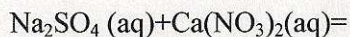


s-solid, g-gas,

Part 1. Aqueous chemistry

1



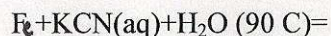
- A. NaNO_3 precipitates; the rest remain in solution
 \rightarrow B. CaSO_4 precipitates; the rest remain in solution
 C. both precipitate

2



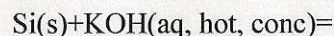
- A. $\text{HgCl}_2(\text{solution}) + \text{H}_2(\text{g})$
 \rightarrow B. no reaction
 C. $\text{Hg}_2\text{Cl}_2(\text{solid}) + \text{H}_2(\text{g})$

3



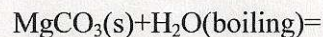
- A. No reaction
 B. $\text{K}_3\text{Fe}(\text{CN})_6 + \text{KOH} + \text{H}_2(\text{gas})$
 \rightarrow C. $\text{K}_4\text{Fe}(\text{CN})_6 + \text{KOH} + \text{H}_2(\text{gas})$

4



- A. no reaction
 \rightarrow B. $\text{K}_2\text{SiO}_3(\text{aq}) + \text{H}_2(\text{g})$
 C. $\text{SiO}_2 + \text{K}_2\text{O} + \text{H}_2(\text{g})$

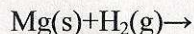
5



- ~~A.~~ $\text{Mg}(\text{HCO}_3)_2$
 \rightarrow B. $\text{Mg}(\text{OH})_2 + \text{CO}_2(\text{g})$
 C. No reaction

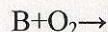
Part 2. Non-aqueous chemistry

6



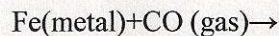
- \rightarrow A. MgH_2
 B. No reaction
 C. Mg_2H

7



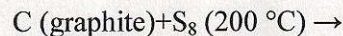
- \rightarrow A. B_2O_3
 B. No reaction
 C. BO_2

8



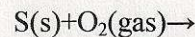
- A. $\text{FeCO}(\text{s})$
 \rightarrow B. $\text{Fe}(\text{CO})_5(\text{l})$
 C. Non of the above

9



- A. CS_2
 \rightarrow B. CS_2 , CS and other products
 C. CS

10



- \rightarrow A. SO_2
 B. SO_3
 C. SO

10
10