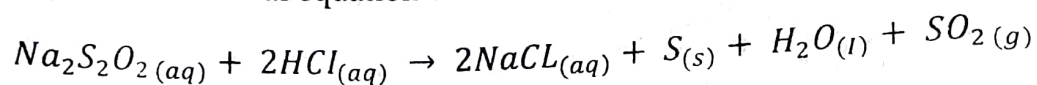


(b)

Item

An insecticide manufacturing factory heavily relies on sulphur as a raw material for production of their insecticides. The sulphur used is obtained from the reaction between sodium thiosulphate solution and hydrochloric acid which are available in the factory stores.

To increase the rate of sulphur formation and meet the high demand for insecticides, the factory chemist decided to increase the temperature of sodium thiosulphate solution. Sodium thiosulphate solution reacts with hydrochloric acid according to the chemical equation below.



You are provided with

BA 1 which is a solution of sodium thiosulphate.

BA 2 which is a solution of hydrochloric acid

Task

(c)

- (a) Design an experiment you would carry out to investigate the suitability of the chemist's intervention.

Handwritten text in Arabic script, consisting of approximately 25 lines of cursive script. The text is mostly illegible due to fading and the cursive style.

(b) Carry out the experiment and record your findings.

(c) What can the factory chemist deduce from your findings?

END