

Name:

Centre/Index No:

School.....

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P525/1

CHEMISTRY

Paper 1

July/August 2024

2 $\frac{3}{4}$ hours



WAKISSHA JOINT MOCK EXAMINATIONS

Uganda Advanced Certificate of Education

CHEMISTRY

Paper 1

2 hours 45 minutes

Instructions to Candidates

- Attempt all questions in section A and any six questions from section B.
- All questions are to be answered in the spaces provided.
- A Periodic Table with relevant atomic masses is supplied at the end of the paper.
- Mathematical tables (3 figures) and non-programmable silent scientific calculators may be used.
- Illustrate your answers with equations where applicable.
- Molar gas volume at s.t.p = 22.4 dm^3

For Examiner's Use Only

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	Total

SECTION A (46 MARKS)

Attempt all questions in this section.

1. Electrode potentials for some half-cells are given below.

Table 1

Half cell	$E^\theta(V)$
$Zn^{2+}(aq)/Zn(s)$	-0.76
$Cr_2O_7^{2-}(aq), H^+(aq), Cr^{3+}(aq)/Pt$	+1.33

- (a) Write the cell notation for the cell formed when the two half-cells are connected. (01 mark)
-
.....
.....

- (b) Write an equation for the overall cell reaction. (1½ marks)
-
.....
.....

- (c) (i) Calculate the free energy change of the cell. ($1F = 96500C$) (02 marks)
-
.....
.....
.....

- (ii) State whether the cell reaction is feasible or not. Give a reason for your answer. (01 mark)
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.....

2. 2-phenylpropane when oxidized in air at 5 atm formed liquid Q which reacts with dilute sulphuric acid to form compounds X and Y. Compound Y forms a crystalline white precipitate on addition of a saturated solution of sodium hydrogensulphite.

- (a) Identify;

(i) X: (01 mark)

.....

(ii) Y: (01 mark)

(iii) White precipitate..... (01 mark)

- (b) Name the reagent that can be used to confirm compound X. (01 mark)
-

- (c) State what would be observed when the reagent named in (b) above is added to X. (01 mark)
-

3. (a) Red lead oxide (Pb_3O_4) was shaken with dilute nitric acid and the resultant mixture filtered. Identify the;
(i) cation in the filtrate. (½ mark)
.....
(ii) residue (½ mark)
.....
- (b) Write an equation for the reaction that took place. (1½ mark)
.....
.....
- (b) Potassium iodide solution was added to the filtrate in (a) above;
(i) State what was observed. (01 mark)
.....
(ii) Write an equation for the reaction that took place. (1½ marks)
.....
.....
4. (a) Write the;
(i) equation for the hydrolysis of ammonium sulphate in water. (01 mark)
.....
.....
- (ii) expression for the hydrolysis constant, K_h for ammonium sulphate. (01 mark)
.....
.....
- (b) (i) The pH of 20 cm^3 of 0.05M ammonium sulphate solution at 25°C was found to be 5.125
Calculate the hydrolysis constant of ammonium sulphate. (03 marks)
.....
.....
.....
.....
.....
- (ii) State the assumptions you have made in b(i) above. (01 mark)
.....
.....

5. State what would be observed and write equation(s) for the reaction(s) that would take place when the following pairs of substances are mixed.
- (a) HO - CH₂COOH and phosphorous (V) chloride. (½ mark)

Observation:

Equation:

(01 mark)

- (b)  and an ice-cold mixture of sodium nitrite and concentrated hydrochloric acid.

Observation:

(½ mark)

Equation:

(01 mark)

- (c) Cobalt (II) sulphate solution and excess concentrated hydrochloric acid.
Observation:

(½ mark)

Equation:

(01 mark)

6. The melting points of some fluorides of period 3 elements of the Periodic Table are shown in table 2.

Table 2:

Formula of fluoride	NaF	AlF ₃	SiF ₄
Melting point (°C)	993	1290	-90.2

- (a) State the trend in the melting points of the fluorides.

(01 mark)

- (b) Explain your answer in (a).

(04 marks)

7. The empirical formula of compound R is CHO . 20 cm^3 of R were mixed with 110 cm^3 of oxygen and the mixture exploded. The residual gas was cooled to room temperature and on absorption by concentrated potassium hydroxide, there was a contraction of 80 cm^3 .

- (a) Calculate the molecular formula of R.

(2½ marks)

(b)

R decolourises bromine water and reacts with sodium carbonate solution with effervescence.

Write the structural formulae and IUPAC names of two geometric isomers of R.

(03 marks)

8.

(a)

Define the term **standard enthalpy of combustion**.

(01 mark)

(b)

The enthalpies of combustion of some substances are shown in table 3.

Table 3:

Substance	Enthalpy of combustion (KJmol^{-1})
Carbondisulphide	-1108.8
Carbon	-393.5
Sulphur	-296.8

Calculate the enthalpy of formation of carbon disulphide.

(03 marks)

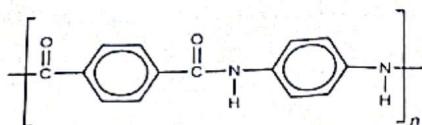
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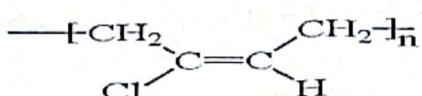
(c) Comment on the stability of carbondisulphide. Give a reason for your answer. (01 mark)

9. The molecular structures of two polymers; Kevlar and Neoprene are shown below.

Kevlar



Neoprene



(a) Name the type of polymer.

(½ mark)

(i) Kevlar

(½ mark)

(ii) Neoprene

(b) Write the structural formula(e) of the monomer(s) of;

(01 mark)

(i) Kevlar

(ii) Neoprene

(½ mark)

(c) State one use of each of the polymers above.

(02 marks)

SECTION B (54 MARKS)

Attempt any six questions from this section.

Any additional question(s) answered will **not** be marked.

10. (a) Zinc is extracted from zinc blende.

(i) State **one** method by which the ore can be concentrated.

(½ mark)

- (ii) Write equation(s) to show how zinc is obtained from the concentrated ore.
.....
.....
- (b) Zinc dust was added to an alkaline aqueous solution containing nitrate ions.
- (i) State what was observed. (01 mark)
.....
- (ii) Write an ionic equation for the reaction that took place. (1½ marks)
.....
.....
- (c) 1.50 g of an ore of zinc were dissolved in excess concentrated ammonia and the resultant solution shaken with trichloromethane. The concentrations of ammonia in the aqueous layer and trichloromethane layer at equilibrium were 0.08 moldm⁻³ and 0.0025 moldm⁻³ respectively at 25 °C.
Calculate the percentage by mass of zinc in the ore.
(Partition coefficient, K_p of ammonia between water and trichloromethane is 25) (04 marks)
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11. (a) Hydrogen peroxide oxidises iodide ions according to the equation.
 $\text{H}_2\text{O}_2(\text{aq}) + 2\text{H}^+(\text{aq}) + 3\text{I}^-(\text{aq}) \longrightarrow 2\text{H}_2\text{O}(\text{s}) + \text{I}_3^-(\text{aq})$
The rate of reaction is independent of the concentration of hydrogen ions and the reaction is first order with respect to hydrogen peroxide.
Given the kinetic data in the table 1 below obtained at 25°C.
- Table 1:**
- | [H ₂ O ₂] moldm ⁻³ | [I ⁻] moldm ⁻³ | [H ⁺] moldm ⁻³ | Initial rate moldm ⁻³ s ⁻¹) |
|--|---------------------------------------|---------------------------------------|--|
| 0.025 | 0.02 | 0.50 | 5.20 x 10 ⁻³ |
| 0.05 | 0.04 | 1.00 | 2.08 x 10 ⁻² |
- (i) State two methods by which the rate of reaction above can be determined. (02 marks)
.....
.....

(ii) Determine the rate equation for the reaction. (02 marks)

.....
.....

(iii) Calculate the value of the rate constant and state its units. (1½ marks)

.....
.....

(b) One of the experiments in Table 1 was repeated at 40 °C using the same concentrations of all reactants.

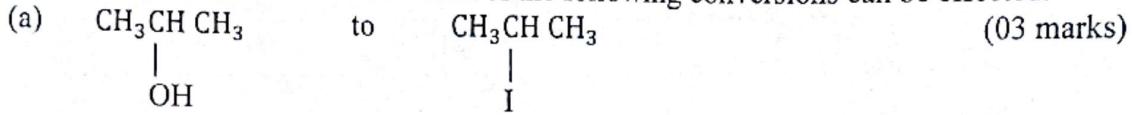
(i) State the effect on the value of the rate constant in a(iii) above. (01 mark)

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.....

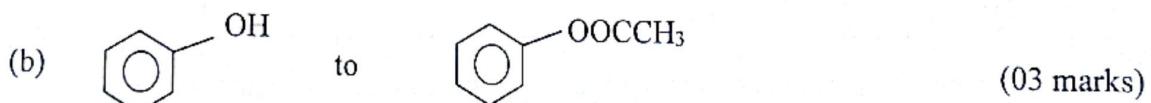
(ii) Explain your answer in b(i) (2½ marks)

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.....
.....

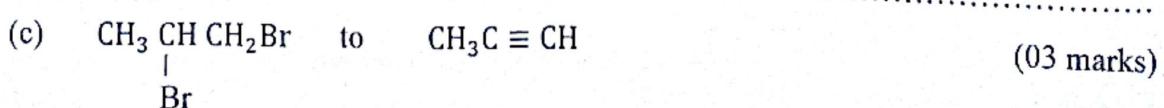
12. Write a mechanism to show how each of the following conversions can be effected.



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13.

Explain each of the following observations.

(a) When anhydrous aluminium chloride is exposed to moist air, misty fumes are formed.

(03 marks)

(b)

Aqueous solutions of Copper (I) salts are colourless while solutions of Copper (II) salts are coloured.

(03 marks)

(d)

When hydrogen sulphide gas is bubbled into acidified ammonium dichromate solution, the orange solution turns green and a yellow precipitate is formed.

(03 marks)

4.

(a)

Fluorine and Iodine belong to group VII of the Periodic Table. Explain why;

(i) Fluorine and Iodine have different physical states at room temperature.
(2½ marks)

(ii) Fluorine reacts directly with carbon while Iodine does not. (02 marks)

Turn Over

(b) Write an equation for the reaction between; (1½ marks)

(i) Fluorine and water.

(ii) Iodine and hot concentrated sodium hydroxide solution. (1 ½ marks)

(iii) hydride of Iodine and excess concentrated sulphuric acid. (1½ marks)

15. Lead (II) Iodide is sparingly soluble in water.

(a) Write the;

(i) equation for the solubility of Lead (II) iodide in water. (01 mark)

(ii) expression for the solubility product, K_{sp} of Lead (II) iodide. (01 mark)

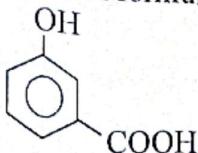
(b) The solubility product of Lead (II) iodide is $1.39 \times 10^{-8} \text{ mol}^3 \text{dm}^{-3}$ at 25 °C. Calculate the solubility of Lead (II) iodide in gdm⁻³. (03 marks)

(c) Explain how the solubility of lead (II) iodide would be affected if to its saturate solution;

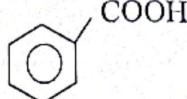
(i) a few drops of potassium iodide are added. (02 marks)

(ii) magnesium ribbon is added. (02 marks)

16. (a) The structural formulae of two aromatic organic compounds are



and



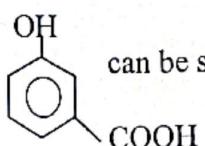
- (i) Name the reagent(s) that can be used to distinguish between the compounds.

(01 mark)

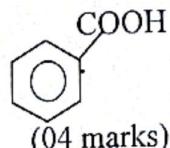
- (ii) State what would be observed when each compound is separately treated with the reagent in a(i) above.

(02 marks)

- (b) Using equations only, show how

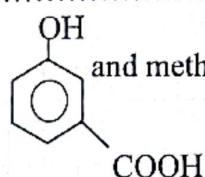


can be synthesized from



(04 marks)

- (c) A mixture of



and methanol was heated in the presence of sulphuric acid.

- (i) State what was observed.

(01 mark)

- (ii) Write an equation for the reaction that took place.

(01 mark)

17. (a) The vapour pressures of acetone and benzene are 30 Kpa and 12.68 kPa at 25 °C respectively. A solution containing 20% acetone and 80% benzene at 25 °C exerted a vapour pressure of 18.00 Kpa.

- (i) Calculate the vapour pressure above the solution assuming it is ideal. (03 marks)

Turn Over

(ii) State the type of deviation from ideal behaviour shown by the solution of acetone and benzene. Give a reason for your answer. (02 marks)

(b) Sketch a well labelled boiling point composition diagram for the mixture of acetone and benzene. (03 marks)

(d) State **one** method by which an azeotropic mixture of acetone and benzene can be separated. (01 mark)

THE PERIODIC TABLE

1	2																			3	4	5	6	7	8																
1 Li 6.9	2 Be 9.0																			1 B 10.8	2 C 12.0	3 N 14.0	4 O 16.0	5 F 19.0	6 Ne 20.2																
3 Na 23.0	4 Mg 24.3																			7 Al 27.0	8 Si 28.1	9 P 31.0	10 S 32.1	11 Cl 35.4	12 Ar 40.0																
19 K 39.1	20 Ca 40.1	21 Sc 45.0	22 Ti 47.9	23 V 50.9	24 Cr 52.0	25 Mn 54.9	26 Fe 55.8	27 Co 58.9	28 Ni 58.7	29 Cu 63.5	30 Zn 65.7	31 Ga 69.7	32 Ge 72.6	33 As 74.9	34 Se 79.0	35 Br 79.9	36 Kr 83.8																								
37 Rb 85.5	38 Sr 87.6	39 Y 88.9	40 Zr 91.2	41 Nb 92.9	42 Mo 95.9	43 Ru 98.9	44 Rh 101	45 Pd 103	46 Ag 106	47 Cd 108	48 In 112	49 Sn 115	50 Sb 119	51 Te 122	52 I 126	53 Xe 127																									
55 Cs 133	56 Ba 137	57 La 139	58 Hf 170	59 Ta 181	74 W 184	75 Os 186	76 Ru 190	77 Ir 191	78 Pt 195	79 Au 197	80 Hg 201	81 Tl 204	82 Pb 207	83 Bi 209	84 Po (209)	85 At (210)	86 Rn (222)																								
87 Fr (223)	88 Ra (226)	89 Ac (227)																	57 La 139	58 Ce 140	59 Pr 141	60 Nd 144	61 Pm (145)	62 Sm 151	63 Sm 150	64 Eu 152	65 Tb 159	66 Dy 162	67 Ho 165	68 Er 167	69 Tm 169	70 Yb 173	71 Lu 175								
																			89 Ac (227)	90 Th 232	91 Pa 233	92 U 238	93 Np 237	94 Pu (240)	95 Am (243)	96 Cm (247)	97 Bk (247)	98 Cf 251	99 Es (254)	100 Fm (257)	101 Md (256)	102 No (254)	103 Lw								

1. Indicates atomic number.

2. Indicates relative atomic mass.

END