Name (Student Number):

**Salt Bridge**

도표이(가) 표시된 사진

자동 생성된 설명

Q. In chemical reactions above, Zn gets (Oxidized / reduced).

Cu gets (oxidized / reduced).

Q. How Zn2+ ions are balanced by the Salt bridge?

Q. What happens when Cu2+ ions are removed by the reduction reaction?

**Electrode Potential**

**See the standard reduction potential below.**

Q. Which one has higher reduction tendency?

Q. If those two are used as electrode, which one is cathode?

Q. Calculate the standard cell voltage (.

**Entropy Concept**

* Entropy is often described as “a measurement of disorder”. This description can explain everything about entropy ( O / X )
* There are numerous ways that the energy can be distributed in the bond of   
  two solids (O / X)
* The options of energy distribution in the bonds are called “microstate”. (O / X)
* All energy configurations of the solid bond have the same probability. ( O / X )
* Entropy is a direct measure of each energy configuration's probability. ( O / X )
* entropy can be thought of as a measurement of the energy spread in the energy configuration. ( O / X )
* Low entropy means the energy is ( spread / concentrated).
* High entropy means the energy is (spread / concentrated).
* In reality, energy stay the same energy configuration ( O / X ).
* In the 6 bonds of solid used in Ted Talks, there is an 0% chance that the hot object would get hotter ( O / X ).
* Higher entropy is always statistically more likely ( O / X ).

**2nd Law of Thermodynamics (Perpetual Machine)**

Q. Please give one example of the perpetual machine.

Q. Why perpetual machine can’t exist? (Use the law of 2nd thermodynamics concept)