12. One limitation of the Brønsted-Lowry Theory of acids and bases is that it

(a) does not explain reactions between acidic and basic oxides, such as $SO_3(g) + MgO(s) \rightarrow MgSO_4(s)$, as they do not involve the transfer of protons.
(b) does not explain the production of a neutral salt solution resulting from the reaction between a strong acid and strong base.
(c) links acids and bases into conjugate acid-base pairs rather than accounting for the transfer of protons.
(d) cannot explain the acidity and basicity of acidic and basic salts.