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Signature	
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CHEMISTRY	
Paper 1	the means exhale change with Lime water
July/Aug 2023	dimme (*)
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JOINT MOCK EXAMINATIONS 2023 Uganda Certificate of Education

CHEMISTRY

Paper One

2 hours 30 minutes

INSTRUCTIONS TO CANDIDATES.

The paper consists of 50 objective type questions.

You are required to write the correct answer A, B,C or D in blue or black ink in the box provided on the All questions are compulsory. right-hand side of each questions.

All workings must be done with in the paper and therefore no extra sheets are allowed.

o not use a Pencil to an	moinups on Examiner's use only	
Marks scored	Initial	
IS OF A COLOR	The continue of oxygen evolved at sign where one potassium chieran	
	End A	

P.T.O

	All questi	ons are compulsory. Answer a	alle	
1.	winch one of the following react	ions represents a chemical chan	ge? A. Melting	ofice
	D. Taporization of Water.	4	Por 11: Intermile	or icc.
	C. Explosion of natural gas with a	air		_
	D. Magnetization of Iron.			
2.	Gas X has the following properties	es.		
	(i) Colourless			
	(ii) No effect on litmus.			Junanaga
	(iii) no observable change with Li	me water		CHISTRE
	(iv) flammable	•		
	Gas X is likely to be			
	A. Hydrogen.			
	B. Chlorine			
	C. Ammonia			
	D. Oxygen	300-000	765 3 Sec.	1 1
3.	Which one of the techniques car	n be employed to separate orange	e dve from blu	e dve in
	black ink? A. Sublimation.	W. 1. 10 2 47	,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,,	o ayo m
	B. Chromatography.	Total Carlo		
	C. Use of a magnet. D. Fractional distillation	MATTARIANT	A THIT ET'S	. ,3
	at a resolution and and and and and and and and and an			
4.	Which one of the following bases	s is insoluble? 🔝 💥 🔾 🖟 🔭 📉	OL	V
	A. $Cu(OH)_2$	Leanned Certificians of Line		
	B. KOH			
	C. NaOH	CHESTAF		-1'
	D. CaO			
5.	When testing for sulphate ions	in solution, dilute nitric acid is	s added before	aqueous
	barium nitrate in order to			,
	A. Change the Sulphate ions i	into sulphite ions B. Eliminat	e any Sulphite	ions or
	Carbonate ions.	a notation of the time		
	C. Acidify the medium for the re	action.		
_	D. Catalyse the reaction.	6.7.2+:	10.24	ten dan kin
6.	An aqueous solution containing			
	with Sodiumhydroxide solution			
ffo	What was the residue that remain			
	Zn(OH) ₂ and Cu(OH) ₂ .	on Uno ALA arms work		
	B. Zn(OH) ₂		nonvoup lastron	ship in a sign
	C. Cu(OFI) ₂	is on all years of the tenni, and or r	in se none with	e de arrangeme
	C. Cu(OH) ₂ . D. [Zn(NH ₃) ₄] ²⁺		way and of hor	D. T. T. Lees a Per
			the state of the s	P. Company of the Com
	Potassium chlorate decomposes			
	2 KClO _{3(S)}	$2KCl_{(s)} + 3O_{2(g)}$	a rie	han scored
	The volume of oxygen evolved a	t stp, when 5.0g of potassium ch		
			The second state and the second	•
	$(\frac{5 \times 22.4}{122.5}) \text{ dm}^3$			
	Δ 122.5			

	В.	$(\frac{2 \times 22.4}{122.5})$) dm ³			entra i bas si is	un in the state
		3 X 22 A	0.52	Or Wells.	uhe (ummilaekan	in Allountin	6 1 2H .
	Ć	(245.4) dm ³	3 10 50 5 X	" win 13	di - di missi	
	Ο,	245				1 /	.11
		5 X 22.4	x3		<u>.</u>	T.	7
	D.	245	$\frac{x3}{}$) dm ³	ř.			
	-,	245	Taittule	that make up dur	minute elements	of operiod earlies ou	White a one
	(C	1 = 35.5, O	= 16, K = 39,	l mole of gas oc	cupies 22.4dm³ a	nt stp)	A. Piacen R. Alamin
8.	W	hich one of	f the following	out of a	nins a dative bond	d Carbon.	C Iron and
	A.	C ₂ H ₄	t the following	substances conta	uns a dative bond	ily Laminos, "EM, mus	man. A. Q
		MgCl ₂	north serution	ble in drime aim	cuipinates is solu	of the following pr	one that W
	_	Cu(NH ₃) ₄	•			1	A
	_	CaO					B. 4. CL.
		CaO					1.70
9	Α	alass rod d	limnad in a 1				CIONER C
•	CO	ntaining di	iipped in a soli	ition of Lead (II) nitrate was thei	n dipped into a boi	ling tube
	VX/I	rich one of	lute sulphuric a	cid.	exide of Y:	the formula of the	Deturmine
	V AA 1	Coloral one of	the following o	bservations was	made?	(31 - 0	
		Colourless			,		0° A
		Yellow Pro					1000
		White Prec					0.0
			white fumes.				
10.	A	iuminium S	ulphate reacts	with Sodium hyd	lrogen carbonate	as below	
	Alz	$2(SO_4)_{3(aq)}$	+ 6NaHCO _{3(aq)}		U. O'L'O'L	$2AI(OH)_{3(S)} + 3Na_2$	$SO_{4(aq)} +$
		$O_{2(g)}$		9.77	Turn 21 etas	malma uniwelli ti i f	tito delidi.v
	Th	e Mass of a	luminium hydr	oxide precipitate	ed by reacting 34.	.2g of aluminium su	ılphate is
	(Al	= 27, S = 3	32, O = 16, H =	arearer than (Le	fing to Seldran	a surprice takes of B	
	Á.	15.6g	B. 21:2g				
	C.	78.0g		level various teo	neren all a T	nai i na oz ada	A1
	D.	156.0g			er of neutons.		d d b as n
11	۸ï	white com	nound T on stre	ng heating prod		es and finally a yell	ow solid
						Veder mid le	
		Zn ²⁺	ne test tube. 11	e possioie eation	10	45	TENDER OF
		Al ³⁺					
		Pb ²⁺			1	117	
	D.	Cu ²⁺	the following	omnounds deco	lourises acidified	Potassium mangar	toto (TIII)
12.	. W	nich one of	the following of	sompounds deed.	lourises actuilled	rotassium mangar	iate (VII)
			om temperature	hine a	inudalu23u natu	aciji promenje Dričenje promenje	
		C ₂ H ₆		Certillani 192	ation of mappings	able E Olygen	
B		C ₂ H ₄			Z TRUUMIN)	arim a til lom	40-5-2
	-	C4 H10					W. P. T. T.
	D.	C7 H16	Note: 100 in	Be . 0580 81 0	ongtont massiss (i	oritoria de la composición dela composición de la composición de la composición de la composición dela composición dela composición dela composición de la composición dela composición	Which of
13.	. Sc	ome Zinc Su	alphate crystals	were neated to c	onstant mass wit	h the following resi	ults. Mass
	of	Crucible =	20.0g	0.5 514	adiologica de la compa	The state of the s	ron I
		- Compai	his and crustal	s = 75.740	the state of the s		

Mass of Crucible and residue = 23.22g	10.0
From the data the value of X in the formula ZnSO ₄ .X H ₂ O	is
A. 1	2 × Ze. 1 × de. 3
B. 2	1177 / 5:1
C. 7	
D. 10	E 22 4 23 4 443
14. Which one of the following contains elements that make u	n duralumin?
1 71 1 7	
A. Zinc and Copper only B. Aluminium and Magnesium only	7 = 25 .5 0 = 16, K = 39 , Lano
D. Aluminium, Magnesium and Copper	the selection of the following subs
15. Which one of the following precipitates is soluble in dilute	e ammonia solution?
A. Fe (OH) ₂	H MgCL
B. AgCl.	(slix)
	- Cab .C
C. Al(OH) ₃	
D. Pb(OH) ₂ 16. 1.60g of the oxide of a metal Y gave 1.44g of the metal when the state of the	hen reduced in hydrogen
Determine the formula of the evide of V	have abunding state and and analysis
Determine the formula of the oxide of Y;	shed one of the following alsor
(Y = 63.5, O = 16)	coitulos realm
A. YO	B. Yellow P. ecipitate
B. YO ₂	C White Proofpiate
C. Y ₂ O ₃	D. D. Dense while lands
D. Y ₂ O	periodic table
17. A and B are atoms of elements in the same period 2 of the	periodic table.
A. is in group II and B is in group III.	
Which of the following statements is true? A. B has one more Proton in its nucleus than A.	the Vilve of aluminium hydroxid
B. The relative atomic mass of B must be one unit greater than	the relative atomic mass
	§2,12.18 50.21 U.
of A. C. A has one electron more than B in the outermost energy le	evel.
C. A has one electron more than b in the outermost energy is	gster all
D. A and B have the same number of neutrons.18. Which of the following is the correct order with which have	logens displace each other from
	the all sides over within hearth.
solutions of their salts? A. I Br C1	A Zeb
A. 1	* C
2,	41.5
$C. Cl \longrightarrow Br \longrightarrow I$	
D. Br Cl I I I I I I I I I I I I I I I I I I	Sodium hyd rovide solution for
19.28.40cm ³ of Sulphuric acid required 25.00cm ³ of 5.12m ³	Solution of cooperation
complete neutralization.	Sale to a design of the second
Calculate the molar concentration of Sulphuric acid.	3
A. 0.0265mol/dm ³ B. 0.0360mol/dm ³ C. 0.0480mol/dm	233
D. 0.0530mol/dm ³	the said also make the of Iron 0
20. Which of the following is the reason why Zinc is used for	r sacrificial protection of front
A. Zinc is less reactive than Iron	save succias apuding data at to
B. Iron is more electro positive than Zinc.	
C. Zinc forms a film of Zinc oxide on the surface of Iro	n. Transfer from the algebra of the second

D. Zinc and Iron react in the same way

1.8.50g of an impure sample of Iron	v way. V required inst 75 cm	3 - F2N / by	drochlori	c acid	to dissolv	vе
it and give neutral solution.	required just 75cm	1, 01 21M H	drocmon	c acid	to disser-	1

The percentage purity of the sample of Iron is (Fe = 56)

A. 63% B. 74%

C. 85% D. 90%

2

22. Ammonium chloride dissolves in water according to the following equation. $NH_4Cl_{(s)} + H_2O_{(l)}$ $\sim NH_4^+(aq) + Cl^-(aq) = +\Delta H$

What does the sign ⁺ΔH mean?

A. Reaction involved is Endothermic

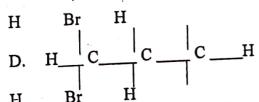
B. Reaction involved is Exothermic

C. Heat content of the Reactants is greater than the heat content of the Produc.'s.

D. Ammonium chloride is only slightly soluble in water.

23. Which one of the following is the structural formula of the compound formed when Ethene reacts with Bromine?





in the

24. What volume of oxygen at stp is required for complete combustion of 40cm3 of Butane according to the equation below?

$$C_{4}H_{10(g)} + \frac{13}{2}O_{2(g)} \longrightarrow 4CO_{2(g)} + 5H_{2}O_{(i)}$$

$$\Rightarrow 4CO_{2(g)} + 5H_{2}O_{(i)}$$

A.	40cm ³
B.	224cm

B. 224cm³ C. 260cm³ D. 448cm³

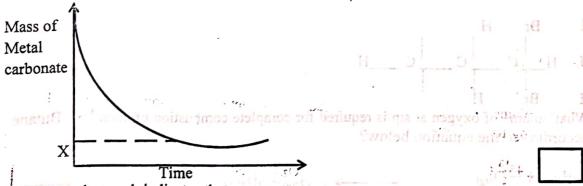
- 25. Which one of the following is Not a synthetic polymer?
 - A. Starch
 - B. Nylon
 - C. Perspex
 - D. Polyethene
- 26. Which one of the following is the correct overall cell equation in an electro-chemical cell made of Magnesium and Copper as electrode in dilute Sulphuric acid as the electrolyte?
 - A. $Mg(s) \rightarrow Mg^{2+}(aq) + 2e$

_{a)} + 2 e	\rightarrow	Cu _(s)
	{q)} + 2 e	${q)} + 2e \longrightarrow$

- C. $Mg(s) + Cu2+(aq) \longrightarrow Mg2+(aq) Cu(s)$
- D. $Mg(s) + 2H+(aq) \longrightarrow Mg^{2+(aq)}H^{2}(g)$.
- 27. Which one of the following elements does not exhibit allotropy?
 - A. Phosphorus
 - B. Sulphur
 - C. Chlorine.
 - D. Carbon.
- 28. Which one of the following observations is made when hydrogen Sulphide is bubbled into a solution of Lead (II) nitrate?
 - A. Yellow deposit
 - B. Dark brown precipitate. C. Colourless solution
 - D. Black precipitate.
- 29. Which one of the following gases is evolved when bleaching powder (Calcium hypochlorite) reacts with Carbondioxide?
 - A. Chlorine
 - B. Carbon-monoxide
 - C. Ammonia.
 - D. Oxygen.

Study the graph below and use it to answer questions 30 and 31.

The graph shows the reaction between metal Carbonate and a fixed volume of hydrochloric acid



30. Point X on the graph indicates that

A. Low yield of Carbondioxide is realized in a longer period of time.

B. Maximum yield of Carbondia	
B. Maximum yield of Carbondioxide is realized because all the acid! C. Mass of metal carbonate continues to decrease with time. Here would the metal of Carbondioxide is realized because all the acid!	has been used up
D. Reaction requires more time to reach completion. 31. How would the rate of reaction be determined to the determined of the acid of the	has been used up.
11 How would the rate of reach completion.	
A. By drawing a tangent of determined using the graph above?	med of all
B. By determining the point on the curve and determine its	slone
B. By determining the point at which the least decrease in mass occu By measuring volume of Carbondioxide over a long register of the	s stope.
C. By measuring volume of Carbondioxide over a long period of time. By determining the average mass of the metal Carbon to the	15.
D. By determining the average mass of the metal Carbonate that reac 32. When 2.5g of solid was heated, 560cm ³ of the gray was produced to the second of the gray was produced to the second of the gray was produced to the second of the gray was produced to the gray was p	te.
32. When 2.5g of solid was heated, 560cm ³ of the gas was produced at room residue of 1.4g was left.	m temperature and
residue of 1.4g was left.	in temperature and
The relative molecular mass of the gas is	******
(1 mole of gas at room temperature occupies 24dm³)	(3): (8) (1
E Super Co.	la maco e lu mm 1)
A. $\left[\frac{(2.5-1.4)X}{560 \times 1000}\right]$ cm ³ subressed accordance of the specific cm ³ subressed accordance cm ³ subressed accorda	non-redect TI over 60 25
A. L 560 X 1000 J	
[14 Y 24 Y 4000]	145 8
B. \[\frac{1.4 \times 24 \times 1000}{560} \] cm ³ \qquad \text{cm} \qquad \text{theoreov} \text{becaucity in theoreov} \text{theoreov} \tex	The form or resisting
B. 2 560 3	A Decreasus he con
	B. Im case on the pres
	C. Increasing concern
the policy of $\left[\frac{560 \times 2.5}{24 \times 1.4}\right]$ dm ³ so $\left[\frac{560 \times 2.5}{24 \times 1.4}\right]$ dm ³ so $\left[\frac{560 \times 2.5}{24 \times 1.4}\right]$	39, What is the new borro
D. L _{24 X 1.4} J dill	- g ki = () (E = 6d) -
[7] Let mole, C. (125 to a)	A Ulifornian
33. Hydrogen Sulphide burns in excess of oxygen according to the equation	on below.
$2H_2S_{(g)} + 3O_{2(g)}$ $\longrightarrow 2H_2O_{(l)} + 2SO_{2(g)}$	saca gulahide completely at
What is the volume of oxygen required to react with 48dm^3 of hydr	ogen surplinde completely at
room temperature and pressure?	
(All volumes are measured at same temperature and pressure)	
A. 24dm ³ And the stranged state of the second state of the secon	a rantizono sili In 12 x
	or we regar the right north
D. 72dm ³ 25g of Lead (II) oxide	according to the equation
D. 72dm ³ 34. What mass of Carbon is required to reduce 2.5g of Lead (II) oxide	at the state of
34. What mass of Carbon is required to reduce $a = b$ below. $2PbO(s) + C(s)$ $2PbO(s) + C(s)$ $2PbO(s) + C(s)$	B. J. Bolb R., d Settler
2PbO(s) + C(s)	HCL 0146 50
(Pb = 207, C = 12, O = 16)	is as no makes of the
(Pb = 207, C = 12, O = 10) A. 0.0335g $(Pb = 207, C = 12, O = 10)$ $(Pb = 207, C = 10)$ $(Pb =$	m a nomiced state of
B. 0.0673g	
C 0.01345g	
U. V.V.J.J.	
D. 0.2690g heating according to the equation.	
D. 0.2690g to the equation.	
D. 0.2690g to the equation.	of Calcium nitrate is
D. 0.2690g 35. Calcium nitrate decomposes on heating according to the equation. 35. Calcium nitrate decomposes on heating according to the equation. 2CaO(S) + 4NO _{2(g)} + O _{2(g)} 2Ca(NO ₃) ₂ (S) 3Ca(NO ₃) ₂ (S)	of Calcium nitrate is
D. 0.2690g 35. Calcium nitrate decomposes on heating according to the equation. 35. Calcium nitrate decomposes on heating according to the equation. 2CaO(S) + 4NO _{2(g)} + O _{2(g)} 2Ca(NO ₃) ₂ (S) 3Ca(NO ₃) ₂ (S)	of Calcium nitrate is
D. 0.2690g 35. Calcium nitrate decomposes on heating according to the equation. 35. Calcium nitrate decomposes on heating according to the equation. 2Ca(NO ₃) _{2(S)} 2Ca(NO ₃) _{2(S)} What is the volume of Nitrogen dioxide evolved at s.t.p when 4.1g of decomposed? A. 1.12dm ³	of Calcium nitrate is
D. 0.2690g 35. Calcium nitrate decomposes on heating according to the equation. 35. Calcium nitrate decomposes on heating according to the equation. 2CaO(S) + 4NO _{2(g)} + O _{2(g)} 2Ca(NO ₃) ₂ (S) 3Ca(NO ₃) ₂ (S)	of Calcium nitrate is

(1 mole of gas occupies 22	2.4dm^3 , $\text{Ca} = 40$, $\text{O} = 16$, $\text{N} = 14$)	
vilat vo	ide contained in 250cm ³ of solution was plume of the acid would be needed to react wit	used to titrate 0.1M
	or the acid would be needed to react wit	n 20cm ³ of the alkali?
B. 15cm ³ C. 30cm ³	- Tarrie a funcional distriction and the second	
D. 45cm ³	t militaria an militaria partital mana ilanggan ya arasa ilanggan ya arasa ilanggan ya arasa ilanggan ya arasa Tanggan ya arasa ilanggan ya arasa ilang	1 3/11
37. Ethane burns in air accordi	ng to the equation of the second	. Into a st
$2C_2 H_{6(g)} + 7O_{2(g)}$	10 the equation.	
Calculate the amount of he	\rightarrow 4CO _{2(g)} + 6H ₂ O _(l) Δ H = 1560KJ/mol at required to liberate 0.46dm ³ of Carbondioxic	nette e.
B. 6.24KJ	at required to liberate 0.46dm ³ of Carbondioxic	de. A. 2.34KJ
C. 12.48KJ	151	BW gr ' Carte
D. 16.02KJ	alər masç işilik gazlisi. — mar temperatum opcupic - 24 d m ³ i	Mill at hours
(1 mole of a gas at s.t.p oc	Curries 22 Adm3)	11 (11) SP 7 (4) COUNT
38. During the haber process N	Nitrogen reacts with hydrogen according to the	
me mader process, 1	vittogen reacts with hydrogen according to the	equation.
$N_{2(g)} + 3H_{2(g)} \sim$	2 NILL.	10001 2
The forward reaction in the	above case is favoured by	(to 2 24 X 1000
A. Decreasing the concentr	ration of repotents	One I of
B. Increasing the pressure		
C. Increasing concentration	of the system.	27-14) X240
D. Not applying a catalyst.		0.03
39. What is the number of mole	es of Sodium ions contained in 17.75g of Sodi	[01100]
(Na = 23, O = 16, S = 32)	es of Bodium ions contained in 17.75g of Sodi	um Sulphate solution?
A. 0.125 moles B. 0.10	60 males C 0 250 males	
D. 0.375 moles	of indices C. 0.250 moles	bide luic lagger shyllt DE 1
	nins 2.4×10^{23} atoms. ($Cu = 64$ Avogadro's c	
A. 20.45g B. 25.23g	atoms. (Cu – 64 Avogadro's c	constant = 6.02×10^{23})
C. 25.51g	lindaeng co	authorized ware
D. 51.02g	that a street terror sare and pressure)	STERSOMERON IN
	5 consists of an assertion (statement) on the	. 1.64
reason on the right hand side	e. Select:	e left – hand side and
A. If both the assertion and	I reason are true statements and the reason is	
of the assertion.	Transfer are true statements and the reason is	a correct explanation
B. If both the assertion and	reason are true statements but the reason is not	gols 4
of the assertion.	3 4	a correct explanation
	ut the reason is not a correct statement.	= 0. TUS = dq1
D If the assertion is not co	rrect but the reason is a correct statement.	
	resort out the rousen is a contest statement.	B12900 H
		86-100 D
•		9000 0 13
	scomposes on hearing according to the equation	Transfer apprairate de
	1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1	

2080(s) + 440(g) + 02(s)

in dioxide evolved at sitp when a lg or cialcies, err. me is

INSTRUCTIONS SUMMARISED		
Assertion	Reason	
A. True	True and reason is a correct explanation	
B. True	True but reason is not a correct explanation	
C. True	Incorrect	
D. Incorrect	True	

					the property of the second second
41	During electrolysis of concentrated sodium Solution, the PH of the the Cathode becomes a	Solution at	BECAUSI	Sodium ions real ions forming	et with excess hydroxide g Sodium hydroxide.
42	2. The heat of combustion is greater than that of F		BECAUSE	Propane has higher	molecular mass
43	Bond breaking is an en Process	dothermic B	BECAUSE	It involves the evolutions	n of heat to the urrounding.
44	4. Covalent compounds a Conduct electricity and Non - electrolytes			they consist of the c	of molecules and do not ontain ions.
4	5. All metals of group I i Periodic table are high electro positive	100	CAUSE	They all accept elector to form positive	trons in their reactions ions.
	In each of the question c	ons 46 – 50, or arefully and t	ne or more then indica	of the answers given the of the correct answer :	nay be correct. according to the followin
	A. If 1, 2 and 3 only as B. If 1 and 3 only are C. If 2 and 4 only are D. If 4 only is correct	correct.			
	D. II 4 only is correct	Ins	tructions s	summarized	
1.5	A	В		C	D
	1,2, and 3 correct	1 and 3 corre	ct 1/3	2 and 4 correct	4 correct

iormed.	igh dilute sodium hydroxide solution, the following
(1 Sodium chlorate.	
(2 sodium chloride (3 Oxygen gas	eng certifolists. Throughpuild so term highlige BECAUSE
(4 Sodium hypochlorite	BELAUSE SUPPLIED TO UP DESCRIPTION OF DESCRIPTION O
(+ Soutum hypochiorite	to noting the surface of the second
8. The following cations form precipitat	es with solutions of alkalis Except.
r (1-Fe ²⁺ To remoderate resident and annual	or instructions of antana language for the property of the pro
(2 Fe ³⁺	near resident menang BECAUSE
(3 Al ³⁺	440 A 348 Sindors to with any
(4 NH ₄ ⁺	•
	re formed when Ammonium nitrate decomposes?
(1 Nitrogen.	demone anaviorion a se m.
(2 Dinitrogen avida	
v. 1	
(3 Ammonia wall to break to be the	in 1916 or a strangeror
(3 Ammonia (4 Water	EUALEE THE LETTER OF THE LETTE
50. Which of the following ions in water i	form precipitates when washing is done with ordinary
50. Which of the following ions in water i	form precipitates when washing is done with ordinary
50. Which of the following ions in water i	form precipitates when washing is done with ordinary
50. Which of the following ions in water for the control of the following ions in water for the control of the following ions in water for the control of the following ions in water for the control of the following ions in water for the control of the following ions in water for the control of the following ions in water for the control of the following ions in water for the control of the following ions in water for the control of the following ions in water for the control of the following ions in water for the control of the following ions in water for the control of	
50. Which of the following ions in water i	form precipitates when washing is done with ordinary
50. Which of the following ions in water to (1 Ca ²⁺) (2 Na ⁺) (3 Mg ²⁺) (4 K ⁺)	form precipitates when washing is done with ordinary
50. Which of the following ions in water for (1 Ca ²⁺) (2 Na ⁺) (3 Mg ²⁺) (4 K ⁺)	form precipitates when washing is done with ordinary
50. Which of the following ions in water for (1 Ca ²⁺) (2 Na ⁺) (3 Mg ²⁺) (4 K ⁺)	form precipitates when washing is done with ordinary
50. Which of the following ions in water for (1 Ca ²⁺) (2 Na ⁺) (3 Mg ²⁺) (4 K ⁺)	form precipitates when washing is done with ordinary
50. Which of the following ions in water to (1 Ca ²⁺ (2 Na ⁺ (3 Mg ²⁺ (4 K ⁺ 1250103 ad gam navigations)	form precipitates when washing is done with ordinary
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