

Attempt ALL questions in 70 minutes. Write your answers on ruled answer sheet provided.

- 1 (a) Define the following terms
- (i) Salt. (1 mark)
 - (ii) Soluble salt. (1 mark)
- (b) Distinguish between the following. Give one example in each.
- (i) Insoluble salt and sparingly soluble salt. (2 marks)
 - (ii) Normal salt and acidic salt. (2 marks)
- (c) State whether the following salts are soluble or not. Copy the table and fill. (2 marks)

Salt.	Soluble or not soluble.
(i) Sodium chloride.	
(ii) Calcium carbonate.	
(iii) Sodium hydrogen carbonate.	
(iv) Lead(II) sulphate.	

- (d) Write short notes about the following methods of preparation of salts.
- (i) Double decomposition method. (2 marks)
 - (ii) Neutralization method. (2 marks)
 - (iii) Direct combination method. (2 marks)
- (e) With the aid of a well labelled diagram, describe how a dry sample of iron(III) chloride is prepared in the laboratory. (5 marks)
- (f) State one use of salts in agriculture. (1 mark)

- 2 (a) What is meant by the following terms.
- (i) Chemical formula. (1 mark)
 - (ii) Radical. (1 mark)
 - (iii) Valency. (1 mark)
- (b) Write the chemical symbols of the following elements. Copy and fill in the table. (2 marks)

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Element	Chemical symbol
(i) Oxygen	
(ii) Aluminium	
(iii) Copper	
(iv) Magnesium	

(c) Write the chemical formula of the following compounds. Copy and fill in the table (4 marks)

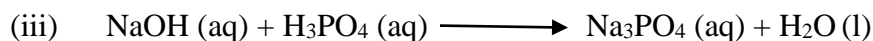
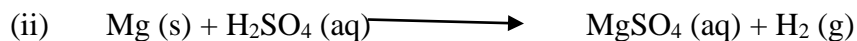
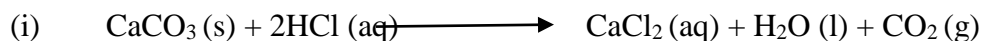
Compound	Chemical formula
(i) Calcium phosphate	
(ii) sodium hydrogen carbonate	
(iii) Ammonium chloride	
(iv) Copper(II) sulphate	

(d) write the name of the following compounds. Copy and fill in the table.

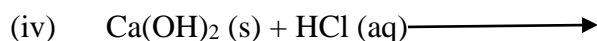
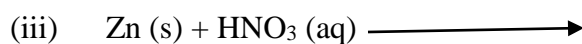
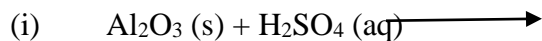
(4 marks)

Chemical formula	Name
(i) H_2SO_4 .	
(ii) $(\text{NH}_4)_2\text{SO}_4$.	
(iii) CaHCO_3 .	
(iv) PbSO_4 .	

(e) State whether the following equations are balanced or not balanced. (3 marks)



(f) Complete the following reaction equations and balance correctly. (4 marks)



END