Answer all the questions (25 X 1 = 25)

Reference electrode and working electrode used in the estimation of Fe(II) 1 point
ions by potentiometry are _____. *

○ Glass electrode and Platinum electrode

○ Platinum electrode and Standard Calomel Electrode

○ Standard Calomel Electrode and Glass electrode

◉ Standard Calomel Electrode and Platinum electrode

Viscosity is due to _____. *   1 point

◉ Resistance to fluid motion

○ Roughness of the surface

○ The pressure difference between the two fluids

○ Potential energy stored in fluid

In Mohr's method the solution needs to be near neutral, because *   1 point

○ Silver precipitates at low pH

◉ Chromate forms H2CrO4 at low pH, which delays the formation of the precipitate.

○ Silver chloride forms at high pH,

○ Potassium chromate dissolves at high pH.

The pH of a liquid solution is a measure of _____. *   1 point

◉ Hydrogen ion activity

○ Dissolved salt content

○ Hydroxyl ion molarity

○ Electrical conductivity

The electrolyte solution within the glass electrode (reference) of the pH   1 point
meter is _____. *

○ Dilute NaCl

○ Dilute HCl

○ Concentrated HCl

◉ Saturated KCl

What is the working principle of conductometry? *   1 point

○ Measurements of pH

○ Measurement of emf.

○ Measurement of potential.

◉ Measurement of conductivity of solution.

When sodium hydroxide is added to HCl, the H+ ions are replaced by   1 point
_____. *

◉ Slow moving Na+ ions

○ Slow moving OH- ions

○ Fast moving OH- ions

○ Fast moving Na+ ions

When a strong base is added to a strong acid after the neutralization point   1 point
*

○ Conductance remains constant

○ Conductance decreases

○ Conductance decreases initially and then increases gradually

◉ Conductance increases

Temporary hardness of water is caused due to the presence of_____ dissolved solids *   1 point

○ Calcium bicarbonates only

○ Magnesium bicarbonates only

○ Sulphates and Chlorides of Calcium or Magnesium

◉ Calcium hydrogen carbonates and Magnesium hydrogen carbonates

---

Conductivity of a solution is directly proportional to *   1 point

○ Current density

○ Volume of the solution

◉ Number of ions

○ Dilution

Estimation of Fe(II) ions by potentiometry is_____titration. *   1 point

○ Complexometric

○ Acid-base

○ Precipitation

◉ Redox

Hardness of water is usually expressed in terms of _____ equivalent   1 point
amount of hardness. *

○ K2CO3

◉ CaCO3

○ MgCO3

○ Na2CO3

Which type of reaction occurs in the following reaction AgNO3 + NaCl →   1 point
AgCl + NaNO3? *

○ placement reaction

◉ Double displacement reaction

○ Single replacement

○ Decomposition

A plot of hsp / C (reduced viscosity) vs C is a _____ for dilute polymer   1 point
solutions *

○ "V" shape curve

○ "S" shape curve

◉ Straight line

○ Triangle

If the ion size decreases in solutions then *   1 point

○ Does not affect the conductance

◉ Conductance increases

○ Conductance decreases

○ First decreases and then increases

The equivalent weight of Sodium Carbonate [Na2CO3] is *   1 point

○ 40

○ 55.85

◉ 53

○ 63

Phenolphthalein color in basic medium is *   1 point

○ Orange

○ Steel Blue

○ Yellow

◉ Pink

Estimation of chloride reaction is a_____. *   1 point

○ Catalytic reaction

○ Equlibrium reaction

○ Redox reaction

◉ Precipitation reaction

Indicator used for the estimation of hardness in EDTA method is
_____. *

1 point

🔘 Eriochrome Black – T

⭕ Phenolphthalein

⭕ Potassium dichromate

⭕ Methyl orange

A buffer solution comprises of _____. *

1 point

⭕ a weak acid in solution

⭕ a weak base in solution

🔘 a weak acid and its conjugate base in solution

⭕ a strong acid in solution

In the Estimation of Fe(II) ions by potentiometry, K2Cr2O7 acts
as_____. *

1 point

⭕ Reducing agent

⭕ Catalyst

🔘 Oxidizing agent

⭕ Indicator

Volume of different concentrations (0.1, 0.2, 0.3 , 0.4 and 0.5 %) of polymer solution used for each viscosity measurement are _____. * 1 point

○ varies with respect to polymer used

⦿ remains fixed

○ varies with respect to the size of the Ostwald viscometer

○ varies with respect to concentration

The unit of Conductance is _____. * 1 point

○ Ampere

○ Volts

⦿ Mho

○ Ohm

Which one of the following is NOT a unit for hardness? * 1 point

○ Degree French

○ Degree clarke

⦿ Degree centigrade

○ Parts per million

When a mixture of sodium carbonate and sodium hydroxide solution is titrated against HCl solution, the Phenolphthalein end point correspond to *

1 point

○ Neutralization of hydroxyl ions and carbonate ions

○ Neutralization of hydroxyl ions only

○ Neutralization of carbonate ions only

⦿ Neutralization of hydroxyl ions and half of carbonate ions

⬤ Send me a copy of my responses.

Back   Submit

Never submit passwords through Google Forms.

Google Forms