In-class Lecture 8 - Diffusion by Master Equation Wednesday, February 9, 2022 We want to solve the concentration profile dynamics due to diffusion simple model of diffusion: Ni-A Ni Ni+A

M boxes

Let's bring up the concentration:
$$\frac{N(i,t)}{V_{boo}} = c(i,t)$$

Let's bing in the \times coordinate: $i \rightarrow \times$
 $i+1 \rightarrow \times + \circ$
 $c(x,t+\Delta t) = c(x,t) + c(x-\alpha,t) + \Delta t + c(x+\alpha,t) + \Delta t$
 $c(x,t) + \Delta t - c(x,t) + \Delta t$

Move
$$c(x,t)$$
 and divide by Δt :
$$\frac{c(x,t)}{\Delta t} = (c(x-a,t)) + c(x+a,t) +$$

$$= ka \frac{\partial c}{\partial x} \left(x + \frac{a}{2}, t\right)$$

$$= -ka \frac{\partial c}{\partial x} \left(x + \frac{a}{2}, t\right)$$

$$= -ka \left[\frac{\partial c}{\partial x} \left(x + \frac{a}{2}, t\right) - \frac{\partial c}{\partial x} \left(x - \frac{a}{2}, t\right)\right] \frac{a}{a}$$

$$= ka^{2} \frac{\partial c}{\partial x} \left(x + \frac{a}{2}, t\right) - \frac{\partial c}{\partial x} \left(x - \frac{a}{2}, t\right)$$

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 $\frac{\partial c}{\partial t} = ko^{2} \frac{\partial^{2} c}{\partial x}$ $\frac{\partial c}{\partial t} = 0$ $\frac{\partial c}{\partial x} = 0$

#Diffusion constant in um^2/s

#Size of our boxes in um

TotalTime=0.01; #Total simulation time in s.

Initial conditions

#Now, we need the simulation parameters. We need our time and space

#intervals to be smaller than any other quantity in the process.

dt=1/k/10; #Temporal interval in s, has be smaller than 1/k

#Size of our cell in um

Ntotal=100; #Number of molecules

dx = 0.01;

 $k=D/dx^2; #D=k*dx^2$

For boxes 1 through 5:

For box M-1 N(N-1,4+14)=N(N-1,4)+N(N-2,4) KA+-N(N-4) KA+ for simulation