

Engineering Optics

Lecture 22

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by

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Problem

$P_1 = 0.25, P_2 = 0.5, P_3 = 0.75$ are the probability of occurrence per second of three different atoms. Which wave will be having more life time at excited state.

Answer:

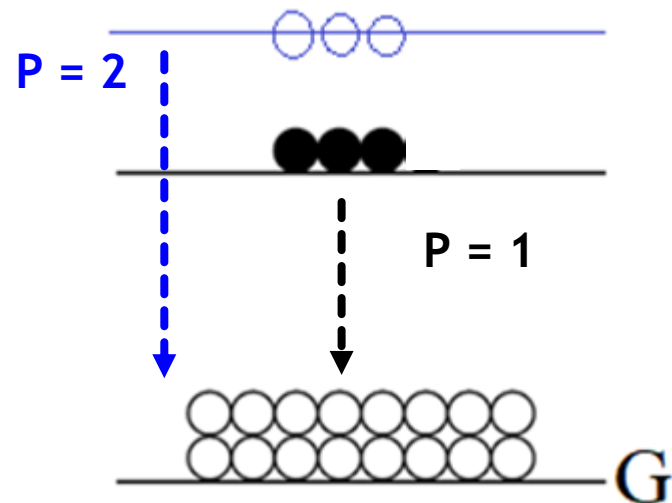
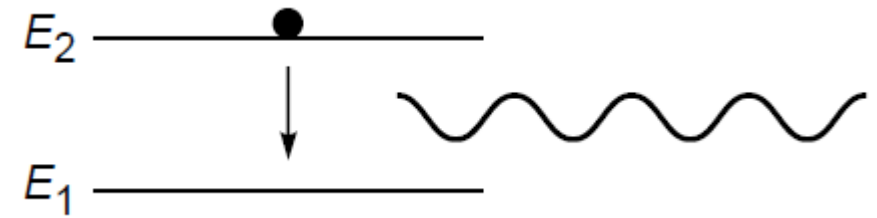
$$P \propto \frac{1}{\tau} \qquad \tau \propto \frac{1}{P}$$

$$P_1 = 0.25, P_2 = 0.5, P_3 = 0.75$$

$\Rightarrow 1^{\text{st}}$ atom will be having more life time

Probability and life time of a state

Keep in mind that the transition rate, the number of atoms making transitions per second, divided by the number of atoms, is the probability of a transition occurring per second, \mathcal{P} . Consequently, the probability per second of spontaneous emission is $\mathcal{P}_{sp} = A_{ji}$.

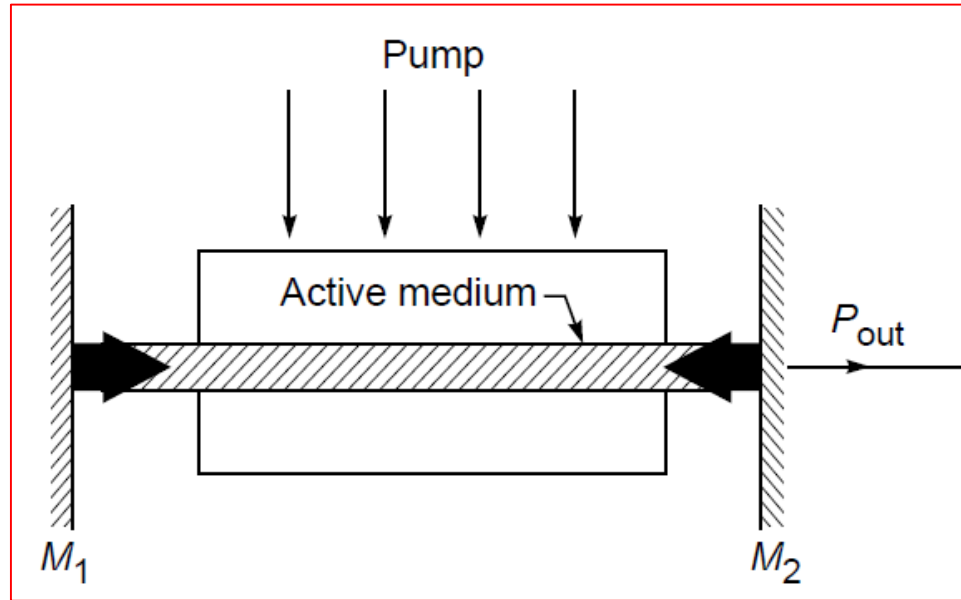


Q: most preferred state?

For a single excited atom making a spontaneous transition to a lower state, the inverse of the transition probability per second is the **mean life or lifetime** of the excited state τ . Thus (operating under conditions that exclude any other mechanism but spontaneous emission), if N atoms are in that excited state, the total rate of transitions, that is, the number of emitted photons per second, is $N\mathcal{P}_{sp} = NA_{ji} = N/\tau$. A low-transition probability means a long lifetime.

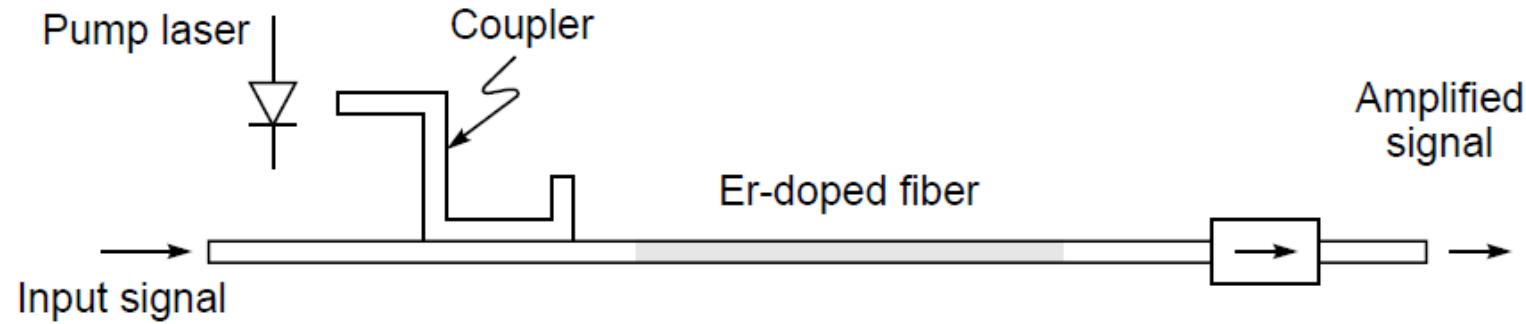
Main Components of the Laser

- ▶ 1. Active medium
- ▶ 2. Pumping source
- ▶ 3. Optical resonator

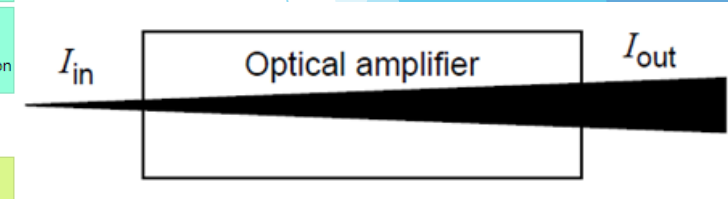


Points to note: Some of the energy is coupled back to the system \rightarrow oscillator.

Understanding Optical Amplification

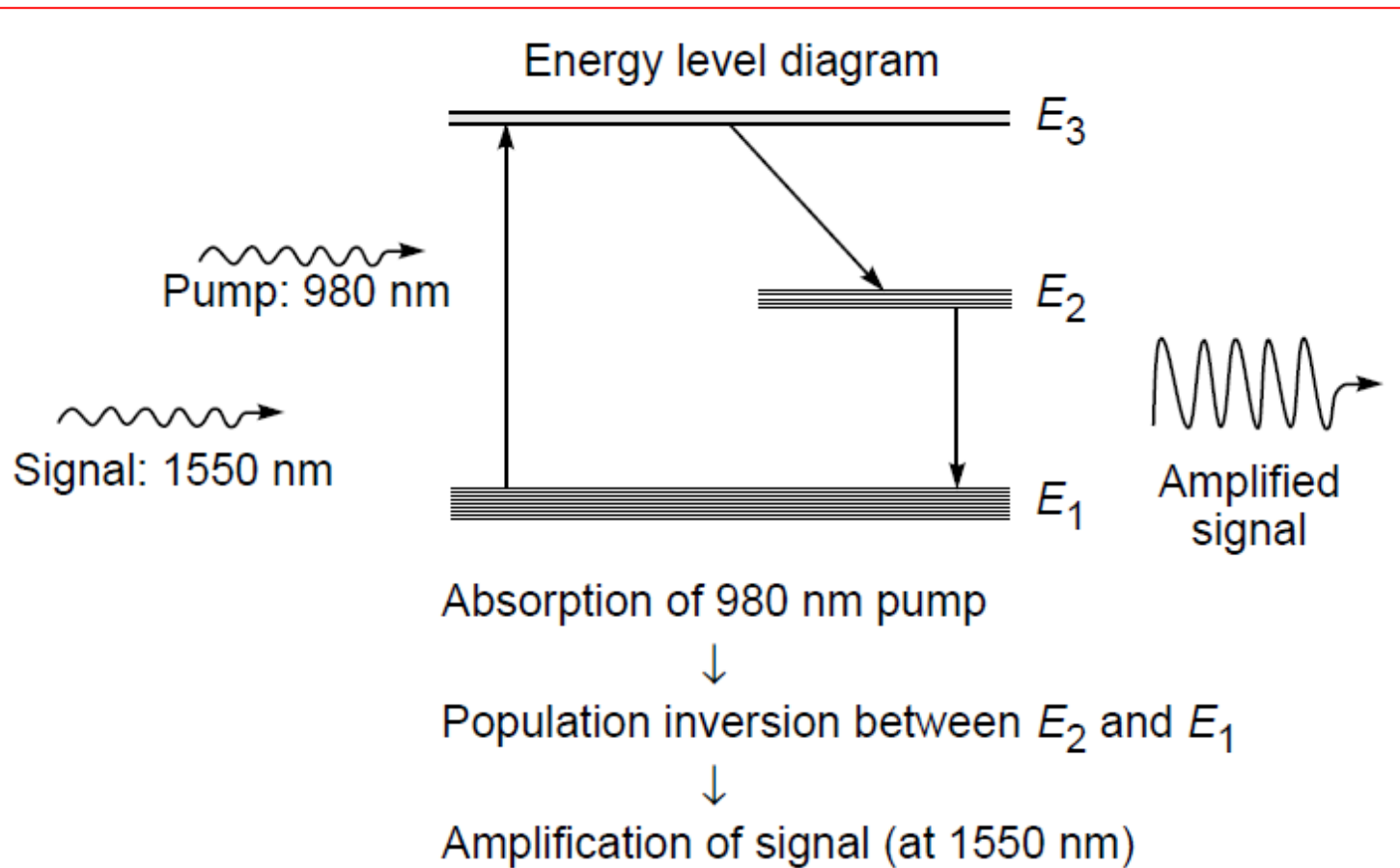


55 Cs Caesium 1	56 Ba Barium 2	57–71	72 Hf Hafnium 4	73 Ta Tantalum 5	74 W Tungsten 4 6	75 Re Rhenium 4	76 Os Osmium 4	77 Ir Iridium 3 4	78 Pt Platinum 2 4	79 Au Gold 3	80 Hg Mercury 1 2	81 Tl Thallium 1 3	82 Pb Lead 2 4	83 Bi Bismuth 3	84 Po Polonium -2 2 4	85 At Astatine -1 1	86 Rn Radon 2	
87 Fr Francium 1	88 Ra Radium 2		89–103	104 Rf Rutherfordium 4	105 Db Dubnium 5	106 Sg Seaborgium 6	107 Bh Bohrium 7	108 Hs Hassium 8	109 Mt Meitnerium 3 4	110 Ds Darmstadtium 4	111 Rg Roentgenium 3	112 Cn Copernicium 2	113 Nh Nihonium 3	114 Fl Flerovium 2 4	115 Mc Moscovium 3	116 Lv Livermorium 2	117 Ts Tennessine 3	118 Og Oganesson 2
Oxidation states 3																		
Oxidation states are the number of electrons added to or removed from an element when it forms a chemical compound.																		
6			57 La Lanthanum 3	58 Ce Cerium 3 4	59 Pr Praseodymium 3	60 Nd Neodymium 3	61 Pm Promethium 3	62 Sm Samarium 3	63 Eu Europium 2 3	64 Gd Gadolinium 3	65 Tb Terbium 3	66 Dy Dysprosium 3	67 Ho Holmium 3	68 Er Erbium 3	69 Tm Thulium 3	70 Yb Ytterbium 3	71 Lu Lutetium 3	
			7			89 Ac Actinium 3	90 Th Thorium 4	91 Pa Protactinium 5	92 U Uranium 6	93 Np Neptunium 5	94 Pu Plutonium 4	95 Am Americium 3	96 Cm Curium 3	97 Bk Berkelium 3	98 Cf Californium 3	99 Es Einsteinium 3	100 Fm Fermium 3	101 Md Mendelevium 3



- ▶ understand optical amplification → EDFA
- ▶ **Erbium doped fiber amplifier**
- ▶ The EDFA: consists of a silica optical fiber the core doped with erbium oxide (Er_2O_3)
- ▶ light is guided through the optical fiber
- ▶ three discrete energy levels of Er^{3+} ion in silica host glass.

EDFA



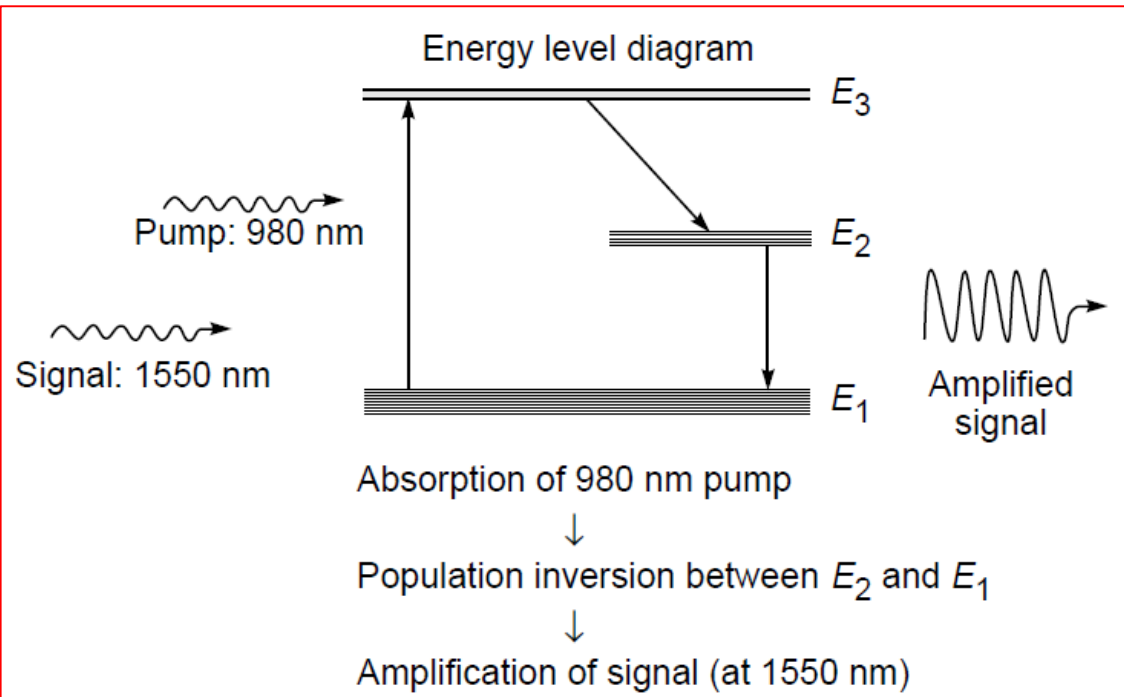
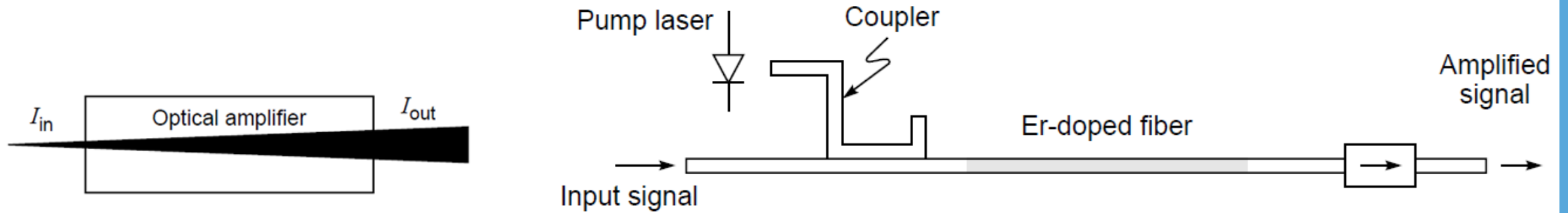
- ▶ $E_3 - E_1 = 1.3 \text{ eV} \rightarrow 980 \text{ nm}$
- ▶ $E_2 - E_1 = 0.81 \text{ eV} \rightarrow 1550 \text{ nm}$
- ▶ When a trigger of 980 nm is fed \rightarrow Er atoms jump to $E_3 \rightarrow$ *pump*
- ▶ Atoms in E_3 jumps to $E_2 \rightarrow$ *heating*
- ▶ State E_2 is **metastable** \rightarrow long lifetime (few milliseconds).
- ▶ large lifetime of state E_2 than $E_3 \rightarrow$ the population of the erbium atoms in state E_2 grows with time
- ▶ Population inversion can be achieved
- ▶ $N_2 > N_1$
- ▶ Send a signal beam (1550 nm) \rightarrow it gets amplified
- ▶ Why? \rightarrow stimulated emission of radiation

Er^{3+} concentration $\approx 7 \times 10^{24} \text{ ions m}^{-3}$ pump power $\approx 5 \text{ mW}$

and, the optimum length of the erbium doped fiber $\approx 7 \text{ m}$.

$$\text{Gain (dB)} = 10 \log \frac{P_{\text{output}}}{P_{\text{input}}}$$

Understanding Optical Amplification through EDFA

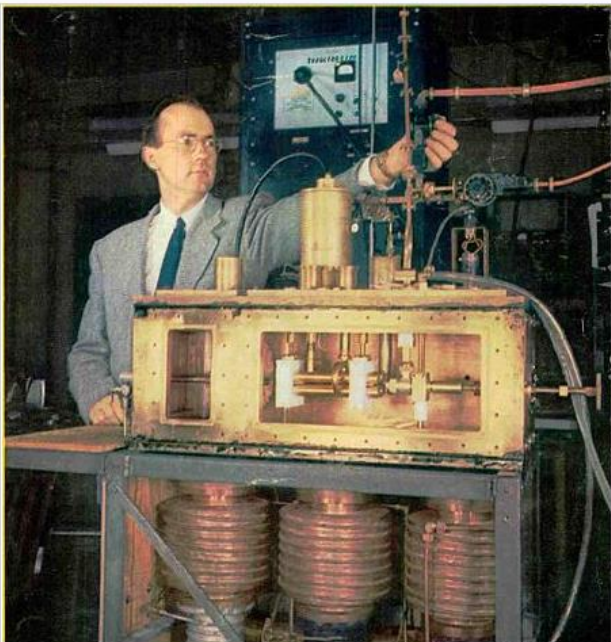


Loss: scattering from the laser medium, absorption at the mirrors

When the laser oscillates in steady state, the losses are exactly compensated for by the gain provided by the medium

→ wave coming out of the laser can be represented as a continuous wave

Do you know?



First prototype ammonia maser and inventor **Charles H. Townes**. The ammonia nozzle is at left in the box, the four brass rods at center are the **quadrupole** state selector, and the resonant cavity is at right. The 24 GHz microwaves exit through the vertical **waveguide** Townes is adjusting. At bottom are the vacuum pumps. [wikipedia](#)

The Nobel Prize in Physics 1964



Photo from the Nobel Foundation archive.
Charles Hard Townes
Prize share: 1/2



Photo from the Nobel Foundation archive.
Nicolay Gennadiyevich Basov
Prize share: 1/4

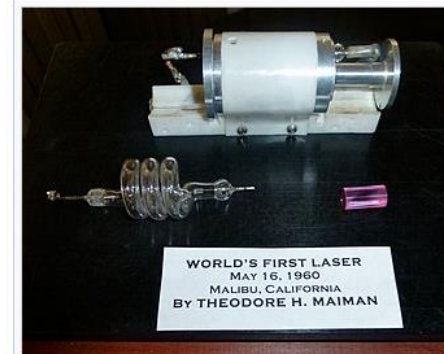


Photo from the Nobel Foundation archive.
Aleksandr Mikhailovich Prokhorov
Prize share: 1/4



Maiman with his laser in July 1960. [aps](#)

<https://www.aps.org/publications/apsnews/2010/05/physicshistory.cfm>



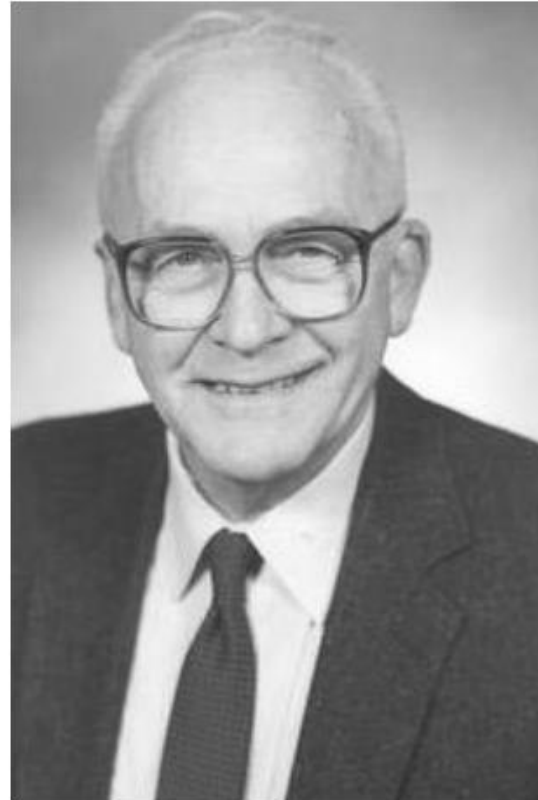
World's first laser [wikipedia](#)

The Nobel Prize in Physics 1964 was divided, one half awarded to Charles Hard Townes, the other half jointly to Nicolay Gennadiyevich Basov and Aleksandr Mikhailovich Prokhorov "for fundamental work in the field of quantum electronics, which has led to the construction of oscillators and amplifiers based on the maser-laser principle."

THE FIBER LASER

- ▶ put the doped fiber between two mirrors (resonator), + appropriate pump → fiber laser
- ▶ 1960: Maiman's demonstration of the first ever laser
- ▶ 1961: Elias Snitzer wrapped a flash lamp around a glass fiber (having a 300 mm core doped with Nd^{3+} ions clad in a lower-index glass) and when suitable feedback was applied, the first fiber laser was born.
- ▶ *Contribution in EDFA*
- ▶ *commercially available in the market*
- ▶ Applications: widespread → in welding, cutting, drilling, and in medical surgery.

Elias Snitzer



OSA Awards & Distinctions

- John Tyndall Award
- Charles Hard Townes Medal
- Fellow

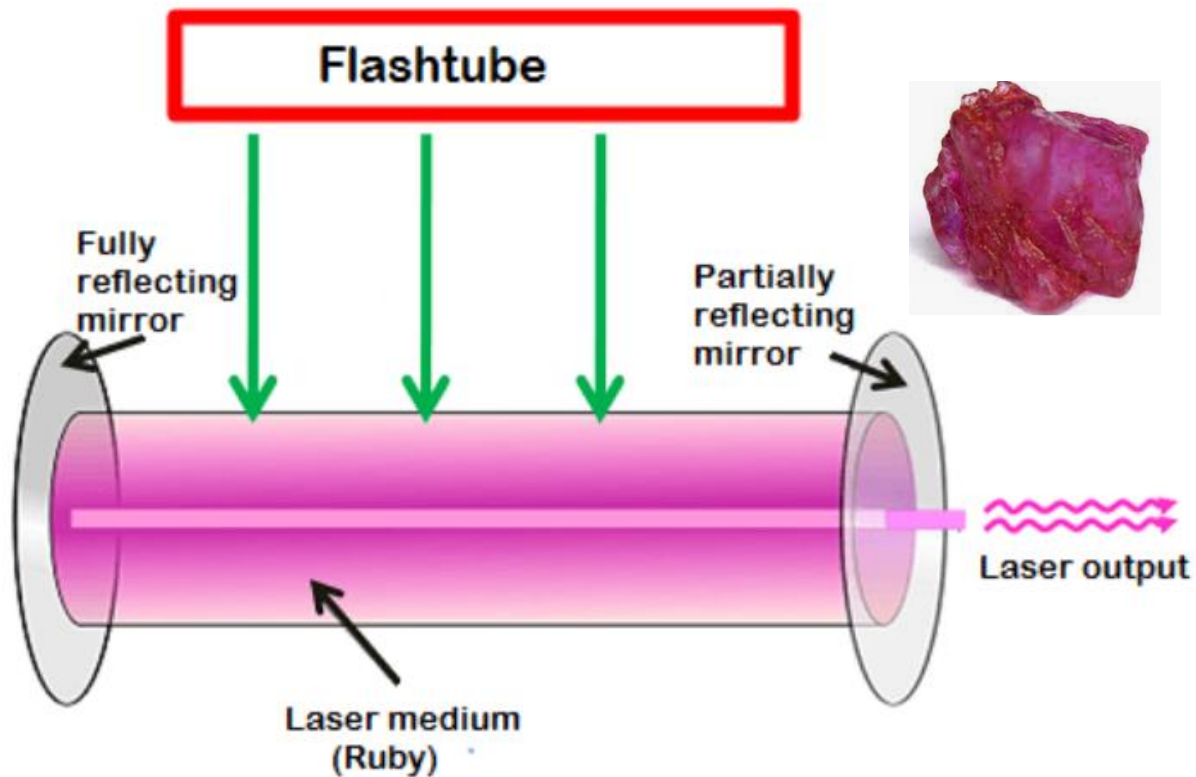
https://www.osa.org/en-us/history/biographies/bios/elias_snitzer/

https://www.photonics.com/Articles/Fiber_Laser_Pioneer_Honored_with_IEEE_Milestone/a52186

Oxidation states are the number of electrons added to or removed from an element when it forms a chemical compound.

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The Ruby LASER



<https://www.physics-and-radio-electronics.com/physics/laser/rubylaserdefinitionconstructionworking.html>

Optics, Ghatak

Maiman's original device → the first operative laser

Active medium → a small, cylindrical, synthetic, pink ruby rod

What is Ruby? Al_2O_3 crystal containing about 0.05 percent (by weight) of Cr_2O_3 .

Helical discharge tube (flashtube) → pump

The rod's end faces were polished flat, parallel and normal to the axis. Then both were silvered to form the **resonator**

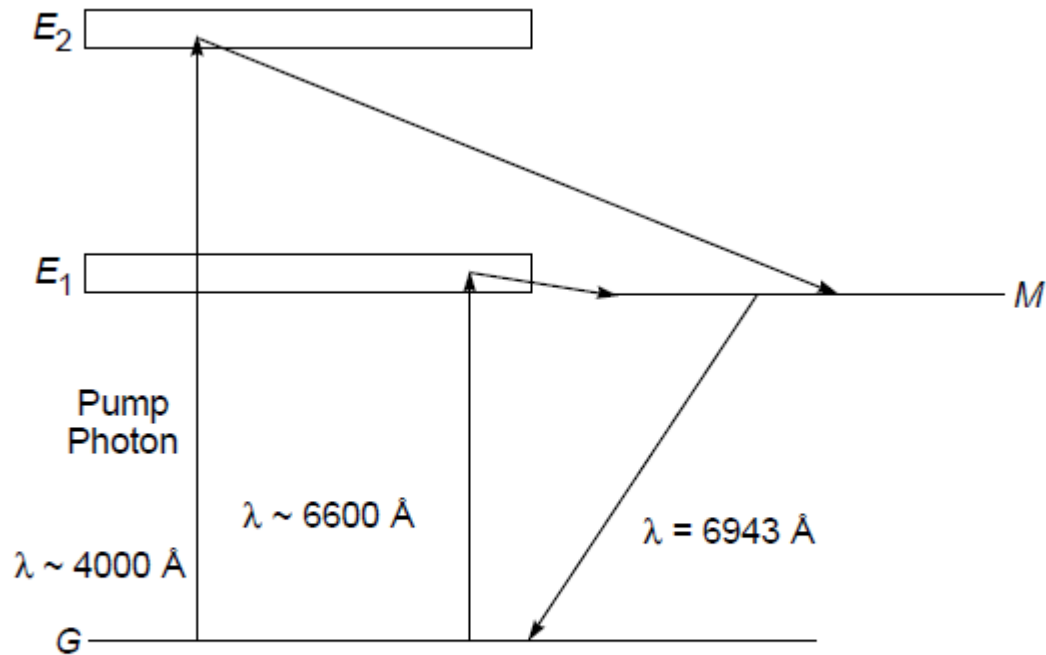
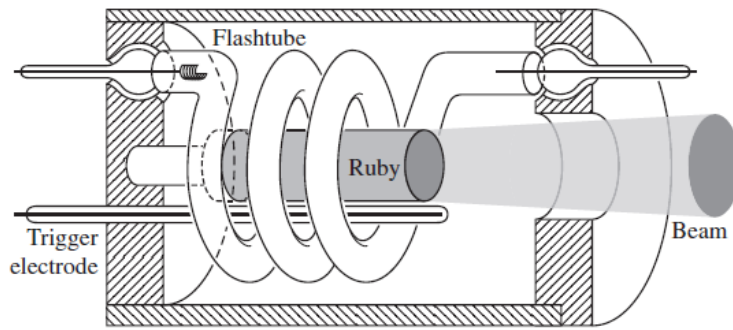


Maiman with his laser in July 1960. 



World's first laser 

Working principle



The energy levels of the chromium ion; G and M represent the ground and metastable states, respectively.

1. photons are produced by the flash lamp
 2. The chromium ion in its ground state can absorb a photon ($\sim 4000 \text{ \AA}$) and make a transition to $E_2 \rightarrow$ by optical pumping OR to E_1 (6600 \AA)
 3. Once in E_2 or $E_1 \rightarrow$ it immediately makes a nonradiative transition (in a time $\sim 10^{-8} \text{ s}$) to the metastable state M (*3 ms lifetime*)
 4. the excess energy (transition from E_2/E_1 to M) is absorbed by the lattice and does not appear as EM radiation.
 5. M has a very long life, the number of atoms in this state keeps increasing and one may achieve population inversion between states M and G .
 6. Once population inversion is achieved, light amplification can take place, with two reflecting ends of the ruby rod forming a cavity.
- ▶ The ruby laser is an example of a three-level laser.
 - ▶ Applications: medical and cosmetic procedures, holography

Why is Ruby red?



- ▶ From blood red to pink → purple
- ▶ Reason? Cr^{3+} concentration in Al_2O_3 ← white
- ▶ chromium atoms absorb certain wavelengths of light (blue, green) and reflect light in the red part of the visible light spectrum.
- ▶ This reflected red light is what your eyes see and gives rubies their distinctive red coloring.

Thank You