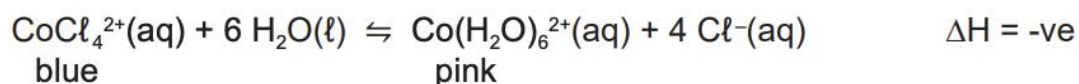


**Question 26****(9 marks)**

Consider the following system that is at equilibrium. Cobalt(II) chloride is dissolved in concentrated hydrochloric acid. The colour of the solution at initial equilibrium is blue.



For each of the applied changes after equilibrium is re-established, predict the:

- shift in equilibrium position (left, right or no change)
- rate of the forward reaction compared to the original rate (increase, decrease or no change)
- colour of the reaction mixture.

Do **not** use arrows to show direction.

Change	Shift in equilibrium position (left, right or no change)	Rate of the forward reaction compared to the original rate (increase, decrease or no change)	Colour of reaction mixture
A few drops of $\text{AgNO}_3(\text{aq})$ are added			
A small volume of water is added			
The system is heated			