Sri Chaitanya IIT Academy, India

A.P, TELANGANA, KARNATAKA, TAMILNADU, MAHARASHTRA, DELHI, RANCHI

A right Choice for the Real Aspirant

ICON CENTRAL OFFICE, MADHAPUR-HYD

Sec: Sr. IPLCO **Date: 17-10-15**

Time: 9:00 AM to 12:00 Noon **RPTM-9** **Max.Marks: 360**

KEY SHEET

MATHS		CHEMISTRY		PHYSICS	
Q.NO	ANSWER	Q.NO	ANSWER	Q.NO	ANSWER
1	4	31	2	61	3
2	2	32	2	62	1
3	3	33	2	63	4
4	3	34	3	64	1
5	3	35	2	65	3
6	2	36	2	66	4
7	3	37	3	67	1
8	4	38	3	68	2
9	1	39	1	69	1
10	4	40	3	70	3
11	1	41	4	71	3
12	1	42	3	72	3
13	1	43	4	73	3
14	1	44	2	74	4
15	2	45	1	75	2
16	3	46	3	76	1
17	1	47	4	77	1
18	4	48	1	78	2
19	4	49	2	79	2
20	4	50	1	80	1
21	1	51	3	81	4
22	2	52	4	82	3
23	1	53	2	83	1
24	2	54	3	84	3
25	2	55	3	85	4
26	4	56	1	86	3
27	2	57	2	87	1
28	3	58	4	88	2
29	3	59	2	89	4
30	2	60	2	90	1

CHEMISTRY

35. $PCl_3 + H_2O \rightarrow H_3PO_3 + HCl$

$AsCl_3 + H_2O \rightarrow H_3AsO_3 + HCl$

$NCl_3 + H_2O \rightarrow HOCl + NH_3$

$BiCl_3 + H_2O \rightarrow BiOCl\downarrow + HCl$

36. CH_4 to PbH_4 – decomposition temperatures decreases

40. $4NH_3 + 5O_2 \rightarrow 4NO + 6H_2O$

44. $2Se_2Cl_2 \rightarrow SeCl_4 + 3Se$

47. $Zn + \underset{(\text{very dilute})}{HNO_3} \rightarrow Zn(NO_3)_2 + NH_4NO_3 + H_2O$

$NH_4NO_2 \xrightarrow{\Delta} N_2 + H_2O$

$(NH_4)_2 Cr_2O_7 \xrightarrow{\Delta} N_2 + Cr_2O_3 + H_2O$

48. SF_4 – gas; SeF_4 – liquid; TeF_4 – solid

51. $Pb(NO_3)_2 \xrightarrow{\Delta} PbO + NO_2 + O_2$

53. $PCl_3 + H_2O \rightarrow H_3PO_3 + HCl$

$H_3PO_3 \xrightarrow{\Delta} H_3PO_4 + PH_3$

55. $AgNO_3 + H_2O + H_3PO_2 \rightarrow Ag + HNO_3 + H_3PO_4$