- 3. Which of the following represents an acid-base reaction in which the underlined species is acting as a Brønsted-Lowry acid?
  - (a)  $2 \text{ CrO}_4^{2-} + 2 \text{ <math>\underline{HCO}_3^{-} \ } \ \ \ \text{Cr}_2\text{O}_7^{2-} + 2 \text{ CO}_3^{2-} + \text{H}_2\text{O}$
  - (b)  $CH_3CH_2CH_2COOH + NaOH \Leftrightarrow CH_3CH_2CH_2COONa + H_2O$
  - (c)  $NH_3 + O^{2-} = NH_2^- + OH^-$
  - (d)  $\overline{HClO_3} + \underline{CH_3NH_2} \iff ClO_3^- + \underline{CH_3NH_3^+}$