Laboratory Slip Test

* Required

Untitled Section

Which indicator is used in Mohr's method? *

- (●) Potassium Chromate
- ( ) Silver Nitrate
- ( ) Potassium dichromate
- ( ) Silver Chromate

Estimation of chloride reaction is *

- ( ) Redox reaction
- ( ) Equlibrium reaction
- (●) Precipitation reaction
- ( ) Catalytic reaction

A precipitation reaction is a double displacement reaction taking place between *

- ○ Acids and bases
- ◉ two aqueous ionic compounds
- ○ two bases
- ○ two acids

Phenolphthalein color in basic medium is *

- ◉ Pink
- ○ Orange
- ○ Yellow
- ○ Colourless

One ppm is equal to _____. *

- ○ 100 mg / L
- ○ 10 mg / L
- ◉ 1000 mg / L
- ○ 500 mg / L

Temporary hardness in water can be removed by: *

- ○ adding soda
- ○ distillation
- ◉ boiling
- ○ adding lime-soda

The color of phenolphthalein indicator in acid solution is *

- ○ Pink
- ○ Yellow
- ◉ Colourless
- ○ Orange

Hardness of water is conventionally expressed in terms of equivalent amount of _____. *

- ○ H2CO3
- ○ MgCO3
- ◉ CaCO3
- ○ Na2CO3

A buffer solution comprises which of the following? *

- ○ a weak acid in solution
- ○ a strong acid in solution
- ○ a weak base in solution
- ◉ a weak acid and its conjugate base in solution

In Mohr's method the solution needs to be near neutral, because *

- ○ Silver chloride forms at high pH,
- ○ Silver precipitates at low pH
- ◉ Chromate forms H2CrO4 at low pH, which delays the formation of the precipitate.
- ○ Potassium chromate dissolves at high pH.

In determination of mixture of bases by titration method, the amount of Sodium Hydroxideis calculated as---. *

- ○ N x Equivalent mass ofSodium Carbonate / 10
- ○ N [OH and CO32- portion] x Equivalent mass of Sodium Hydroxide and Sodium carbonate / 10
- ◉ N [OH portion] x Equivalent mass ofSodium Hydroxide / 10
- ○ N [CO32- portion] x Equivalent mass of Sodium carbonate /10

The equivalent weight of Sodium Carbonate [Na2CO3] is *

- ○ 40
- ◉ 53
- ○ 55.85
- ○ 63

Temporary hardness of water is caused due to the presence of dissolved *

- ○ calcium hydrogen carbonates only
- ○ magnesium hydrogen carbonates only
- ○ Sulphates and chlorides of calcium or magnesium
- ◉ calcium hydrogen carbonates and magnesium hydrogen carbonates

Which of the following does not cause the permanent hardness in water? *

- ○ Nitrates
- ○ Sulphates
- ○ Chlorides
- ◉ Bicarbonates

Which of the following is not a unit of hardness? *

- ○ Parts per million
- ◉ Degree centigrade
- ○ Degree clarke
- ○ Degree French

Which type of reaction occurs in the following reaction $AgNO_3 + NaCl \rightarrow AgCl + NaNO_3$? *

- ○ Displacement reaction
- ○ Single replacement
- ○ Decomposition
- ◉ Double displacement reaction

What is the role of chromate ions in chloride estimation? *

- ○ It acts as a reducing agent
- ○ It acts as a buffer
- ◉ It acts as an indicator
- ○ It acts as an oxidizing agent

Which of the following is not a primary standard? *

- ○ NaCl
- ○ Anhydrous Na2CO3
- ◉ AgNO3
- ○ Oxalic acid

Soft water + Buffer + EBT ----------à *

- ○ Appearance of wine-red colour
- ◉ Appearance of steel blue colour
- ○ Formation of weak complex
- ○ Formation of brown precipitate

When pH is below 8.5 the indicator ---- is colourless. *

- ○ EBT
- ○ Methyl orange
- ◉ Phenolphthalein
- ○ K2CrO4

When basic solution is titrated against HCl in the burette with Methyl orange indicator, the end point is the color change from *

- ○ Yellow to Violet
- ○ Orange to Yellow
- ○ Appearance of Pink color
- ◉ Yellow to Orange

A neutralization reaction is a --------- reaction taking place between the acids and the bases. *

- ◉ double displacement
- ○ Displacement
- ○ Substitution
- ○ Addition

When mixture of sodium carbonate and sodium hydroxide solution is titrated against HCl solution, the Phenolphthalein end point correspond to *

- ○ Neutralization of OH- ions and CO32- ions
- ○ Neutralization of OH- ions only
- ○ Neutralization of CO32- ions only
- ◉ Neutralization of OH- ions and half of CO32- ions

What is the advantage of Mohr's method? *

- ○ A Very clear colour change
- ◉ Simple method
- ○ Capability for different PH
- ○ Must be 1M nitric acid solution

What is the indicator used for estimation of hardness? *

- ○ Phenolphthalein
- ○ Methyl orange
- ◉ Eriochrome Black – T
- ○ Potassium dichromate

Permanent hardness of water cannot be removed by *

- ○ Adding soda
- ○ Adding lime soda
- ○ Distillation
- ◉ Boiling

In the EDTA method, the purpose of adding a buffer is _____. *

- ○ to maintain the pH of 6-8 range
- ◉ to maintain the pH of 8-10 range
- ○ to maintain the pH of 4-6 range
- ○ to maintain the conc. of the reagent

Methyl orange is *

- ◉ Pink in acidic medium, yellow in basic medium
- ○ Yellow in acidic medium, pink in basic medium
- ○ Colourless in acidic medium, pink in basic medium
- ○ Pink in acidic medium, colourless in basic medium

What is the pH range in which chloride determination using Mohr's method is conducted? *

- ○ < 3
- ○ 5
- ○ > 12
- ◉ 6 -9

Why do we have to standardize the AgNO3 solution? *

- ○ To find the normality of NaCl
- ○ To calculate the normality of AgCl
- ◉ To find the normality of AgNO3
- ○ To calculate the volume of NaCl

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