Aim:

An Experiment to determine the Enthalpy Change when Z ----

Hypothesis: heat gain / heat loss

Endothermic / Exothermic reaction

Variables: dependent · temperature
independent: Time (stop watch)
Controlled: Volume of water (50, 60, 70 cm³)

Materials: Beaker
thermometer
metal Q & Z
stop clock
distilled water

Procedure:

Q/R was dissolved in 50, 60 OR 70 cm³ of distilled was water tomake a solution BA7 & temperature measured (27°C)
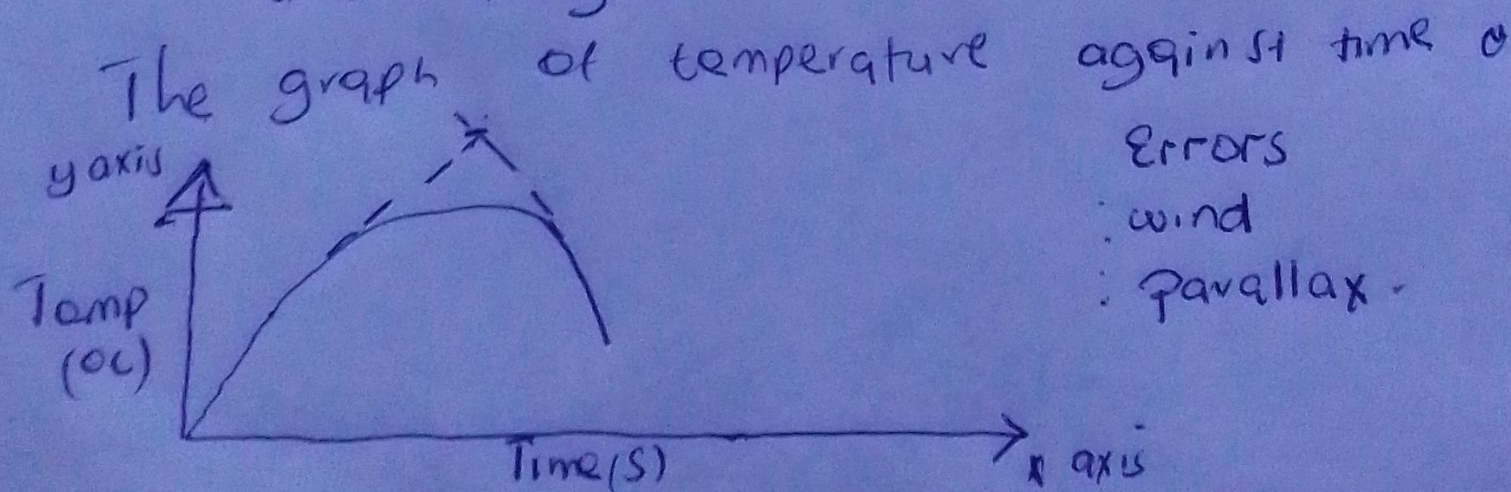
Metal Q/R was added and the mixture stirre using athermometer · temperature was recorded at an interval of 30 seconds the results was recorded in the table below

Table of Results

Time (°C)	0	30	60	90	120	150	180
Temp (°C)	27						

data analysis

Tittle: The graph of temperature against time

Errors
· wind
· Parallax.

Conclusion: heat was given out when Q & Z was reacted with solution BA7
OR
heat was lost when Q & R was reacted with solution BA