Gas-Liq-Solids Three-laws-Thermo

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Outline

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General Notice

- unit: Kelvin(K) & $J \cdot K^{-1} \cdot mol^{-1}$

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Gas

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Why do we study kinetic molecular theory?

- Graham's law of effusion?

Understanding from A New Point of View

Here we will discuss the **ideal gas equation** from a new point of view, *i.e.*, **kinetic molecular theory**(KMT).

Understanding from A New Point of View

First we should get aware of the **prerequisite** of KMT.

[1] Sun, Ting, *CHEM2100J-FA21-Ch5-6*, pp. 35.

Understanding from A New Point of View

First we should get aware of the **prerequisite** of KMT. Recall what has been taught in lectures.

1. A gas is in continuous random motion
2. Gas molecules are infinitesimally small
3. They move in straight lines until collision
4. Gas molecules do not influence one another except during collisions
5. The collisions are elastic

Prerequisites of KMT shown in slides[1]

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Liquid

Outline

An example of blocks

example

This is an example of block.

An example of blocks

example

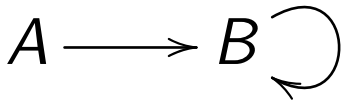
This is an example of block.

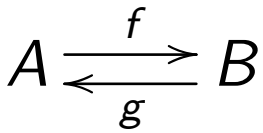
This is another block.

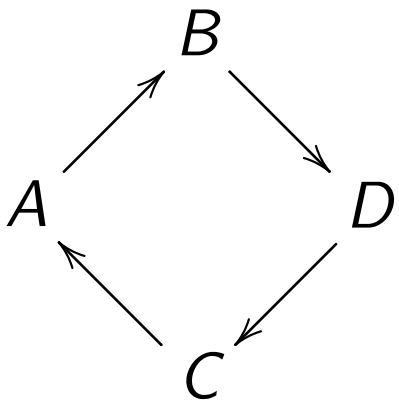
Examples of figures and tables

Figure: An example of figure

Examples of Graphs







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- Smoot Salute! `http://web.mit.edu/spotlight/smoot-salute`.