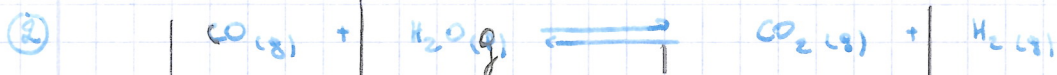




$n_i$	a	b	0	0
$n_j$	(a-c)	(b-c)	c	3c
$p$	$\frac{(a-c)RT}{V}$	$\frac{(b-c)RT}{V}$	$\frac{cRT}{V}$	$\frac{3cRT}{V}$
$a = \frac{p}{p^\circ}$	$\frac{(a-c)RT}{Vp^\circ}$	$\frac{(b-c)RT}{Vp^\circ}$	$\frac{cRT}{Vp^\circ}$	$\frac{3cRT}{Vp^\circ}$



$n_i$	c	b-c	0	0
$n_j$	(c-d)	(b-c-d)	d	d
$p$	$\frac{(c-d)RT}{V}$	$\frac{(b-c-d)RT}{V}$	$\frac{dRT}{V}$	$\frac{dRT}{V}$
$a = \frac{p}{p^\circ}$	$\frac{(c-d)RT}{Vp^\circ}$	$\frac{(b-c-d)RT}{Vp^\circ}$	$\frac{dRT}{Vp^\circ}$	$\frac{dRT}{Vp^\circ}$

$$K_{c1} = \frac{\frac{RT}{Vp^\circ} \left( \frac{3cRT}{Vp^\circ} \right)^3}{\frac{(a-c)RT}{Vp^\circ}} = \frac{\frac{3c^2 RT^2}{V^2 p^{\circ 4}} \cdot \frac{Vp^\circ}{(a-c)RT}}{\frac{(a-c)RT}{Vp^\circ}} = \frac{3c^2 RT}{(a-c)Vp^\circ}$$

$$K_{c2} = \frac{\frac{dRT}{Vp^\circ} \cdot \frac{dRT}{Vp^\circ}}{\frac{(c-d)RT}{Vp^\circ}} = \frac{\frac{d^2 RT^2}{V^2 p^{\circ 2}} \cdot \frac{Vp^\circ}{(c-d)RT}}{\frac{(c-d)RT}{Vp^\circ}} = \frac{d^2 RT}{(c-d)Vp^\circ}$$

$$3c + d = 0,6948 \text{ mol}$$

$$\left[ \begin{array}{l} n_{\text{H}_2} = 3c + d \\ n_{\text{CO}} = \cancel{c-d} \quad c-d \\ n_{\text{CH}_4} = a-c \\ n_{\text{H}_2\text{O}} = \cancel{b-c-d} \quad b-c-d \\ n_{\text{CO}_2} = d \end{array} \right.$$

$$\left[ \begin{array}{l} P_{\text{H}_2} = \frac{(3c+d)RT}{V} \\ P_{\text{CO}} = \frac{(c-d)RT}{V} \\ P_{\text{CH}_4} = \frac{(a-c)RT}{V} \\ P_{\text{H}_2\text{O}} = \frac{(b-c-d)RT}{V} \\ P_{\text{CO}_2} = \frac{dRT}{V} \end{array} \right.$$

$$\left[ \begin{array}{l} a_{\text{H}_2} = \frac{(3c+d)RT}{Vp^\circ} \\ a_{\text{CO}} = \frac{(c-d)RT}{Vp^\circ} \\ a_{\text{CH}_4} = \frac{(a-c)RT}{Vp^\circ} \\ a_{\text{H}_2\text{O}} = \frac{(b-c-d)RT}{Vp^\circ} \\ a_{\text{CO}_2} = \frac{dRT}{Vp^\circ} \end{array} \right.$$

	$\text{CH}_4(g) + \text{H}_2\text{O}(g) \rightleftharpoons \text{CO}(g) + 3\text{H}_2(g)$					$\text{CO}(g) + \text{H}_2\text{O}(g) \rightleftharpoons \text{CO}_2(g) + \text{H}_2(g)$			
$n_1^0$	a	b	0	0	$n_1$	c	b-c	0	0
$n_2$	a-c	b-c	c	3c	$n_2$	c-d	b-c-d	d	d
$p^{\text{eq}}$	$\frac{(a-c)RT}{V}$	/	$\frac{cRT}{V}$	$\frac{3cRT}{V}$	Lois de (2)				
$\frac{p^{\text{eq}}}{p^0} = a$	a-								

$$p_{\text{CO}} = \frac{d^2}{(b-c-d)^2} \cdot \frac{RT}{V}$$

$$K = \frac{p_{\text{CO}} \cdot p_{\text{H}_2}^3}{p_{\text{CH}_4} \cdot p_{\text{H}_2\text{O}}}$$

$$K_{\text{eq}} = \frac{p_{\text{CO}} \cdot p_{\text{H}_2}^3}{p_{\text{CH}_4} \cdot p_{\text{H}_2\text{O}}}$$

$$K_{\text{eq}} = \frac{p_{\text{CO}} \cdot p_{\text{H}_2}^3}{p_{\text{CH}_4} \cdot p_{\text{H}_2\text{O}}}$$

$$K_{\text{eq}} = \frac{p_{\text{CO}} \cdot p_{\text{H}_2}^3}{p_{\text{CH}_4} \cdot p_{\text{H}_2\text{O}}}$$

$$K_{\text{eq}} = \frac{p_{\text{CO}} \cdot p_{\text{H}_2}^3}{p_{\text{CH}_4} \cdot p_{\text{H}_2\text{O}}}$$

$$K_{\text{eq}} = \frac{a_{\text{CO}} \cdot a_{\text{H}_2}^3}{a_{\text{CH}_4} \cdot a_{\text{H}_2\text{O}}} = \frac{(c-d)RT}{p^0} \cdot \frac{(3c+d)^3 RT^3}{V^3 p^0^3} \cdot \frac{V}{(a-c)RT} \cdot \frac{V}{(b-c-d)RT}$$

$$= \frac{(c-d)(3c+d)^3 R^2 T^3}{(a-c)(b-c-d) V^2 p^0^2}$$

$$K_{\text{eq}} = \frac{a_{\text{H}_2} \cdot a_{\text{CO}_2}}{a_{\text{CO}} \cdot a_{\text{H}_2\text{O}}} = \frac{(3c+d)RT}{p^0} \cdot \frac{dRT}{p^0} \cdot \frac{V}{(c-d)RT} \cdot \frac{V}{(b-c-d)RT} \cdot \frac{3c+d}{(c-d)(b-c-d)}$$

$$3c+d = 0,6848 \text{ M}$$

$$b-c-d = 0,2684 \text{ M}$$