Form #16 Please record the form number for your records. DO NOT WRITE YOUR NAME s-solid, g-gas, Part 2. Non-aqueous chemistry Part 1. Aqueous chemistry Na_2SO_4 (aq)+ $Ca(NO_3)_2$ (aq)= $Mg(s)+H_2(g) \rightarrow$ A. NaNO₃ precipitates; the rest remain in solution →A. MgH₂ →B. CaSO₄ precipitates; the rest remain in solution B. No reaction C. both precipitate C. Mg₂H 7 2 $B+O_2 \rightarrow$ Hg(1)+HCl(aq)= \rightarrow A. B₂O₃ A. HgCl₂ (solution)+H₂(g) B. No reaction →B. no reaction C. Hg₂Cl₂ (solid) +H₂(g) C. BO₂ 3 8 $F_e + KCN(aq) + H_2O(90 C) =$ Fe(metal)+CO (gas)→ A. FeCO(s) A. No reaction -B. Fe(CO)5(1) B. K₃Fe(CN)₆+KOH+H₂(gas) C. Non of the above -C. K₄Fe(CN)₆+ KOH H₂(gas) 9 C (graphite)+S₈ (200 °C) \rightarrow Si(s)+KOH(aq, hot, conc)= A. no reaction A. CS₂ \rightarrow B. K₂SiO₃(aq)+H₂(g) - B. CS₂, CS and other products C. $SiO_2+K_2O+H_2(g)$. C. CS 10 $S(s)+O_2(gas) \rightarrow$ $MgCO_3(s)+H_2O(boiling)=$ AMA. Mg(HCO₃)₂ -A. SO₂ \rightarrow B. Mg(OH)₂+CO₂(g) B. SO₃ C. No reaction C. SO