

# ORGANIC CHEMISTRY

## COURSE OUTLINE:

- Definition of organic Chemistry
- Properties of organic compounds
- Functional groups of organic compounds
- Molecular empirical structural formulae
- Homologous series (Groups of organic compounds)
- Alkanes
- Alkenes
- Alkynes
- Alkyl halides (Halogeno alkanes)
- Alcohols
- Ethers
- Phenols
- Benzene
  - ✓ Methy/Benzene
  - ✓ Nitrobenzene
  - ✓ Chlorobenzene
- Carbons/compounds
  - ✓ Aldehydes (Alkanals)
  - ✓ Ketones (Alkanones)
- Carboxylic acids
  - ✓ Acidic halides (Alkanoy chlorides)
  - ✓ Amides
- Esters
- Amines
- Diazonium salts
- Applied organic chemistry
- Practical organic chemistry

## Reference books:

Principles of Organic Chemistry  
A' level Chemistry – Ramsden.

## INTRODUCTION:

Definition of Organic Chemistry:

Organic Chemistry is the Chemistry of Carbon and related elements. It does not involve inorganic compounds of carbon like  $\text{CO}_2$ ,  $\text{CO}_3^{2-}$ s,  $\text{HCO}_3^-$ s.

Organic Chemistry started when the first organic compounds were largely obtained from living things. This therefore makes carbon as the most abundant element in living organisms.

### PROPERTIES OF ORGANIC COMPOUNDS:

Organic compounds are characterized by the following properties:

- They are made up of carbon and other elements like  $\text{H}_2$ ,  $\text{O}_2$ ,  $\text{N}_2$ , S, P and  $\text{Cl}_2$ .
- They are complex in structure e.g. some are made up of long chains of carbon, highly branched chain, rings of carbon.
- They exhibit isomerism a condition in which several structures can be written from one molecular formula.
- All organic compounds are bonded by covalent bonding.
- Most organic compounds undergo slower reactions.
- From the above properties, organic chemistry places a lot of interest in the behaviour of carbon.

### Carbon as an element in organic chemistry:

$\text{C}(6) = 2:4 \quad 1\text{S}^2 2\text{S}^2 2\text{P}^2$   
Electronic configuration

From the above configuration, carbon has got two principle quantum numbers, a total of 6 electrons.

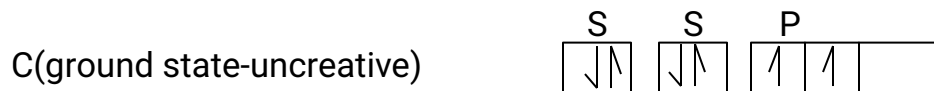
2 electrons in the S-orbital (inner energy level).

4 electrons in the 2<sup>nd</sup> principle quantum number distributed as 2 in the S-orbital and 2 in P-orbital.

When carbon undergoes a chemical reaction, it uses all the four electrons in the P-quantum number which gives it a tetravalent.

Since bonding in organic compounds occurs in covalent bond (sharing electrons), it means each carbon atom should contribute four (2 pairs of electrons). Carbon then forms 4 covalent bonds by molecular mixing of orbitals a process known as Hybridization.

### HYBRIDIZATION OF ORBITALS IN CARBON

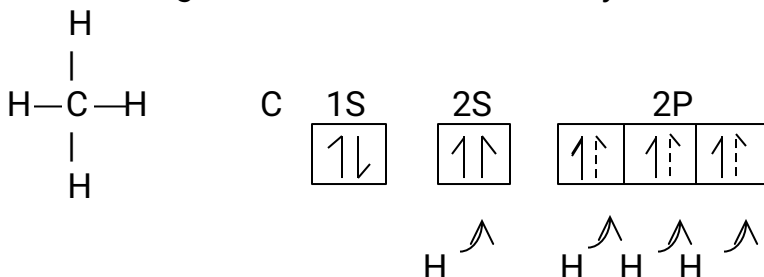


When carbon atom is excited, one electron from the 2S – orbital is sent to the 2P giving in total four unpaired electrons which now become the tetravalent of carbon.

Bonding in carbon thereof involves the mixing of the S – and the P-orbitals.

(i) Formation of alkanes (Example methane).

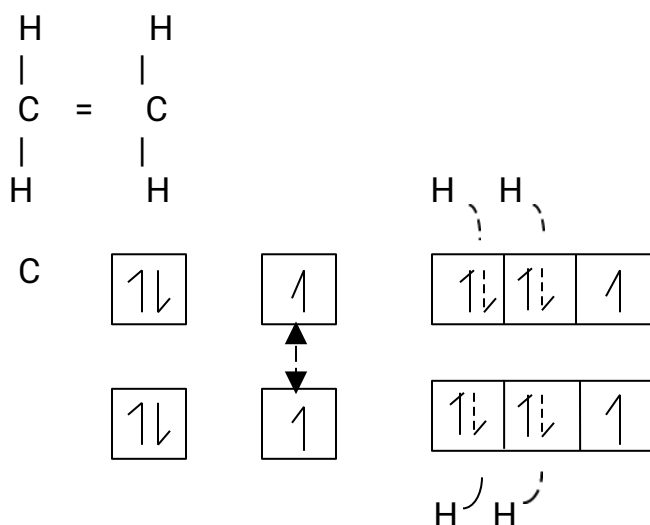
Bonding in alkanes involves  $2SP^3$  hybridization.



(ii) **Bonding in alkenes**

In alkenes, bonding involves 2Ps hybridization.

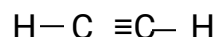
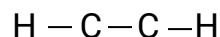
Ethene:



(iii) **Bonding in alkynes**

$2SP$  hybridization e.g. ethyne

i.e. 1 electron in 2S of each carbon atom combines to form a normal sigma bond while each hydrogen combines with each of the 2SP electron leaving 2 unpaired electrons in the 2P which will overlap to form  $2\pi$  bonds.

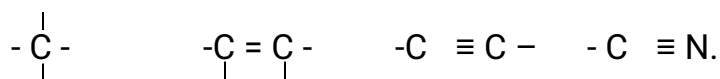


### UNIQUE PROPERTIES OF CARBON

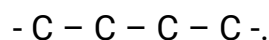
Carbon is generally a unique element and it displays the following properties:

- (i) If forms multiple bonds with other elements or itself. E.g.
- double bond  $\text{C} = \text{C}$  (alkene);  $\text{C} = \text{O}$  (carbons)
  - triple bond  $\text{C} \equiv \text{C}$  (alkynes);  $\text{C} \equiv \text{N}$  (Nitrates)

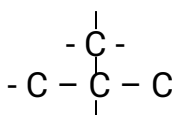
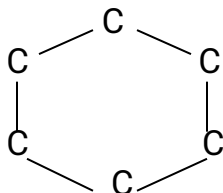
Carbon forms up to 4 covalent bonds and this is always exhibited in all its compounds.



Carbon forms long chain compounds a process known as catenation, e.g.



### Carbon rings and branched compounds.



### FUNCTIONAL GROUP:

Is an atom or groups of atoms in an organic compound that determines the chemical properties of that compound.

Most of the chemical reactions of organic compounds are determined by their functional group.

Families of organic compounds known as homologous series are also determined by functional group.

No.	Homologous series (class)	Functional group	Examples
1.	Alkanes $C_nH_{2n+2}$ (R – H)	$\begin{array}{c}   \quad   \\ -C - C- \\   \quad   \end{array}$ (Carbon single bond)	$H_4$ (Methane)
2.	Alkenes ( $C_nH_{2n}$ )	$\begin{array}{c}   \quad   \\ -C = C- \\   \quad   \end{array}$ (Carbon double bond)	$C_2H_4$ (Ethene)
3.	Alkynes ( $C_nH_{2n+2}$ )	$-C \equiv C-$ (Carbon triple bond)	$C_2H_2$ (Ethyne)
4.	Alkyl halides/halogen alkanes (R–x) $C_nH_{2n+1}X$ Where X is a halogen e.g. $Cl_2$ , Br.	Halogen(X)	$C_2H_5Cl$ (Chloro ethane)
5.	Alkanols (alcohols) ( $C_nH_{2n+1}OH$ )(R –OH)	Hydroxyl group (–OH)	$C_3H_7OH$ (Propanol)
6.	Ethers (R–OR)	– O – (Oxygen)	$C_2H_6O$ $\begin{array}{c} H \quad H \\   \quad   \\ H - C - O - C - H \\   \quad   \\ H \quad H \end{array}$ (di methyl) (ether of methoxyl)
7.	Carbonyl Compounds (i) Alkanals (aldehydes)	$\begin{array}{c} O \\    \end{array}$	$C_2H_4O$ $\begin{array}{c} H \quad O \\   \quad    \\ H - C - C - H \end{array}$

		- C – H (Carbony/Carbon & H <sub>2</sub> )	$\begin{array}{c} \text{H} \\ \text{(Ethanal)} \end{array}$
	(ii) Alkanones (Ketone) $\begin{array}{c} \text{O} \\    \\ \text{R} - \text{C} - \text{R} \end{array}$	$\begin{array}{c} \text{O} \\    \\ - \text{C} - \end{array}$ (Carbony/Carbon)	$\begin{array}{c} \text{C}_3\text{H}_6\text{O} \text{ (Propanone)} \\ \begin{array}{c} \text{H} \quad \text{H} \\   \quad   \\ \text{H} - \text{C} - \text{C} - \text{C} - \text{H} \\   \quad   \quad   \\ \text{H} \quad \text{O} \quad \text{H} \end{array} \end{array}$
8.	Carboxylic acid $\text{R} - \text{COOH}$	$\begin{array}{c} \text{O} \\    \\ - \text{C} - \text{O} - \text{H} \end{array}$ Carboxylic acid group	$\begin{array}{c} \text{C}_2\text{H}_4\text{O}_2 \text{ (Ethanoic acid)} \\ \begin{array}{c} \text{H} \quad \text{O} \\   \quad    \\ \text{H} - \text{C} - \text{C} - \text{O} - \text{H} \\   \\ \text{H} \end{array} \end{array}$
9.	Esters (ROOR)	$\begin{array}{c} \text{O} \\    \\ - \text{C} - \text{O} - \\ / \end{array}$	$\begin{array}{c} \text{C}_3\text{H}_6\text{O}_2 \text{ (Methyl ethanoate)} \\ \begin{array}{c} \text{H} \quad \text{O} \quad \text{H} \\   \quad    \quad   \\ \text{H} - \text{C} - \text{C} - \text{O} - \text{C} - \text{H} \\   \quad   \quad   \\ \text{H} \quad \text{H} \quad \text{H} \end{array} \end{array}$
10.	Nitro compound $\text{R} - \text{NO}_2$	-NO <sub>2</sub> (Nitro group)	$\begin{array}{c} \text{C}_2\text{H}_4\text{NO}_2 \text{ (Nitro ethane)} \\ \begin{array}{c} \text{H} \quad \text{H} \\   \quad   \\ \text{H} - \text{C} - \text{C} - \text{NO}_2 \\   \quad   \\ \text{H} \quad \text{H} \end{array} \end{array}$

11.	Nitriles (Cyanohydrins) $R - C \equiv N$	$-C \equiv N$ (Nitrile group)	Ethanitrile $C_2H_3N$ $\begin{array}{c} H \\   \\ H - C - C \equiv N \end{array}$
12.	(Acido halides) $\begin{array}{c} O \\    \\ R - CO - X \end{array}$	$\begin{array}{c} O \\    \\ -C - X \end{array}$	$C_2H_3O_4Cl$ $\begin{array}{c} H \quad O \\   \quad    \\ H - C - C - X \end{array}$
13.	Amides $\begin{array}{c} O \\    \\ R - C - NH_2 \end{array}$	$\begin{array}{c} O \\    \\ -C - NH_2 \end{array}$ (Amide group)	$C_2H_5 - ON$ $\begin{array}{c} H \quad O \\   \quad    \\ H - C - C - NH_2 \end{array}$
14.	Amines $(R - NH_2)$	$-NH_2$ (Amino group)	$\begin{array}{c} H \\   \\ H - C - NH_2 \end{array}$

## EMPIRICAL MOLECULAR AND STRUCTURAL FORMULAE

Empirical formula is the simplest formula which expresses the ratio of the number of atoms present in a molecular or compound. E.g.

A certain organic compound has 39.13% by mass Carbon 52.23%, Oxygen and the rest hydrogen, determine its empirical formula.

$$39.13 + 52.23$$

$$100 - 91.36 = 8.64$$

Element	C	H	O
Percentage composition	39.13	8.64	52.23
<u>Composition</u>	<u>39.13</u>	<u>8.64</u>	<u>52.23</u>
Rfm	12	1	16

Number of moles	3.2608	8.64	3.2644
Divide by smallest	$\frac{3.2608}{3.2608}$	$\frac{8.64}{3.2608}$	$\frac{3.2644}{3.2608}$
Mole ratio	1 x 2	2.6 x 2	1 x 2    1 x 3    2.6 x 3
1 x 3	1	5.3	1 $C_3H_8O_3$

$\therefore$  Empirical formula =  $C_2H_5O_2$ .

Molecular formula expresses the true or actual number of each atom present in 1 mole of a compound. It is also a multiple of an empirical formula.

$$\frac{\text{Molecular formula mass}}{\text{Empirical formula mass}} = n = \text{simple whole number.}$$

Molecular formula = [Empirical formula].

### MOLECULAR FORMULA MASSES:

They can be determined from;

- (i) Vapour density:  
When a compound is vapourized and weighed, its mass is twice the molecular formula mass.  
 $2 \times V.P = MFM.$

- (ii) Rates of diffusion (Graham's law)

$$MFM_A = \left( \frac{R_B}{R_A} \right)^2 MFM_B$$

- (iii) Ideal gas equation

$$PV = nRT; PV = \frac{mRT}{mR}$$

$$Mr = \frac{mRT}{PV}$$

### GASEOUS HYDROCARBONS (EUDIOMETRY)

Eudiometry is a technique used to determine molecular formula of gaseous hydrocarbons.

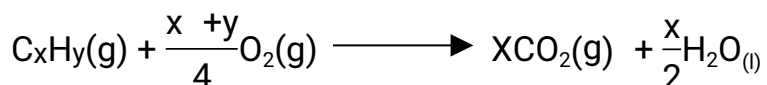


When a gaseous hydrocarbon is exploded with excess oxygen, it readily burns forming  $\text{CO}_2$  and  $\text{H}_2\text{O}$  as the only product.

If the gases are passed over an alkali e.g. (KOH or NaOH), all the  $\text{CO}_2$  formed is absorbed and the diminution (reduction) in the volume of the gases is due to  $\text{CO}_2$ .

The residual gas at that point will now be excess  $\text{O}_2$ .

The volume of  $\text{H}_2\text{O}$  formed is regarded as negligible.



From the above equation, every one mole of a hydrocarbon is equivalent to X moles of  $\text{CO}_2$  produced.

Similar when the volumes are related, the volume of  $\text{CO}_2$  is equivalent to (X) x volume of hydrocarbon.

$X(\text{Volume of HC}) = \text{Volume of } \text{CO}_2 \text{ produced.}$

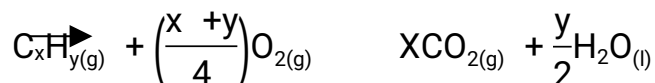
Then the volume of  $\text{O}_2$  used can also be related to the volume of the hydrocarbon by volume of;

$$\text{Oxygen used} = \left(\frac{x+y}{4}\right)(\text{Volume of hydrocarbon}).$$

### Example:

$15\text{cm}^3$  of a gaseous hydrocarbon were exploded with  $105\text{cm}^3$  of oxygen in a sealed vessel after cooling, the residual volume occupies  $75\text{cm}^3$ . On addition of caustic potash, there was a final diminution/decrease of volume to  $30\text{cm}^3$ . Determine the molecular mass of a hydrocarbon.

Let the hydrocarbon be  $\text{C}_x\text{H}_y$ .



Volume of  $\text{O}_2$  used =  $105 - 30 = 75\text{cm}^3$

Volume of  $\text{CO}_2$  produced =  $75 - 30 = 45\text{cm}^3$

$\text{C}_x\text{H}_y = \text{volume of } \text{CO}_2 \text{ produced.}$

$$= 45.$$

$$\frac{45}{15} = 3.$$

$$(\text{Volume of HC}) \left( \frac{x+y}{4} \right) = \text{Volume of O}_2 \text{ used.}$$

$$15 \left( 3 + \frac{y}{4} \right) = 75$$

$$3 + \frac{y}{4} = 5$$

$$3 + \frac{y}{4} = \frac{75}{15}$$

$$\frac{y}{4} = 2$$

$$y = 8.$$

Molecular formula is  $\text{C}_3\text{H}_8$

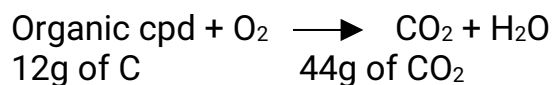
$$12 \times 3 + 8 \times 1 = 44$$

## DETERMINATION OF MOLECULAR FORMULA OF ORGANIC COMPOUND USING MASS OF $\text{CO}_2$ and $\text{H}_2\text{O}$ .

When an organic compound is burnt, the masses of the products, carbon dioxide and  $\text{H}_2\text{O}$  together with the initial mass of the organic compound can be used to determine the molecular formula.

### Example:

0.464g of an organic compound Q when burnt gave 1.32g of  $\text{CO}_2$ , 0.315g of  $\text{H}_2\text{O}$ . When 0.2325g of Q were separately burnt,  $7.08 \text{ cm}^3$  of  $\text{N}_2$  gas at s.t.p was produced. Determine the molecular formula of Q and draw its structure. Vapour density = 46.5.



2g of H                      18g of H<sub>2</sub>O.

12g of carbon produces 44g of CO<sub>2</sub>.

Xg of carbon produce 1.32g of CO<sub>2</sub>.

$$\left(\frac{12}{44} \times 1.32\right) \text{g of carbon inorganic compound.}$$
$$= 0.36\text{g.}$$

$$\text{Percentage of carbon} = \frac{0.36}{0.464} \times 100$$
$$= 77.6\%$$

2g of H<sub>2</sub> in Q produces 18g of H<sub>2</sub>O.

y of H<sub>2</sub> in Q produces 0.315g of H<sub>2</sub>O.

$$y = \left(\frac{2}{18} \times 0.315\right) \text{g of H}_2$$
$$= 0.035\text{g.}$$

$$\text{Percentage of H in Q} = \frac{0.035}{0.464} \times 100$$
$$= 7.5\%$$

$$\text{Mass of Nitrogen} = 0.36 + 0.035 = 0.395$$
$$0.464 - 0.395$$
$$= 0.069\text{g of N}_2.$$

$$\text{Percentage of N}_2 = 77.6 + 7.5 = 85.1\%$$
$$(100 - 85.1)$$
$$= 14.9\%$$

1mole of a gas contains 22400cm<sup>3</sup> at s.t.p.

Xmoles of N<sub>2</sub> contains 7.08cm<sup>3</sup>.

$$\frac{7.08}{22400} = 3.16 \times 10^{-4} \text{ moles.}$$

0.2325g of Q produces 3.16 x 10<sup>-4</sup> moles of N<sub>2</sub>.

Xg of Q produced by 1mole of N<sub>2</sub>.

Elements	C	H	N
Percentage composition	77.6	7.5	14.9

Relative atoms	$\frac{77.6}{12}$	$\frac{7.5}{1}$	$\frac{14.9}{14}$
Moles	6.47	7.5	1.06
Simple ratio	$\frac{6.47}{1.06}$	$\frac{7.5}{1.06}$	$\frac{1.06}{1.06}$
Simple ratio	6	7	1

Empirical formula  $C_6H_7N$

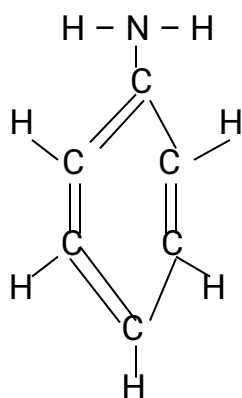
$$(C_6H_7N)_n = 46.5 \times 2$$

$$n(12 \times 6) + 7 \times 1 + 14 = 93$$

$$93n = 93$$

$$n = 1.$$

Molecular formula =  $C_6H_7N$

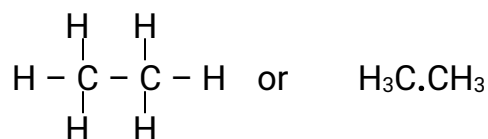


Phenyl-amine.

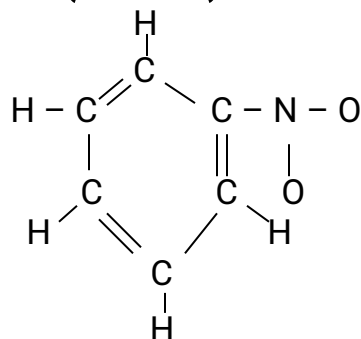
## STRUCTURAL FORMULA

It shows the different atoms and how they are connected in one molecule of a compound e.g.

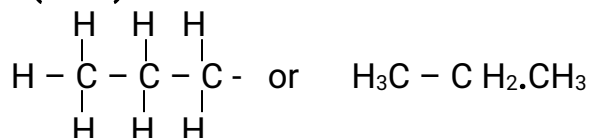
1. **Ethane structure.**



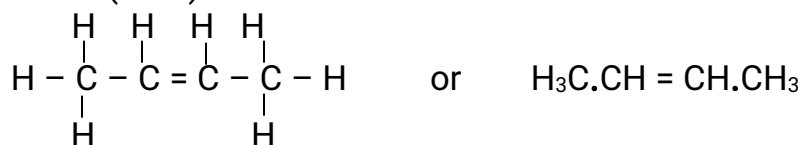
2. **Nitrobenzene (C<sub>6</sub>H<sub>5</sub>NO<sub>2</sub>).**



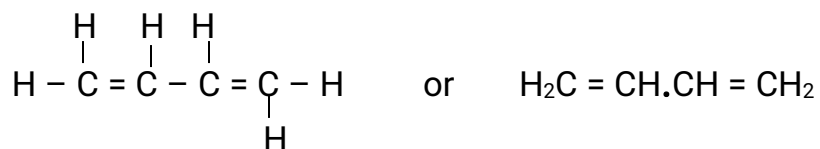
3. **Propane (C<sub>3</sub>H<sub>8</sub>)**



4. **But-2-ene (C<sub>4</sub>H<sub>8</sub>)**



5. **Buta-1,3-diene**



## ORGANIC COMPOUNDS


Organic compounds are generally classified according to the functional groups they do have and it is very important that all the organic compounds are given names.

## NOMENCLATURE OF ORGANIC COMPOUNDS

Naming of organic compounds follows a strict system of nomenclature where names are assigned.

## NAMING OF HYDRO CARBONS

They are named using “stem” names that indicate the alkyl groups. E.g.

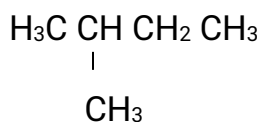
Alkyl group:	Name:	
$\text{---}^+\text{CH}_3$	1	Methyl
$\text{---}^+\text{CH}_2\text{CH}_3$	2	Ethyl
$\text{---}^+\text{CH}_2\text{CH}_2\text{CH}_3$	3	Propyl
$\text{---}^+\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$	4	Butyl
$\text{---C}_6\text{H}_5$ or 	6	Phenyl

## RULES GOVERNING NOMENCLATURE (Guidelines)

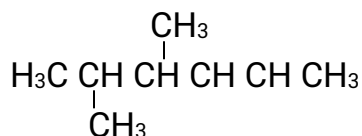
- Hydrocarbons are named as derivatives of alkanes. Where the suffix -ane is added to the stem name.
- Always identify the longest continuous chain in the compound and this gives the parent name. Other things attached to that chain are regarded as branches. E.g.



- Number – the carbon parent chain from one end to the other end giving the position of the branch the lowest number. E.g.



- If there are more than one substituent branch, the numbering has to consider the lowest sum of the positions. E.g.



R – L – sum of branches = 4 + 5 = 9.

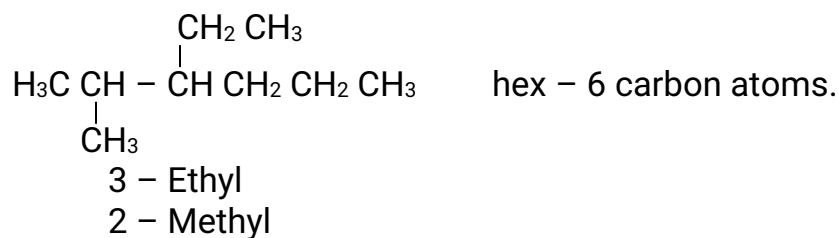
L – R, sum of branches = 2 + 3 = 5  $\longrightarrow$  selected.

The sum of the locants moving from left to right is the lowest and therefore it is preferred.

- (v) All the positions of the locants are identified and they are put as prefix separated with a (-) in the branch. E.g.  
 2 – methyl  
 3 – methyl

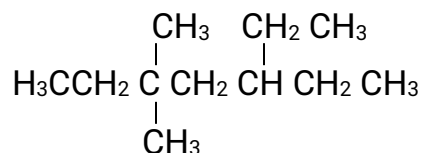
- (vi) When there are more than one substituent, the following is noted:

- (a) Similar substituents are given a prefix di – 2, tri – 3, tetra – 4 ---- and their locants are also indicated before the prefix but the locant separated by a comma(.). E.g. 2, 3-dimethyl.  
 The numerical order of the locants is followed.



The alphabets in the beginning letter of the names of the substituent has got to be followed.

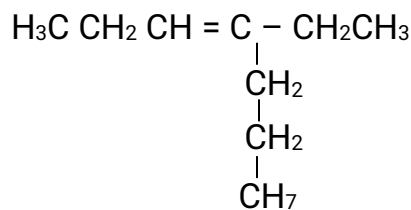
- (vii) The last name to be written is the parent name which indicates the number of carbon atoms in the chain.



3, 3-dimethyl -5-ethyl heptanes.

If a multiple bond (triple or double bond) is present, then the longest chain identified must involve the (double or triple) bond.

The multiple bond must be given the lowest number within the stem name.



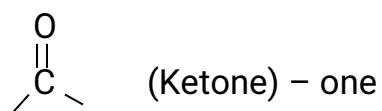
4-ethyl hept-3-ene.

- (viii) When there are other functional groups, their positions, names are indicated. E.g. Halogen (C/, F, Br, I)= Chloro, Fluoro, Bromo, Iodo.

OH = (alcohol) – O/

C = C (double bond) – ene

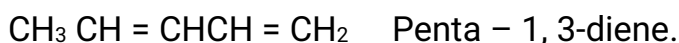
C  $\equiv$  C (triple bond) – yne.



-NH<sub>2</sub> (amine) – amine.

- (ix) **VOWEL/CONSONANT RULE**

When naming a compound where a locant is included within the stem name, the two letters close to the locant must be a vowel and a consonant and not all vowels or all consonants. E.g.



- (x) When naming compounds with very many functional groups (poly functional compounds), the functional groups are assigned an order depending of the seniority order and this order is the carboxylic acid, carbonyl, cyanide, ketone, alcohol, amines, double bond, triple bond, nitro group and halogen.





-COOH > C-H > C≡N >  $\overset{\text{Ketone}}{\text{C}} > \overset{\text{Hydroxyl}}{\text{C}} > \text{-OH} > \text{-NH}_2 > \text{-C=C-} >$

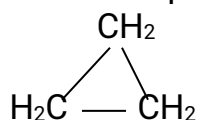
C≡C > NO<sub>2</sub> > X(halogen).

E.g.  $\text{CH}_3\underset{\text{C/}}{\text{CH}}\text{CH}_2\text{COOH}$  3-chloro butanoic acid.

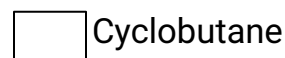
$\text{CH}_3\underset{\text{C/}}{\text{CH}}-\underset{\text{NH}_2}{\text{CH}}\text{CH}=\text{CH}_2$  4-chloro-3-amino pentene.

$\text{H}_2\text{C}=\text{CHCH}_2\text{CH}_2\text{C}\equiv\text{CH}$  Hexane-5-yne.

For the case of cyclic compounds (ringed compounds), the prefix "cyclo" is added to the parent alkane name.



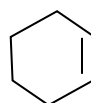
Cyclopropane



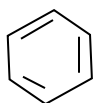
Cyclobutane



Cyclopentane



Cyclohexene



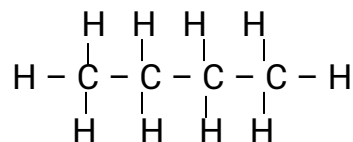
Cyclohexa-1, 3, 5 - triene.  
Or  
Benzene (Phenyl group)

## STRUCTURES OF ORGANIC COMPOUNDS FROM NAMES:

Structures can be written from the names.

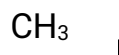
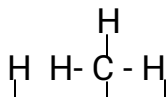
**Example:**

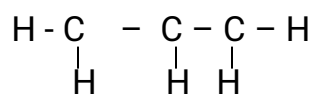
**Butane**



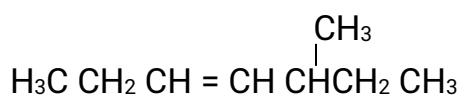
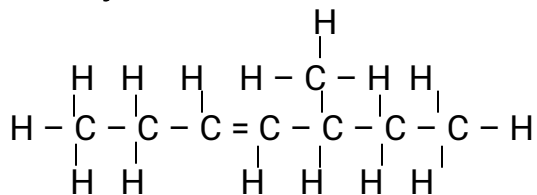
Or  $\text{H}_3\text{CCH}_2\text{CH}_2\text{CH}_3$

**2-Methyl propane.**

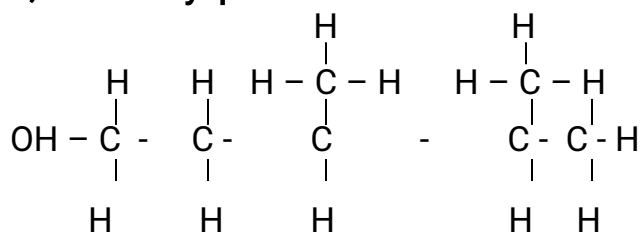




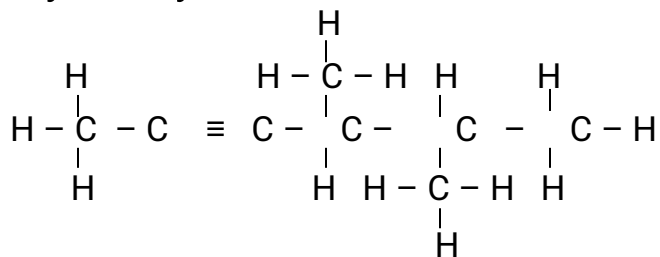
**5-methyl hept-3-ene.**



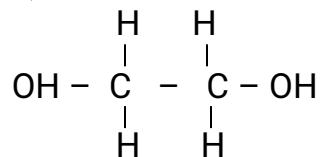
**3, 4-dimethyl pentanol.**



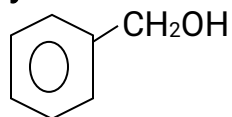
**4, 5-dimethyl hex-2-yne.**



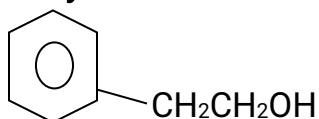
**Ethane-1, 2 – diol**



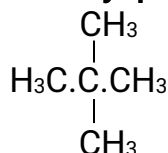
**Phenyl methanol**



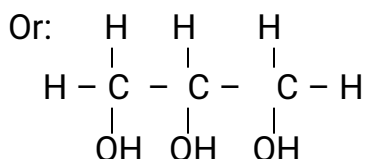
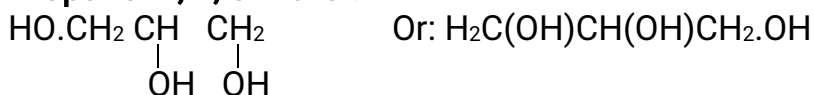
### 2-phenyl ethanol



### 2, 2 –dimethyl propane



### Propane- 1, 2, 3 – triol.



### *International Union of Pure and Applied Chemistry.*

**N.B:** Name organic compounds should always consider the IUPAC system. Where names are systematically written and this is the most universally known way of naming. However, trivial names may be found in a number of books and these should not be used for exam purposes.

E.g.  $\text{CH} \equiv \text{CH}$  (Ethyne) = (Acetylene) trivial name.  
 $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$  Propanol = (Propyl alcohol) trivial name.

## ISOMERISM

This is the existence of compounds with the same molecular formula but different structural formulae. These compounds are known as **Isomers**.

### Types of isomerism

There are two types of isomerism with a number of subtypes.

#### 1. **Structural isomerism**

The type of isomerism where compounds differ in the arrangement of atoms within the molecule i.e. isomers under structural isomerism differ in the carbon skeleton.

#### 2. **Stereo isomerism**

Type of isomerism where compounds of the same molecular formula and

structural formula differ in the spatial arrangement of atoms.

## STRUCTURAL ISOMERISM

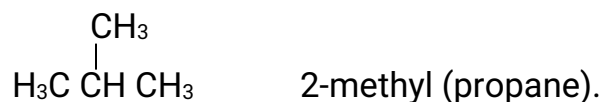
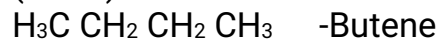
Structural isomers can be further subdivided into:

(i) **Chain structural isomerism**

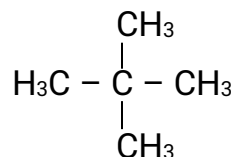
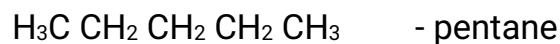
This is where compounds differ in the arrangement of carbon atoms in a molecule.

**Note:** The position of the functional group in the chain must remain the same.

E.g. ( $C_4H_{10}$ )



$C_5H_{12}$



(ii) **Functional isomerism**

Type of isomerism where the isomers differ by the functional group. Such isomers also differ in both physical and chemical properties. E.g. Alcohols.

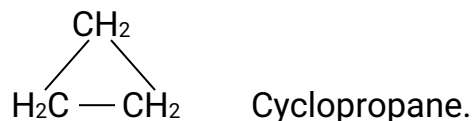
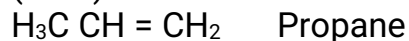
- **Ethers are isomeric.**

$C_2H_6O$

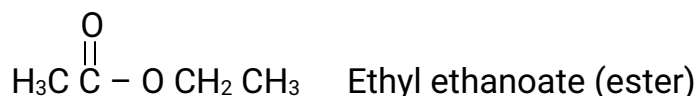
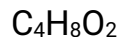


- **Alkenes and cycloalkanes**

$(C_3H_6)$



- **Carboxylic acids and esters.**



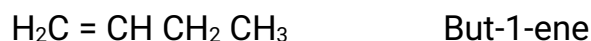
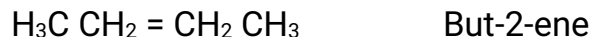
(iii) **Positional isomerism**

Type of isomerism where the compounds with the same molecular formula differ in position of the functional group within the same carbon skeleton.

- **(C<sub>3</sub>H<sub>7</sub>OH)**



- **C<sub>4</sub>H<sub>8</sub>**

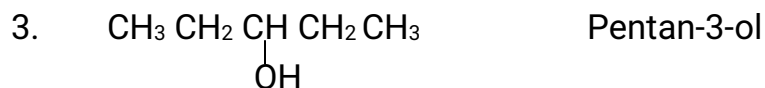
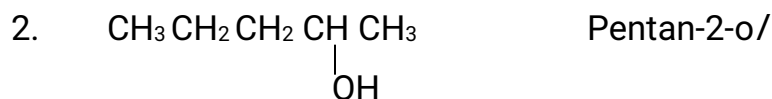


- **(C<sub>4</sub>H<sub>9</sub>Cl)**



**Question:**

Write all the isomers of the compounds with molecular formula C<sub>5</sub>H<sub>12</sub>O.



- CH<sub>3</sub>
5.  $\begin{array}{c} \text{CH}_3 \\ | \\ \text{CH}_3 - \text{C} - \text{CH}_2\text{OH} \\ | \\ \text{CH}_3 \end{array}$  2, 2-dimethyl propanol
6.  $\begin{array}{c} \text{CH}_3 \text{ CH}_2 \text{ CH} \text{ CH}_2 \text{ OH} \\ | \\ \text{CH}_3 \end{array}$  2-methyl butanol
7. CH<sub>3</sub> CH<sub>2</sub> O CH<sub>2</sub> CH<sub>2</sub> CH<sub>3</sub> Ethoxy propane.
8. CH<sub>3</sub> O CH<sub>2</sub> CH<sub>2</sub> CH<sub>2</sub> CH<sub>3</sub> Methoxy butane.

Note: 1, 2, 3 are positional isomers.  
 4, 5, 6 are chain isomers.  
 7 and 8 are functional isomers.

### STEREO-ISOMERISM

The compounds differ by spatial arrangement of atoms. There are two types i.e. geometrical stereoisomerism and optical stereoisomerism.

### GEOMETRICAL STEREOISOMERISM

Is where isomers have the same structural formula but differ in the spatial arrangement of the groups around a double bond.

Geometrical isomers arise as a result of a double bond which is planar and does not allow free rotation (restricted rotation) unlike a single carbon bond which freely allows rotation.

Geometrical isomers can therefore be;

(i) **Cis-geometrical isomer.**

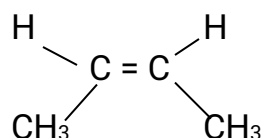
Is when identical groups of atoms are on the same plane of the double bond.

(ii) **Trans-geometrical isomer.**

Is when identical atoms are rotated on opposite planes of the double bond.

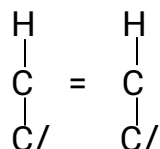
But-2-ene



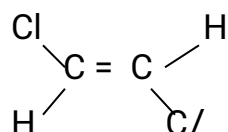


Bis but-2-ene

1, 2-dichloro ethene.

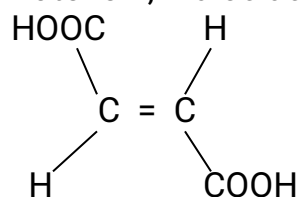


Cis 1,2 dichloro ethene.

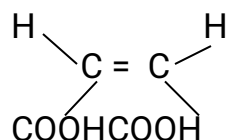


Trans 1, 2 dichloro ethene

Butene-1, 4-dioic acid



Trans butane-1, 4-dioic acid.



Cis butane- 1, 4-dioic acid

## OPTICAL ISOMERISM

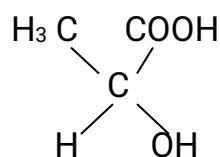
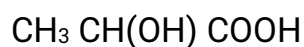
Type of isomerism that arises when two or more compounds with the same molecular or structural formula have got different optical properties in rotating the plane of polarized light.

The optical isomers are usually not super imposable onto each other and they have got marked differences in the physical and chemical properties.

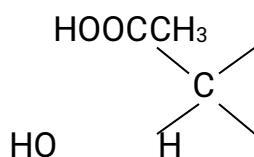
For any optical isomers to exist, they must have a mirror image of each other and rotating a plane of polarized light either to the right or to the left.

When an isomer rotates light to the right, it is known as **DEXTROROTARY** (D-isomer) (+ve isomer).

When it rotates light to the left, it is known as **LEVOROTARY** (L-isomer or –ve isomer). E.g. lactic acid (2-hydroxy propanoic acid).



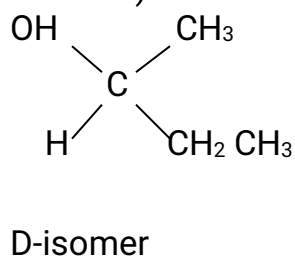
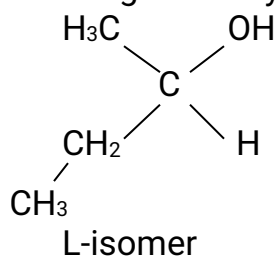
D-isomer



(L-isomer)

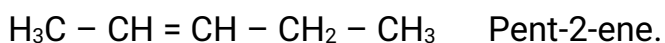
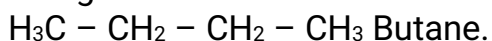
### CONDITIONS OF GEOMETRICAL ISOMERS

A compound must have asymmetrical centre. This is a centre joined by four different groups of atoms and that carbon centre is said to be asymmetric carbon. E.g. 2-methyl butan-2-ol (amyl alcohol).

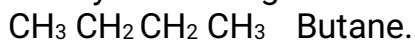


### Ways of writing organic compounds

- (i) Using carbon-carbon bonds.



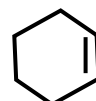
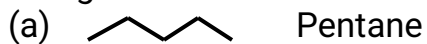
- (ii) Plainly illustrating atom.



- (iii) Using the molecular formula.

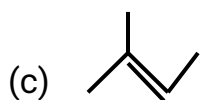


- (iv) Using skeletal structure.

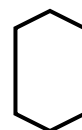


Cyclohexene.





2-methyl but-2-ene



Cyclohexane.

**Note:** No. (iv) is not commonly used at A' level.

## ORGANIC REACTIONS

Generally most organic reactions are slow compared to inorganic reactions. At the same time, they yield different product depending on the conditions of reaction. When considering a particular organic reaction, conditions at which it takes place must be specified.

### TYPES OF ORGANIC REACTIONS.

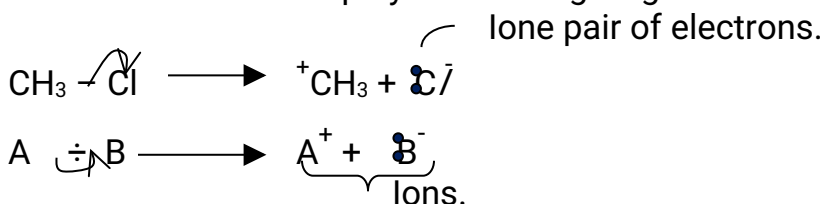
There are several types of organic reactions, categorized under different ways.

#### Bond breaking in organic reaction.

There are two ways in which bonds can be broken in organic reactions.

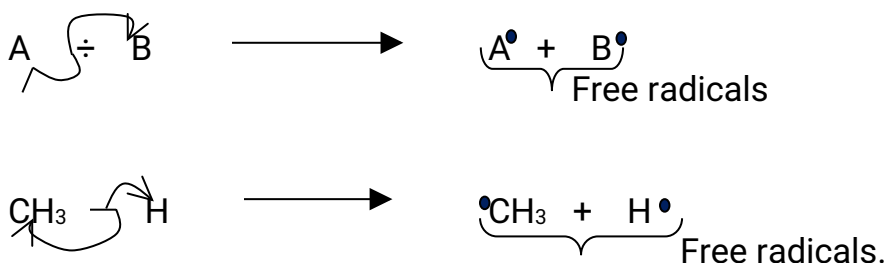
##### (a) Heterolytic fission.

Here an organic bond is broken and the resulting electrons within a broken bond are then taken up by one atom giving rise to ions. E.g.



##### (b) Homolytic fission

Type of bond breaking where by the resulting electrons in the broken bond are equally shared by the two atoms. This results into atoms with unpaired electrons which are called free radicals. E.g.



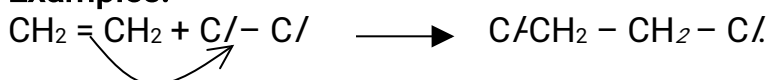
## TYPES OF ORGANIC REACTIONS

### 1. ELECTROPHILIC ADDITION REACTIONS

Electrophilic addition reactions are reactions in which an electrophile is added to a reacting substrate (molecule).

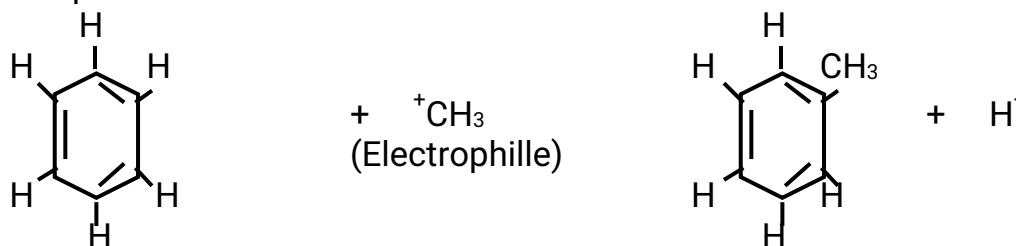
An electrophile is an electron seeking species or atoms. Electrophiles are usually positively charged. E.g.  $^+\text{NO}_3$ ,  $^+\text{C}/$ ,  $^+\text{CH}_3$ .

**Examples:**



### 2. ELECTROPHILIC SUBSTITUTION REACTION.

Type of organic reaction where an electrophile is substituted by another electrophile.



$^+\text{CH}_3$  substitutes  $\text{H}^+$ .

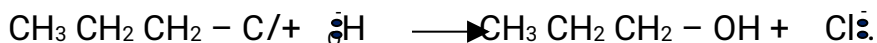
### 3. SUBSTITUTION NUCLEOPHILIC REACTIONS

Type of organic reaction where a nucleophile replaces another nucleophile in a reacting molecule.

A nucleophile is a species or group of atoms that seek for the positive centre of the reacting molecule. They either be negatively charged e.g.

$\text{:O}^-$ ,  $\text{Cl}^-$ ,  $\text{Br}^-$  or neutral (uncharged). E.g.  $\text{H}_2\text{O}$ ,  $\text{NH}_3$ .

All nucleophiles must have a lone pair of electrons.



Nucleophilic substitution reactions can either be;

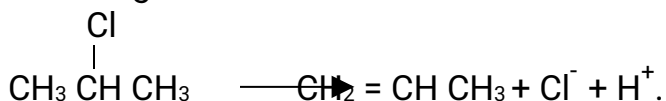
(i) Substitution nucleophilic bimolecular ( $\text{S}_\text{N}2$ ). Here two molecules are involved in the rate determining step of the reaction.

OR:

(ii) Substitution nucleophilic unimolecular ( $\text{S}_\text{N}1$ ). Here only one molecule is involved in the rate determining step of the rxn.

### 4. Elimination reactions (E).

These are reactions where an atom is completely removed from the molecule. This always results into the formation of a double or multiple bond. E.g.



**NOTE:** Elimination reactions can also be bimolecular ( $E_2$ ) or unimolecular ( $E_1$ ).

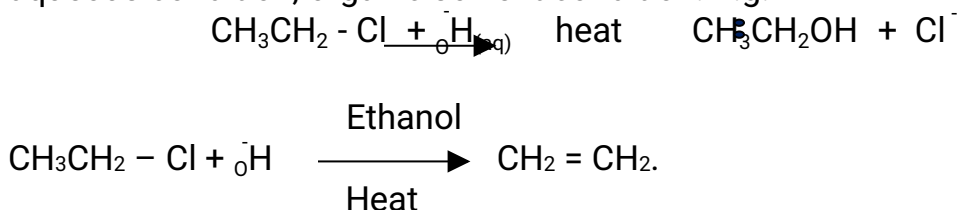
## FACTORS THAT AFFECT ORGANIC REACTIONS.

### (i) Conditions of temperature.

Temperature affects a number of organic reactions and therefore when writing those reactions, temperature has got to be specified.

### (ii) Nature of the solvent.

The nature of the solvent used affects the type of organic reaction e.g. aqueous condition, organic solvent condition. E.g.

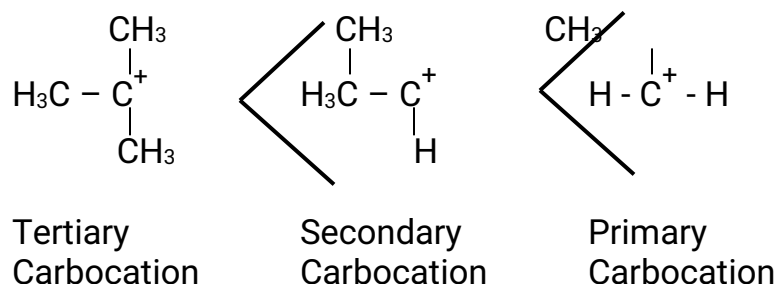


### (iii) The structure of the reacting molecule.

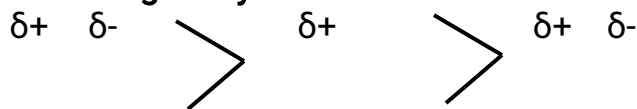
The structure of the reacting molecule can affect the organic reaction in several ways like.

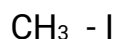
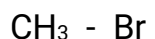
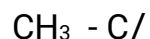
#### (a) Steric hindrance.

A condition in which the surrounding alkyl groups prevent a reagent from attaching the reaction centre e.g.



#### (b) Electronegativity





(c) **Bond length.**

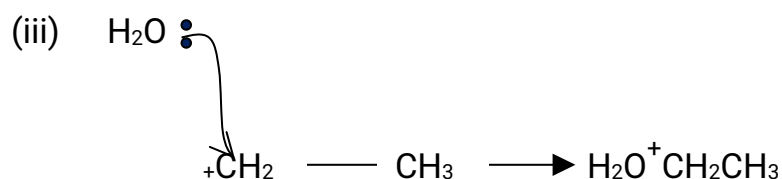
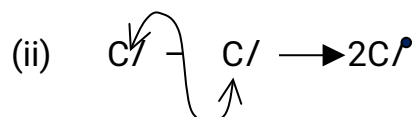
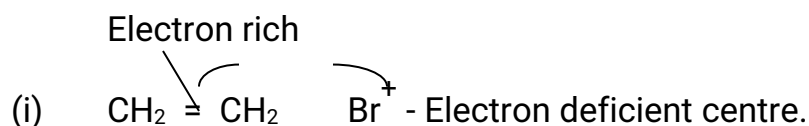
Usually longer bonds are easily broken than the shorter ones. If a reaction proceeds by breaking a particular bond, a longer bond can then easily be broken than a smaller one.

## MECHANISM OF REACTIONS

A mechanism of an organic reaction is an illustration by showing the flow of electrons from one centre to another. This flow is normally illustrated by means of curled arrows.

When a double headed arrow is used, it implies two electrons flowing i.e.  $\curvearrowright$  and when a single headed arrow is used, it implies one electron i.e.  $\rightarrow$ .

The arrow must always begin at the centre of the double bond where electrons are positioned and ends where electrons are deficient or it must begin at the centre of a lone pair of electrons. E.g.



## ALKANES

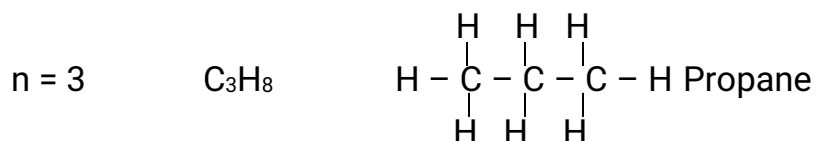
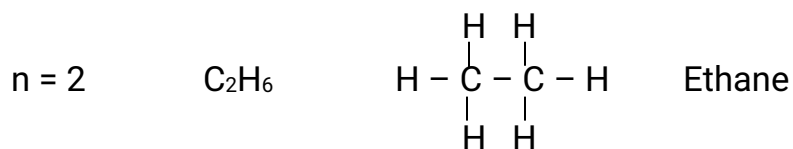
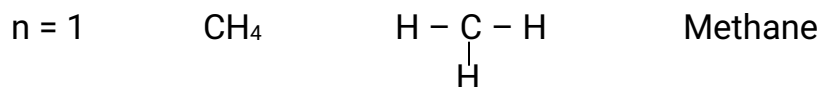
These are saturated hydrocarbons containing carbon and hydrogen elements only joined by a single covalent bond.

The open chain hydrocarbons (alkanes) have got a general formula of  $\text{C}_n\text{H}_{2n+2}$  where  $n$  = a whole number.

Functional group of all alkanes is a carbon - carbon "C- C" single bond where all the four valency electrons of carbon are involved in the bonding.

**Examples:**





## NOMENCLATURE

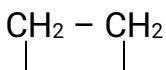
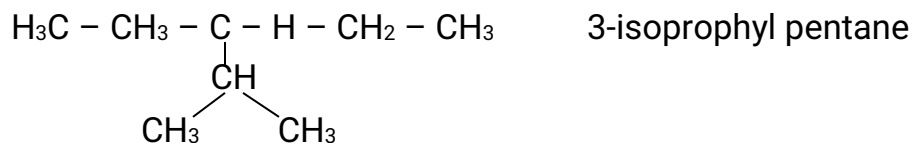
Alkanes are named as saturated hydrocarbons by adding the suffix-ane on the stem name of the parent hydrocarbon.

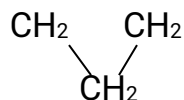
The parent hydrocarbon indicates the number of carbon atoms.

<u>Number of carbon atoms</u>	<u>Parent name</u>	<u>Alkane</u>
n = 1	Meth	Methane
n = 2	Eth	Ethane
n = 3	Prop	Propane
n = 4	But	Butane
n = 5	Pent	Pentane
n = 6	Hex	Hexane

The branches on the parent chain must be indicated together with their positions.

For cyclic alkanes, the word cyclo has got to be written before the parent name.  
H<sub>3</sub>C CH<sub>2</sub> CH<sub>2</sub> CH<sub>3</sub>      n – Butane.

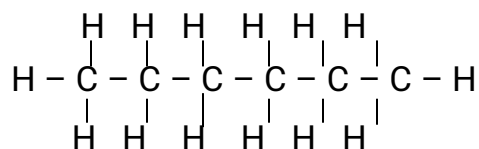




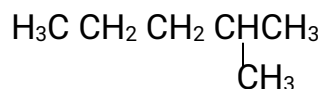
Cyclo pentane

## ISOMERISM IN ALKANES

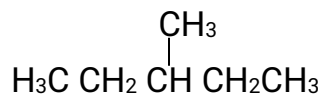
Alkanes exhibit structural isomerism specifically chain isomerism where the isomers differ in the chain.



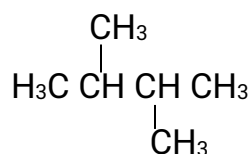
n - hexane



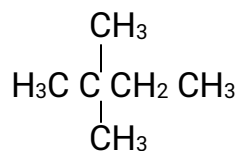
2-methyl pentane



3-methyl pentane



2, 3-dimethyl butane



2, 2-dimethyl butane

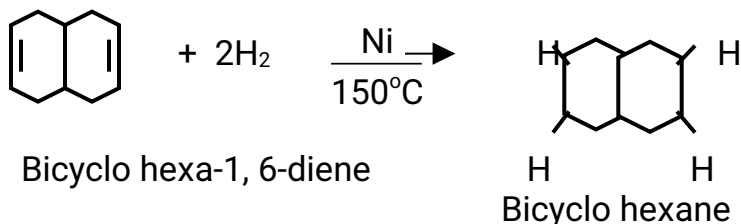
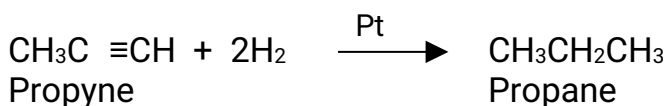
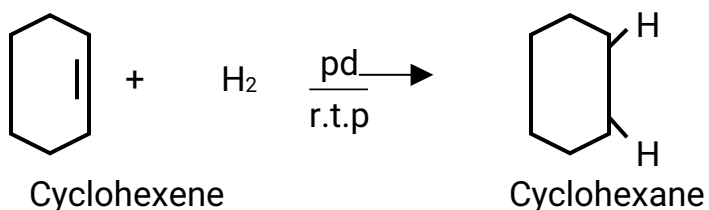
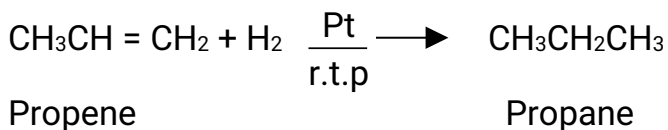
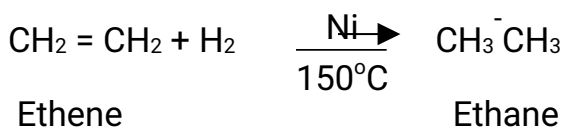
## PREPARATIONS OF ALKANES

### 1. FROM UNSATURATED HYDROCARBONS

Alkanes are prepared by reduction of unsaturated hydrocarbons e.g. alkenes and alkynes.

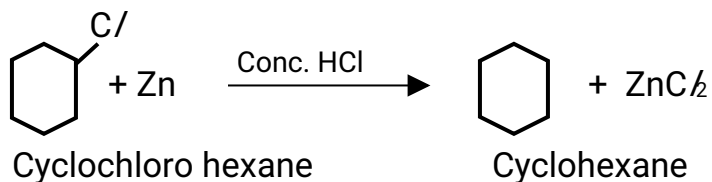
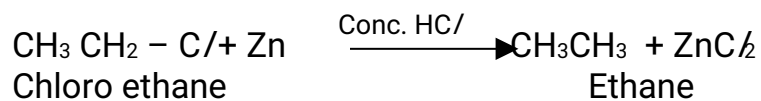
When an alkene is reacted with  $\text{H}_2$ , in the presence of a suitable catalyst like nickel at  $150^\circ\text{C}$ , platinum (Pt) and paradium (Pd) at room temperature, an alkane is formed.

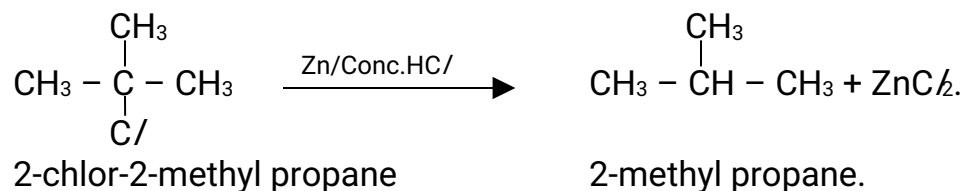
Alkynes also undergo the same reaction except that excess  $H_2$  is required.



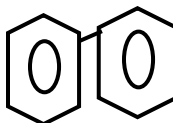
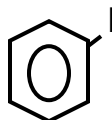
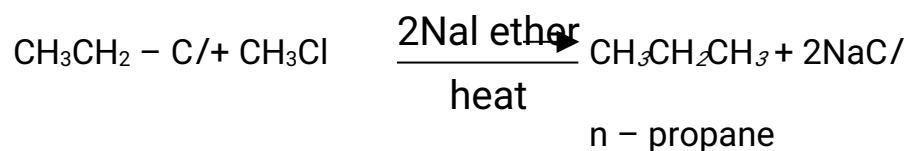
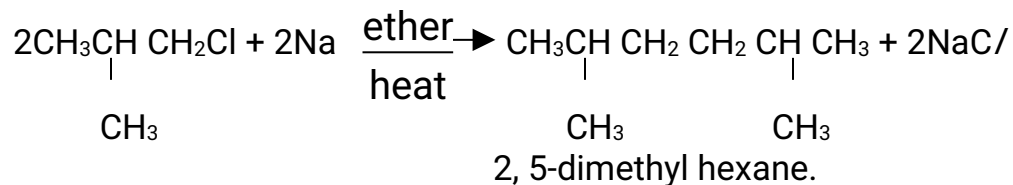
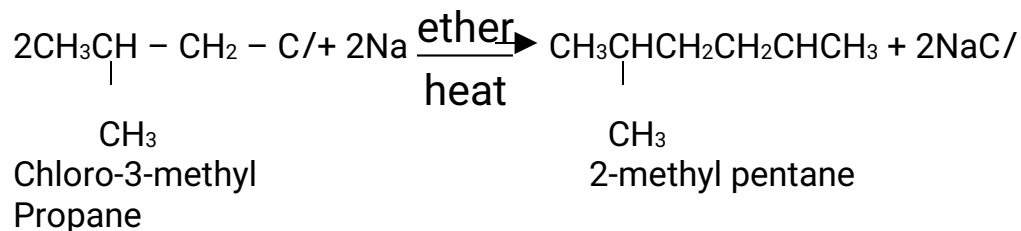
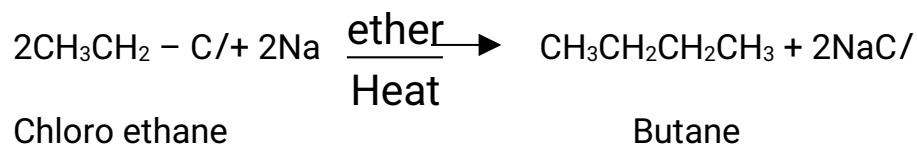
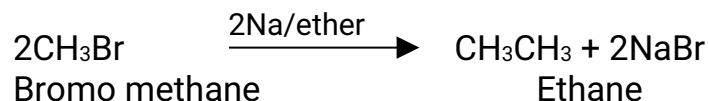
## 2. FROM HALOGENO ALKANES (ALKYL HALIDES) (R-X)

- (i) When alkyl halides are reduced with zinc in the presence of conc. HC/acid, alkanes are formed.





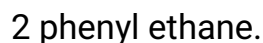
- (ii) When alkyl halides are heated with sodium metal in ether (diethyl ether) alkanes are formed but with increase in carbon chain (doubling the carbon chain). This reaction is known as **Wurtz** reaction and its usually important in organic synthesis when doubling the carbon chain.





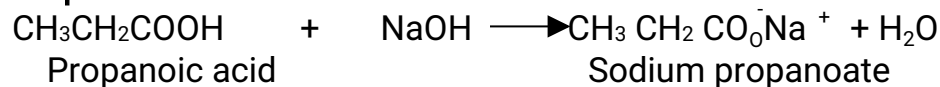


Alky halides can be reduced using zinc-copper couple in presence of ethanol to give alkane.

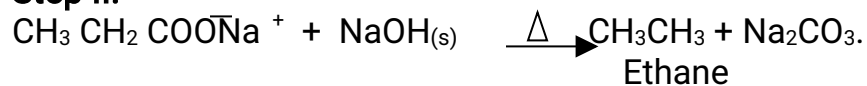


Alkanes can be prepared from carboxylic acids or salts of carboxylic acids by reacting with fused sodium hydroxide or soda lime.

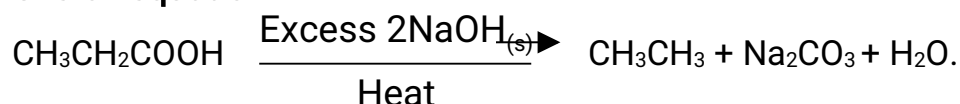
### Step 1:



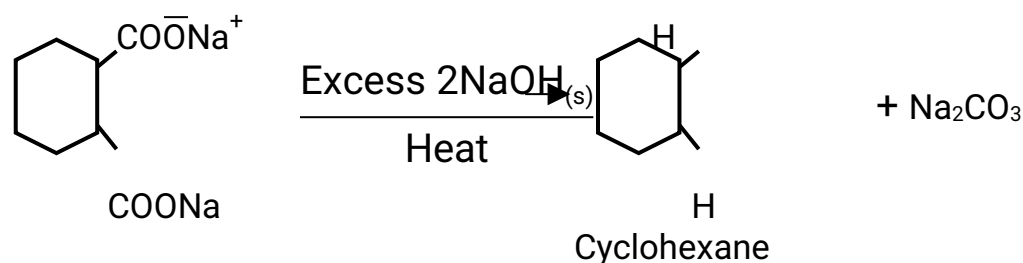
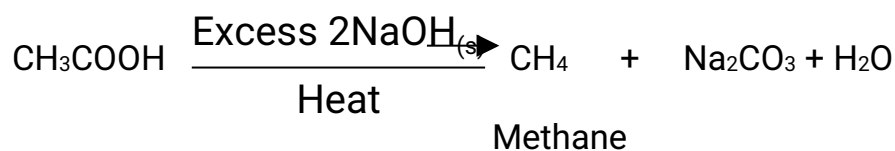
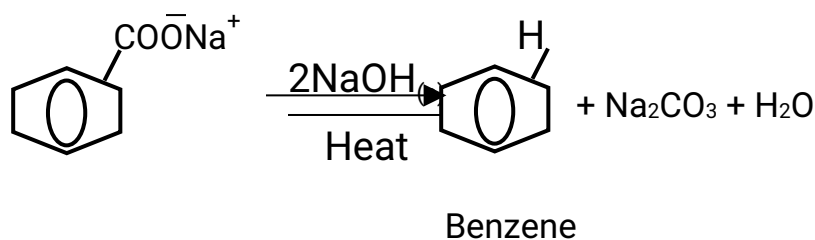
### Step II:



**Overall equation:**



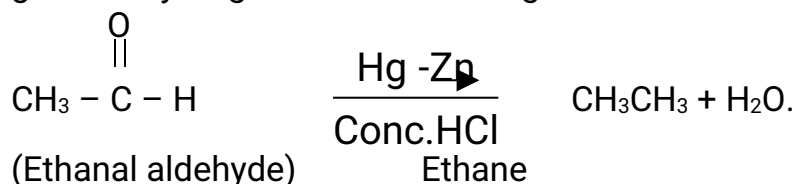
**Note:** This reaction is very important in organic synthesis by reduction of the carbon chain length by one carbon atom.

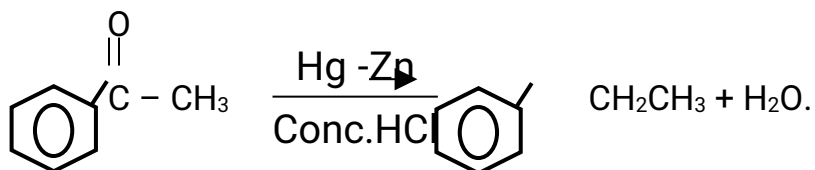
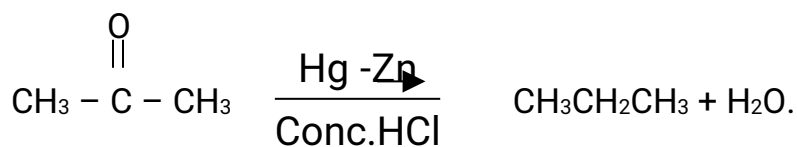


#### 4. FROM CARBONYL COMPOUNDS

Reduction of carbonyl compounds using amalgamated zinc produces alkanes in the presence of conc. HCl/

Carbonyl compounds are aldehydes and ketones containing an oxygen atom in the carbonyl carbon, (C = O) which is removed by zinc and then given to hydrogen to form H<sub>2</sub>O. e.g.

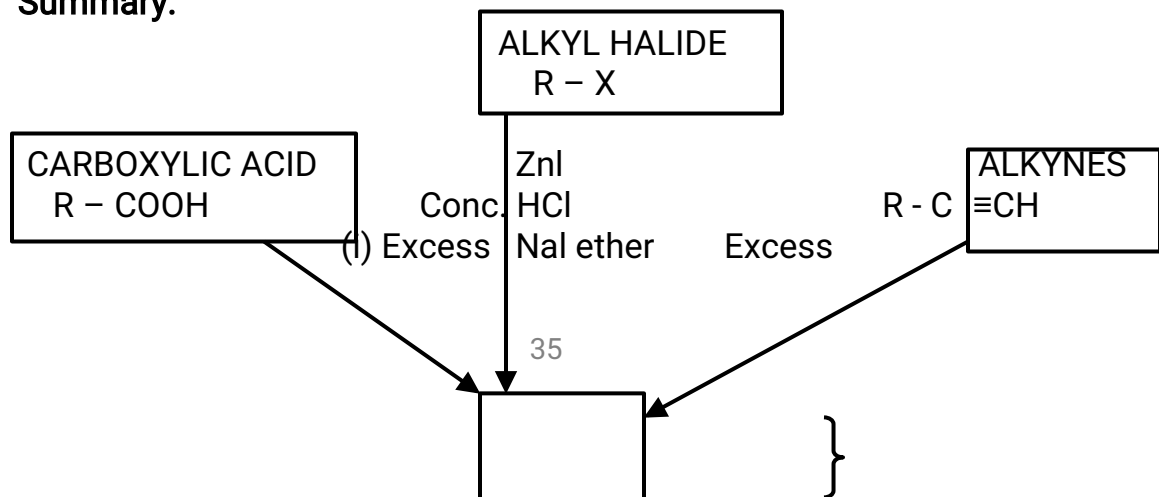


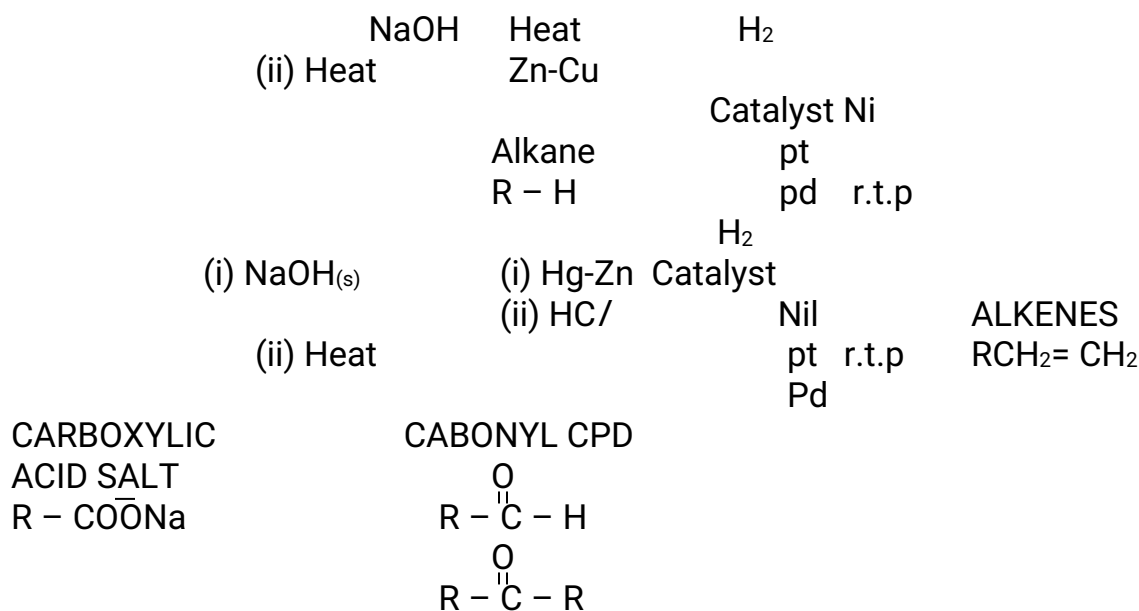


## 5. CRACKING OF HIGHER ALKANES

When higher alkanes are heated under high temperatures, they break down into shorter alkanes and sometimes alkenes are also formed.

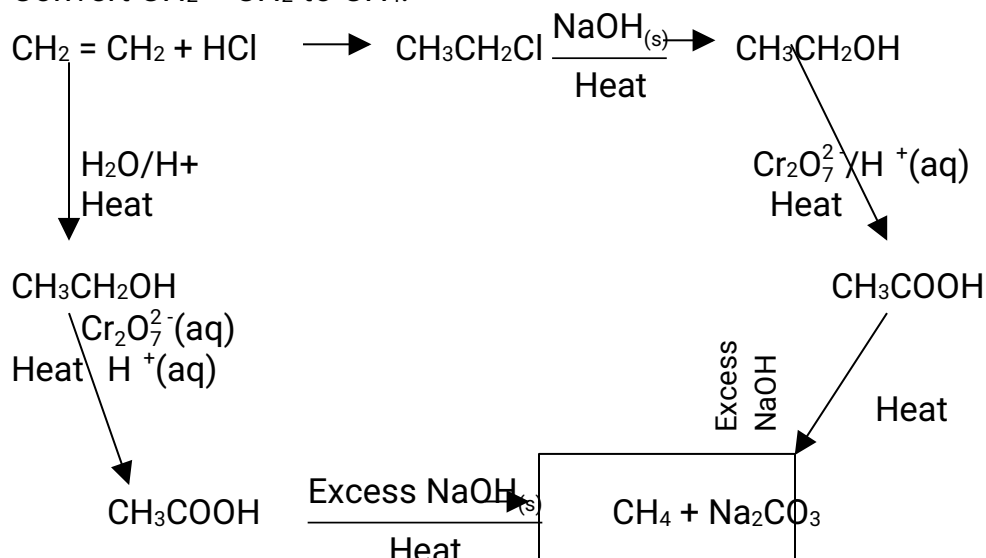
**Summary:**



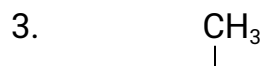
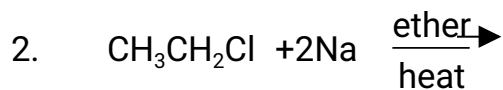
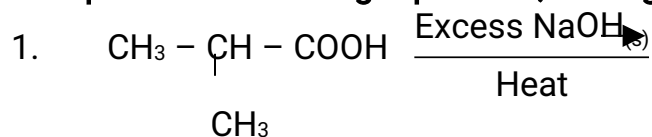


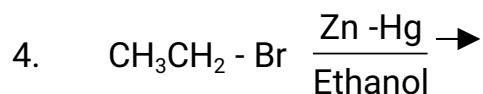
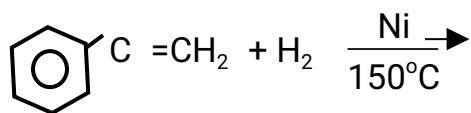
**Question:**

Convert CH<sub>2</sub> = CH<sub>2</sub> to CH<sub>4</sub>.

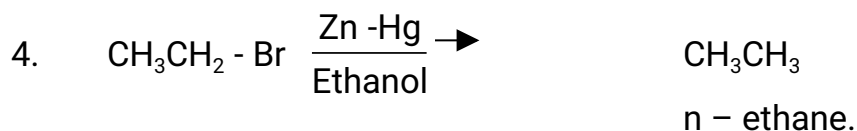
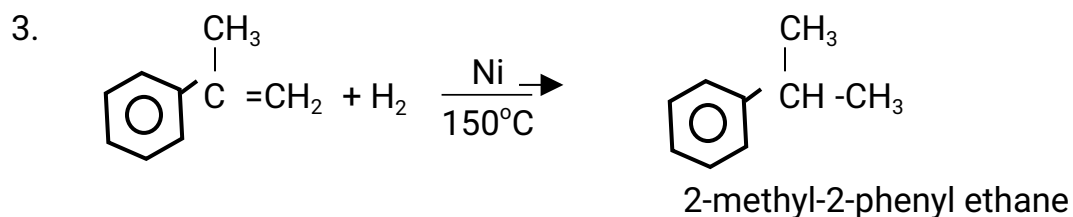
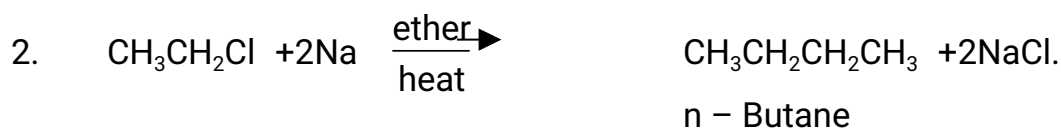
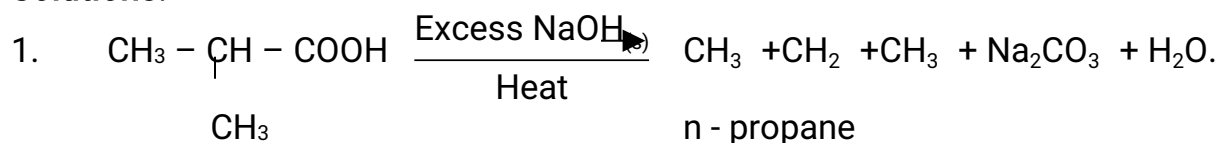


Complete the following equations, writing the IUPAC name of the main product.





### Solutions:



## PROPERTIES OF ALKANES

### PHYSICAL PROPERTIES

#### (a) Nature:

Lower alkanes exist as gases ( $\text{C}_1 - \text{C}_3$ ) exist as gases at room temperature.

This is because they have got low boiling points.

The medium ( $C_4$  -  $C_9$ ) exist as liquids. Higher alkanes are greater than  $C_{10}$  exist as solids at room temperature.

(b) **Boiling and melting points:**

Generally the boiling points and melting points of alkanes are lower than corresponding compounds like alcohols, carboxylic acids.

**Explanation:**

Alkane molecules are held by weak Vander Waal's forces of attraction while corresponding alcohols and carboxylic acids are held by relatively strong hydrogen bonding.

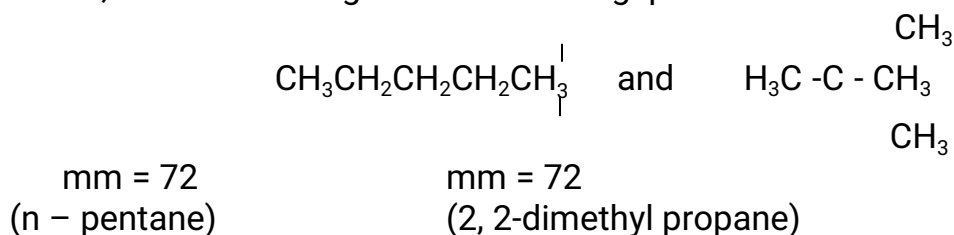
Within alkanes, the boiling point generally increases with the increase in molecular weight.

Ethane (30) boils at lower temperature than Butane (56).

**Explanation:**

Increase in the molecular weight increases the size of the molecules which correspondingly increases the Vander Waal's forces of attraction. That is why butane (56) boils at a higher temperature than ethane (30).

However, when branching is considered e.g. pentane.



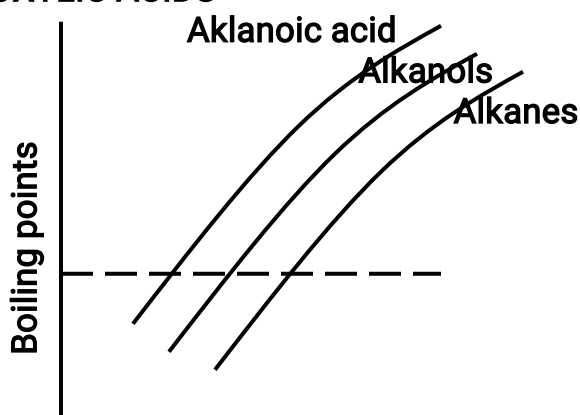
2, 2-dimethyl propane boils at a lower temperature than n-pentane yet both are of the same molecular weight.

**Explanation:**

2, 2-dimethyl propane is highly branched giving it a spherical shape which reduces its size and weaker Vander Waal's forces of attraction will exist between the molecules. While normal pentane is a straight chain molecule which gives the molecule an extended structure resulting into relatively stronger Vander Waal's forces.

**A SKETCH SHOWING BOILING POINTS OF ALKANES, ALCOHOL AND**

## CARBOXYLIC ACIDS



### Assignment:

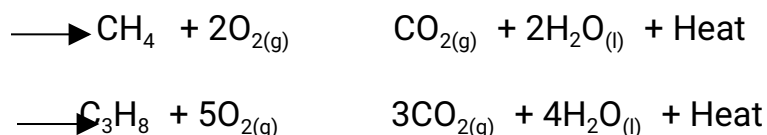
The boiling points of alkanes with even and odd number of carbon atoms are not the same.

- Identify how they differ
- Explain the above difference.

## COMBUSTION OF ALKANES

Alkanes usually burn in air to give  $\text{CO}_2$  and  $\text{H}_2\text{O}$  only and a lot of heat is produced. This explains why alkanes are used as fuels in domestic and industrial use. E.g. methane is an example of a bio gas that explodes readily to give heat.

Gasoline usually contains propane or butane used in domestic and industrial heating and also to run auto-mobile.

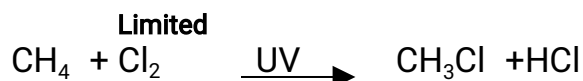
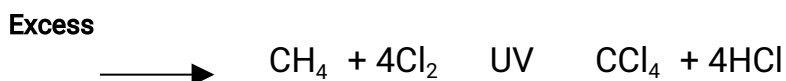


## REACTIONS WITH HALOGENS IN PRESENCE OF (UV) LIGHT OR HEAT

Alkanes are generally unreactive but in the presence of UV radiations (like sunlight), alkanes react with chlorine to form chloro alkane. This is an example of a substitution reaction where  $\text{H}_2$  atoms are substituted by chlorine atoms following a free radical mechanism.

A free radical mechanism is a type of reaction where free radicals are involved.

A free radical is a chemical species with unpaired electrons and it is always very unstable but reactive. E.g.



The above reaction is a chain reaction where the products formed become the reactants of the next step.

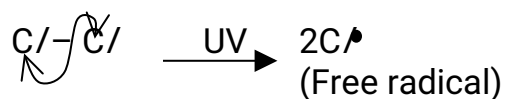
Therefore, the reaction can then be divided into three;

- (i) Initiation step
- (ii) Propagation step
- (iii) Termination step

### INITIATION STEP

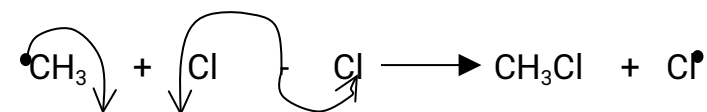
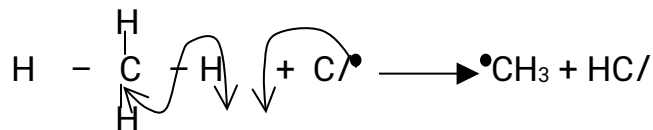


Formation of free radicals using UV. i.e.

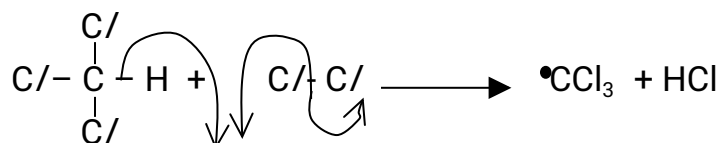
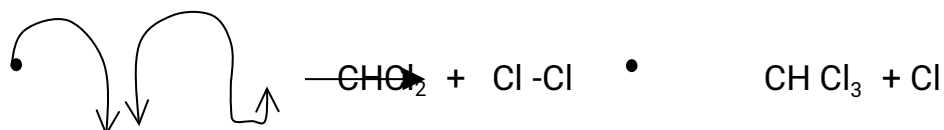
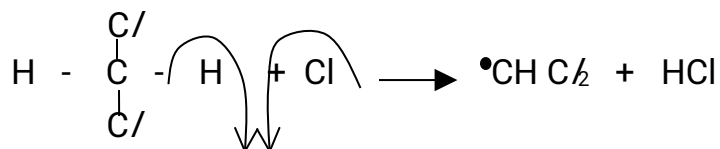
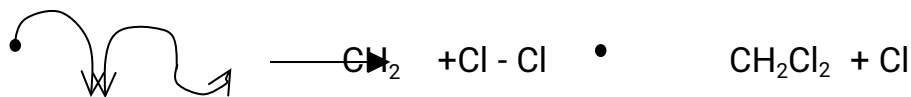
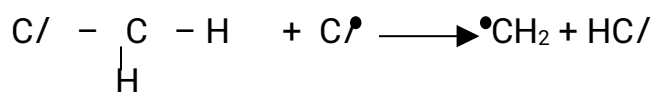


### PROPAGATION STEP

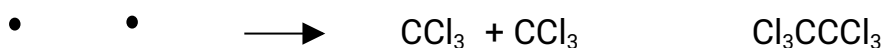
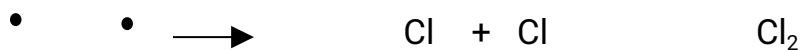
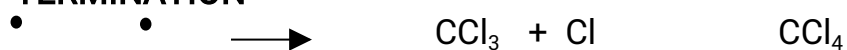
Here the free radicals generated react with neutral molecules to form new radicals.







### TERMINATION

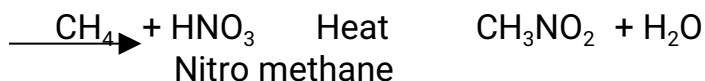


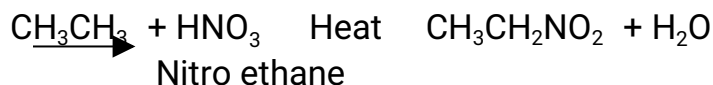
**Note:** The above reaction yields a mixture of products and therefore it might be very good in the synthesis of organic compounds. E.g. the above produces a mixture of compounds like chloro methane, Di chloro methane, tri chloro methane, tetra chloro methane.

However, the extent of reaction is determined by the amount of chlorine present.

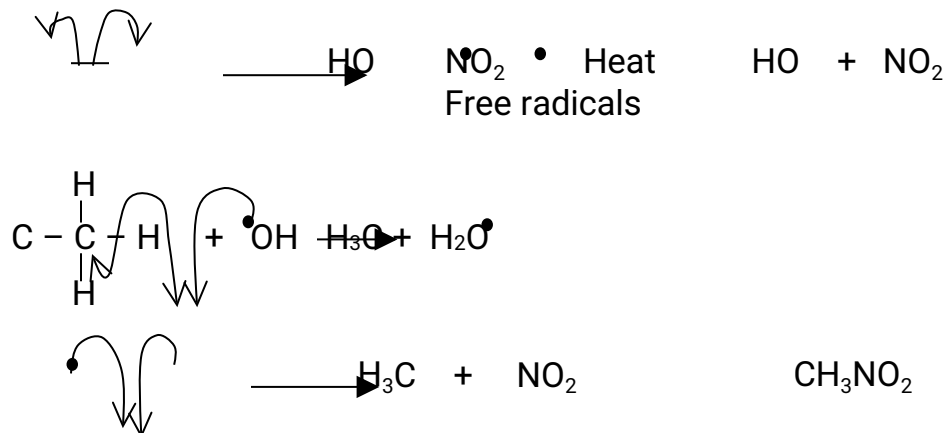
### NITRATION

Alkanes react with  $\text{HNO}_3$  acid vapour (heated  $\text{HNO}_3$  acid) to form Nitro alkanes, unlike chlorination, in this reaction only one hydrogen atom is substituted by the nitro group.





## MECHANISM



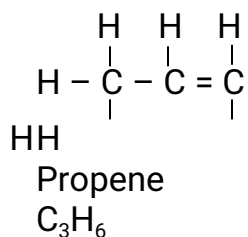
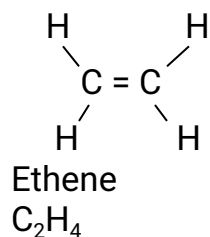
## Petroleum industry

- How petrol is formed (distribution of petroleum)
- Cracking of alkanes (thermal cracking)
- Products of a petro – chemical industry.
- Uses of alkanes – fuels, solvents.

## ALKENES

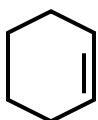
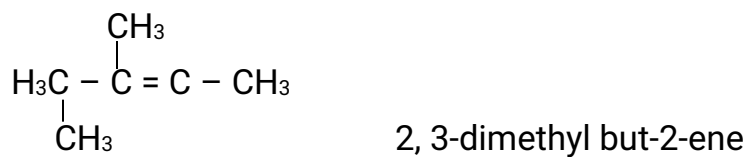
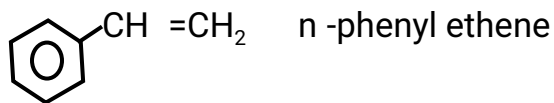
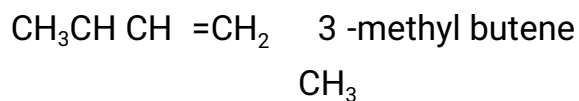
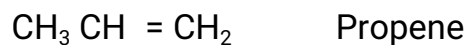
Alkene are unsaturated hydro carbons consisting of a carbon-carbon double bond (- C = C -) as the functional group.

All members in alkene homologous series conform to the general formula  $\text{C}_n\text{H}_{2n}$ .

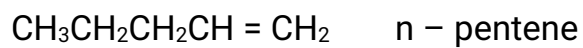
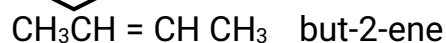


## NOMENCLATURE

Alkenes are named as derivations of alkanes by removing the suffix “ane” and replacing it with “ene”. E.g.



Cyclo Hexene

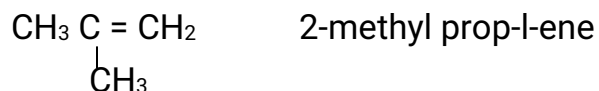


## ISOMERISM

Alkenes exhibit four types of isomerism i.e.

### (i) Chain isomerism

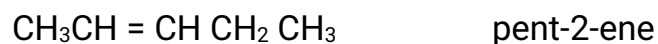
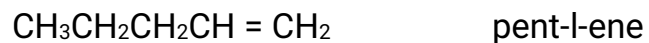
Examples:  $\text{C}_4\text{H}_8$



### (ii) Position isomerism

These differ by the position of the double bond.

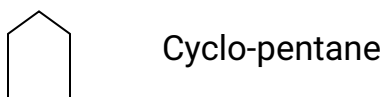
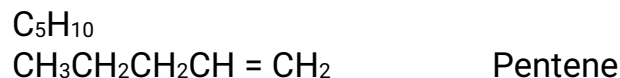
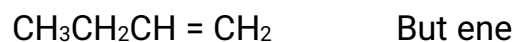
E.g.  $\text{C}_5\text{H}_{10}$ .



### (iii) Ring isomerism

Isomers differ by the shape of the carbon skeleton. Alkene and cycloalkanes are isomeric.

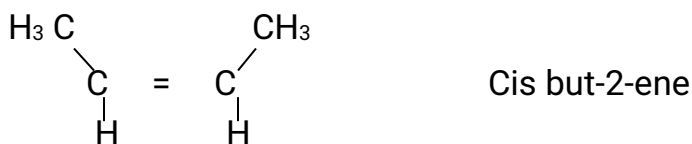
E.g.  $C_4H_8$



(iv) **Geometrical isomerism**

Compounds differ by the different ways in which the atoms are arranged by the double bond. (restricted rotation of a double bond).

Example: But-2-ene.



**Exercise:**

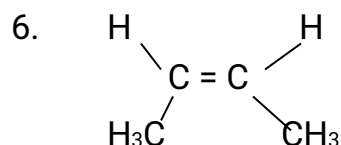
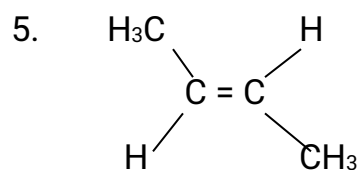
Write all the isomers of  $C_4H_8$

1.  $CH_3CH_2CH = CH_2$  n - but-1-ene

2.  $CH_3 - \underset{\begin{array}{c} | \\ CH_3 \end{array}}{C} = CH_2$  2-methyl prop-1-ene

3.  $CH_3CH = CHCH_3$  but-2-ene

4.  $CH_2CH_2CH_2CH_2$



1 and 2 are chain isomers  
 3 and 4 are position isomers  
 5 is a ring isomers  
 6 is geometrical isomers

## METHODS OF PREPARATION OF ALKENES

Alkenes are prepared from a number of ways:

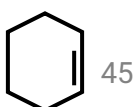
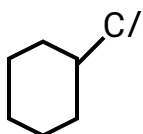
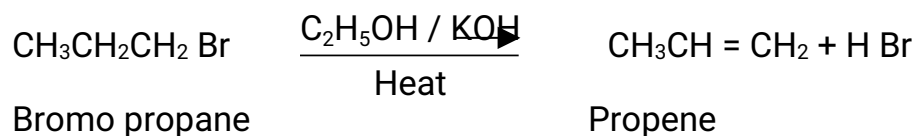
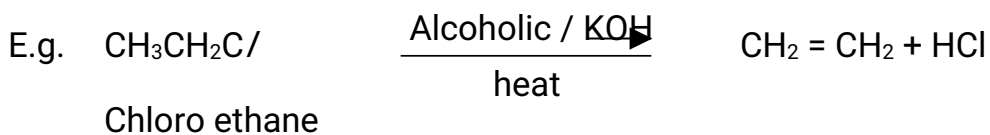
### (i) From Alkyl halides (Halogeno alkanes)

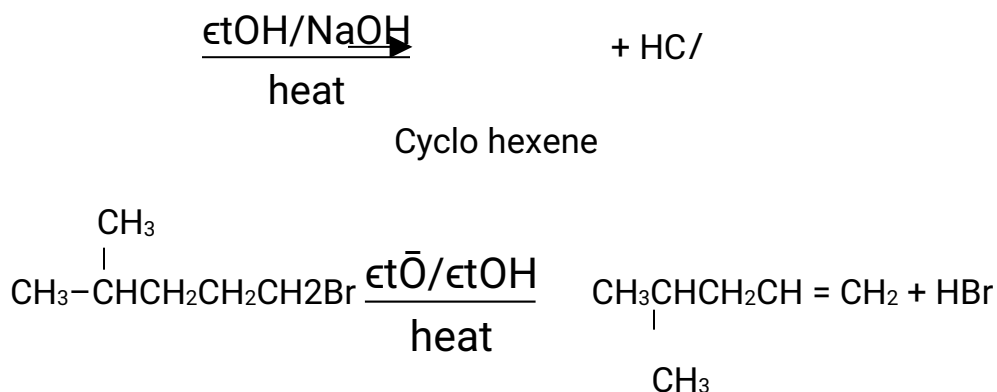
When halogeno alkanes are heated with an alcoholic alkaline solution like NaOH or KOH, this reaction is an elimination reaction when both H<sub>2</sub> and the halogen are removed from an alkyl halide to form an alkene as a major product and hydrogen halide as a minor product.

#### Conditions:

Heat

Alcoholic alkali (NaOH or KOH)





(ii) **Dehydration of alcohols**

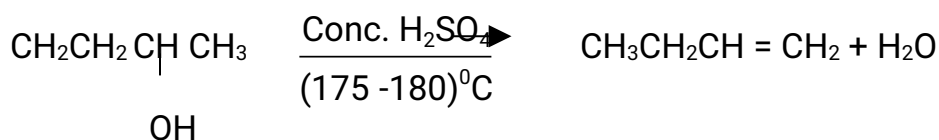
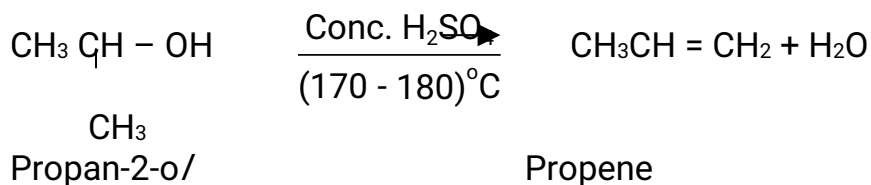
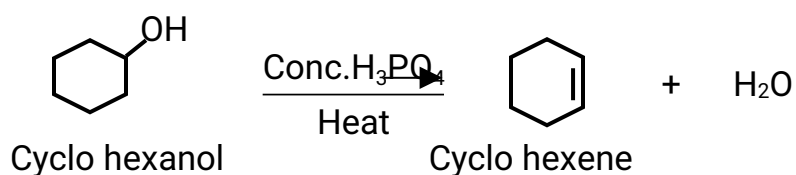
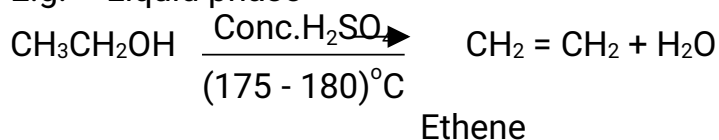
When alcohol is heated with conc.  $\text{H}_2\text{SO}_4$  or  $\text{H}_3\text{PO}_4$  acid in either liquid or vapour phase, an alkene is formed.

Conditions:

Heat at appropriate temperature depending on the type of alcohol ( $170 - 180^\circ\text{C}$ ).

Concentrated acid

E.g. Liquid phase

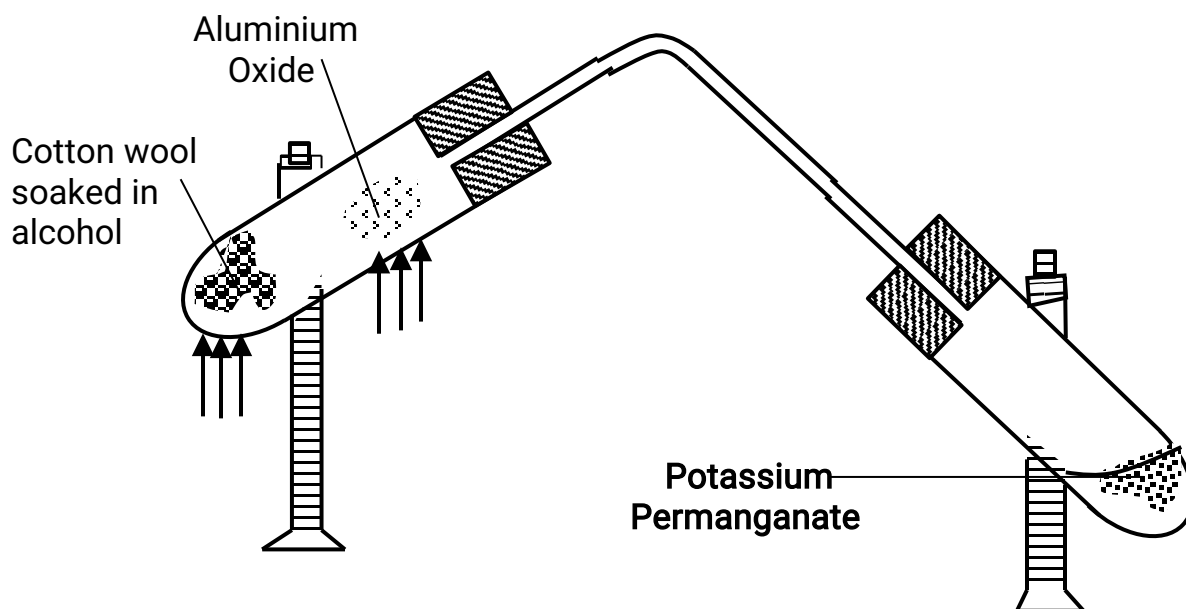


Butanol

butene

### VAPOUR PHASE:

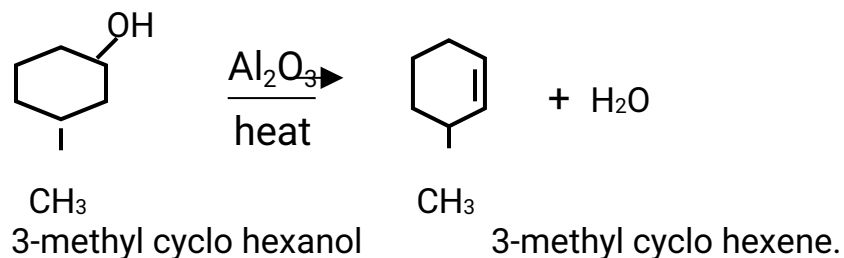
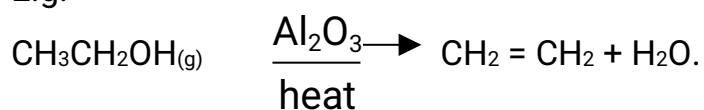
When the vapour of an alcohol generated by heating, the alcohol is passed over heated  $\text{Al}_2\text{O}_3$  as a catalyst, an alkene is formed.



### Observation:

The purple solution of potassium permanganate is decolourised.

E.g.

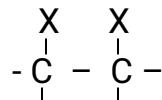


The above reaction is also known as elimination reaction because an alcohol loses a water molecule and it forms an unsaturated compound.

### (iii) Dehydro Halogenations of Vicinal Dihalides

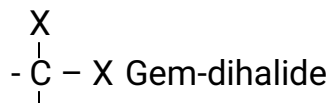
This is a reaction where there is removal of both hydrogen and halogen atoms from one molecule to form unsaturated compound.

A Vicinal dihalide is a compound with two halogen atoms positioned at adjacent carbon atoms. The structure is shown below.

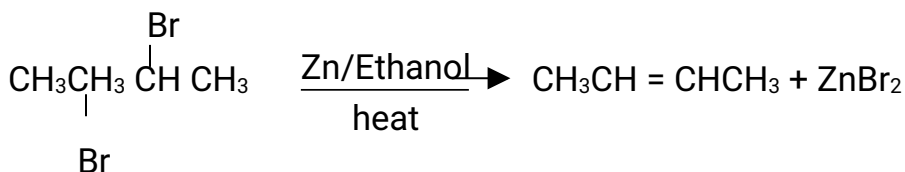
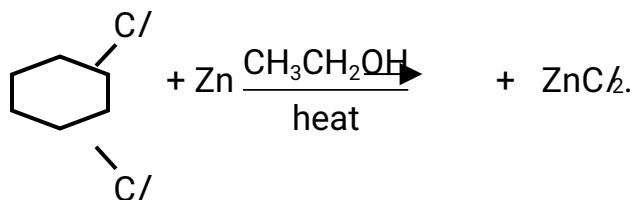
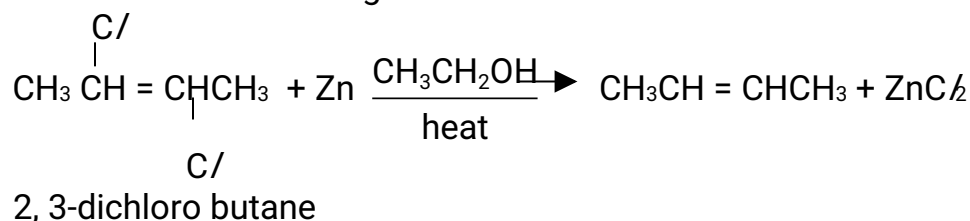


Vicinal is different from a Gem-dihalide.

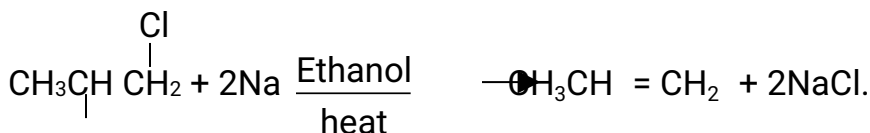
A Gem-dihalide has got two halogen atoms positioned within the same carbon atom. E.g.



- (a) When a vicinal dihalide is heated with zinc in presence of an alcohol, an alkene is formed. E.g.



**Note:** Sodium metal can also be used to give the same products that atoms of sodium are required. E.g.

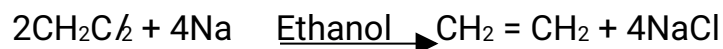
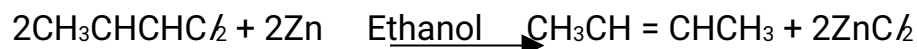




Cl

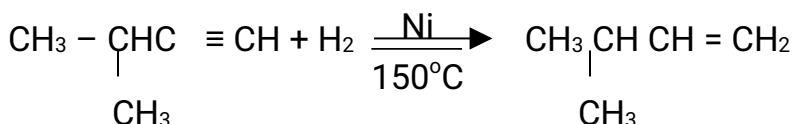
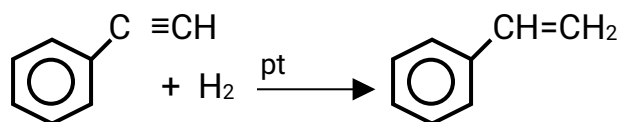
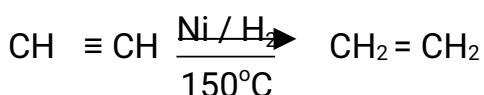
(b) **Using gem dihalides**

Higher alkenes are produced but with an increased number of carbon atoms and this reaction is useful in increasing the carbon length during organic synthesis.



(iv) **Partial reduction of alkynes**

When Alkynes are reacted with hydrogen in equi-molar ratio in presence of a heated catalyst like nickel at 150°C or platinum and palladium at room temperature, an alkene is formed.



**CHEMICAL REACTIONS OF ALKENES**

Alkenes are generally very reactive compared to alkanes. This is so because alkenes have a double bond consisting of  $\pi$  electrons which can be easily donated to reactants (electrophiles).

A double bond in alkenes is also shorter and stronger; that means it releases high energy compared to a single bond.

Bond	Energy KJ mol <sup>-1</sup>	Length (nm)
C – C	346	0.154
C = C	598	0.134

The above make alkenes reactive to electrophiles.

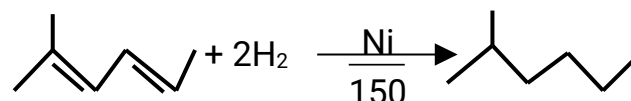
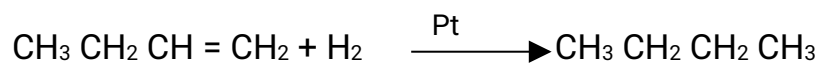
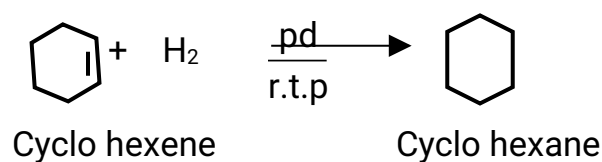
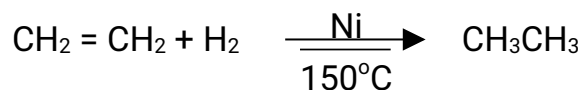
**TYPES OF REACTIONS**

(a) **Addition reactions**

Alkenes undergo addition reactions where a double bond is converted to a single bond.

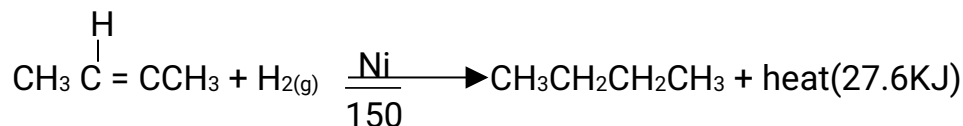
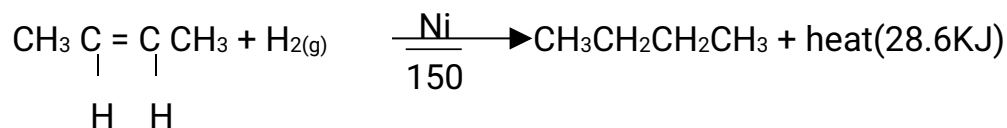
(i) Hydrogenation (Reduction)

Alkenes react with  $H_2$  in presence of heated catalyst of Nickel at  $150^\circ C$  or platinum and palladium of at room temperature to form alkanes. In this reaction all the hydrogen atoms are added across a double bond.

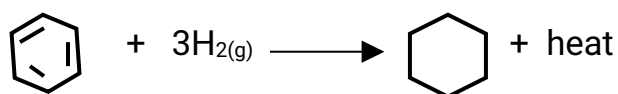


This reaction forms the basis of synthesizing saturated compounds from unsaturated e.g. formation of margarine when unsaturated fats are heated with  $H_2$  in presence of a catalyst to form a fat (margarine).

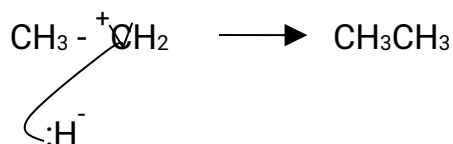
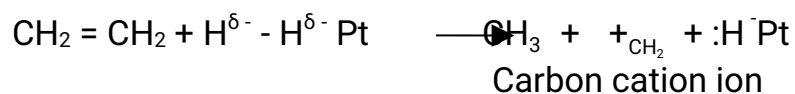
Determination of stability of compounds like Cis and trans-forms of alkenes, benzene and cyclo - 1, 3, 6 - triene.



The Cis form is unstable compared to the trans form.

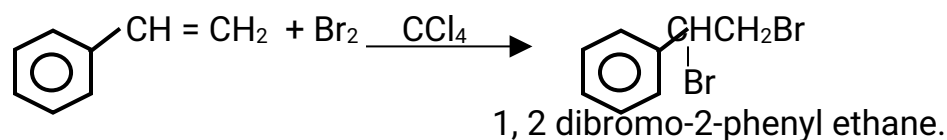
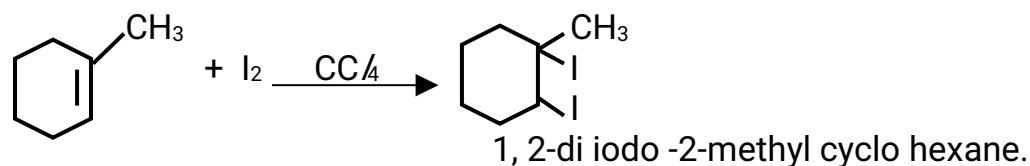
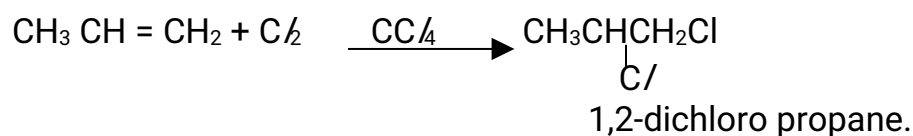


**Mechanism:**

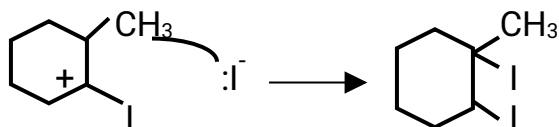
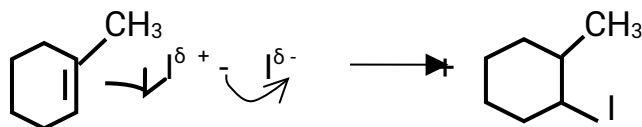
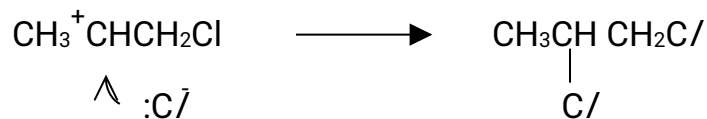
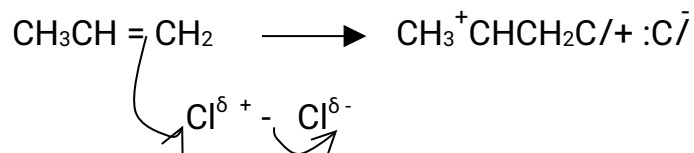


(ii) Halogenation

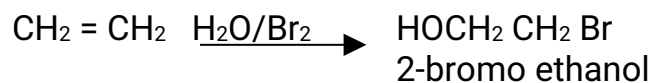
Alkenes react with halogens like  $\text{Cl}_2$ ,  $\text{Br}_2$ ,  $\text{I}_2$  in presence of an organic solvent like carbon tetra-chloride to form a dihalide compound.



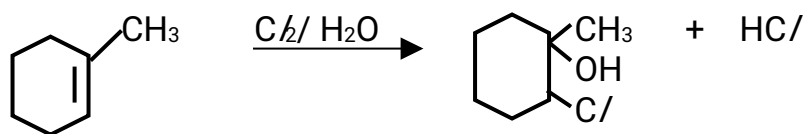
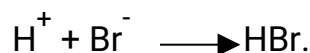
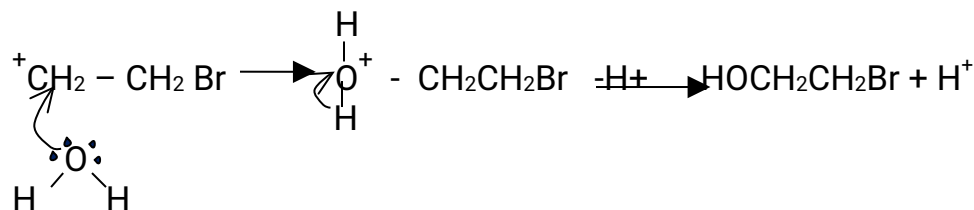
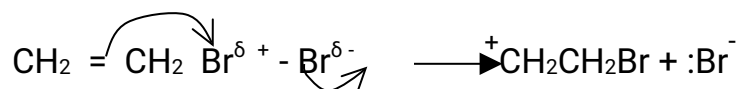
**Mechanism:**



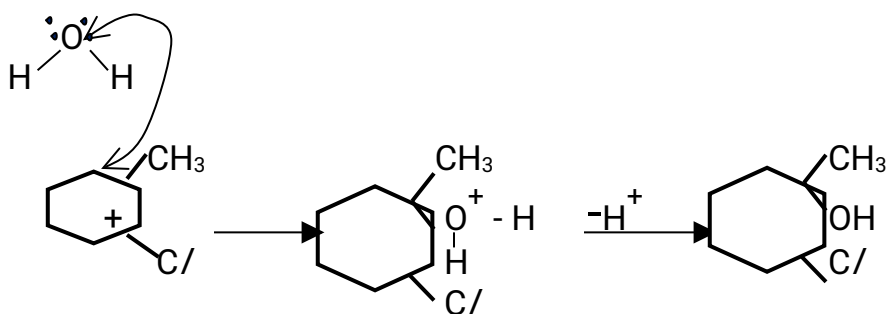
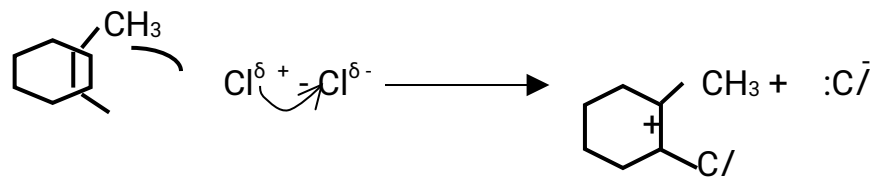
If the same reaction is done in the presence of  $\text{H}_2\text{O}$ , then the product formed is an hydroxyl halogeno alkane. This is because the OH generated from water is a better nucleophile than the halogen in addition to its presence in excess.

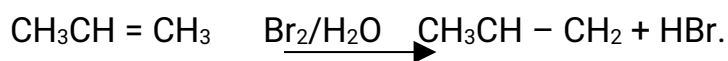
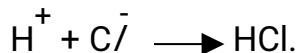


### Mechanism

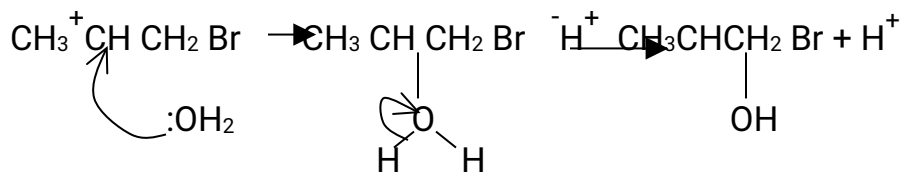
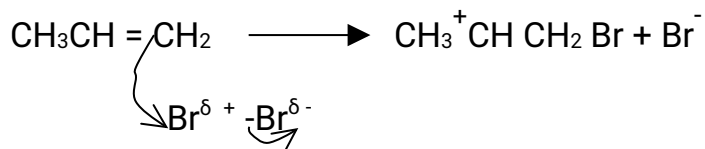


### Mechanism:

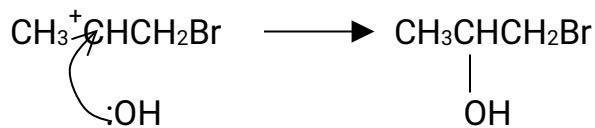
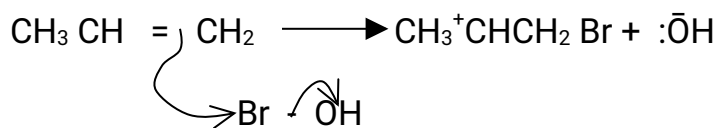




**Mechanism:**



**Alternatively:**



## Reactions with halogen acids (HX)

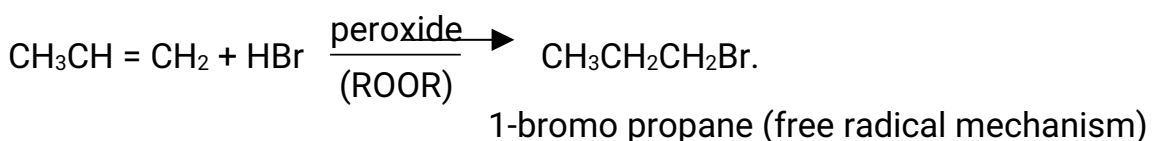
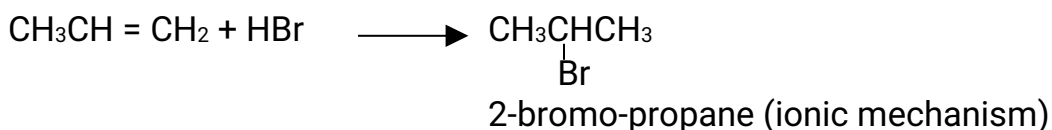
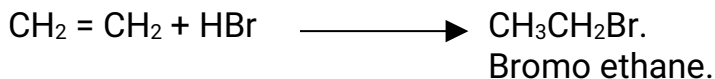
Alkenes react with halogen acids to form alkyl halides. This reaction is an addition reaction that can proceed into two ways according to the conditions of reaction.

(i) **Electrophillic addition**

This reaction occurs readily in the absence of a peroxide following an ionic mechanism.

(ii) **Free radical mechanism**

This occurs in the presence of a peroxide. In both types of reaction, the products formed are different.



**MARKWONIKOFF RULE**

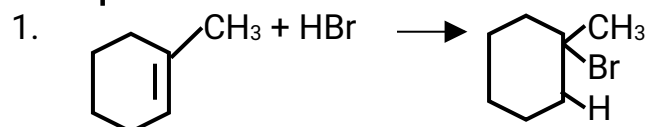
Markwonikoff rule predicts the products formed when halogen acid is reacted with unsymmetrical alkene.

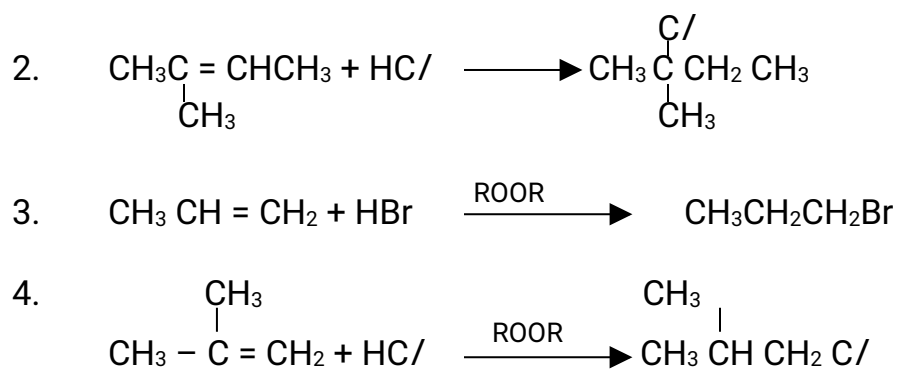
The rule states that addition of an halogen acid to unsymmetrical alkene gives a product where the hydrogen is added to the carbon atom with the lower number of hydrogens and the halogen atom is added to the hydrogen atom across a double bond to a carbon atom with few H<sub>2</sub> atoms.

This reaction is followed in the absence of a peroxide. E.g.  
CH<sub>3</sub>OOCH<sub>3</sub> (peroxide) or ROOR.

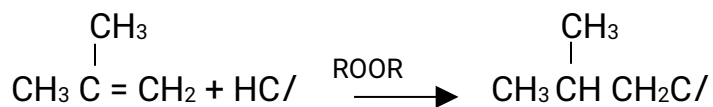
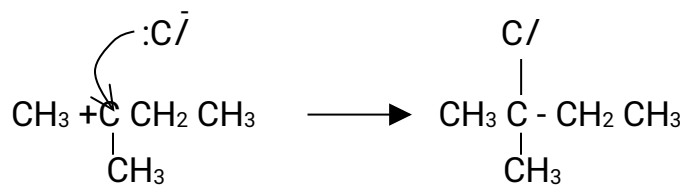
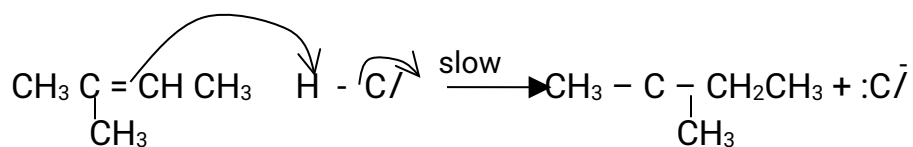
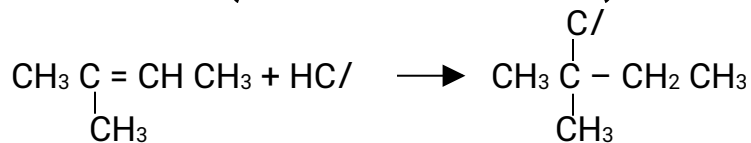
In the presence of a peroxide, anti Makwonikoff's rule applied where the hydrogen from the halogen acid is added to the carbon atom across the double bond with the highest hydrogen atom and the halogen goes to the carbon atom with the least hydrogens.

**Examples:**

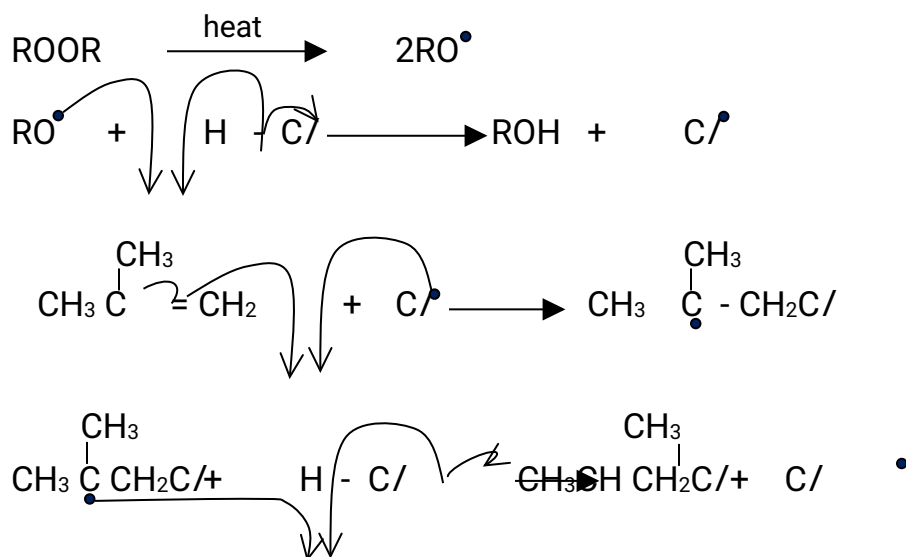




### MECHANISM (MARKWONIKOFF RULE)



### MECHANISM (ANTI MARKWONIKOFF)



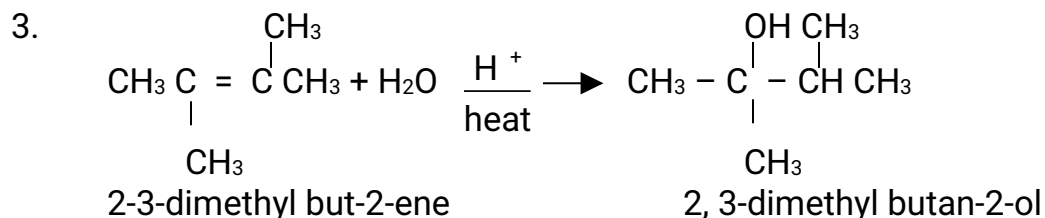
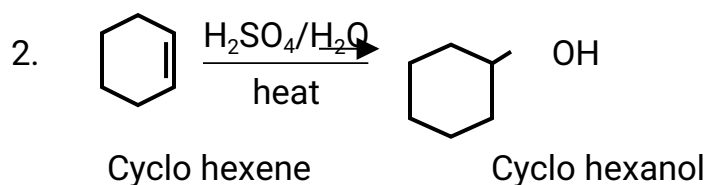
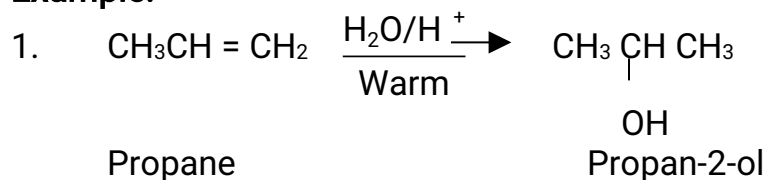
The reactions follow a more stable carbo cation ions in the order of  $3^\bullet > 2^\bullet > 1^\bullet$ .

## REACTIONS OF ALKENES WITH H<sub>2</sub>O IN PRESENCE OF AN ACID(H<sub>2</sub>O/H<sup>+</sup>(aq)

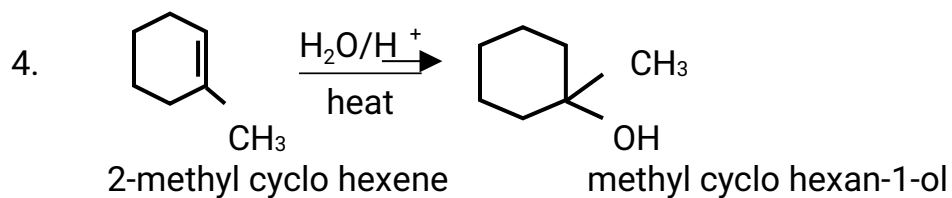
### MINERAL ACID WATER

When an alkene is heated with H<sub>2</sub>O in presence of H<sub>2</sub>SO<sub>4</sub> acid, an alcohol is formed. This reaction is known as hydration of alkenes and is used in the preparation of alcohols.

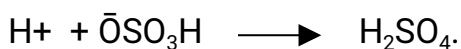
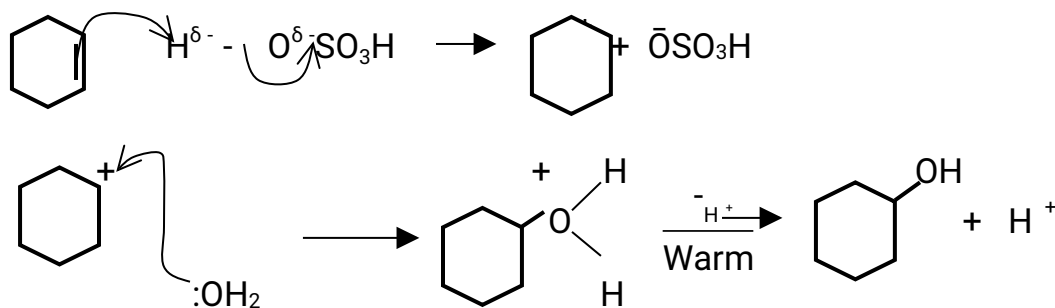
#### Example:



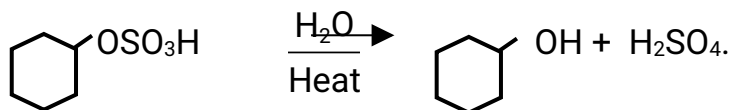




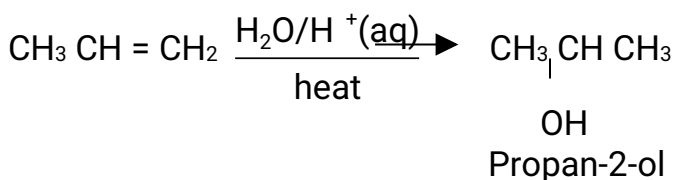
**Mechanism:**



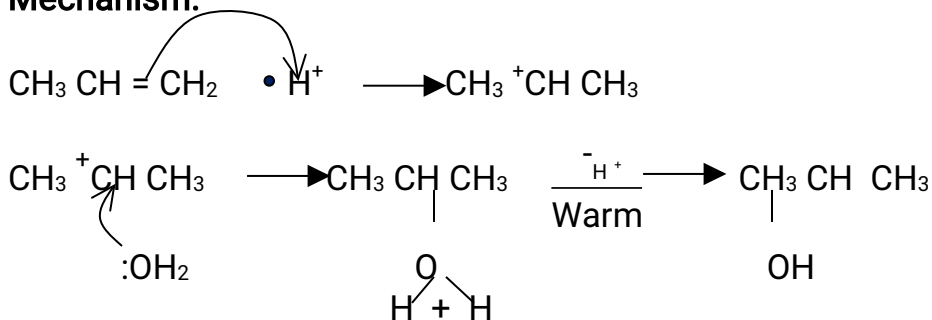
**NOTE:** If the reaction is done by first reacting the alkene with the sulphuric acid, the intermediate formed is (Alkyl hydrogen sulphate).



On diluting and warming, the product formed is the alcohol.

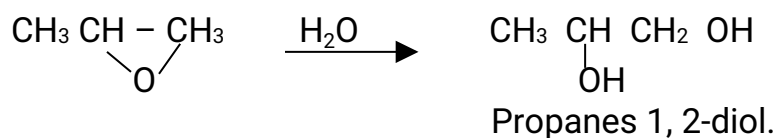
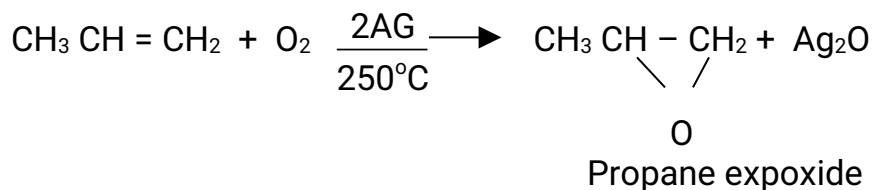


**Mechanism:**

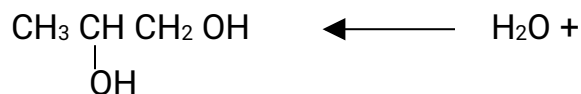
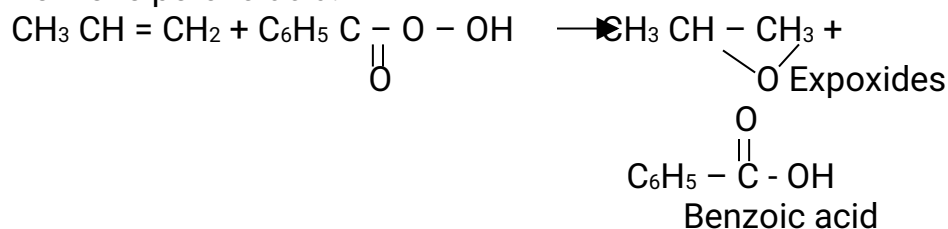


## REACTIONS WITH MOLECULAR O<sub>2</sub>.

Alkenes react with molecular O<sub>2</sub> in presence of a heated silver catalyst to form Expoxides. Hydrolysis of expoxides with water produced d-ols.



This reaction can also occur, when alkenes are reacted with peroxo acids. E.g. Benzene peroxo acid.

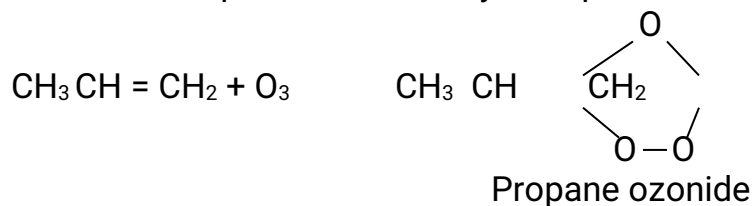


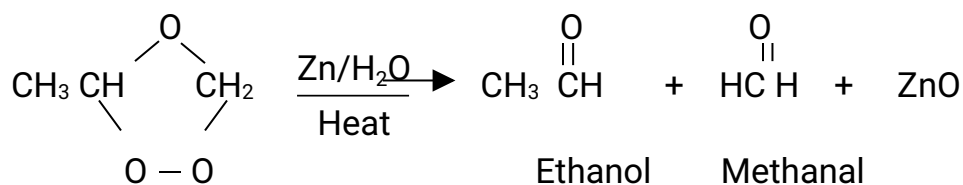
## REACTIONS WITH OZONE (O<sub>3</sub>)

Alkenes react with ozone to form ozonides hydrolysis of ozonides in the presence of;



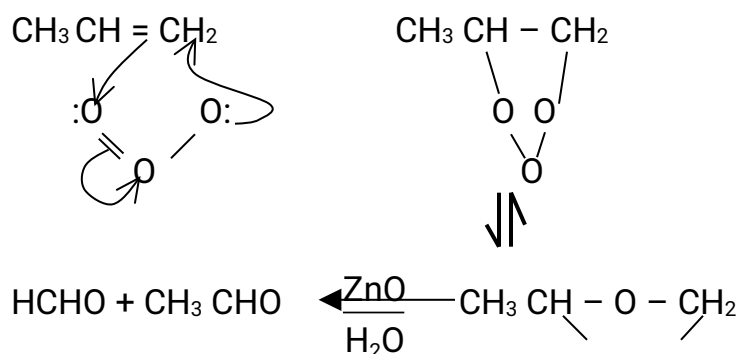
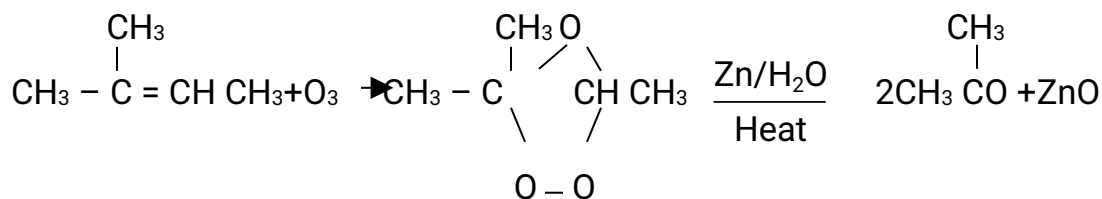
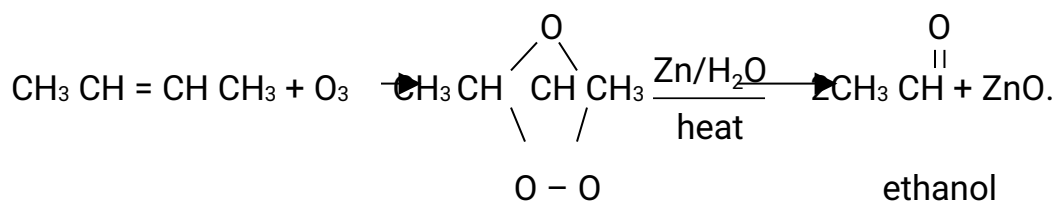
Zinc and H<sub>2</sub>O produces carbonyl compounds.

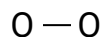




The above reaction is very important in two main ways.

- (i) In organic synthesis of converting an alkene to a carbonyl compounds of aldehyde and ketone.
- (ii) It is important in the analysis of the structure of the original alkene. If one product is formed after the hydrolysis of the ozonide, then the alkene used is asymmetrical where the double bond is at the centre. But if two products are formed, then the alkene is unsymmetrical.



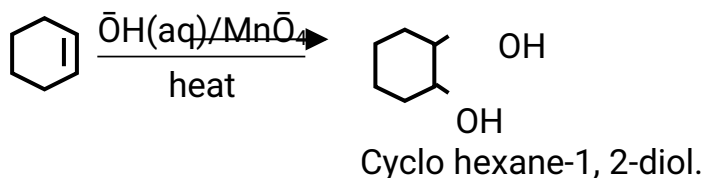
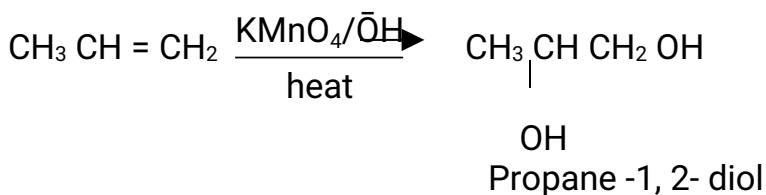
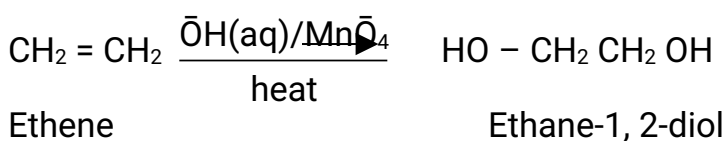


## OXIDATION REACTION OF ALKENES

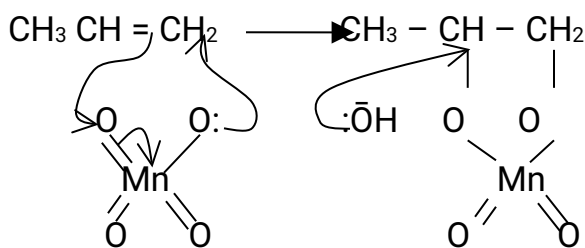
Alkenes undergo oxidation reaction using mild oxidizing agents like Alkaline potassium permanganate to form di-ols.

During this reaction, observation made is that the purple solution mixture, turns colourless (purple solution decolourised).

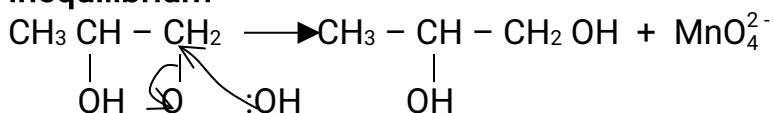
This test indicates the presence of a double bond which is converted to a saturated compound.

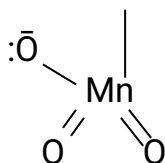


## MECHANISM:



## Inequilibrium





## POLYMERISATION

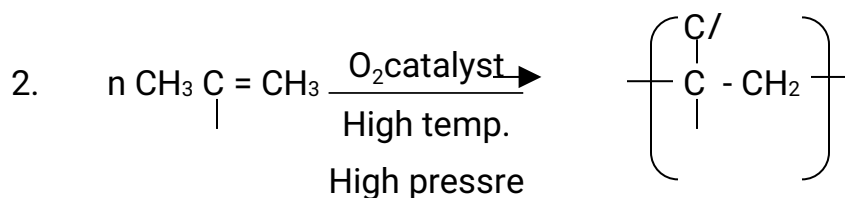
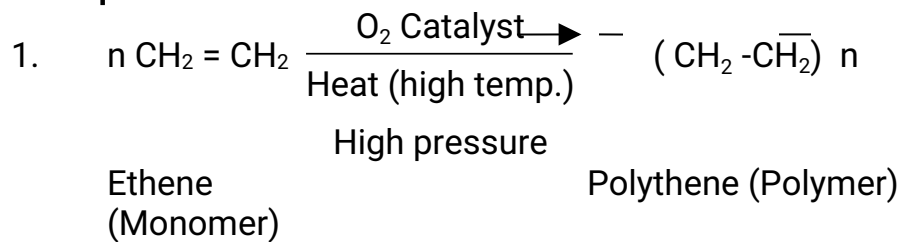
Alkenes undergo polymerisation known as addition polymerisation due to the presence of a double bond, enabling the monomer units to repeatedly combine to form high molecular mass products known as polymers.

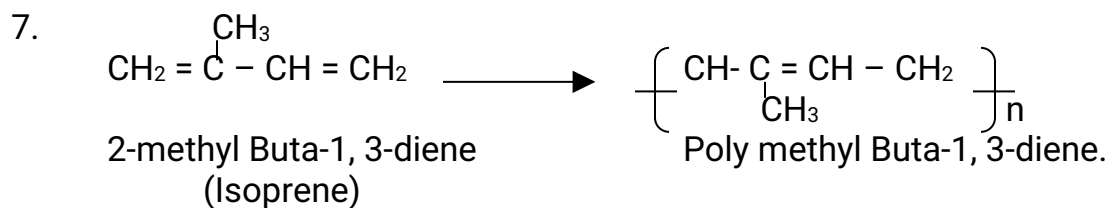
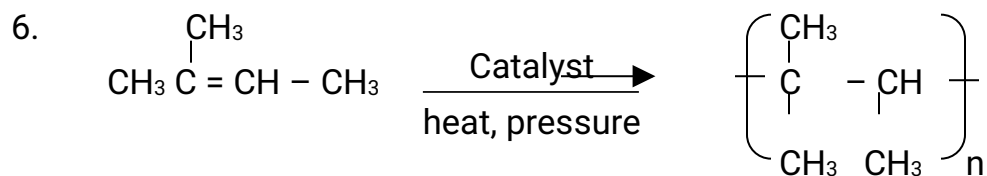
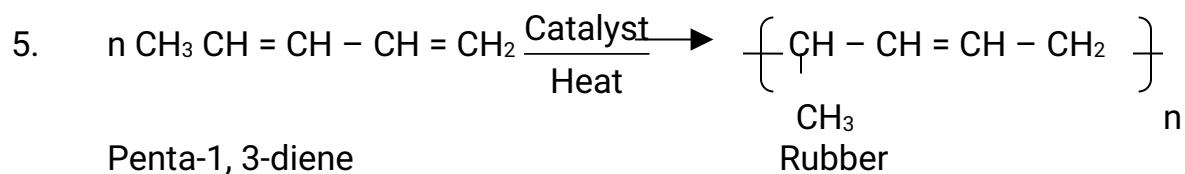
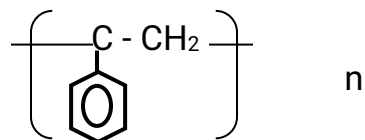
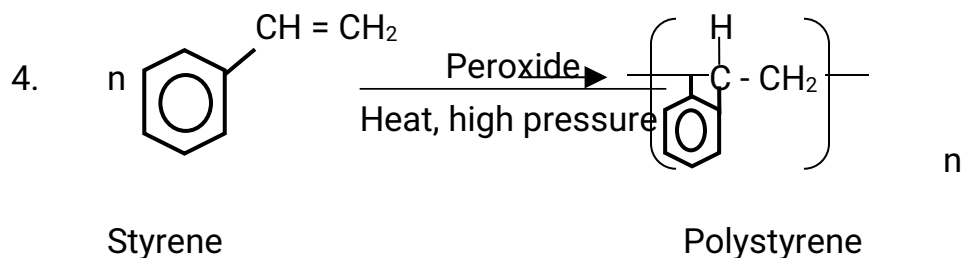
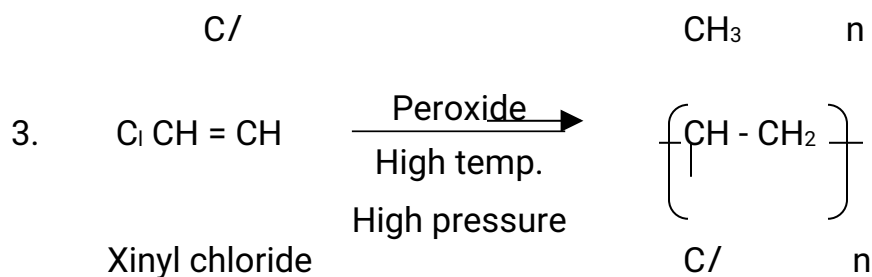
During polymerisation, a catalyst is required in order to initiate the formation of free radicals.

Catalysts popularly used include;

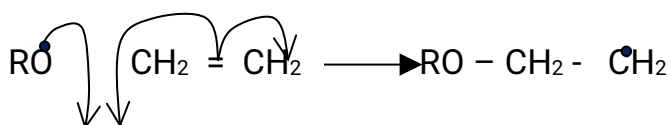
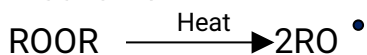
- (i) Molecular oxygen/peroxide
- (ii) Metallic catalyst.

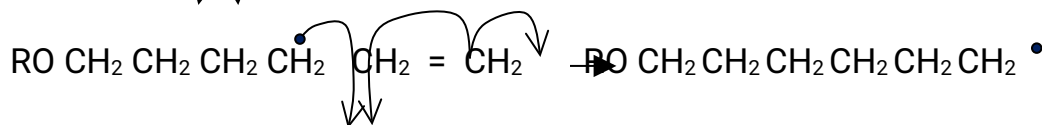
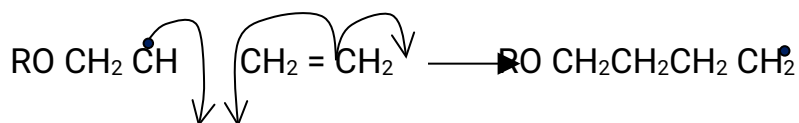
### Example:



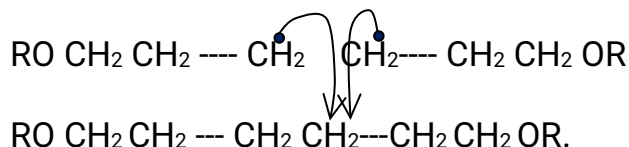


### Mechanism:





**Termination:**



## ALKYNES

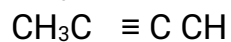
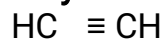
Alkynes are unsaturated hydro carbons which contain a triple bond in the carbon skeleton as their functional group.

They have a general formula  $\text{C}_n\text{H}_{2n-2}$  where  $n \geq 2$ .

## NOMENCLATURE AND ISOMERISM OF ALKYNES

The name of alkyne is formed by replacing the end n, from the corresponding "ane" with "yne" as shown below.

### Alkyne

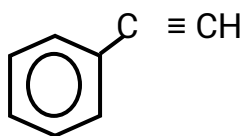


### IUPAC name:

Ethyne

Propyne

Butyne



Phenyl ethyne.



Buta-1, 3- diyne.



Alkynes show chain, functional and position isomerism.

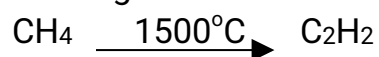
### PREPARATION OF ALKYNES

(a) Only ethyne can be prepared by the following methods:

(i) Partial combustion of methane, yields ethyne.



(ii) Heating of methane in absence of air at  $1500^\circ\text{C}$  also yields ethyne.



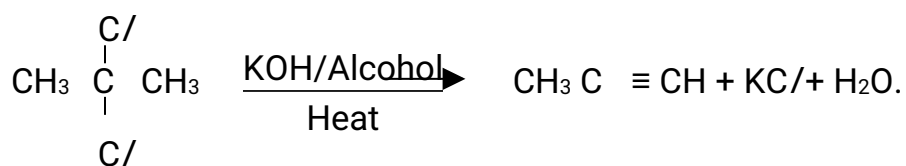
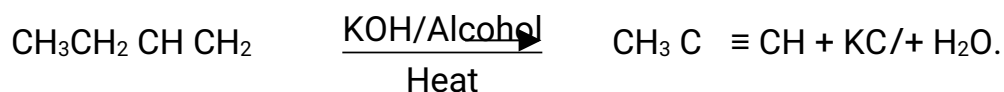
(iii) Hydrolysis of calcium carbide also yields ethyne.



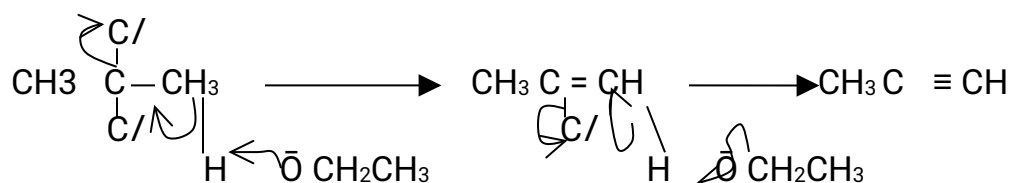
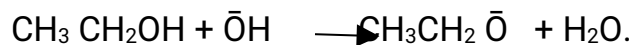
### GENERAL METHODS FOR PREPARATION OF ALKYNES

#### DIHALOGENATION OF DIHALOGENATED ALKENES

This is done by using alcoholic KOH. It is an elimination reaction in which two moles of halogen acids (HX) are lost e.g.



Mechanism:



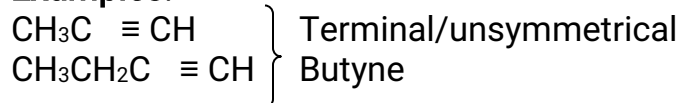


## TYPES OF ALKYNES

Alkynes are classified according to the position of the triple bond.

- (i) Terminal
- (ii) Symmetrical
- (iii) Unsymmetrical

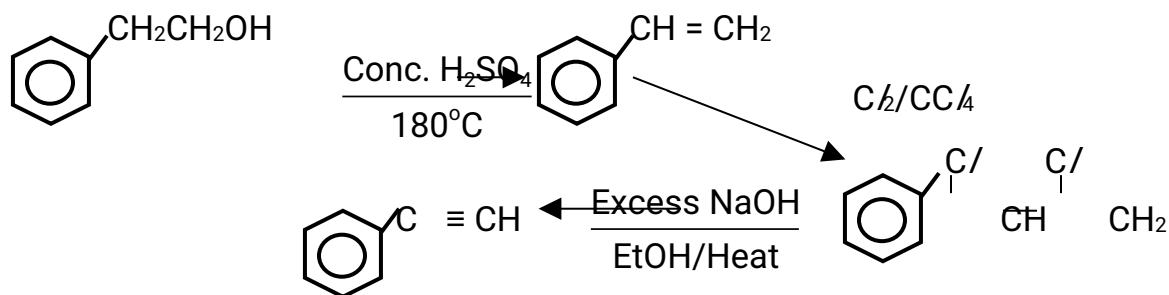
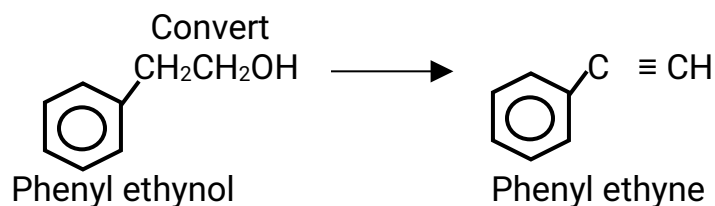
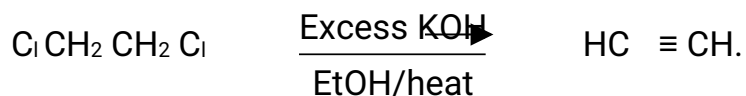
### Examples:

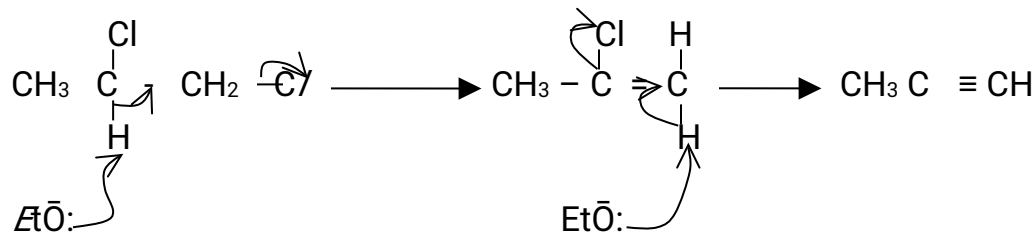
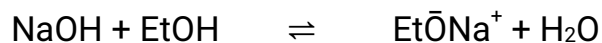
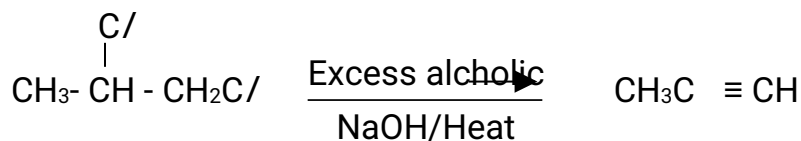


## PREPARATION OF ALKYNES

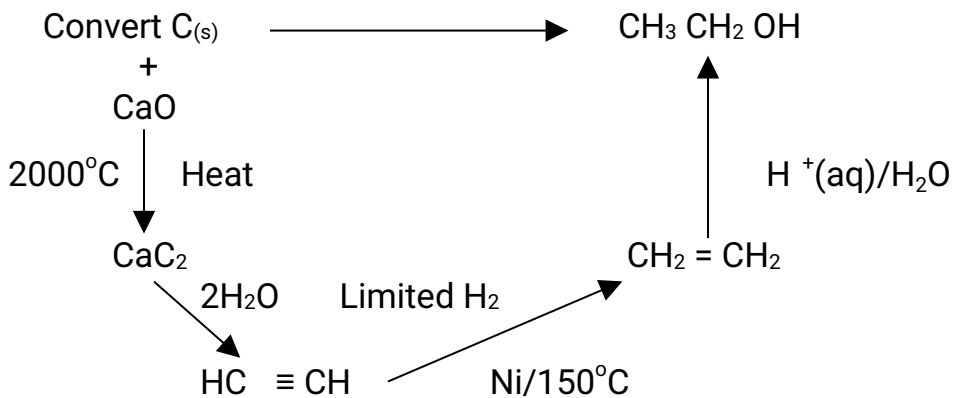
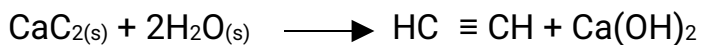
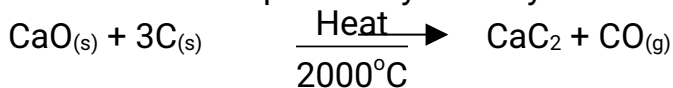
### FROM DIHALOGENO ALKANES

When dihalogeno alkanes are heated with excess alkali in the presence of an organic solvent like ethanol, dehydrohalogenation occurs giving an alkyne.



**Mechanism:****(ii) From CaO.**

This method is specifically for ethyne.

**(iii) From alkane (CH<sub>4</sub> → CH<sub>2</sub>)**

Partial oxidation of alkanes.



1500°C

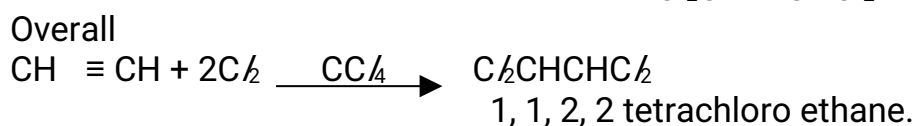
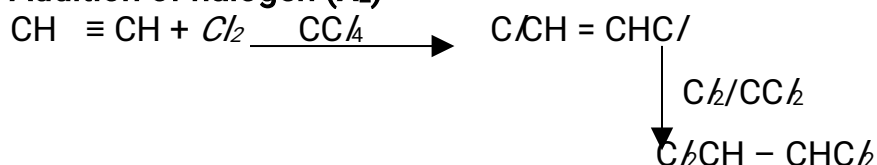
Thermal  
cracking

## REACTIONS OF ALKYNES

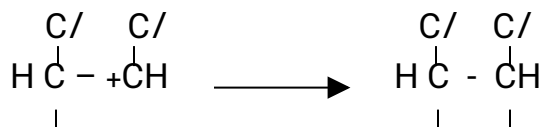
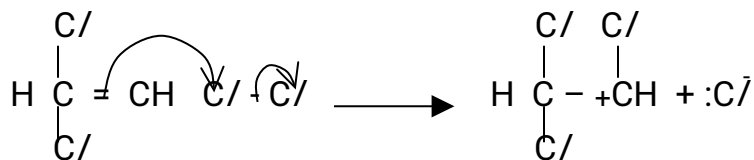
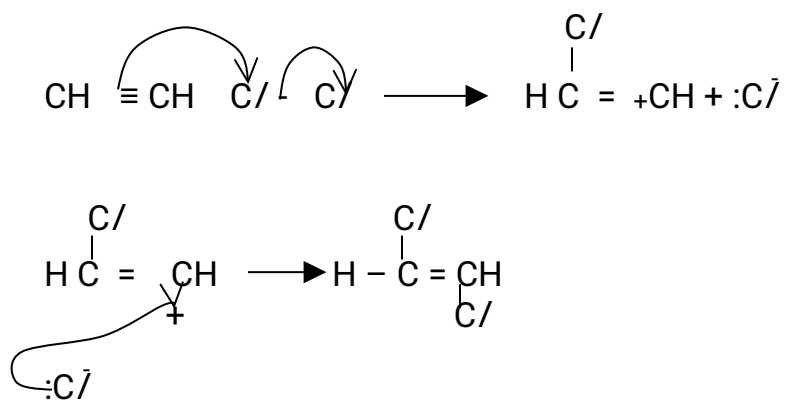
### (i) Electrophilic addition

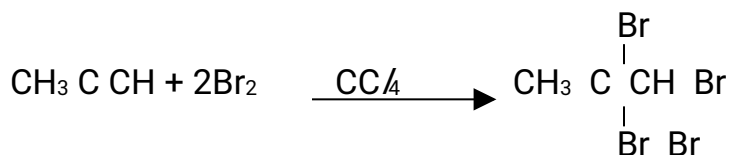
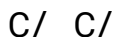
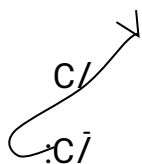
Is where an electron seeking species is added across a triple bond. Alkynes are very reactive because of the presence of  $\pi$  electrons in a triple bond. This reaction is essentially like that of alkenes, however, 2moles of the electrophile are required to saturate an alkyne.

### (ii) Addition of halogen ( $X_2$ )



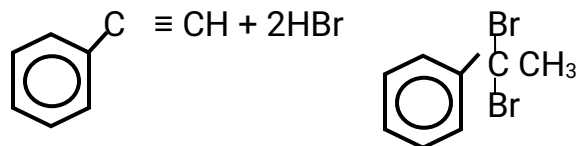
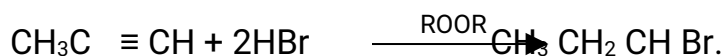
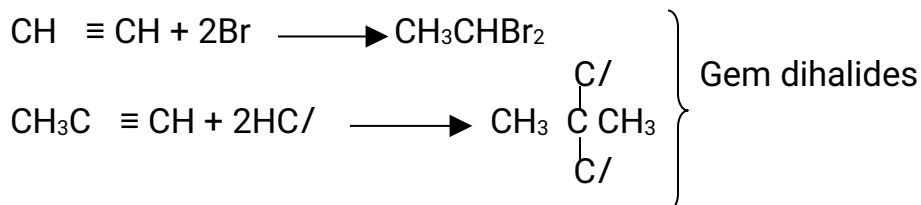
Mechanism:





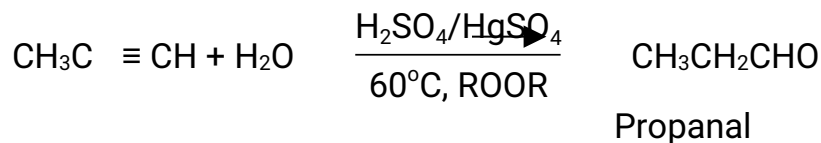
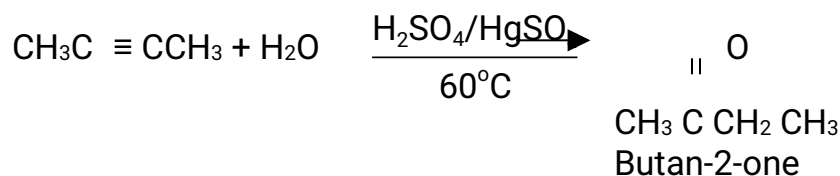
(iii) **Halogen acids (HX)**

Alkynes react with halogen acids to give dihalides depending on the structure of an alkyne, the dihalide can either be a gem or vicinal due to Markovnikov's rule.



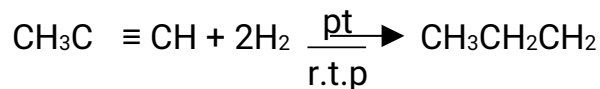
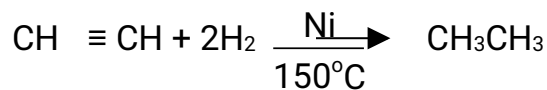
(iv) **With H<sub>2</sub>O.**

Alkynes react with H<sub>2</sub>O in the presence of H<sub>2</sub>SO<sub>4</sub> acid and HgSO<sub>4</sub> with a temperature of about 60°C to form carbonyl compounds. A symmetrical alkyne forms a ketone.



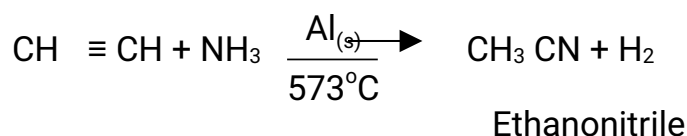
(v) **Reactions with H<sub>2</sub>.**

Alkynes just like alkenes react with  $H_2$  when passed over a suitable catalyst like Ni/ $150^\circ C$ , Pt/r.t.p to form an alkane.



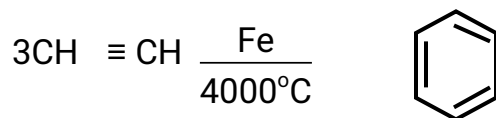
**(vi) Nitrile formation**

When little  $NH_3$  is passed over an alkyne heated over an Al catalyst at  $573^\circ C$ , a nitrile is formed.



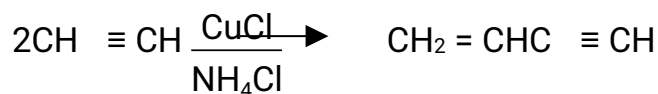
**(vii) Polymerisation**

Due to the presence of a triple bond, alkynes polymerise to form cyclic or aromatic compounds. This occurs when an alkyne is heated in a glass tube containing Fe catalyst ( $4000^\circ C$ ).

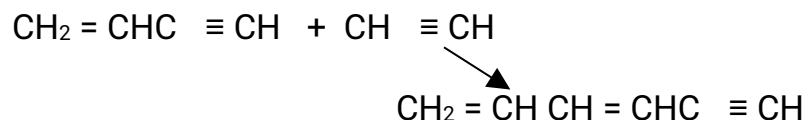


When two molecules of ethyne combine by passing them through a saturated solution of copper(I) chloride in ammonium chloride, vinyl ethyne is formed.

Vinyl ethyne is a good starting material for linear polymers.

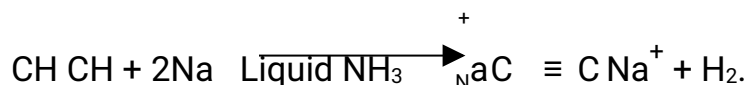


Vinyl ethyne.



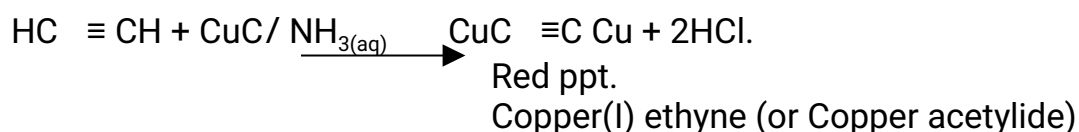
A terminal proton of an alkyne can be substituted by other electropositive metals. This indicates that, such alkynes are acidic.

**NOTE:** Only terminal alkynes behave this way. Non terminal do not.



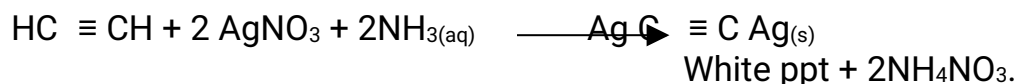
### WITH AMMONIACAL COPPER(I) CHLORIDE SOLUTION

Similarly,  $\text{H}_2$  in alkynes can be substituted when its passed in solution of  $\text{CuCl} \cdot \text{NH}_3$  giving a red precipitate of  $\text{Cu}^+$  alkyne derivative. This reagent is called ammoniacal copper(I) chloride solution,  $\text{CuC}/\text{NH}_{3(\text{aq})}$ .



### WITH AMMONIACAL SILVER NITRATE SOLUTION.

When a terminal alkyne is bubbled through a solution of ammoniacal  $\text{AgNO}_3$ , a white precipitate of silver dicarbide is formed.



Differentiate between  $\text{CH}_3\text{C} \equiv \text{CH}$  and  $\text{CH}_3\text{C} \equiv \text{CCH}_3$

Reagent: Ammoniacal copper(I) chloride solution.

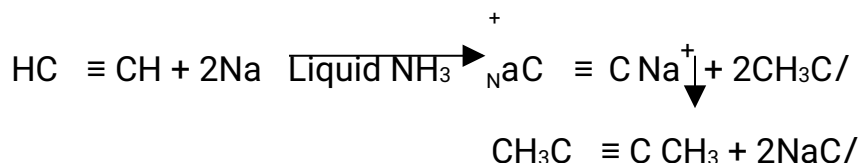
With  $\text{CH}_3\text{C} \equiv \text{CH}$  : No observable change.

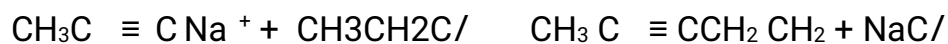
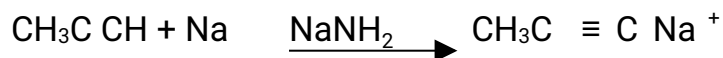
With  $\text{CH}_3\text{C} \equiv \text{CCH}_3$  : A red precipitate is observed.

The above reactions are used for distinguishing between terminal alkynes and non terminal

### SYNTHESIS OF HIGHER ALKYNES

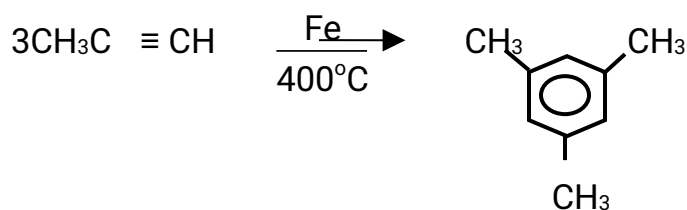
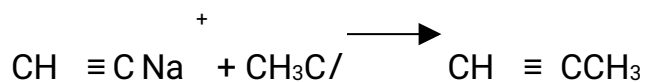
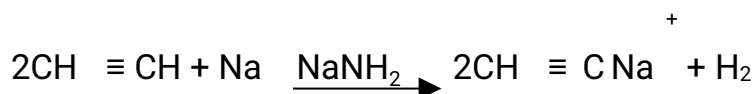
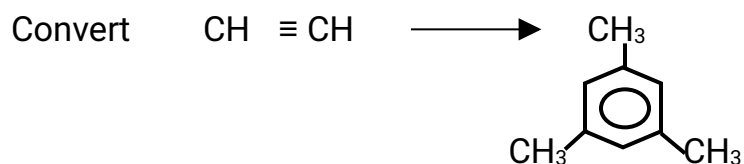
When an alkyne is passed through liquid  $\text{NH}_3$  in presence of Na metal and the derivative formed is reacted with an alkyl halide, the chain is increased giving a higher alkyne.





### Physical properties of alkynes

- They are insoluble in  $\text{H}_2\text{O}$ .
- They are very soluble in organic solvents like benzene,  $\text{CCl}_4$ .
- They are less dense than  $\text{H}_2\text{O}$ .
- Their boiling points increase in number of carbon atoms. E.g. Ethyne boils at  $-75$ , propyne  $-43$ , butyne at  $91$ .
- They have got almost the same boiling points as those of corresponding alkenes.
- They have got the same vanderwaal's forces of attraction.



## AROMATIC COMPOUNDS (ARENES)

The term aromatic is derived from the Greek word Aroma to mean sweet smell. This term has been carried forward by modern Chemists used as aromatic to mean or describe the study of compounds with benzene ring.

Benzene is a parent compound of aromatic compounds.

### STRUCTURE AND BONDING OF BENZENE

Structure and bonding of benzene was made clear by analysis made by Chemists. They found out that a clear colourless compound was isolated from distillation of crude oil and on analysis, it was found to contain 92.3% carbon and 7.7% hydrogen only.

When 0.250g of this liquid was vapourized at 100°C, they found out that the liquid occupies 98cm<sup>3</sup>. The results were subjected to analysis.

Element	C	H
Percentage composition	92.3	7.7
Relative atoms	$\frac{92.3}{12}$	$\frac{7.7}{1}$
Moles	7.69	7.70
Mole ratio	$\frac{7.69}{7.69}$	$\frac{7.70}{7.69}$



1                      1

Empirical formula = CH

$$\begin{aligned} \text{Volume of gas at } 100(373)\text{K} \\ &= \frac{24 \times 373}{293} \\ &= 30.04 \text{ dm}^3 \end{aligned}$$

98cm<sup>3</sup> of liquid contain 0.25g at 100°C.

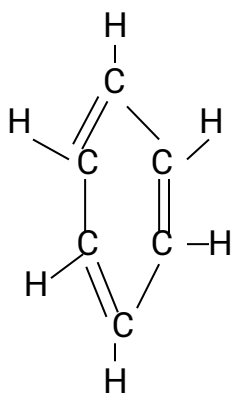
$$\begin{aligned} \text{Mass of 1mol } \left( \frac{6.25 \times 30040}{98} \right) \text{g} \\ &= 76.6 \quad 77\text{g}. \end{aligned}$$

$$\begin{aligned} (\text{CH})_n &= 77 \\ 12n + n &= 77 \\ n &= \frac{77}{13} \\ n &= 5.92 \approx 6. \end{aligned}$$

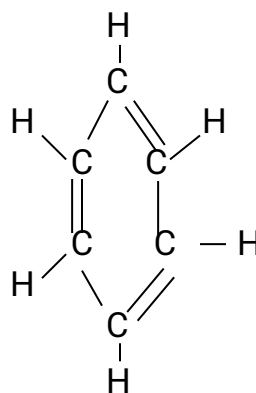
Molecular formula is C<sub>6</sub>H<sub>6</sub>.

The structure proposed was C<sub>6</sub>H<sub>6</sub> with all its isomers.

In 1865, Kekule proposed that the structure of benzene molecule is not a straight chain but an hexagonal ring consisting of alternating double bonds.



OR



But a number of evidences proved out that the above structure, proposed by Kekule was not true. Benzene does not have true double bonds since it does not

undergo addition reaction with bromine in darkness.

### EVIDENCE PUT TO EXPLAIN BENZENE STRUCTURE.

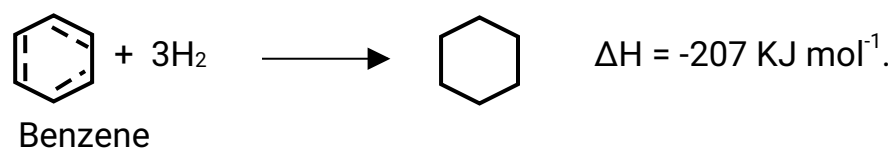
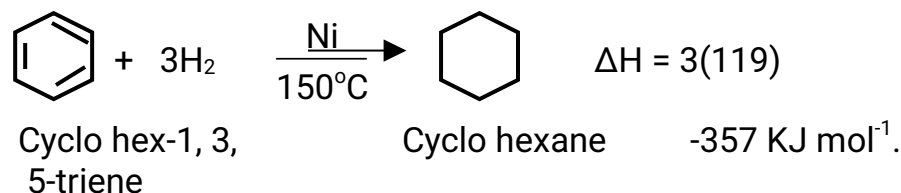
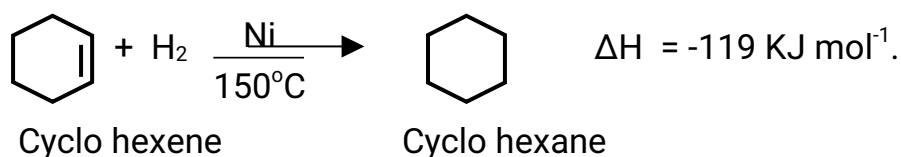
- (i) X-ray structure determination to measure the length of carbon-carbon bonds.

Compound	Bond	Length/nm
Cyclo hexane	C – C	0.154
Cyclo hexene	C = C	0.134
Benzene	C $\equiv$ C	0.140

The measurements of the bond length showed that, the bond length of benzene are between single and double, therefore are not true double bonds ruling out Kekule's structure of double bonds.

- (ii) **Thermodynamic measurements:**

- (a) Enthalpy of hydrogenation.



### Conclusion:

The double bonds of benzene are not true double bonds.

Benzene is more stable than cyclo hex-1, 3, 5, triene by an extra 150 KJmol<sup>-1</sup>

1

- (b) **Enthalpy of formation of benzene.**

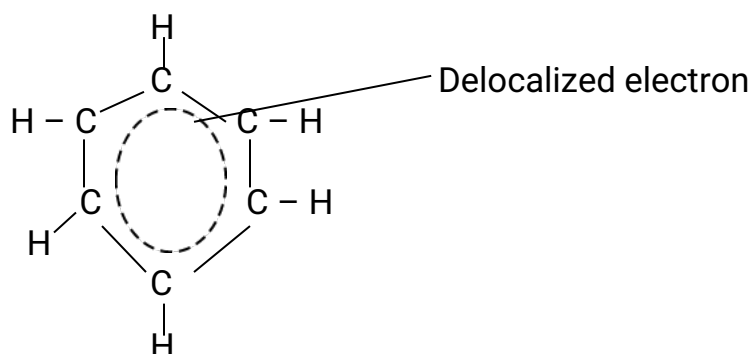
The theoretical/calculated enthalpy of formation of benzene is 252 KJ mol<sup>-1</sup> but the practical value of this energy is 82 KJ mol<sup>-1</sup>. This means benzene is more stable by 170 KJ mol<sup>-1</sup> of energy.



**(c) Reactions of benzene**

Benzene does not undergo addition reactions but undergoes substitution reactions implying it does not have double bonds but a system of delocalized electrons or the pie ( $\pi$ ) system).

**DELOCALIZATION OF BENZENE STRUCTURE**

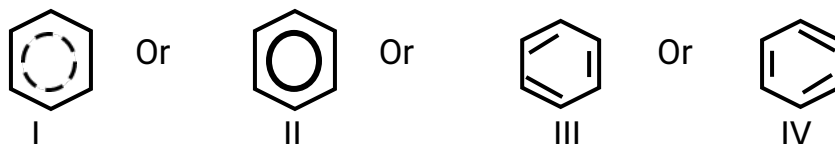


Benzene structure has got six  $\sigma$  carbon – carbon bonds.

It has got 3- $\pi$  carbon – carbon bonds.

The  $\pi$  bonds are not localized but they are shared by the six carbon atoms giving a  $\pi$  system of electrons and this is called delocalization.

**STRUCTURE OF BENZENE.**



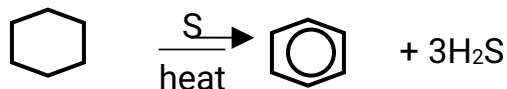
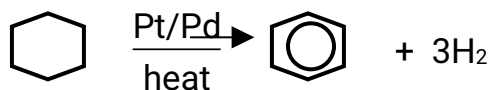
**BENZENE**

Is the parent aromatic compound with a hexagonal ring of carbon atoms having a  $\pi$  system of electrons.

**FORMATION OF BENZENE**

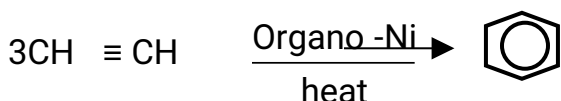
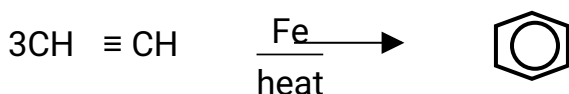
(i) From dehydrogenation of cyclo hexane when heated with Pt/Pd or heated

with sulphur. Cyclo hexane loses H<sub>2</sub> atoms to form benzene.



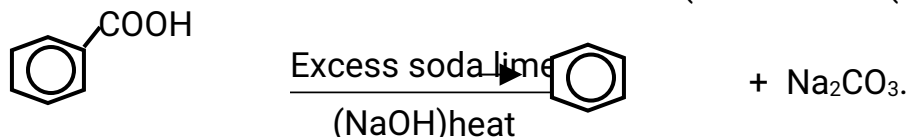
(ii) **Polymerisation of ethyne**

When ethyne is heated with an organo iron or Nickel, it undergoes cyclisation to form benzene.

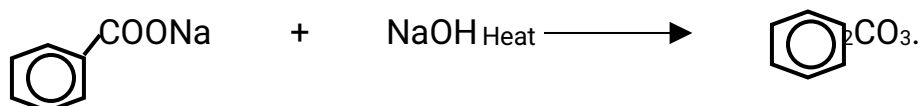
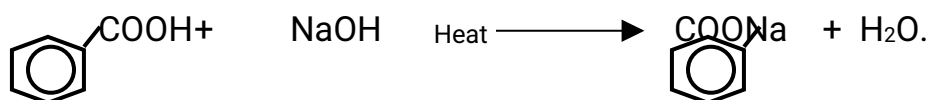


(iii) **Decarboxylation of benzoic acid**

When benzoic acid is heated with soda lime (a mixture of (NaOH and CaO).



Two step reaction

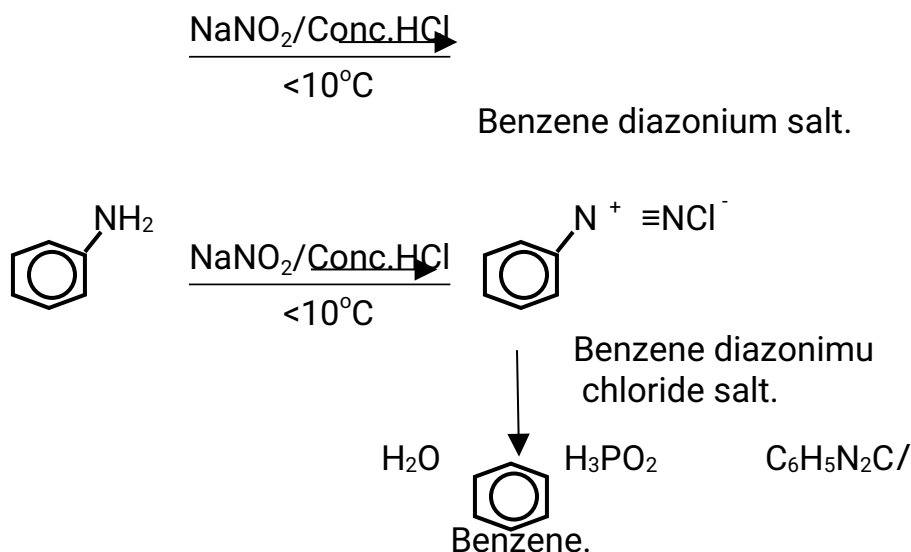


(iv) **From benzene diazonium salt.**

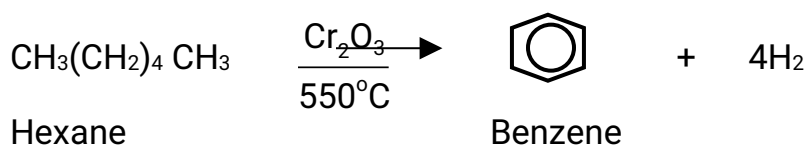
When benzene diazonium salt is reacted with pyrophosphorus acid in presence of water, benzene is formed.

The formation of benzene diazonium salt can be started from phenyl amine which is reacted with a mixture of NaNO<sub>2</sub> and HCl/ (HNO<sub>2</sub> acid) at a temperature <10.

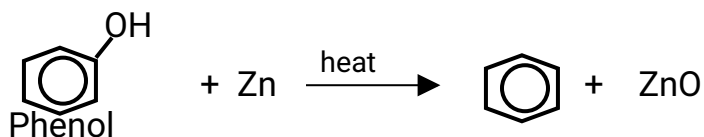




- (v) Catalytic reformation of petroleum under heat of about  $550^\circ\text{C}$  in the presence of  $\text{Cr}_2\text{O}_3/\text{Al}_2\text{O}_3$ .



- (vi) **From phenol**  
When phenol is heated with zinc dust, benzene is formed.



## PROPERTIES OF BENZENE

### PHYSICAL PROPERTIES

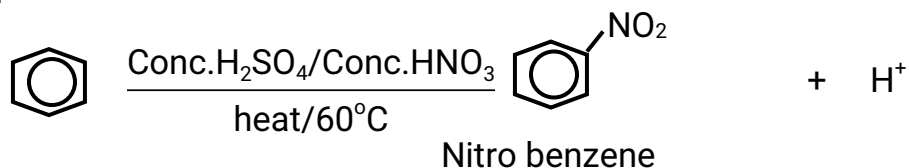
- ✓ It is a colourless liquid.
- ✓ It has got a x-tic smell.
- ✓ It is insoluble in water but dissolves in organic solvents e.g.  $\text{CCl}_4$ , methyl benzene.
- ✓ It burns with a highly sooty flame.

### CHEMICAL PROPERTIES

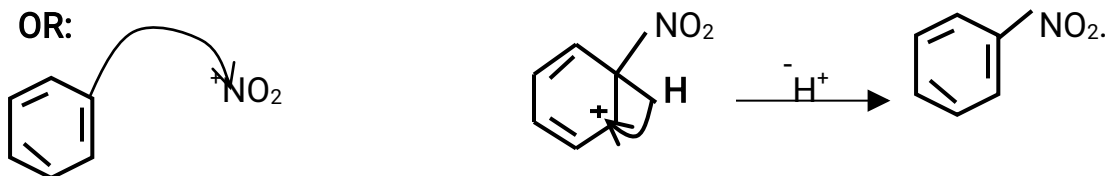
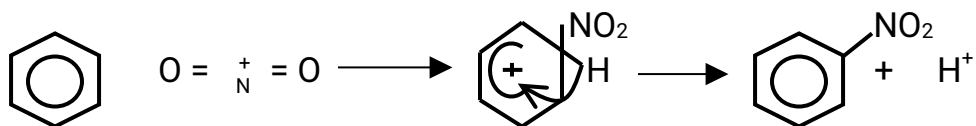
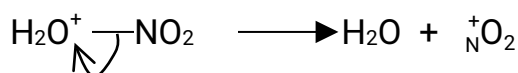
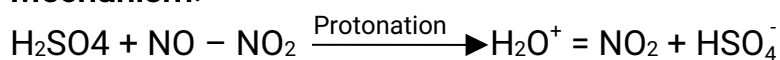
Benzene undergoes electrophilic substitution reaction where the hydrogen atom on the ring is replaced by another electrophille. Benzene therefore reacts mainly with electrophilles (Positively charged species) to form substituted products on the benzene ring.

(i) **Nitration reaction**

Is a reaction where the nitronium ion  ${}^+\text{NO}_2$ , ( $\text{O} = +\text{N} = \text{O}$ ) substitutes the hydrogen in the benzene ring. The electrophile of the nitronium ion is generated from a mixture of Conc.  $\text{H}_2\text{SO}_4$  and Conc.  $\text{HNO}_3$  acid heated at a temp. of  $60^\circ\text{C}$ .



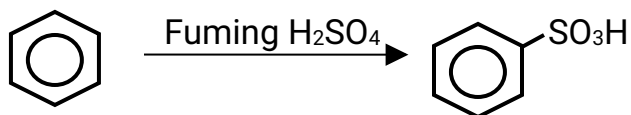
**Mechanism:**



(ii) **Sulphonation:**

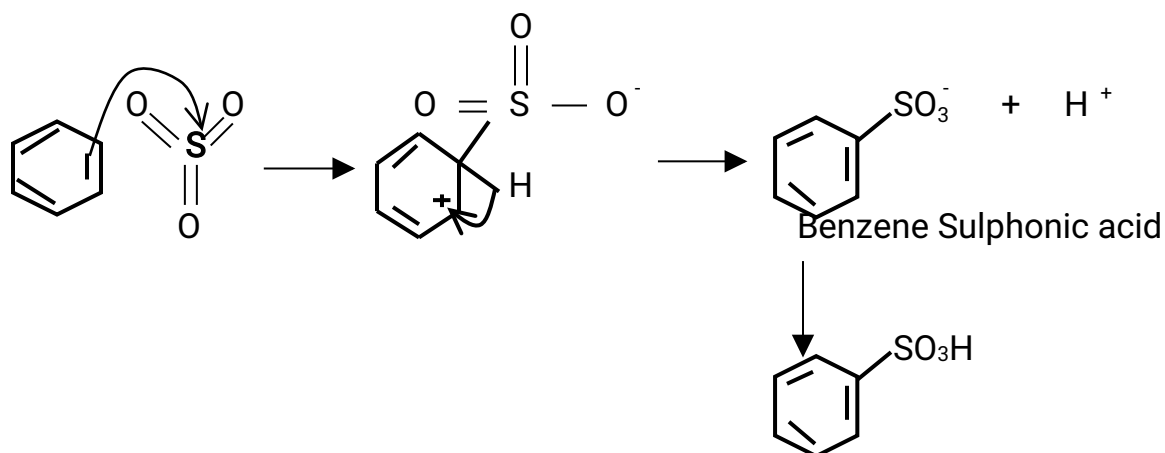
This is a reaction where benzene reacts with Conc. Or fuming  $\text{H}_2\text{SO}_4$  acid to form benzene sulphuric acid.

The electrophille in this reaction is  $\text{SO}_3$  which is a neutral electrophille generated from fuming sulphuric acid.



**Mechanism:**





### (iii) Halogenations

Benzene reacts with halogens like  $\text{Cl}_2$ ,  $\text{Br}_2$  only in the presence of halogen carrier which polarizes the halogen molecule by accepting the lone pair of electrons so that the electrophile is generated.

In the absence of the halogen carrier, benzene does not react with halogens.

Examples of halogen carriers used are:

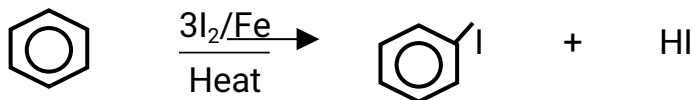
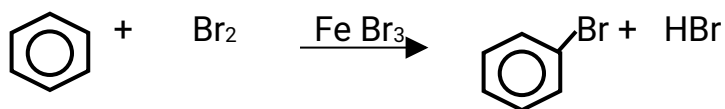
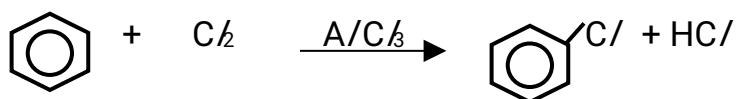
$\text{AlCl}_3$

$\text{CH}_3\text{COOH}$

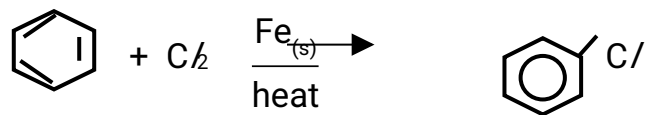
$\text{FeCl}_3$

$\text{BF}_3$

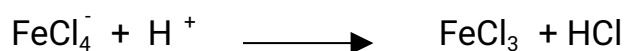
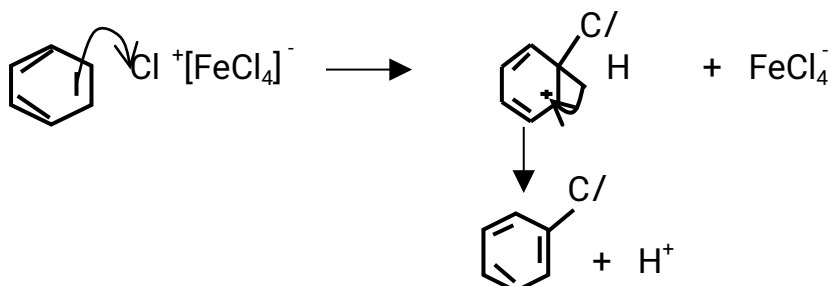
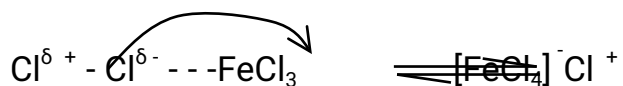
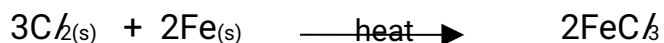
$\text{Fe}$ ,  $\text{Cl}_2$ /Heat



**Mechanism:**



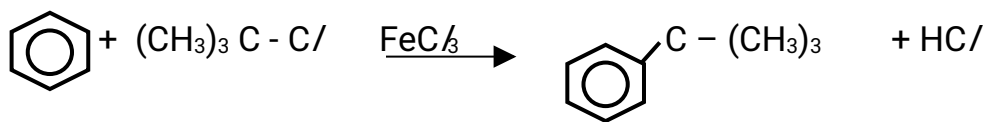
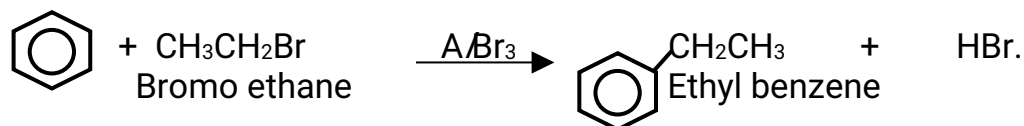
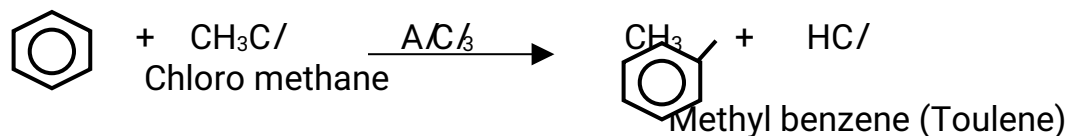
**Mechanism:**



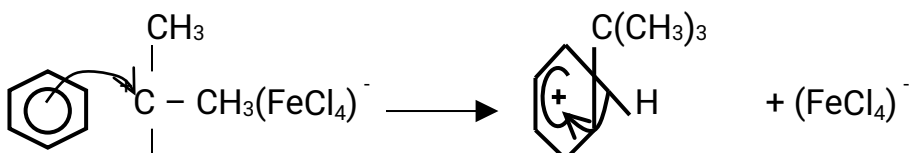
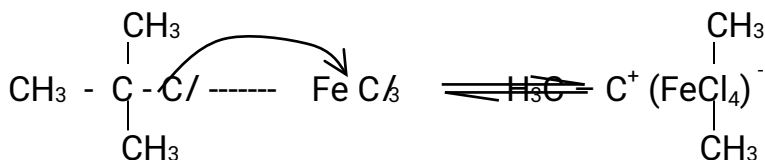
### Alkylation of Benzene

This is the reaction where an aromatic compound reacts with an alkyl halide in the presence of a halogen carrier to form an alkyl aromatic compound.

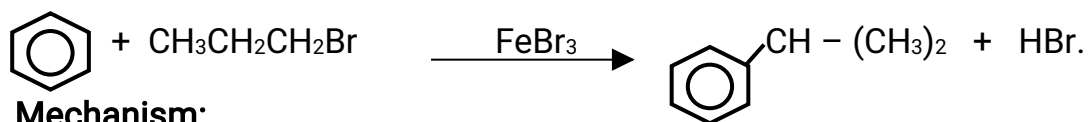
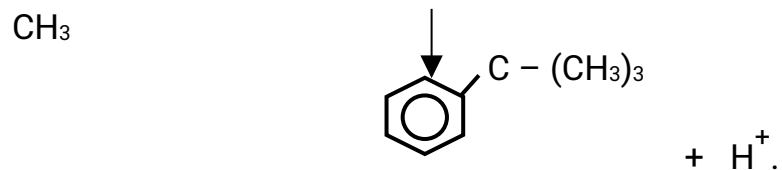
This reaction was first identified by **Friedel craft**. It is also called Friedel Craft alkylation.



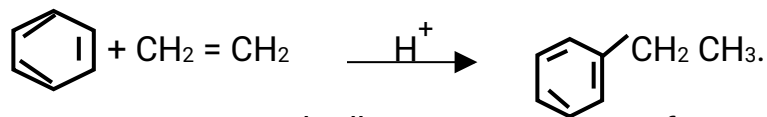
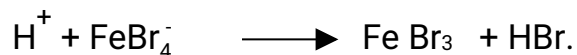
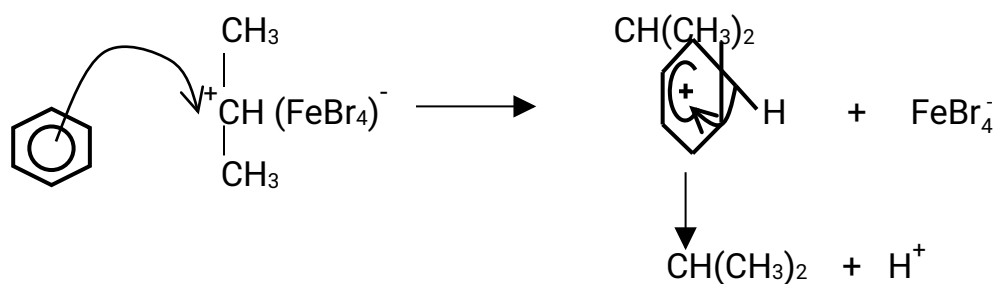
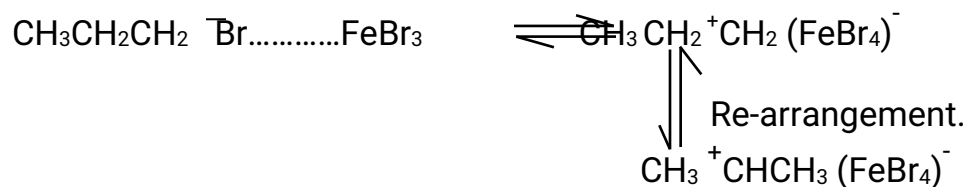
### Mechanism



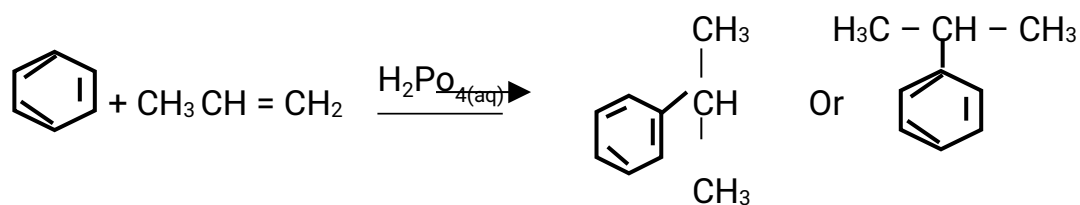




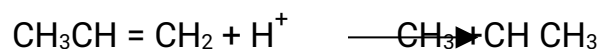
**Mechanism:**

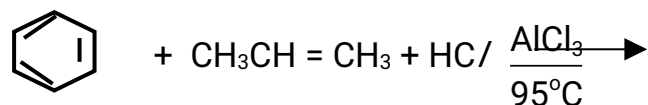
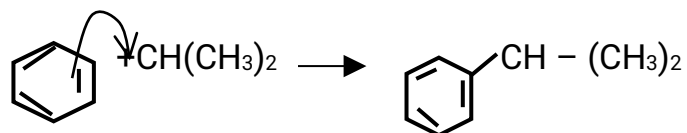


Benzene reacts with alkenes in presence of a mineral acid.  
The mineral acid is necessary to form an electrophile.



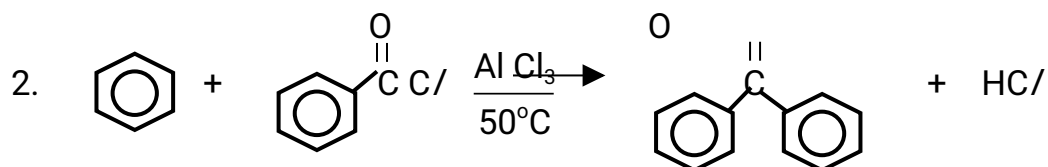
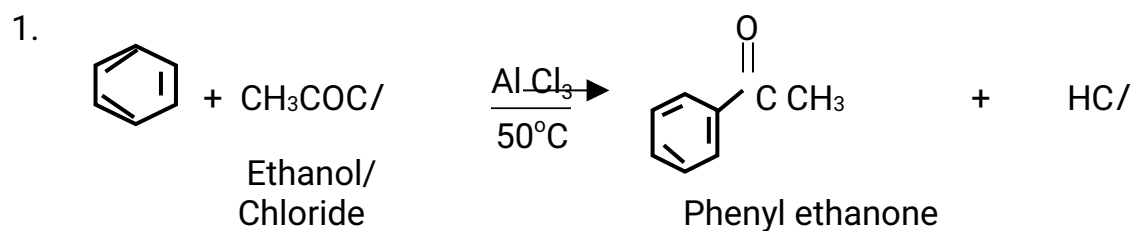
**Mechanism:**



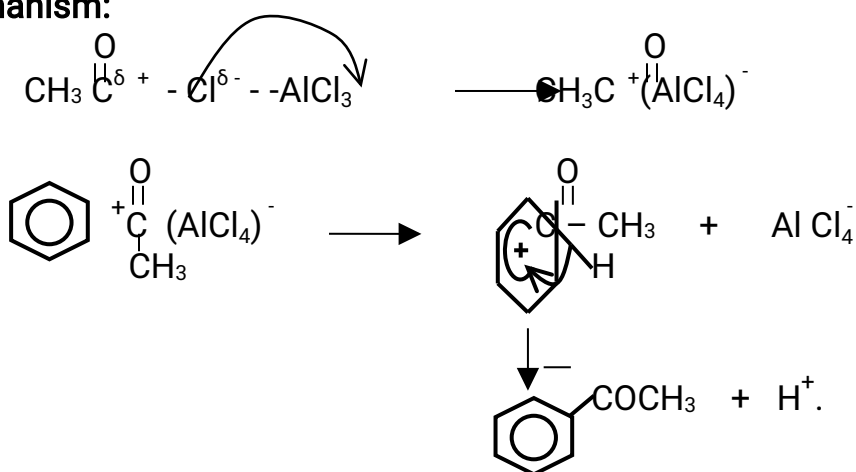


### Acylation

Is a reaction of benzene with acid halides in the presence of halogen carriers to form aromatic ketones at a temperature of  $50^\circ\text{C}$ .

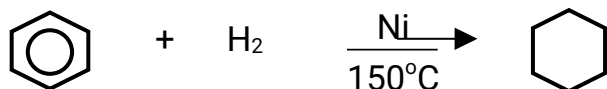


### Mechanism:



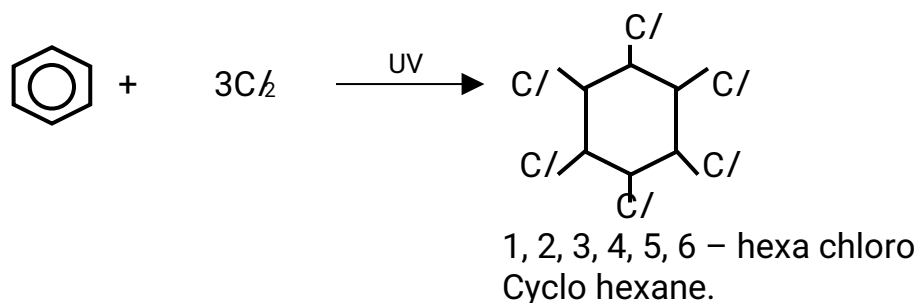
### Hydrogenation:

Benzene reacts with hydrogen in the presence of Ni catalyst when heated at 150°C to form cyclohexane.



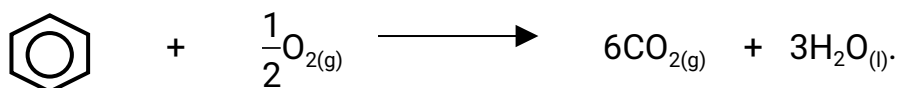
### Halogen in presence of light (uv).

When chlorine is mixed with benzene in the presence of sunlight or UV rays, the benzene ring becomes highly substituted with chlorine atoms to form 1, 2, 3, 4, 5, 6 hexachloro benzene.

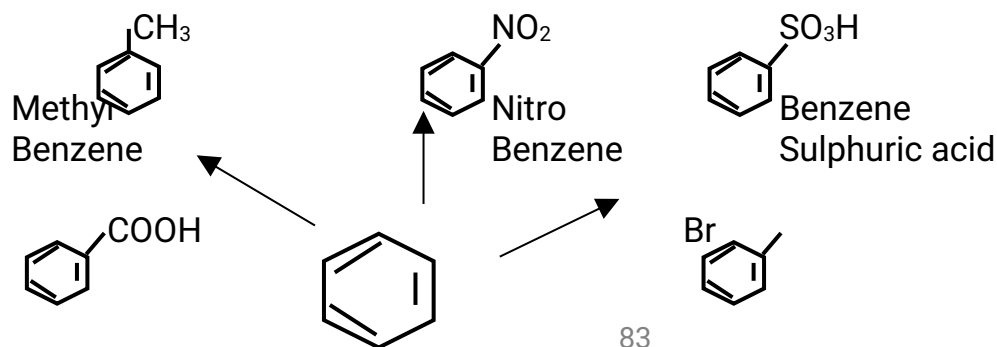


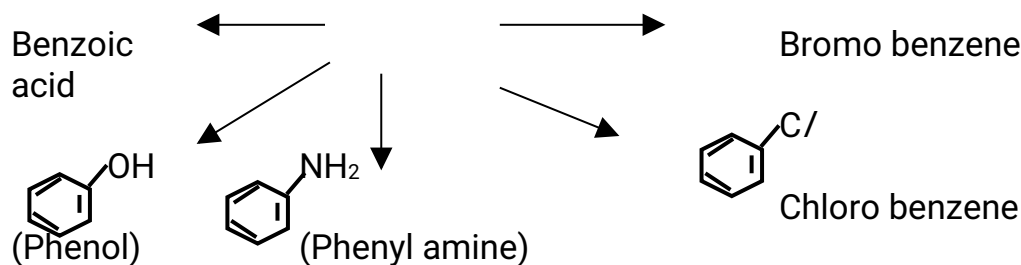
### Combustion

Benzene burns in air with a sooty flame to form CO<sub>2</sub> and H<sub>2</sub>O.



### DERIVATIVES OF BENZENE

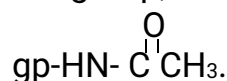




Derivatives of benzene react differently depending on the behaviour of the group of atom(s). The chemical behaviour of these groups are divided into three groups.

(i) The group which substitutes the benzene ring directing any incoming group to ortho and para positions. E.g.

(a) OH-group, ethoxide group OCH<sub>3</sub>, CH<sub>3</sub>-gp, NH<sub>2</sub>-nitro group, amides

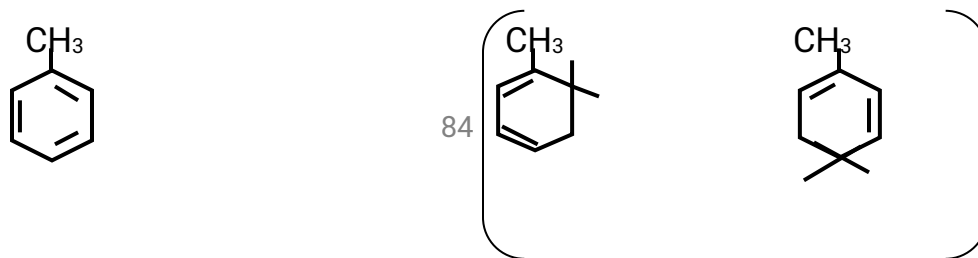
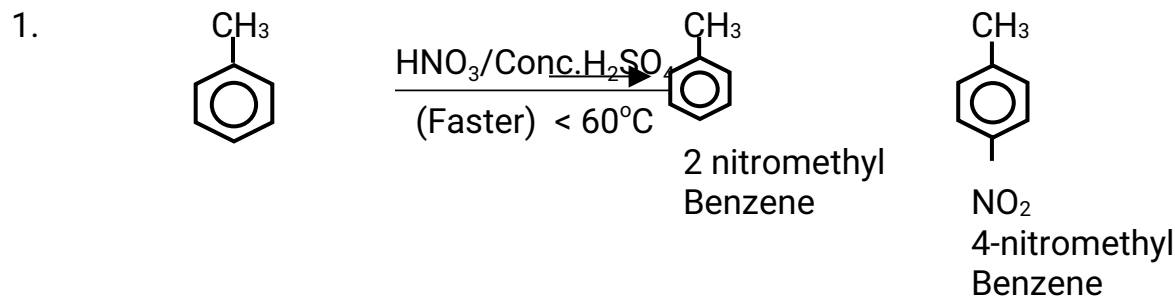


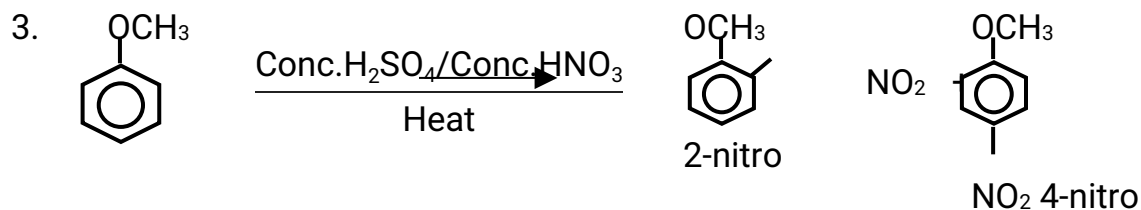
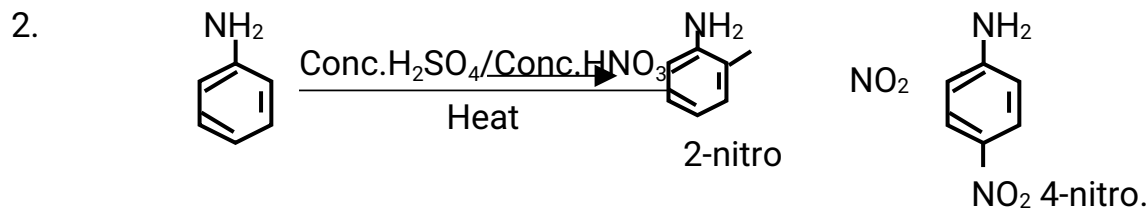
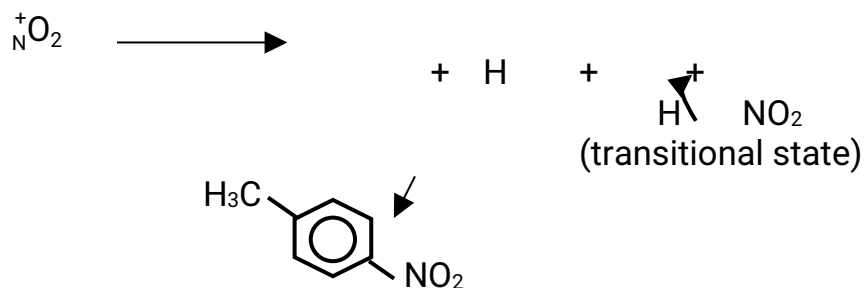
These groups have got positive inductive effect by pushing the electrons towards the ring, activating the benzene ring so that the electrons are available at positions 2, 4, and 6 where the incoming electrophile are added.

By activating the ring, they increase the electron density on the ring so that the electrophiles are added faster.

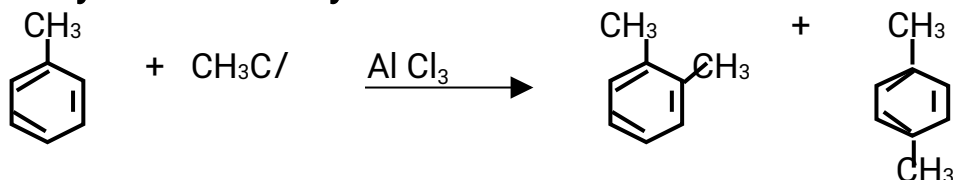
Usually two products are given, i.e. position 2 and position 4 products.

## METHYL BENZENE

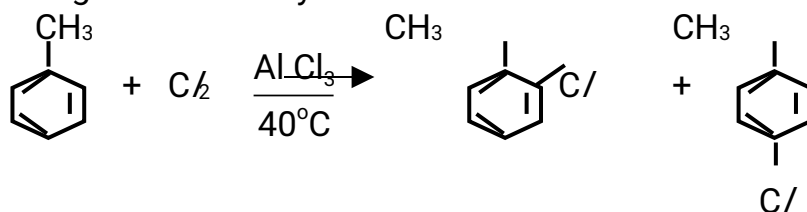




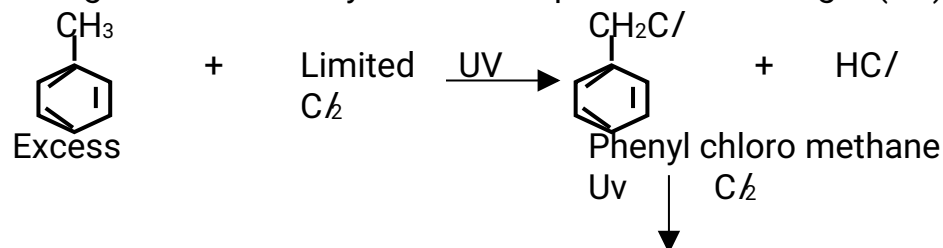
### Alkylation of methyl benzene.

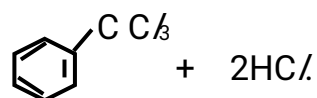


### (a) Halogenation of methyl benzene.



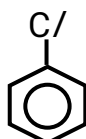
### (b) Halogenation of methyl benzene in presence of sunlight (UV).



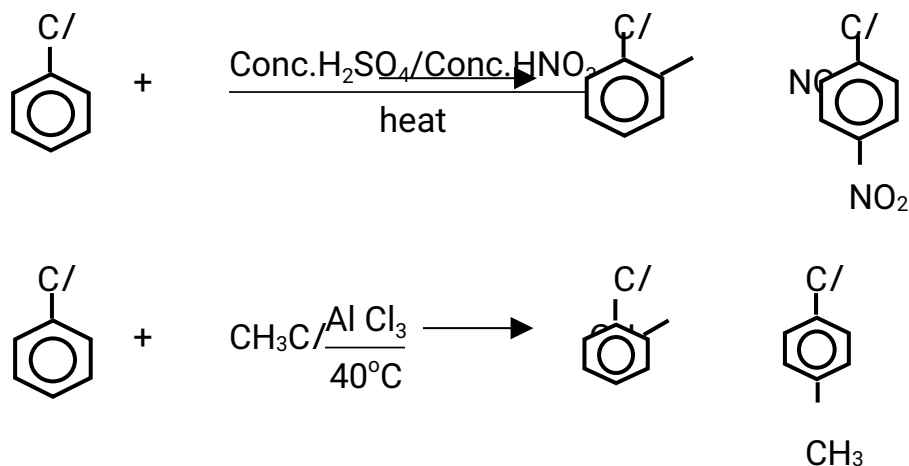


This reaction follows a free radical mechanism.

## CHLORO BENZENE



Chloro benzene or other halogeno benzene compounds can substitute the ring at position 2 and 4 but at a much slower rate. This is because halogens have got a negative inductive effect (electronegative) and withdrawal electrons from the ring towards themselves so that the pi electrons on the ring can be availed to electrophiles at position 2 or 4 where there is a high electron density.



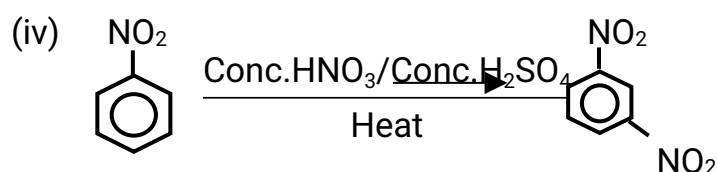
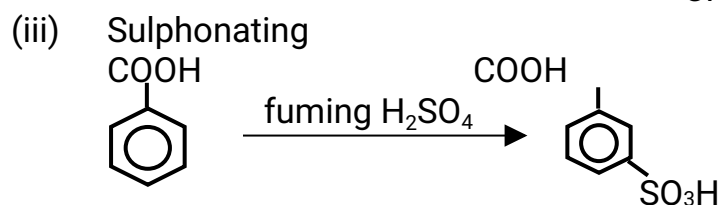
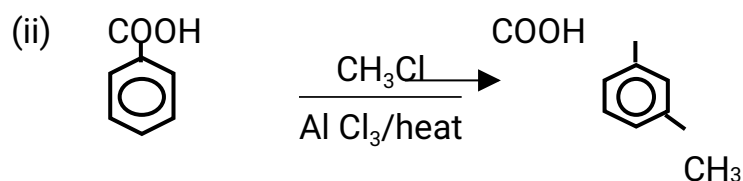
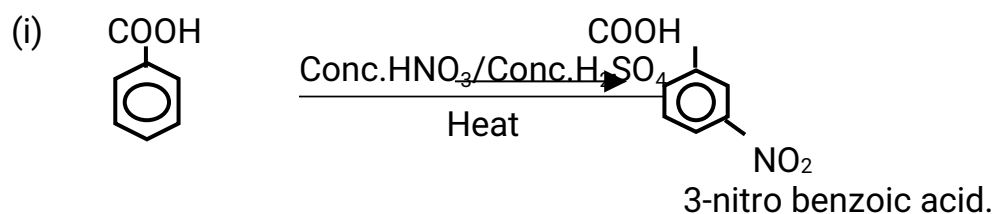
Groups that direct further substitution at position 3 of the ring do so at a much slower rate including COOH, NO<sub>2</sub>, CN, SO<sub>3</sub>H.

## BENZOIC ACID



Meta(3) position directors.

Carboxylic acid on benzoic acid directs the incoming substituents at position 3 so that one major product is formed.

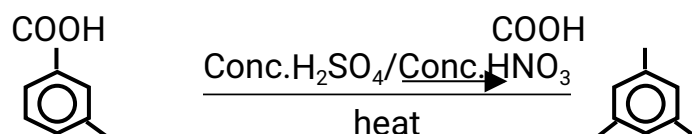
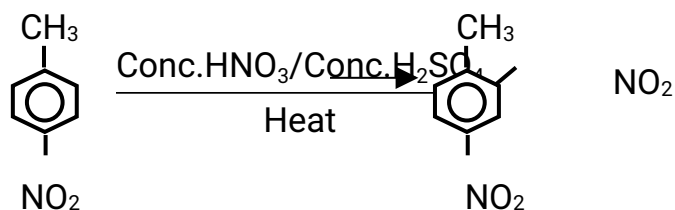


The above are Meta (3) directors because they pull electrons away from the ring and decreasing electron density at position 2 or 4 and the only place where the electrons are available is the Meta position.

## DISTRIBUTED COMPOUNDS

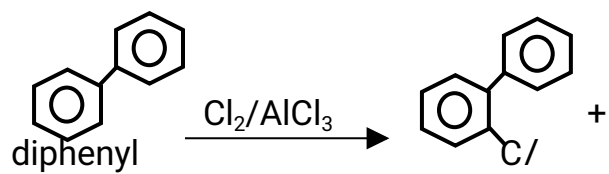
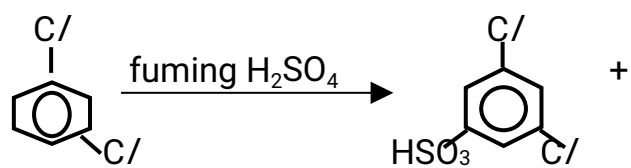
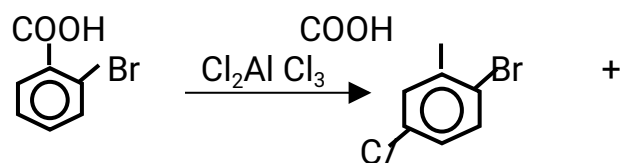
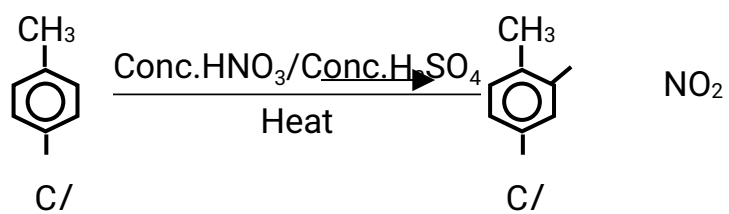
When there are two substituents on the benzene ring, the orientation effect of these two will be determined or predicted by their reinforcing effects or the order of their superiority. The following order is normally observed.

- OR > NR<sub>2</sub> > Alkyl group > halogen > Meta directors.



SO<sub>3</sub>H

NO<sub>2</sub> SO<sub>3</sub>H



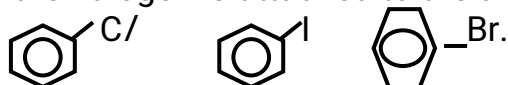


## ALKYL HALIDES (Halogeno alkanes)

Alkyl halides are compounds in which halogen atoms are directly attached onto the hydro carbon chain or aromatic ring.

R – X where R = Alkyl or Aryl group.  
X = Halogen atom (C/, Br, I, F)

When the halogen is attached to the aromatic ring, it is called aromatic halide. e.g.



When the halogen is directly attached to the hydro carbon, it is called alkyl halide / halogen alkane. E.g.

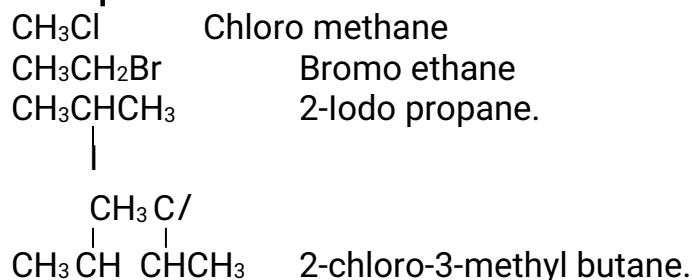


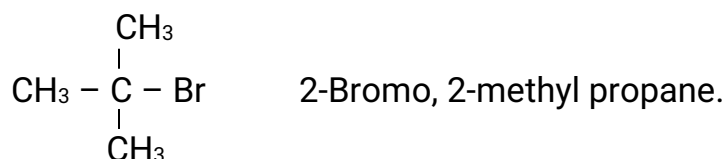
## ALKYL OR HALOGENO ALKANES

### Nomenclature:

Alkyl halides are named as derivatives of corresponding alkanes by inserting the number and the prefix, chloro, Bromo, Iodo, Fluoro, to the alkane name.

### Example:





## ISOMERISM

Alkyl halides exhibit two types of isomerism.

- (i) Chain isomerism
- (ii) Positional isomerism

### CHAIN ISOMERISM:

These arise as a result of difference in the arrangement of carbon atoms giving different carbon skeleton but the position of the halogen atom remains the same.  
C<sub>4</sub> H<sub>9</sub> Br.

1. CH<sub>3</sub> CH<sub>2</sub> CH<sub>2</sub> CH<sub>2</sub> Br                      1-Bromo butane.
2. CH<sub>3</sub> CH CH<sub>2</sub> Br                      1-Bromo-2-methyl propane.  
 $\begin{array}{c} \text{CH}_3 \end{array}$

### POSITIONAL ISOMERISM:

These arise as a result of the different positions taken by the halogen atom on the same chain. E.g.  
C<sub>4</sub> H<sub>9</sub> Br.

- $$\begin{array}{ll}
 \text{CH}_3 \text{ CH}_2 \text{ CH}_2 \text{ CH}_2 \text{ Br} & \text{1-Bromo butane} \\
 \text{CH}_3 \text{ CH}_2 \text{ CH} \text{ CH}_3 & \text{2-Bromo butane} \\
 | & \\
 \text{Br} &
 \end{array}$$


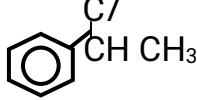

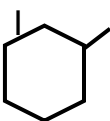
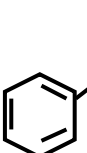
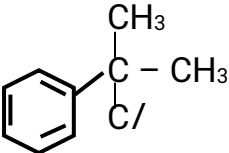
### METHODS OF PREPARATION:

Alkyl halides are prepared from the following:

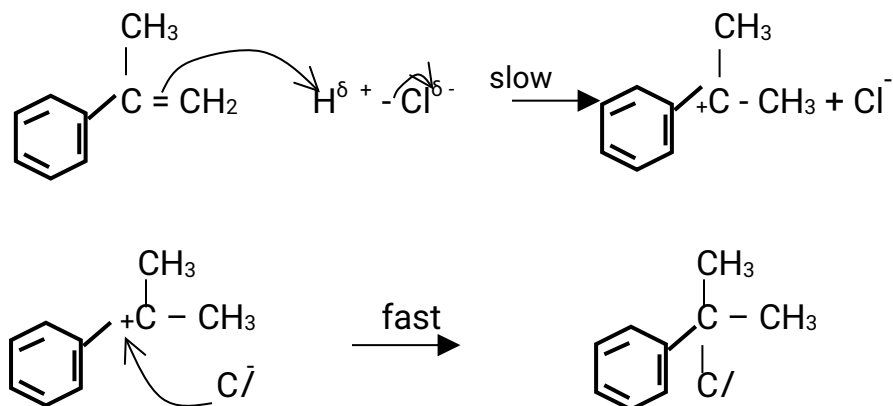
#### (i) From Alkenes:

Addition of an halogen acid to an alkene at room temperature gives an alkyl halide. When unsymmetrical alkene is used, addition of an halogen acid gives a product predicted by Markovnikov's rule. But if a peroxide used, then the product will be that of anti Markovnikov's rule.

1. CH<sub>3</sub>CH = CH<sub>2</sub> + HBr                      r.t.p                       $\begin{array}{c} \text{Br} \\ | \\ \text{CH}_3 \text{ CH} \text{ CH}_3 \end{array}$

2.  $\text{CH}_3\text{CH}=\text{CH}_2 + \text{HBr} \xrightarrow[\text{Heat}]{\text{ROOR}}$   $\text{CH}_3\text{CH}_2\text{CH}_2\text{Br}$
3.   $\text{CH}=\text{CH}_2 + \text{HC}/ \longrightarrow$  
4.   $+ \text{HI} \longrightarrow$   Iodo cyclo hexane.
5.   $+ \text{HC}/ \longrightarrow$  

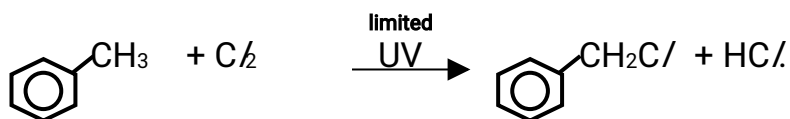
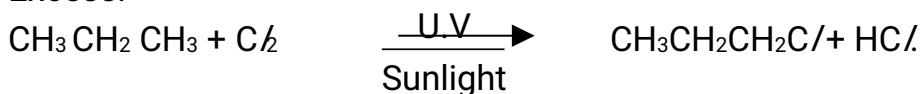
**Mechanism:**



(ii) **From Alkanes**

When alkanes are reacted with a halogen molecule in presence of UV rays of sunlight. A mixture of alkyl halides are formed but if the reaction conditions are controlled by using excess alkane or limited halogen molecule, then the product formed would be one type of alkyl halide.

Excess:



(iii) From Alcohols:

Alcohols are the most important source of preparing alkyl halides in reaction where hydroxyl groups of alcohol are replaced by halogen atoms which could be derived from:

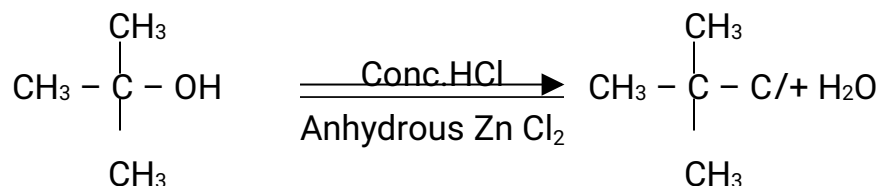
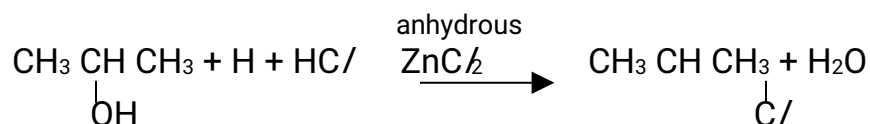
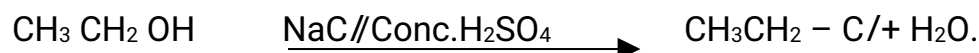
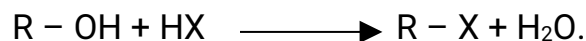
- halogen acids (HX).
- Phosphorus tri or penta halides (PX<sub>3</sub>/PX<sub>5</sub>)
- Thionyl chloride (SOCl<sub>2</sub>).

(a) **Action of halogen acid on the alcohol:**

When a halogen acid is heated with an alcohol in presence of a dehydrating agent e.g.

- ZnCl<sub>2</sub> (anhydrous)
- Conc. H<sub>2</sub>SO<sub>4</sub>.
- Al<sub>2</sub>O<sub>3</sub> (Dry)

An alkyl halide is formed. Sometime the halogen acid is generated "insitu."



A solution of anhydrous ZnCl<sub>2</sub> in Conc.HCl/ (Lucas reagent) is an important reagent for distinguishing the three classes of alcohols i.e.

- Primary.
- Secondary and tertiary alcohols.

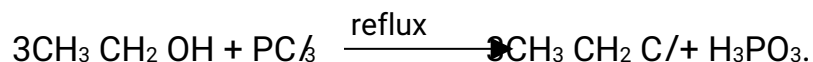
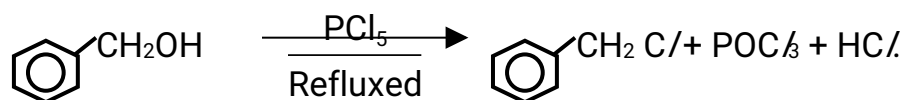
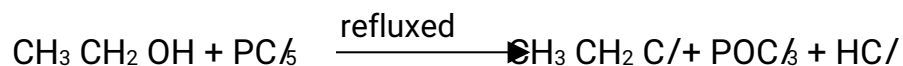
The tertiary alcohol gives an immediate cloudy solution of an alkyl halide. (0.5 minutes).

The secondary alcohol gives cloudy solution between 5 and 10 minutes.

The primary alcohol does not give a cloudy solution at room temperature.

**(b) Action of phosphorous halides:**

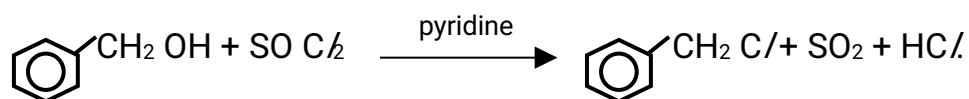
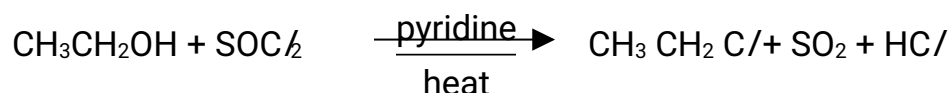
When alcohols are refluxed with phosphorous halide like phosphorous tri chloride or phosphorous penta chloride, the halogeno alkanes or alkyl halides are formed.



**(c) Action of thionyl chloride (SOCl<sub>2</sub>):**

When alcohols are refluxed in presence of thionyl chloride in the presence of an organic base pyridine to neutralize the mixture of toxic gases, an alkyl halide is formed.

This is a very convenient way of preparing alkyl halides.

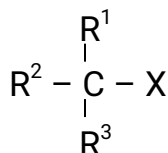


## REACTIONS OF ALKYL HALIDES

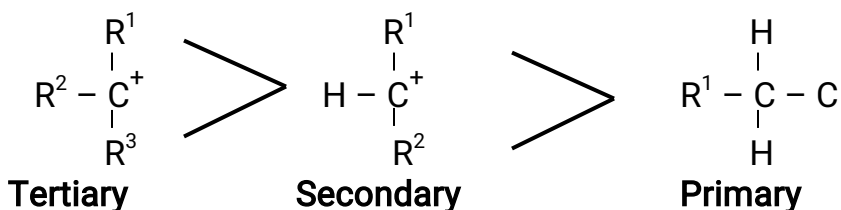
(i) **Primary Alkyl halides  $\text{RCH}_2\text{CH}_2 - \text{X}$ .**

(ii) Secondary alkyl halide.  $\text{R}-\text{CH}-\text{R}^2$   
 $\quad \quad \quad | \quad \quad |$   
 $\quad \quad \quad \text{H} \quad \quad \text{X}$

(iii) Tertiary alkyl halide:

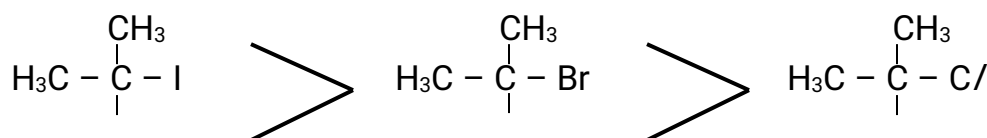
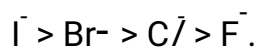


The three classes of alkyl halides given above react differently due to the stability of the carbo cation ion which is formed as an intermediate during the reaction. The stability is governed by the order that tertiary > secondary > primary.



Lower alkyl halides are gases at room temperature.

The boiling points of alkyl halides with the same number of carbon atoms is determined by the atomic size of the halogen which later is as a result of the Van de Waal's forces of attraction.

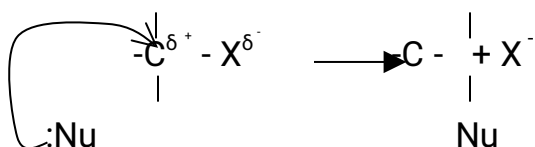




### CHEMICAL PROPERTIES:

Halogen atoms being more electronegative than carbon means that the C – X bond is highly polar due to the presence of partial induced charge where the halogen is partially negative and carbon is partially positive.

The polarity of the C – X bond makes alkyl halides to be very reactive to the nucleophile thus nucleophilic substitution.



The type of halogen present also determines the reaction.

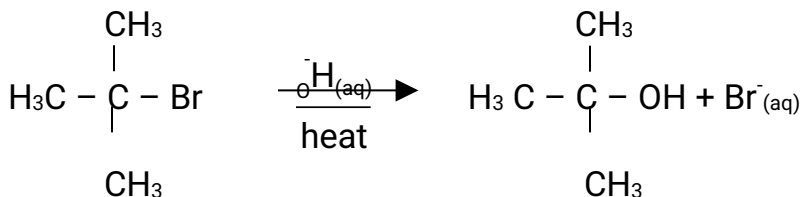
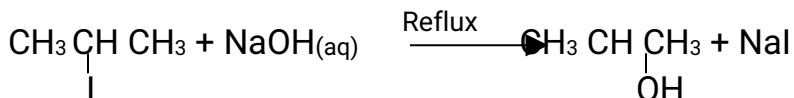
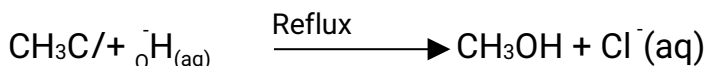
The electronegativity difference between the halogen together with the bond length determines overall the reactivity. Down the group of the halogens there is decrease in electronegativity and at the same time increase in the bond length.

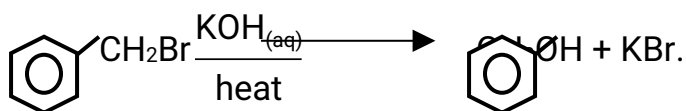
The reaction becomes faster with Iodo-alkanes than Bromo chloro fluoro alkanes because of low electronegativity in the iodides.

### NUCLEOPHILIC SUBSTITUTION REACTIONS

#### (i) Reaction with alkalis:

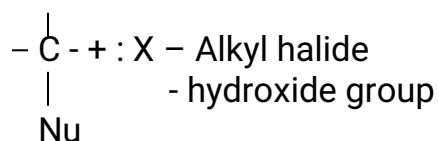
Alkalis react with alkyl halides to form alcohols when they are refluxed.



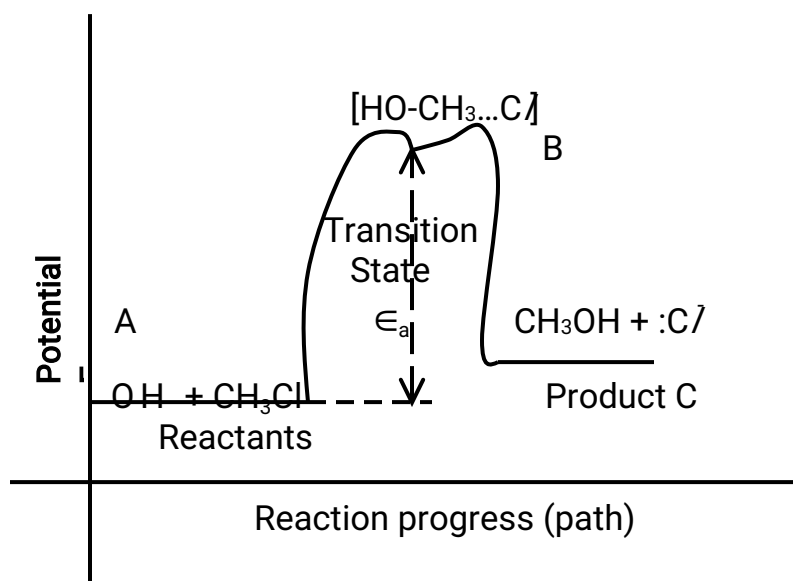
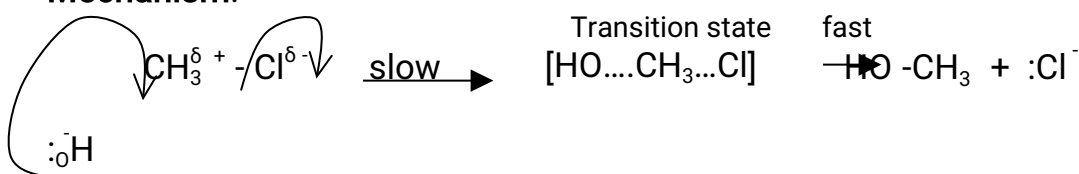


### Primary alkyl halides with aqueous alkali.

Primary alkyl halides follow nucleophilic substitution bimolecular ( $\text{S}_\text{N}2$ ) mechanism. Is a type of reaction where two molecules involve the rate determining step.



### Mechanism:



The nucleophile which is  $\text{OH}^-$  approaches the carbon atom carrying halogen from the opposite side.

This partially forms a nucleophile carbon bond and at the same time, a



carbon halogen bond partly breaks. This leads to the increase in potential energy from A to B.

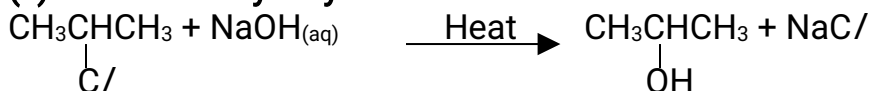
At B, there is partial halogen bond broken and partial nucleophile bond formed. This state is known as **activation complex** or **transition state**.

The energy level between the reactants and the activation complex is called the **activation energy**.

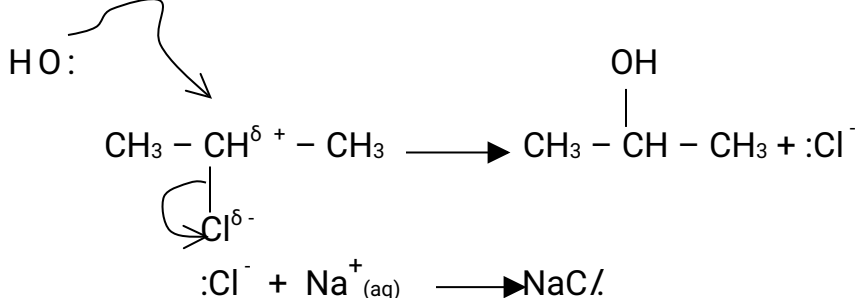
Later, the potential energy of the system increases when all the carbon halogen bonds are broken and the carbon nucleophile bonds are formed.

The energy of the system finally decreases until point C where the product of the alcohol is finally formed. This type of reaction is categorized as  $S_N2$  reactions mechanism because there are two molecules involved at the activated complex.

**(ii) Secondary alkyl halide**

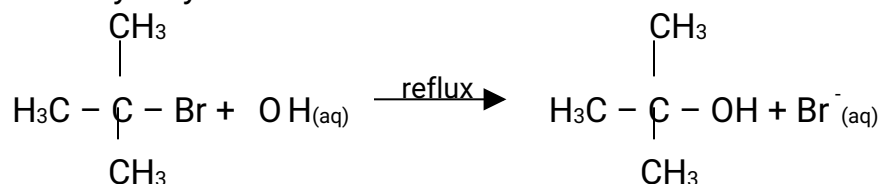


**Mechanism:**



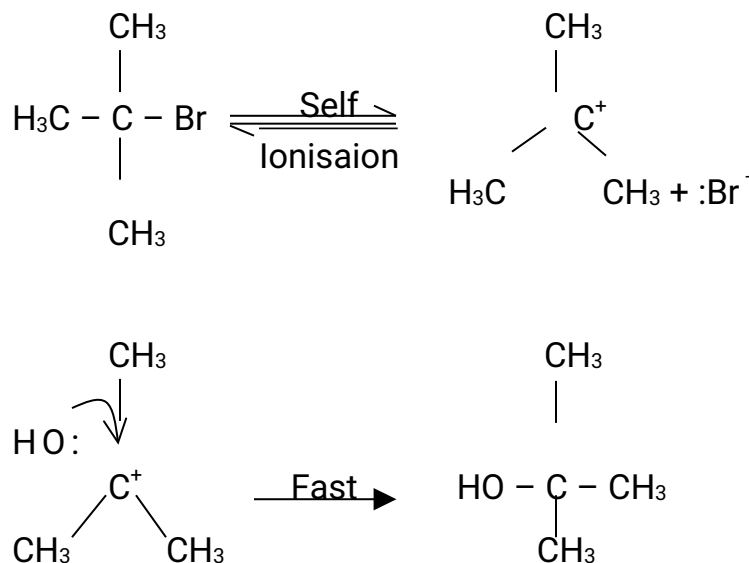
The mechanism followed by secondary alkyl halides is always between  $S_N1$  and  $S_N2$ .

**(iii) Tertiary alkyl halides**



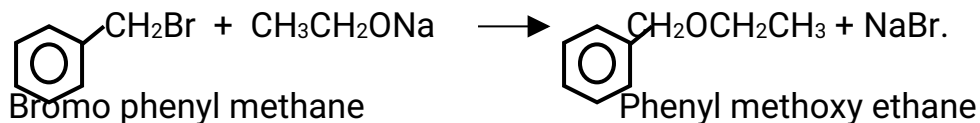
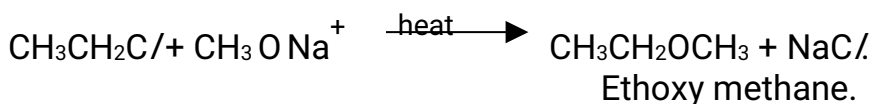
Tertiary alkyl halides follow  $\text{SN}_1$  mechanism. Here only the alkyl halide molecule is involved in the activated complex and thus its concentration alone determines the order of reaction.

**Mechanism:**

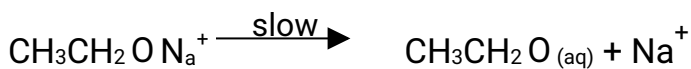


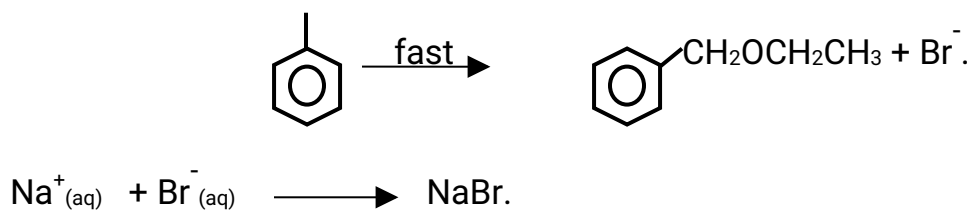
## REACTIONS WITH ALKA OXIDES (Na, K)

Alkyl halides react with sodium or potassium alkali oxides to form ethers under heat.



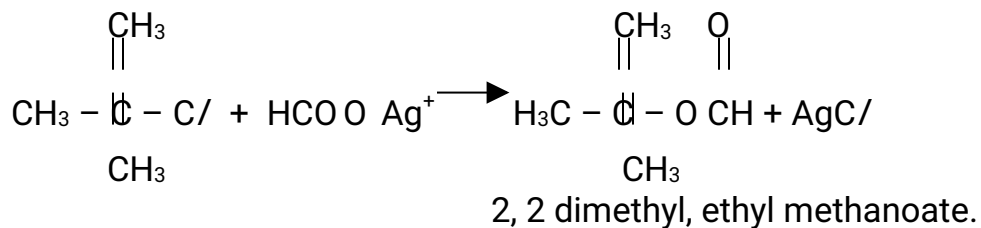
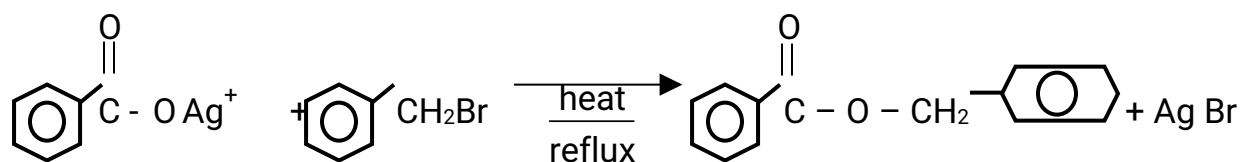
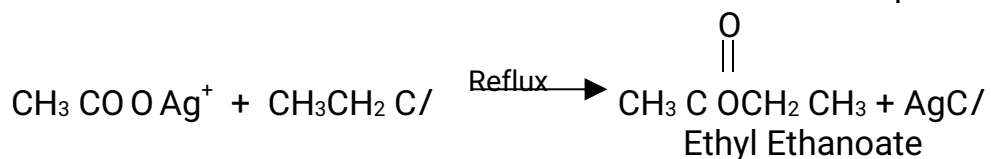
**Mechanism:**



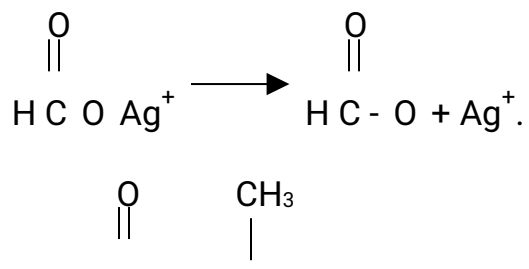
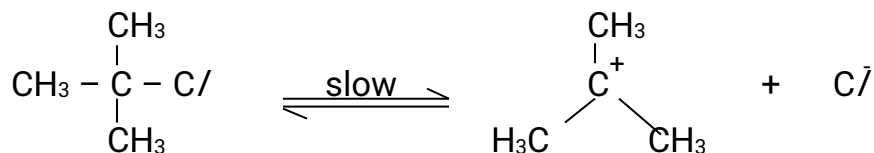


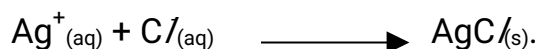
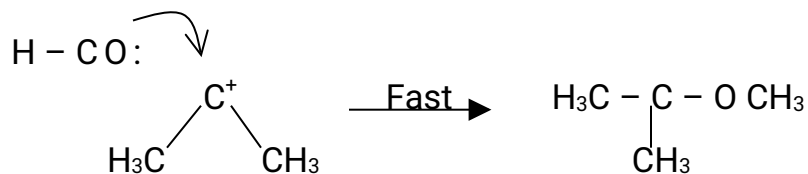
## REACTIONS WITH SILVER SALTS OF CARBOXYLIC ACIDS

Alkyl halides react with silver salts of carboxylic acids when refluxed to form esters. This reaction uses the alkanoate ion as the nucleophile.



### Mechanism:





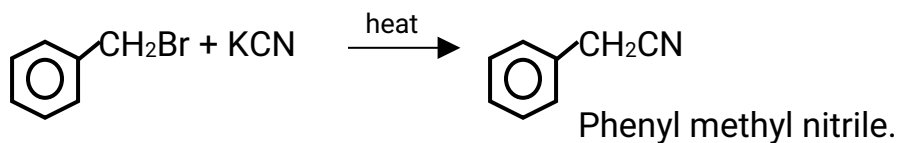
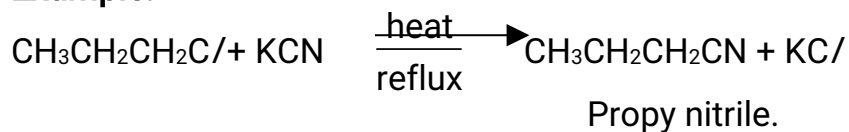
## REACTIONS WITH POTASSIUM CYANIDE (KCN)

Alkyl halides react with KCN in presence of an alcohol when refluxed to form alkyl nitriles.

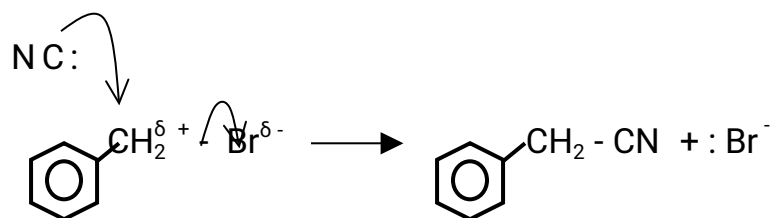
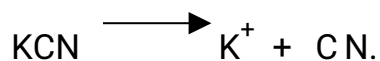
Condition (reflux) – alcohol present.

Nucleophile  $\text{C} \equiv \text{N}$  or  $\text{CN}^-$ .

**Example:**



**Mechanism:**



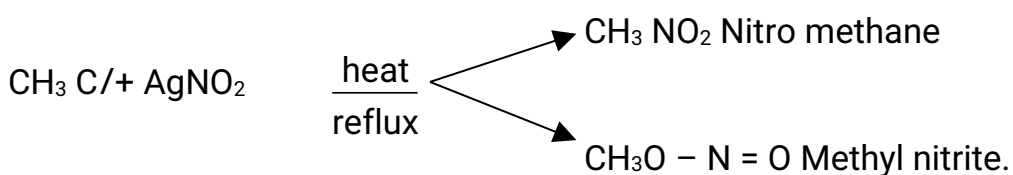
**Note:**

- (i) The above reaction is important inorganic synthesis, more especially increasing the carbon length (chain) by a single carbon.
- (ii) Silver cyanide can also be used instead of potassium cyanide.

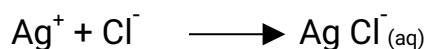
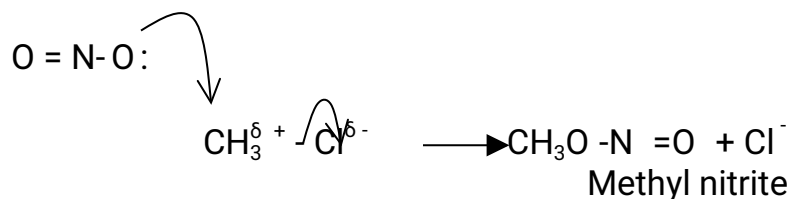
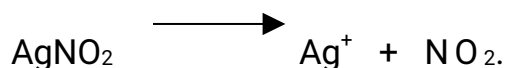
Convert  $\text{CH}_3\text{COOH}$  to  $\text{CH}_3\text{CH}_2\text{COOH}$ .

## REACTIONS WITH SILVER NITRITE ( $\text{AgNO}_2$ ).

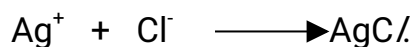
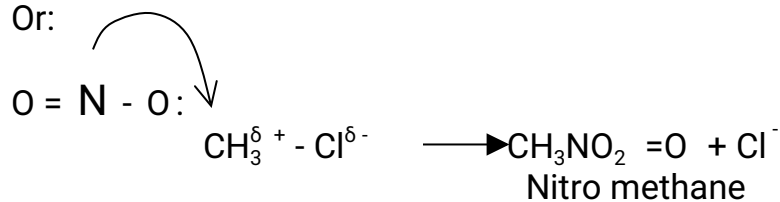
Alkyl halides react with  $\text{AgNO}_2$  when refluxed to form a mixture of nitro alkanes and alkyl nitrite. Such types of reactions are not important in synthesis since they yield a mixture of products.



### Mechanism:

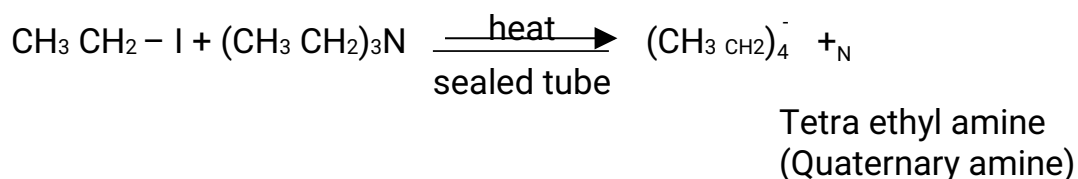
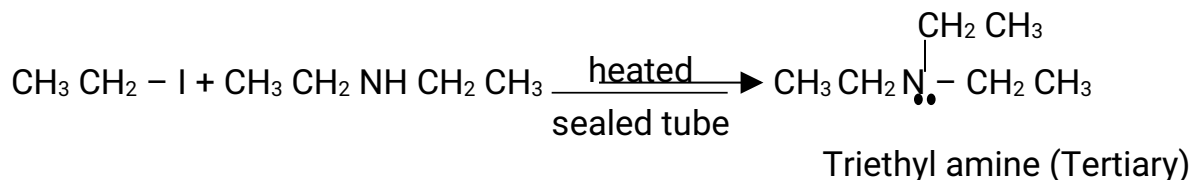
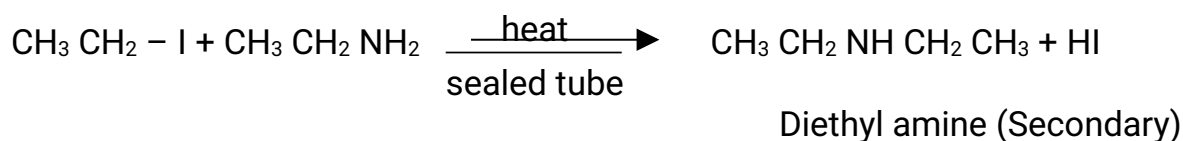
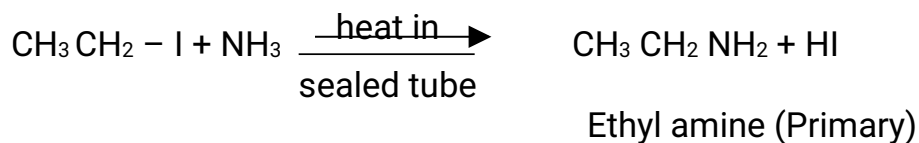


Or:



## REACTIONS WITH AMMONIA AND AMINES

Alkyl halides react with conc.  $\text{NH}_3$  to form a mixture of amines when heated in a sealed tube. A mixture of amines are produced because a product at one stage become a nucleophile for the next stage. The nucleophile is ammonia and the amines.



## ELIMINATION REACTIONS OF ALKYL HALIDES

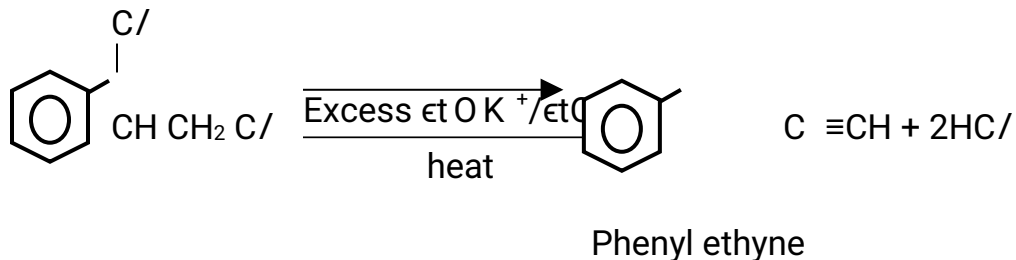
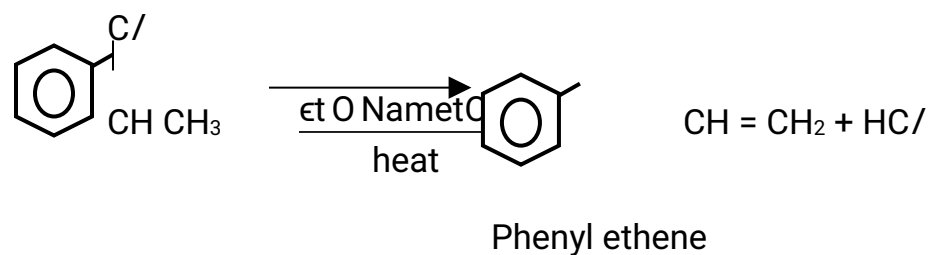
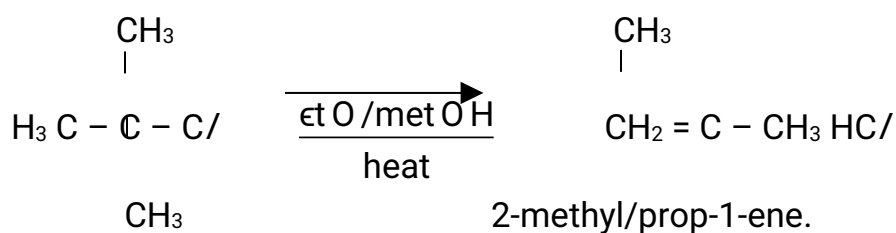
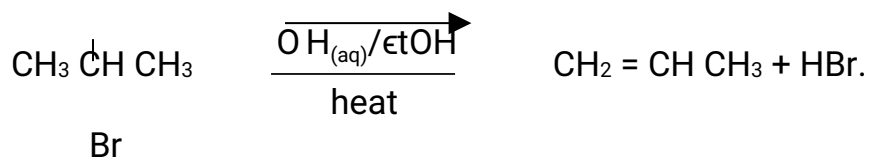
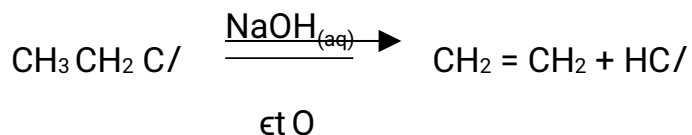
These are reactions that will result into the formation of unsaturated compound with elimination of a water molecule. Alkyl halides when refluxed or heated with an alkali (KOH, NaOH) in the presence of an alcohol, they form alkene. The reaction can also be effected by using a strong base which is an alkoxide.

Strong base (Alkoxide)

Ethoxide /  $\text{EtO}^-$

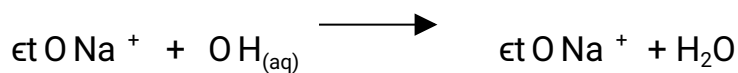
Methoxide /  $\text{MeO}^-$

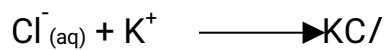
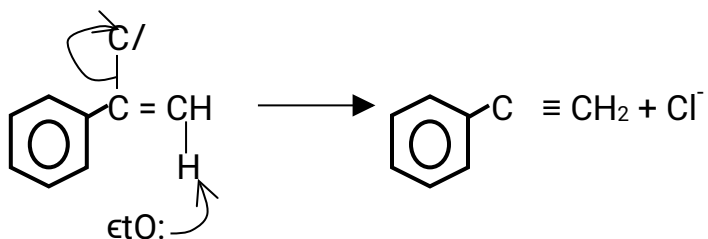
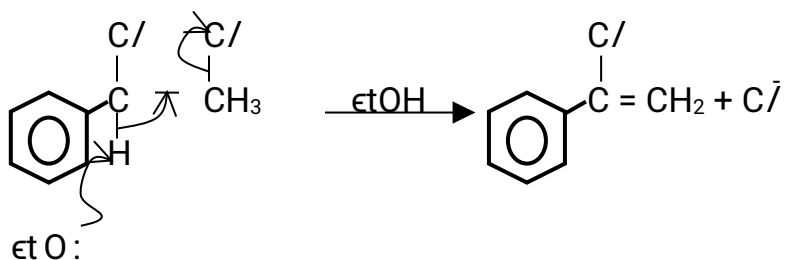
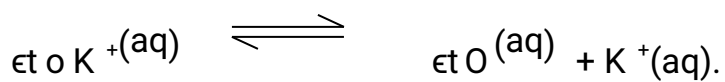
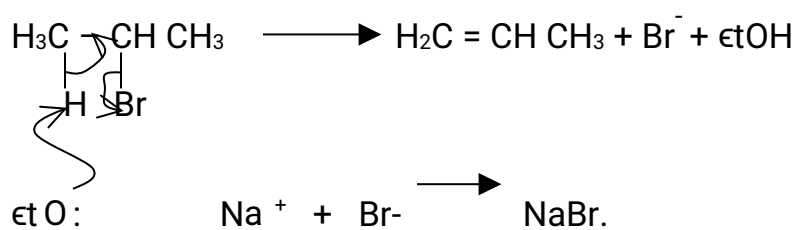
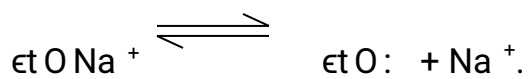
## DEHYDROHALOGENATION



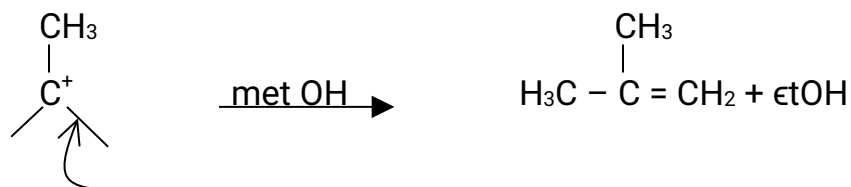
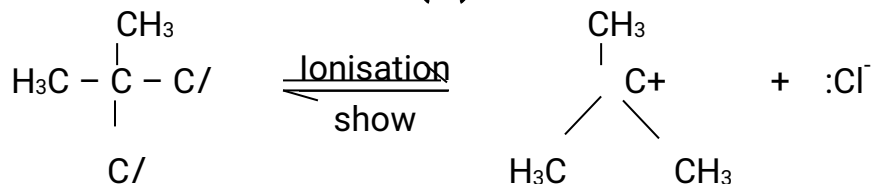
### Mechanisms:

Elimination bimolecular ( $\epsilon_2$ ).

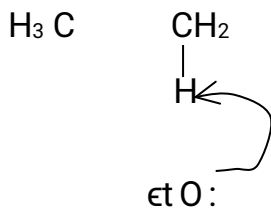




### Elimination unimolecular (E<sub>1</sub>)



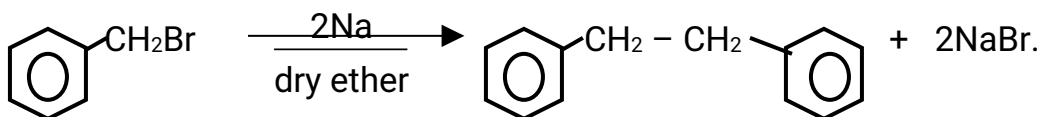
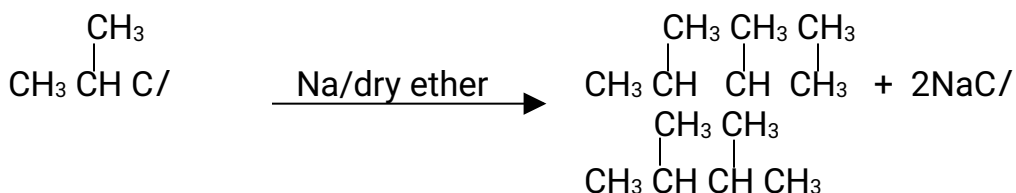
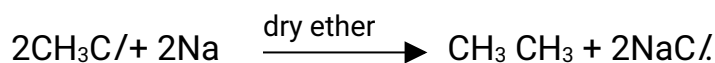
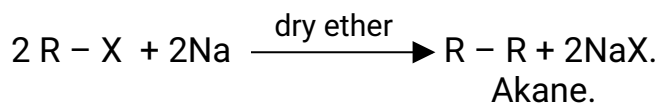




**Research:** Discuss all the reactions of chloro ethane with NaOH. (25mks)

### Wurtz reaction:

When alkyl halides are reacted with Na metal in the presence of ether, alkanes are formed. The product will have an increased carbon atom by 2 i.e. the carbon number doubles.



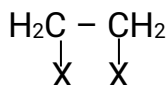
## DIHALIDES COMPOUNDS

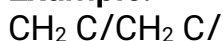
There are compounds that have got two halogen atoms within the same carbon chain.

There are two types of dihalides;

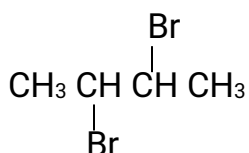
### (i) Vicinal dihalide

This contains two halogen atoms located on any adjacent carbon atom i.e.

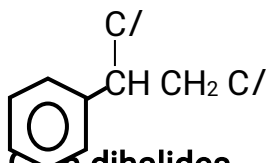


**Example:**

1, 2-dichloro ethane.



2, 3-dibromo butane.



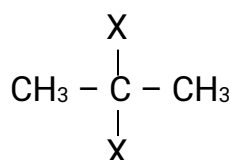
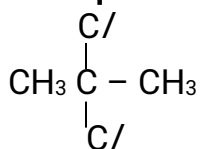
1, 2-dichlorophenyl ethane.

(ii)

**Gem dihalides**

These contain the two halogen atoms located on the same carbon atoms.

i.e.

**Example:**

2, 2 dichloro propane

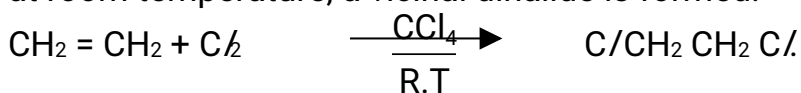


1, 1-dichloro ethane.

**PREPARATION OF DIHALIDES**

Dihalides are prepared from the following reactions:

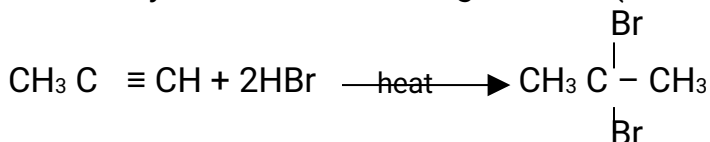
(i)

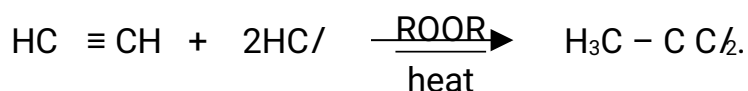
**From Alkenes**When alkenes are reacted with halogens in presence of  $\text{CCl}_4$  (organic base) at room temperature, a vicinal dihalide is formed.

(ii)

From alkynes and halogen acids

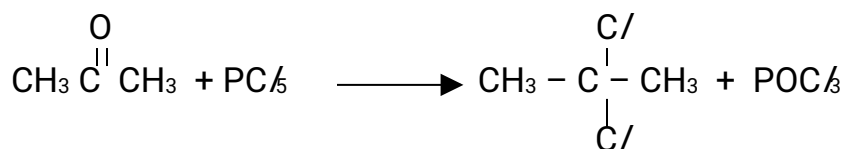
When alkynes react with halogen acids (excess), a gem dihalide is formed.





(iii) **From carbonyl compounds (Aldehydes and Ketones).**

When a carbonyl compound is reacted with phosphorous pentachloride, a dihalide is formed. Using aldehydes, will produce a gem dihalide and ketones also produce gem dihalide.

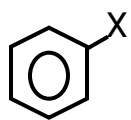


## POLYHALIDES

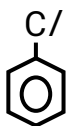
These are halogen compounds with more than two halogen atoms on the same carbon chain.

## AROMATIC HALIDES

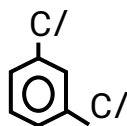
These are compounds with one or more halogen atoms directly attached to the aromatic ring.



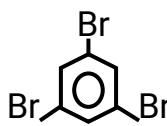
or  $\text{Ar} - \text{X}$



Chloro benzene

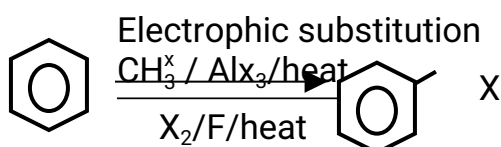


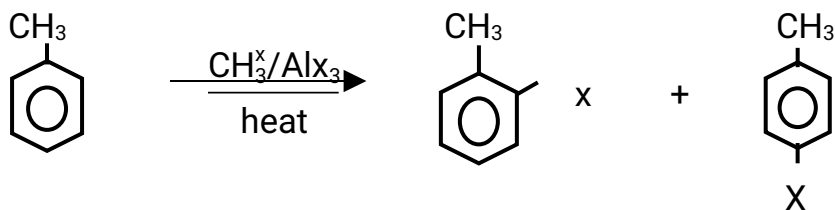
1, 3-dichloro benzene



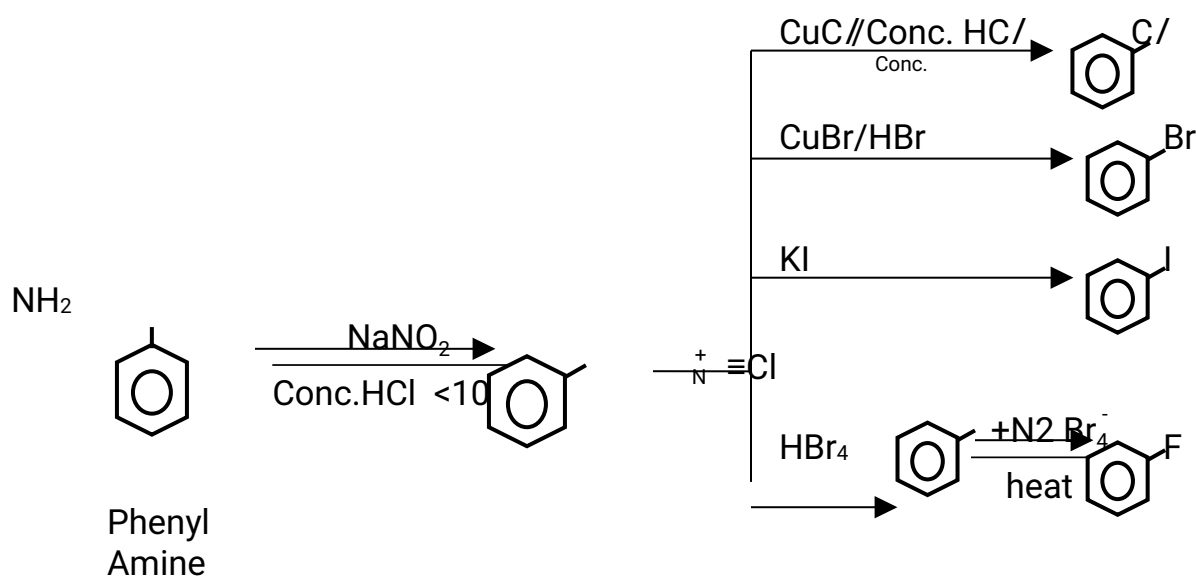
1, 3, 5-tribromo benzene.

## PREPARATION





### From benzene diazonium salt



## PHYSICAL PROPERTIES OF AROMATICS

They are colourless liquids or solids with characteristic flame.

They are insoluble in  $\text{H}_2\text{O}$  but soluble in organic solvents like  $\text{CCl}_4$ .

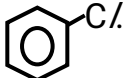
### NOTE:

They do not easily undergo nucleophilic substitution unlike alkyl halide.

Distinguish between c1ccccc1C# and c1ccccc1CC#

Reagent: Hot aqueous  $\text{NaOH}$  in dil  $\text{HNO}_3$  and  $\text{AgNO}_3$ .

Observations: - A white precipitate forms with c1ccccc1CC#

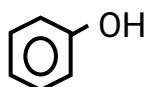
- No observable change with 

## ALCOHOLS AND PHENOLS:

These are compounds that contain hydroxyl group. Therefore the functional group is hydroxyl group. The difference between alcohol or alkanols and phenol is that the hydroxyl group is directly attached to the aromatic ring in phenols.

Alkanol  
 $R - OH$

Phenol  
 $Ar - OH$



## ALKANOLS OR ALCOHOLS

Alcohols are organic compounds derived from hydro carbons but where one or more hydrogens is/are replaced by hydroxyl group. The general formula is  $R - OH$  where R is an alkyl group or simply represented as  $C_nH_{2n+1}OH$ . Where n = simple number.

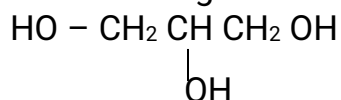
They are also called alkanols simply because a hydroxyl group replaces a hydrogen.

## TYPES OF ALCOHOLS

There are basically three types of alcohols:

- (i) Monohydric alcohol  
This is one that has got one hydroxyl group.
- (ii) Dihydric alcohol  
This is one that has got two hydroxyl groups.
- (iii) Polyhydric alcohol

These have got more than two hydroxyl groups e.g.



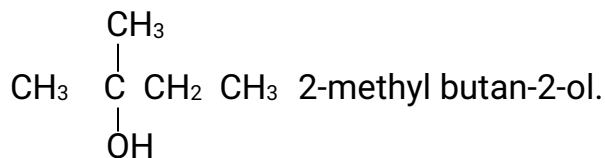
## NOMENCLATURE

Alcohols are named as alkanols according to the IUPAC. This is done by replacing the last "e" in alkane name with suffix "ol" (functional group name).

The position of the functional group has to be indicated just before the suffix "ol".

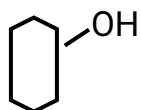
If the stem name has got a vowel, then a consonant letter must be added just before the position of the functional group.

E.g.  $\text{CH}_3 \text{CH}_2 \text{OH}$  Ethano-1-ol

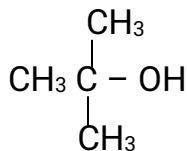


$\text{HO} - \text{CH}_2 \text{CH}_2 \text{OH}$  Ethane-1, 2 – diol

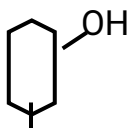
$\text{HO} - \text{CH}_2 \text{CH}(\text{OH}) \text{CH}_2 \text{OH}$  Propane – 1, 2, 3 – triol.



Cyclohexan-1-ol

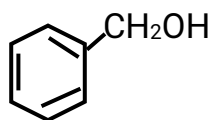


2-methyl propan-2-ol



Cyclohexan-1, 3-diol  
Xyclohexane-1, 3-diol

OH



Phenyl methan-1-ol.

**NOTE:**

Loot at isomerism in alcohols.

Types – structural:

(i) Chain

(ii) Position

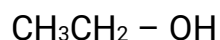
- Functional (alcohol and others are isomeric).

## CLASSES OF ALCOHOLS

Monohydric alcohols are classified into three classes.

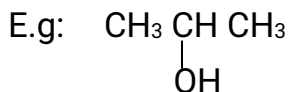
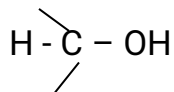
(i) **Primary alcohol**

This has one alkyl group attached to the carbon atom carrying the –OH group. E.g.



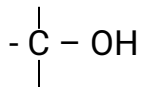
(ii) **Secondary alcohol**

This has got two alkyl groups attached to the carbon atom carrying the OH group.

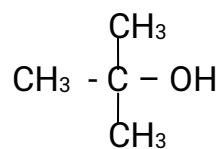


(iii) **Tertiary alcohol**

This has three alkyl groups attached to the carbon atom carrying the OH group.



E.g.

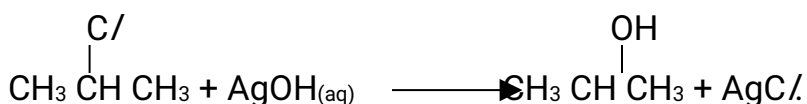
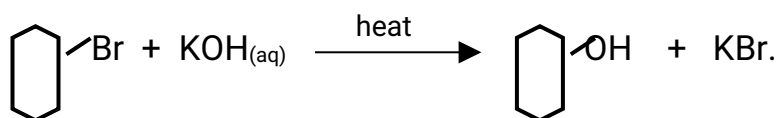
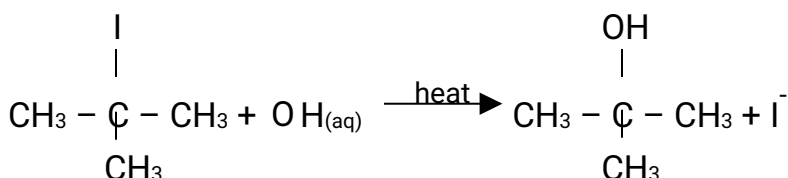
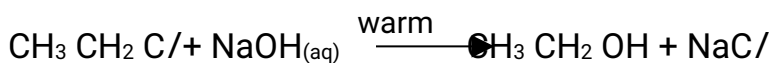


## METHODS OF PREPARATION

### (i) From alkyl halide (SN reaction)

When alkyl halides are refluxed with aqueous alkali (KOH or NaOH) or with moist silver oxide, alcohols are formed.

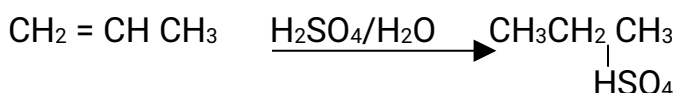
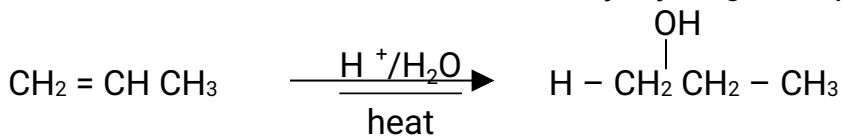
This is a hydrolysis reaction that occurs by nucleophilic substitution reaction.



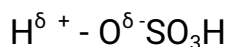
### (ii) From alkenes

When alkenes are reacted with dilute mineral acid water and heated, an alcohol is formed. They usually used acid is  $\text{H}_2\text{SO}_4$  acid.

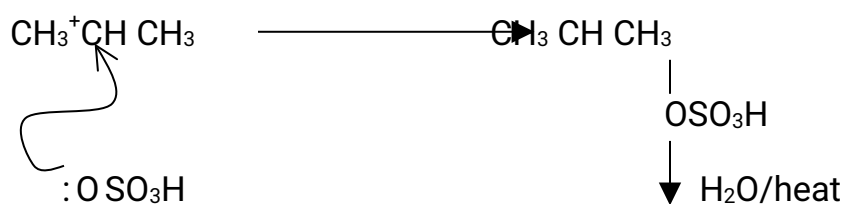
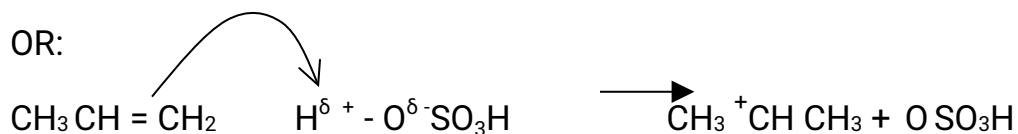
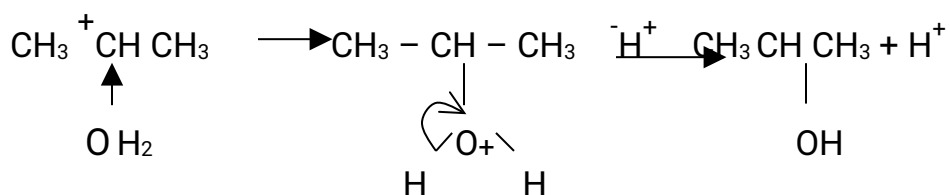
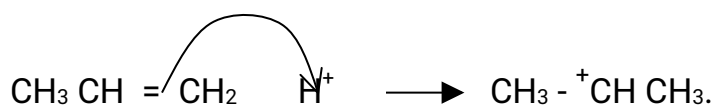
If the mixture is not heated, then an alkyl hydrogen sulphate is formed.



**Mechanism:**

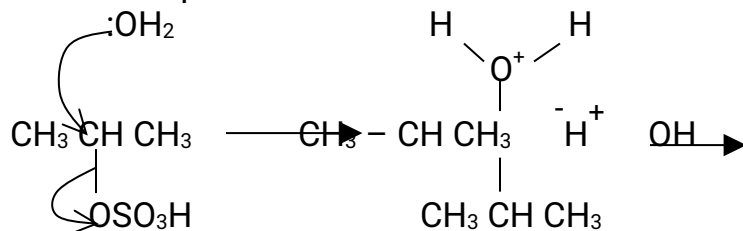




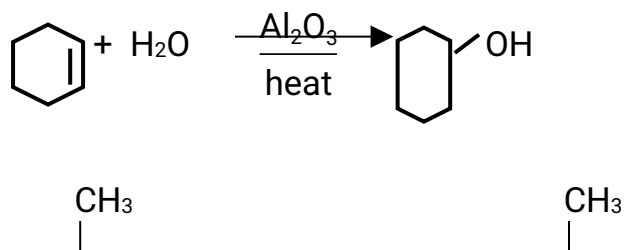
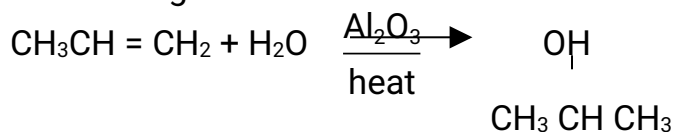


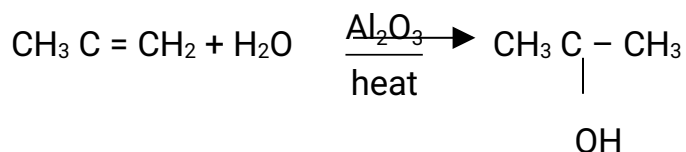
**NOTE:**

H<sub>2</sub>O acts as a nucleophile that will react with the alkyl hydrogen sulphate in the last step to form an alcohol. i.e.



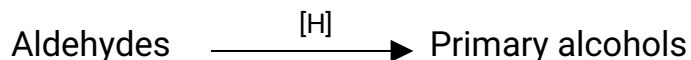
When alkenes are reacted with H<sub>2</sub>O in presence of Al<sub>2</sub>O<sub>3</sub>, an alcohol is formed. E.g.





(iii) **From carbonyl compounds (Aldehydes and ketones)**

Carbonyl compounds are reduced in the presence of a suitable reducing agent to alcohols.



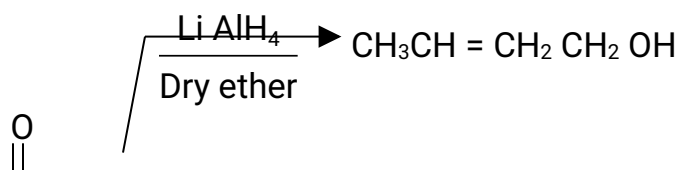
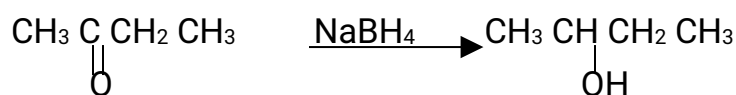
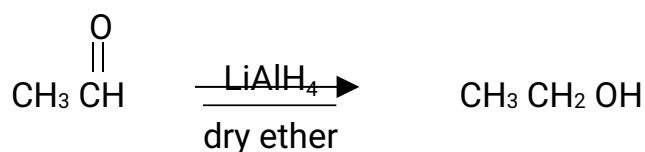
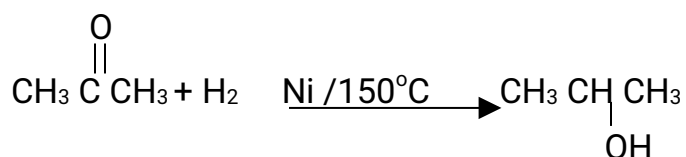
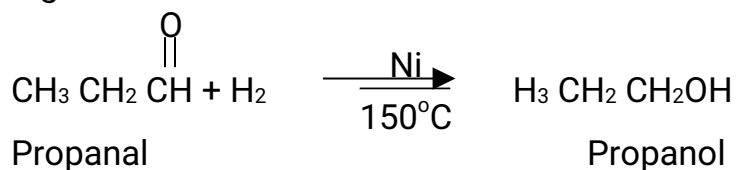
**Reducing agents normally used:**

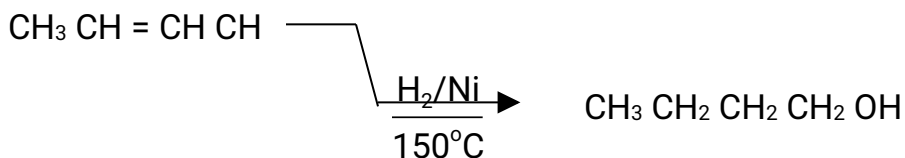
Hydrogen in presence of catalyst, Ni/150, Pt/pd, r.t.p.

Using Li Al H<sub>4</sub> (Lithium, Aluminium, Tetra hydride) in presence of dry ether.

Sodium boron tetrahydride (NaBH<sub>4</sub>)

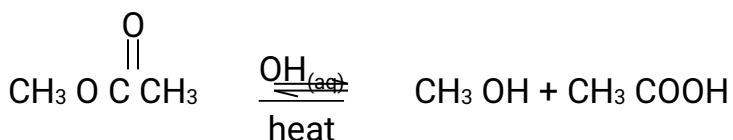
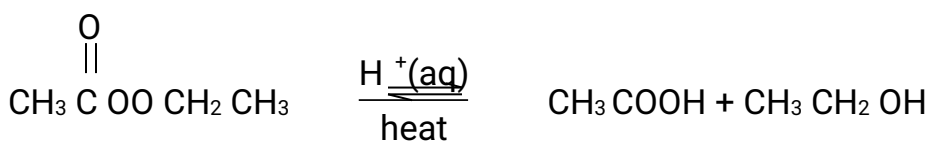
E.g.





(iv) **Hydrolysis of esters**

Esters are hydrolyzed in presence of mineral acids or alkalis to form corresponding alcohols and carboxylic acids. This reaction is not very useful for synthesizing alcohols because of the mixture of alcohol and acid.

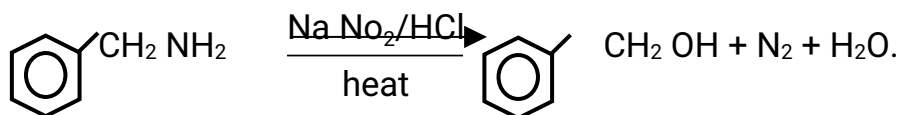
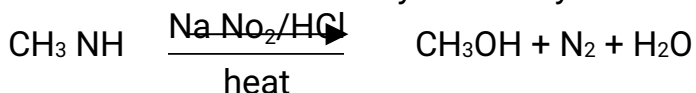


Ethyl ethanoate

**Mechanism:** (look for acid/base hydrolysis of esters).

(v) **From primary amines**

They react with  $\text{HNO}_2$  acid which is generated "insitu" by reacting sodium nitrite and conc.  $\text{HCl}$  to form alcohols. This reaction is only for primary amines and not secondary or tertiary.



(vi) **From fermentation of carbohydrates.**

## PROPERTIES OF ALCOHOLS

### PHYSICAL PROPERTIES

Lower members are liquids, higher members are solids at r.t.p with x-tic smell.

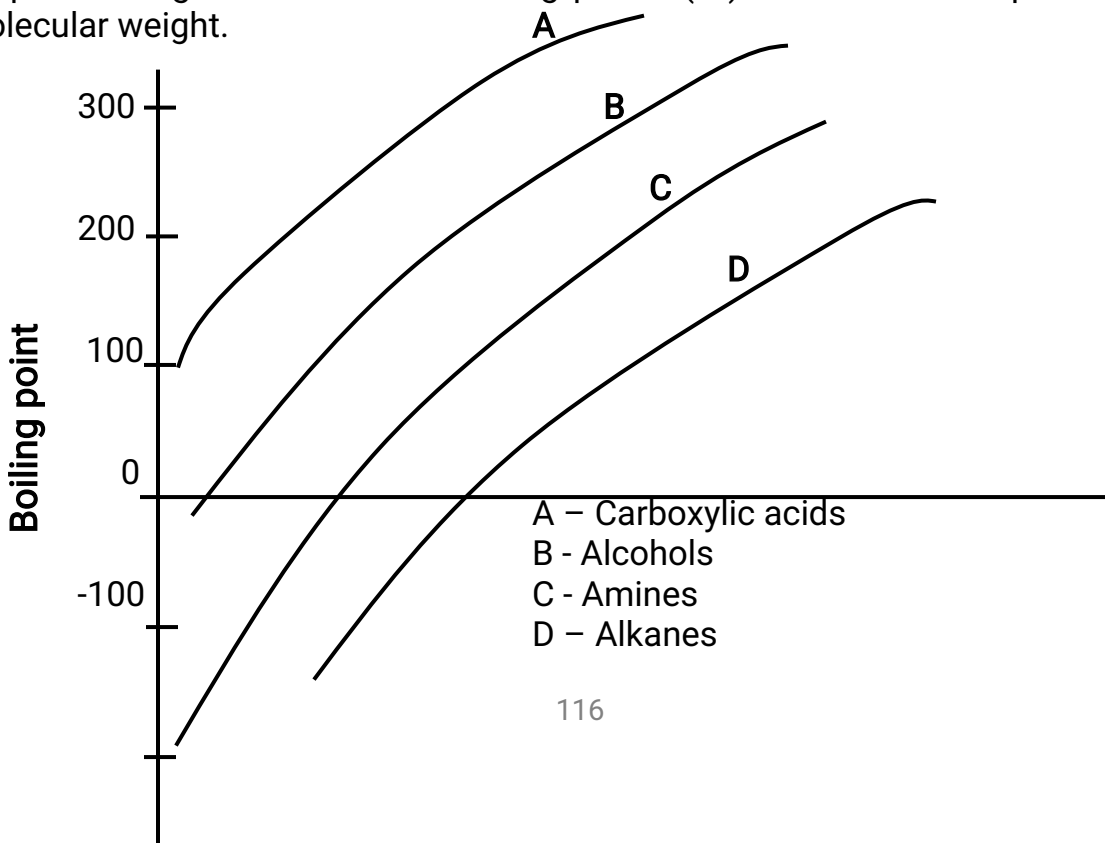
The lower members are very soluble in H<sub>2</sub>O but the solubility decreases with the molecular mass.

Alcohols have got relatively higher boiling points compared to similar hydrocarbons of approximately the same molecular mass.

#### Examples:

	Molecular weight	Boiling points
Ethane	30	-42
Methanol	32	46
Butane	58	-0.5
Propan-ol	60	98

Graph showing the variation of boiling points (°C) of different compounds with molecular weight.



## Molecular weight

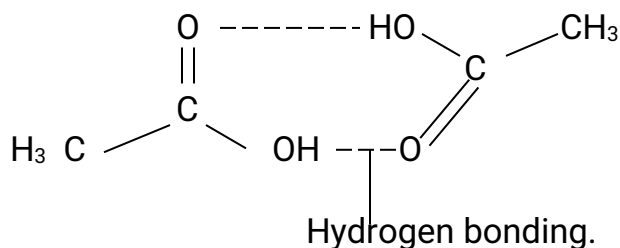
### Note:

The boiling points increase generally with increase in molecular weight. Increase in molecular weight increases the Van Der Waal's forces of attraction. This makes the compound stronger hence high boiling points.

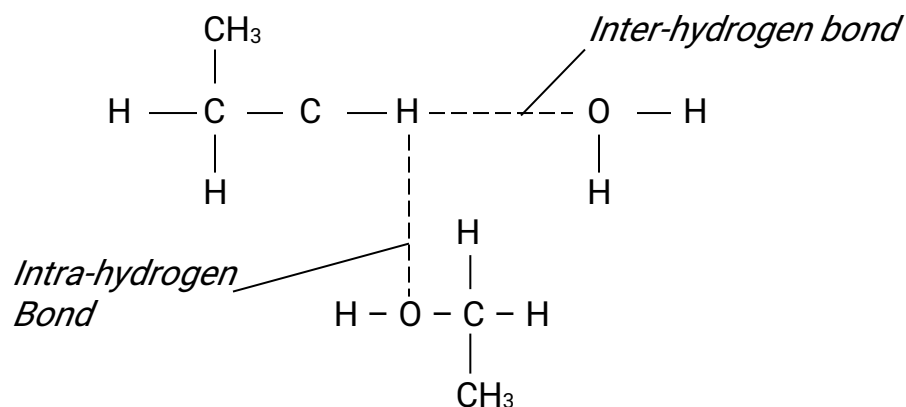
### Explanation:

The boiling points of the compounds above on the graph are determined by hydrogen bonding.

Carboxylic acids have got the highest boiling points due to extensive hydrogen bonds which makes molecules to dimerise in a liquid hence difficult to separate them during boiling so that they escape to vapour.



Alcohols' boiling point is higher than that of alkanes but lower than that of carboxylic acid because of inter and intramolecular hydrogen bonding. Molecules will interact in liquids and aqueous phases and high boiling point.



Amines boil at a lower temperature than alcohols of the corresponding molecular weight because nitrogen is less electronegative than oxygen so the hydrogen

bonds formed in amines are weaker than in alcohols thus amines boil at a lower temperature.

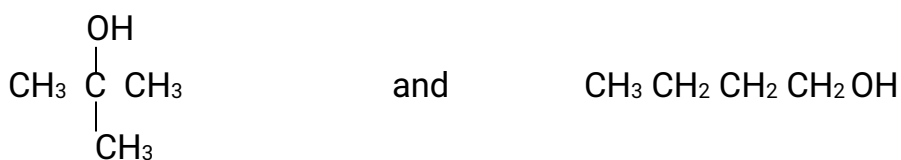
Alkanes have the least boiling point because of the absence of hydrogen with only Van der Waal's forces of attraction which are weaker and easily broken during boiling.

**In summary:**

The solubility in water and the boiling points of alcohols are due to hydrogen bonding.

**Questions:**

1. Methyl amine (mm = 31) boils at  $-6.3^{\circ}\text{C}$  while methanol (mm = 32) boils at  $46^{\circ}\text{C}$ . Explain.
2. 2-methyl propan-2-ol boils at a lower temperature than butan-1-ol yet they all have the same molecular weight.



2 methyl propan-2-ol is highly branched giving it a spherical shape which decreases on the size and weakens the Van der Waal's forces of attraction existing between the molecules.

Butan-1-ol is a straight chain molecule which gives it an extended structure resulting in relatively stronger Van der Waal's forces.

## CHEMICAL PROPERTIES OF ALCOHOLS

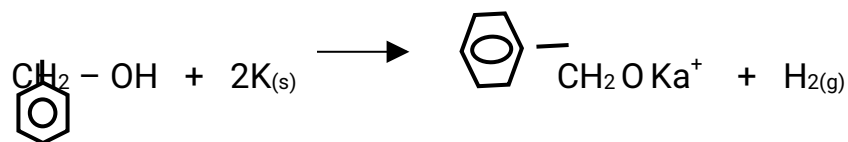
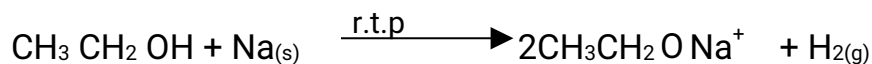
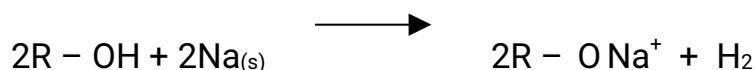
Alcohols consists of three major bonds that are involved in a chemical reaction.

- (i) Functional group – OH group.  
Where the O – H bond is involved in the reaction.
- (ii) Oxygen-carbon bond which is highly polarized.
- (iii) Alkyl groups – R group.

### REACTION INVOLVING CLEAVAGE OF THE O – H BOND

#### Reactions with electropositive metals:

When an alcohol is reacted with a metal like sodium or potassium, an alkoxide and  $H_2$  gas are formed.

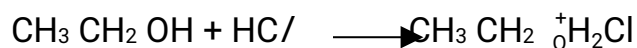
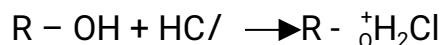


#### Observation:

Effervescence of a colourless gas that burns with a pop sound.

#### Reactions with mineral acids:

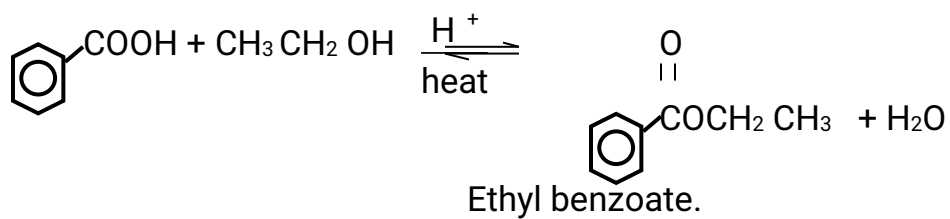
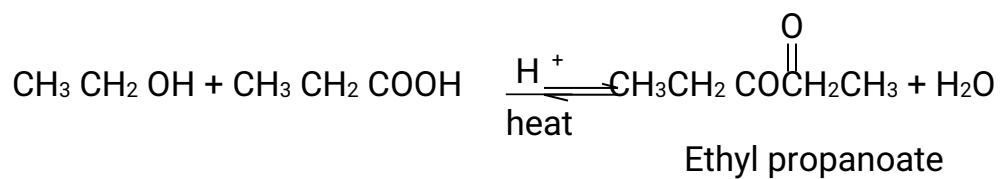
When an alcohol is reacted with a mineral acid, a salt is formed.



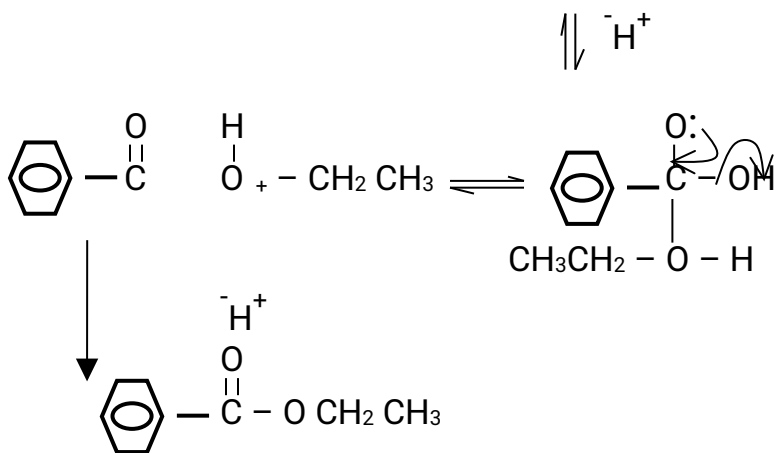
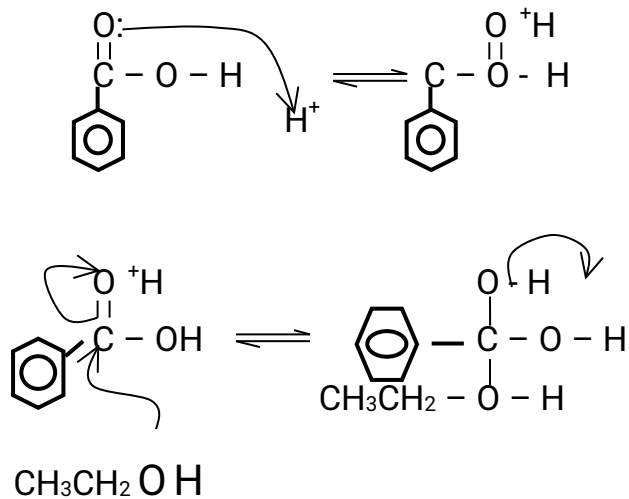
From the two reactions above, alcohols are regarded as amphoteric because they react with acids and bases to show both acidic and basic.

#### Esterification

Alcohols react with carboxylic acids in presence of mineral acids to form esters. This reaction is known as esterification.

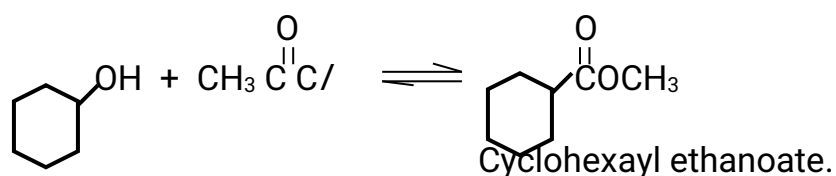
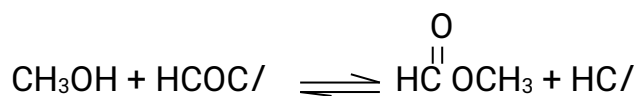
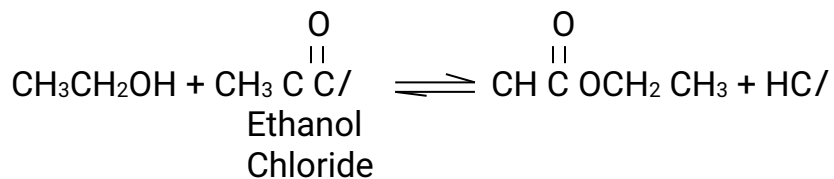


**Mechanism:**

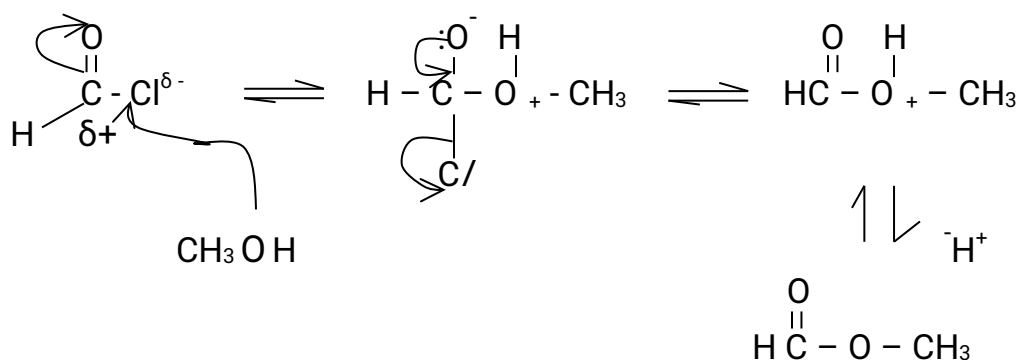




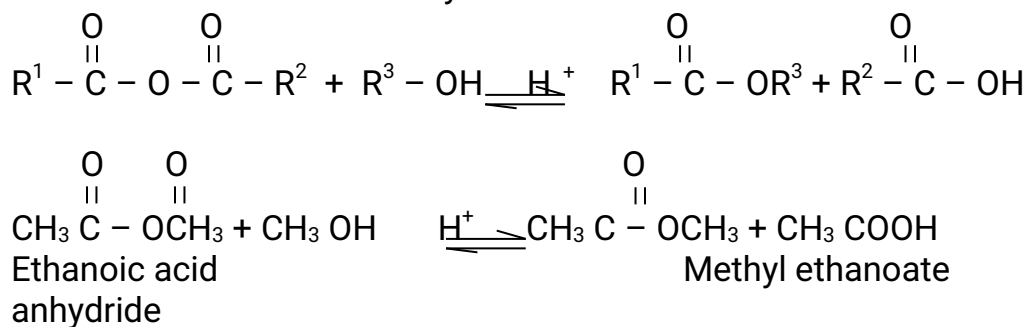
Alcohols react with acid chlorides to form esters.



**Mechanism:**

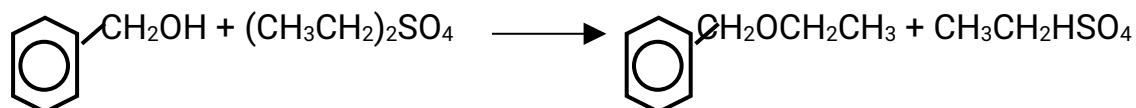
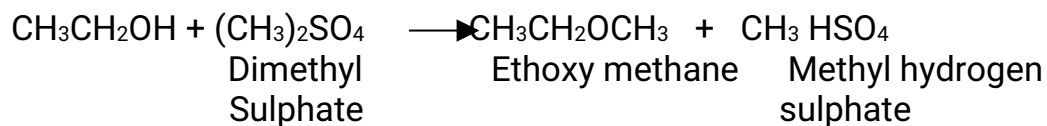


Alcohols react with acid anhydrides to form esters.



### Alkylation of alcohols

Alcohols react with dialkyl sulphates to form ethers and alkyl derivatives. This reaction involves replacement of hydrogen in the alcohol with alkyl group.

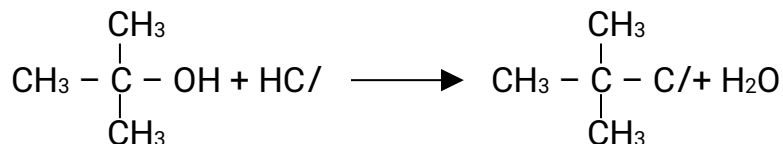
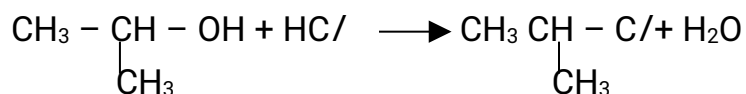
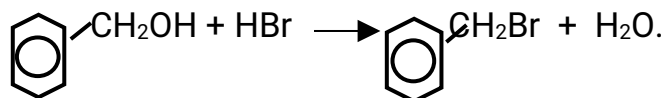
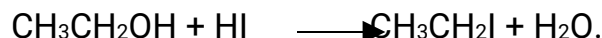
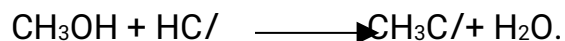


### REACTIONS INVOLVING CLEAVAGE OF A CARBON OXYGEN BOND

In these reactions, we are removing both oxygen and hydrogen atoms from the alcohol.

#### Reactions with halogen acids (Hx)

Alcohols react with halogen acids to form alkyl halides.

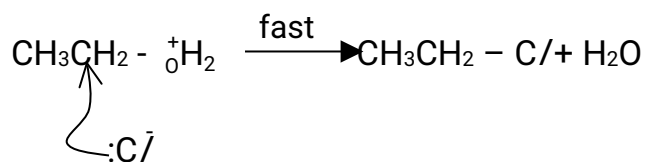
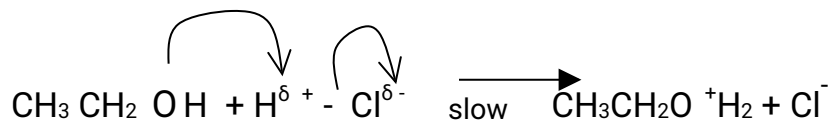


#### Mechanism:

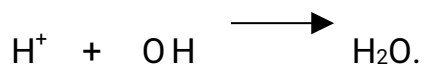
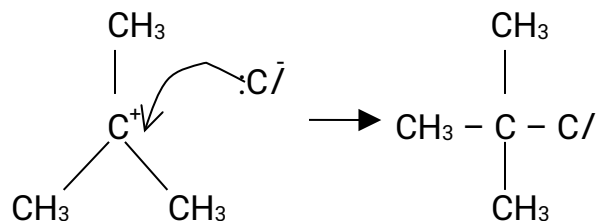
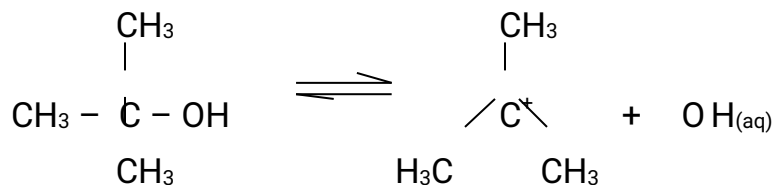
Primary and secondary alcohols follow  $SN_2$  mechanism because of the fairly

unstable carbon cation ion, while the tertiary alcohols follow  $\text{SN}_1$  because of the stability of the carbon cation ion formed.

**$\text{SN}_2$ :**



**$\text{SN}_1$ :**



## PRACTICAL IMPORTANCE OF THE REACTION

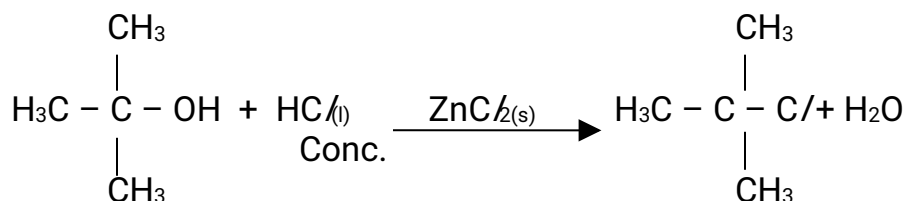
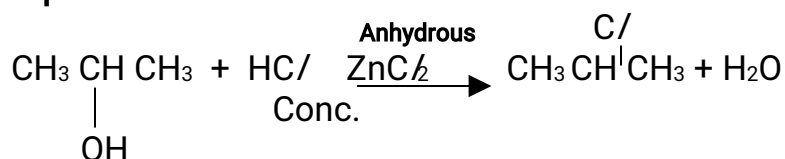
The practical importance of the above reaction is to distinguish between the three classes of alcohol.

**Reagent:** Anhydrous zinc chloride in concentrated hydrochloric acid.

**Observation:**

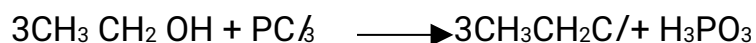
- An immediate cloudy solution at room temperature forms with a tertiary alcohol.
- Within 5 minutes at room temperature, a cloudy solution is formed with a secondary alcohol.
- No cloudy solution forms at room temperature with primary alcohol.

**Equations:**



### REACTIONS WITH PHOSPHORUS HALIDES

Alcohols react with  $\text{PX}_5$  and  $\text{PX}_3$  to form alkyl halides where X is a halogen.



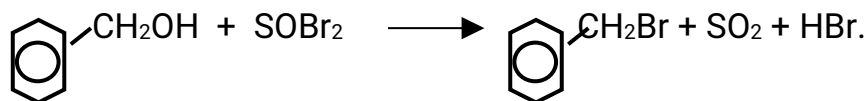
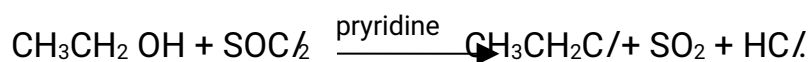
Others:  $\text{PI}_3/\text{PI}_5$   
 $\text{PBr}_3/\text{PBr}_5$ .

**NOTE:**

Evolution of  $\text{HCl}$  or fuming when  $\text{PCl}_5$  is added to a compound suggests the presence of OH group in that compound.

### REACTION WITH THINLY CHLORINE

Alcohols react with  $\text{SOCl}_2$  or  $\text{SOBr}_2$  to form alkyl halides. An organic base like pyridine must be included to neutralize toxic and poisonous gases liberated.



**Mechanism:**

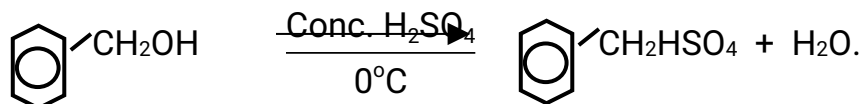
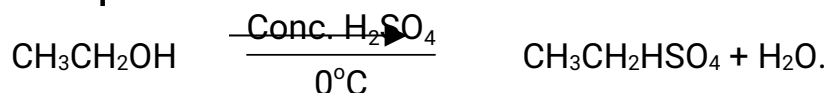
## REACTION WITH H<sub>2</sub>SO<sub>4</sub> ACID

Alcohols react with H<sub>2</sub>SO<sub>4</sub> giving different products depending on the conditions of reaction.

**Conditions:**

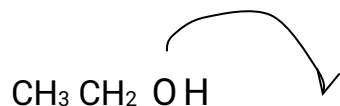
- A. (i) Conc. H<sub>2</sub>SO<sub>4</sub> Products got are substituted products.  
 (ii) Excess alcohol  
 (iii) Low temperatures.  
 Substituted products.  
 Alky hydrogen sulphate – 0°C.  
 Ether – warm (140°C)
- B. (i) Conc. H<sub>2</sub>SO<sub>4</sub>.  
 (ii) High temperatures (heat) - Elimination product.  
 (iii) Limited alcohol.

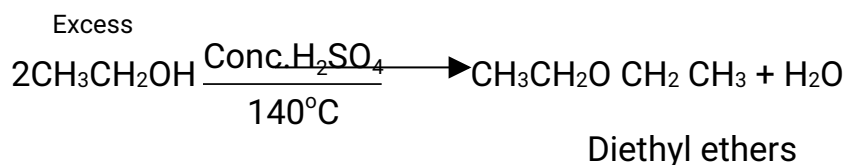
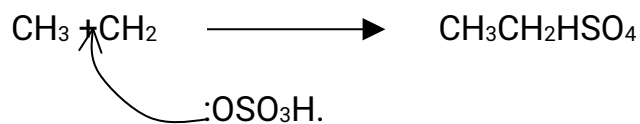
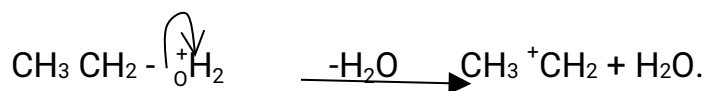
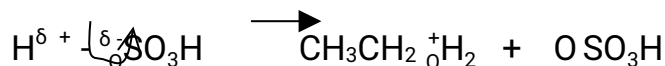
**Example:**



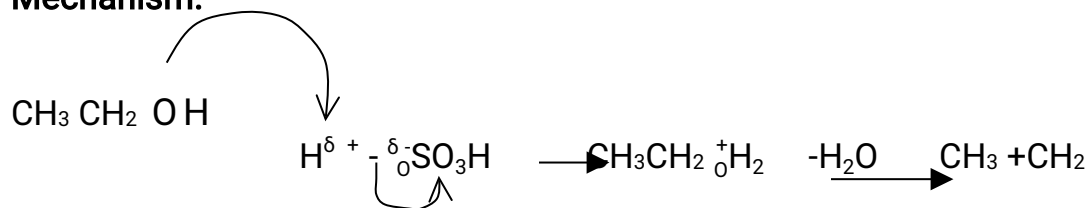
Phenyl methyl  
Hydrogen sulphate.

**Mechanism:**

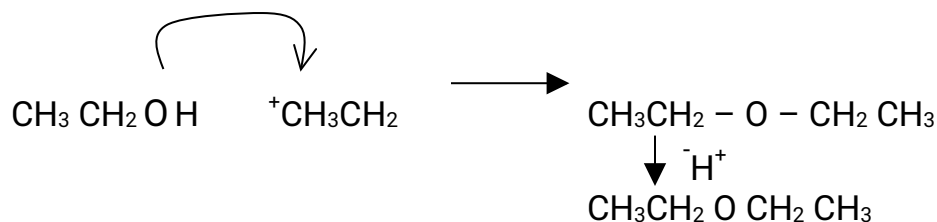




**Mechanism:**



$[\text{CH}_3\text{CH}_2\text{OH}] \gg [\text{H}-\text{O}-\text{S}(\text{OH})_2]$  therefore  $\text{CH}_3\text{CH}_2\text{OH}$  is a better nucleophile.



Alcohols are dehydrated when heated with conc.  $\text{H}_2\text{SO}_4$  and  $\text{H}_3\text{PO}_4$  acid to form alkenes in a liquid phase or when the alcohol is passed over heated  $\text{Al}_2\text{O}_3$  at  $300^\circ$  in a vapour phase.

Water molecule is eliminated and therefore this type of reaction is known as elimination or dehydration.

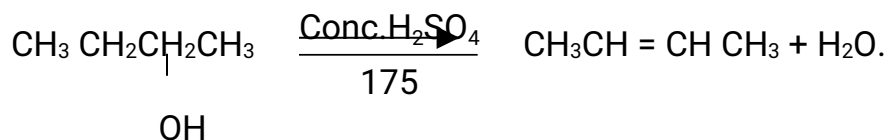
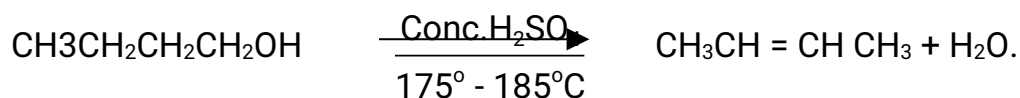
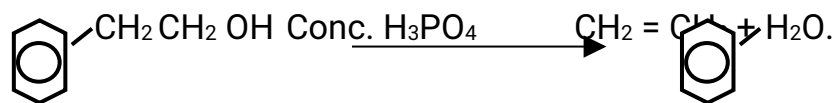
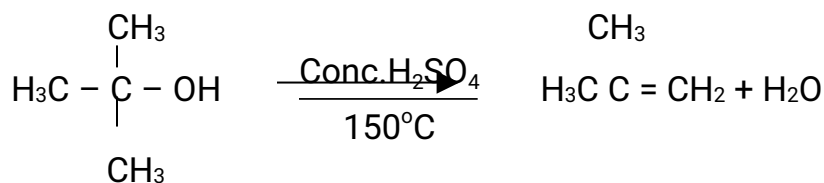
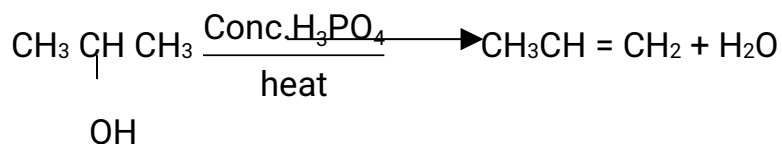
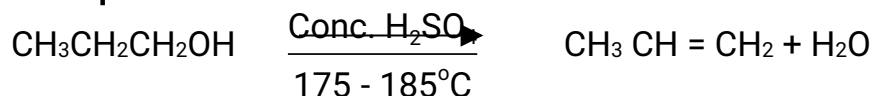
The reactivity in the liquid phase is determined by the type of carbon cation ion formed.

Elimination reactions occur under the following conditions:

- (i) Heat:                      [170 – 185°C]      Primary alcohol  
    [150 – 180°C]      Secondary alcohol  
    [90 – 150°C]      Tertiary alcohol.

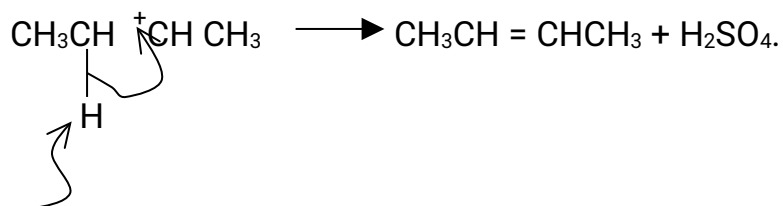
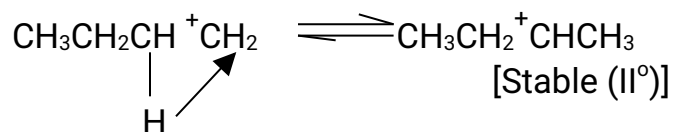
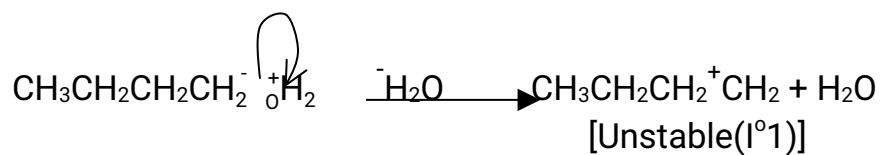
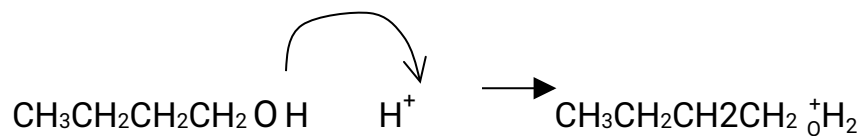
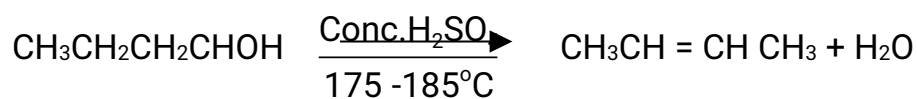
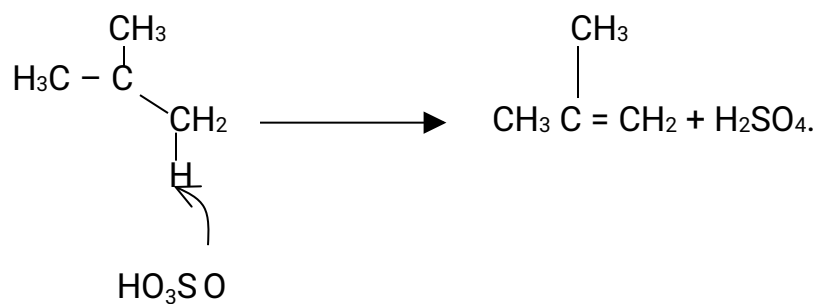
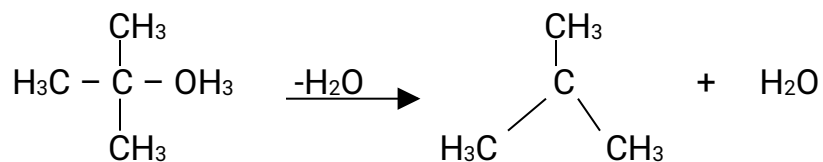
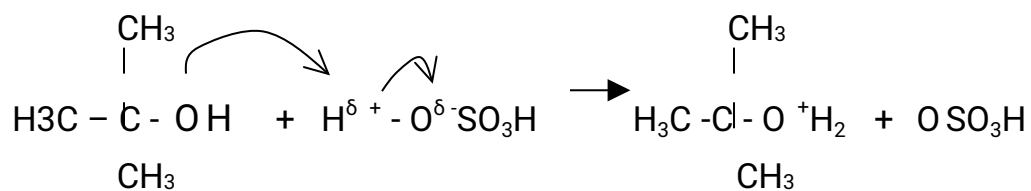
- (ii) Conc. H<sub>2</sub>SO<sub>4</sub>:            [>60%].

**Example:**



**Mechanism:**

1. Protonation of alcohol
2. Loss of H<sub>2</sub>O to form carbo cation ion.
3. Re-arrangement of the carbo cation to a more stable form.
4. Loss of the proton to form alkene.

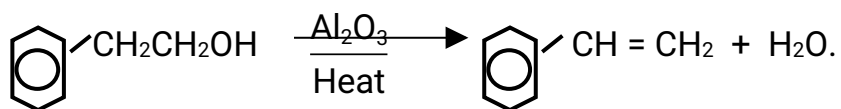
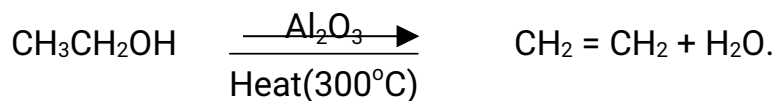
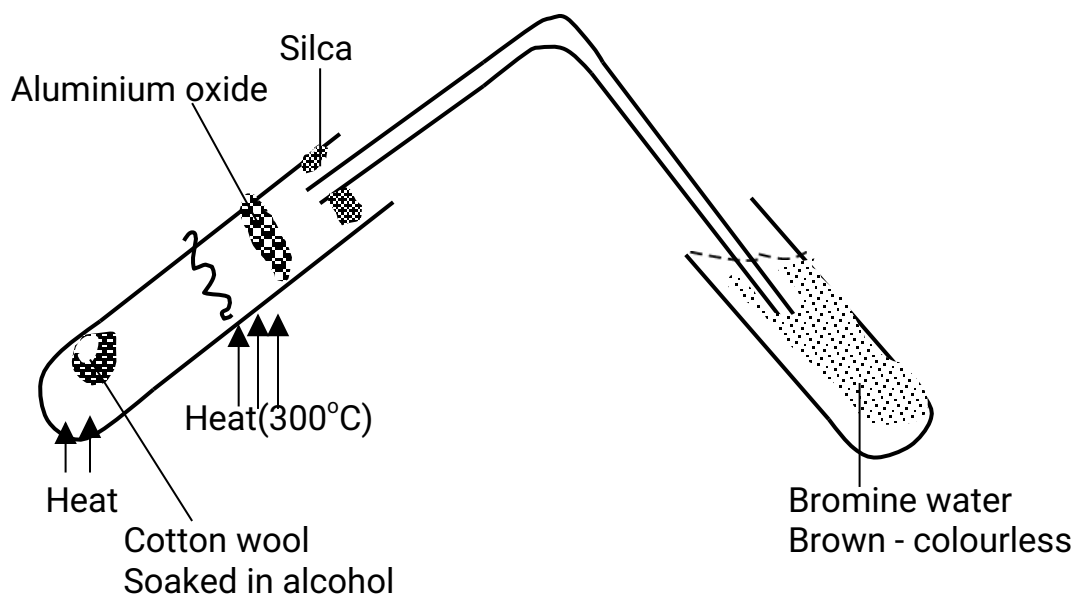




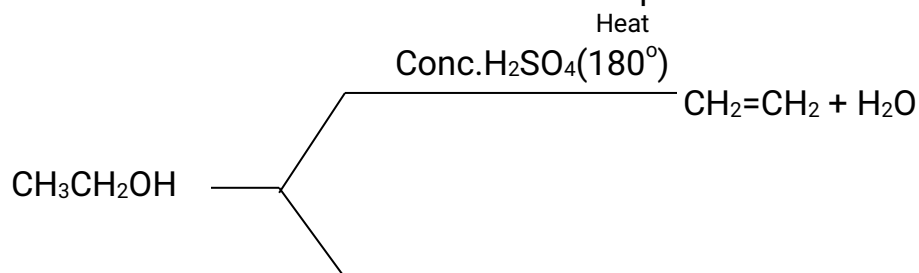
HO<sub>3</sub>SO:

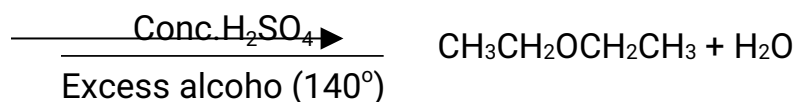
## DEHYDRATION IN VAPOUR PHASE

When an alcohol is heated and its vapour is passed over heated Al<sub>2</sub>O<sub>3</sub> at 300°, an alkene is formed which is detected by turning bromine water colourless.

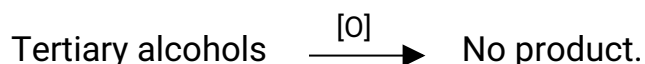


N.B: Elimination reactions in alcohols compete with substitution reactions. The difference is due to the conditions provided.





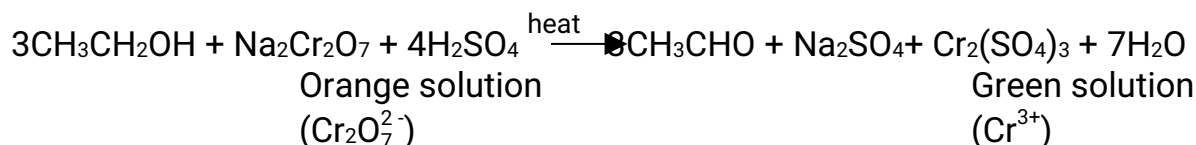
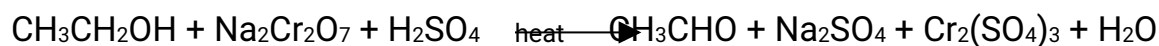
## OXIDATION REACTIONS



Oxidation of alcohols is effected by oxidizing agents such as:

- (i) Acidified  $\text{K}_2\text{Cr}_2\text{O}_7 / \text{H}^+_{(\text{aq})}$
- (ii) Acidified  $\text{Na}_2\text{Cr}_2\text{O}_7 / \text{H}^+_{(\text{aq})} / \text{Cr}_2\text{O}_7^{2-}_{(\text{aq})}$
- (iii) Chromic acid,  $\text{CrO}_3$ .

Acidification is done by use of sulphuric acid.

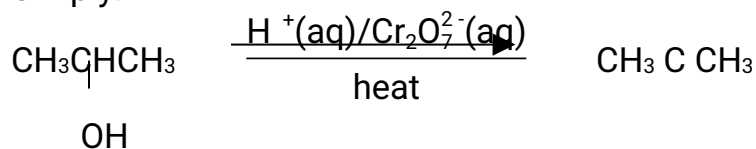


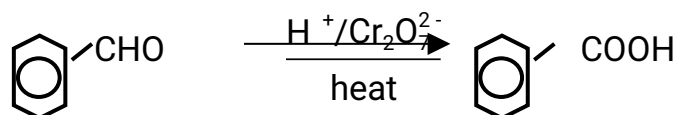
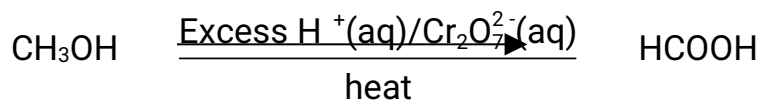
During the process, the solution mixture turns from orange to green indicating the conversion of the dichromate to chromium(III) ion.

Excess oxidizing agent.



Simply:

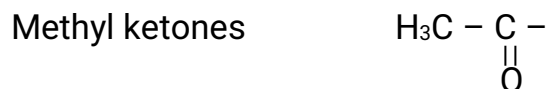
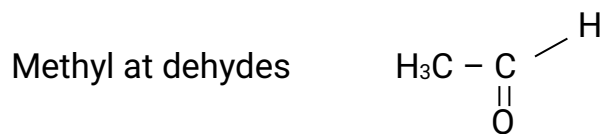
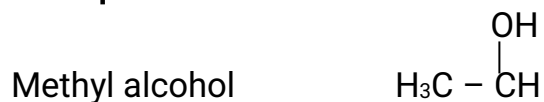




### ODO FORM/HALO FORM TEST

This test is usually performed to confirm for the presence of compounds that have got a methyl group.

#### Example:

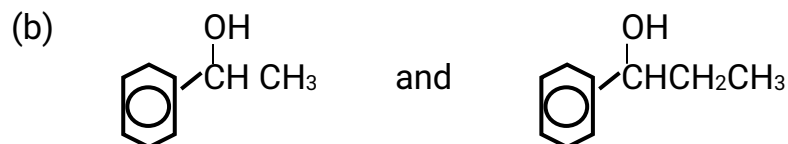


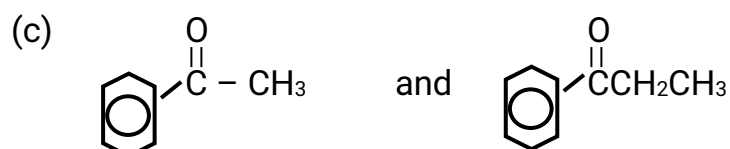
Product formed  $\text{CHI}_3$  - Tri iodo methane (Yellow ppt)

This test confirms / differentiates methyl / alcohols from the rest. Methyl alcohols form tri iodo methane when oxidized which is observed as a yellow ppt, while the rest of the alcohols no observable change.

Reagent: Hot NaOH in Iodine solution.

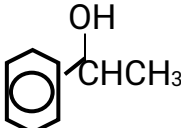
**Question:** Give a reagent that can be used to distinguish between pairs of the compounds below and state what is observed in each case. Write the equation for the reaction(s) that take place.

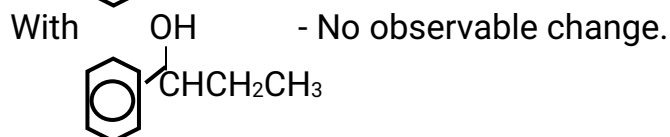


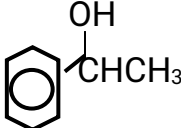


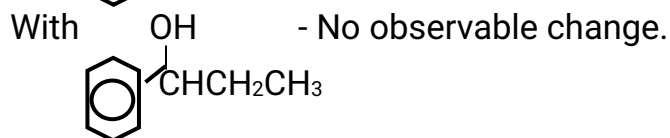
**Solutions:**

- (a) Hot sodium hydroxide solution.  
 With  $\text{CH}_3\text{CH}_2\text{OH}$  - A yellow precipitate is observed.  
 With  $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$  - No observable change.

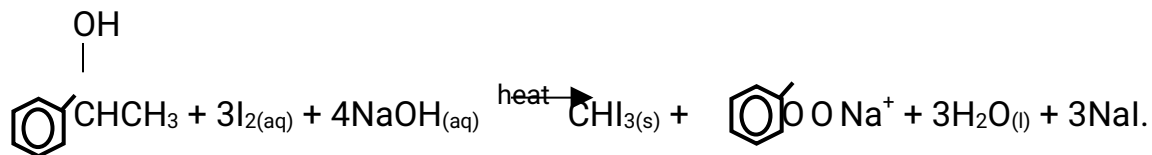
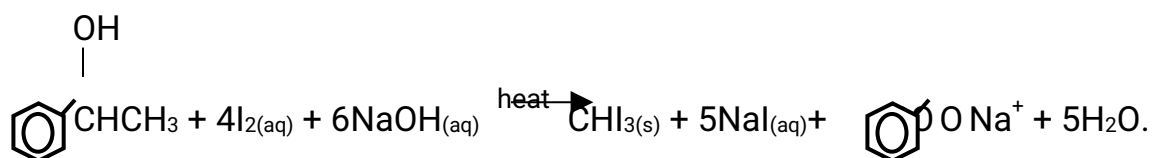
- (b) Reagent: Hot sodium hydroxide in iodine solution.  
 With  - A yellow precipitate is observed.



- (c) Reagent: Hot sodium hydroxide in iodine solution.  
 With  - A yellow precipitate is observed.



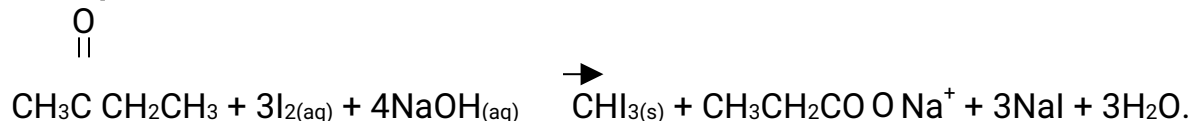
**Equations:**



Rule: Alcohols;  $4\text{I}_2$ ;  $6\text{NaOH}$

Aldehydes/ketones,  $3I_2$ ;  $4NaOH$ .

**Example:**

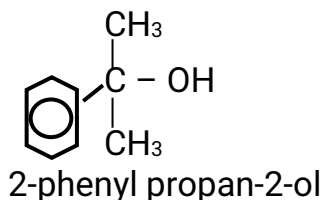
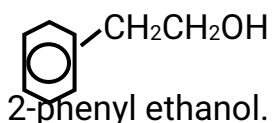
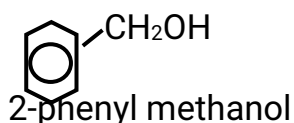


## AROMATIC ALCOHOLS

These are aromatic compounds with hydroxyl groups. They are divided into two: Aromatic alcohols and phenols.

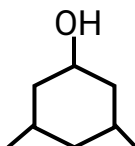
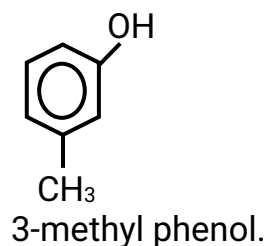
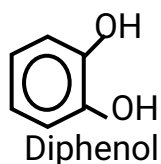
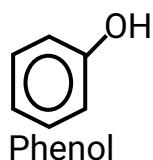
(i) **Aromatic Alcohols:**

These contain an aromatic ring and then the hydroxyl group which is not directly attached to a ring but attached via the chain on the ring.



(ii) **Phenols:**

These contain an aromatic ring with one or more OH group directly attached to the ring.



HO

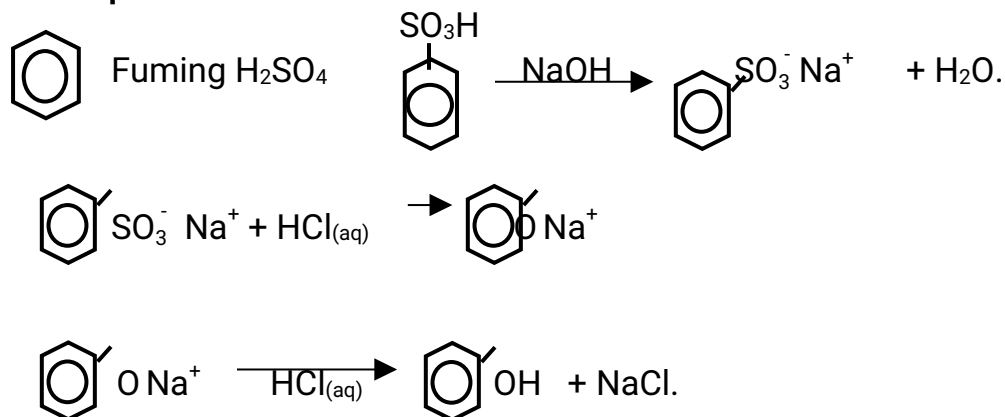
OH

## PHENOLS

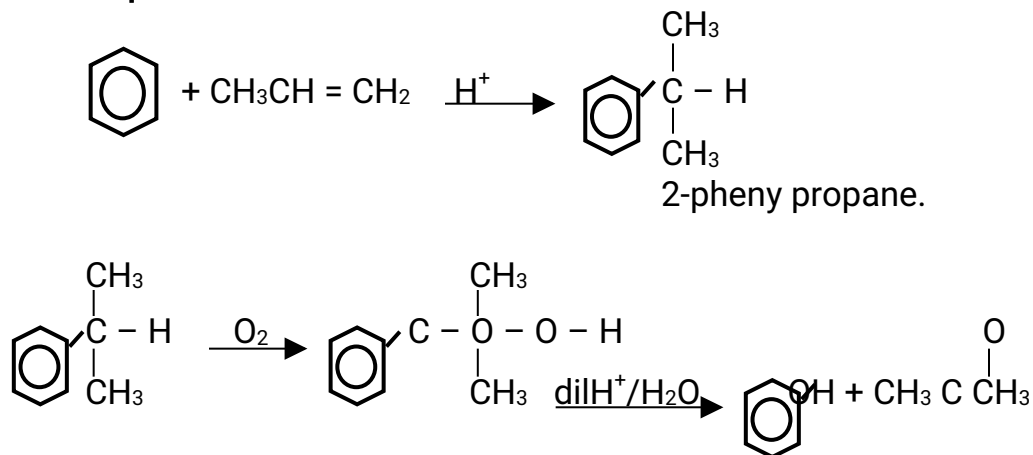
Is a class of compounds named as phenols and one or more OH groups are directly attached to the ring.

### METHODS OF PREPARATION

#### 1. Benzene sulphonic acid



#### 2. Cumene process



#### 3. Aromatic halides

In cumene process, Benzene is reacted with propene in presence of an acid or halogen carrier like  $\text{AlCl}_3$  to form two phenyl propane.

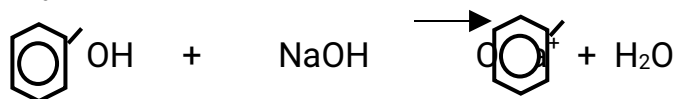
Molecular  $\text{O}_2$  is blown into this product and the intermediate is acid hydrolysed to give a phenol.

When an aromatic halide is treated with alkali at extremely harsh conditions like  $300^\circ\text{C}$ , 200 atm of pressure and copper, a phenol is formed.

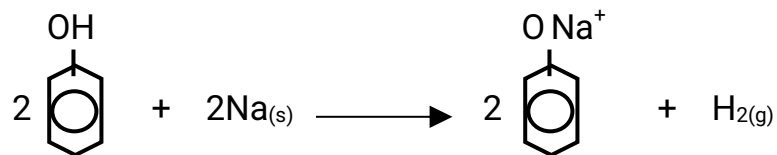
This subsequently strengthens the C – O bond by developing a partial double bond between C and O and at the same time weakening the O – H bond by decreasing the electron density and releases a proton easily. This does not occur in aliphatic aromatic alcohols as ordinary alcohols.

Subsequently, phenols being slightly acidic, under the following:

(a) Very soluble in NaOH or KOH

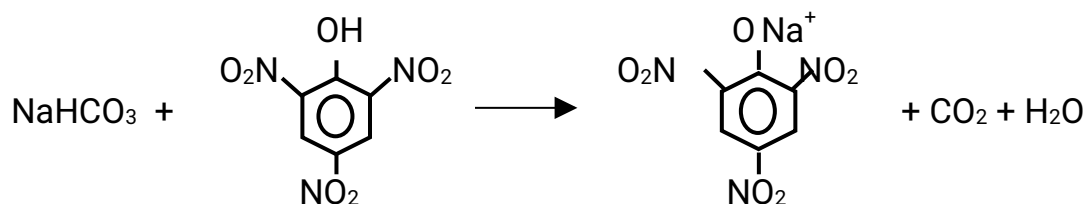


(ii) Phenols react with very reactive metals like Na to form salt and H<sub>2</sub>.



Acid properties of phenols can be improved by;

Adding electrons with a.....which de-active the ring. Such groups include the Nitro groups. 2, 4, 6-trinitrophenol is more acidic than phenol. It liberates CO<sub>2</sub> from saturated NaHCO<sub>3</sub>.



Carboxylic acids are more acidic than phenols. [Give the reactions that distinguish acidic character of phenols and carboxylic acids.]

### COMPLEX FORMATION WITH NEUTRAL FeCl<sub>3</sub>.

Phenols form a complex with aqueous neutral FeCl<sub>3</sub> and the colour of the complex product is violet colouration (purple). This reaction is important in confirming the presence of a phenolic group and it is also useful in distinguishing between phenol and aromatic alcohols.

Question: Distinguish between HCOOH and CH<sub>3</sub>COOH.  
(We use the Iodo form test).

OR:

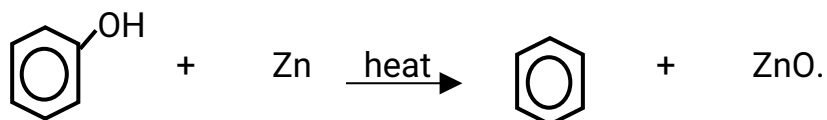
Using neutral (iron III) chloride) solution.

A purple colouration with HCOOH.

No observable change with CH<sub>3</sub>COOH.

### REACTION WITH ZINC DUST:

Phenol when heated with Zinc dust forms benzene.



### ALCOHOL REACTIONS OF PHENOLS RESEMBLING

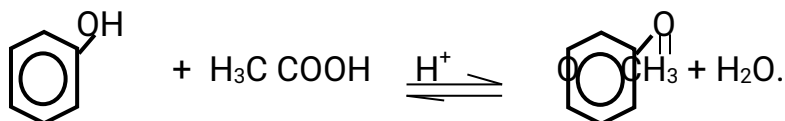
This involves cleavage of C - O bond.

Phenols undergo some reactions resembling alcohol where the C - O bond is involved.

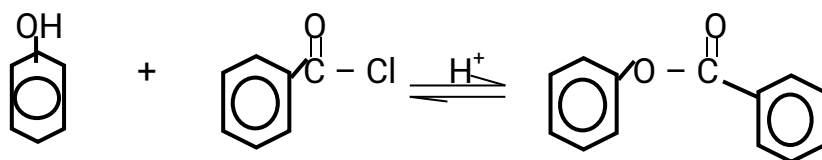
#### Esterification:

Phenols react with carboxylic acids, halides and acid anhydrides to form an ester.

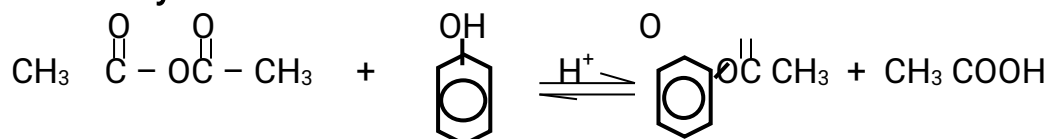
#### Carboxylic acids:



#### Acid halides:



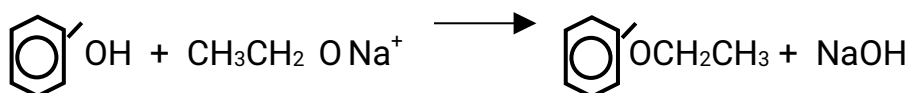
#### Acid anhydride:



#### Ether formation:

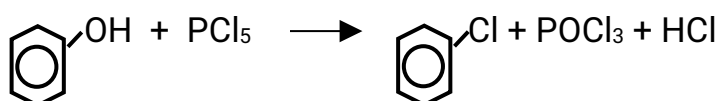
Phenols react with alkali oxides to form ethers.





### Phosphorous penta halides:

Phenols react with  $\text{PCl}_5$  liberating  $\text{HCl}$  gas which fumes in air.



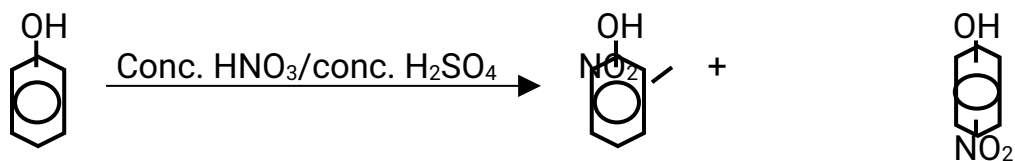
### Reactions with liquid $\text{NH}_3$ .

## REACTIONS OF THE RING IN PHENOL

Due to the presence of the  $\text{OH}$  group in phenol, phenol undergoes a number of reactions where delocalized electrons on the ring are involved.

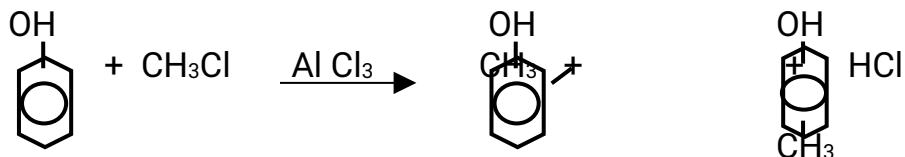
### Nitration:

Phenol reacts with conc.  $\text{H}_2\text{SO}_4$  and conc.  $\text{HNO}_3$  acids to form 2-nitro phenol and 4-nitrophenol.



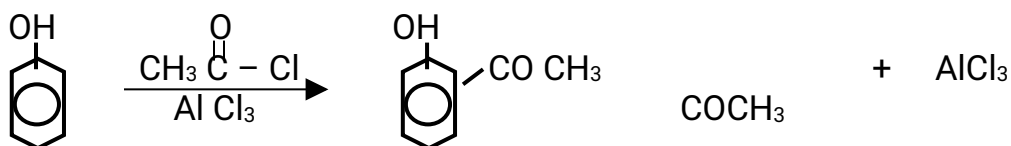
### Alkylation:

Phenol reacts with alkyl halides in presence of a halogen carrier to form alkylated product phenol.



Acylation  $\left( \begin{smallmatrix} \text{O} \\ \parallel \\ \text{C} \end{smallmatrix} \right)$

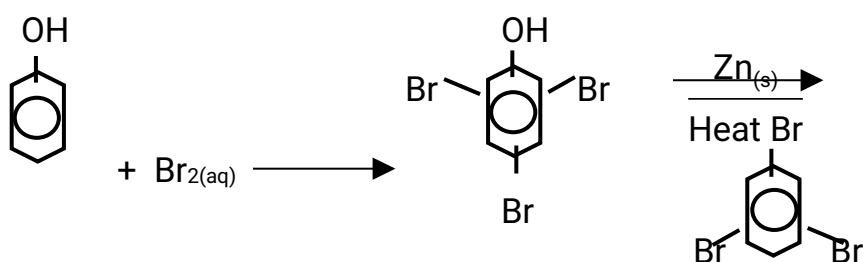
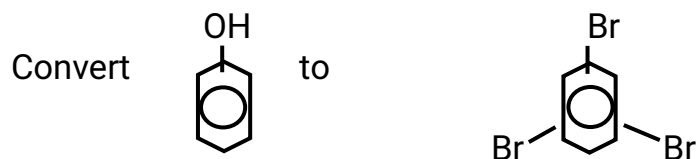
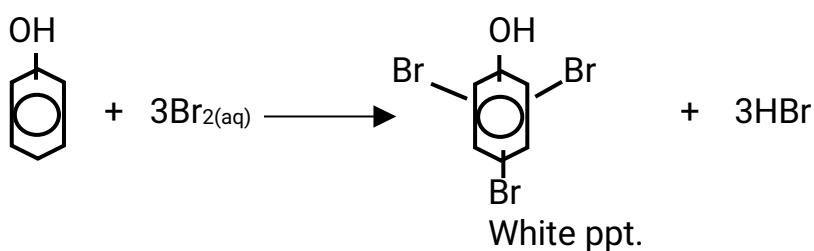




### Bromine H<sub>2</sub>O.

When reacted with bromine, H<sub>2</sub>O phenol forms a substituted poly product of 2, 4, 6 tri bromo phenol which is a white precipitate.

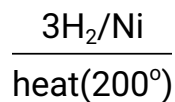
The presence of OH makes phenol to be very reactive. This reaction is very important in confirming the presence of phenol together with the neutral iron (III) chloride test.



### Hydrogen in presence of heated Ni catalyst.

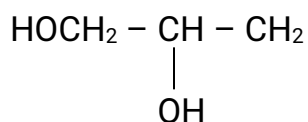
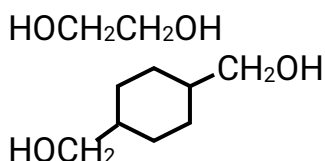
Phenol can be hydrogenated into cyclo hexanol when reacted with H<sub>2</sub> in the presence of a heated catalyst.





## POLYHYDRIC ALCOHOLS

These are alcohols with more than one OH group attached to the same carbon chain as:



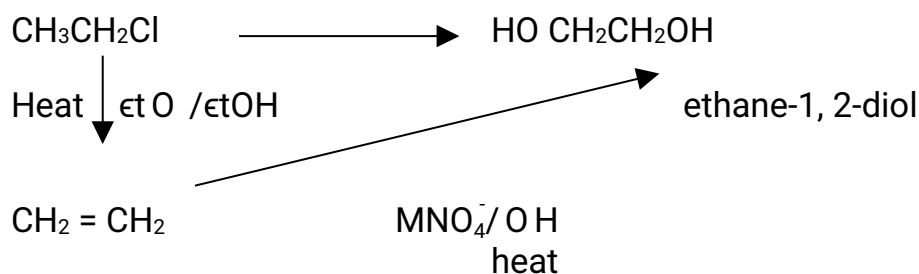
The chemical and physical properties of these types of alcohols are slightly the same; however, there are slight differences.

Ethane-1, 2-diol boils at a higher temperature.

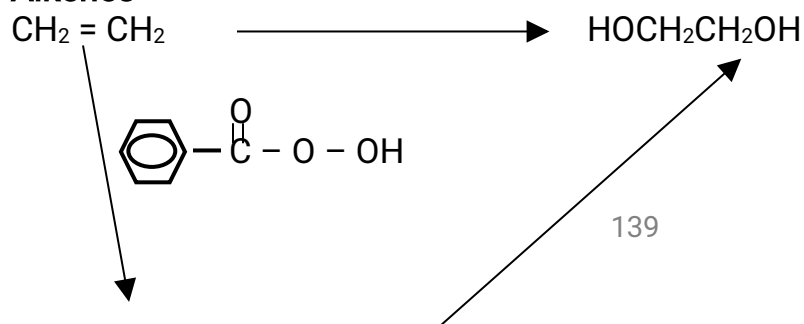
Ethane-1, 2-diol has two OH group which results into formation of extensive hydrogen bonds which is difficult to break hence high boiling point.

## PREPARATION OF DIOLS

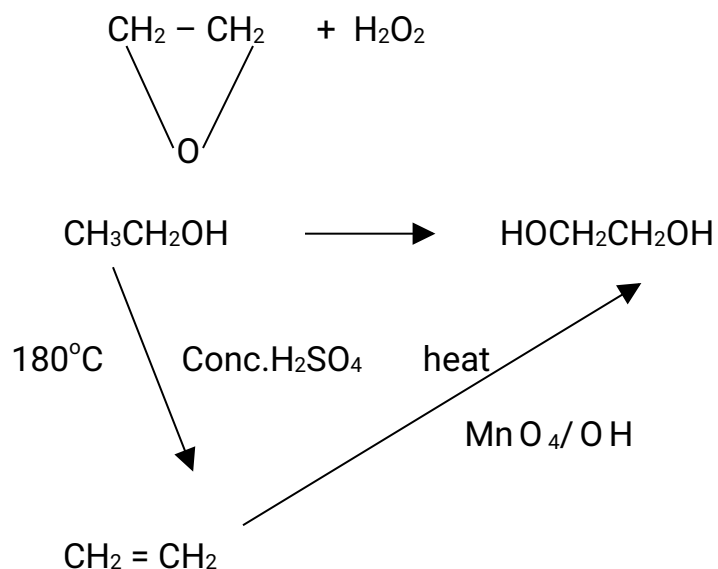
They can be prepared from alkyl halides.



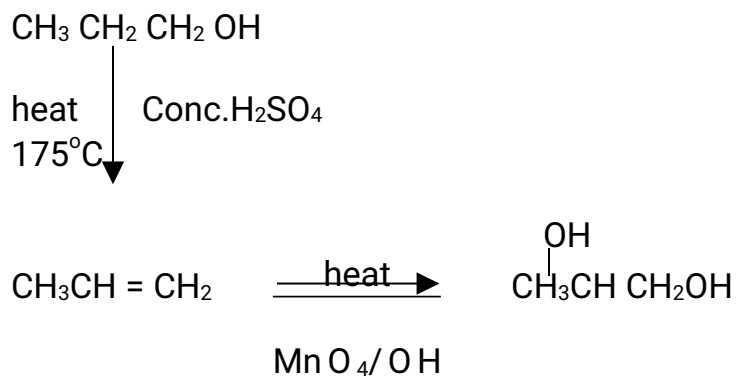
### Alkenes



Benzene peroxo acid      200°C/pressure



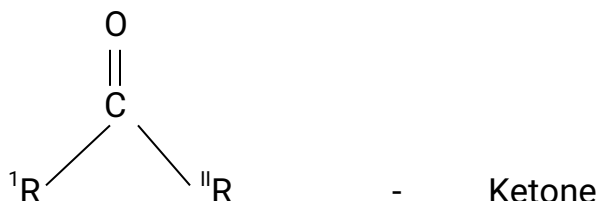
Reactions of diols are the same as those of monohydric alcohols. The difference is that excess of the reagent is used.



## CARBONYL COMPOUNDS

These are compounds containing a carbonyl functional group. They are alkanal (Aldehydes) and alkanones (Ketones).

These two belong to a group of organic compounds known as carbonyl compounds with a general formula  $C_nH_{2n}O$ .

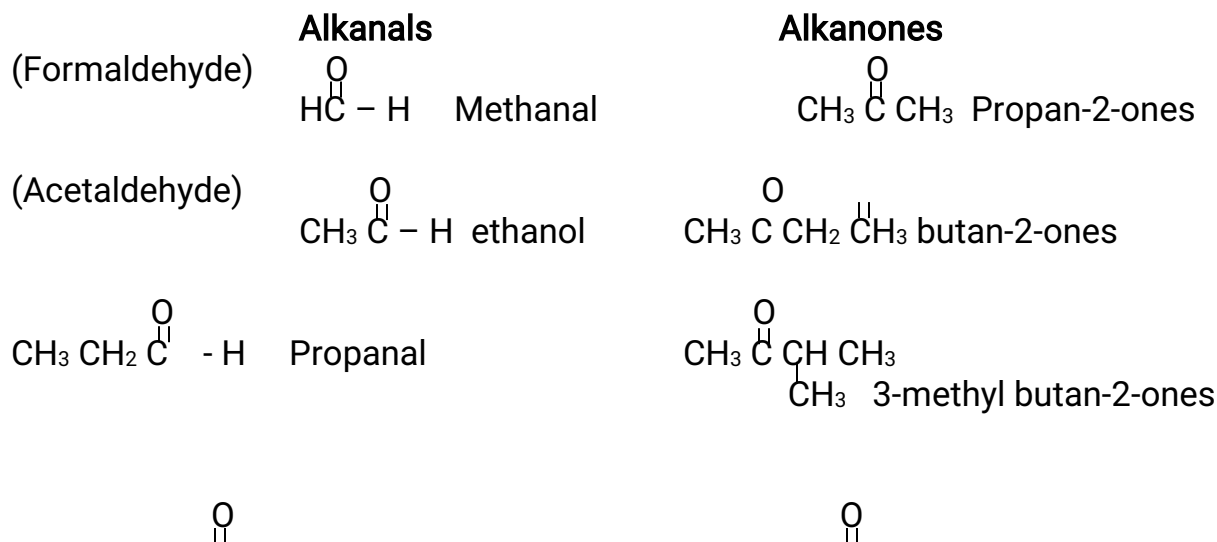


The major difference between aldehyde and ketone is in the presence or absence of  $H_2$  attached onto the carbonyl carbon.

In aldehydes, "H" is to the carbonyl carbon while in a ketone, similar or different alkyl groups are attached to the carbonyl carbon as shown above.

## NOMENCLATURE

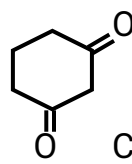
IUPAC system is used in the naming of both aldehydes and ketones. Generally, aldehydes are named as alkanals by replacing "e" in the alkane name with –"als" while ketones are named as alkanones by replacing –"e" with –"ones".



$\text{CH}_3\text{CH}_2\text{CH}_2\text{C} - \text{H}$  – Butanal

$\text{CH}_3\text{CH}(\text{CH}_3) - \text{C} - \text{CH}_2\text{CH}_2\text{CH}_3$   
2-methyl hexan-3-one

$\text{CH}_3\text{CH}(\text{CH}_3)\text{CHO}$  2-methyl propanal

  
Cyclohexa-1,3-dione.

## ISOMERISM

Both aldehydes and ketones exhibit structural and functional isomerism.

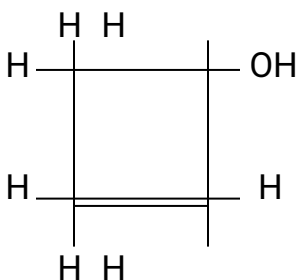
In structural isomerism, they show position and chain isomers.

In functional isomerism, both aldehydes and ketones are isomeric within themselves and with other classes of organic compounds like alcohols and cyclic ethers.

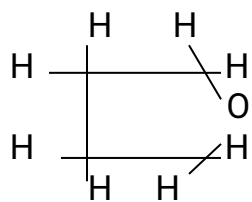
$\text{C}_4\text{H}_8\text{O}$

$\text{CH}_3\text{CH}_2\text{CH}_2\text{CHO}$  Butanal

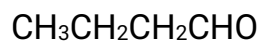
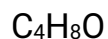
$\text{CH}_3\text{CH}(\text{CH}_3)\text{COCH}_3$  Butan-2-one



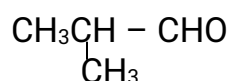
Cyclic butanol (cyclic alcohol)



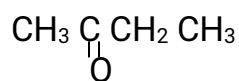
Cyclic ether (Tetacfuran)



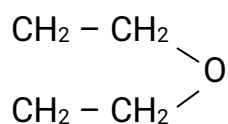
Butanal



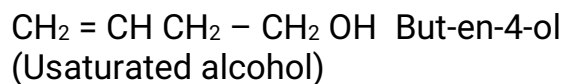
2-methyl propanal (Aldehyde)



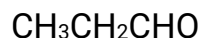
Butan-2-one (Ketone)



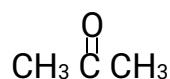
Tetrahydrofuran (furan)



Write all the isomers of the compounds with the molecular formula  $\text{C}_3\text{H}_6\text{O}$ .



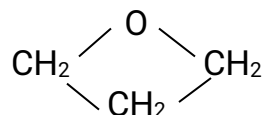
Propanal (Aldehyde)



Propanone (Ketone)



Prop-1-en-3-ol.



Trihydrofuran

## PREPARATION OF CARBONYL COMPOUNDS

They can be prepared from the following:

- i) Mild oxidation of alkanol (alcohols)
- ii) Carboxylic acids or their salts.
- iii) Ozonolysis of alkenes followed by hydrolysis.

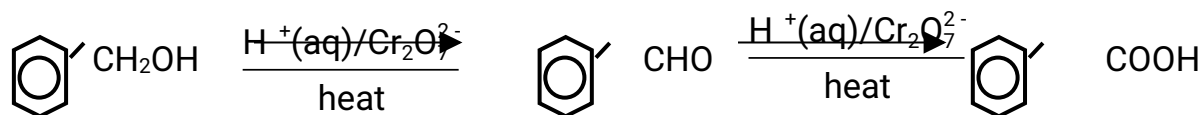
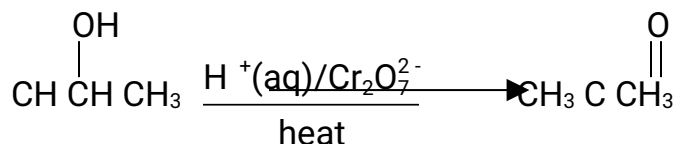
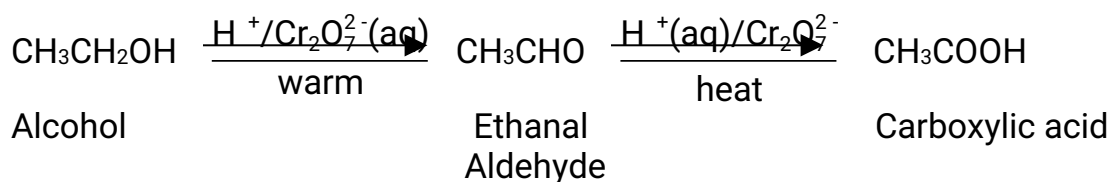
- iv) Acid hydrolysis of alkynes.
- v) Hydrogenation of acid chlorides.
- vi) Acylation (aromatic ketones)

## OXIDATION OF ALCOHOLS

Primary and secondary alcohols are oxidized in liquid phase using  $\text{K}_2\text{Cr}_2\text{O}_7 / \text{H}^+$  or  $\text{Na}_2\text{Cr}_2\text{O}_7 / \text{H}^+$ ,  $\text{CrO}_3 / \text{H}^+$  to form aldehydes and ketones respectively.

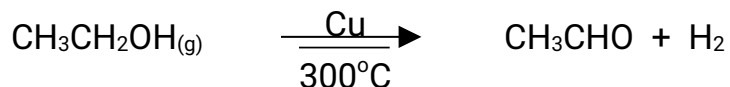
During oxidation process, a colourless solution turns to green solution.

### Examples:

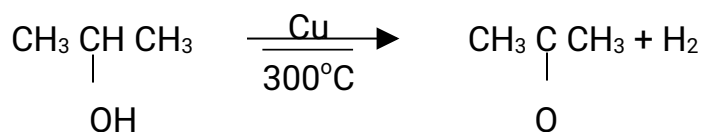


During the oxidation process, especially for primary alcohols, there will be further oxidation to carboxylic acid which renders this method ineffective but this further oxidation can be avoided by carrying out a reaction under low heat (temperatures) or limiting the amount of the oxidant used.

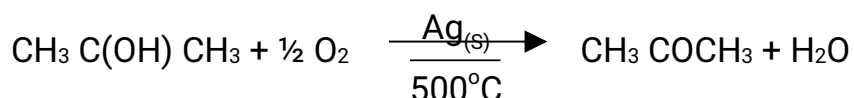
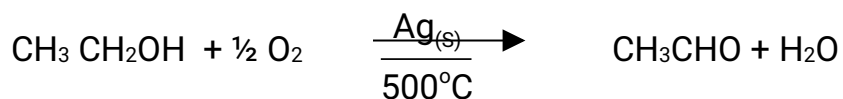
Oxidation can also be effected under vapour phase. When an alcohol vapour is passed over heated Cu at  $300^\circ\text{C}$ , it is oxidized to an aldehyde or ketone.





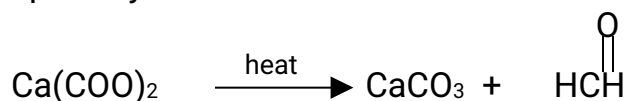


Oxidation under vapour phase can also be brought about when an alcohol together with limited  $\text{O}_2$  is passed over heated Ag at  $500^\circ\text{C}$ .



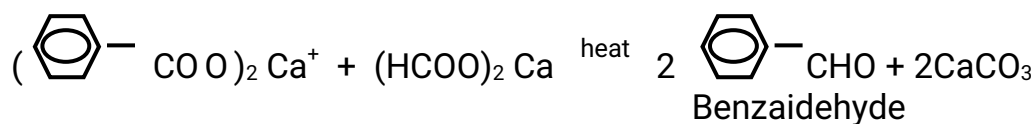
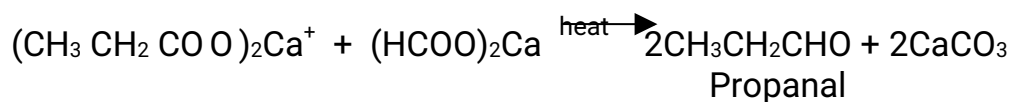
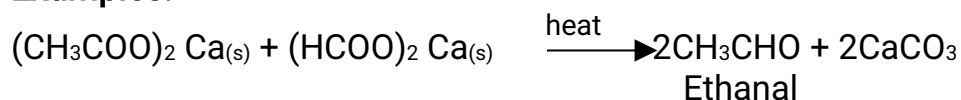
## CARBOXYLIC ACIDS

When a carboxylic acid salt of calcium is heated, an aldehyde is formed. More specially methanol is formed.

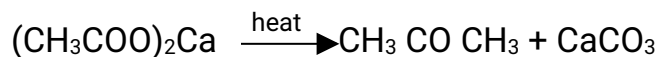


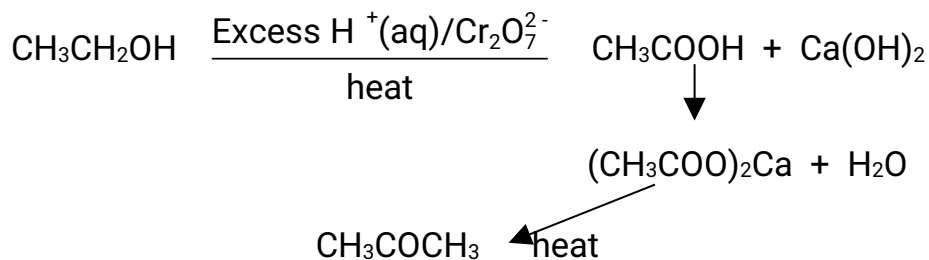
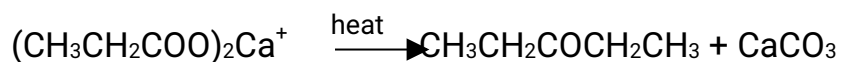
With other higher aldehydes, Ca salt of a carboxylic acid is heated with calcium methanoate.

### Examples:



Ketones can be prepared by heating Ca salts of higher carboxylic acids.

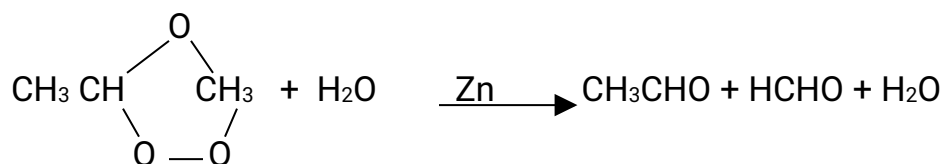
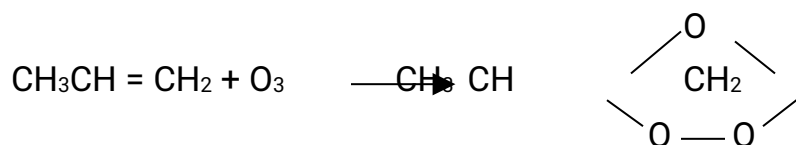
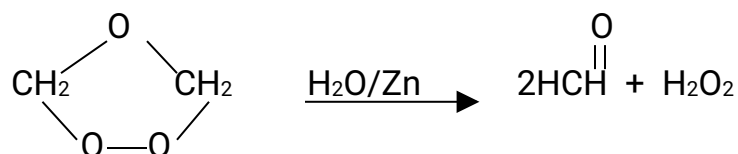
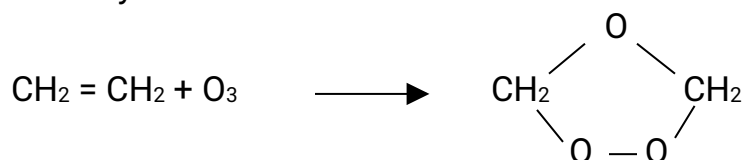


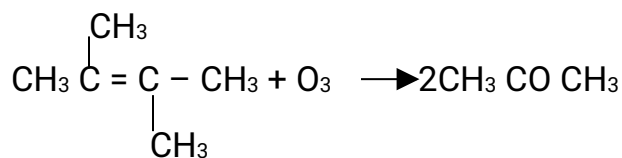


## OZONOLYSIS

When an alkene is reacted with ozone, an ozonide is formed. When the ozonide is diluted in  $\text{H}_2\text{O}$  in presence of some zinc, a carbonyl compound is formed. Zinc dust is used to decompose  $\text{H}_2\text{O}_2$ .

**Note:** Aldehydes formed depends on the types of alkene used. Symmetrical alkene gives one type of aldehyde while unsymmetrical will give you both carbonyls.

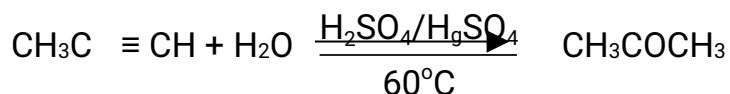
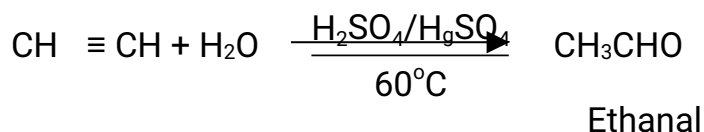




The above method has the disadvantage of yielding a mixture of products which may not be good in organic synthesis.

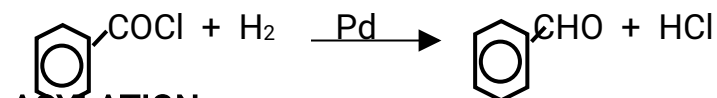
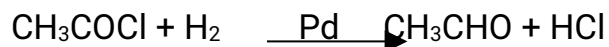
### HYDROLYSIS (ACID) OF ALKYNES

When an alkyne is passed through hot  $\text{H}_2\text{SO}_4$  in presence of mercury sulphate, the carbonyl carbon is formed.



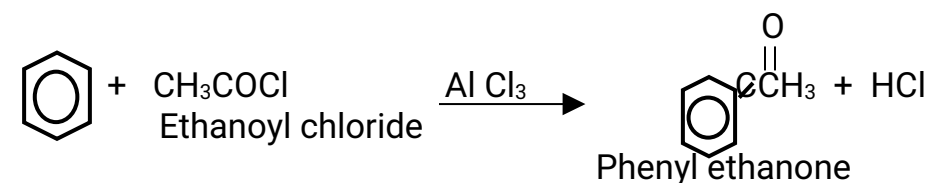
### HYDROGENATION OF ACID CHLORIDES

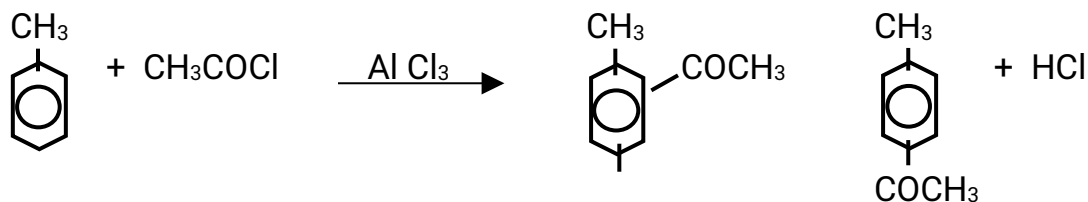
Acid chlorides are reduced by hydrogen in presence of catalysts like Pd to form aldehydes.



### ACYLATION

Alkyl acylation is a reaction between acid halides with a benzene ring compound. This reaction produces aromatic ketones when it is carried out in presence of a halogen carrier ( $\text{AlCl}_3$ ,  $\text{FeCl}_3$ )





## PHYSICAL PROPERTIES OF CARBONYL COMPOUNDS

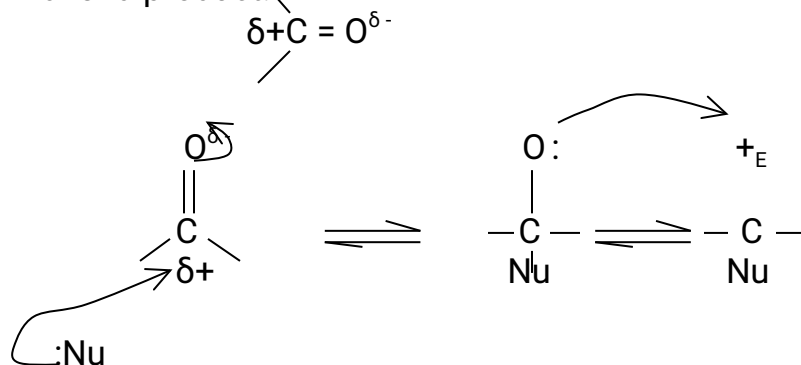
Lower members of aldehydes and ketones are liquids and are miscible with  $\text{H}_2\text{O}$  forming neutral solutions. The miscibility in  $\text{H}_2\text{O}$  is due to the polar carbonyl carbon.

They have higher boiling points than their hydrocarbon counterparts. This is because of the polar nature of the carbonyl carbon that results into the intermolecular attractions but their boiling points are lower than those of alcohols and carboxylic acids due to the lack of  $\text{H}_2$  bonding.

## CHEMICAL PROPERTIES OF CARBONYL COMPOUNDS

Carbonyl compounds chemically react because of the property of the carbonyl carbon which is a common functional group present in both aldehydes and ketones.

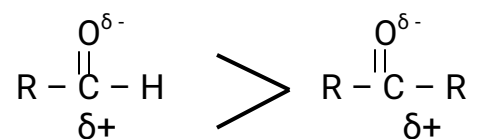
A carbonyl carbon contains more electronegative oxygen bonded to carbon which pulls electrons towards itself and attains a  $\delta^-$  charge while the carbon remains  $\delta^+$ . Therefore, the carbon oxygen bond is polarised creating a good condition for nucleophilic reagents to be added across the carbonyl carbon. The intermediate formed with a negative on oxygen reacts with an electrophile to form the final end product.



Carbonyl group in aldehyde is more reactive than in ketones.

**Explanation:**

In ketones, the presence of two alkyl groups having a positive inductive effect highly neutralizes the positive charge in the carbon resulting into a less attraction of a nucleophile.

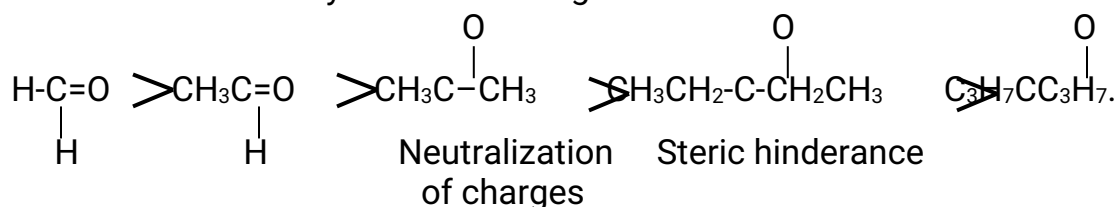
***Methanol is more reactive than other aldehydes. Explain.***

In methanol only hydrogen is bonded to the carbonyl carbon that the partial positive charge formed is not neutralized making methanol more reactive.

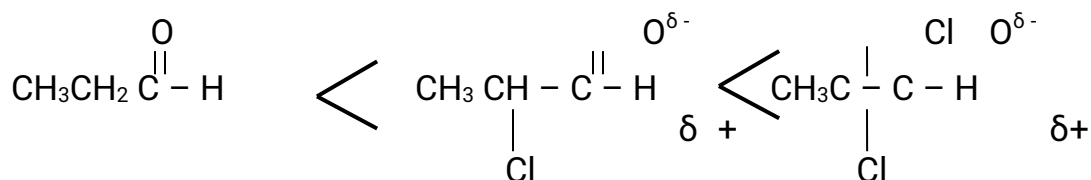
Therefore, the more number of alkyl groups added to a carbonyl carbon, the less the reaction due to the following reasons:

- (i) The alkyl groups have got electron pushing (positive inductive) effect which neutralizes the partial positive charge on the carbon.
- (ii) Several alkyl groups will have a crowding effect preventing a nucleophile from being attached to the carbon.

Order of the reactivity of the following:

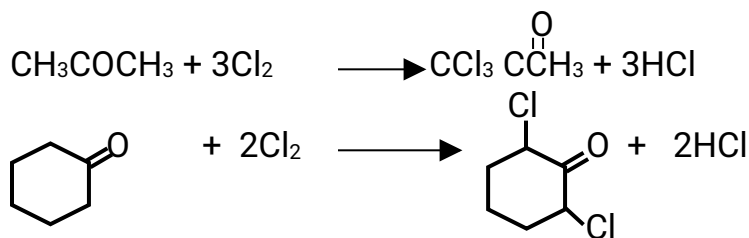
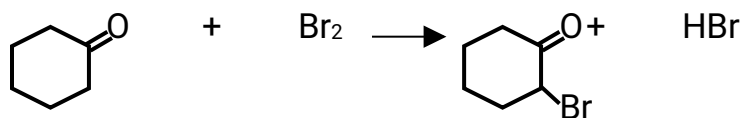
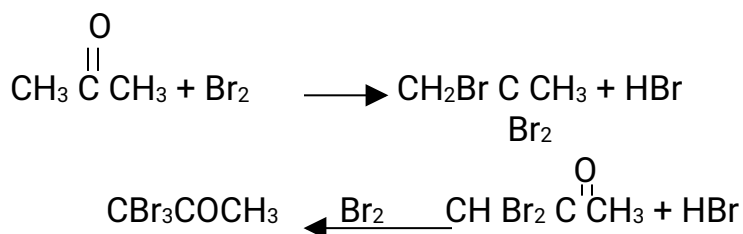
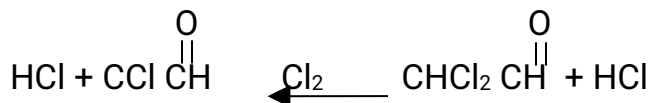
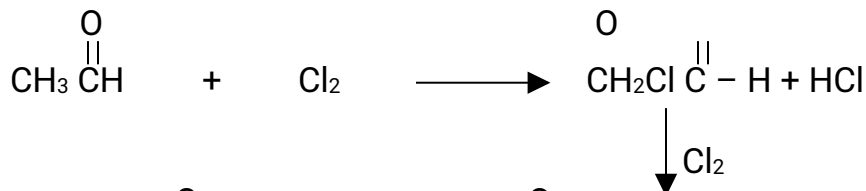


When other atoms of a more electronegative effect are added next to the carbonyl group, the reactivity of the carbonyl carbon compound increases. E.g. if a hydrogen on the carbon atom next to the carbonyl group is replaced with a halogen, the reactivity increases.

**REACTION OF CARBONYL COMPOUNDS****GENERAL REACTION (BOTH KETONES AND ALDEHYDES)**

- (i) Reactions of alkyl groups:

Both aldehydes and ketones containing alpha hydrogen i.e. hydrogens on the carbons next to the carbonyl group undergo halogenation reaction, when reacted with a halogen to form substituted products. Substitution proceeds until the alpha hydrogens are replaced.

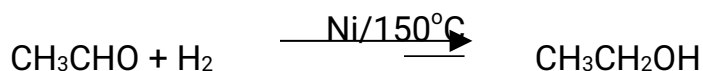


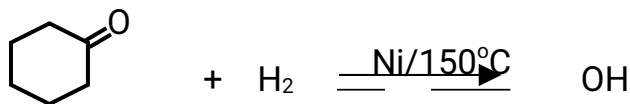
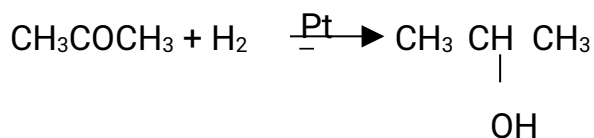
(ii) **Reactions due to carbonyl group:**

Both aldehydes and ketones containing a carbonyl group undergo the following reactions:

**Reduction (Reaction with  $\text{H}_2$ )**

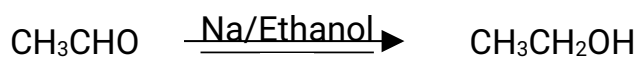
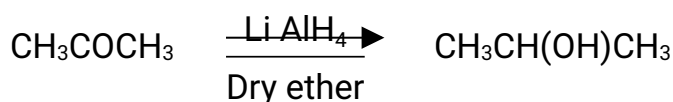
Aldehydes and ketones are reduced by  $\text{H}_2$  in presence of a suitable catalyst e.g. Ni/ $150^\circ\text{C}$ , Pt/Pd into primary and secondary alcohols respectively.



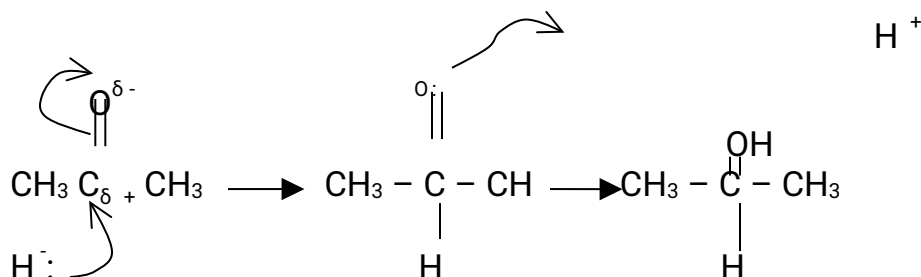


### Using reducing agents:

Similarly strong reducing agents such as Aluminium, Na, Lithium tetrahydride, Na/Ethanol can also be used to effect reduction.

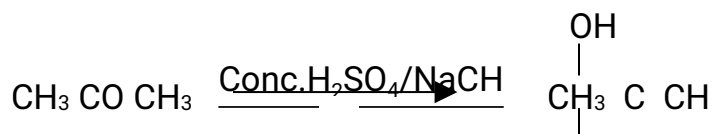
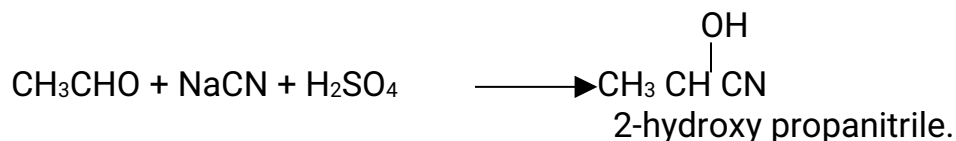


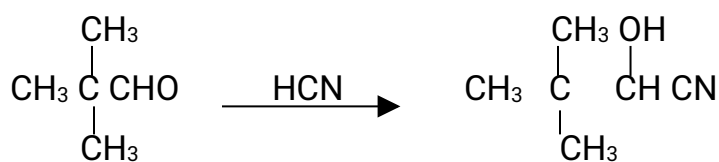
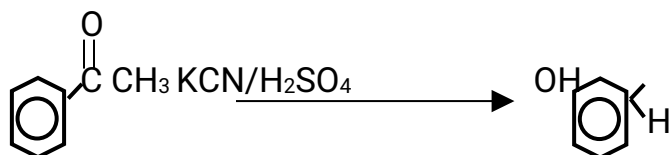
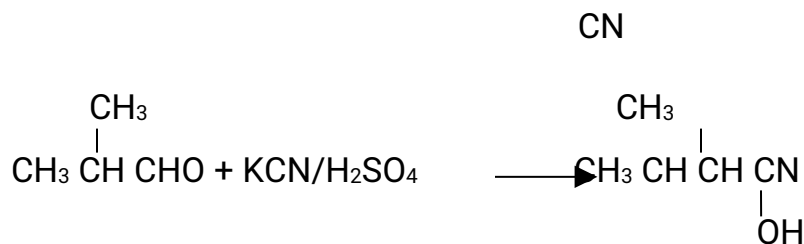
### Mechanism of reduction:



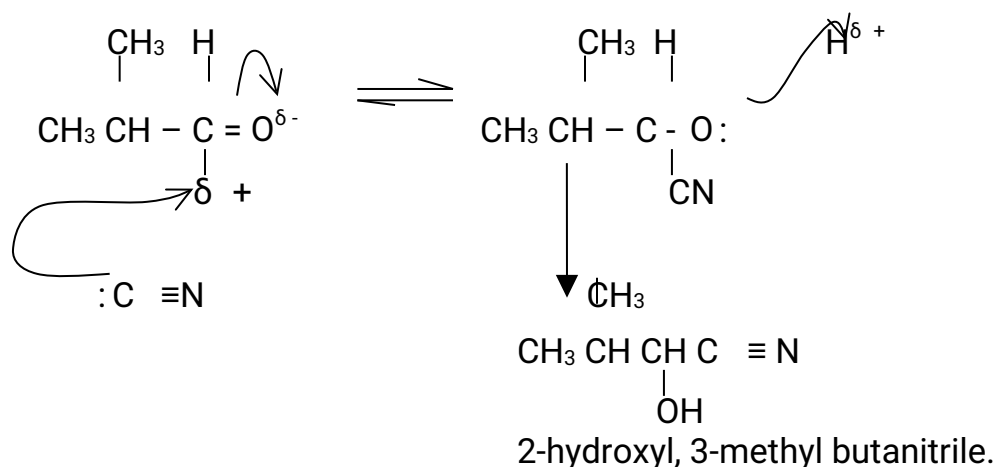
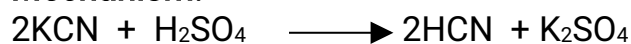
### Addition of HCN (hydrogen cyanide)

Both carbonyl compounds react with HCNs to form 2 hydroxyl nitrates or cyanohydrins. The HCN is generated insitu from the reaction between NaCN and Conc. H<sub>2</sub>SO<sub>4</sub>.



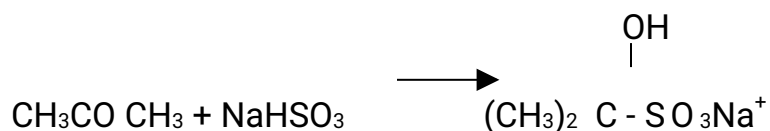
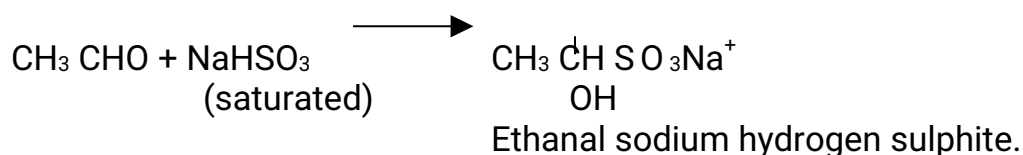


### Mechanism:

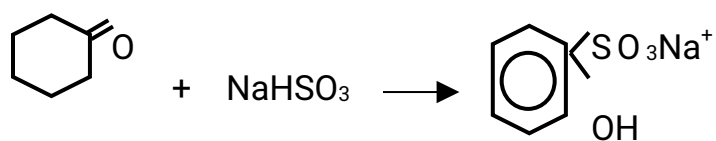


### REACTION WITH NaHSO<sub>3</sub>

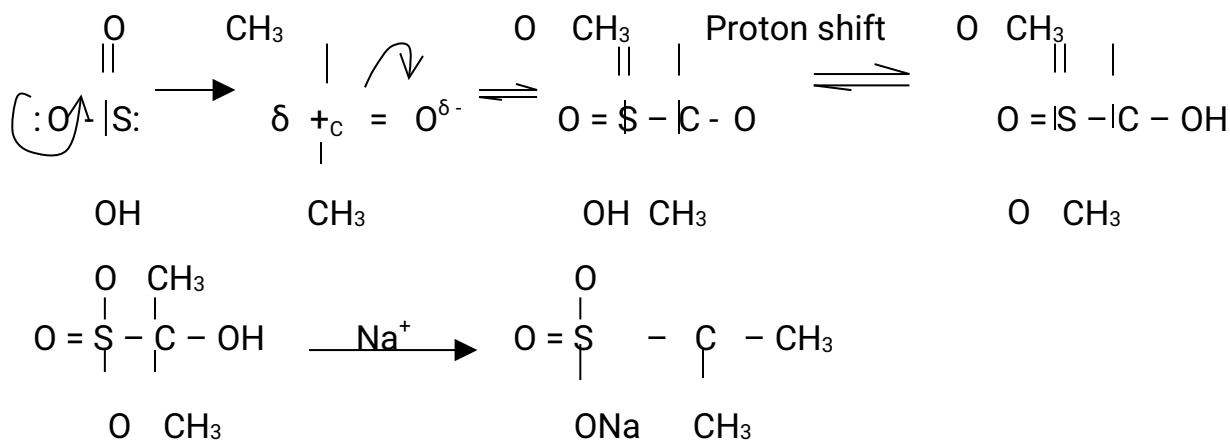
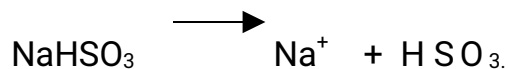
Both carbonyl compounds react with a saturated solution of NaHSO<sub>3</sub> to give crystalline solids of NaHSO<sub>3</sub> derivatives.





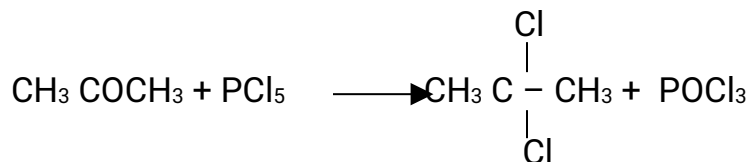
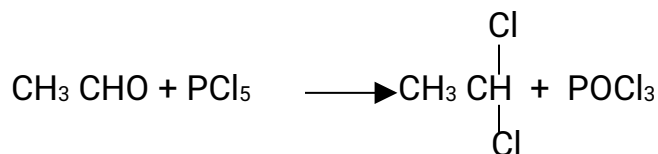


### Mechanism:



### ADDITION REACTION WITH PCl<sub>5</sub>.

Both carbonyl compounds react with PCl<sub>5</sub> to give gem-dihalides with no steaming in moist air i.e. (No hydrogen halide formed) which distinguishes between alcohols and carbonyl compounds.

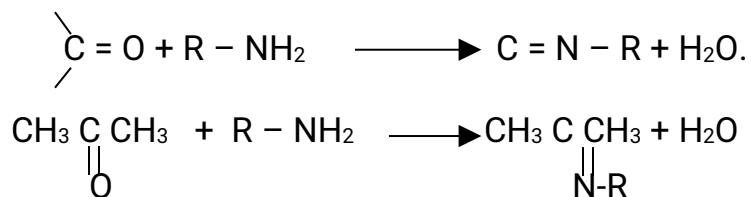




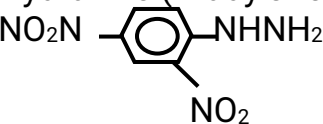
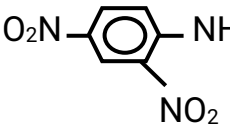
### CONDENSATION REACTIONS

Aldehydes and ketones react with compound containing amino groups to form condensations products and loss of H<sub>2</sub>O molecule. These products formed have got sharp melting points which are used in identification of the original

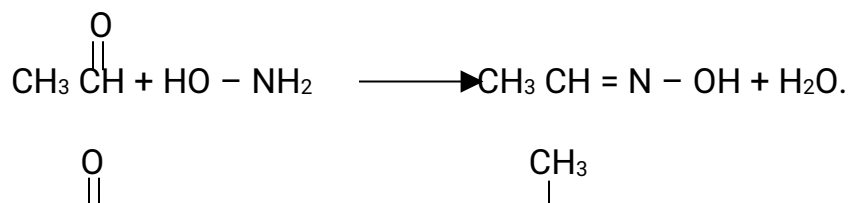
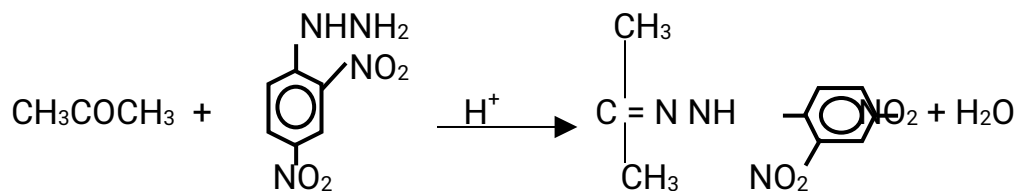
compounds.

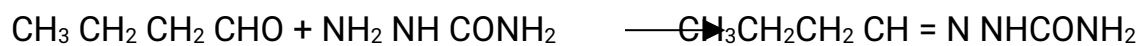
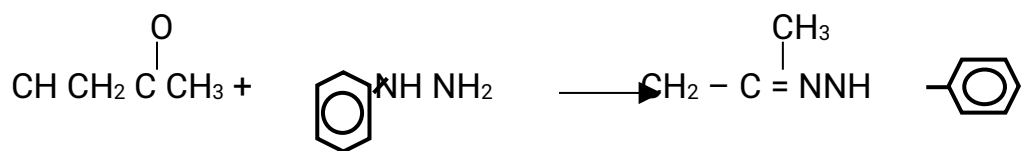
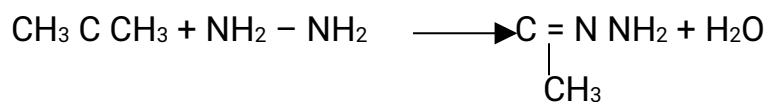
These types of reactions are condensation because there is elimination of a  $\text{H}_2\text{O}$  molecule.



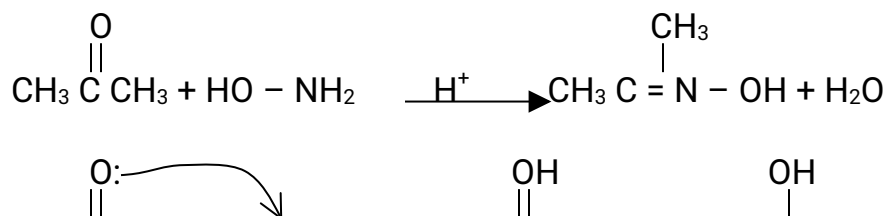
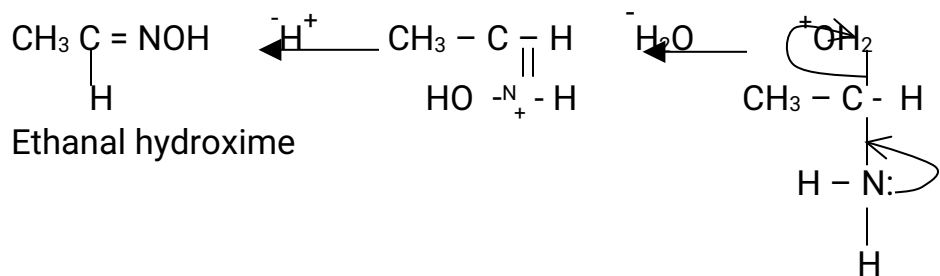
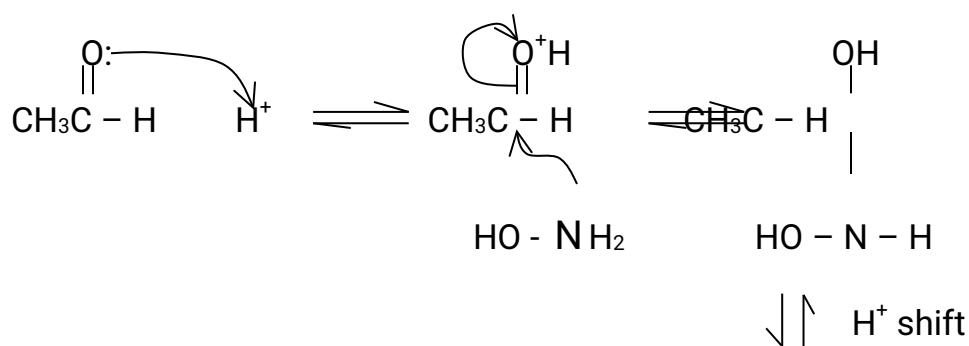
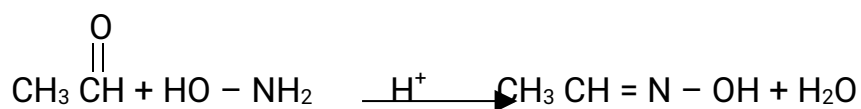
	Reagents	Product	R – group
1.	Hydroxyl amine $\text{HO} - \text{NH}_2$	Oxime $\text{C} = \text{N} - \text{OH}$	$-\text{OH}$
2.	Hydrazine $\text{NH}_2 - \text{NH}_2$	Hydra zone $\text{C} = \text{N} \text{NH}_2$	$-\text{NH}_2$
3.	Phenyl hydrazine 	Phenyl hydra zone $\text{C} = \text{N} \text{NH} - \text{C}_6\text{H}_5$	
4.	Semi carbazine $\text{NH}_2 \text{NH} \text{CONH}_2$	Semi carbazone $\text{C} = \text{N} - \text{NHCONH}_2$	$\text{NH} - \text{CONH}_2$
5.	2, 4 dinitrophenyl Hydrazine (Brady's reagent) 	2, 4 dinitrophenyl hydrazone $\text{C} = \text{N} \text{NH} - \text{C}_6\text{H}_3(\text{NO}_2)_2$	

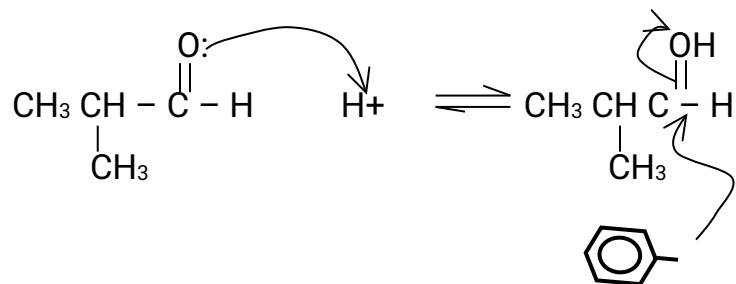
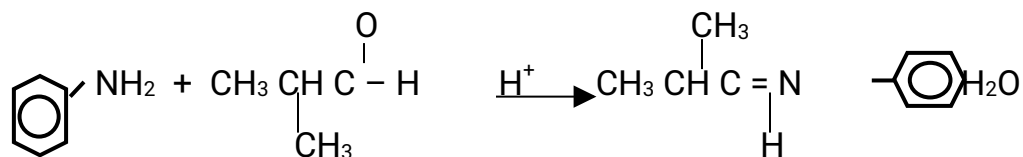
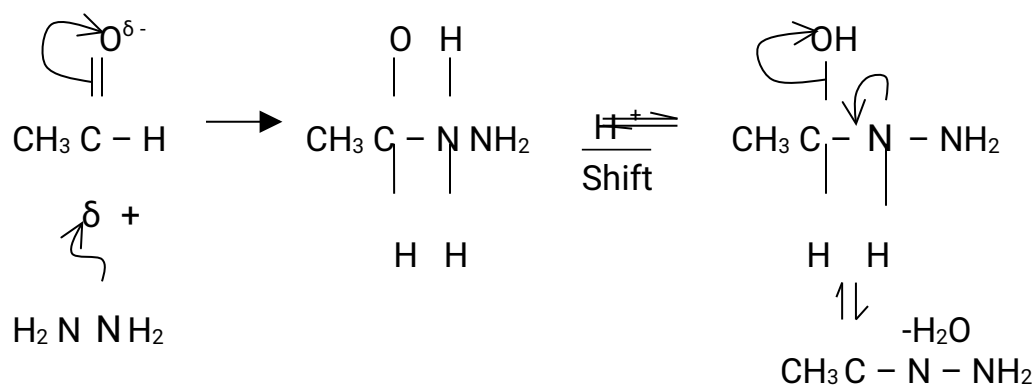
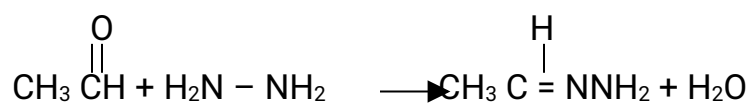
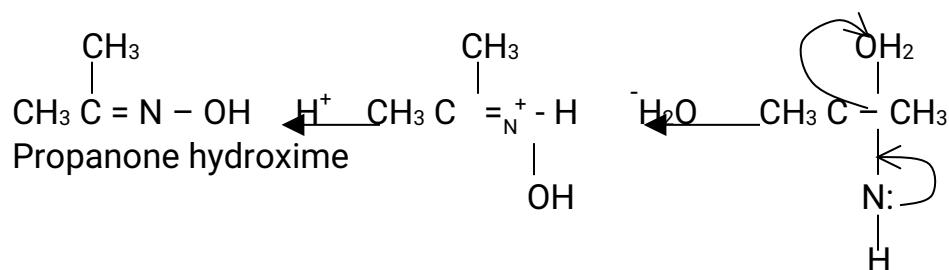
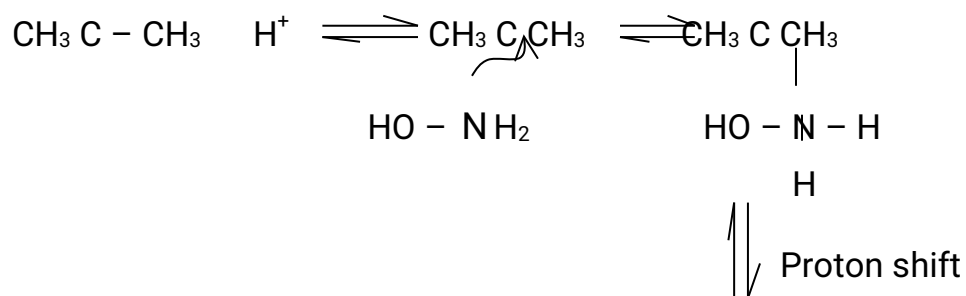
**Mechanism:**

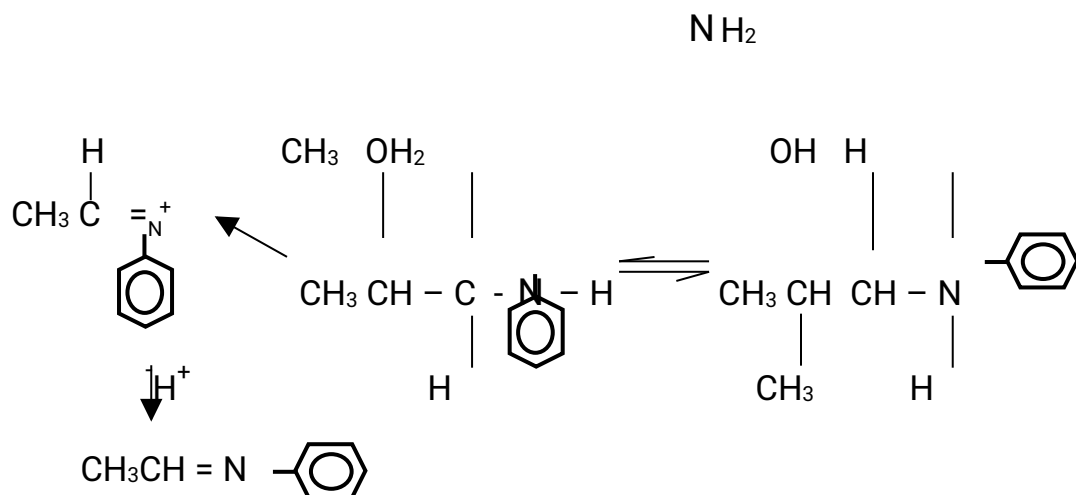




## Mechanisms of condensation reactions of carbonyl compounds



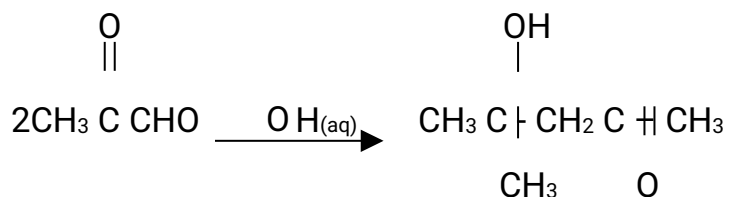
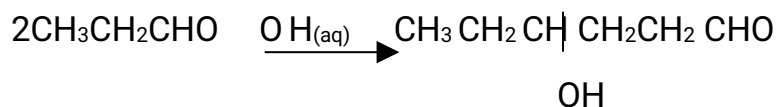
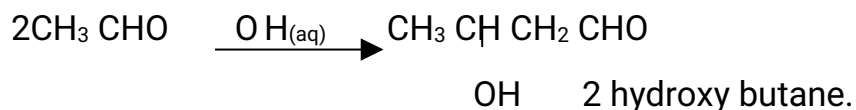




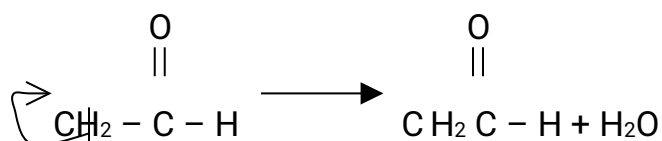
### Reactions with dilute NaOH

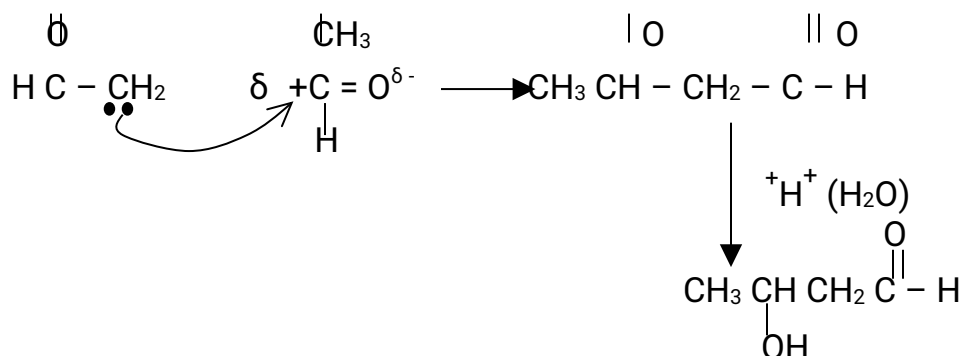
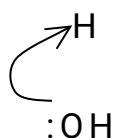
Aldehydes containing  $\alpha$ -hydrogens and ketones react with dilute NaOH to form hydroxyl alkanals or alkanones which are compounds containing both hydroxyl group and carbonyl carbon but with doubling of the chain.

When these compounds are heated, H<sub>2</sub>O is lost and a double bond is formed.



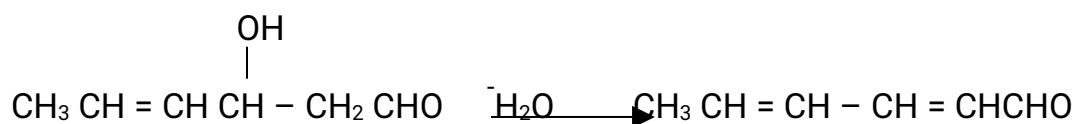
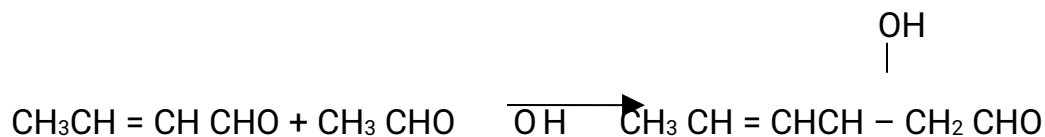
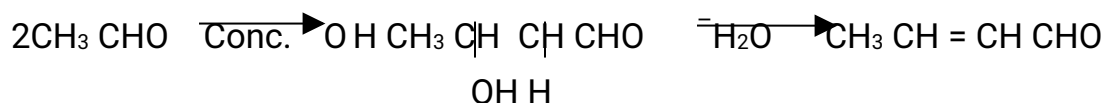
### Mechanism:



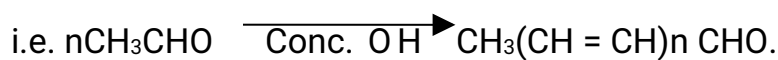


### Reactions with conc. NaOH

This reaction is possible only with aldehydes having  $\alpha$ -hydrogens and not with ketones to give brown resinous compounds formed by a series of condensation.



### Polymerization:



### Reaction that distinguish between aldehydes and ketones.

#### Oxidation:

Aldehydes are oxidized by oxidizing agents e.g.  $\text{H}^+_{(\text{aq})}/\text{K}_2\text{Cr}_2\text{O}_{7(\text{aq})}$   $\text{H}^+/\text{CrO}_3$  to

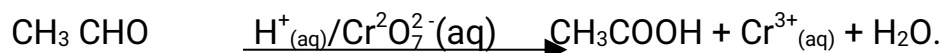
carboxylic acid.

Ketones are resistant to oxidation and therefore they never yield any product with the same reagents.

**Test:**

Add 2 drops of an aldehyde into 2cm<sup>3</sup> of H<sup>+</sup><sub>(aq)</sub>/K<sub>2</sub>Cr<sub>2</sub>O<sub>7(aq)</sub> and warm.

An orange solution turns to green, the resultant solution becomes acidic.

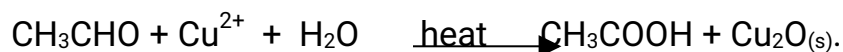


**Benedicts or Fehling's solution:**

Increases a solution containing Cu<sup>2+</sup> ions. This solution is blue in colour but when reacted with an aldehyde under heat, the aldehyde decreases Cu<sup>2+</sup> to Cu<sup>+</sup> resultant into a red-brown precipitate of Cu<sub>2</sub>O<sub>(s)</sub>.

This property of aldehyde is reduction.

Ketones do not give a red-brown precipitate.

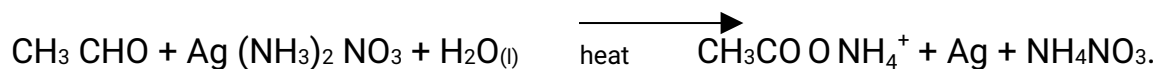


**Ammoniacal silver nitrate (Tollens reagents)**

Ammoniacal silver nitrate solution is a solution of silver nitrate in NH<sub>3</sub>.

Aldehydes unlike ketones are oxidized by this mild oxidizing agent to a carboxylic acid.

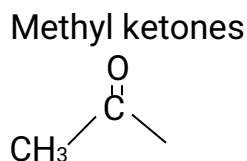
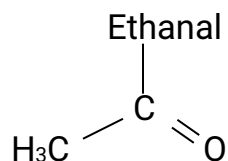
The Ag<sup>+</sup> in the AgNO<sub>3</sub> is reduced to Ag metals. Therefore, the observation made in practicals is; a grey deposit or silvery coating on the sides of the test tube.



Note that atmospheric O<sub>2</sub> can also oxidize aldehydes. This is why samples of aldehydes are normally contaminated when exposed to the atmosphere.

**Iodoform / haloform reaction:**

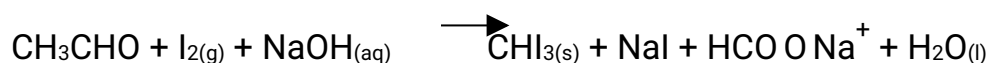
This reaction is only possible for methyl aldehydes and ketones. Since we have only one methyl aldehyde, ethanal, the reaction is strictly positive for ethanol and methyl ketones with the following structure.



The reaction is useful in practicals in determining the structure of carbonyl compounds.

### Test:

When iodine and NaOH is added to ethanol or a methyl ketone and warmed, a yellow precipitate of tri iodo methane is formed.

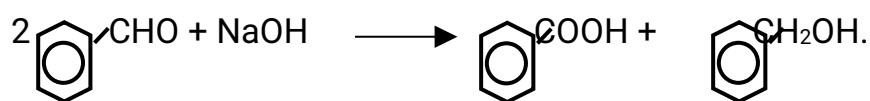


### CANNIZARO REACTIONS:

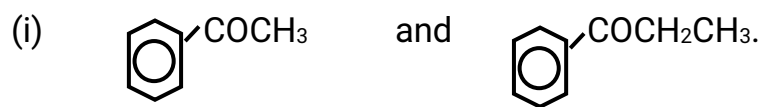
This reaction is only possible for aldehydes without  $\alpha$ -hydrogen ketones.

When NaOH is reacted with an aldehyde without an  $\alpha$ -hydrogen of general formula.

$\text{HCHO}$ ,  $\text{C}_6\text{H}_5\text{CHO}$ , the aldehyde is oxidized to carboxylic acid and the other half is decreased to the alcohol.





Name the reagent that can be used to distinguish between the following pairs of compounds and in each case, state what would be observed when the reagent is separately treated with each compound.



**Reagent:** Hot iodine solution in aqueous sodium hydroxide.



With  : A yellow precipitate.



With  : No observable change.

(ii)  $\text{CH}_3\text{COCH}_3$  and  $\text{HCHO}$

Reagent:


With  $\text{CH}_3\text{COCH}_3$

With  $\text{HCHO}$

(iii)   $\text{CHO}$  and   $\text{COCH}_2\text{CH}_3$

Reagent:

With   $\text{CHO}$

With   $\text{COCH}_2\text{CH}_3$

(iv)  $\text{CH}_3\text{CH}_2\text{COCH}_3$  and  $\text{HCHO}$

Reagent:

With  $\text{CH}_3\text{CH}_2\text{COCH}_2\text{CH}_3$


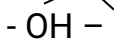
With  $\text{HCHO}$

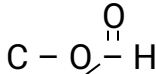
## CARBOXYLIC ACIDS

### CARBOXYLIC ACIDS/ALKANOIC ACIDS

Carboxylic acids are organic compounds with a general formula,  $\text{R} \overset{\text{O}}{\underset{\text{||}}{\text{C}}} - \text{OH}$

( $\text{C}_n\text{H}_{2n}\text{O}_2$ ) if they are saturated. They contain two functional groups:

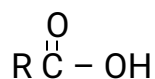
 Carbonyl carbon group  
 - OH - hydroxyl group

Carboxylic acid group  
  $\text{C} - \text{O} \overset{\text{O}}{\underset{\text{||}}{\text{C}}} - \text{H}$

Carboxylic acids can be roughly categorized into the following:

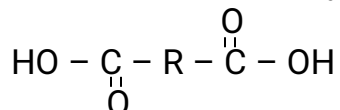
(i) **Monobasic acid:**

Contains a single carboxylic acid group.



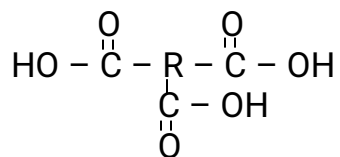
(ii) **Diabasic acid**

Contains two carboxylic acid groups joined to the same carbon chain.



(iii) **Tribasic acid**

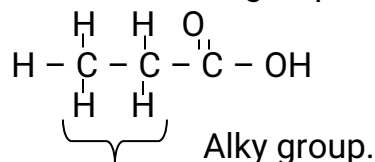
Contains three carboxylic acid groups joined to the same carbon chain.



**Nature of the alkyl group**

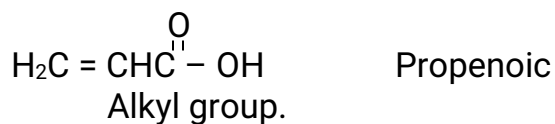
**Saturated carboxylic acid**

Contains the R group with single carbon-carbon bond.



**Unsaturated carboxylic acid**

Contains atleast a double bond in the alkyl chain.



**NOMENCLATURE**

According to the IUPAC system, carbonxylic acids are named after their corresponding alkanes. This is usually done by replacing a suffix-ane by OIC.

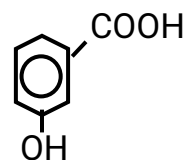
HCOOH                      Methanoic acid

CH<sub>3</sub>COOH                  Ethanoic acid

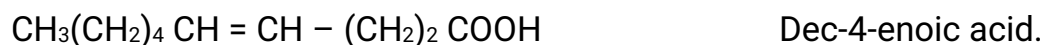
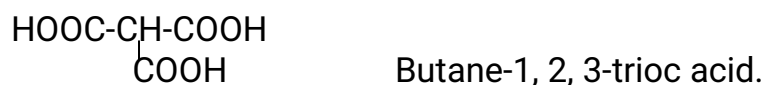
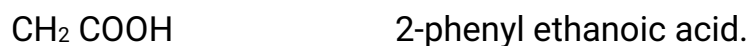
CH<sub>3</sub>CH COOH              2-methyl propan-1-oic  
CH<sub>3</sub>



Benzoic acid



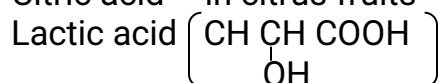
3-hydroxy benzoic acid



## MONOBASIC CARBOXYLIC ACIDS

These contain a single carboxylic acid group. They are largely found in nature. E.g.

Citric acid – in citrus fruits



Formic acid ( $\text{HCOOH}$ ) – insect bites.

## ISOMERISM

Monobasic acids show both structural and functional isomerism.

### STRUCTURAL

#### Chain isomerism:

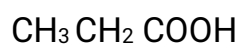
Isomers differ from nature of the carbon chain  $\text{C}_5\text{H}_{10}\text{O}_2$ .



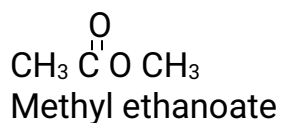


## FUNCTIONAL

Monocarboxylic acids are isomeric with esters e.g.  $\text{C}_3\text{H}_6\text{O}_2$



Propanoic acid



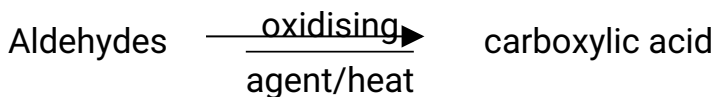
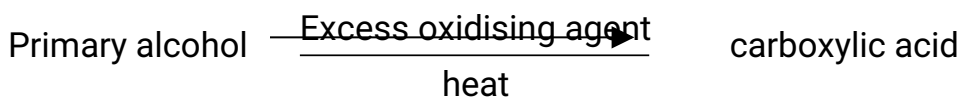
## METHODS OF PREPARATION OF MONOCARBOXYLIC ACIDS

Monocarboxylic acids can be prepared from the following:

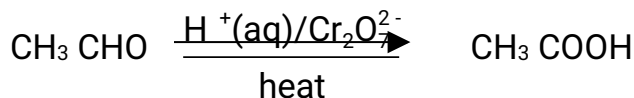
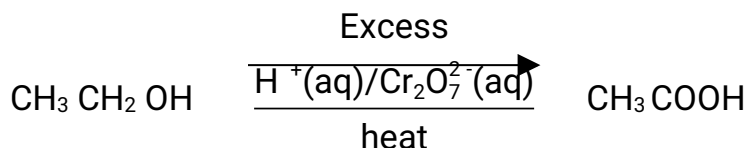
(i) **Oxidation of primary alcohols and aldehydes:**

Primary alcohols are oxidized by excess oxidizing agents until carboxylic acids are formed.

Aldehydes are also oxidized to carboxylic acids.

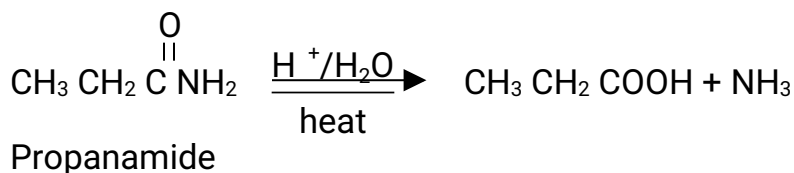
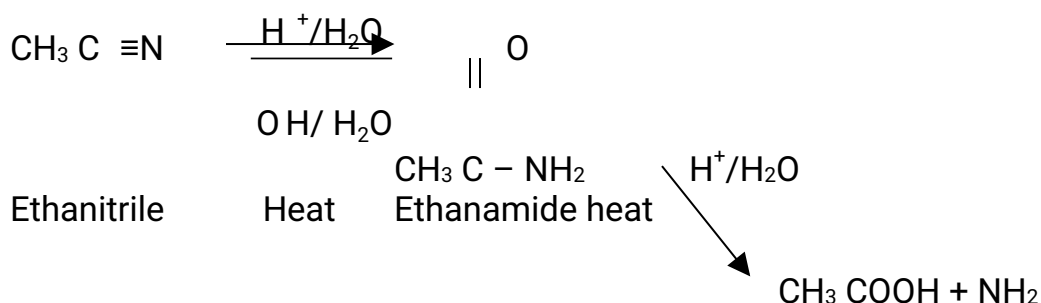


The main lab. Oxidizing agents used in this case are acidified  $\text{K}_2/\text{Na}_2\text{Cr}_2\text{O}_7/$   
 $\text{H}^+_{(\text{aq})}$  /  $\text{CrO}_{3(\text{aq})}$ ,  $\text{H}^+_{(\text{aq})}$  /  $\text{KMnO}_{4(\text{aq})}$

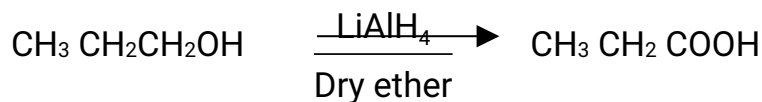
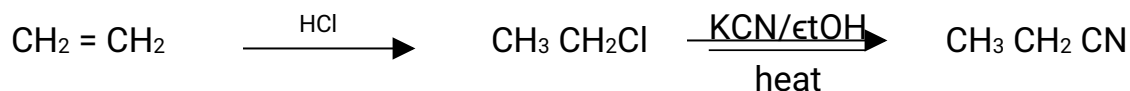
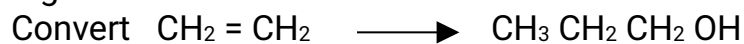


(ii) **Hydrolysis of acid nitriles and acid amides with alkaline or acid.**

When a nitrile is heated under reflux with mineral acid or alkali, an amide is first formed which later is further hydrolyzed to a carboxylic acid.

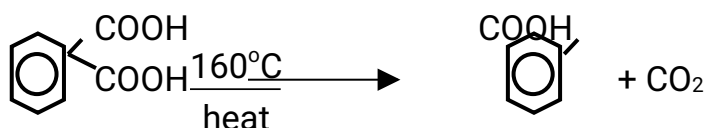
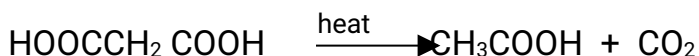
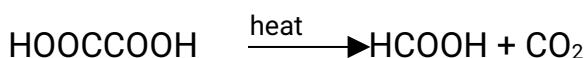


The usefulness of this reaction is, increase of the carbon chain of one carbon.  
E.g.



(iii) **From dicarboxylic acids**

When dicarboxylic acids are heated, they are decomposed by loss of  $\text{CO}_2$  to form a monocarboxylic acid. This reaction is important in reducing the carbon chain by one carbon.



## PROPERTIES OF CARBOXYLIC ACIDS

### Physical properties:

Lower members ( $\text{C}_1 - \text{C}_4$ ) are colourless liquids very soluble in  $\text{H}_2\text{O}$ .

Medium members ( $\text{C}_5 - \text{C}_9$ ) are solids only partially soluble in  $\text{H}_2\text{O}$ .

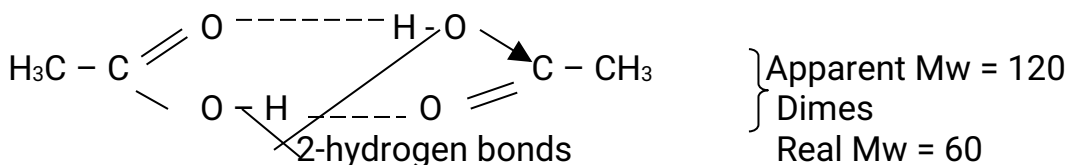
### SOLUBILITY OF CARBOXYLIC ACIDS IN $\text{H}_2\text{O}$ .

Solubility of carboxylic acids decreases with the increase in the carbon chain. This is because the more number of carbon chains which are hydrophobic outweighs the carboxylic acid group, which is responsible for forming  $\text{H}_2$  bonds with  $\text{H}_2\text{O}$  hence low solubility.

### Melting and boiling points:

Carboxylic acids show high values of melting and boiling points compared to their corresponding alcohols and alkanes. This is because a single molecule of a carboxylic acid dimerises via  $\text{H}_2$  bonds which makes their melting and boiling points to be higher.

By them undergoing dimerisation, their apparent molecular masses when determined by **cryoscopic method** to be twice the actual molecular mass.



The formation of  $\text{H}_2$  bonds also make them to be very soluble in  $\text{H}_2\text{O}$ .

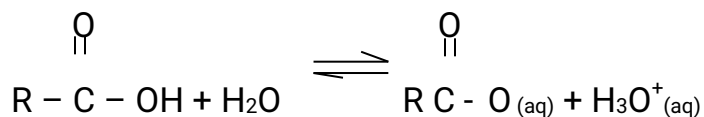
## CHEMICAL PROPERTIES

### Acid nature:

Carboxylic acids are acidic in nature. They are more acidic than phenol but less

acidic compared to mineral acids. Aqueous solutions of carboxylic acids turn litmus paper red.

When dissolved in H<sub>2</sub>O, they dissociate. An equation is established as shown.



$$K_a = \frac{[\text{RCOO}^-][\text{H}_3\text{O}^+]}{[\text{RCOOH}]} \text{ Mol dm}^{-3}.$$

The K<sub>a</sub> value is then used to denote the strength of the acid. The bigger the K<sub>a</sub> value, the more stronger the acid.

Acid	K <sub>a</sub>	PK <sub>a</sub>
Methanoic, HCOOH	1.77 x 10 <sup>-4</sup> mol dm <sup>-3</sup>	3.75
Ethanoic, CH <sub>3</sub> COOH	1.75 x 10 <sup>-5</sup> mol dm <sup>-3</sup>	4.82

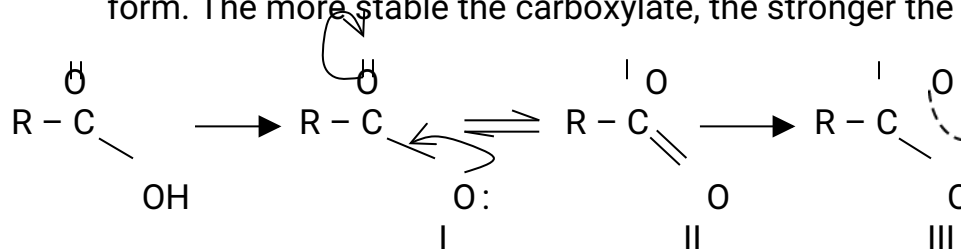
Another important value is PK<sub>a</sub>, therefore the higher the PK<sub>a</sub> value, the weaker the acid.

### Explanation/reason for acid strength

Acidity of carboxylic acids is determined by a number of factors;

#### (i) Resonance of the carboxylic ion

When the carboxylic acid group ionizes, it forms the carboxylic ion and this ion may show several resonance structures in order to attain a more stable form. The more stable the carboxylate, the stronger the acid.



I, II and III are resonance structures of carboxylic acid.

#### (ii) Inductive effect:

Inductive effect is a tendency of an atom or groups of atoms to donate or

withdraw electrons from a certain group.

Adjacent electron withdrawing substituents increase the acidity by further stabilizing the carboxylate. E.g.

Acids	Structure	PKa
Ethanoic acid	CH <sub>3</sub> CO <sub>2</sub> H	4.7
Fluoro ethanoic	FCH <sub>2</sub> CO <sub>2</sub> H	2.6
Dichloro ethanoic	ClCH <sub>2</sub> CO <sub>2</sub> H	1.3
Chloro ethanoic	Cl <sub>2</sub> CHCO <sub>2</sub> H	2.9
Tri-chloro ethanoic	Cl <sub>3</sub> CCO <sub>2</sub> H	0.9
Nitro ethanoic	NO <sub>2</sub> CH <sub>2</sub> CO <sub>2</sub> H	1.7

The above table shows that the presence of an electron withdrawing atom pulls electrons from the bond decreasing electron density of that bond (O-H) and the ease of a proton release.

More atoms of such groups decreases the electron density of the O – H bond further and stabilizing the carboxylate thus high acidity.

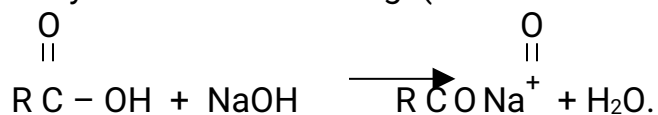
The data below shows the PKa of the following acids. Explain.

Acid	PKa
CH <sub>3</sub> COOH	4.7
CH <sub>3</sub> CH <sub>2</sub> COOH	4.9

## REACTIONS OF CARBOXYLIC ACIDS

They react as weak acids.

They react with bases e.g. (KOH and NaOH) to form a salt and H<sub>2</sub>O.

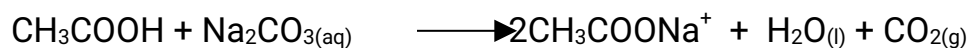




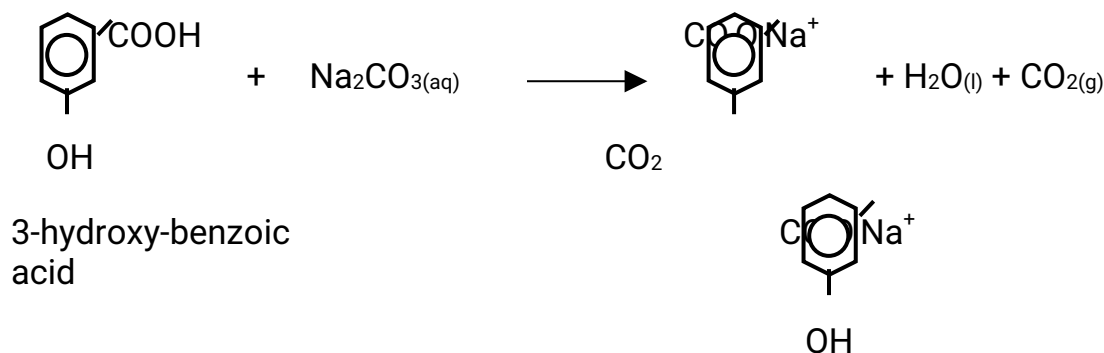


### REACTION WITH SODIUM CARBONATE ( $\text{Na}_2\text{CO}_3$ )

Carboxylic acids unlike phenol react with  $\text{Na}_2\text{CO}_3$  very slowly to liberate  $\text{CO}_{2(g)}$ . This reaction is of practical importance in distinguishing carboxylic acids from phenols. The reagent used is saturated  $\text{Na}_2\text{CO}_3$ .



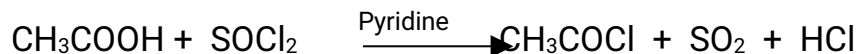
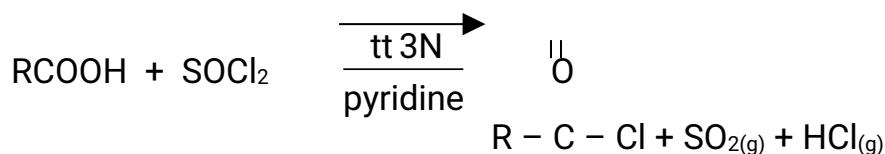
Similarly,  $\text{CO}_2$  can also be evolved from saturated  $\text{NaHCO}_3$ .



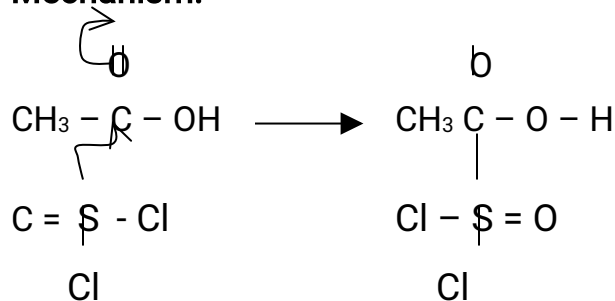
### OTHER REACTIONS:

#### With $\text{SOCl}_2$ .

Carboxylic acids react with  $\text{SOCl}_2$  in presence of an organic base e.g. pyridine to form an acid chloride and  $\text{SO}_2$ . This is one useful way of preparing acid chlorides.

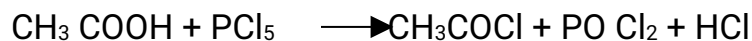
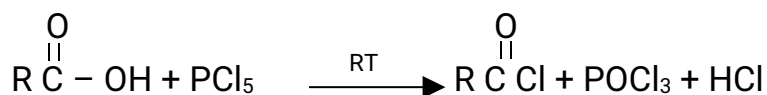


### Mechanism:



### Reaction with PCl<sub>5</sub>.

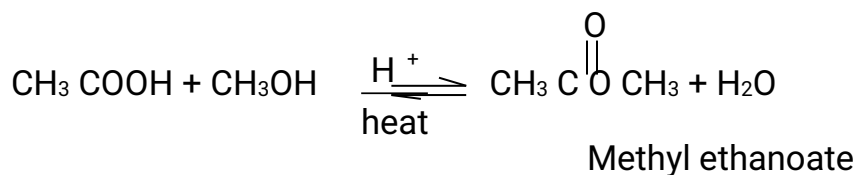
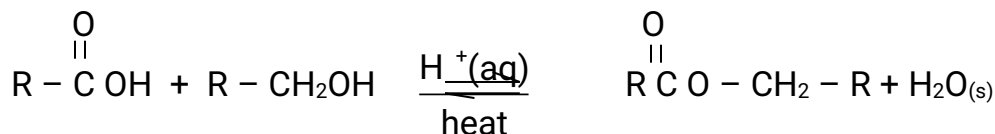
Carboxylic acids react with PCl<sub>3</sub> or PCl<sub>5</sub> to form acid halides.

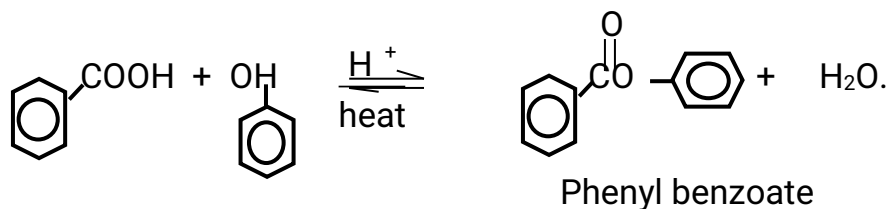
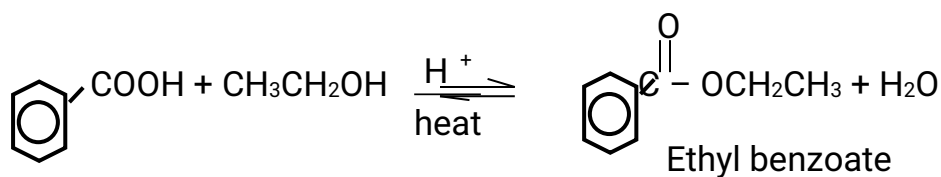


### Reaction with alcohols

React with alcohols in presence of a mineral acid to form an ester under heat. This is called esterification where an ester is formed by refluxing a carboxylic acid and alcohol in presence of an acid catalyst.

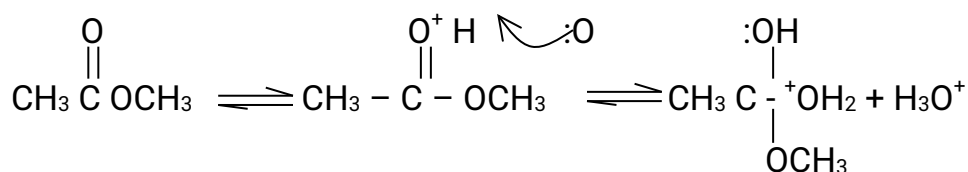
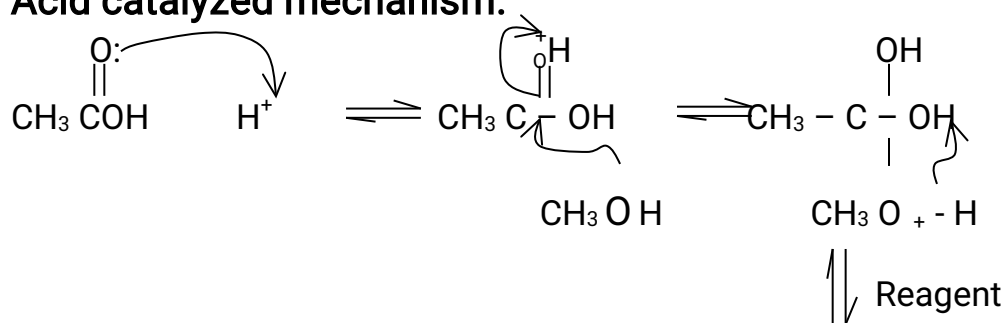
To complete the reaction, one of the components either an acid or an alcohol has to be in excess or to remove an ester at a certain interval once it is formed.





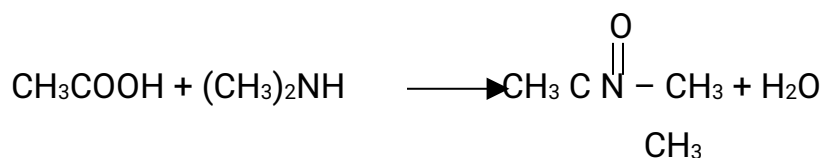
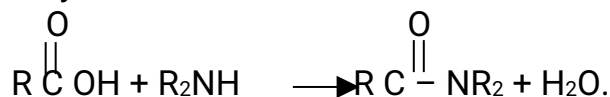
The mechanism for this reaction depends on the catalyst used/base catalysed or acid catalysed.

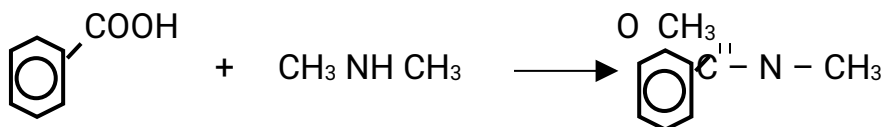
### Acid catalyzed mechanism.



### Reaction with amides:

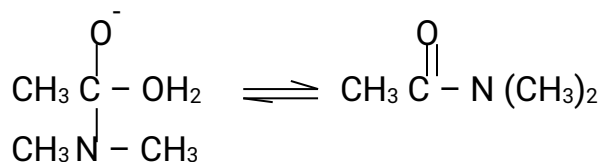
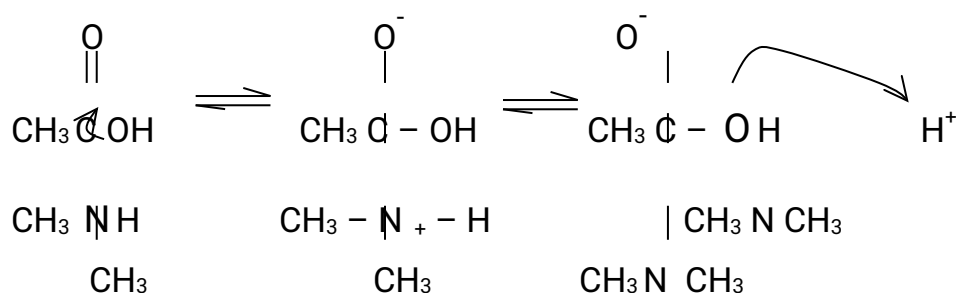
Carboxylic acids react with amides to form substituted amines. With amines, they form amides.





This reaction occurs because of the presence of a proton on the amine which are substituted by the carboxylate.

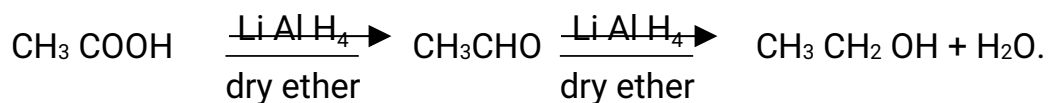
Mechanism:



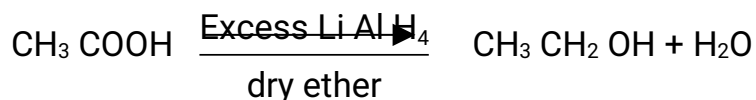
## REDUCTION OF CARBOXYLIC ACIDS

Carboxylic acids are usually reduced by reducing agents e.g.  $\text{LiAlH}_4$ /ether (THF) to form aldehydes and later alcohols (primary).

**Note:**  *$\text{NaBH}_4$  is not commonly used to decrease carboxylic acids because it is less reactive.*

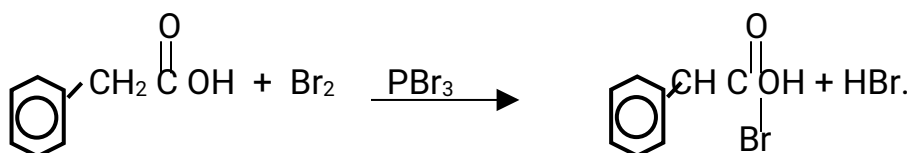
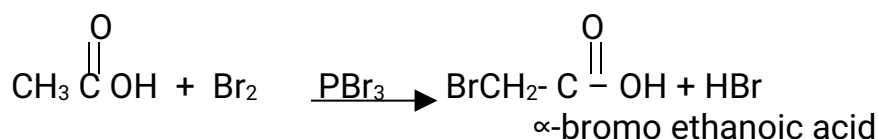


OR:



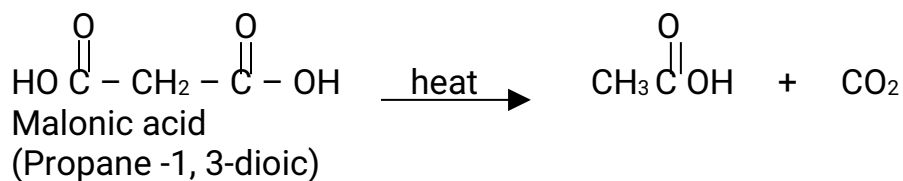
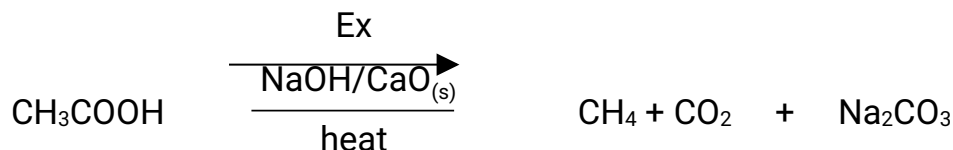
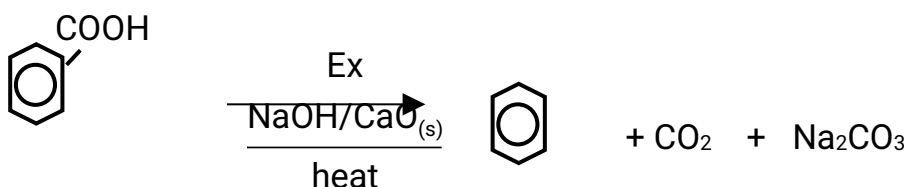
### REACTIONS WITH HALOGENS ( $\alpha$ -HALOGENATIONS)

Carboxylic acids can be halogenated at the carbon atom adjacent to the carboxylic group when reacted with reagents like  $\text{Br}_2$ ,  $\text{PCl}_5$ . This reaction depends on the character of the carbonyl compound where the product of the reaction known as  $\alpha$ -bromo carboxylic acid, is converted to  $\alpha$ -hydroxy or  $\alpha$ -amino carboxylic acid.



### DECARBOXYLATION

Carboxylic acids with a carbonyl group at the third position readily undergoes thermal decarboxylation where  $\text{CO}_2$  is lost to form a simple alkane when heated in the presence of dry soda lime.

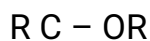


Carboxylic acid derivatives

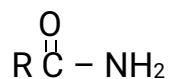
Esters

||

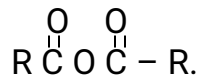
O



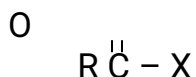
Amides



Anhydrides

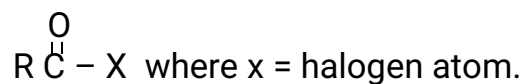


Acid halides



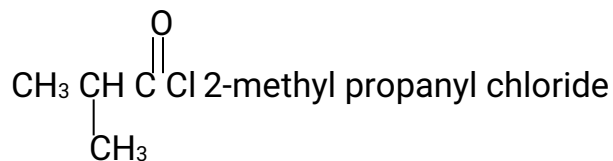
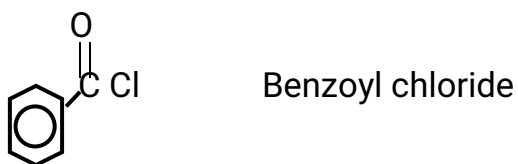
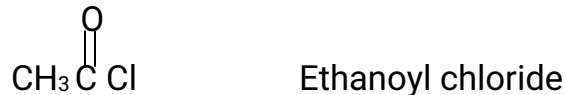
## ACID HALIDES

These are compounds derived from carboxylic acids by reacting an acid with a halogen. The commonest examples are acid chlorides.



### Nomenclature:

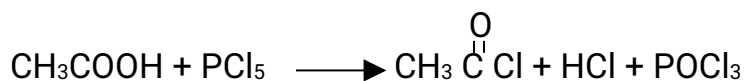
Acid chlorides are named by replacing the suffix -ic in acids with -oyl.



## PREPARATION

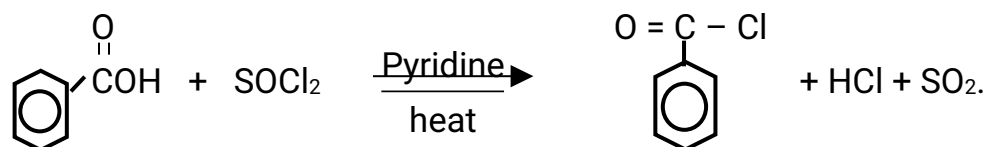
i) **Reacting carboxylic acids with phosphorous halides.**





ii) **Reaction of carboxylic acid with thionyl chloride**

This reaction should occur in a fume cupboard due to evolution of toxic acids or should occur in the presence of pyridine which absorbs.



## PROPERTIES OF ACID HALIDES

### Physical:

They are colourless volatile liquids with irritating smells. They fume easily in moist air due to evolution of HCl gas.

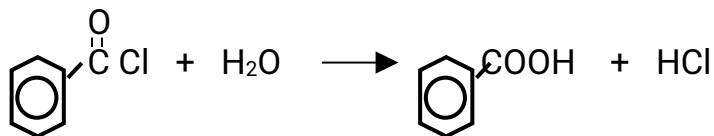
Have low boiling points than the corresponding acids.

### Chemical properties:

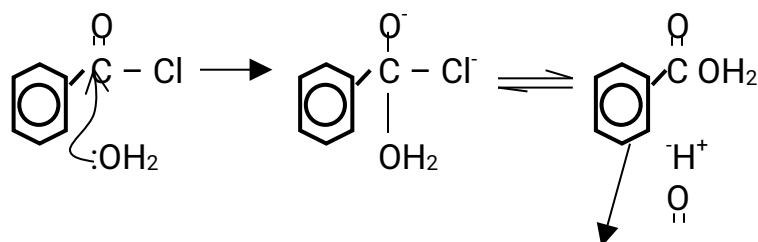
They undergo a number of reactions making them suitable for organic synthesis.

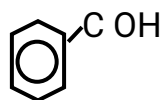
### Hydrolysis using H<sub>2</sub>O.

Acid halides react with H<sub>2</sub>O to form their parent carboxylic acids.



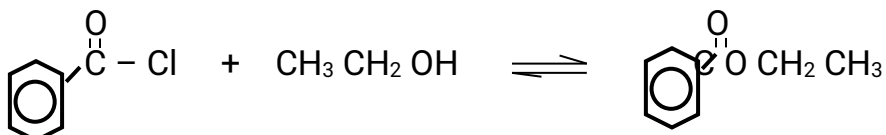
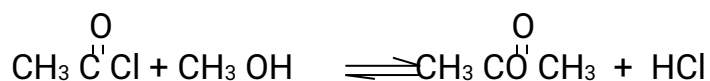
### Mechanism:



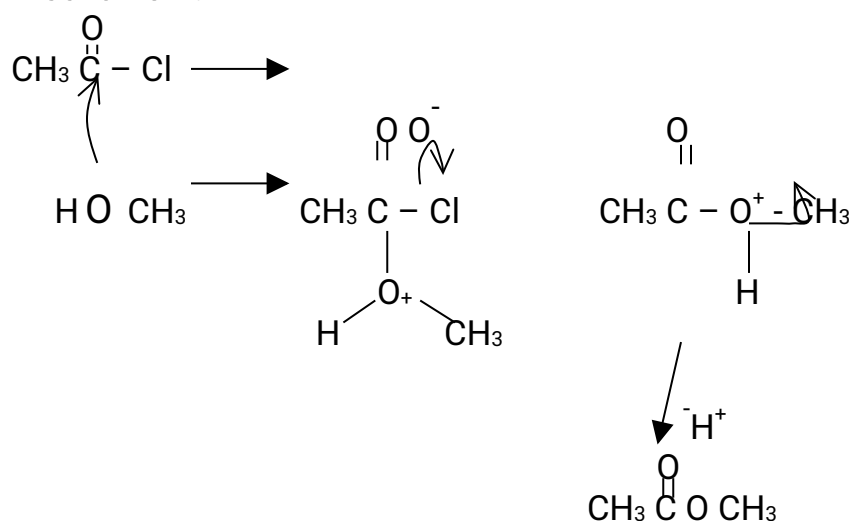


### Reaction with alcohols:

Acid chlorides react with alcohols to form esters.

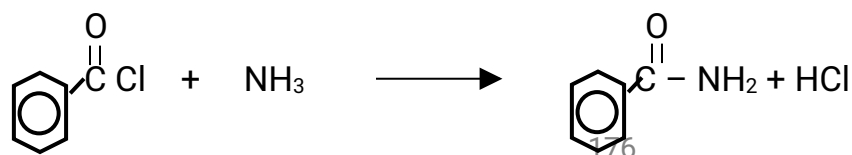
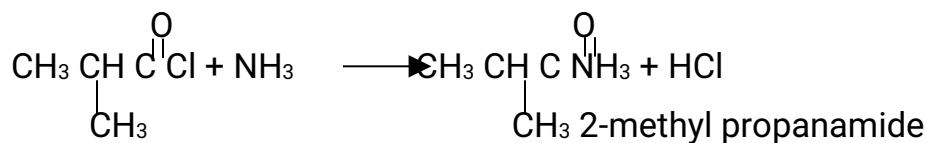
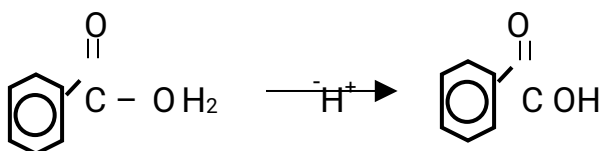


### Mechanism:



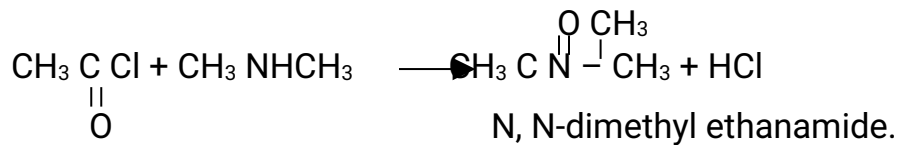
### Reaction with $\text{NH}_3$ and amines:

Acid chlorides react with  $\text{NH}_3$  to form amides. This reaction is responsible with primary/secondary amines to form amides.

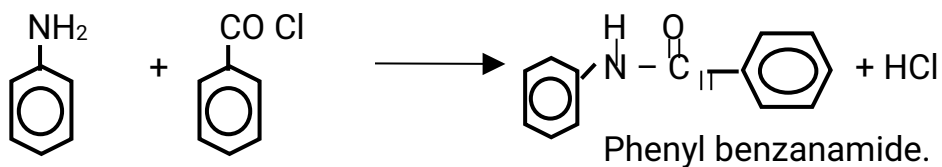
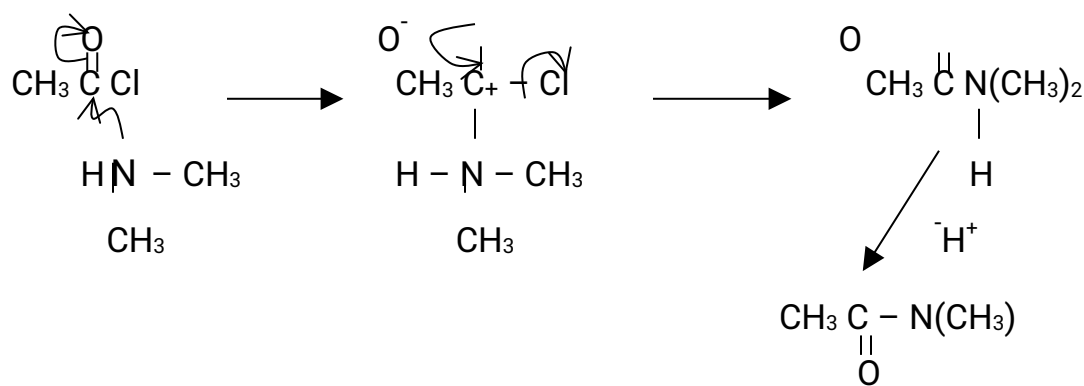




### Benzanamide

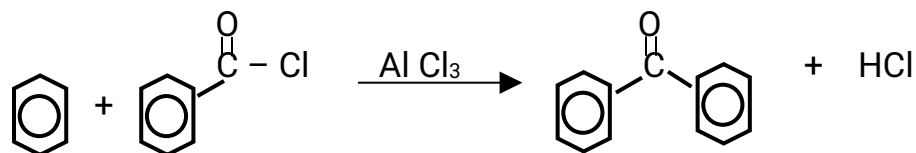
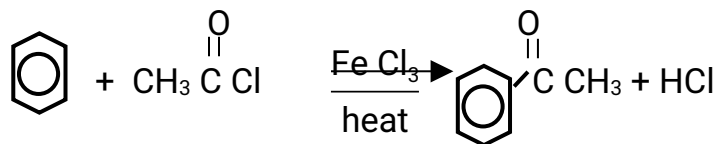


### Mechanism:



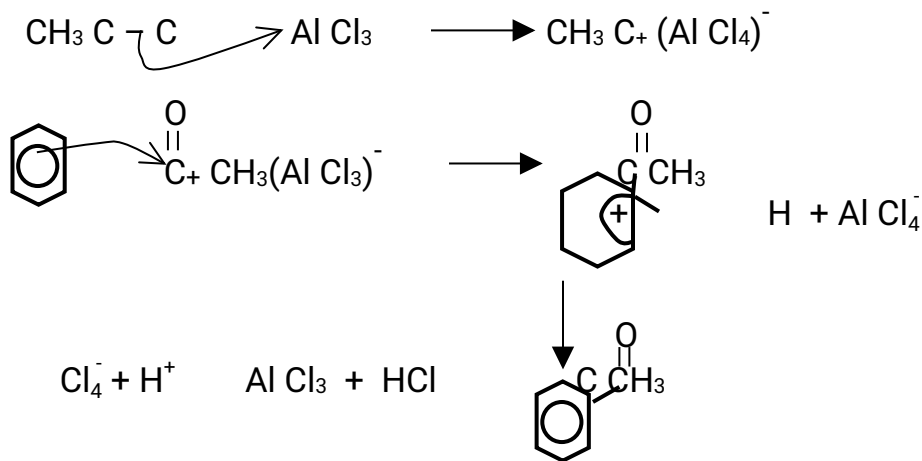
### Reaction with benzene:

Acid halides react with benzene in presence of a hydrogen carrier.



### Mechanism:





## AMIDES

Amides are compounds derived from carboxylic acid and nitrogen containing compound like amine or  $\text{NH}_3$  with a general formula  $\text{R}-\overset{\text{O}}{\underset{\text{||}}{\text{C}}}-\text{NH}_2$ .

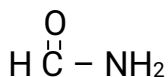
Unlike amines, they contain a carbonyl carbon directly attached to the nitrogen.

Naturally, such compounds are found in proteins.

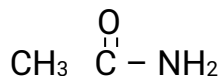
### Nomenclature:

Amides are named as derivatives of parent hydrocarbons alkane, replacing -e - amide.

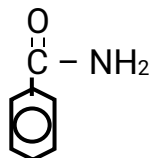
E.g.



Methanamide



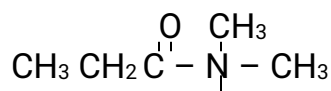
Ethanamide



Benzenamide



N, N-dimethyl ethanamide

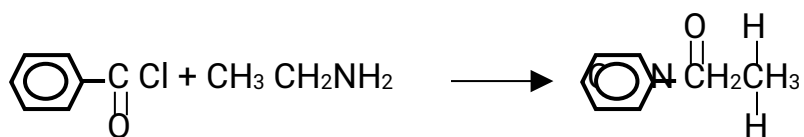


N, N , N – trimethyl propanamide.



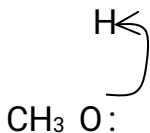
All amides are colourless crystalline solids except ethanamide. They have got higher boiling points than corresponding carboxylic acids due to the formation of H<sub>2</sub> bonding.

### Reaction between acid halides and $\text{NH}_3$ or amines.

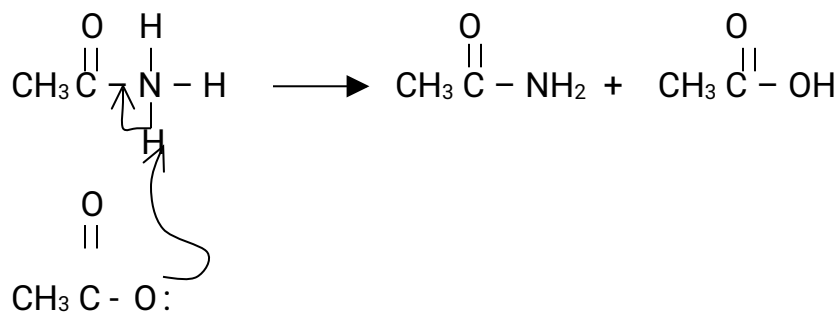
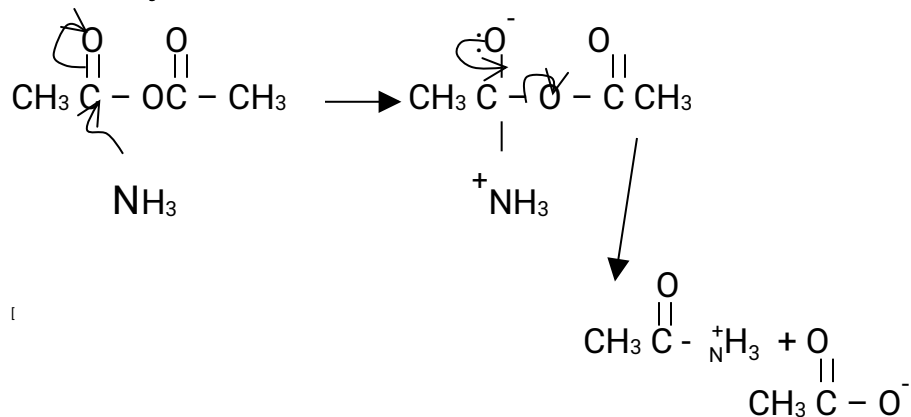

$$\text{CH}_3 \overset{\text{O}}{\parallel} \text{C} \text{O} \text{CH}_3 + \text{NH}_3 \xrightarrow{\text{heat}} \text{CH}_3 \overset{\text{O}}{\parallel} \text{C} \text{NH}_2 + \text{CH}_3 \text{OH}$$
$$\text{CH}_3\overset{\text{O}}{\parallel}\text{C}-\text{O}-\overset{\text{O}}{\parallel}\text{C}\text{CH}_3 + 2\text{NH}_3 \xrightarrow{\text{heat}} 2\text{CH}_3\overset{\text{O}}{\parallel}\text{C}\text{NH}_2 + \text{H}_2\text{O}$$

Ethanoic acid anhydride.

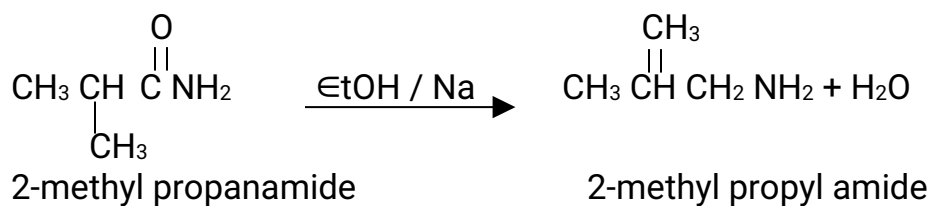
$$\begin{array}{c}
 \text{CH}_3\text{C}(=\text{O})\text{OCH}_3 + \text{NH}_3 \longrightarrow \text{CH}_3\text{C}(\text{O}^-)(\text{NH}_3^+)\text{OCH}_3 \longrightarrow \text{CH}_3\text{C}(=\text{O})\text{NH}_3^+ + \text{O}^-\text{CH}_3 \\
 \text{O} \\
 \text{CH}_3\text{C}(=\text{O})\text{NH}_3^+ + \text{H}_2\text{O} \longrightarrow \text{CH}_3\text{C}(=\text{O})\text{NH}_2 + \text{CH}_3\text{OH}
 \end{array}$$



## With anhydrides



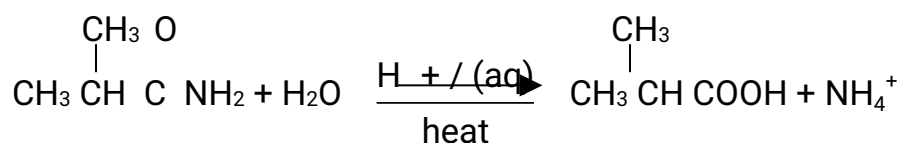
Amides are reduced by reducing agents such as sodium tetrahydride borate,  $\text{LiAlH}_4$ .



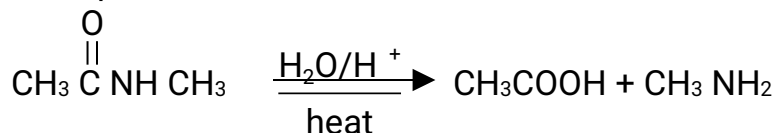
## Hydrolysis

Amides are hydrolysed in presence of a mineral acid or an alkali to form a carboxylic acid.

For example:



Example:



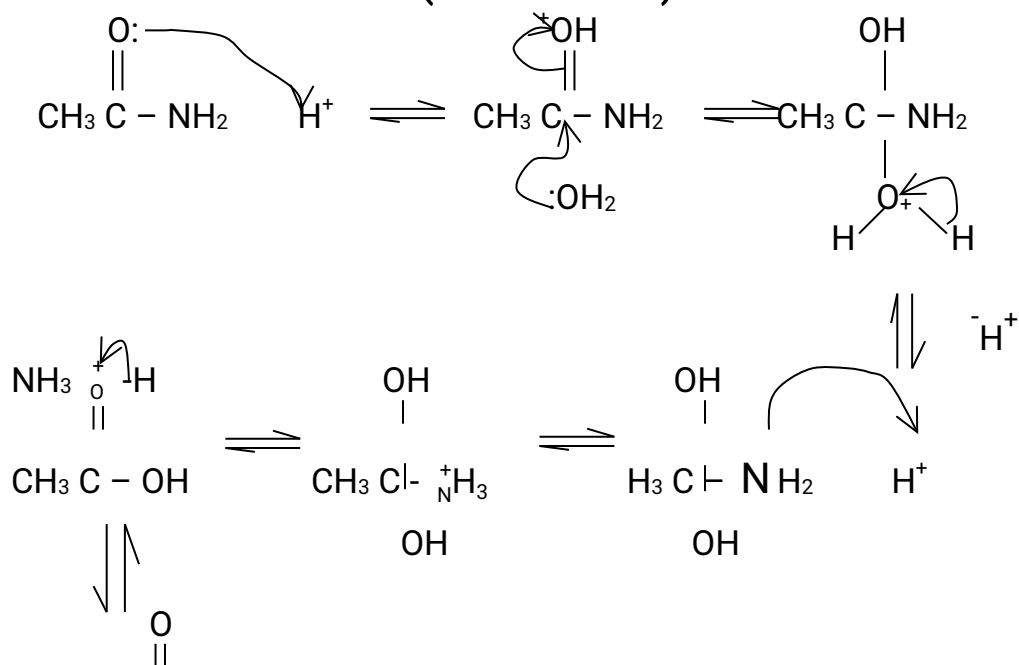
Primary amine

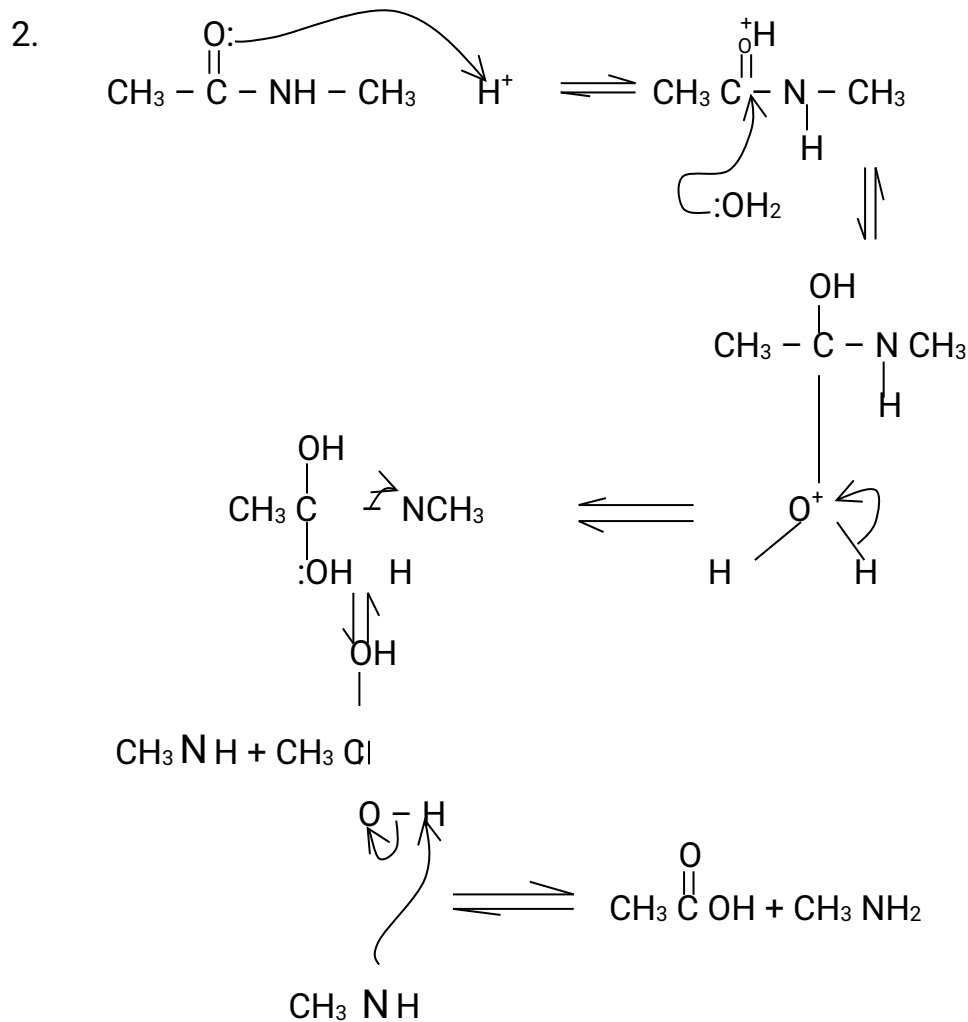
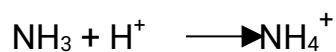
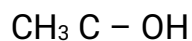
$$\text{CH}_3\text{CON}(\text{CH}_3)_2 + \text{H}_2\text{O} \xrightarrow[\text{heat}]{\text{OH}^-} \text{CH}_3\text{COOH} + \text{NH}(\text{CH}_3)_2$$

Secondary amine.

## ACID CATALYZED HYDROLYSIS (MECHANISMS)

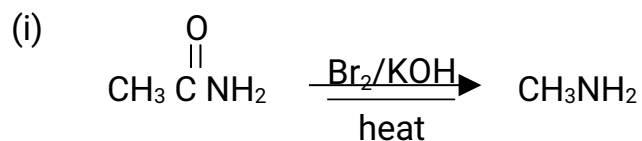
1.



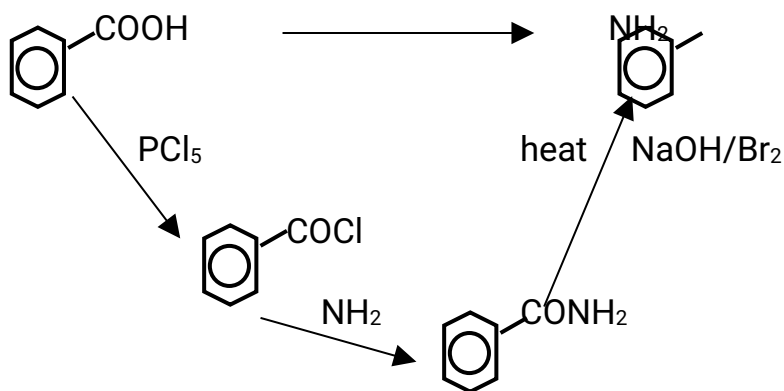
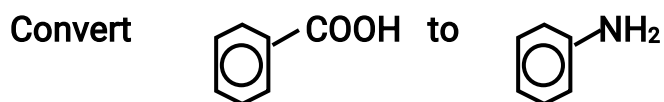
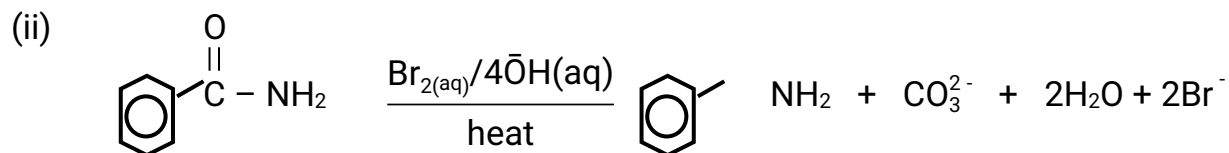
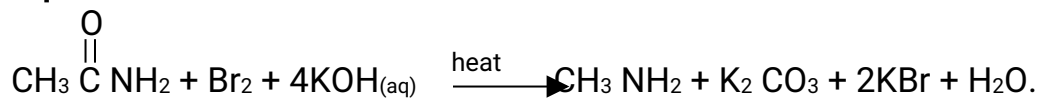


## HOFFMAN'S DEGRADATION

When an amide is heated with  $\text{Br}_2$  and an alkali, a primary amine is produced which is less than one carbon from the original amide; this reaction is known as Hoffman's degradation because it involves reduction of the carbon chain by one carbon atom.

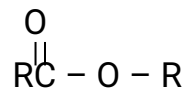


**Equation:**



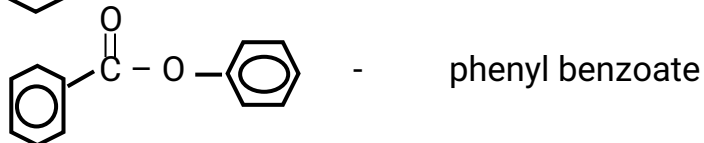
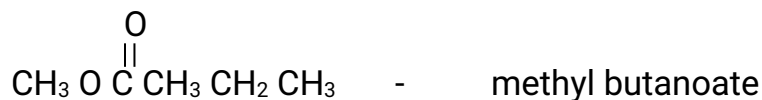
## ESTERS

Esters are derivatives of carboxylic acids. Are highly volatile compounds with a fruity smell. They conform to a general formula:



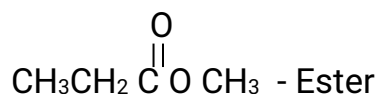
## NOMENCLATURE

Esters are named using their parent acid name but adding the suffix –ate.



### Isomerism

Esters show structural isomerism and functional isomerism. Functionally esters are isomeric with carboxylic acids.

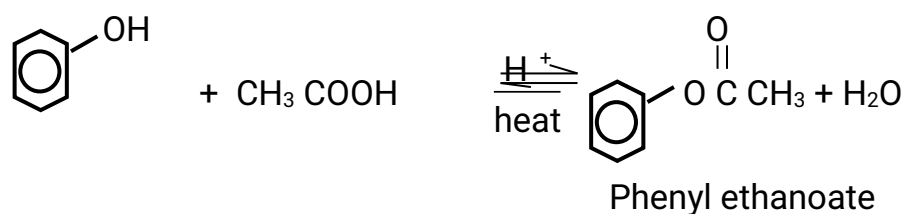
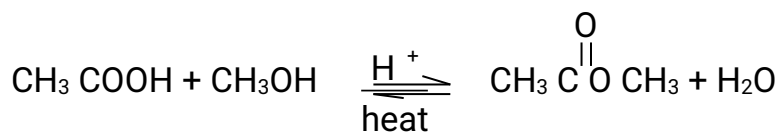


## PREPARATION OF ESTERS

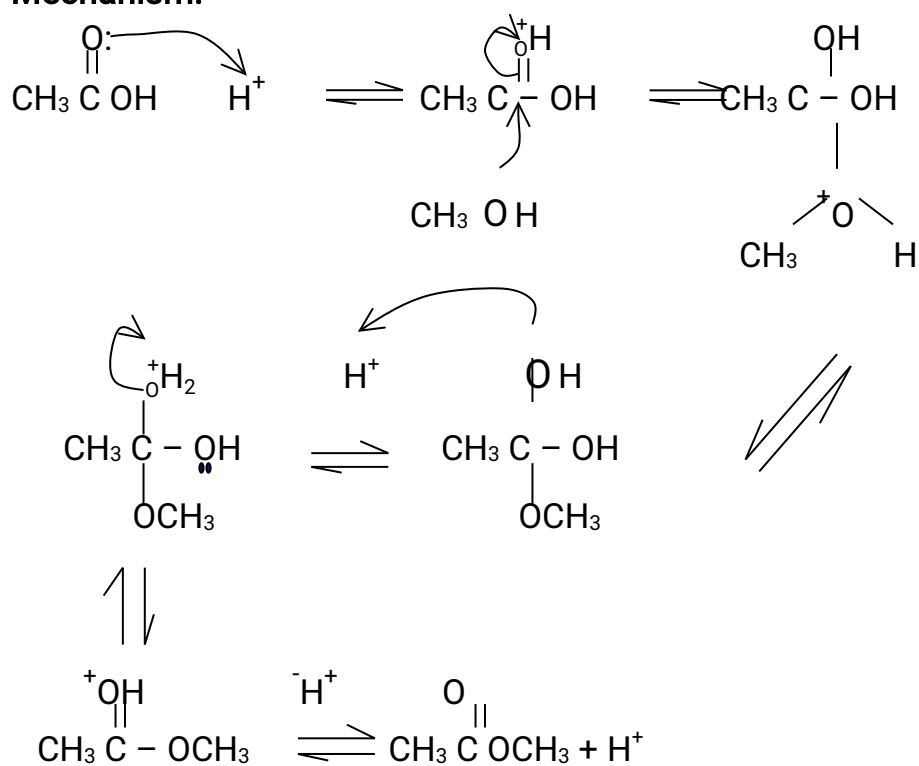
### (i) Esterification:

From carboxylic acids and an alcohol. This reaction is catalyzed by a mineral acid or an alkali.



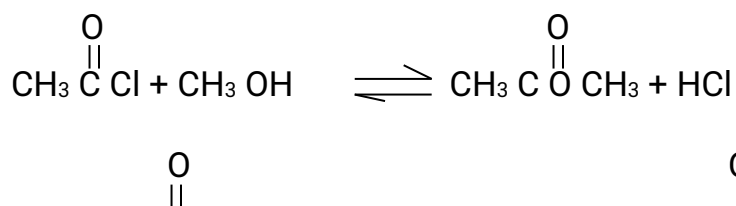


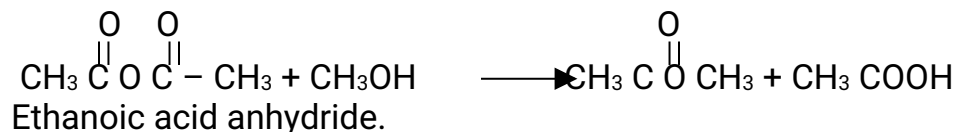
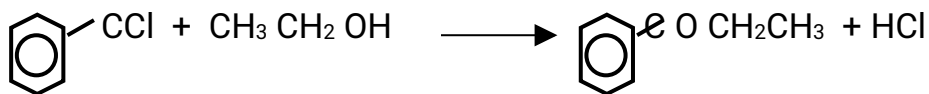
### Mechanism:



### From acid chlorides and acid anhydrides

When alcohols are reacted with acid chlorides or acid anhydrides, an ester is formed.



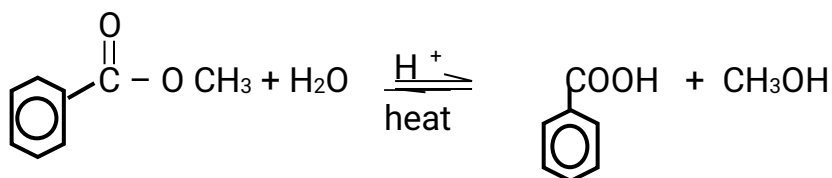
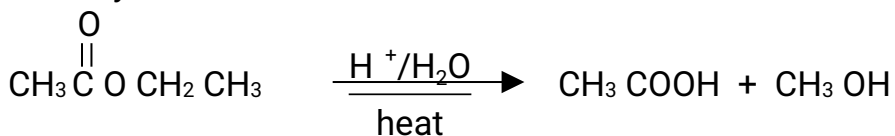


## REACTIONS OF ESTERS

Esters undergo the following reactions:

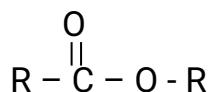
### i) Hydrolysis

Esters are hydrolyzed in presence of a mineral acid to a corresponding carboxylic acid and alcohol.



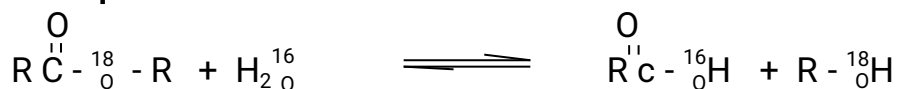
### NOTE:

Hydrolysis involves cleavage at



The position of this cleavage is useful in radiolabelling to trace for the reaction mechanisms.

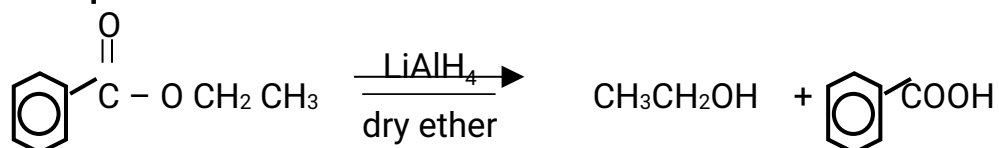
### Example:



### ii) Reduction

Esters are reduced with  $\text{LiAlH}_4$  in presence of ether to carboxylic acid and the alcohol.

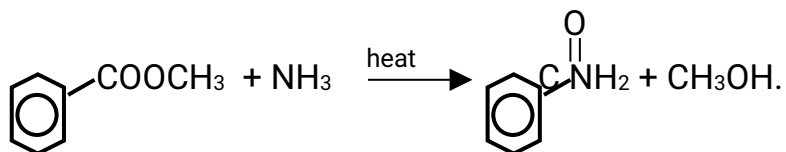
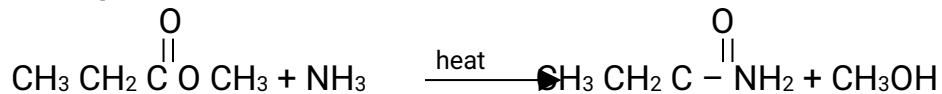
### Example:



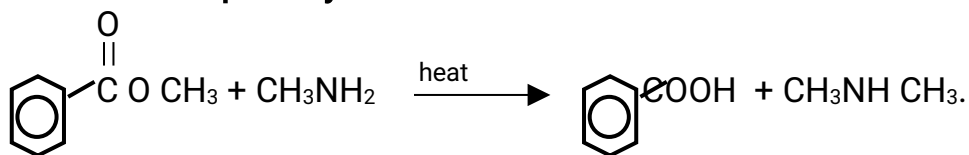
iii) **Reaction with NH<sub>3</sub>**

Esters react with NH<sub>3</sub> to give amides and alcohols.

**Example:**



iv) **Reaction with primary amines**



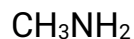
## AMINES

These are compounds that are derived from ammonia base by replacing one hydrogen with an amino group. They therefore have a general formula.  
RCH<sub>2</sub> – NH<sub>2</sub>.

The functional group of amines is – NH<sub>2</sub> (amino group).

### NOMENCLATURE

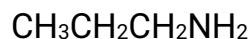
Amines are named as derivatives of alkanes by adding amine suffix to the stem name.



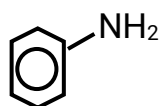
methyl amine



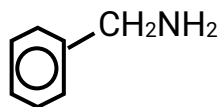
ethyl amine



propyl amine



phenyl-amine (aniline)



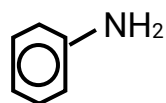
phenyl methyl amine

## CLASSES OF AMINES

Amines are classified into four groups or four classes.

### i) **Primary amines**

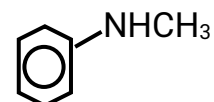
Is formed when only one hydrogen is replaced from the  $\text{NH}_3$ .



Phenyl amine (primary aromatic)

### ii) **Secondary amines**

Is formed when two hydrogens of  $\text{NH}_3$  are replaced by alkyl or anyl groups.



Phenyl methyl amine

### iii) **Tertiary amines**

Are formed when all the hydrogens in the  $\text{NH}_3$  are replaced by the alkyl group.



### iv) **Quartenary amines**

Are salts formed when the lone pair on the Nitrogen is donated to an alkyl group.



## Isomerism

Amines exhibit all the three types of isomerism, chain, positional and functional.

Write all isomers of  $\text{C}_4\text{H}_{11}\text{N}$ .



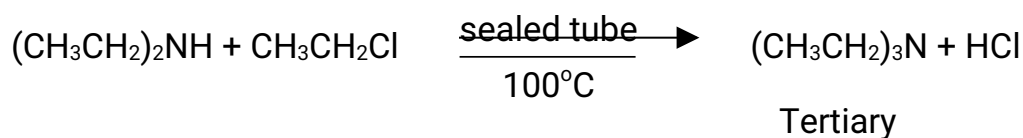
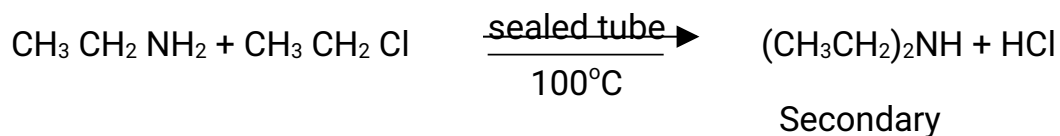
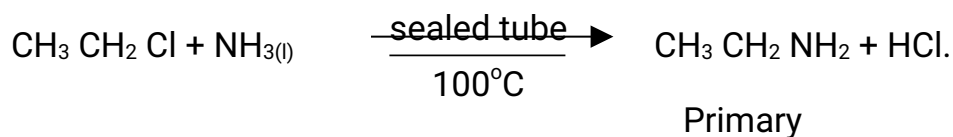
2.  $\begin{array}{c} \text{CH}_3 \text{ CH CH}_2 \text{ NH}_2 \\ | \\ \text{CH}_3 \end{array}$  2-methyl propylamine
3.  $\begin{array}{c} \text{CH}_3 \\ | \\ \text{CH}_3 - \text{C} - \text{NH}_2 \\ | \\ \text{CH}_3 \end{array}$  2-amino-2-methyl propane.
4.  $\text{CH}_3 \text{ CH}_2 \text{ CH CH}_3$  2-amino butane
- 1 and 2 are chain isomers.  
3 and 4 are position isomers.
5.  $\begin{array}{c} \text{H} \\ | \\ \text{CH}_3 \text{ CH}_2 \text{ N CH}_2 \text{ CH}_3 \end{array}$  N-diethyl amine (functional)
6.  $\begin{array}{c} \text{CH}_3 \\ | \\ \text{CH}_3 \text{ N CH}_2 \text{ CH}_3 \end{array}$  (Functional)

## METHODS OF PREPARATION

### From alkyl halides

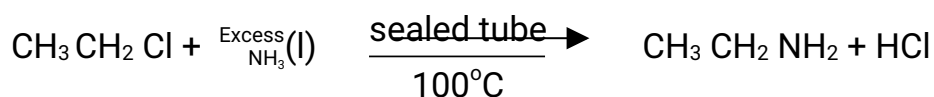
When alkyl halides are heated with  $\text{NH}_3$  in a sealed tube at a temperature of  $100^\circ\text{C}$ , the reaction gives a mixture of amines. This method is not suitable for producing a specific amine because of a mixture of amines.

Although the products produced may be controlled by using excess  $\text{NH}_3$  and at the same time separating each product by functional distillation.

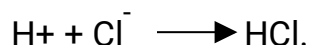
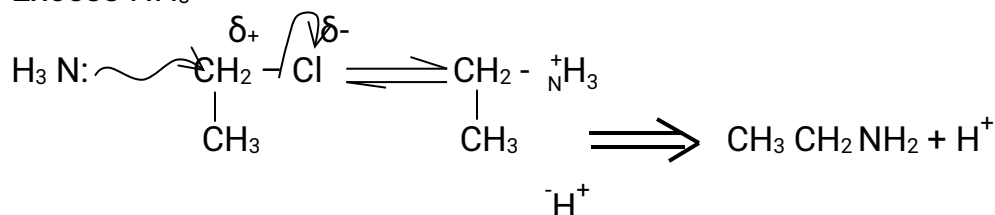


**NOTE:**

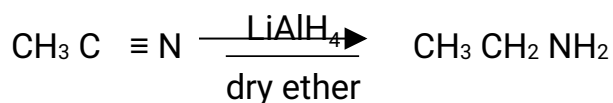
When excess  $\text{NH}_3$  is used,

**Mechanism:**

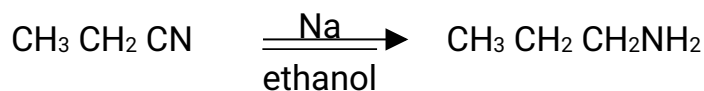
Excess  $\text{NH}_3$

**From alkyl cyanides**

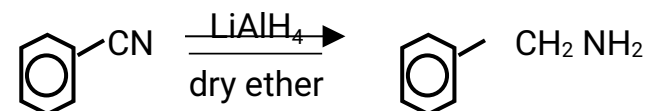
When an alkyl cyanide also known as a nitrile is reduced with reducing agents such as  $\text{LiAlH}_4$  or  $\text{NaBH}_4$ . Or simply sodium in an alcohol, dry ether a primary amine is formed.



Ethanitrile



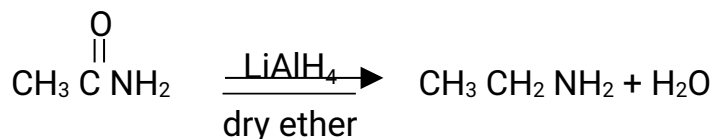
Propanitrile

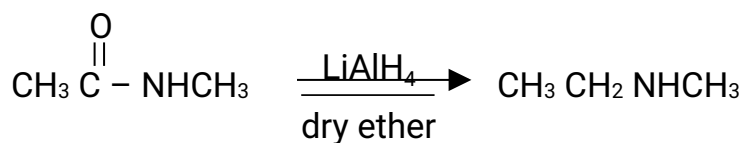


Benzenitrile

**From acid amides**

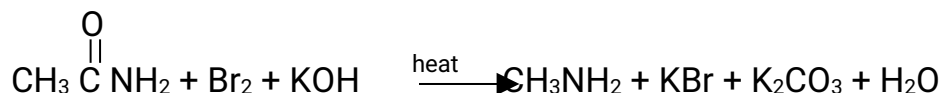
When an acid amide is reduced using reducing agents like  $\text{LiAlH}_4$ /dry ether, primary amines are formed.





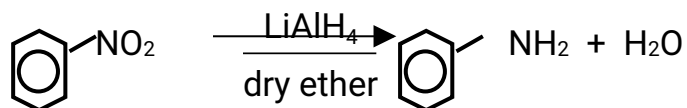
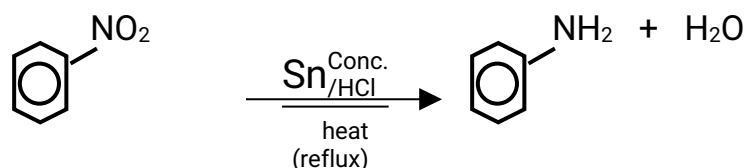
### HOFFMAN'S DEGRADATION

When amines are reacted with Br<sub>2</sub> in a solution of KOH or NaOH, a primary amine is formed which is one carbonless from the parent amide. This reaction shortens the chain by one carbon and it is known as Hoffman's degradation.



### From nitro compounds

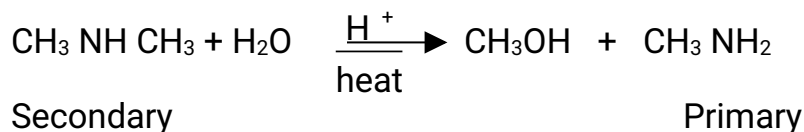
When a nitro compound preferably aromatic nitro compound is reacted with tin in conc. HCl or LiAlH<sub>4</sub>/ether, aromatic amine is formed.

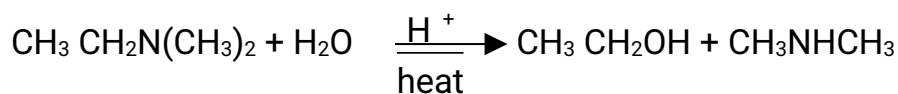


This method is suitable for preparing aromatic amines.

### HYDROLYSIS OF TERTIARY AND SECONDARY AMINES

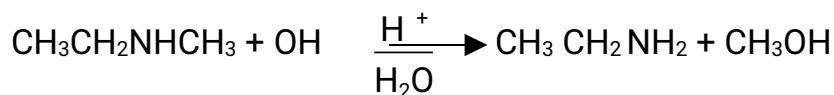
This produces primary and secondary amines respectively. When they are heated with mineralized acid H<sub>2</sub>O or a base.





Tertiary

Secondary



Mechanism: (Base catalysed)

### PHYSICAL PROPERTIES OF AMINES

- Lower amines up to  $\text{C}_2$  are gases while those with  $\text{C}_3 = \text{C}_{11}$  are liquids, higher amines are solids at room temperature.
- Lower amines have got a fishy ammoniacal smell or odour.
- Boiling points of amines are higher than those of corresponding alkanes.
- Lower members are soluble in  $\text{H}_2\text{O}$  but the solubility decreases with increase in molecular weight.

### Explanations for boiling points and solubility

Question:

Methyl amine $\text{CH}_3\text{NH}_2$	Molecular weight 31	Boiling point $^\circ\text{C}$ -7
$\text{CH}_3\text{CH}_3$	30	-89

Methyl amine boils at a high temperature than ethane because of the presence of intramolecular hydrogen bonding which holds methylamine molecules together in the liquid phase and difficult to escape to vapour phase during boiling and more heat required to boil while in ethane, the weak Van der Waal's forces are easily broken and less heat required to boil.

Comparing three classes of amines, primary amine boils at a higher temperature than the secondary followed by tertiary.

Amine	Molecular weight	Boiling point $^\circ\text{C}$
$\text{CH}_3\text{CH}_2\text{CH}_2\text{NH}_2$	59	49
$\text{CH}_3\text{CH}_2\text{NHCH}_3$	59	35
$\text{CH}_3 - \underset{\text{CH}_3}{\text{N}} - \text{CH}_3$	59	3.5



Primary amine forms atleast two intramolecular hydrogen bonding due to the presence of more hydrogen atoms on the nitrogen and hence boiling at a higher temperature.

The secondary amine forms only one hydrogen bond since it has only one hydrogen in the nitrogen while the tertiary amine has not hydrogen on the nitrogen and therefore forming no hydrogen bonding and ends boiling at the lowest temperature.

### Assignment (Research work)

Carboxylic acids boil at higher temperatures than alcohols which also boils at a higher temperature than an amine and an alkane has the least boiling point. Explain.

## BASICITY OF AMINES

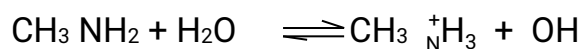
Amines are bases which form weak alkaline solutions when dissolved in H<sub>2</sub>O. Such solutions turn litmus paper to blue.

The basic strength of an amine is determined by the extent of ionization denoted as K<sub>b</sub>. The greater the K<sub>b</sub> value, the more basic the amine. The basic character of an amine can be explained by the presence of lone pair of electrons on the nitrogen atom which are always donated.

The more available these lone pair of electrons are, the more basic an amine is.

The availability of electrons on the nitrogen atom is determined by the atoms that are bonded onto it.

Groups of atoms that have got positive inductive effect will push electrons towards nitrogen making them available while those with negative inductive effect will lower the electron density on the nitrogen atom making them less available.



$$K_b = \frac{[\text{CH}_3\overset{+}{\text{N}}\text{H}_3][\text{OH}^-]}{[\text{CH}_3\text{NH}_2]}$$

### Questions:

1. Methylamine is a more stronger base than ammonia.

- Phenylamine is a weaker base than methylamine.
- Ethylmethanamine is a stronger base than trimethylamine.

Trend:

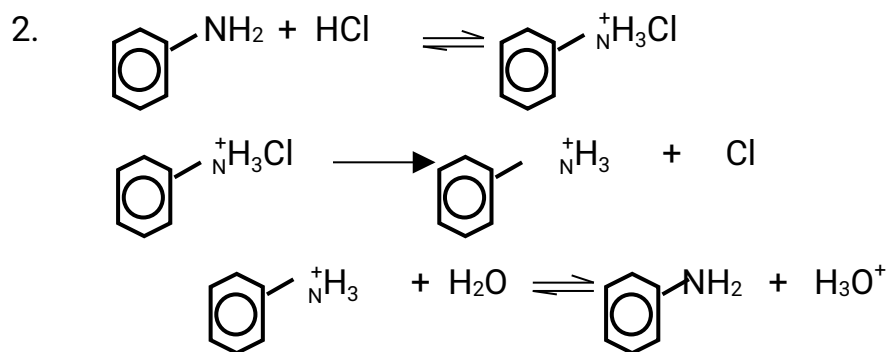
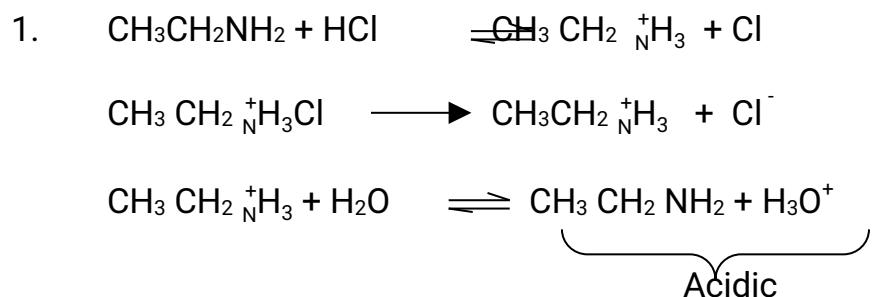
Secondary amine > Primary > Tertiary > ammonia

- Amino benzene is a weaker base than aliphatic primary amine. Explain.

The final of Basicity will be;

Secondary > Primary > Tertiary > Ammonia > Aromatic

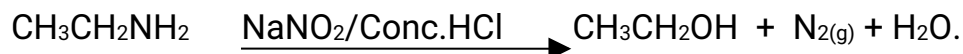
Being basic amines, they react with mineral acids to produce salts. These salts are strong electrolytes because they fully get ionized and their pH is slightly less than 7.



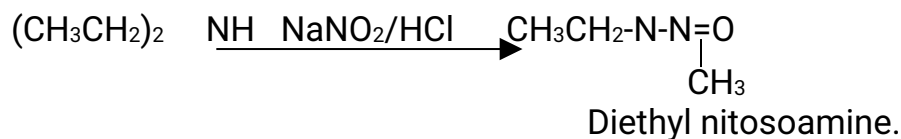
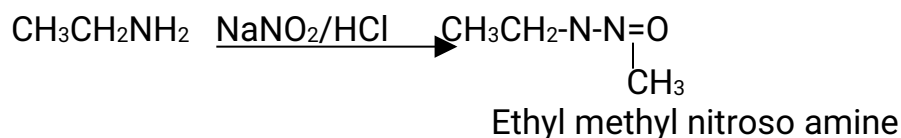
## REACTIONS OF AMINES

### 1. With nitrous acid:

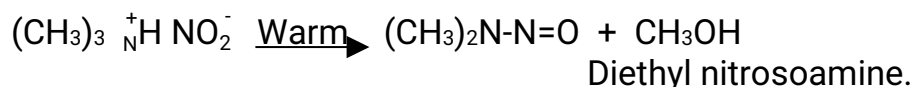
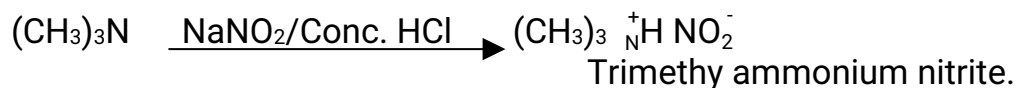
Nitrous acid being liable is generated in situ. Primary amines react with nitrous acid to form a colourless solution of alcohol and a colourless gas of nitrogen.



Secondary amines will produce a yellow oily substances known as nitroso amines.

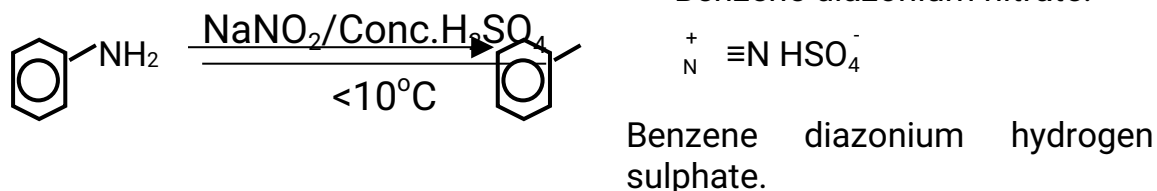
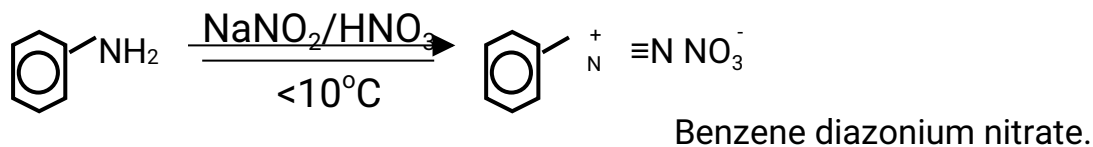
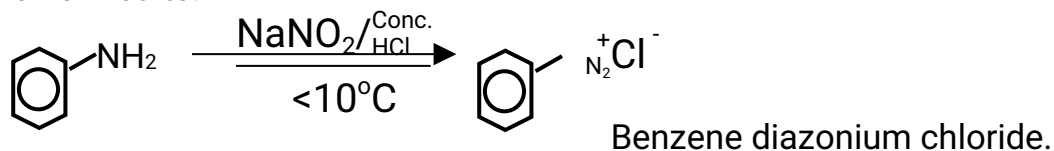


Tertiary amines when dissolved in HNO acid to form  $\text{NH}_4\text{NO}_2$  which when warmed forms a nitrosoamine and an alcohol by decomposition.

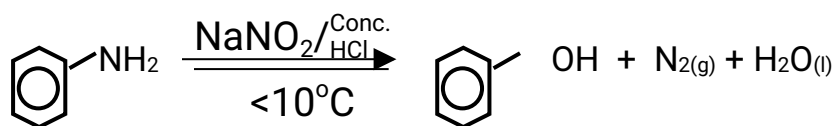


Aromatic primary amines react with nitrous acid giving different products depending on the temperatures.

At temperatures below  $10^\circ\text{C}$ , aromatic primary amines with nitrous acid to form diazonium salts.




At temperatures above ten ( $>10^{\circ}\text{C}$ ), the major product formed is a phenol.



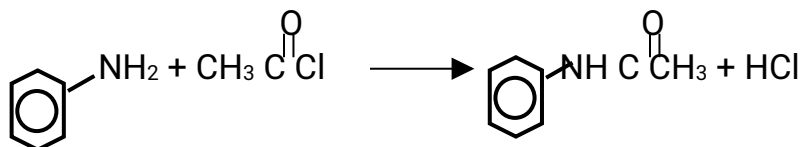
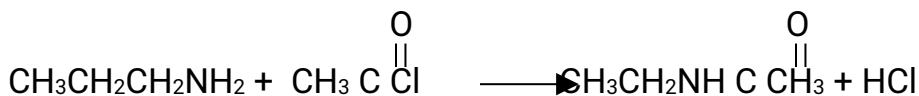
The above reaction is of practical importance in distinguishing the classes of amines.

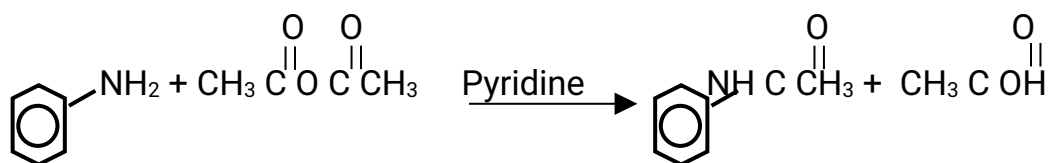
In summary:

- i)  $1^{\circ}$  amines : Yield a colourless solution and effervescence of a colourless gas.
- ii)  $2^{\circ}$  amines : Yield a yellow oily substance which when phenol and conc.  $\text{H}_2\text{SO}_4$  is added and the mixture made alkaline. This yields an *intense blue colour*.
- iii)  $3^{\circ}$  amines : A yellow oily substance persists.
- iv) Aromatic amine : A brightly coloured compound known as an azo dye is produced when 2-naphthol () is added.

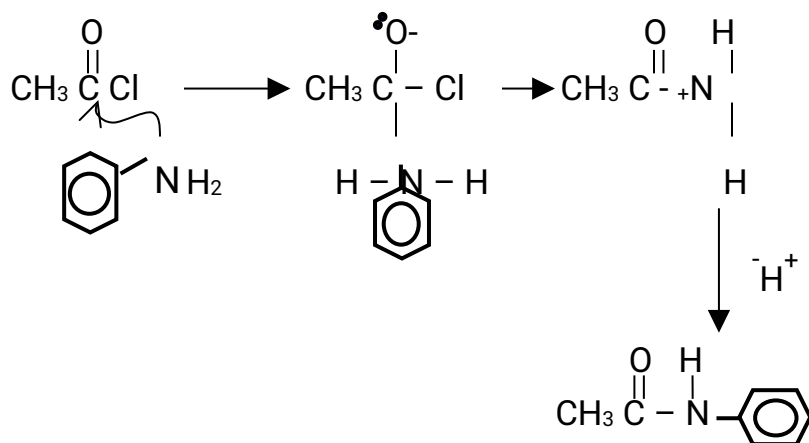
## 2. Reactions with acid chlorides and acid anhydrides.

$1^{\circ}$  and  $2^{\circ}$  amines react with acid chlorides and acid anhydrides to form amides.





**Mechanism:**



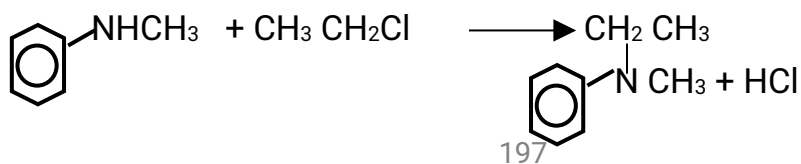
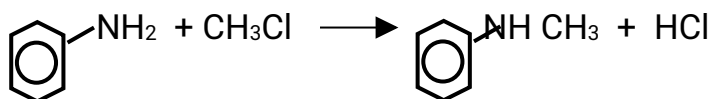
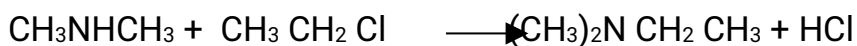
### 3. Reactions with sodium metal

Amines react with Na when heated to produce  $\text{H}_2$ . This reaction is only possible with  $1^\circ$  and  $2^\circ$  amines which do have  $\text{H}_2$ . The tertiary ones without  $\text{H}_2$  do not.



### 4. Reactions with alkyl halides

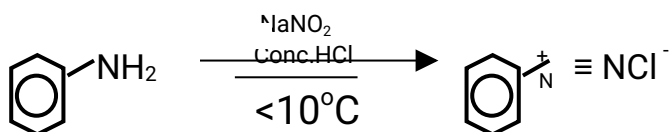
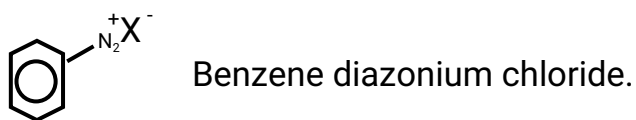
Aromatic and alkyl amines react with alkyl halides to form higher amines.



## DIAZONIUM SALTS

These are salts formed from primary aromatic amine when reacted with an inorganic mineral acid. The process of forming these salts is diazotization.

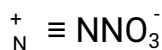
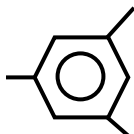
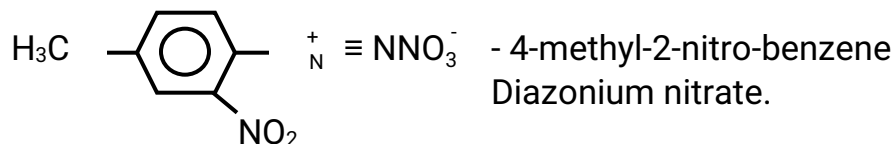
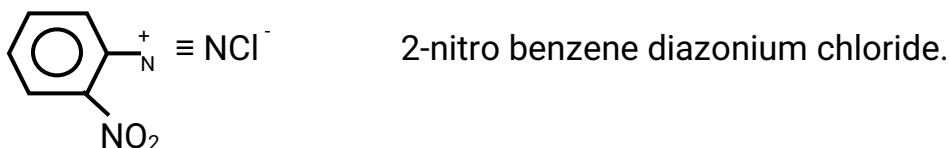
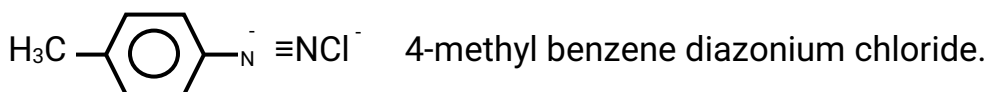
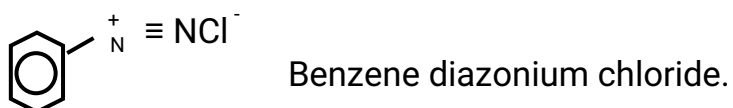
Temperature is critical in the formation of a diazonium salt and the temperature always has to be  $<10^{\circ}\text{C}$ . Where X is the halogen or  $\text{NO}_3^-$ ,  $\text{HSO}_4^-$ .



What determines the anion is the acid used.

### Nomenclature

Diazonium salts are named as benzene diazonium.



H<sub>3</sub>C

3, 5-dimethyl benzene diazonium  
Hydrogen sulphate.

CH<sub>3</sub>

### Physical properties of diazonium salts

Are a colourless crystalline solids which turn brown on exposure to air.

Their solutions are neutral to litmus but very soluble in H<sub>2</sub>O to give an ionic solution that has got good electrolytic conductivity.

### Chemical reactions

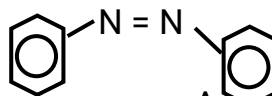
Diazonium salts are very useful in organic synthesis reacting with different substances to form different products.

### Coupling reaction

Benzene diazonium salts react with phenols to form brightly coloured compounds known as azo compounds by adding an aromatic ring to the nitronium ion a reaction known as coupling.

Ar – N = N – Ar

or

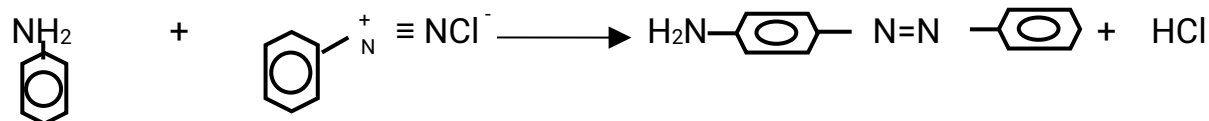
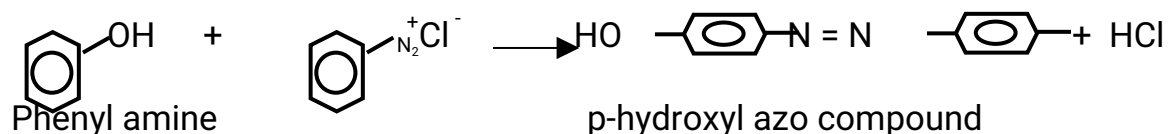


Azo compounds.

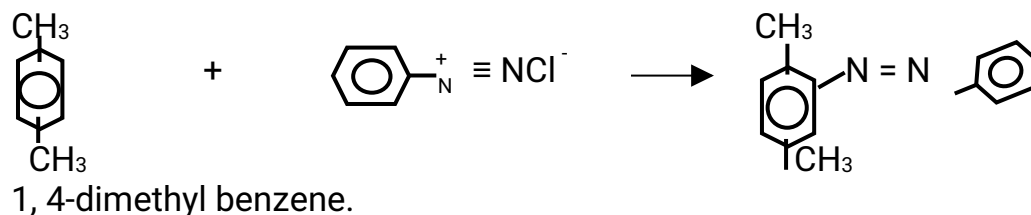
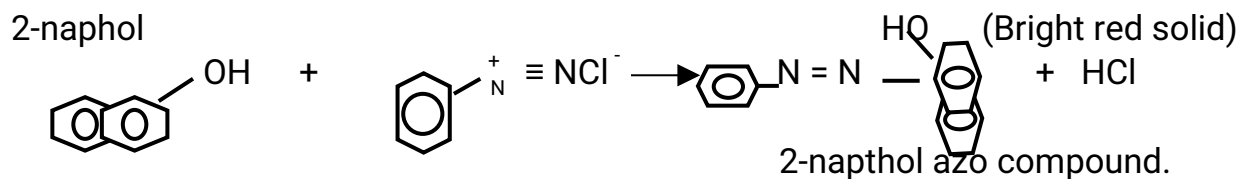
For coupling to occur, the aromatic ring must have a strong electron releasing groups such as hydroxyl, NH<sub>2</sub> – (amino), OR group and NHR group.

These groups will activate the ring making the electrons available at the para position for diazotization.

If the para position is already occupied, then diazoti-phenol.



p-amino azo compound.

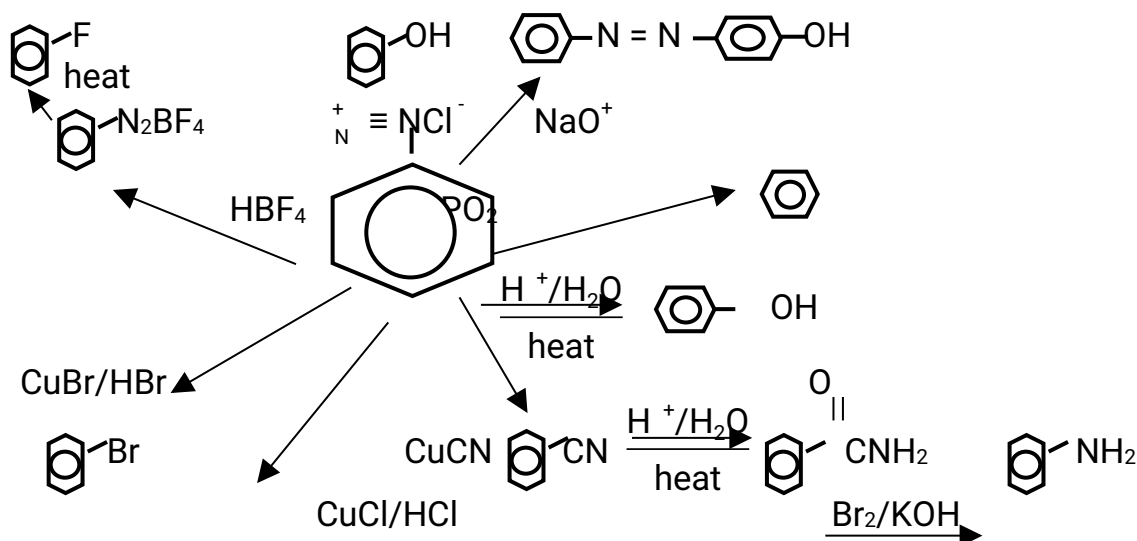
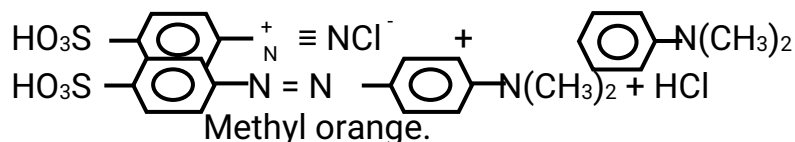


### The importance of coupling

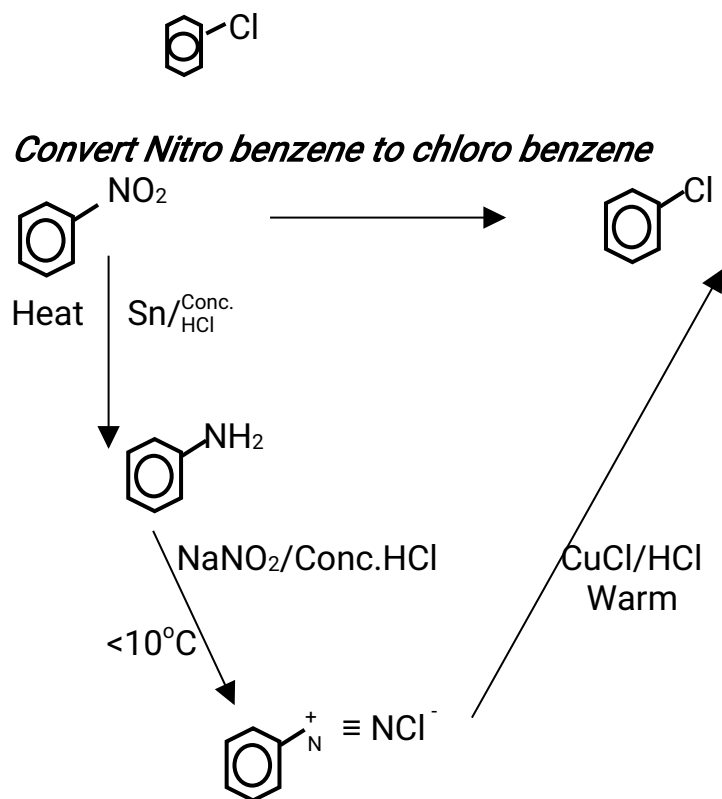
#### i) Test for primary aromatic amines:

If a primary aromatic amine is treated with a mixture of  $\text{NaNO}_2$  and  $\text{HCl}$ , 2-naphthol, a bright red solid is formed due to the reaction of diazonium salt with 2-naphthol at a temperature below  $10^\circ\text{C}$ .

ii) It is used in preparation of dyes (coloured compounds) e.g. methyl orange indicator is prepared by coupling 4-hydrogen sulphite benzene, diazonium chloride and dimethyl phenyl amine.







## APPLIED ORGANIC CHEMISTRY

1. Polymers (plastic)
2. Lipids – Soap and detergents
3. Hydro carbon – petroleum

### POLYMER AND POLYMERISATION

#### Definition:

A polymer is a compound of high molecular mass made up of small repeated units known as monomers. These molecules are formed by a process known as polymerization.

Polymerization is process of combining low molecular weight molecules repeatedly in some regular manner to form a high molecular mass compound. In polymerization, the low molecular mass molecules are always known as

monomers.

These are simple molecules which when combined can form a dimer, trimer or a polymer when many.

### **POLYMERS:**

There are two types of polymers depending on how they are formed.

i) **Natural polymers**

Are naturally occurring polymers not made by man e.g. cotton, wool, protein, rubber, starch e.t.c.

ii) **Artificial polymers (Synthetic)**

These are man-made polymers.

Example: Polythene

Polyester

Nylon e.t.c.

Poly vinyl chloride (PVC)

Terylene

### **Advantages of synthetic polymers over natural**

- They can be produced in large quantities within a short period of time.
- Are recyclable
- Have got improved properties e.g. cross linkages for toughness, high density for tensile strain.

### **Advantages of natural polymers over artificial polymers**

- They are biodegradable
- They are cheap since no process involved in producing them.

### **Properties of polymers**

A polymer is judged by some of these properties:

i) **The length of a polymer**

This determines its strength and the melting point. The longer the length, the stronger the polymer would be and a high melting point it will have.

ii) **Cross linkages**

A cross link is a force that joins molecules within a polymer. When a polymer has got several cross linkages, it will have an increased strength.

iii) **Branching**

Increased branches within a polymer decreases its strength and lowers its melting point.

iv) **Intermolecular forces**

If a polymer has got several intermolecular forces within itself, then it will have increased strength than when there are few intermolecular forces.

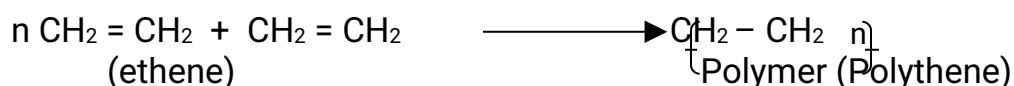
## CLASSIFICATION OF POLYMER

There are several ways of classifying polymers:

i) **Addition polymers**

These are polymers made by directly adding monomers without loss of any simple molecule.

Example:



ii) **Condensation polymers**

These are polymers made by combining monomer units but splitting out a small molecule which is usually H<sub>2</sub>O. E.g. Terylene

PET (Polythene terephthalate)

iii) **Thermoplastic polymers**

These are polymers that can soften and flow when heated and harden when they cool. In other words, they can be remolded into any shape by heat. E.g. polythene.

iv) **Thermosetting polymers**

these are polymers which when heated cannot melt but instead burn away. Such polymers cannot be remolded since they are initially set to a solid. E.g. Formica.

**Other terminologies include:**

i) **Copolymer**

A copolymer is a type of polymer made by having one or more different monomers. E.g. SBR (Styrene Butadiene Rubber). This is made from styrene and buta-1, 3-diene.

ii) **A fibre**

Is a form of polymer consisting of strings or bundles of strings that settle up during its formation.

iii) **Elastomers**

Is a type of polymer which when stress is applied, and released, it springs back to its original position e.g. rubber.

## ADDITION POLYMERIZATION

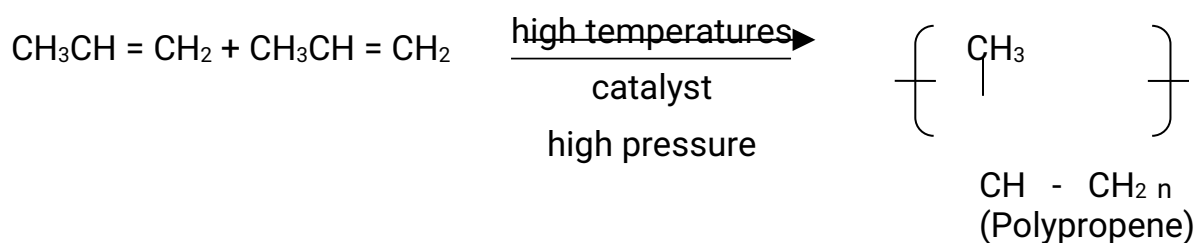
### Addition polymers

Definition: Addition polymerization


#### Examples:

Monomers	Formula	Polymer and trade name	Uses
Ethene (ethylene)	$\text{CH}_2 = \text{CH}_2$	Polythene (Polyethylene)	Making bags, squeeze bottles, films, toys, moulded objects, shoes, electrical insulators e.t.c.
Propene (propylene)	$\text{CH}_3\text{CH} = \text{CH}_2$	Poly propene (polypropylene) (Vectra)	Making bottles, films, indoor and outdoor carpets.

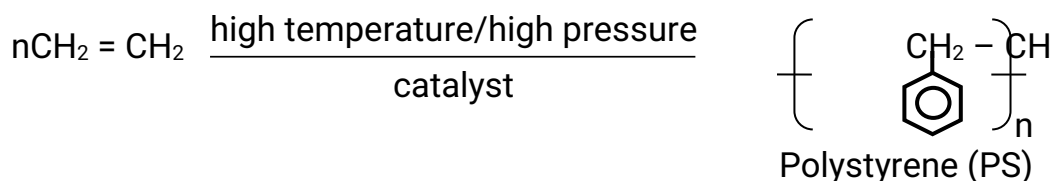
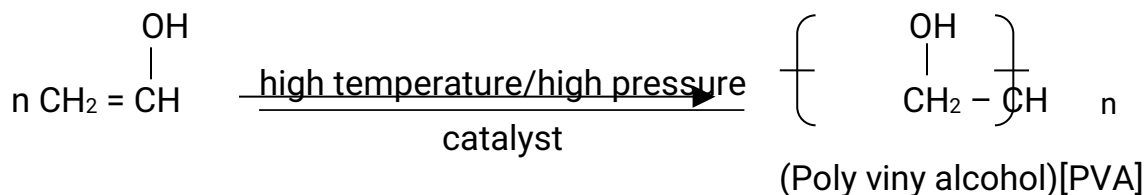
Polypropene is formed when propene is heated at various pressures and catalytic temperatures.



Vinyl chloride (Chloroethene)	$\text{CH}_2 = \text{CHCl}$	Poly vinyl chloride (PVC)	- Making of floor tiles.  - Making pipes, raincoats, insulator tapes.
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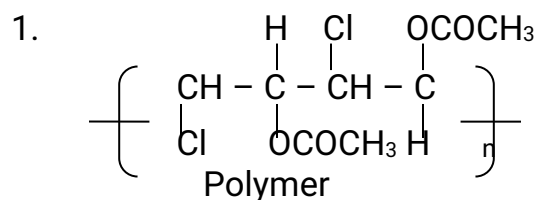
By modifying ethene, several polymers with better properties can be formed by adding groups of atoms with either  $-\text{OH}$  group (polar) or  (non polar, giving

polymers such as poly vinyl alcohol and polystyrene respectively. i.e.



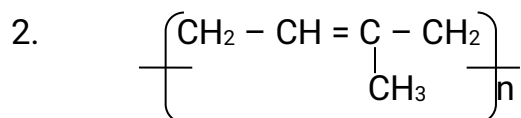
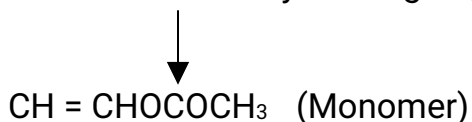
Styrene	$\text{CH}_2 = \underset{\text{C}_6\text{H}_5}{\text{CH}}$	Polystyrene (Styre form) or styrone	Making food and drink containers, coolers, construction materials, mattresses e.t.c.
Acryonitrile	$\text{CH}_2 = \text{CHCN}$	Polyacrylonitrile (Acrilan)	Fabrics, rugs e.t.c.
Vinyl acetate	$\text{CH}_2 = \text{CHOCOCH}_3$	Poly vinyl acetate (PVA)	Latex paint, coatings, textiles, adhesives.

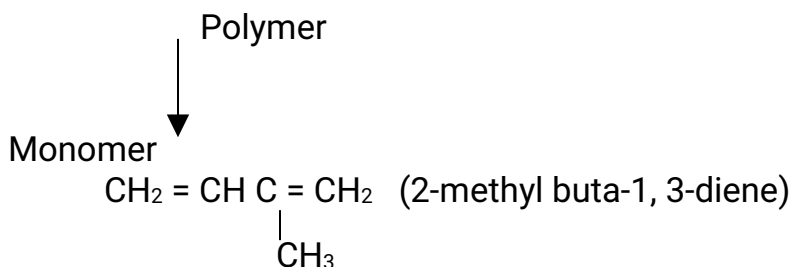
Emphasis: Structure of the polymer and its monomer.



Rules of writing monomers from polymers

Examine whether you are going to see the repeating units.





## HIGH DENSITY POLYTHENE AND LOW DENSITY POLYTHENE (HDPE & LDPE)

On the other hand, if ethene is polymerized giving branches, it will end up producing a low density polymer of 0.92g/cm<sup>3</sup> and this polymer will be known as LDPE which is soft and flexible.

## CONDENSATION POLYMERS

These are formed by splitting monomers to eliminate a small molecule which is usually H<sub>2</sub>O. This reaction is known as condensation reaction.

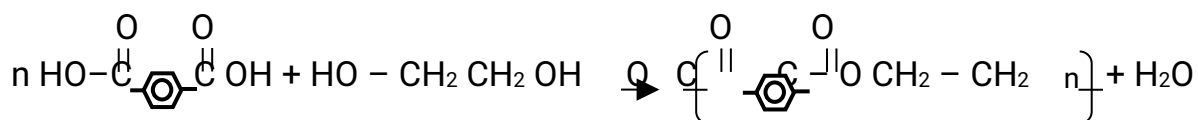
The empirical formula of the polymer in condensation polymerization is not the same as that of a monomer.

### Example:

#### i) Polyester (Dacron)

Are formed by reacting a carboxylic acid and an alcohol monomers to form a polyester and H<sub>2</sub>O molecule. E.g.

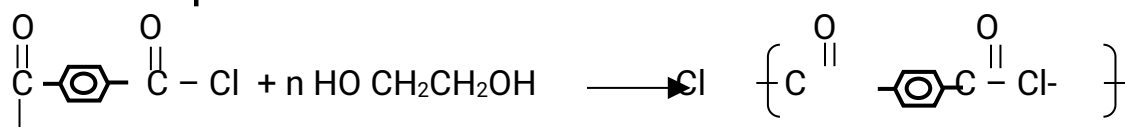
Benzene-1, 4-dioic (Terephthalic acid) and Ethane-1, 2-diol.



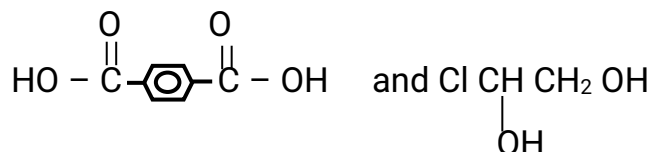
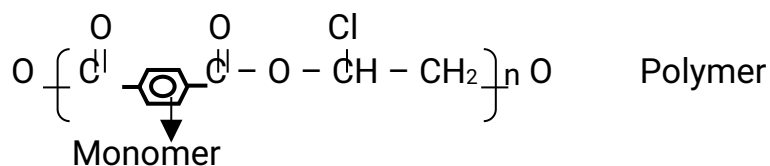
### Uses of polyesters

- Used in making of textile fabrics.
- Making films, magnetic coated films using audio and video tapes.
- In heart/cardiac operations to replace some heart blood vessels because it is non toxic, inert and non inflammatory.

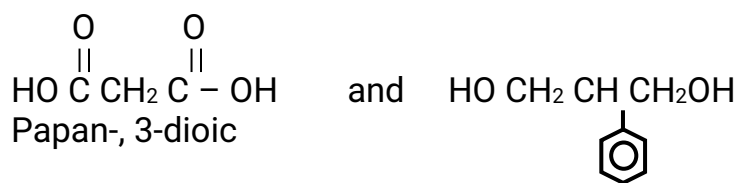
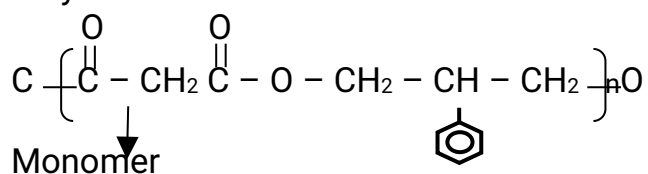
### Other examples:



Cl



Polymer

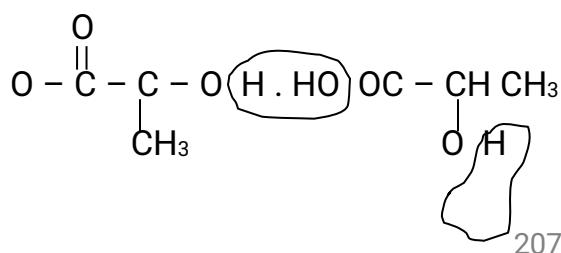
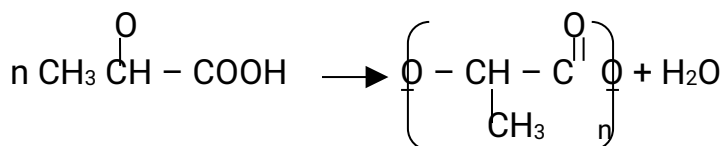


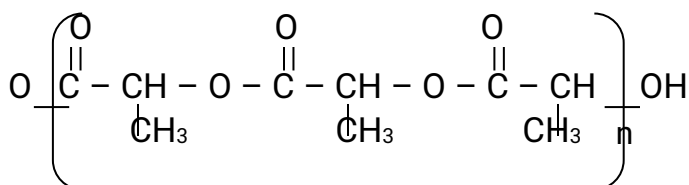
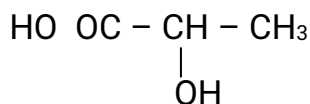
2-phenyl propane -1, 3-diol.

ii) **Lactic acid (2-hydroxy propanoic acid) polymer**

Is a polymer made up of one monomer of 2-dydroxy propanoic acid which undergoes self condensation. The uses of this polymer are of great importance in surgery and stitching broken deep wounds.

The advantage is that the polymer threads get dissolved in the body.





iii) **Polyamides**

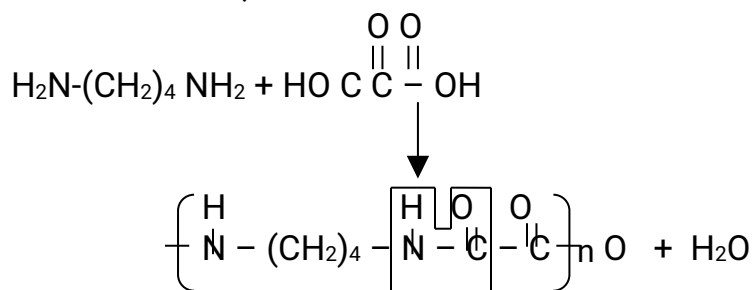
These are polymers formed by condensing a dicarboxylic acid or carboxylic acid chloride with a diamine to form an amide bond or link.

**Examples:**

Ethane-1, 4-dicarboxylic acid

Hexane-1, 6-diamine

Ethane-1, 2-dioic

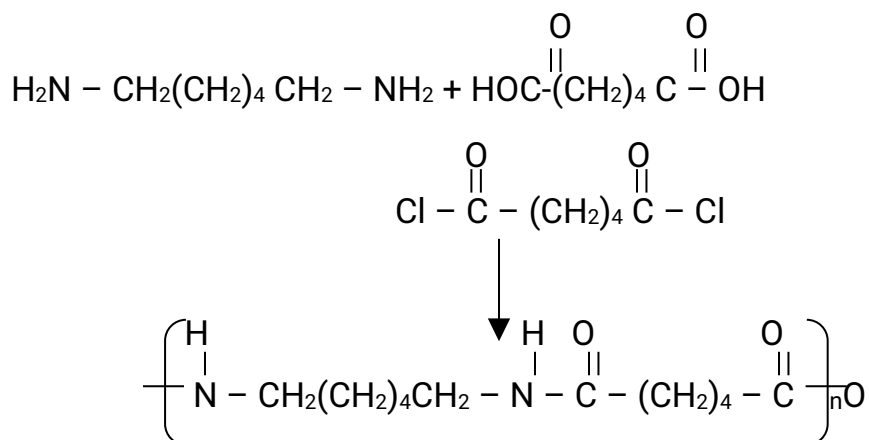


Peptide or amide bond.

**Other examples**

i) **Nylon-6 6:**

Is a polymer of Hexane-1, 6-dioic or its acid chloride and Hexane-1, 6-diamine. It is called so because its monomers are made up of 6 carbon atoms each.



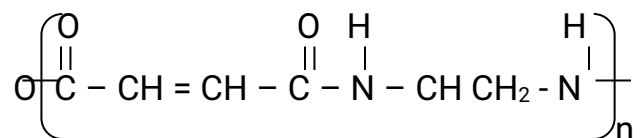


### Uses:

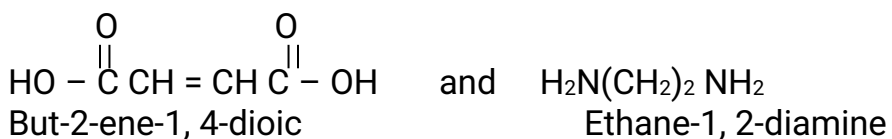
- Made of high quality fabric for the making of clothes.
- Used in parachutes.
- Used in military gear.

The advantage of nylon-6, 6 is that it has got a good tensile strength, H<sub>2</sub>O repellant, however, it has got a disadvantage of burning easily when put on fire.

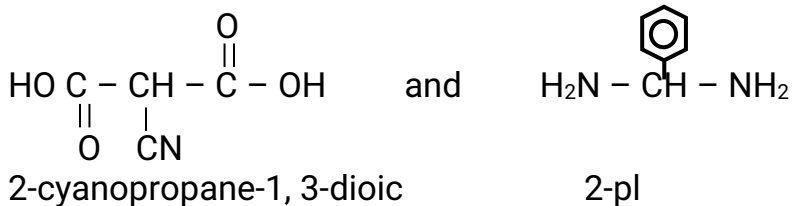
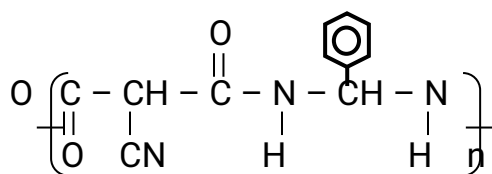
### Polymer



### Monomer



### Polymer



## NATURAL POLYMERS

These are polymers not made by man.

**Examples:**

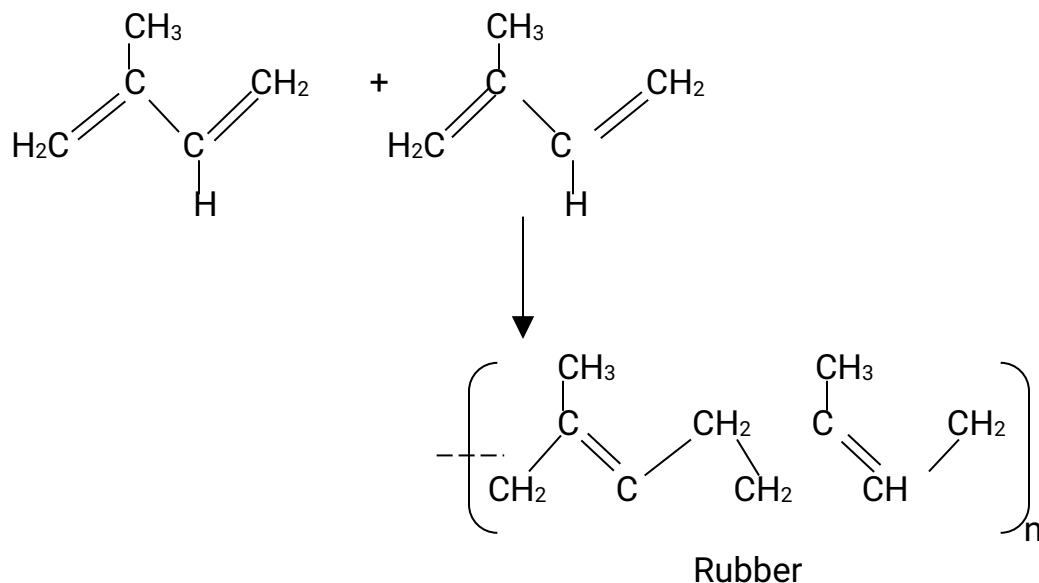
Cotton	}	(Polyamides)
Wool		
Proteins		
Carbohydrates		(Polysaccharide)
Lipids		(Polyesters)
Rubber		

## RUBBER

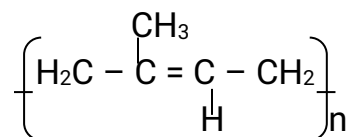
### NATURAL RUBBER

Is a polymer of 2-methyl buta-1, 3-diene which is simply known as isoprene unit.

Natural rubber is obtained from the rubber tree. Naturally, the rubber tree polymerizes the isoprene units by linking carbon 1 and carbon 4 and this leaves a double bond between carbon 2 and carbon 3 which double bond is in a cis – configuration.



OR:



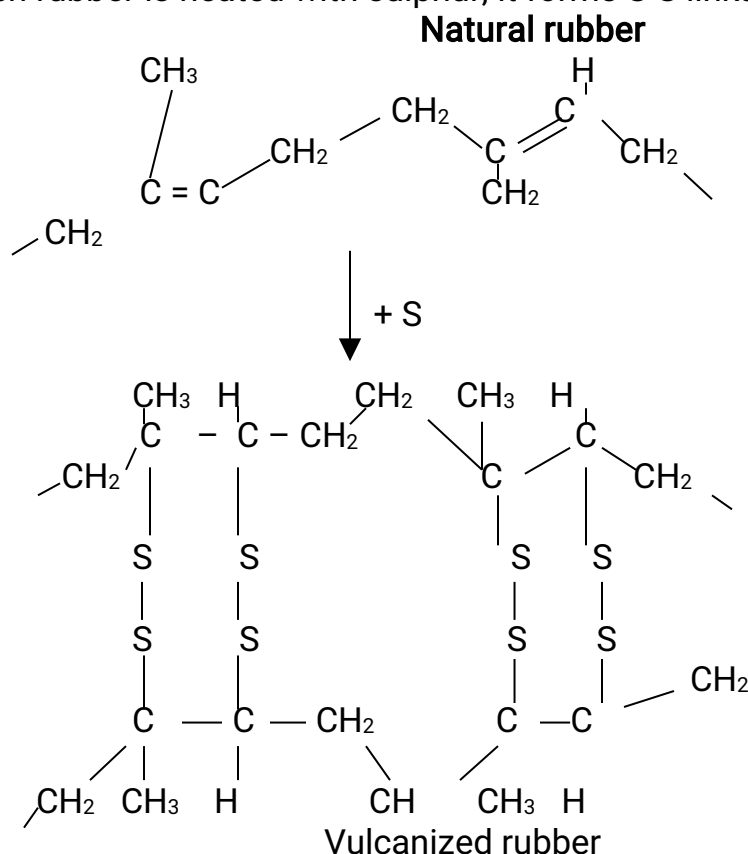
Natural rubber is weak due to a few cross linkages and therefore it is less elastic

but it can be improved by the process called Vulcanization.

### Definition:

Vulcanization is a process of making rubber, hard, tough and resistant to wearing so that it is made useful. This is done by cross linking monomers using disulphide bonds, S-S-bridge linkage.

When rubber is heated with sulphur, it forms S-S linkages which make it tough.



Vulcanized rubber is useful in the following ways:

- Manufacturing of car tyres
- Manufacture of balls, bags, shoe soles e.t.c.
- Used in insulation e.t.c.

### Synthetic rubber

The knowledge of composition and structure of rubber has given rise to synthetic rubber. This has made polymer by polymerizing elastomers to give rise to a copolymer which has got some properties of natural rubber.

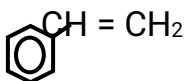
One of the examples is:

Styrene buta-diene rubber (SBR) which is made by polymerizing buta-1, 3-diene

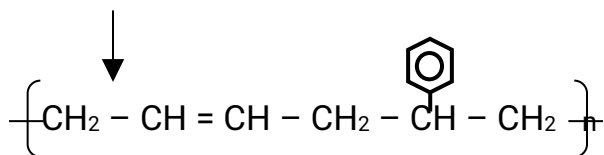
and styrene.



buta -1, 3-diene



Phenyl ethene



## FATS AND OILS

- Triglycerides / trimesters
- Differences between the oils and fats
- Components of a triglyceride (glycerol and fatty acid)
- Saponification
- Additives of soap
- Properties of soap.

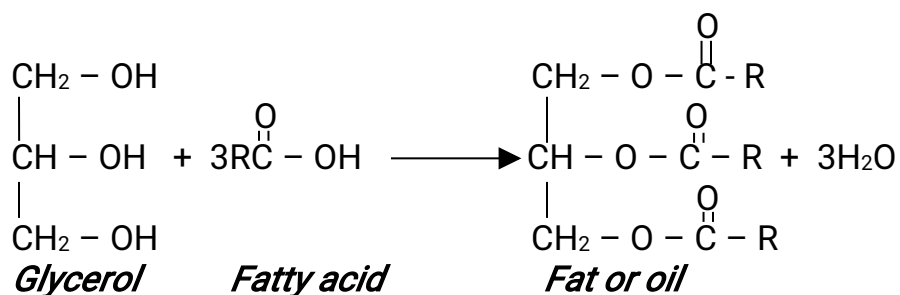
### FATS AND OILS

Fats and oil are trimesters made from glycerol (propane-1, 2, 3-triol) and long chain carboxylic acids called fatty acids. The R – groups of the acids, which can be the same or different within the same molecule, can be saturated or unsaturated, i.e. they may contain one or more carbon-carbon double bonds.

### COMMON FATTY ACIDS

Name	Number of carbon atoms	Formula
<b>Saturated acids</b>		
Butanoic	C <sub>4</sub>	CH <sub>3</sub> CH <sub>2</sub> CH <sub>2</sub> CO <sub>2</sub> H
Lauric	C <sub>12</sub>	CH <sub>3</sub> (CH <sub>2</sub> ) <sub>10</sub> CO <sub>2</sub> H
Myristic	C <sub>14</sub>	CH <sub>3</sub> (CH <sub>2</sub> ) <sub>12</sub> CO <sub>2</sub> H
Palmitic	C <sub>16</sub>	CH <sub>3</sub> (CH <sub>2</sub> ) <sub>14</sub> CO <sub>2</sub> H
Stearic	C <sub>18</sub>	CH <sub>3</sub> (CH <sub>2</sub> ) <sub>16</sub> CO <sub>2</sub> H
<b>Unsaturated acids</b>		
Oleic	C <sub>18</sub>	CH <sub>3</sub> (CH <sub>2</sub> ) <sub>7</sub> CH=CH(CH <sub>2</sub> ) <sub>7</sub> CO <sub>2</sub> H
Linolenic	C <sub>18</sub>	CH <sub>3</sub> CH <sub>2</sub> CH=CHCH <sub>2</sub> CH=CHCH <sub>2</sub> CH=CH(CH <sub>2</sub> ) <sub>7</sub> CO <sub>2</sub> H

A fatty acid – R group may be unsaturated, monounsaturated or polyunsaturated, depending on whether one or more double bonds is present.



### TRIGLYCERIDES (Triesters)

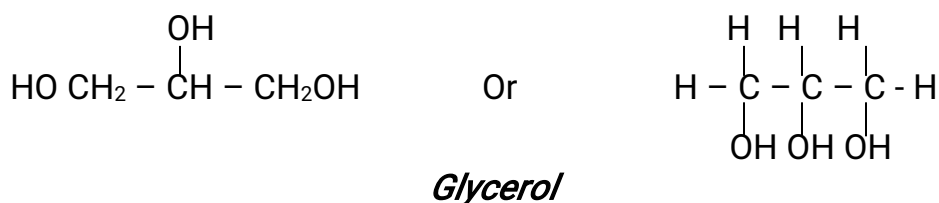
Common fats and oils made of a glycerol and fatty acids (act as components). Triglycerides consist of three fatty acids and one glycerol.

### FATTY ACIDS AND GLYCEROL

(Components of a triglyceride)

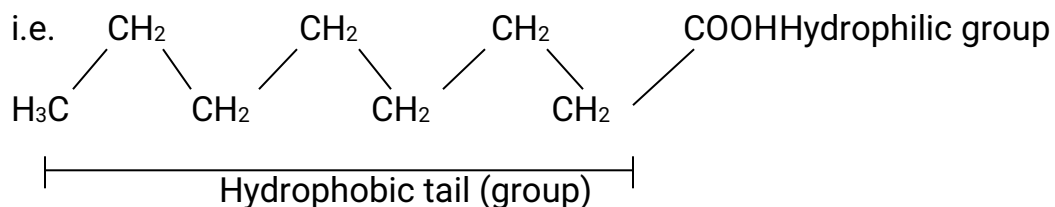
#### GLYCEROL

Is a polyhydric alcohol made up of three carbon atoms and 3 hydroxyl groups. It is therefore a small molecule.



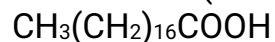
#### FATTY ACIDS

A fatty acid is a long carbon chain molecule consisting of a carboxylic acid group at one end. The long carbon chain makes it to be hydrophobic and the carboxylic acid group makes it to be hydrophilic.



If there are no double bond in the hydrocarbon chain, then the fatty acid is said to be ***saturated fatty acid***.

Example: Stearic acid (Octadecanoic acid)

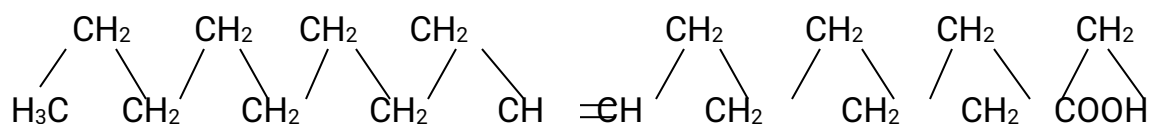


Saturated fatty acids from straight chains and their physical state is largely solid at room temperature because of their high melting points.

On the other hand, fatty acids that have got one or more double bonds within the hydrocarbon chain are said to be ***unsaturated fatty acids***.

**Example:**

Oleic acid (Octadec-cis-a-enoic acid)

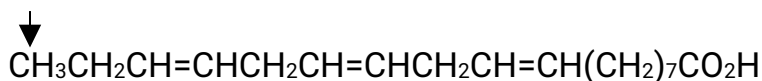


The presence of a double bond in the hydrocarbon chain results into the formation of a bend, which prevents the molecules from close packing resulting into the fatty acids having low melting points.

Those with trans double bond will tend to have fairly straight or linear carbon chain and the molecules will closely pack easily with high melting points making them solids at room temperature.

Fatty acids with more than one double bond are called **polyunsaturated fatty acids (PUFAS)**.

**Example:** Linolenic acid –  $\text{C}_{18}$ .



## SPONIFICATION

Saponification is the process of making soap.

## SOAPS AND DETERGENTS

### SOAP

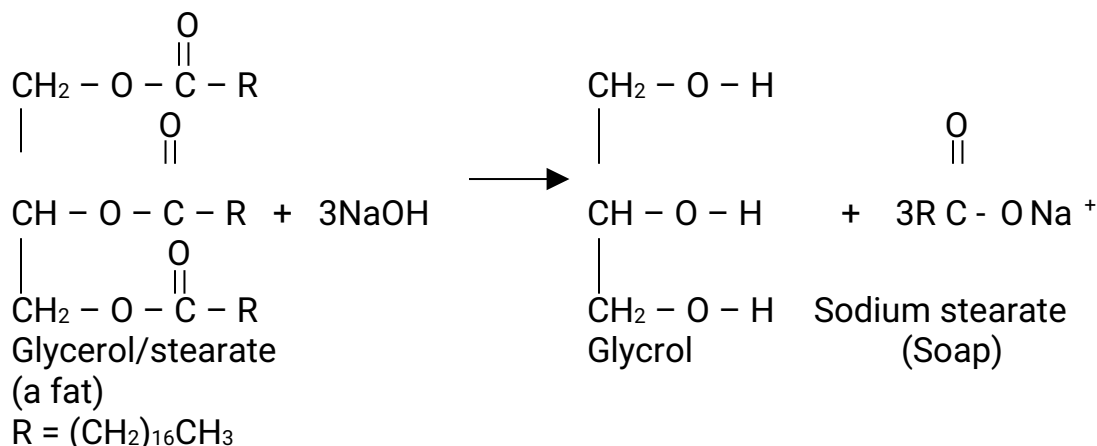
Soap is sodium or potassium salt of long chain carboxylic acids obtained by boiling fat or oil with NaOH or KOH solution.

### MANUFACTURE OF SOAP

Oil or fat is put on a pan, sodium hydroxide solution is added and the mixture boiled.

Sodium chloride (Common salt) is added to precipitate out the soap a process known as salting out. Soap precipitates out as a hard cake.

It is then removed and processed into various shapes. Perfumes and dyes are also added to add value to the soap.



### SOURCES OF OILS

#### i) **Vegetable oils:**

These can be obtained from cotton seeds, sunflower, simsim, castor, groundnuts e.t.c.

In extraction of vegetable oil after removing the husks, the seeds are then crushed in power form and then boiled with H<sub>2</sub>O for some time. The oil floats on top of the H<sub>2</sub>O.

#### ii) **Animal oil:**

These are esters obtained from animals e.g. from whale, fish (sardines) and lard oil.

## SOURCES OF FATS:

### i) **Vegetable fats:**

Are obtained from plants like coconut or palm trees.

### ii) **Animal fats:**

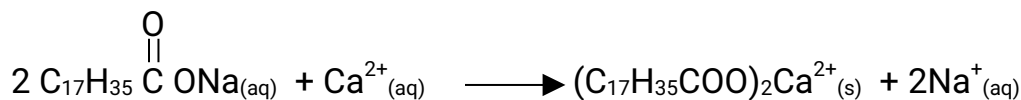
Are obtained from animals e.g. butter from milk and lard.

## Uses of fats and oils

- Are used for food.
- In the manufacture of soap, candles and hair oils.
- In the manufacture of paints and vanishes.
- Ground nut oil is used to make vegetable ghee.

## DISADVANTAGES OF USING ORDINARY SOAP

Soap forms scum with hard H<sub>2</sub>O. Hard H<sub>2</sub>O contains dissolved Ca<sup>2+</sup> or Mg<sup>2+</sup> which reacts with soap forming insoluble Ca<sup>2+</sup> salt called scum.



Formation of scum is expensive in that much soap has to be used to produce enough later for washing.

It also stains clothes.

## THE CLEANING ACTION OF SOAP

A molecule of soap contains a water loving part which is polar CO O called hydrophilic part and the non polar water insoluble (dirt soluble) part R – CH<sub>2</sub> called the hydrophobic part.

During washing, these molecules are dispersed and they form a spherical cluster around grease or dirt called micelle with the polar end attracted by H<sub>2</sub>O. The dirt combines with the hydrophobic part while H<sub>2</sub>O dissolves in the hydrophilic part.



As H<sub>2</sub>O molecules attract the polar end, the surface tension of H<sub>2</sub>O is lowered. Repulsion between the polar end of soap and rinsing help remove the dirt.

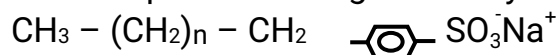
## DETERGENTS

Synthetic detergents solve the problem of scum formation.

Synthetic detergents are surface active agents and are called “**soapless soaps.**”

Like soap, detergents contain both hydrophilic (H<sub>2</sub>O soluble) and hydrophobic (oil solution) parts.

An example of a detergent is alkyl benzene sulphonates i.e.



However, the major disadvantage of alkyl benzene sulphonate detergents is that, they contain so many branches which are non-biodegradable and this can be dangerous if re-supplied into drinking H<sub>2</sub>O.

This can be solved by making detergents with no branching (i.e. linear alkyl group).

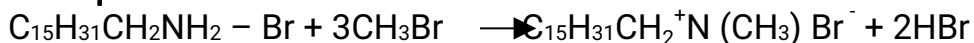
## TYPES OF DETERGENTS

### i) CATIONIC DETERGENTS

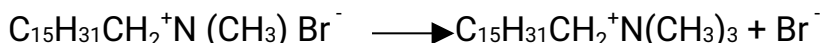
These carry positive heads e.g. hexadecyl trimethyl ammonium bromide. They are made by reacting amines with alkyl halides.

They are used in making air conditioners.

#### Example:



In H<sub>2</sub>O it behaves as;



### ii) ANIONIC DETERGENTS

These carry negative heads. They are made from benzene by alkylation, sulphonation and neutralization reaction e.g. sodium-4-dodecyl benzene sulphonate.

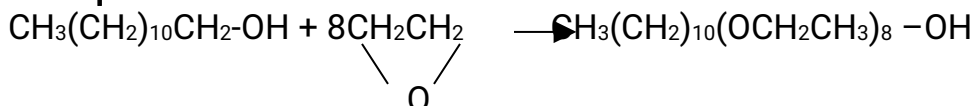
These are used in the manufacture of both pastes and shampoos.

Anionic detergents can also be made from straight chain alcohol ( $C_{10} - C_{14}$ )

### iii) **NON-IONIC DETERGENTS**

These are neutral in  $H_2O$  i.e. the head carries no charge i.e. they are not dissociated e.g.  $CH_3(CH_2)_{10}CH_2 - (OCH_2CH_2)_8 - OH$ , ethoxylate made from long chain alcohols with ethoxides.

#### **Example:**



They are used as liquid detergents.

The action of detergents can be improved by addition of the following which constitute the bulk of the detergents.

- Inorganic phosphates:** Are added to soaps to remove any soluble  $Ca^{2+}$  or  $Mg^{2+}$  in  $H_2O$  as  $Ca^{2+}$  or  $Mg^{2+}PO_4^{3-}$ .
- Sodium peroxoborate:** This is added to act as a bleaching agent. They make clothes appear bright. They release  $H_2O_2$  in  $H_2O$  which bleaches.
- Sodium sulphate:** This is added mainly to increase the bulk of the powder.

#### **Note:**

- The major disadvantage of detergents containing phosphates is that they cause pollution in rivers, dams, and lakes. (Phosphates are used as food by algae so their presence promote the growth of algae causing foaming of  $H_2O$  thus polluting it.
- Detergents are better than soaps because they can be used in any type of  $H_2O$  i.e. there is no formation of scum.

**END**