

## (Chapter 12)(Aldehydes Ketones and Carboxylic Acids)

### XII

#### Question 12.1:

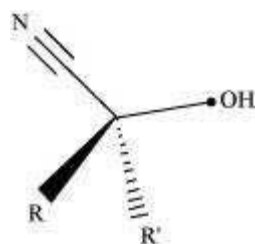
What is meant by the following terms? Give an example of the reaction in each case.

- (i) Cyanohydrin (ii) Acetal  
(iii) Semicarbazone (iv) Aldol  
(v) Hemiacetal (vi) Oxime  
(vii) Ketal (viii) Imine  
(ix) 2,4-DNP-derivative (x) Schiff's base

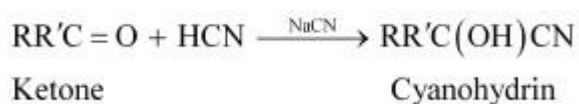
Answer

#### (i) Cyanohydrin:

Cyanohydrins are organic compounds having the formula  $RR'C(OH)CN$ , where R and R' can be alkyl or aryl groups.



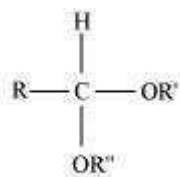
Aldehydes and ketones react with hydrogen cyanide (HCN) in the presence of excess sodium cyanide (NaCN) as a catalyst to form cyanohydrin. These reactions are known as cyanohydrin reactions.



Cyanohydrins are useful synthetic intermediates.

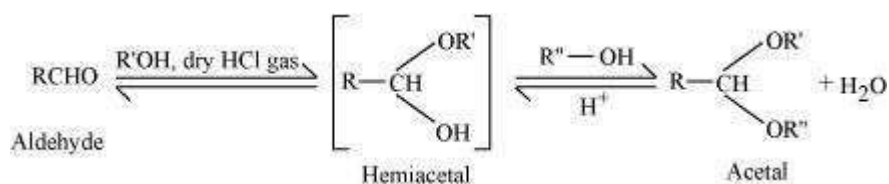
#### (ii) Acetal:

Acetals are gem-dialkoxy alkanes in which two alkoxy groups are present on the terminal carbon atom. One bond is connected to an alkyl group while the other is connected to a hydrogen atom.



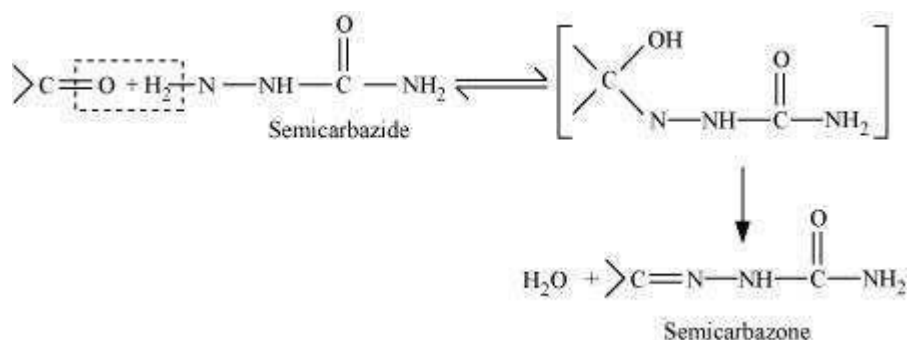
General structure of an acetal

When aldehydes are treated with two equivalents of a monohydric alcohol in the presence of dry HCl gas, hemiacetals are produced that further react with one more molecule of alcohol to yield acetal.



### (iii) Semicarbazone:

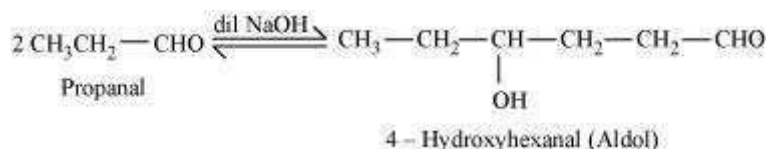
Semicarbazones are derivatives of aldehydes and ketones produced by the condensation reaction between a ketone or aldehyde and semicarbazide.



Semicarbazones are useful for identification and characterization of aldehydes and ketones.

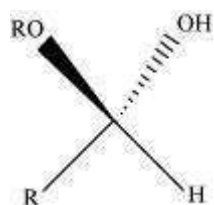
### (iv) Aldol:

A  $\beta$ -hydroxy aldehyde or ketone is known as an aldol. It is produced by the condensation reaction of two molecules of the same or one molecule each of two different aldehydes or ketones in the presence of a base.



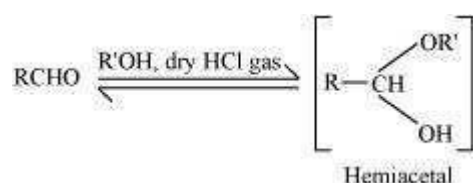
### (v) Hemiacetal:

Hemiacetals are  $\alpha$ -alkoxyalcohols



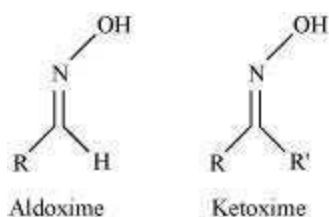
General structure of a hemiacetal

Aldehyde reacts with one molecule of a monohydric alcohol in the presence of dry HCl gas.

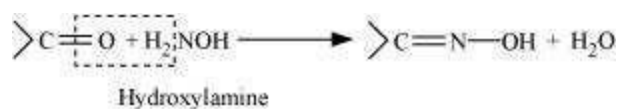


#### (vi) Oxime:

Oximes are a class of organic compounds having the general formula  $\text{RR'CNOH}$ , where R is an organic side chain and R' is either hydrogen or an organic side chain. If R' is H, then it is known as aldoxime and if R' is an organic side chain, it is known as ketoxime.

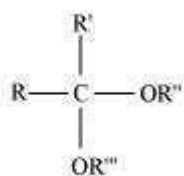


On treatment with hydroxylamine in a weakly acidic medium, aldehydes or ketones form oximes.



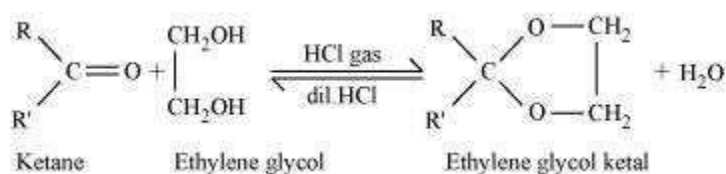
#### (vii) Ketal:

Ketals are gem-dialkoxyalkanes in which two alkoxy groups are present on the same carbon atom within the chain. The other two bonds of the carbon atom are connected to two alkyl groups.



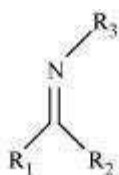
General structure of a ketal

Ketones react with ethylene glycol in the presence of dry HCl gas to give a cyclic product known as ethylene glycol ketals.



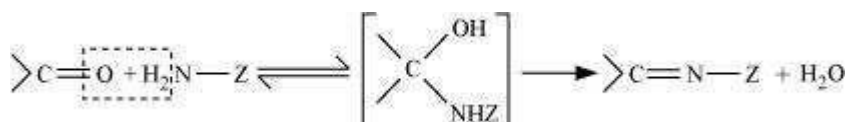
#### (viii) Imine:

Imines are chemical compounds containing a carbon nitrogen double bond.



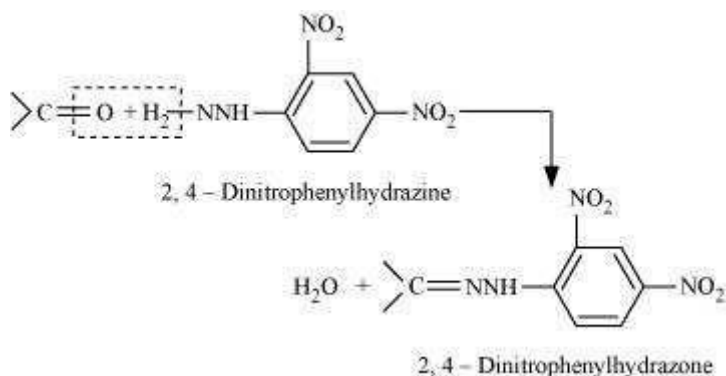
General structure of an imine

Imines are produced when aldehydes and ketones react with ammonia and its derivatives.



#### (ix) 2, 4-DNP-derivative:

2, 4-dinitrophenylhydrazones are 2, 4-DNP-derivatives, which are produced when aldehydes or ketones react with 2, 4-dinitrophenylhydrazine in a weakly acidic medium.



To identify and characterize aldehydes and ketones, 2, 4-DNP derivatives are used.

**(x) Schiff's base:**

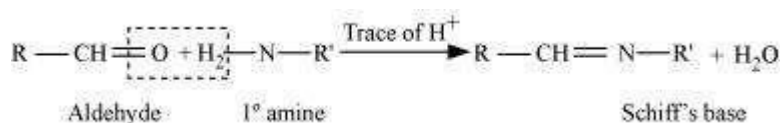
Schiff's base (or azomethine) is a chemical compound containing a carbon-nitrogen double bond with the nitrogen atom connected to an aryl or alkyl group-but not hydrogen. They have the general formula  $R_1R_2C = NR_3$ . Hence, it is an imine.

It is named after a scientist, Hugo Schiff.



General structure of schiff's base

Aldehydes and ketones on treatment with primary aliphatic or aromatic amines in the presence of trace of an acid yields a Schiff's base.



**Question 12.2:**

Name the following compounds according to IUPAC system of nomenclature:

- (i)  $\text{CH}_3\text{CH}(\text{CH}_3)\text{CH}_2\text{CH}_2\text{CHO}$
  - (ii)  $\text{CH}_3\text{CH}_2\text{COCH}(\text{C}_2\text{H}_5)\text{CH}_2\text{CH}_2\text{Cl}$
  - (iii)  $\text{CH}_3\text{CH}=\text{CHCHO}$
  - (iv)  $\text{CH}_3\text{COCH}_2\text{COCH}_3$
  - (v)  $\text{CH}_3\text{CH}(\text{CH}_3)\text{CH}_2\text{C}(\text{CH}_3)_2\text{COCH}_3$
  - (vi)  $(\text{CH}_3)_3\text{CCH}_2\text{COOH}$
  - (vii)  $\text{OHCC}_6\text{H}_4\text{CHO}$ -*p*
- Answer
- (i) 4-methylpentanal
  - (ii) 6-Chloro-4-ethylhexan-3-one
  - (iii) But-2-en-1-al
  - (iv) Pentane-2,4-dione
  - (v) 3,3,5-Trimethylhexan-2-one

**(vi)** 3,3-Dimethylbutanoic acid

**(vii)** Benzene-1,4-dicarbaldehyde

**Question 12.3:**

Draw the structures of the following compounds.

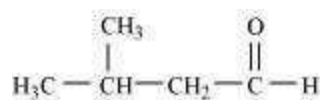
**(i)** 3-Methylbutanal **(ii)** *p*-Nitropropiophenone

**(iii)** *p*-Methylbenzaldehyde **(iv)** 4-Methylpent-3-en-2-one

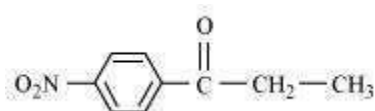
4-Chloropentan-2-one **(vi)** 3-Bromo-4-phenylpentanoic acid  
**(vii)** *p,p'*-Dihydroxybenzophenone **(viii)** Hex-2-en-4-ynoic acid

Answer

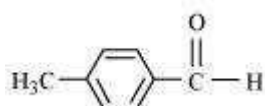
**(i)**



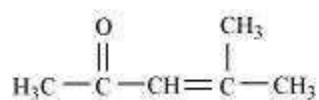
**(ii)**



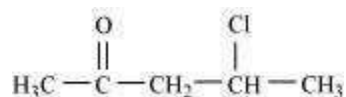
**(iii)**



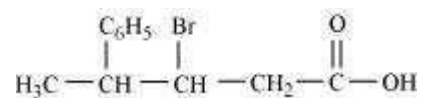
**(iv)**



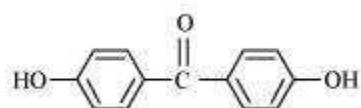
**(v)**



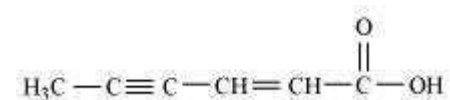
**(vi)**



**(vii)**



**(viii)**



**Question 12.4:**

Write the IUPAC names of the following ketones and aldehydes. Wherever possible, give also common names.

**(i)**  $\text{CH}_3\text{CO}(\text{CH}_2)_4\text{CH}_3$  **(ii)**  $\text{CH}_3\text{CH}_2\text{CHBrCH}_2\text{CH}(\text{CH}_3)\text{CHO}$

**(iii)**  $\text{CH}_3(\text{CH}_2)_5\text{CHO}$  **(iv)**  $\text{Ph}-\text{CH}=\text{CH}-\text{CHO}$



**(vi)**  $\text{PhCOPh}$

Answer

**(i)**  $\text{CH}_3\text{CO}(\text{CH}_2)_4\text{CH}_3$

IUPAC name: Heptan-2-one

Common name: Methyl n-propyl ketone

**(ii)**  $\text{CH}_3\text{CH}_2\text{CHBrCH}_2\text{CH}(\text{CH}_3)\text{CHO}$

IUPAC name: 4-Bromo-2-methylhexanal

Common name: ( $\gamma$ -Bromo- $\alpha$ -methyl-caproaldehyde)

**(iii)**  $\text{CH}_3(\text{CH}_2)_5\text{CHO}$

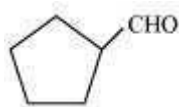
IUPAC name: Heptanal

**(iv)**  $\text{Ph}-\text{CH}=\text{CH}-\text{CHO}$

IUPAC name: 3-phenylprop-2-enal

Common name:  $\beta$ -Phenylacrolein

**(v)**



IUPAC name: Cyclopentanecarbaldehyde

**(vi)**  $\text{PhCOPh}$

IUPAC name: Diphenylmethanone

Common name: Benzophenone

**Question 12.5:**

Draw structures of the following derivatives.

**(i)** The 2,4-dinitrophenylhydrazone of benzaldehyde



**(ii)** Cyclopropanone oxime

**(iii)** Acetaldehydedimethylacetal

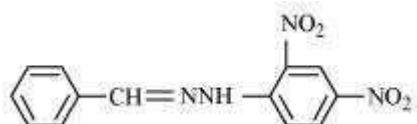
**(iv)** The semicarbazone of cyclobutanone

The ethylene ketal of hexan-3-one

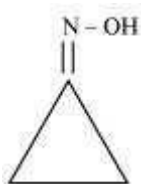
**(vi)** The methyl hemiacetal of formaldehyde

Answer

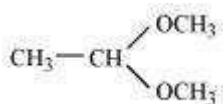
**(i)**



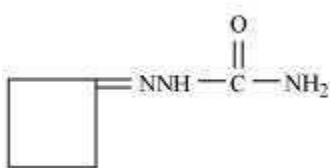
**(ii)**



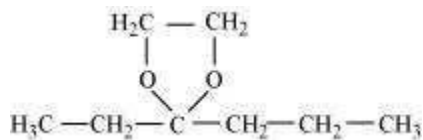
**(iii)**



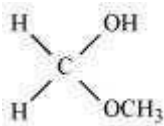
**(iv)**



**(v)**



**(vi)**



**Question 12.6:**

Predict the products formed when cyclohexanecarbaldehyde reacts with following reagents.

(i)  $\text{PhMgBr}$  and then  $\text{H}_3\text{O}^+$

(ii) Tollens' reagent

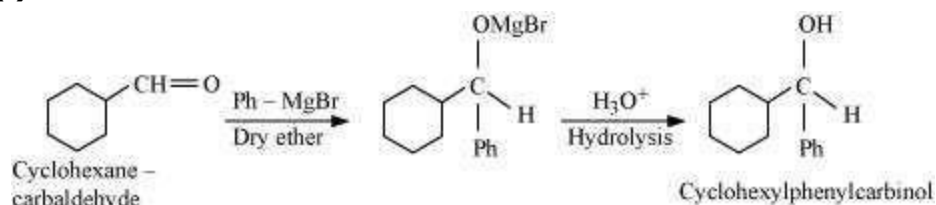
(iii) Semicarbazide and weak acid

(iv) Excess ethanol and acid

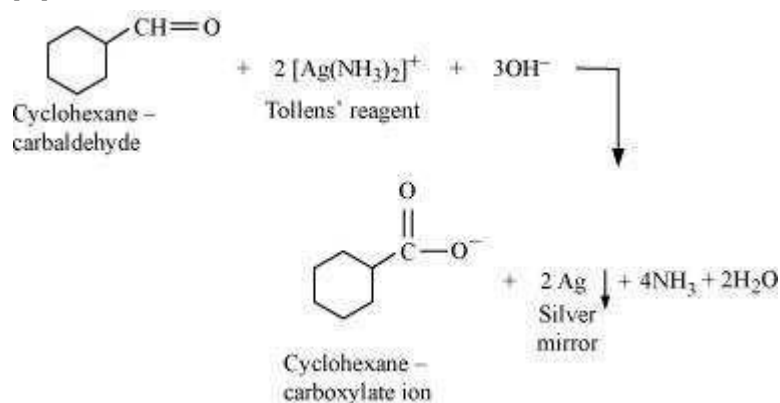
(v) Zinc amalgam and dilute hydrochloric acid

Answer

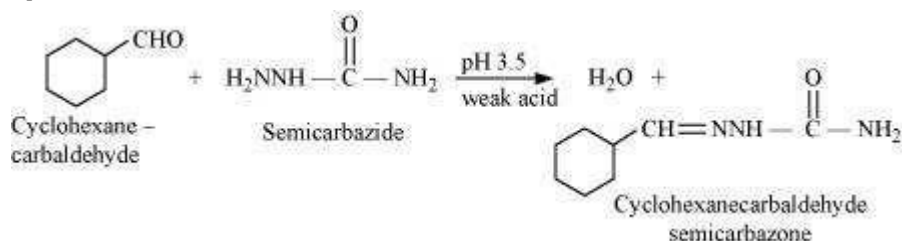
(i)



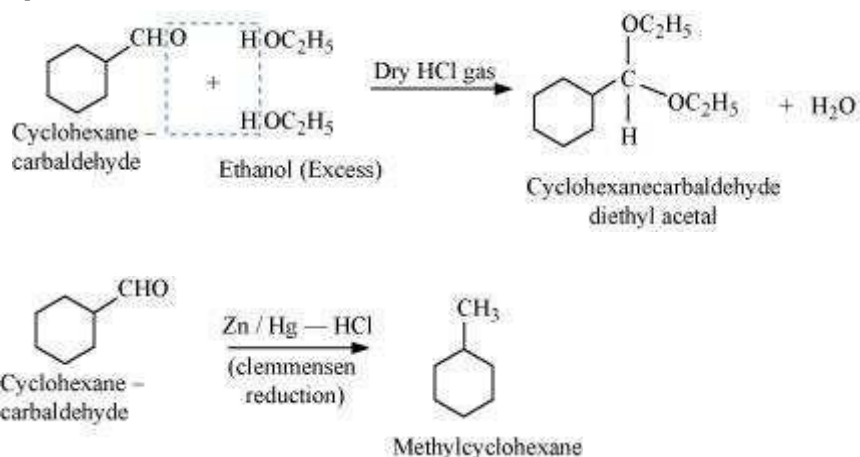
(ii)



(iii)



(iv)



**Question 12.7:**

Which of the following compounds would undergo aldol condensation, which the Cannizzaro reaction and which neither? Write the structures of the expected products of aldol condensation and Cannizzaro reaction.

- (i) Methanal (ii) 2-Methylpentanal  
(iii) Benzaldehyde (iv) Benzophenone  
(v) Cyclohexanone (vi) 1-Phenylpropanone  
(vii) Phenylacetaldehyde (viii) Butan-1-ol  
(ix) 2, 2-Dimethylbutanal

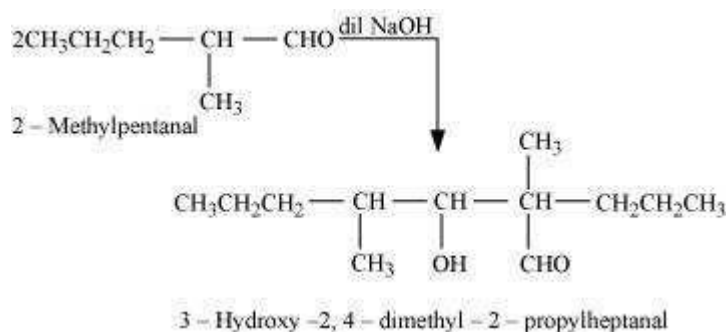
Answer

Aldehydes and ketones having at least one  $\alpha$ -hydrogen undergo aldol condensation. The compounds (ii) 2-methylpentanal, (v) cyclohexanone, (vi) 1-phenylpropanone, and (vii) phenylacetaldehyde contain one or more  $\alpha$ -hydrogen atoms. Therefore, these undergo aldol condensation.

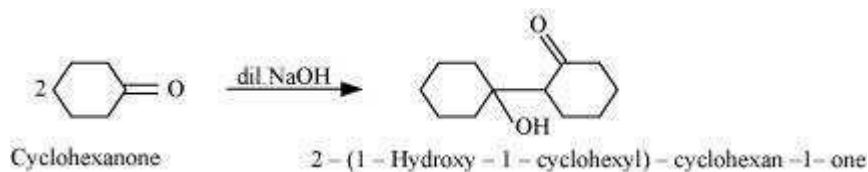
Aldehydes having no  $\alpha$ -hydrogen atoms undergo Cannizzaro reactions. The compounds (i) Methanal, (iii) Benzaldehyde, and (ix) 2, 2-dimethylbutanal do not have any  $\alpha$ -hydrogen. Therefore, these undergo cannizzaro reactions.

Compound (iv) Benzophenone is a ketone having no  $\alpha$ -hydrogen atom and compound (viii) Butan-1-ol is an alcohol. Hence, these compounds do not undergo either aldol condensation or cannizzaro reactions.

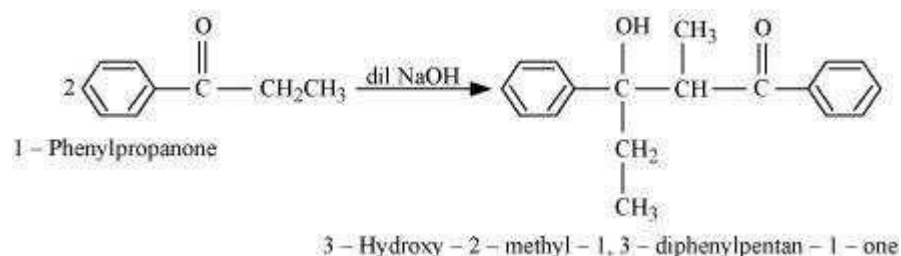
### Aldol condensation (ii)



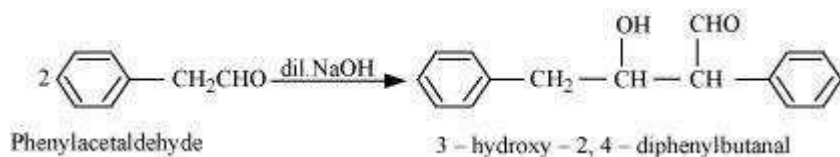
(v)



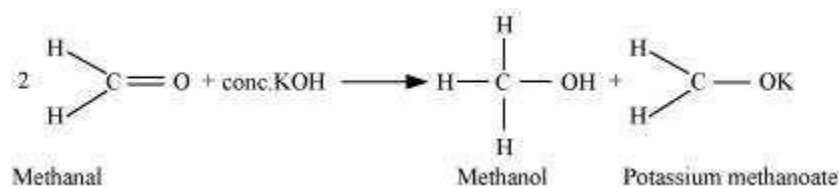
(vi)



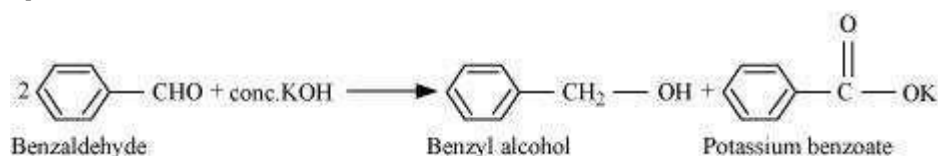
(vii)



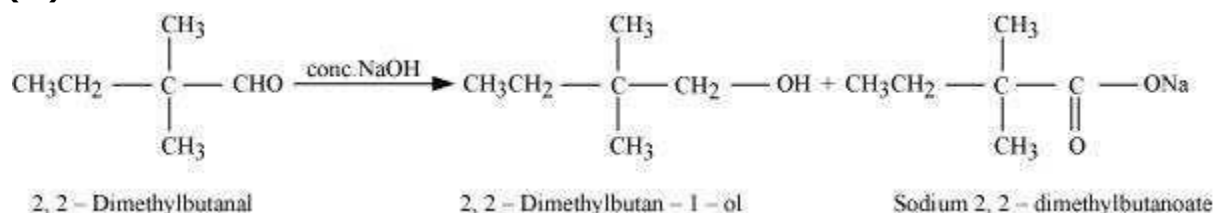
### Cannizzaro reaction (i)



(iii)



(ix)



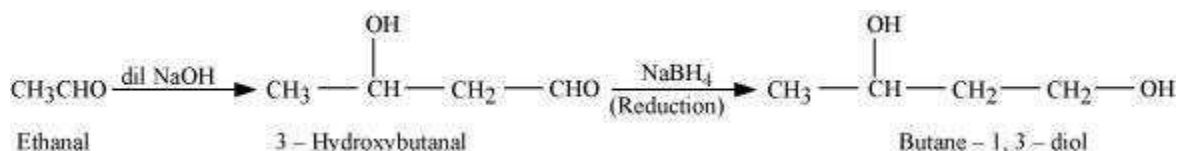
**Question 12.8:**

How will you convert ethanal into the following compounds?

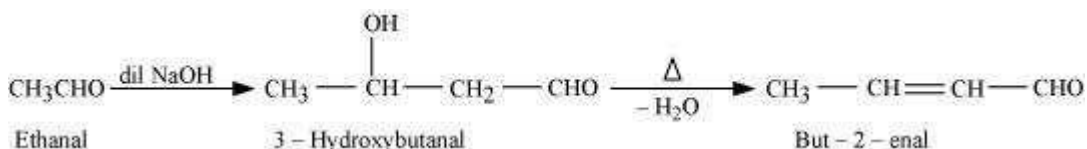
**(i)** Butane-1, 3-diol **(ii)** But-2-enal **(iii)** But-2-enoic acid

Answer

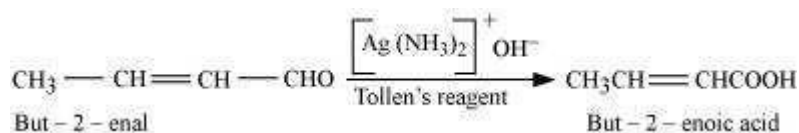
**(i)** On treatment with dilute alkali, ethanal produces 3-hydroxybutanal which on reduction gives butane-1, 3-diol on reduction.



**(ii)** On treatment with dilute alkali, ethanal gives 3-hydroxybutanal which on heating produces but-2-enal.



**(iii)** When treated with Tollen's reagent, But-2-enal produced in the above reaction produces but-2-enoic acid .

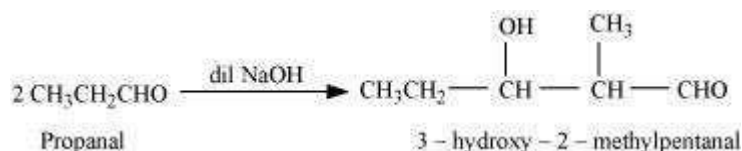


### Question 12.9:

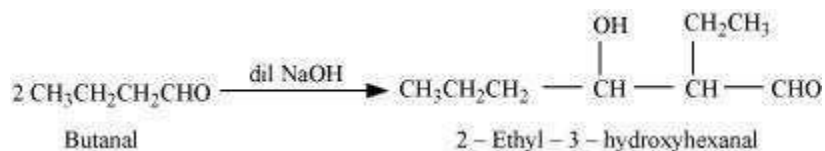
Write structural formulas and names of four possible aldol condensation products from propanal and butanal. In each case, indicate which aldehyde acts as nucleophile and which as electrophile.

Answer

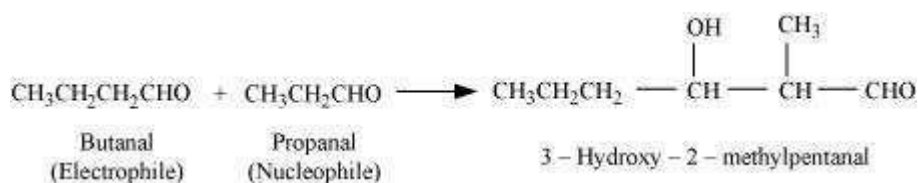
- (i) Taking two molecules of propanal, one which acts as a nucleophile and the other as an electrophile.



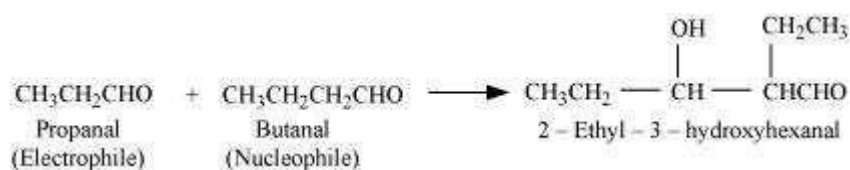
- (ii) Taking two molecules of butanal, one which acts as a nucleophile and the other as an electrophile.



- (iii) Taking one molecule each of propanal and butanal in which propanal acts as a nucleophile and butanal acts as an electrophile.



- (iv) Taking one molecule each of propanal and butanal in which propanal acts as an electrophile and butanal acts as a nucleophile.

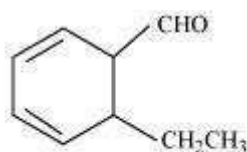


**Question 12.10:**

An organic compound with the molecular formula  $C_9H_{10}O$  forms 2, 4-DNP derivative, reduces Tollens' reagent and undergoes Cannizzaro reaction. On vigorous oxidation, it gives 1, 2-benzenedicarboxylic acid. Identify the compound.

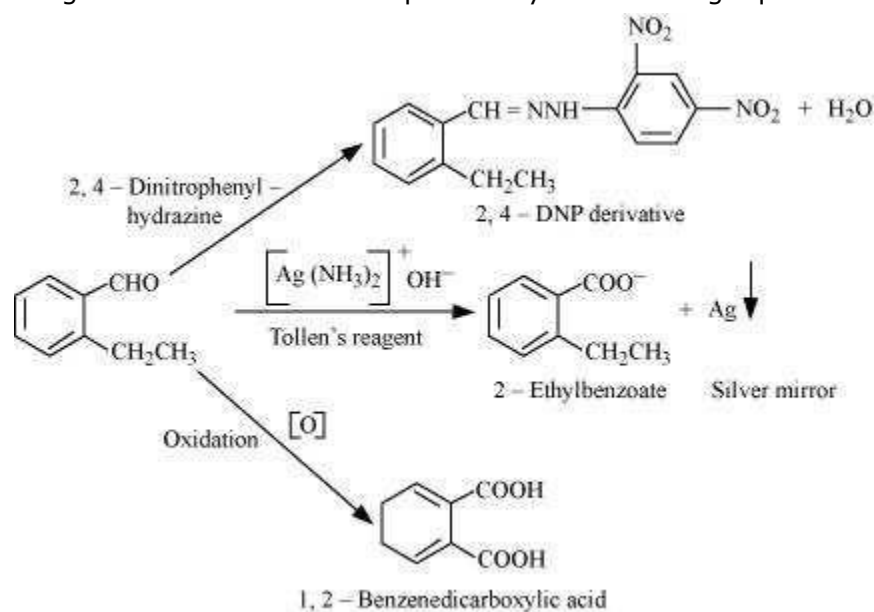
Answer

It is given that the compound (with molecular formula  $C_9H_{10}O$ ) forms 2, 4-DNP derivative and reduces Tollen's reagent. Therefore, the given compound must be an aldehyde. Again, the compound undergoes cannizzaro reaction and on oxidation gives 1, 2benzenedicarboxylic acid. Therefore, the  $-CHO$  group is directly attached to a benzene ring and this benzaldehyde is ortho-substituted. Hence, the compound is 2ethylbenzaldehyde.



2 - Ethylbenzaldehyde

The given reactions can be explained by the following equations.



**Question 12.11:**

An organic compound (A) (molecular formula  $C_8H_{16}O_2$ ) was hydrolysed with dilute sulphuric acid to give a carboxylic acid (B) and an alcohol (C). Oxidation of (C) with chromic acid produced (B). (C) on dehydration gives but-1-ene. Write equations for the reactions involved.

Answer

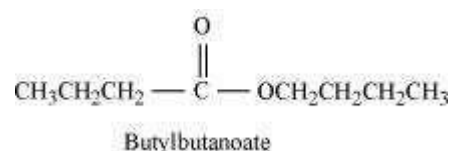
An organic compound A with molecular formula  $C_8H_{16}O_2$  gives a carboxylic acid (B) and an alcohol (C) on hydrolysis with dilute sulphuric acid. Thus, compound A must be an ester. Further, alcohol C gives acid B on oxidation with chromic acid. Thus, B and C must contain equal number of carbon atoms.

Since compound A contains a total of 8 carbon atoms, each of B and C contain 4 carbon atoms.

Again, on dehydration, alcohol C gives but-1-ene. Therefore, C is of straight chain and hence, it is butan-1-ol.

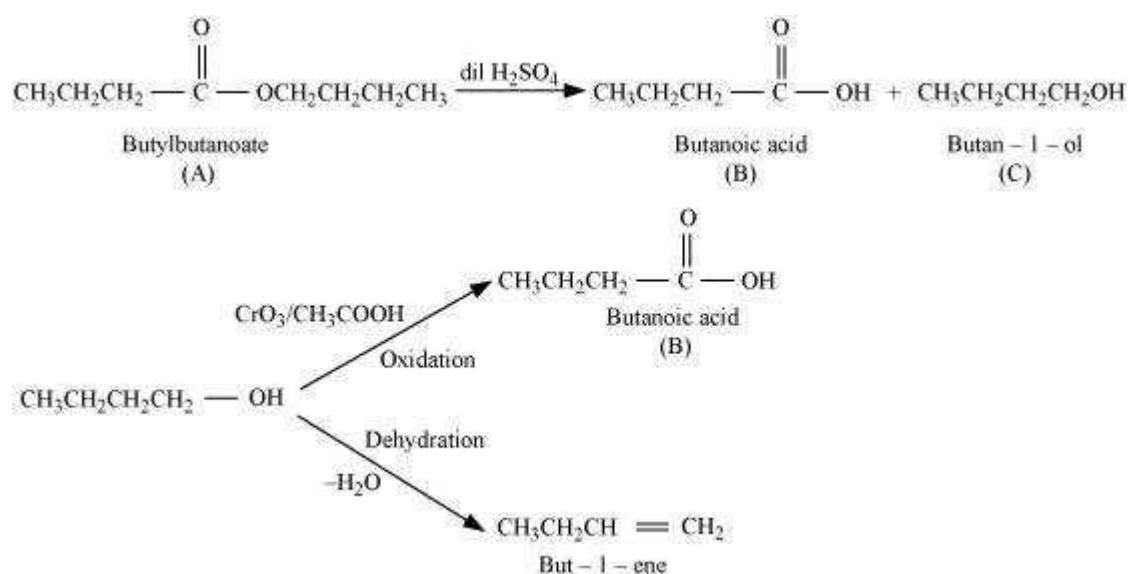
On oxidation, Butan-1-ol gives butanoic acid. Hence, acid B is butanoic acid.

Hence, the ester with molecular formula  $C_8H_{16}O_2$  is butylbutanoate.



All the given reactions can be explained by the following equations.





**Question 12.12:**

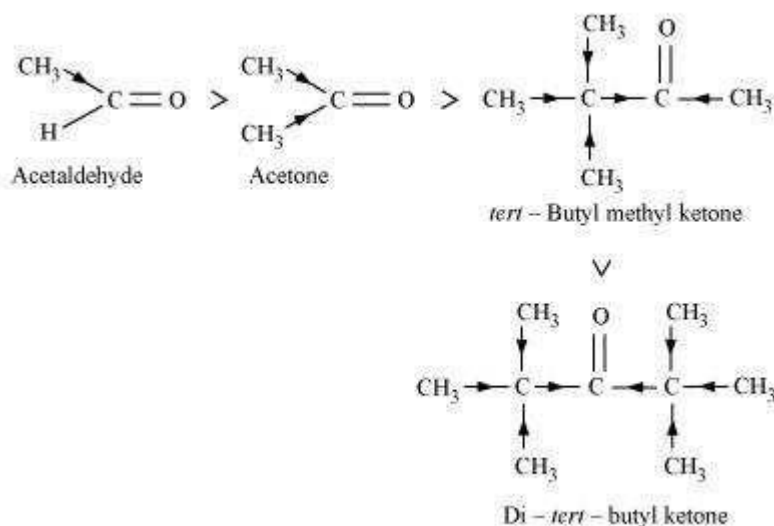
Arrange the following compounds in increasing order of their property as indicated: **(i)** Acetaldehyde, Acetone, Di-*tert*-butyl ketone, Methyl *tert*-butyl ketone (reactivity towards HCN)

**(ii)**  $\text{CH}_3\text{CH}_2\text{CH}(\text{Br})\text{COOH}$ ,  $\text{CH}_3\text{CH}(\text{Br})\text{CH}_2\text{COOH}$ ,  $(\text{CH}_3)_2\text{CHCOOH}$ ,  $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH}$  (acid strength)

**(iii)** Benzoic acid, 4-Nitrobenzoic acid, 3,4-Dinitrobenzoic acid, 4-Methoxybenzoic acid (acid strength)

Answer

**(i)** When HCN reacts with a compound, the attacking species is a nucleophile,  $\text{CN}^-$ . Therefore, as the negative charge on the compound increases, its reactivity with HCN decreases. In the given compounds, the +I effect increases as shown below. It can be observed that steric hindrance also increases in the same



Hence, the given compounds can be arranged according to their increasing reactivities toward HCN as:

Di-*tert*-butyl ketone < Methyl *tert*-butyl ketone < Acetone < Acetaldehyde

**(ii)** After losing a proton, carboxylic acids gain a negative charge as shown:



Now, any group that will help stabilise the negative charge will increase the stability of the carboxyl ion and as a result, will increase the strength of the acid. Thus, groups having +I effect will decrease the strength of the acids and groups having -I effect will increase the strength of the acids. In the given compounds, -CH<sub>3</sub> group has +I effect and Br<sup>-</sup> group has -I effect. Thus, acids containing Br<sup>-</sup> are stronger.

Now, the +I effect of isopropyl group is more than that of *n*-propyl group. Hence, (CH<sub>3</sub>)<sub>2</sub>CHCOOH is a weaker acid than CH<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>COOH.

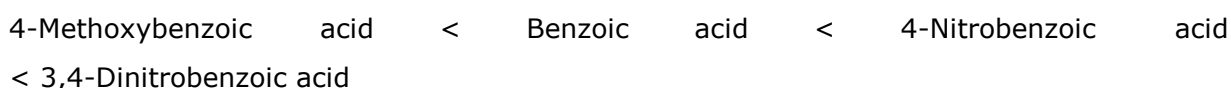
Also, the -I effect grows weaker as distance increases. Hence, CH<sub>3</sub>CH(Br)CH<sub>2</sub>COOH is a weaker acid than CH<sub>3</sub>CH<sub>2</sub>CH(Br)COOH.

Hence, the strengths of the given acids increase as:

(CH<sub>3</sub>)<sub>2</sub>CHCOOH < CH<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>COOH < CH<sub>3</sub>CH(Br)CH<sub>2</sub>COOH < CH<sub>3</sub>CH<sub>2</sub>CH(Br)COOH

**(iii)** As we have seen in the previous case, electron-donating groups decrease the strengths of acids, while electron-withdrawing groups increase the strengths of acids. As methoxy group is an electron-donating group, 4-methoxybenzoic acid is a weaker acid

than benzoic acid. Nitro group is an electron-withdrawing group and will increase the strengths of acids. As 3,4-dinitrobenzoic acid contains two nitro groups, it is a slightly stronger acid than 4-nitrobenzoic acid. Hence, the strengths of the given acids increase as:



### Question 12.13:

Give simple chemical tests to distinguish between the following pairs of compounds.

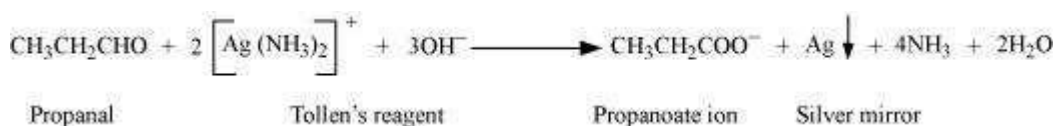
- (i) Propanal and Propanone
- (ii) Acetophenone and Benzophenone
- (iii) Phenol and Benzoic acid
- (iv) Benzoic acid and Ethyl benzoate
- (v) Pentan-2-one and Pentan-3-one
- (vi) Benzaldehyde and Acetophenone
- (vii) Ethanal and Propanal

Answer

(i) Propanal and propanone can be distinguished by the following tests.

#### (a) Tollen's test

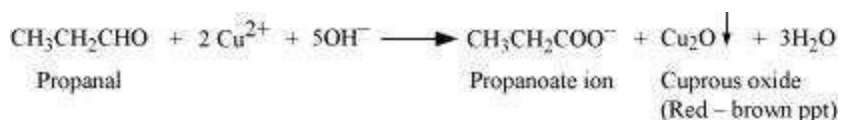
Propanal is an aldehyde. Thus, it reduces Tollen's reagent. But, propanone being a ketone does not reduce Tollen's reagent.



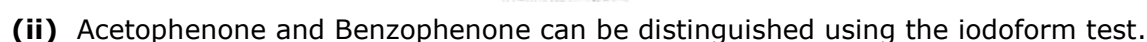
#### (b) Fehling's test

Aldehydes respond to Fehling's test, but ketones do not.

Propanal being an aldehyde reduces Fehling's solution to a red-brown precipitate of  $\text{Cu}_2\text{O}$ , but propanone being a ketone does not.



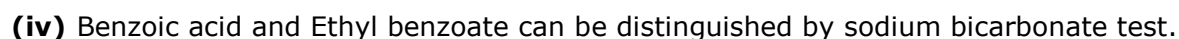
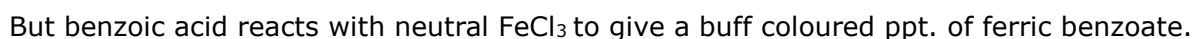
Aldehydes and ketones having at least one methyl group linked to the carbonyl carbon atom respond to iodoform test. They are oxidized by sodium hypoiodite (NaOI) to give iodoforms. Propanone being a methyl ketone responds to this test, but propanal does not.



Methyl ketones are oxidized by sodium hypoiodite to give yellow ppt. of iodoform.

**(iii)** Phenol and benzoic acid can be distinguished by ferric chloride test.

Phenol reacts with neutral  $\text{FeCl}_3$  to form an iron-phenol complex giving violet colouration.



Acids react with  $\text{NaHCO}_3$  to produce brisk effervescence due to the evolution of  $\text{CO}_2$  gas.

$$\text{C}_6\text{H}_5\text{COOC}_2\text{H}_5 + \text{NaHCO}_3 \longrightarrow \text{No effervescence due to evolution of CO}_2 \text{ gas}$$

**Iodoform test:**

$$\begin{array}{l} \text{CH}_3\text{CH}_2\text{CH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{CH}_3 + 3 \text{ NaOI} \longrightarrow \text{CH}_3\text{CH}_2\text{CH}_2\text{COONa} + \text{CHI}_3\downarrow + 2\text{NaOH} \\ \text{Pentan-2-one} \qquad \qquad \qquad \text{Sodium butanoate} \qquad \text{Iodoform} \\ \qquad \qquad \qquad \qquad \qquad \qquad \qquad \qquad \qquad \qquad \qquad \qquad \qquad \text{(yellow ppt)} \\ \\ \text{CH}_3\text{CH}_2-\overset{\text{O}}{\parallel}{\text{C}}-\text{CH}_2\text{CH}_3 + \text{NaOI} \longrightarrow \text{No yellow ppt of iodoform} \\ \text{Pentan-3-one} \end{array}$$

### (a) Tollen's Test

$$\text{C}_6\text{H}_5\text{CHO} + 2[\text{Ag}(\text{NH}_3)_2]^+ + 3\text{OH}^- \longrightarrow \text{C}_6\text{H}_5\text{COO}^- + \text{Ag} \downarrow + 4\text{NH}_3 + 2\text{H}_2\text{O}$$

Benzaldehyde                  Tollen's reagent                  Benzoate ion    Silver mirror

[illegible]

### Iodoform test

$$\text{CH}_3\text{CHO} + 3\text{NaOI} \longrightarrow \text{HCOONa} + \text{CHI}_3 + 2\text{NaOH}$$

Ethanal    Sodium      Iodoform  
methanoate (yellow ppt)

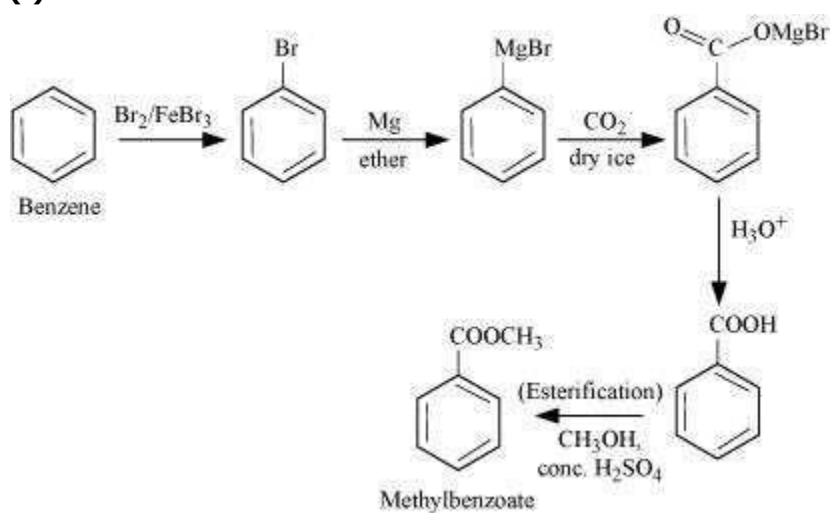
**Question 12.14:**

How will you prepare the following compounds from benzene? You may use any inorganic reagent and any organic reagent having not more than one carbon atom

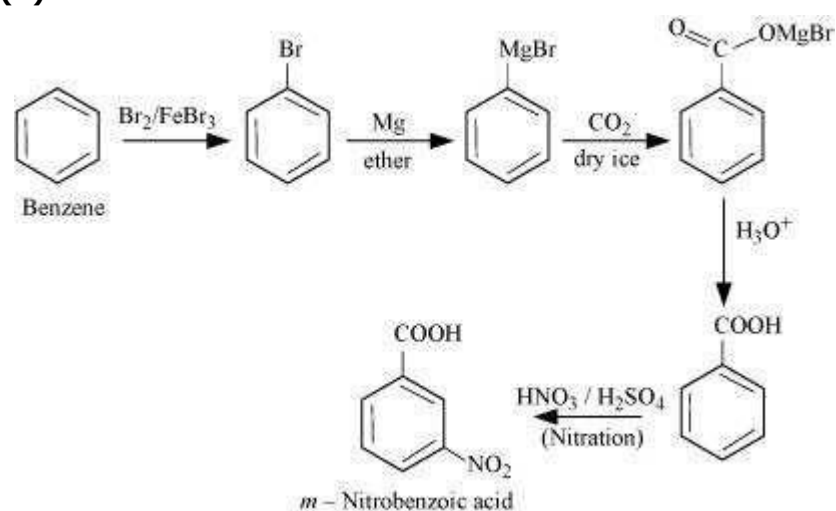
- (i) Methyl benzoate (ii) *m*-Nitrobenzoic acid  
(iii) *p*-Nitrobenzoic acid (iv) Phenylacetic acid  
(v) *p*-Nitrobenzaldehyde.

Answer

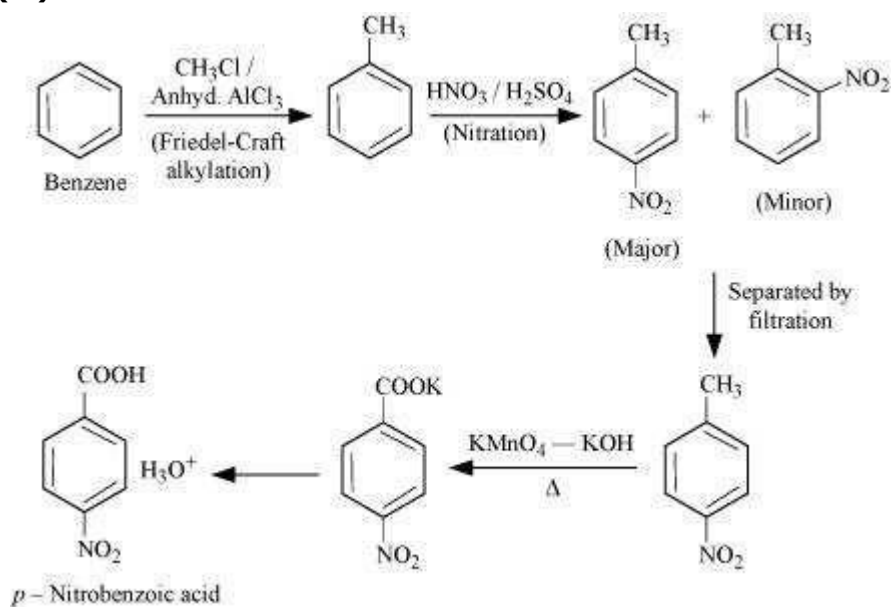
(i)



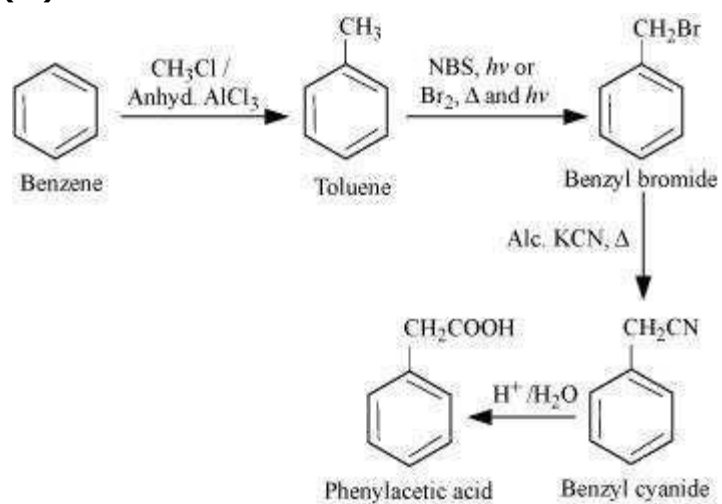
(ii)



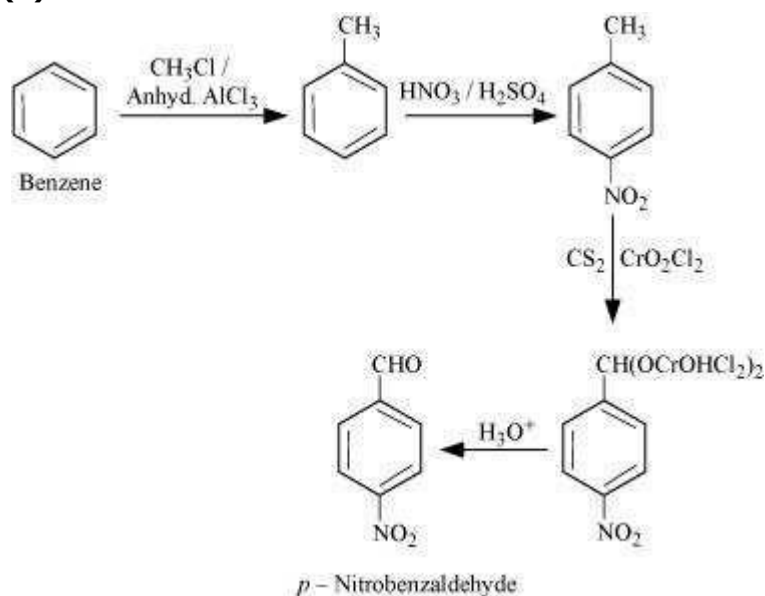
(iii)



(iv)



(v)



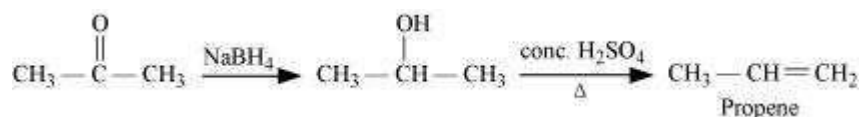
**Question 12.15:**

How will you bring about the following conversions in not more than two steps?

- (i) Propanone to Propene
- (ii) Benzoic acid to Benzaldehyde
- (iii) Ethanol to 3-Hydroxybutanal
- (iv) Benzene to *m*-Nitroacetophenone
- (v) Benzaldehyde to Benzophenone
- (vi) Bromobenzene to 1-Phenylethanol
- (vii) Benzaldehyde to 3-Phenylpropan-1-ol
- (viii) Benzaldehyde to  $\alpha$ -Hydroxyphenylacetic acid
- (ix) Benzoic acid to *m*- Nitrobenzyl alcohol

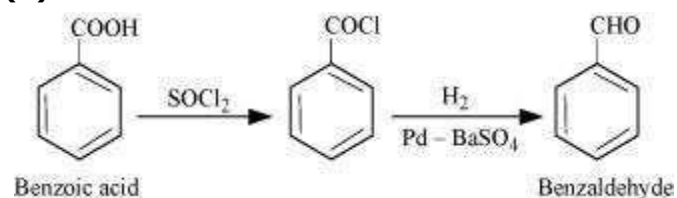
Answer

(i)

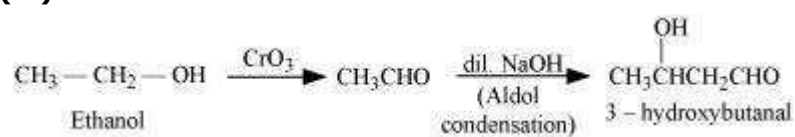




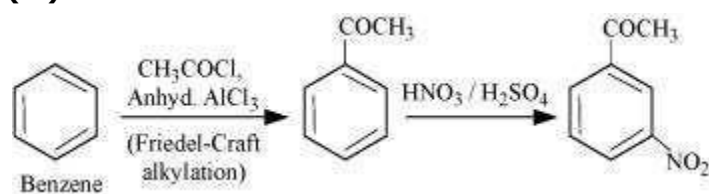
(ii)



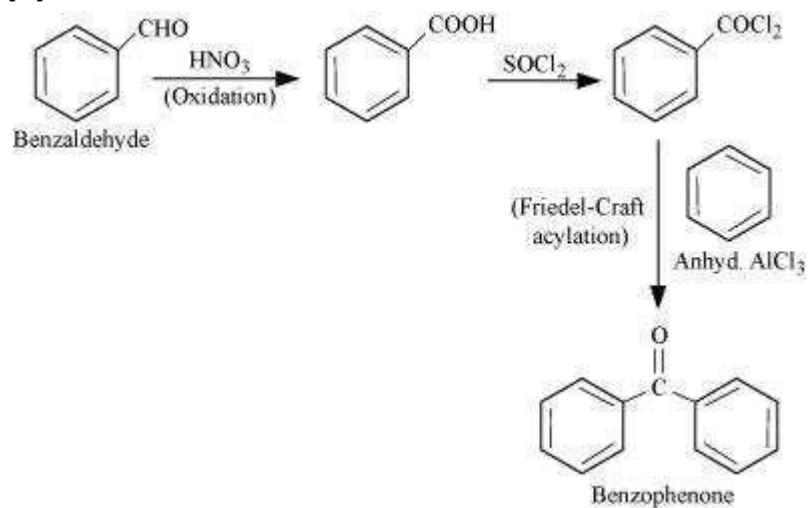
(iii)



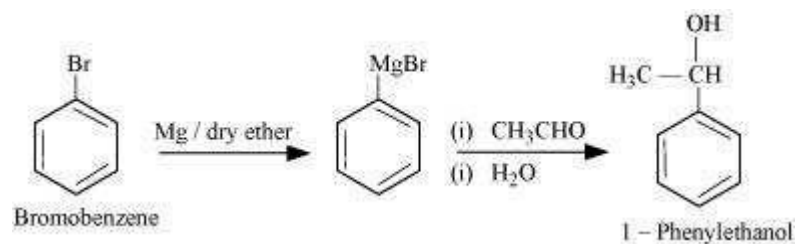
(iv)



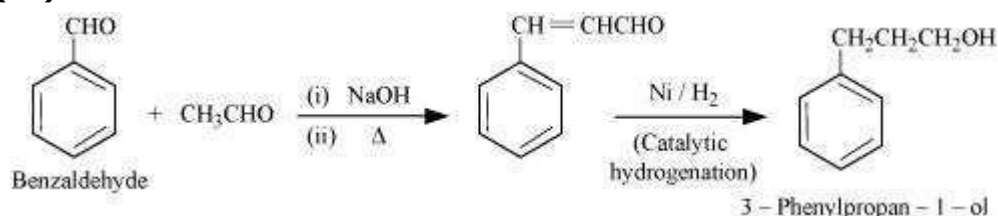
(v)



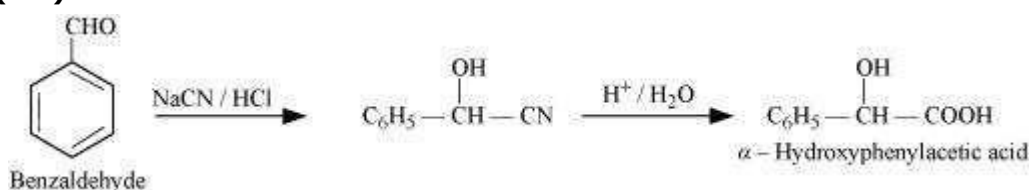
(vi)



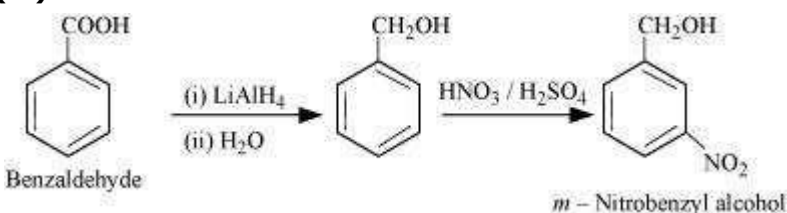
(vii)



(viii)



(ix)



**Question 12.16:**

Describe the following:

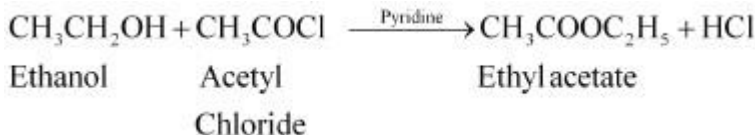
- (i) Acetylation
- (ii) Cannizzaro reaction
- (iii) Cross aldol condensation
- (iv) Decarboxylation

Answer

**(i) Acetylation**

The introduction of an acetyl functional group into an organic compound is known as acetylation. It is usually carried out in the presence of a base such as pyridine, dimethylaniline, etc. This process involves the substitution of an acetyl group for an active hydrogen atom. Acetyl chloride and acetic anhydride are commonly used as acetylating agents.

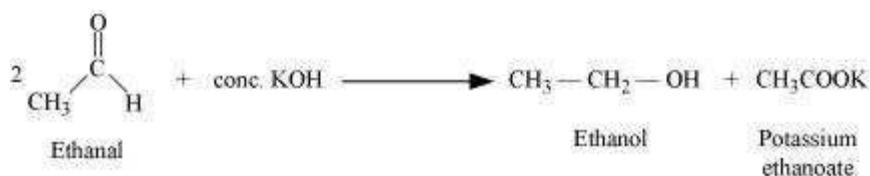
For example, acetylation of ethanol produces ethyl acetate.



**(ii) Cannizzaro reaction:**

The self oxidation-reduction (disproportionation) reaction of aldehydes having no  $\alpha$ -hydrogens on treatment with concentrated alkalis is known as the Cannizzaro reaction. In this reaction, two molecules of aldehydes participate where one is reduced to alcohol and the other is oxidized to carboxylic acid.

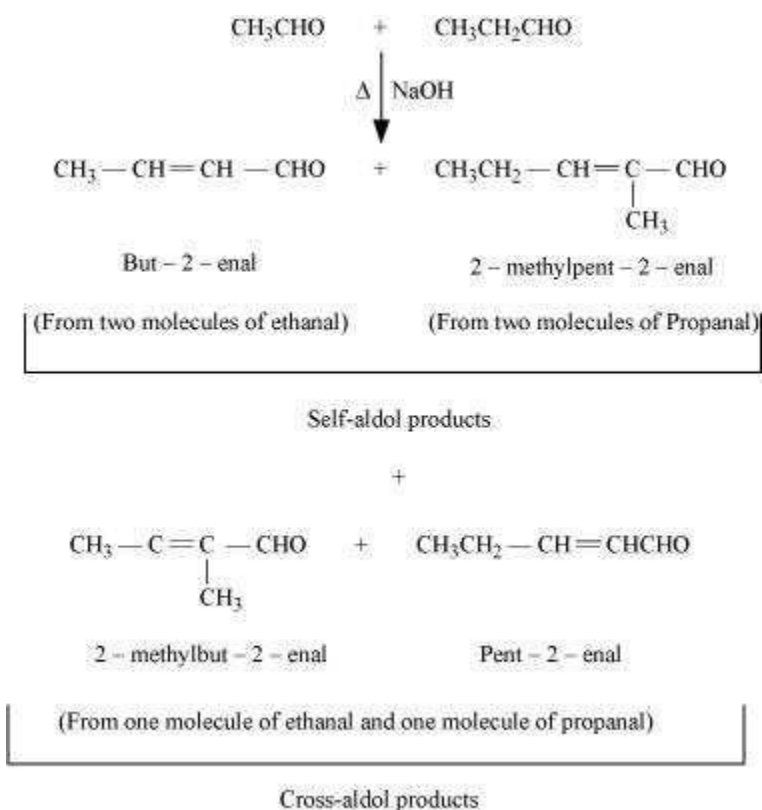
For example, when ethanol is treated with concentrated potassium hydroxide, ethanol and potassium ethanoate are produced.



**(iii) Cross-aldol condensation:**

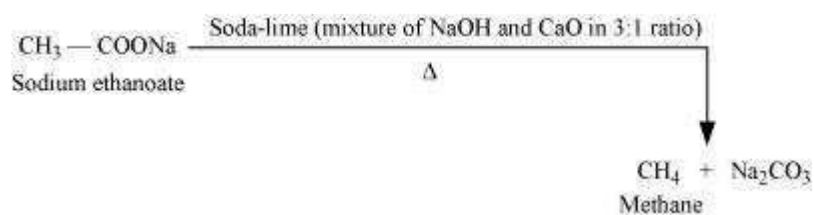
When aldol condensation is carried out between two different aldehydes, or two different ketones, or an aldehyde and a ketone, then the reaction is called a cross-aldol condensation. If both the reactants contain  $\alpha$ -hydrogens, four compounds are obtained as products.

For example, ethanal and propanal react to give four products.



#### (iv) Decarboxylation:

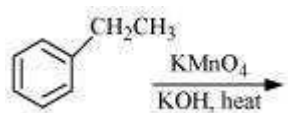
Decarboxylation refers to the reaction in which carboxylic acids lose carbon dioxide to form hydrocarbons when their sodium salts are heated with soda-lime.



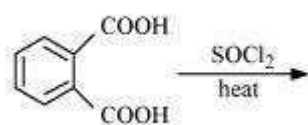
Decarboxylation also takes place when aqueous solutions of alkali metal salts of carboxylic acids are electrolyzed. This electrolytic process is known as Kolbe's electrolysis.

#### Question 12.17:

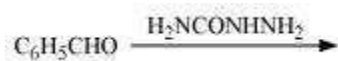
Complete each synthesis by giving missing starting material, reagent or products **(i)**



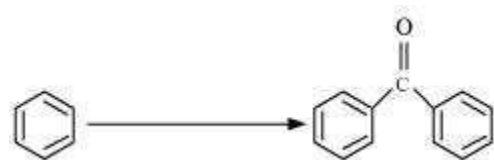
(ii)



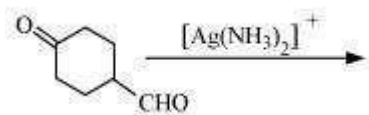
(iii)



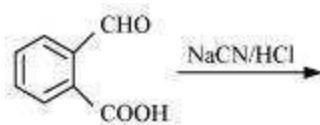
(iv)



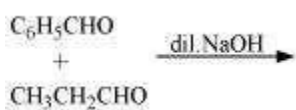
(v)



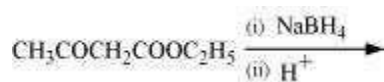
(vi)



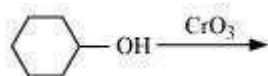
(vii)



(viii)



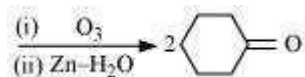
(ix)



(x)

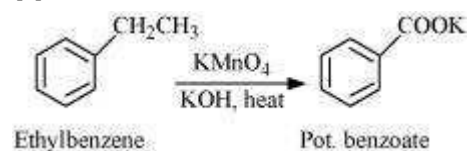


(xi)

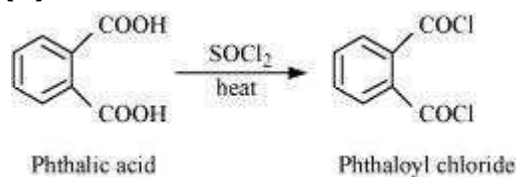


Answer

(i)



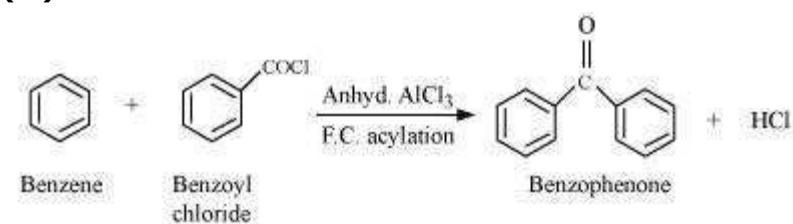
(ii)



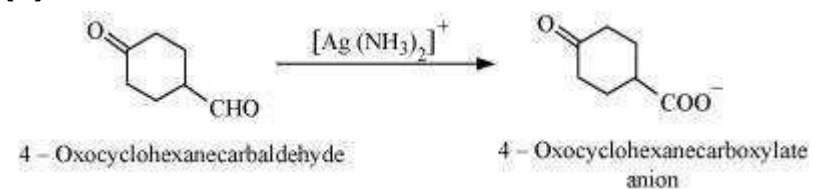
(iii)



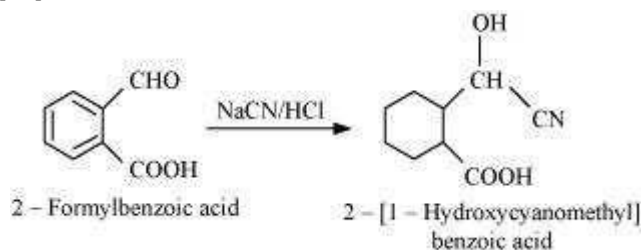
(iv)



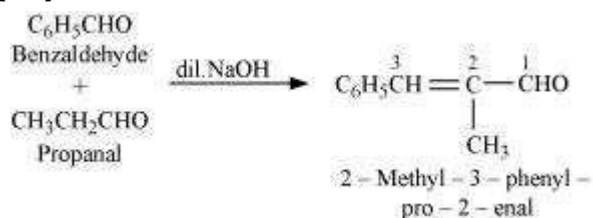
(v)



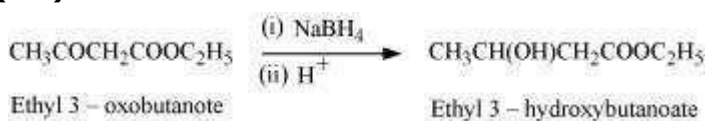
(vi)



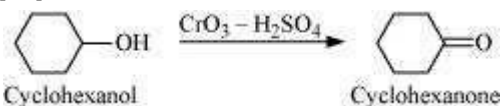
(vii)



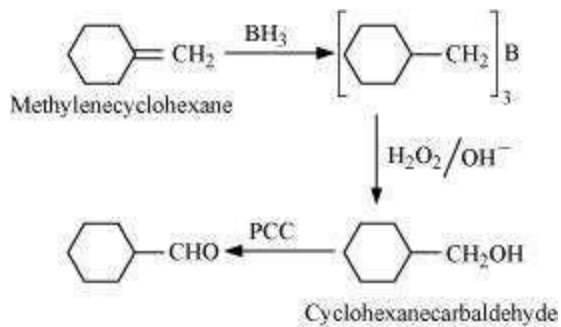
(viii)



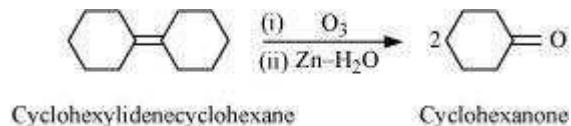
(ix)



(x)



(xi)



Question 12.18:

Give plausible explanation for each of the following:

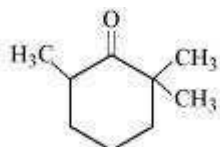
- (i) Cyclohexanone forms cyanohydrin in good yield but 2, 2, 6 trimethylcyclohexanone does not.
- (ii) There are two  $\text{-NH}_2$  groups in semicarbazide. However, only one is involved in the formation of semicarbazones.
- (iii) During the preparation of esters from a carboxylic acid and an alcohol in the presence of an acid catalyst, the water or the ester should be removed as soon as it is formed.

Answer

- (i) Cyclohexanones form cyanohydrins according to the following equation.



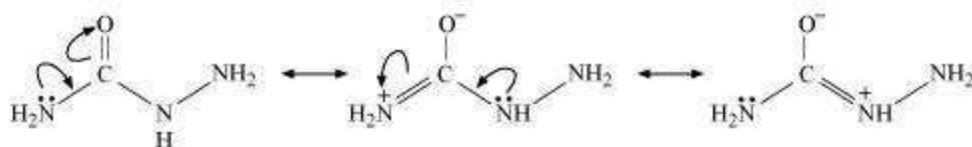
In this case, the nucleophile  $\text{CN}^-$  can easily attack without any steric hindrance. However, in the case of 2, 2, 6 trimethylcyclohexanone, methyl groups at  $\alpha$ -positions offer steric hindrances and as a result,  $\text{CN}^-$  cannot attack effectively.



2, 2, 6 – Trimethylcyclohexanone

For this reason, it does not form a cyanohydrin.

- (ii) Semicarbazide undergoes resonance involving only one of the two  $\text{-NH}_2$  groups, which is attached directly to the carbonyl-carbon atom.

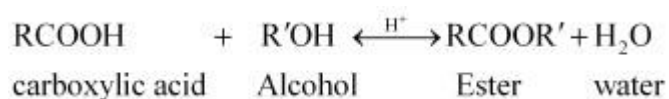


Therefore, the electron density on  $\text{-NH}_2$  group involved in the resonance also decreases. As a result, it cannot act as a nucleophile. Since the other  $\text{-NH}_2$  group is not involved in



resonance; it can act as a nucleophile and can attack carbonyl-carbon atoms of aldehydes and ketones to produce semicarbazones.

**(iii)** Ester along with water is formed reversibly from a carboxylic acid and an alcohol in presence of an acid.



If either water or ester is not removed as soon as it is formed, then it reacts to give back the reactants as the reaction is reversible. Therefore, to shift the equilibrium in the forward direction i.e., to produce more ester, either of the two should be removed.

#### Question 12.19:

An organic compound contains 69.77% carbon, 11.63% hydrogen and rest oxygen. The molecular mass of the compound is 86. It does not reduce Tollens' reagent but forms an addition compound with sodium hydrogensulphite and give positive iodoform test. On vigorous oxidation it gives ethanoic and propanoic acid. Write the possible structure of the compound.

Answer

% of carbon = 69.77 %

% of hydrogen = 11.63 %

% of oxygen = {100 – (69.77 + 11.63)}%  
= 18.6 %

Thus, the ratio of the number of carbon, hydrogen, and oxygen atoms in the organic compound can be given as:

$$\begin{aligned} \text{C:H:O} &= \frac{69.77}{12} : \frac{11.63}{1} : \frac{18.6}{16} \\ &= 5.81 : 11.63 : 1.16 \\ &= 5 : 10 : 1 \end{aligned}$$

Therefore, the empirical formula of the compound is C<sub>5</sub>H<sub>10</sub>O. Now, the empirical formula mass of the compound can be given as:

$$5 \times 12 + 10 \times 1 + 1 \times 16$$

= 86

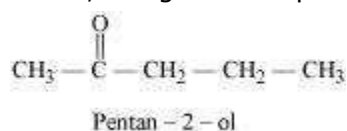
Molecular mass of the compound = 86

Therefore, the molecular formula of the compound is given by  $C_5H_{10}O$ .

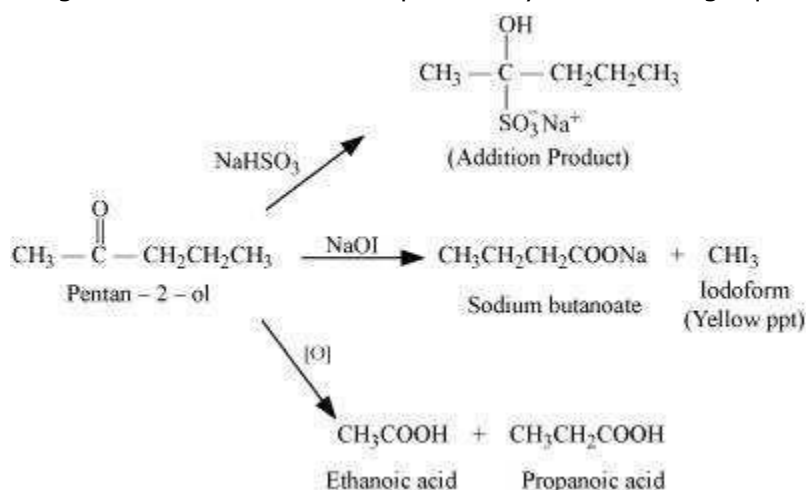
Since the given compound does not reduce Tollen's reagent, it is not an aldehyde. Again, the compound forms sodium hydrogen sulphate addition products and gives a positive iodoform test. Since the compound is not an aldehyde, it must be a methyl ketone.

The given compound also gives a mixture of ethanoic acid and propanoic acid.

Hence, the given compound is pentan-2-one.



The given reactions can be explained by the following equations:

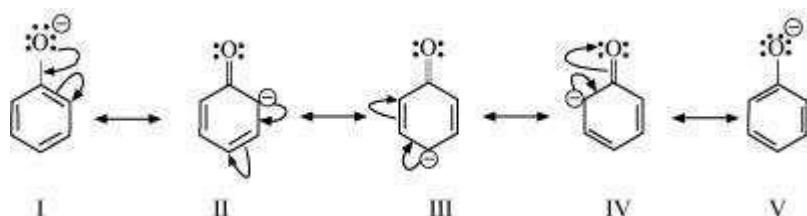


#### Question 12.20:

Although phenoxide ion has more number of resonating structures than carboxylate ion, carboxylic acid is a stronger acid than phenol. Why?

Answer

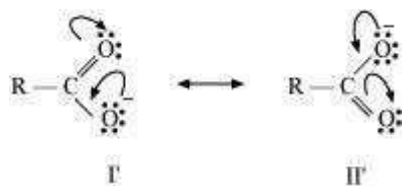
Resonance structures of phenoxide ion are:



It can be observed from the resonance structures of phenoxide ion that in II, III and IV, less electronegative carbon atoms carry a negative charge. Therefore, these three structures contribute negligibly towards the resonance stability of the phenoxide ion.

Hence, these structures can be eliminated. Only structures I and V carry a negative charge on the more electronegative oxygen atom.

Resonance structures of carboxylate ion are:



In the case of carboxylate ion, resonating structures I' and II' contain a charge carried by a more electronegative oxygen atom.

Further, in resonating structures I' and II', the negative charge is delocalized over two oxygen atoms. But in resonating structures I and V of the phenoxide ion, the negative charge is localized on the same oxygen atom. Therefore, the resonating structures of carboxylate ion contribute more towards its stability than those of phenoxide ion. As a result, carboxylate ion is more resonance-stabilized than phenoxide ion. Hence, carboxylic acid is a stronger acid than phenol.