amphibians gave rise to cotylosaurs, from which arose dinosaurs, turtles, lizards, and therapsids.

Class Reptilia

amniotic egg allowed freedom from water, shelled egg. (Amnion: protection). Reptiles have four extra-embryonic membranes:

- Amnion: supports aquatic environment inside egg in fluid sac.
- Allantois: allows gas exchange and elimination of wastes.
- Chorion: gas exchange
- Yolk sac: only one of the four left over from amphibian ancestor

Reptiles cold-blooded, or ectothermic, meaning that their heat come from their environment. Sometimes defined as all amniotes that are not birds or mammals.

Reptiles can be classified by skull structure into four groups:

- Anapsid
- Synapsid
- Diapsid
- Euryasid

Refers to number of holes in the skull. Cotylosaurs had Anapsid skull

Dermatocranium: from bony outer skull structure, precursor to human cranium.

Subclass Anapsidia

Subclass Testudinata
turtles[1], terrapins

Subclass Diapsida
dinosaurs[2], snakes[3], most stuff

Subclass Synapsida

Order Therapsids

Subclass Diapsida
includes Ichthyosaurs, marine reptiles convergent on dolphins; Plesiosaurs, ancient sea monsters; Squamates, including lizards and snakes; and Thecodonts, which gave rise to

- birds
- dinosaurs
- crocodilians

Dinosaurs: broken into two groups, based on hip structure

- Saurischia: lizard hips (gave rise to birds [!]), ancestrally bipedal
- Ornithischian: bird hips, ancestrally quadripedal

Crocodilians: come from archosaurs, the only extant (still living today) archosaur descendant. Ancestrally bipedal, secondarily quadripedal.

Synapsids: refers to joined (Greek syn-, together with) parts of skull. Led eventually to mammals. Synapsid pelycosaur >> therapsid >> mammals

Pelycosaur: Sail-backed dinosaur, legs not spread out like lizard but more pillar-like and under body, allowing greater activity and competence in motion, pendulum like rather than constant push-up. Teeth differentiated into different types, for pre-processing of food needed by higher metabolism. Skull changes, bone histology, suggestions of warm-bloodedness.

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1   http://en.wikibooks.org/wiki/turtle
2   http://en.wikibooks.org/wiki/dinsosaur
3   http://en.wikibooks.org/wiki/snake

Class Aves

arose late Jurassic, early Cretaceous. Feathers, skeleton modified for flight. Feathers: epidermal derivative, made of keratin (like fingernails). Carpometacarpis: bears primary flight feathers, parallel to hand parts. Keeled sternum: breastbone, powerful one needed to support flight muscles. Strong, light, occasionally hollow bones. All birds lay eggs (as contrasted to reptiles, which have developed live birthing over 100 independent times.) Why are there no live-bearing birds? Early birds had teeth, lost them. With mammals, only exothermic animals.

Archaeopteryx: "ancient wing", Jurassic bird-reptile, very dinosaur-like. Good fossils found in Zolenhoffen, German sandstone mine with fine sand, shows feathers clearly, found shortly after Darwin's publication and used to support his hypothesis. Thick, heavy bones and no sternum, bony tail, not a good flyer but did have primary flight feathers.

Archaeornithes: includes archaeopteryx.

Paleognathae: gave rise to Australian flightless birds.

Neognathae: remaining live birds.

Class Mammalia

Two unique characteristics, or synapomorphies:

- Hair
- Mammary glands

(don't fossilize well)

Three skeletal characteristics (fossilize)

- Lower jaw only one bone, the dentary (several in reptiles)
- Three bones in middle ear: malleus, incus, stapes (reptiles have one or two, never three)
- Joint between upper and lower jaws between dentary and squamosal of skull (in reptiles this joint is between other bones)

Mammals basically have a synapsid skull design inherited from ancestor

Non diagnostic characteristics (not unique to mammals):

- Warm-blooded
- Skin glands: sweat glands and oil-producing sebaceous glands
- Large nasal cavities (because of high metabolism) Clean, warm and humidify air
- Heterodonty (differentiated teeth)
- Diphiodonty: two sets of teeth: baby and adult ("deciduous" teeth, drop out) (reptile teeth are continually replaced)

Subclass Protheria

monotremes (Greek mon-, one; and trema, hole), or egg-laying mammals, have one opening for excretion and urination.

Subclass Theria

Metatheria: Marsupials (opossum, kangaroo. . .) Eutheria: Placental mammals (all common mammals)

Marsupium: (from Greek marsypion, purse or pouch). Gestation period much shorter than in Eutherian mammals, but after leaving the uterus the tiny offspring crawls into a pouch where it completes development latched onto a teat.

Recent molecular (read: genetic) evidence suggests that two different mammal groups may have developed live-bearing ability separately. Instead of being a "rough draft" for placental-style live bearing, perhaps the marsupial pouch approach is another solution to the same problem. Advantage: in tough times the parent can pitch out the offspring and increase its own chance of survival.

25 Tissues and Systems

26 Epithelial tissue

Comes from various sources, **ectodermal** or **endodermal** material. Cell sheet lines a surface or body cavity. One side, called **freesurface or Apical**, is exposed to

- animal **interior** (forming the **lumen**) or
- **exterior** of its **body cavity**.

The other side rests on the **basal layer**.

Epithelial tissue is not penetrated by blood vessels.

Two categories:

- **sheets**
- **glands**

Classified on two features:

- **simple**, (a single layer of cells),
- **stratified**, (more than one cell layer.)

Cell shape at free surface:

- **squamous** (broad and flat)
- **cuboidal** (spherish)
- **columnar** (tall and thin)

Simple squamous epithelium

usually lines body cavities and vessels,alveoli, glomeruli of kidney; in blood and lymph vessels called endothelium; in body cavities called mesothelium (serosae): parietal serous membranes line body wall, visceral serous membranes cover organ

Simple cuboidal epithelium

in ducts like kidney and salivary glands.

Simple columnar epithelium

nonciliated type lines digestive tract, ciliated type lines some regions of uterine tubes and lungs

Stratified squamous epithelium

(important) lines mouth, esophagus,and vagina. Cells sometimes dead, flat and keratinized, making them resistant to abrasion. Stratified squamous epithelium changes to columnar squamous epithelium progressively down esophagus to the stomach.

Epidermis

from epithelium. Below this is **dermis**, thicker and with blood vessels.

Two specialized epithelia:

- **pseudostratified**
- **transitional**

Pseudostratified epithelia

lines the trachea (where it is ciliated)and the male urethra (where it is non ciliated), looks stratified but not.

Transitional epithelia

found only in bladder and urinary system. As it stretches it appears to go from 6 to 3 cell layers deep.

Glandular epithelia

(**gland**: group of cells that excretes something.. mostly derived from **epithelium**. Glands are classified into **endocrine** and **exocrine** by where they excrete.

Endocrine glands

secrete hormones into the **blood** without use of ducts.

Exocrine glands

secrete onto the **body surface** or into a **cavity**, thru a **duct**. Exocrine substances include sweat, mucous, oil, and saliva. An exocrine gland is the **liver**, which secretes bile.

27 Connective tissue

This is a "grab bag" category of diverse tissue types. Functions include binding and supporting. Types include **bone**, **cartilage**, **fibrous connective tissue**, **blood** and **adipose** (fat) tissue.

If you took away everything in the body except the connective tissue, you'd still be able to see the basic form of the body.

Form: distinctive cells surrounded by a **cell matrix** made of **extra-cellular fiber** grounded in a **ground substance** (excluding blood)

Types:

1. **connectile connective tissues** (can be 1. **loose** or 2. **dense**)

2. **special connective tissue** (includes blood, bones and cartilage).

Fibroblasts form **connective tissue proper;**

chondoroblasts form **cartilage;**

osteoblasts form **bone;**

and **blood** is formed from various sources.

Ground substance: "unstructured" material that fills space between cells and contains fibers. Made of

1. **interstitial fluid** (bathes cells)

2. **proteoglycans** (protein core with attached polysaccharides, glycoaminoglycans or GAGs such as chondroitin sulfate, keratin sulfate, and hyalronic acid, whose consistency is syrupy to gelatin-like)

3. **cell-adhesion proteins** (connect connective tissue cells to the fibers).

Fibers of **connective tissue**:

1. **Collagen** (flexible protein resistant to stretching, tensile strength, most abundant protein in animals, white)

2. **elastin** (rubbery, resilient protein, in dermis, lungs, blood vessels, yellow when fresh)

3. **andreticulin** (like collagen).

Loose connective tissue: found beneath skin, anchors muscles,nerves etc. Include **fibroblasts**, **macrophages**, **mast cells**,and **adipose cells**. **Fibers** include **collagen** and **elastic fibers**. Ground substance is "syrupy". Adipose included.

Dense connective tissue: largely densely packed fibers of **collagen** or **elastin** regularly or irregularly arranged. Forms **tendons** and **ligaments**, coverings of muscles, capsules around organs and joints, and **dermis** of skin.

Cartilage vs. bone

Feature	Bone	Cartilage
cell type	osteocytes	chondrocytes
ground substance	calcium phosphate	chondroitin sulfate
vascularization	vascular	avascular
micro architecture	highly ordered	less organized
units called	osteons	
fibrous sheath	peristeum	perichondrium

Cartilage: There are three cartilage types:

1. **hyaline** cartilage

2. **fibrocartilage** (fibrous cartilage)

3. **elastic** cartilage

Hyaline cartilage: most widespread cartilage type, in adults forms articular surfaces of long bones, rib tips, rings of trachea, and parts of skull. Mostly **collagen**, name refers to **glassy** appearance. In embryo, bones form first as **hyaline** cartilage, later **ossifies**. Found in tracheal rings. Few collagen fibers.

Fibrous cartilage: have lots of **collagen** fibers. Found in intervertebral discs, pubic symphesis. Grades into dense **tendon** and **ligament** tissue.

Elastic cartilage: springy and elastic. Found in internal support of external ear and in epiglottis, yellow when fresh.

Chondrocites (cartilage cells) rely on diffusion for nutrients, as cartilage has no direct blood supply, and no enervation (nerves). Can be loaded with calcium salts.

Bone: Specialized connective tissue, calcium phosphate arranged in highly ordered unit called **osteon**, or **Hyvercian system**. Concentric rings around central canal with blood vessels and enervation (nerves). Bone varied, not all vertebrate bone is even cellular. Our concern: simple pattern for mammals.

Lacuna (spaces in which osteocytes found); **canaliculi** (little canals) bigger diagonal cells, layers of bone called **lamellae.**

Three types of bone cells, ending in

-**blast**, (mend bone)

-**cyte** (fortify bone)

-**clast** (tear down bone)

Classified by

1. **appearance** (**spongy** vs. **hard**)

2. **where found** (outside or inside)

3. **how it is formed** (endochondral cartilage model forms first and then is ossified, and entramembranous, bone forms directly without cartilage precursor)

Example of **endochondral bone formation**: long bone begins to ossify from center shaft, calcified region expands and cuts off diffusion of nutrients as bone replaces cartilage. In young mammals, secondary ossification centers then form at bone ends, growth has stopped by sexual maturity as all primary bone is ossified. In other animals, bones continue growing throughout their lifetime.

Three types of **intramembrous bone**:

1. **dermal** bone

2. **sessamoid** bone

3. **perichondral** bone.

Dermal bone forms skull, shoulder/pectoral girdle, and integument, descended from dermal armor of ancestor. Comes from **mesoderm**, in dermis of skin.

Sessamoid bones: form directly in tendons. Example: kneecap, also in wrist. Deals with stress.

Perichondral bone means **"around cartilage,"** forms around cartilage or bone. Functions in bone repair and in ossification of endochondral bone.

Bone remodeling and repair: bone has mineral structure, and develops tiny fractures, which, under stress, can lead to larger fractures. To combat this, bone is constantly replaced. **Osteoclasts** channel through existing bone, tear down and leave behind **osteoblasts** and **lacuna**, leaving **osteocytes**. Continually resets mineral structure of bone, and is preventative maintenance.

When bone broken, callus forms in open ends, **periosteum** gives rise to new bone with calcium and new bone matrix, leaves irregular mend. Later, osteoblasts continue fixing over time and slowly removing imperfection.

28 Muscle tissue

Mesodermal in origin, muscle has several functions: **supply force for movement**, **restrain movement**, proper **posture**, act on viscera (internal organs) for **peristalsis** (moving food down digestive tract), give body **shape**, form **sphincters**, (such as in esophagus, between stomach and intestine, large and small intestine, in anus), in **sheets** of muscles, affect air **flow** in and out of lungs, line **blood vessels** and play vital role in **circulation**.

Secondary roles: **heat production** (shivering a specialized heat production to supplement metabolism).

Muscles co-opted to other non-original functions: sharks detect electrical field created by fish muscles. Some fish formed electric organs, create current strong enough to repel predators or stun prey. Other fish can use field as "radar" to see things and communicate with other animals. (Evolved independently in different groups).

Different classifications: by **color**, (red or white) **location**, nature of **nervous system control** (voluntary or involuntary), **embryonic origin**, or by **general microscopic appearance** (striated, smooth, and cardiac.)

Striated muscle (or **skeletal** muscle): under **voluntary control**. Individual cells called **fibers**, grouped into **fascicle**. **Myofibrils** founding one cell made of even smaller **myofilaments**. Each striated cell very long and **multi-nucleated**. Fibers joined end to end to form longer **composite fibers**. **Sarcomeres**: repeating units make up **myofibrils**. Two kinds of **myofilaments**, thick kind made up of **myosin** and thin of **actin**. Striations visible in light microscope, smaller part only with electron microscope.

Cardiac muscle: occurs only in **heart**. Light banding visible under light microscope. Each band short, principally mononucleate (occasionally dinucleate) often branched, joined together with intercollated discs. Involuntary. Waves of contraction spread through intercollated discs. Initiated by nerve stimulation or can originate in the heart itself (useful in heart transplants.)

Smooth muscle: no striations visible with light microscope. Almost entirely **visceral** function: digestion, sphincters, urogenital tracts, piloerectory muscles (make hairs stand up), lungs. Non-voluntary control. Slow and sustained action. Each cell mononucleate, short, fusiform (spindly) in shape, cells usually uniform in size.

Striated muscle contraction: Muscle broken into units called fascicles, in units of myofibrils. Repeating units called sarcomeres, consisting of two kinds of myofilaments:

1. thick, **myosin** filament

2. thin, **actin** filament.

Sarcomere: Thick and thin filaments interspersed in ordered grid.

Sliding filament theory: thick and thin filaments move past each other in opposite direction, shortening length. Longer muscles contract more rapidly than short ones (see cell bio for details).

Myosin molecule: two polypeptides twisted together with two globular heads at end.

Myosin filament: many slender myosin molecules together.

Actin filament: chain of actin single, **tropomyosin** strands with repeated **globular troponin**, and with **actin**. All play role in muscle contraction. Myocin heads have sites that bind to actin. Actin filaments have many regular sites that can bind to myosin.

Troponin has **four site**s:

1. one to bind **myosin**

2. one for **actin**

3. one for **tropomyocin**

4. one for **calcium ions**

Nerve signal reaches muscle, triggers release of chemical signal called neurotransmitter, that diffuses across cell membrane (**sarcolimic reticulum**) and binds to receptors in it. Receptor is **acetylcholine**, ACH. When there is enough nerve signal, the message travels through t-line to sarcoplasmic reticulum to release calcium ions.

Lacking calcium, **tropomyosin** site blocked. In calcium, myosin binding sites exposed and heads bind to **actin** molecules, delivering force to move **fibers** in relation to each other. Myocin head then interacts with **ATP** to get "recocked", if myosin still exposed then it fires again and results in further muscle contration. If there is no further nerve signal, **sarcoplasmic reticulum** sequesters Ca+ ions again and no recocking occurs.

Quirari (or curare): known from movies, used in South America, blocks **acetylcholine receptors** in cell and causes skeletal paralysis. Victim dies of asphyxiation because he can't breathe.

Duchenne's muscular dystrophy: degeneration of **sarcolema**, plasma membrane of muscle cell unable to release signal and quickly atrophies.

Fast and **slow twitch fibers**: vertebrate muscle fiber. Terms relative within one group of animals. Differences related to differences in **enervation**, type of **myocin**, and **actin activation**.

Two parts of force generated by muscle: 1. **active** component 2. **elastic** component (energy stored in muscle when stretched by gravity or another force. Stored in muscle elastic tissue around tendons. Especially important in **limb oscillation**, like running, or trunk twisting, like fish swimming. Up to 90% of stored elastic energy can be recovered.)

How does a muscle match its power to its job? Two ways:

1. **rate modulation**, derived from frequency of nervous stimulation of muscle, force increases as frequency of stimulation increases up to point of tetanus.

2. **selective involvement of motor units**, a given neuron enervates a fixed number of muscle cells, (a motor unit), and force is increased by recruiting more motor units. Motor units may be small, such as in eye, or larger, like in leg muscle.

How do muscles grow stronger?

1. add more **myofilaments**, increases cross sectional area by up to 50%, more little ratchets working

2. proliferation in **blood vessels** and **connective tissue** around muscle

Muscle strength is relative to **cross sectional area**, not length. Not always feasible to add more cross sectional area.

Pinnate fibers: oriented obliquely (Y-shaped) to minimize muscle mass, in certain circumstances, like calf muscle. Spreads muscle out.

Velocity of shortening greater in long muscle than short. Why? Contraction tied to relation between fibers, and to total length of muscle. Both long and short muscles reach same percentage of contraction in same unit time, but distance covered by the longer muscle is greater.

Synergist muscles: muscles work together to produce motion in same general direction. Bicep shares work with brachialis.

Antagonist muscles: muscles that oppose each other. Bicep pulls forearm in, triceps pulls it back out.

Origin vs. insertion: origin is the end of the muscle that more fixed in its attachment to the body. The more movable end called insertion.

Fixators: muscles that act to stabilize a joint or lever system. Like upper arm when you clench your fist hard.

Flexors and extensors: applied mainly to limbs. Flexor bends one part relative to another about limb, extensor straightens it.

Adductor and abductor: adductor draws a limb toward the ventral surface. Abductor moves limb away from ventral surface. (Adduct: drawn toward; abduct: carry away).

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29 Vertebrate digestive system

Functions to break down food into molecules small enough to absorb, or pass across digestive membrane.

Digestive tract: tube extending from lips of mouth to anus or cloacae in bird, reptile or monotreme.

Lumanal glands: empty into inner body cavity (lumen: inner surface).

Tract divided into three main regions: 1. buccal cavity 2. pharynx 3. alimentary canal

Alimentary canal divided into four regions: 1. esophagus 2. stomach 3. small intestine 4. large intestine

Accessory digestive glands, outside digestive tract proper, secrete into lumen of tract through ducts. Includes the **salivary glands**, **liver** and **pancreas**.

Buccal cavity, which includes **palate** and **tongue**, develops from infolding of **stomadeum**, or second opening of blastula, whereas the rest of the digestive tract develops from the **primitive** gut.

Teeth: capture and hold prey. In mammals in particular further process and break down food into small particles, increasing surface area available for enzymatic action.

Tooth anatomy: 1. **crown** projects above gum, 2. **root** below gum, 3. **enamel** is outer coating of crown, hardest surface in body, of epideral origin 4. **dentin**, below enamel, bone-like and forms bulk of tooth, is harder than bone and contains nerves and blood vessels. (Remember that mammals are **heterodontic**, with different types of teeth).

Pharynx: air passage for adult, gill slits in embryo. Important in lower vertebrates, site of gills. Features derived from **pharyngeal** region: first pharyngeal pouch gives rise to parts of the ear, other pouches give rise to various other structures.

Alimentary canal: epithelium lines lumen, glands secrete into lumen, longitudinal and circular muscles help digestive movements (**peristalsis**).

Esophagus: tube carries food from mouth to stomach. Expands to fit large **bolus** (lump of chewed food). Secretes mucus for lubrication. Birds have **crop** for storage, enlargement of esophagus.

Epiglottis: keeps food out of air tube, an evolutionary "kludge," or fix.

Stomach. Absorbs water, alcohol, nutrients, uses gastric juice with enzymes, mucous, HCl, released by **chief** and **parietal cells** (release protein enzymes) in gastric pits. **Rugae**: folds of stomach, disappear when full. **Sphincter** at both ends of stomach, control food passage.

Chyme: semi-digested food released to small intestine.

Small intestine: three regions, **duodenum**, **jejunum**, and **ileum**.

Duodenum site of most intestinal digestion. **Jejunum** and **ileum** do most of intestinal absorption. Ileum ends with another sphincter, **ileocolic valve** or ileosecal valve. Structure: Circular folds covered with **villi** (singular is villus).

Villi: finger-like cellular projections, covered with **microvilli**, tiny projections which increase surface area. Increases surface area by 900x, speeds **digestion** (break down) and **absorption** (taking in nutrients).

Large intestine: larger diameter, shorter length than small intestine. No villi. In mammals, forms large gentle loop, colon, empties into straight region, rectum, empties into outside world through anal sphincter. Colon: absorbs water left over, also absorbs vitamins released by bacteria which live there (vitamin K).

Food: made up of 1. **proteins**, 2. **fats**, 3. **carbohydrates** 4. **fibrous material**.

Digestive system breaks foods down. **Proteins** must be broken to **amino acids** to be absorbed. **Polysaccharides** to **monosaccharides**, **lipids** to **fatty acids** and **monoglycerides** to absorb.

Salivary glands in mouth, saliva contains mucous, salt and a few enzymes (amalase, begins starch breakdown). Snake venom from oral gland, mixture of toxins and digestive enzymes. Breaks down blood vessels and disables nervous system.

Stomach enzymes: released in inactive form, **zymogene**, converts to active form in lumen of gut. Transformation is triggered by another enzyme, or the stomach's low pH. **Pepsin** secreted as **pepsinogen** (-ogen means primitive form). Stomach glands secrete up to two or three liters a day of gastric juice, which is reabsorbed.

Chyme released to **duodenum**.

Small intestine has two major accessory glands: 1. **pancreas** 2. **liver**

Pancreas has **endocrine** and **exocrine** functions, releases large amounts of carbonate to neutralize acidic chyme, as intestinal enzymes work in neutral pH, and stuff to break down lipids and starch (zymogens, like tripsin)

Liver releases **bile**. Bile made from cholesterol, stored in **gall bladder**, released in duodenum, emulsifies fats.

Emulsify: keeps fats in tiny drops, which are suspended, increasing surface area and action of lipases. Protein and carbohydrates absorbed in intestine, taken to liver for processing. Fatty acids go to lymphatic system

Appendix: vestigial remnant. Much variation in digestive systems within mammals: herbivore, carnivore, insectivore, non-ruminant herbivore.

Rumen: four-chambered stomach of animals like cows (ruminant herbivores). Cellulose resistant to digestion, rely on microorganisms to break down cellulose. Some bacteria, protists and fungi can break down cellulose, almost no animals can. Bacteria break down cellulose in rumen, to be taken back to the mouth to chew their cud (ruminate). Later cow swallows to proceed with digestion. (Horses not like this).

Coprophagy: rabbits and other animals eat their own feces for the nutritious products of the cecum.

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30 Circulatory system

Circulatory system functions

1. **Transportation**

```
   a. Respiration: gas exchange (O2 and CO2), overcomes limited rate
of
```

diffusion

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   b. Nutrition:<br>
   c. Excretory: (remove metabolic wastes)<br>
```

2. **Regulation**

```
a. Transport hormones<br>
b. Regulate body temperature<br>
c. Protection<br>
   i.  Blood clotting<br>
   ii.  Immune system (carries white blood cells)
```

Vasodilation: allows heat loss across epidermis, as seen in elephant ears, takes more blood to surface of body, sweating may accompany

Countercurrent heat exchange: used by dolphins in fins to conserve heat in cold water. Veins surround an artery, and blood returning to body absorbs heat from blood traveling out from body to fin, minimizing heat loss. Used by dogs in feet, etc.

Blood made of 1. **plasma** and 2. **formed cellular elements** (red and white blood cells, and platelets).

Plasma makes up 55% of blood volume. Cellular elements make up the other 45%.

Plasma makeup: 90% water, 7-8% soluble proteins (**albumin** maintains blood osmotic integrity, others clot, etc.) 1% electrolytes 1% elements in transit

Red blood cell (**erythrocyte**): contains **hemoglobin**, functions in oxygen transport. In mammals, red blood cells lose nuclei on maturation, and take on biconcave, dimpled, shape. No self repair, live 120 days. About 1000x more red blood cells than white blood cells. About 7-8 micrometers in diameter.

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Hematocrit: proportion of blood volume that is occupied by cells, about 43% in humans on average. 48% for men and 38% for women.

White blood cells (**leukocytes**): Nucleated, about 10-14 micrometers in diameter, commonly amoeboid, escape circulatory system in capillary beds. Include basophils, eosinophils, neutrophils, monocytes, B- and T-cell lymphocytes.

Platelets (**thrombocytes**) Membrane bound cell fragments in mammals, no nucleus. In non-mammals, platelet role replaced by nucleated cells. Accumulate at site of broken blood vessels, form clots. Bud off special cells in bone marrow. 1-2 micrometers in diameter. 7-8 day life span, 1/10 or 1/20 as abundant as white blood cells.

Arteries: carry blood **away** from heart. Smallest tubes called **arterioles**, feed blood to capillaries.

Veins: return blood to heart. Smallest veins called **venules**.

Structure of arteries and veins, listed from inside (lumen) out: 1. **epithelium** (endothelium), 2. **elastic connective tissue fibers**, 3. **smooth muscle**, 4. **connective tissue**. Arteries have thicker elastic layer than do veins.

Capillaries, where exchange of materials occurs, are very thin and narrow, and red blood cells pass through single file. Capillaries are tiny but numerous, and their total volume is greater than that of supplying arteries.

Blood velocity drops in capillaries, picks back up in veins. Pressure highest in arteries, lower in capillaries and arteries.

Osmotic pressure draws interstitial fluid from blood in arterioles, but replaces it in venules.

One-way valves mean that blood can flow only one way, works with residual blood pressure and compression by skeletal muscles. Low pressure in thoracic cavity caused by breathing also helps move blood.

Lymphatic system: part of the immune system, a one-way, or **open**, system. Takes up **interstitial fluid** not taken up by venules.

Lymphatic structures:

1. lymphatic **capillaries**

2. lymphatic **vesicles**

3. **lymph nodes**

4. lymphatic **organs** (**spleen** and **thymus**)

Lymph: movement in mammals through one-way valves, similar to blood movement in veins. (Some non-mammals have lymphatic hearts of unknown embryonic origin. Frogs and salamanders have several.) Lymph rejoins cardiovascular system into a large vein near the heart via single large thoracic duct.

As lymph passes through system, passes **lymphocytes**, second part of immune system.

Heart: pumps blood, design varies between animals. In adult mammal,four chambers form two separate circulations

1. **pulmonary circulation** to and from lungs and

2. **systemic circulation** to and from tissues of body.

Everything in the heart comes in pairs: 2 atria, 2 ventricles (left and right).

Diagrams usually drawn as though animal were on its back.

Pattern of blood flow through heart: blood returning from **major veins** (vena cava) enters right **atrium**, contraction there delivers blood to right **ventricle** through a **tricuspid valve**, one of atrial ventricular valves (AV valve). Contraction of right ventricle drives blood through **semi lunar valve** into **pulmonary circuit** and to **lungs**.Blood return to heart in **pulmonary veins**, is oxygenated. Goes to left **atrium**, which contracts and delivers blood to left **ventricle**by way of **aortic semi-lunar valve**, then goes to **systemic circulation**.

Both **atria** and **ventricles** contract in unison, left is more powerful than right (to all system vs. just lungs).

Systole: heart contraction, **diastole**: heart relaxed

Timing of heart contraction: **ventricles** rebound to relaxed shape (diastole), and **semi-lunar valves** close. Both **atria**(singular: atrium) fill with blood coming from pulmonary and systemic circulations.Pressure rises in the atria and blood begins to move into the **ventricles**.The atria then contract, forcing more blood into the ventricles. There is a pause, then ventricles contract. This raises ventricle pressure, **atrio-ventricular(AV) valves** shut and semi-lunar valves open, forcing blood from the left ventricle into the major arteries and from the right ventricle into the **aorta**.

Control for this action doesn't rely on nervous stimulation, has intrinsic rhythmicity, called **myogenic**. This is the case in mammal as well as in mollusk hearts. Other animals have **neurogenic** hearts that rely on nervous stimulation for heart action, originating in the **cardiac ganglion**.

The rhythmicity of mammalian heart relies on the **sino-atrial (SA)node**, or pacemaker. This is a **phylogenic** (based on evolutionary history) remnant of an early vertebrate heart that had one more chamber than modern hearts.

How the heart contracts: waves of **depolarization** start in **SA node** and spread through **atria**. Connectile tissue pauses the spread of depolarization at the atrial ventricular node. Signal continued by **bundle branches** to **lower ventricle**, begins to stimulate heart to contract. Contraction starts at bottom of heart at **heart apex**,then signals spread through heart.

Medulla (in the brain) controls autonomic nervous system. (The medulla is part of the brain, is continuous with the spinal cord, and controls involuntary actions of the body). **Sympathetic cardiac accelerator**connects to **spinal cord**, uses **norepinephrine** to signal. **Parasympathetic cardio-inhibitory center** reaches heart through **Vagus**

nerve, uses**acetylcholine** to signal. Hyperpolarizes membrane to inhibit heart contraction. (Autonomic nervous system: two parts working in contra to control from both sides.) Dominant effect here is **inhibitory**. If we cut Vagus nerve, heart rate promptly rises about 25 bpm.

31 Respiratory system

In humans and other animals, for example, the anatomical features of the respiratory system include **airways**, **lungs**, and the respiratory muscles.

Other animals, such as insects, have respiratory systems with very simple anatomical features, and in amphibians even the skin plays a vital role in gas exchange.

Plants also have respiratory systems but the directionality of gas exchange can be opposite to that in animals. The respiratory system in plants also includes anatomical features such as holes on the undersides of leaves known as **stomata**.

In mammals, the **diaphragm** divides the body cavity into the

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abdominal cavity: contains the viscera (e.g., stomach and intestines)
```

```
thoracic cavity: contains the heart and lungs.
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Respiratory tree: terminates in alveolus, alveoli. Respiratory bronchioles branch into alveolar ducts and into alveoli. Alveolus: microscopic air sacs, 300 million of these in human lungs. Total surface area large. Gas diffuses micrometer, very tiny distance.

Nervous System

Composed of tissues designed to integrate sensory information and direct a coordinated response to the environment.

Basic unit of the nervous system is the neuron[1], a highly specialized cell that uses both electrical and chemical processes to communicate. Neurons "listen" to sensory organs or other neurons, and can simultaneously "hear" from 1 to hundreds of inputs simultaneously. Likewise, a neuron can "talk" to other neurons or cells that can create an action, like muscle cells or glands.

Neurons are outnumbered in the nervous system by *glia*. Glia were once thought to only play a supportive role in helping neurons survive; today we know that they also are important participants in the communication process. Glial cells include astrocytes, ependymal cells, and a cell that has a macrophage like function. There are also oligodendocytes and Schwann cells that provide a myelin sheath.

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1   http://en.wikipedia.org/wiki/Neuron

31.1 Neuron structure
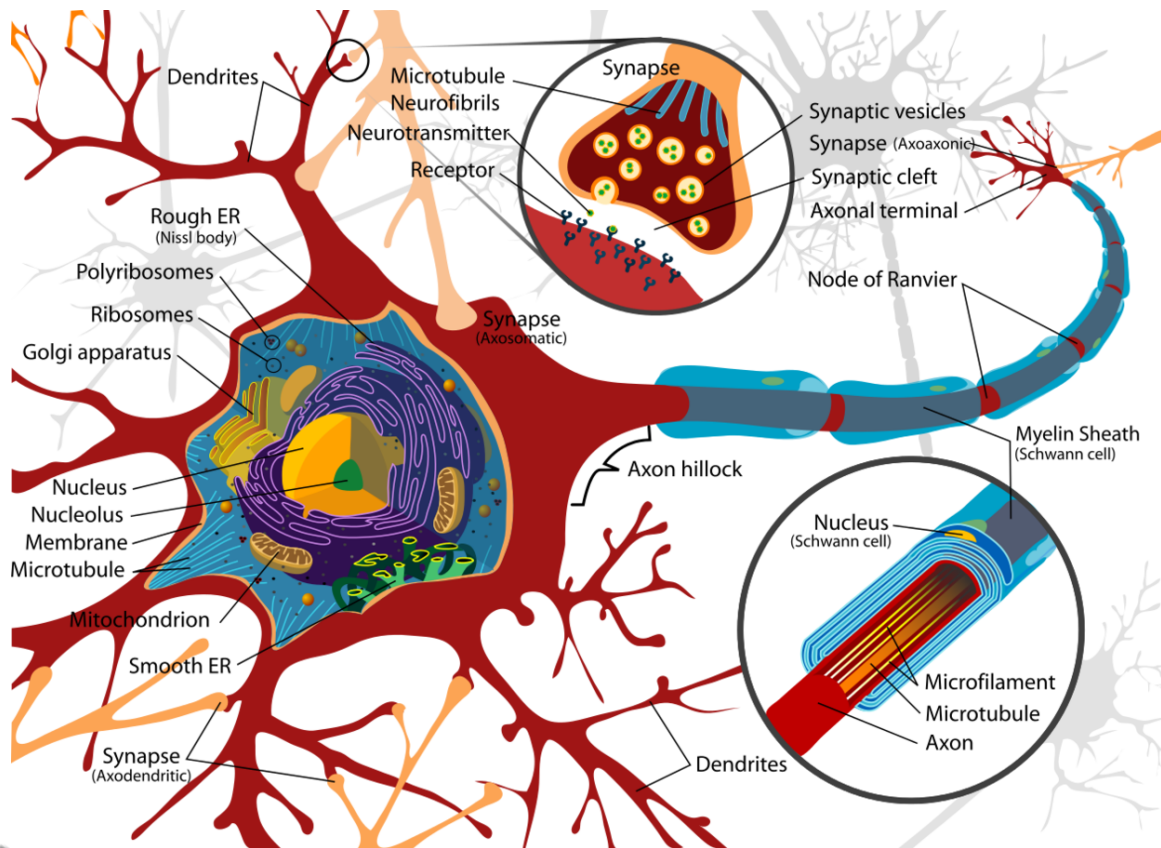
Figure 9   Neuron

- Dendrite will receive information from other axons
- Stoma is the neuron body and contains typical cell parts including the mitochondria and nucleus. This is where neurotransmitters are synthesized.
- Axon hillock is where the cell body and the axon meet.
- Axon is surrounded by myelin.
  - Contains nodes of Ranvier

Types of axons include unipolar, bipolar, pseudopolar.

Action potentials results from excitatory stimuli received from the dendrites that results in a signal that travels down the axon. On the cellular level, there is a Na+ influx via channels causing a depolarization of the cell. Once those channels close, the slower opening K+ channels will open resulting in hyperpolarization by the cell.

Synapses

Neurotransmitters allow a chemical signal to be sent from one neuron to the other. Neurotransmitters must bypass a physical gap called the synapse. Examples of neurotransmitters include acetylcholine, epinephrine, and glutamate.

Example : Neuromuscular synapses

Circuits / Nuclei / Ganglia

31.2 Central nervous system

Central nervous system includes the brain and the spinal cord.

Brain There are four lobes of the brain.

- Frontal lobe - decision making
- Parietal lobe
- Temporal lobe
- Occipital lobe - vision

The areas of the brain are also dedicated to different functions.

- Precentral gyrus
- Postcentral gyrus

Protection The CNS is protected by three layers - the pia mater, the arachnoid mater, and the dura mater. Protection also comes from the circulation of cerebral spinal fluid (CSF). CSF helps to float the brain and also provide nutrients to both the brain and the spinal cord. CSF is synthesized from the choroid plexus of the lateral ventricles. In total, there are four ventricles - 2 lateral ventricles, a 3rd, and 4th ventricle.

31.3 Peripheral nervous system

The peripheral nervous system is broken down into two sub-systems, the somatic nervous system and the autonomic nervous system.

ANS - Autonomic Nervous System The ANS has two components - parasympathetic and sympathetic.

The sympathetic nervous system is the "fight or flight" or fright response and results in an increased heart rate, increased rate of breathing, and an elevated blood glucose level. There is also decreased digestion. The second neurotransmitter is epinephrine. In this case, the first neuron is short and the second neuron is long.

The parasympathetic nervous system is the rest and digest system.

Drugs must be able to pass the blood brain barrier to have an effect on the CNS. Drugs act by affecting the neuron and how likely it is to fire an action potential.

Stroke occurs when there is a blood clot that goes to the brain and prevent blood flow.

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32 Sensory systems

Categorized by

1. **nature of stimulus**, such as **mechanical**, **chemical**or **light** stimulus, and

2. **where stimulus received**, such as **outside** (**exteroceptors**,such as the **eye** and **skin temp receptors**) or **inside**body (**interoceptors**, such as **blood body temperature receptors**).

Transduction of sensory input into signal. Means to "carry across",signal transduced, or carried, from environment into nervous signal.

Three sensory processes we cover

1. **taste** and **smell** (chemoreception)

2. **gravity** and **movement**

3. **light**

32.1 Taste and smell (chemoreception)

Found in mammal nose and mouth, fly feet, fish bodies, moth antennae.

Papilla: bumps on tongue, contain taste buds down between. Sweet, sour, salty and bitter. Some act directly by ion channels, others act indirectly. Other "taste" sensations really smell.

Smell: received in **nasopharynx**. Airborne molecules go into solution on moist epithelial surface of nasal passage. Approximately 1000 genescode for sensory neuron receptors. "Fried onions" odor not one receptor but a mixture of many odors registered in our mind as one. Very sensitive, habituates rapidly (don't notice a smell after a bit). Odor sensation has relatively unfiltered root to higher brain centers.

Snakes more chemosensory focused than us.

32.2 Response to gravity and movement

Registered in **inner ear**. Three semicircular canals loop in three planes at right angles to each other, responsible for transduction of movement messages. Method: hair cells

deformed by gelatinous membrane. **Vestibular apparatus**, gives us perception of gravity and movement. Due to **physical response**, not chemical binding.

Cochlea: bony, coil shaped part of inner ear, where hearing occurs.

Sound enters through **auditory canal**, vibrates **tympanic membrane**,moving three bones of middle ear (**malleus**, **incus**, and **stapes**)against oval window opening in front of cochlea. Cochlea has three fluid filled ducts, one of these the **organ of Corti**. Sound waves in air go to vibration in organ of Corti; fluid tickles **hair cells**, which register the movement along **basilar membrane** in cochlea. Different sound frequencies move different portions of basilar membrane. Hearing loss due to loss of hair cells.Humans normally smell more than 300 odors in a day(Facts and Truth).

Transduction of sound accomplished throgh **physical deformation**,not chemical binding.

32.3 Vision

Light enters **pupil**, ***focused by*** *lens***onto** *retina*.

Sclera: hardened part behind retina.

Optic nerves and **neurons** attached to retina. Blind spot where optic nerve attaches, has no receptors.

Two types of photoreceptors

1. **rods** - black and white low-light vision, 100 million in each retina in humans.

2. **cones** - color vision, work best under better illumination. 3 million in each retina.

Fovia: region of most acute vision, has most of the cones, few rods.

Transduction process of light to signal a molecular change, to light absorbing molecule called photopigment. Located in outer parts of rods and cones in pigment discs. The rod photopigment is called **rhodopsin**,cone has three photopigments, called **photopsins**. This molecular change initiates pathways to result in action potential in downstream neuron leading to vision center in brain.**Parul Godika**

Each of the three photopsins has a different peak of sensitivity: **blue**,**green** or **red**, and changes **isometric form** (from cisto trans) based on light from a particular wavelength range. Color blindness:inherited lack of one or more types of these cones. Gene carried on X chromosome, therefore more common in men than women.

32.4 Homeostasis

Is a very important part of everyone's and everything's lives. Defined as **dynamic constancy of internal environment**, maintenance of a relatively stable environment inside an organism usually involving **feedback regulation**.

Homeostasis is maintained in face of

1. a **varying external environment**, or
2. a **non-ideal, constant external environment** (as with the penguin).

Deals with **temperature**, **pH**, **chemical concentrations**,**pressure**, **oxygen levels**.

Occurs through **negative feedback loops**.

Various forms: simple thermostat in house turns off heater when above a certain temperature and on when below a certain temperature Involves **stimulus**, **sensor**,**integrating center**, **effector** and **response**.

More efficient control has **two sensors** and **two effectors**. Can be antagonistic to each other, such as, one cools, the other heats.

Precise control through **proportional control**, not all-or-none, furnace comes on a little bit if the house a bit cold. Examples in humans: vasoconstriction, change in metabolic rate, shivering. Physiological responses for high body temp: blood goes to body surface, sweating, behavioral changes (get out of sun).

Positive feedback loop: effector increases deviation from set point. Amplifies reaction. Like blood clotting process, uterine contraction during childbirth. Negative feedback must exist at some point for control.

32.5 Osmotic environments and regulations

1. **Marine invertebrates**
   a) fully marine invertebrates (not intertidal or estuarine) osmoconformers (set internal environment same as environment, no net flow of ions) in a stenohaline (narrow non-changing salt level) environment
   b) Coastal, intertidal, estuarine (ion levels fluctuate) invertebrates.Partly osmo-confomers, partly osmoregulators in a euryhaline (wide salt level variation) environment (ex: shore crab, regulates sometimes when salt levels in environment get real low).
2. **Freshwater animals**. Here, environment has lower solute concentrations than do living organisms so water tends to flow in and solutes out.
   a) Freshwater fish (bony) dilute urine, and gills actively take up ions (NaCl)
   b) Freshwater invertebrates: same situation as freshwater fish but with different structures
   c) Freshwater amphibians: active uptake of salts across their skin
3. **Marine fishes**: Here the environment has a higher solute concentration than does the organism so water tends to flow out and ions in.
   a) Bony fishes: actively secrete salts (NaCl) across gills, absorb water across gut wall, their kidney (unlike mammalian kidney) is unable to generate concentrated urine so glomerulus is reduced, active tubular secretion of $MgSO_4$
   b) cartilaginous fishes (and coelacanth): blood retains urea and trimethylamineoxide to increase its osmolality to that of seawater

4. **Terrestrial animals**: here problem is loss of water to a drier environment, and regulation of salt levels.
   a) water loss adaptations
   b) concentrated exception of salts and nitrogenous wastes

Hypoosmotic: having less osmotic potential than nearby fluid

Hyperosmotic: having more osmotic potential than nearby fluid

Isoosmotic: having equal osmotic potential than nearby fluid

Glomerulus: reduces volume of kidney

Fish started in salt water, spread to fresh water, later reinvaded salt-water environment.

Terrestrial animal water sources:

1. drinking

2. moist foods

3. from breakdown of metabolic molecules like fats. (Desert kangaroo rats get 90% of their water from metabolism.)

Secretion of nitrogenous wastes: from metabolism of amino acids, amino group has to be removed in one of three basically interchangeable chemical forms:

1. **ammonia** (aquatic life)

2. **urea** (mammals)

3. **uric acid** (birds)

Ammonia very toxic, soluble, and cheap to produce. Easy to expel for bony fishes.

Urea: low toxicity, good solubility, more costly to lose as it contains other groups on it. Must be released in solution, water cost.

Uric acid (white part of bird poo) low toxicity, insoluble, secreted with little water loss, more costly side groups lost than the others.

Mammalian kidney: Structure: fist-sized organ in lower back. About 1/5 of blood from aorta at any time is passing through kidneys. Blood passes through kidney many times a day.

Nephron: structural and functional unit of kidney.

Bowmans capsule: funnel-like opening, contains primary filter, the **glomerulus**.

Proximal convoluted tubule: receives stuff from Bowmans capsule.

Loop of Henle: descends and ascends.

Vasa recta: capillaries that surround the Loop of Henle.

Glomerulus: main filter of the **nephron**, located within the **Bowman's capsule**

Kidney properties and processes important to its function

1. **Active transport of solutes** from one fluid to another against a concentration gradient, Na^+ actively transported out of filtrate by cells of the thick ascending loop of Henley into the interstitial fluid

2. **Passive movement of solutes** and water from one fluid to another(down a concentration gradient), movement of water and NaCl out of descending loop of Henley into interstitial fluid.

3. **Differential permeability of cells** in different regions of the nephron to movement of water and solutes, ascending thick look is impermeable to water, descending portion is permeable to water

4. **Hormonal control of that permeability**, antidiuretic hormone(ADH) increases permeability of collecting due to water, resulting in reduced volume of filtrate and thus more concentrated urine.

5. **Increasing solute concentration in the interstitial fluid of the kidney**, from the cortex to the deepest medulla, maintained by a countercurrent multiplier mechanism

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33 Additional material

- Francis Crick[1] *chemist and molecular biologist, discovered structure of DNA molecule*
- Charles Darwin[2] *the father of the science of evolutionary biology*
- Richard Dawkins[3] *zoologist and biology populariser*
- Stephen Jay Gould[4] *paleontologist and science populariser*
- J.B.S. Haldane[5] *geneticist and evolutionary biologist, founded population genetics and the modern synthesis*
- Bill Hamilton[6] *formulated theory of inclusive fitness and kin selection*
- Thomas Huxley[7] *"Darwin's Bulldog", early evolutionary biologist and science populariser*
- Lynn Margulis[8] *introduced the theory of eukaryotic cell origin through endosymbiosis*
- Barbara McClintock[9] *geneticist and molecular biologist, discovered transposons*
- Gregor Mendel[10] *discovered the basic rules of heredity*
- Ernst Mayr[11] *evolutionary biologist and science populariser*
- Mark Ridley[12] *science populariser*
- Fred Sanger[13] *founder of DNA and protein sequencing techniques*
- John Maynard Smith[14] *evolutionary biologist and science populariser*
- Alfred Russel Wallace[15] *evolutionary biologist*
- James Watson[16] *molecular biologist, discovered structure of DNA molecule*
- Edward Wilson[17] *founded "sociobiology"*

This book is intended as a compilation of biographies describing the lives and work of influential biologists.

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1   http://en.wikipedia.org/wiki/Francis%20Crick
2   http://en.wikibooks.org/wiki/General%20Biology%2FGallery%20of%20Biologists%2FCharles%20Darwin
3   http://en.wikipedia.org/wiki/Richard%20Dawkins
4   http://en.wikipedia.org/wiki/Stephen%20Jay%20Gould
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16   http://en.wikipedia.org/wiki/James%20Watson
17   http://en.wikipedia.org/wiki/Edward%20O.%20Wilson

Additional material

33.1 External Links

- EvoWiki: List of Biologists[18]

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18   http://wiki.cotch.net/index.php/List_of_biologists

34 Glossary

- **Autotroph**[1]: an organism which can make its own energy
- **Cell**[2]: Fundamental structural unit of all living things
- **Ether**[3]:
- **Eukaryote**[4]: an organism[5] with a nucleus
- **Exoenzyme**[6]: an enzyme used to break down organic molecules[7] outside the body
- **Glycerol**[8]:
- **Heterotroph**[9]: an organism which can not make its own energy
- **Hydrocarbon**[10]: an organic compound that contains carbon[11] and hydrogen[12] only.
- **Lipid**[13]: fatty acid[14] esters[15] which form the basis of cell membranes
- **Nucleus**[16]: Membrane[17]-bound organelle[18] which contains the chromosomes[19]
- **Prokaryote**[20]: an organism with no nucleus
- **Seed**:
- **Flower**:
- **Tracheid**:
- **Haploid**: A cell with a single set of chromosomes (23 in humans), in humans this is usually in gametes. This is commonly represented by n.
- **Diploid**: A cell with two sets of chromosomes (46 in humans). This is commonly represented by 2n.
- **Sporangium**:

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2   http://en.wikibooks.org/wiki/Cell
3   http://en.wikibooks.org/wiki/Ether
4   http://en.wikibooks.org/wiki/Eukaryote
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18   http://en.wikibooks.org/wiki/organelle
19   http://en.wikibooks.org/wiki/chromosome
20   http://en.wikibooks.org/wiki/Prokaryote

The majority of the modules making up this book are based on notes very generously donated by Paul Doerder, Ph.D.[21] and Ralph Gibson, Ph.D.[22] both currently of the Cleveland State University[23].

The book was initiated by Karl Wick[24], who donated many of his own class notes for other modules, and who is fleshing out the outline format of Dr. Doerder's notes into text.

34.1 Users

Alsocal[25]

Darren Hess[26] MD/PhD recent grad, enjoys teaching, hopes to help work up the Nervous System Tissue section.

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24   http://en.wikibooks.org/wiki/User%3AKarl%20Wick
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