

Lecture Notes on Atmospheric Thermodynamics and Cloud Physics

Harshit Prahant Dhanwalkar (SC21B164)^{1*}

Abstract

Notes of Lectures plus addional information from books.

¹MTech, Earth Sciences System (ESS), 1st year, Department of Physics, Indian Institute Of Spacesscience and Technology (IIST)

*email: harshitpd1729@gamil.com

Contents

1 Lecture 1 09/08/2024	3	8 Lecture 8 30/08/2024	17
1.1 Equilibrium State	3	Pressure profiles in the idealized atmosphere	17
1.2 Transformation	3	8.1 Constant density atmosphere	17
1.3 Exact and Inexact differentiable	3	8.2 Constant temperature atmosphere	18
1.4 Extensive and Intrincive variables	3	8.3 Constant lapse rate atmosphere	18
1.5 Laws of Thermodynamics	4	9 Lecture 9 04/09/2024	19
2 Lecture 2 14/08/2024	5	9.1 1 st law of thermodynamics and it's application	19
2.1 Kinetic molecular theory of gases	5	10 Lecture 10 05/09/2024	20
2.2 Ideal gas	5	10.1. 1 st law of thermodynamics and it's application	20
2.3 Early experiments and laws	5	10.2 Heat capacity	20
3 Lecture 3 16/08/2024	7	10.3 Heat capacity at constant volume	21
3.1 Another form of Gay-Lussac's Law	7	11 Lecture 11 06/09/2024	22
3.2 Boyle's Law	7	11.1. Specific heat capacity	22
3.3 Avagadro's Law	7	11.2. Specific heat at constant volume	22
3.4 Ideal gas Law	7	11.3. Specific heat at constant pressure	22
3.5 Van Der Waal's equation	8	11.4. Special forms of 1 st laaw of thermodynamics	22
3.6 Meteorological form of Ideal gas	8	12 Lecture 12 11/09/2024	24
3.7 Composition of Earth's Atmosphere	8	12.1. Poisson's equation for adiabatic transformation	24
4 Lecture 4 19/08/2024	9	13 Lecture 13 12/09/2024	26
4.1 Pressure	9	13.1. Adiabatic transformation - Poisson's equation	26
4.2 Mass of the Atmosphere	9	13.2 Adiabatic Lapse Rate	26
5 Lecture 5 21/08/2024	11	13.3. Dry Adiabatic Lapse Rate (DALR)	26
5.1 Dalton's law of partial pressure	11	13.4. Possion's equations	26
5.2 Humidity	11	14 Lecture 14 20/09/2024	28
5.3 Ideal gas equation for moist gas	11	14.1. Heat Engines	28
6 Lecture 6 28/08/2024	13	14.2 Efficiency	28
6.1 Archimedes Principle	13	14.3. Carnot Cycle	28
6.2 Buoyancy	13	14.4. Skew T - log P diagram	29
6.3 Hydrostatic equation	14	15 Lecture 15 03/10/2024	33
7 Lecture 7 29/08/2024	15	15.1. Cyclones	33
7.1 Ideal gas equation with Hydostatic equation	15	15.2 Specific enthaphy	33
7.2 Geopotential	15		

16 Lecture 16 07/10/2024	35
16.1. Lesson learnt from carnot cycle	35
16.2 Entropy	35
17 Lecture 17 10/10/2024	36
17.1. Entropy	36
17.2 <i>2nd</i> law of themodynamics	36
18 Lecture 18 10/10/2024	38
18.1. Special Forms of <i>2rd</i> law of Thermodynamics	38
18.2 Moist Processes	38
18.3. Relative humidity	38
18.4. Dew point temperature	39
18.5. Latent heat	39

1. Lecture 1 09/08/2024

- System: Specific chunk of matter we are interested in.
- Surrounding: Everything else in the universe outside of the system.
- Air parcel: System with the following assumptions:
 - Properties within the air parcel are uniform.
 - Boundaries are closed, so that matter/mass is not exchanged with the surroundings.

Question 1.1: How do system interact with surrounding?

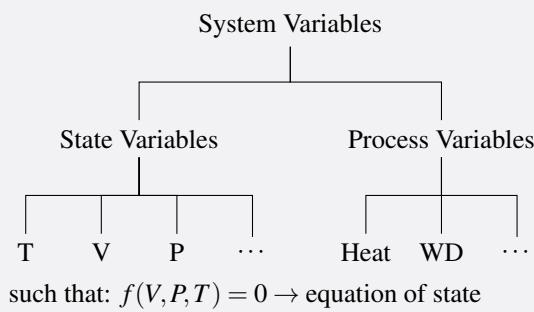
- Matter
- Radiation

Question 1.2: Classify system on the basis of how matter interact with surround?

⇒ We can classify system into 3 types on the basis of how matter in system interact with it's surrounding:

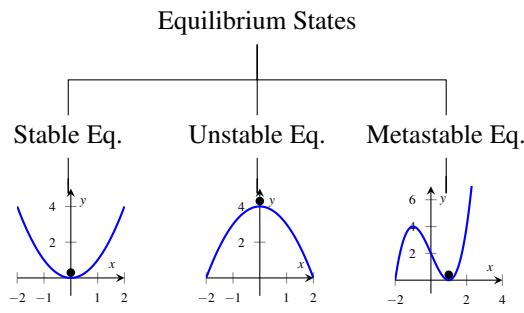
- Open
- Closed
- Isolated

Question 1.3: How do we characterize a system?



1.1 Equilibrium State

Equilibrium: No change in the system if the surrounding doesn't change.



1.2 Transformation

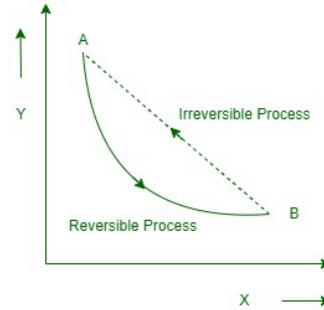
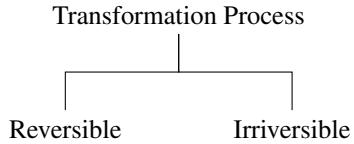


Figure 1. Reversible and Irreversible Process

1.3 Exact and Inexact differentiable

Let z be a function of x and y , exact differentiable:

$$dz = \left(\frac{\partial z}{\partial x} \right) dx + \left(\frac{\partial z}{\partial y} \right) dy \quad (1)$$

$$dz = \int_i^f \delta z = z(x_f, y_f) - z(x_i, y_i) \quad (2)$$

∴ dT, dP, dV are exact differentiable.

$\delta Q, \delta W$ are inexact differentiable.

line integral:

$$\oint \delta z = 0$$

→ z is stable variable and it is exact differentiable iff it's reversible

$$\therefore \oint dT = 0, \oint dP = 0, \oint dq \neq 0, \oint dw \neq 0$$

1.4 Extensive and Intrincive variables

- **Extensive:** Depends on size of system. E.g. Volume.
- **Intrincive:** Independent of size of system. E.g. Temperature.

Any variable divided by mass gives intrincive.

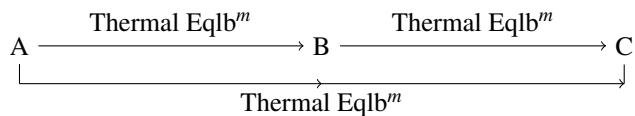
For e.g. specific volume (α) = $\frac{\text{volume}}{\text{mass}}$.
Similarly,

$$p = \frac{P}{m}, q = \frac{Q}{m}, w = \frac{W}{m}$$

1.5 Laws of Thermodynamics

1. 0th law of thermodynamics:

Temperature → Quantity that determines the direction of heat flow. If two objects are in thermal contact and there is no net heat transfer, then the system is said to be in **thermal equilibrium**.



2. Lecture 2 14/08/2024

Bulk properties → microscopic properties → which can be linked to microscopic properties

For e.g. P is exerted due to random motion and collision of molecules with each other and on the walls of container.

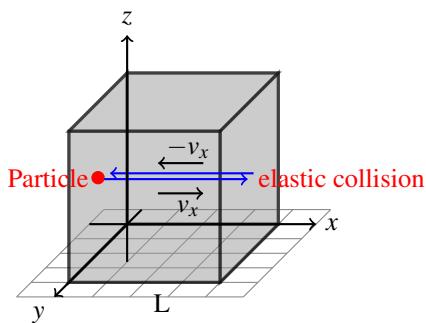
2.1 Kinetic molecular theory of gases

Conditions:

- Molecules are in **random motion**.
- Collisions between molecules with each other and wall of container are **elastic** in nature ensuring **no K.E. loss**.

Derivation of av. K.E.

Particles travel distance equal to length of box (L) with velocity V_x, V_y and V_z .



Force exerted by the molecules on the face:

$$F = \frac{\Delta P}{\Delta t} = \frac{\Delta mv}{\Delta t} = m \frac{\Delta v}{\Delta t} \quad (3)$$

Change in velocity

$$\Delta v_x = v_x - (-v_x) = 2v_x \quad (4)$$

$$\Delta v_y = 0 \quad (5)$$

$$\Delta v_z = 0 \quad (6)$$

$$\implies \Delta V = 2v_x \quad (7)$$

$$\therefore F = \frac{m(2v_x)}{\Delta t} \quad (8)$$

We also have,

$$\Delta t = \frac{2L}{v_x} \quad (9)$$

from Eq. (8) and Eq. (9), we get:

$$F = \frac{m(2v_x^2)}{2L} = \frac{mv_x^2}{L} \quad (10)$$

For N number of molecules and average velocity of all molecules moving in x -direction \bar{v}_x

$$\therefore F = \frac{Nm(\bar{v}_x^2)}{L} \quad (11)$$

Pressure (P) exerted on walls of container:

$$P = \frac{F}{A} = \frac{Nm(\bar{v}_x^2)}{L \times L^2} = \frac{Nm(\bar{v}_x^2)}{V} \quad (12)$$

$$\implies PV = Nm(\bar{v}_x^2) \quad (13)$$

When we consider the velocity of molecules in all directions (v_{tot}).

$$v_{tot}^2 = \bar{v}_x^2 + \bar{v}_y^2 + \bar{v}_z^2 \quad (14)$$

$$\implies \bar{v}_{tot}^2 = 3\bar{v}_x^2 \quad (15)$$

\therefore The Pressure(P) becomes:

$$PV = \frac{1}{3}Nm\bar{v}_{tot}^2 \quad (16)$$

$$3PV = Nm\bar{v}_{tot}^2 \quad (17)$$

$$\frac{3}{2}PV = \frac{1}{2}Nm\bar{v}_{tot}^2 \quad (18)$$

$$\frac{3}{2}PV = N \times \left(\frac{1}{2}m\bar{v}_{tot}^2 \right) \quad (19)$$

$$\implies \frac{3}{2}PV = N \times (K.E.)_{av} \quad (20)$$

$$(21)$$

Using Ideal gas equation : $PV = Nk_B T$, where k_B is boltzmann constant

$$\frac{3}{2}(Nk_B T) = N \times (K.E.)_{av} \quad (22)$$

$$\implies (K.E.)_{av} = \frac{3}{2}k_B T \quad (23)$$

2.2 Ideal gas

- Molecules are in random motion.
- During the motion of **molecules do not exert force, except when they collide with each other or the walls of container**. This can also be stated as there is **no force of attraction between molecules**.
- The collisions between molecules are **elastic**.
- Sum of the volume of molecules is negligible** compared to volume of container.

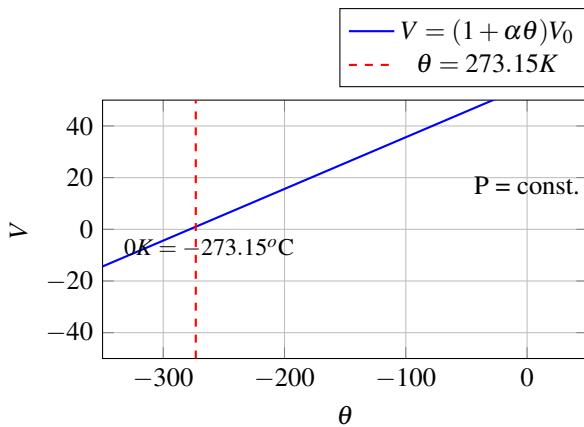
2.3 Early experiments and laws

1st Law of Gay-Lussac

Increase of volume of an ideal gas at **constant pressure** is proportional to increase in temperature and also to the volume occupied by the gas at 0°C

$$dV \propto V_0 d\theta \quad (24)$$

$$dV = \alpha V_0 d\theta \quad (25)$$

**Figure 2.** Plot of V vs. θ

where α is volume coefficient of thermal expansion.

$$\alpha = \frac{1}{d\theta} \frac{dV}{V_0} = \frac{1}{273} \quad (26)$$

$$\int_{V_0}^V dV = \int_{0^\circ C}^{\theta} \alpha V_0 d\theta \quad (27)$$

$$V - V_0 = \alpha V_0 (\theta - 0^\circ C) \quad (28)$$

$$\Rightarrow V = (1 + \alpha\theta)V_0 \quad (29)$$

2nd Law of Gay-Lussac

An Ideal gas kept at constant volume, then the increase in pressure is proportional to increase in temperature and pressure at $0^\circ C$.

$$dP \propto P_0 d\theta \quad (30)$$

$$dP = \beta P_0 d\theta \quad (31)$$

where β is pressure coefficient of thermal expansion.

$$\beta = \frac{1}{d\theta} \frac{dP}{P_0} = \frac{1}{273} \quad (32)$$

$$\int_{P_0}^P dP = \int_{0^\circ C}^{\theta} \beta P_0 d\theta \quad (33)$$

$$P - P_0 = \beta P_0 (\theta - 0^\circ C) \quad (34)$$

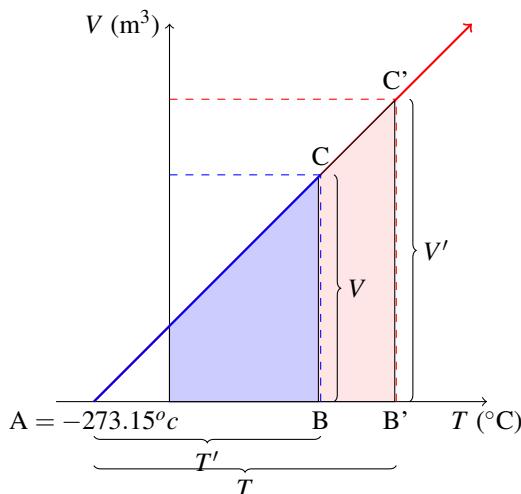
$$\Rightarrow P = (1 + \beta\theta)P_0 \quad (35)$$

3. Lecture 3 16/08/2024

Question 3.1: Why do we need Kinetic theory of gases?

⇒ Kinetic theory of gases connects microscopic properties to macroscopic properties of gases.

3.1 Another form of Gay-Lussac's Law

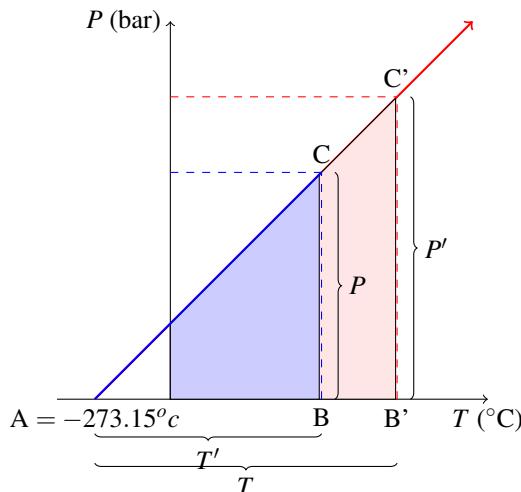


From similarity of triangle;

$$\begin{aligned} \frac{BC}{AB} &= \frac{B'C'}{AB'} \\ \frac{V}{T} &= \frac{V'}{T'} \\ \frac{V}{V'} &= \frac{T}{T'} \end{aligned}$$

for P is Constant

Similarly for Pressure at constant Volume



From similarity of triangle;

$$\frac{BC}{AB} = \frac{B'C'}{AB'}$$

$$\frac{P}{T} = \frac{P'}{T'}$$

$$\frac{P}{P'} = \frac{T}{T'}$$

for V is Constant

3.2 Boyle's Law

Boyle's Law states that, at constant temperature, the pressure of a given amount of gas is inversely proportional to its volume. Mathematically, it is expressed as:

$$P \propto \frac{1}{V} \quad (36)$$

$$PV = P'V' \quad (37)$$

$$PV = \text{const.} \quad (38)$$

where P is the pressure of the gas, and V is its volume. This implies that if the volume of a gas increases, its pressure decreases, and vice versa, as long as the temperature and the amount of gas remain constant.

3.3 Avagadro's Law

Avogadro's Law states that, at the same temperature and pressure, equal volumes of all gases contain the same number of molecules. Mathematically, it is expressed as:

$$V \propto n \quad (39)$$

where V is the volume of the gas, and n is the number of moles of the gas. This implies that the volume of a gas is directly proportional to the number of moles, provided temperature and pressure are constant.

For one mole of gas contains $6.023 \times 10^{23} \text{ mol}^{-1}$ of particles.

3.4 Ideal gas Law

System defined by (P, V, T) undergo change shown in following Fig.3

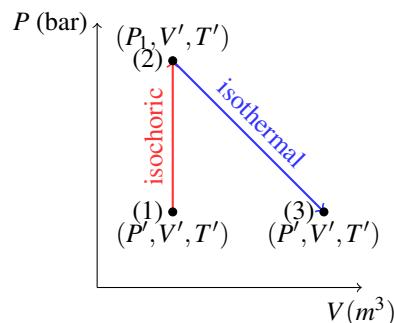


Figure 3. Diagram illustrating different thermodynamic processes in a P - V diagram.

Apply Gay-Lussac law (1) → (2):

$$P_1 = P \frac{T'}{T} \quad (40)$$

Apply Boyle's law (2) → (3):

$$P'V' = P_1 V$$

from eq.40 we get

$$P'V' = \left(\frac{PT'}{T} \right) V$$

after rearranging we get:

$$\frac{P'V'}{T'} = \frac{PV}{T}$$

$$\frac{PV}{T} = \text{const}$$

$$PV = AT$$

where $A = nR^*$

R^* is universal gas constant ($8.314 \text{ J K}^{-1} \text{ mol}^{-1}$)

$$\Rightarrow PV = nR^*T \quad (41)$$

3.5 Van Der Waal's equation

$$\left[P + a \left(\frac{n}{V} \right)^2 \right] \left[V - (nb) \right] = nR^*T \quad (42)$$

Where

a = coeff. which characterises the intermolecular forces = $1.35 \times 10^5 \text{ J m}^3 \text{ K}^{-1} \text{ mol}^{-2}$

b = coeff. which accounts for effective volume occupied by molecules = $3.64 \times 10^{-2} \text{ m}^3 \text{ K}^{-1} \text{ mol}^{-1}$

When $nb \rightarrow 0$ or $\frac{n}{V} \rightarrow 0$ Van Der Waal equation → Ideal gas equation.

3.6 Meteorological form of Ideal gas

Let there be n -kilomols particles/molecules of gas.

Therefore the combination of i^{th} component of the gas can be given by:

$$n = \sum_{i=1}^k n_i \quad (43)$$

The total mass of sample in kg:

$$M = \sum_{i=1}^k n_i m_i \quad (44)$$

where m_i represents molar mass of i^{th} particle/molecule in a sample.

Using Ideal gas equation:

$$PV = nR^*T \quad (45)$$

$$\frac{PV}{M} = \frac{nR^*T}{M} \quad (46)$$

$$P\alpha = \frac{n}{M} R^* T \quad (47)$$

since $\frac{V}{M} = \alpha$, and called specific volume.

$$P\alpha = \frac{\sum_{i=1}^k n_i}{\sum_{i=1}^k n_i m_i} R^* T \quad (48)$$

$$P\alpha = \frac{R^* T}{\bar{m}} \quad (49)$$

$$\Rightarrow P\alpha = R_d T \quad (50)$$

where \bar{m} is mean molar mass and given by $\frac{\sum_{i=1}^k n_i m_i}{\sum_{i=1}^k n_i}$ and $R_d = \frac{R^*}{\bar{m}}$ and unit $\text{JKg}^{-1} \text{K}^{-1}$

Since we know $\alpha = \frac{1}{\rho}$, where ρ is density of gas.

$$P\alpha = R_d T \quad (51)$$

$$\Rightarrow P = \rho R_d T \quad (52)$$

Question 3.2: what is R_d ?

$\Rightarrow R_d \rightarrow$ is a specific gas constant where d stands for dry air and this constant is not universal, varies with time and conditons over a particular place.

$$R_d = \frac{R^*}{\bar{m}}$$

and unit $\text{JKg}^{-1} \text{K}^{-1}$

3.7 Composition of Earth's Atmosphere

Gas	Fraction/Volume	Molecular Mass
N_2	78.1%	28.01
O_2	20.9%	31.999
Ar	0.93%	39.9

Table 1. Composition of Earth's Atmosphere.

We haven't take water vaours($H_2O(v)$), Carbon diaoxide(CO_2) and Ozone(O_3) becasue these gaes are highly variable w.r.t time and geography.

Question 3.3: Find R_d for Earth Atmosphere?

$$\begin{aligned} \Rightarrow R_d &= \frac{R^*}{\bar{m}} \\ &= \frac{8.314 \times 10^3}{28.96} \\ &= 0.287085 \\ &= 287.085 \text{ J kg}^{-1} \text{ K}^{-1} \end{aligned}$$

4. Lecture 4 19/08/2024

$CO_2 \rightarrow$ radiation trapping process \rightarrow absorb radiation in IR region of spectrum.

Question 4.1:

- Determine the gas constant for the atmosphere of Venus which consists of 95% CO_2 and 5% N_2 by volume.
- The mean surface temperature of Venus is 740K as compared to 288K surface temperature of Earth. The surface pressure on Venus is 90 times that on Earth. By what factor the density of near surface Venus atmosphere is greater than the Earth?

\Rightarrow Solution:

- Mean molar mass of gases of atmosphere of Venus:

$$\bar{m} = \frac{\sum_{i=1}^k n_i m_i}{\sum_{i=1}^k n_i}$$

$$\bar{m} = 0.95 \times 46 + 0.05 \times 28$$

$$\bar{m} = 43.2$$

Calculate Specific Molar const. (R_V) for Venus:

$$R_V = \frac{R^*}{\bar{m}}$$

$$\Rightarrow R_V \approx 192 \text{ } JK^{-1} \text{ } Kg^{-1}$$

- Surface Temperature of Venus = $T_V = 740K$
Surface Temperature of Earth = $T_E = 288K$
Surface Pressure of Earth = P_E
Surface Pressure of Venus = $P_V = 90 \times P_E$

$$P_V = \rho_V R_V T_V$$

$$P_E = \rho_E R_E T_E$$

$$\frac{P_V}{P_E} = \frac{\rho_V R_V T_V}{\rho_E R_E T_E}$$

$$90 = \frac{\rho_V}{\rho_E} * \frac{192}{287} * \frac{740}{288}$$

$$\frac{\rho_V}{\rho_E} = 52.36$$

Venus's atmosphere is 52.36 times denser than Earth's atmosphere.

$$\Rightarrow \rho_E = 1.23 \text{ } kgm^{-3} \text{ and } \rho_V = 65.97 \text{ } kgm^{-3}$$

Question 4.2: Why CO_2 have radiation trapping affinity but not gases like N_2 ?

$\Rightarrow CO_2$ has radiation trapping affinity because its molecular structure allows it to absorb and re-emit IR radiation, contributing to the greenhouse effect. This is due to its vibrational modes that change the molecule's dipole moment. In contrast, N_2 , with its symmetric diatomic structure, cannot absorb infrared radiation effectively, as its vibrations do not change the dipole moment, making it non-contributory to the greenhouse effect.

4.1 Pressure

Units of Pressure

Pressure are usually expressed in the following units:

$$1 \text{ bar} = 1.013 \times 10^5 \text{ Pa}$$

$$10^5 \text{ Pa} = 1000 \text{ hPa} = 1000 \text{ mbar}$$

Question 4.3: Why does ozone layer depletion happen primarily over the South Pole in Antarctica?

\Rightarrow Ozone layer depletion happens primarily over the South Pole because of the following reasons:

- Polar Stratospheric Clouds (PSCs):** During the Antarctic winter, temperatures drop below -75°C , leading to the formation of PSCs. These clouds facilitate chemical reactions that convert inactive chlorine compounds into reactive forms, which destroy ozone.
- Isolation of the Polar Vortex:** The strong polar vortex over Antarctica isolates air, keeping temperatures low and trapping ozone-depleting chemicals within the vortex.
- Sunlight and Ozone Destruction:** In Antarctic spring, returning sunlight provides energy for reactions between chlorine radicals and ozone, leading to significant ozone depletion and the formation of the "ozone hole."
- Comparison with the Arctic:** The Arctic has a weaker and less stable polar vortex, resulting in less dramatic ozone depletion compared to Antarctica.

4.2 Mass of the Atmosphere

At any point in the atmosphere, atmosphere above will exert a downward force due to gravitational force

$$F = \rho g \quad (53)$$

$$F = \int_0^{\infty} \rho g dz \quad (54)$$

We know $P = \text{Force}/(\text{unit area})$, assuming gravity g_0 remain constant, we get:

$$\Rightarrow P_s = g_0 \int_0^\infty pdz \quad (55)$$

where P_s is vertically integrated and have unit kgm^{-2}

Question 4.4: Globally average surface pressure of earth is 985hpa. Estimate mass of atmosphere.

\Rightarrow Given $985\text{hpa} = 985 \times 10^2 \text{Pa}$

surface area(sa) = $4\pi r^2$

where r is radius of earth = 6400km

we know, pressure(p) = $\frac{\text{force}}{\text{sa}}$

$$p = \frac{mg}{4\pi r^2}$$

$$986 \times 10^2 = \frac{m \times 9.81}{4 \times \pi \times (6400 \times 10^3)^2}$$

$$m = 5.168172908 \times 10^{18} \text{kg}$$

\therefore The approximate mass of atmosphere is equal to 5.1708×10^{15}

Question 4.5: The average atmospheric pressure on surface of Mars is 6hPa and raduis 3400km. Find mass of Mars.

\Rightarrow Given Surface pressure of Mars $P = 6\text{hPa} = 6 \times 10^2 \text{pa}$

Radius of Mars $R = 3400\text{km} = 3.4 \times 10^6 \text{m}$

surface area(sa) = $4\pi r^2$

where r is radius of earth.

we know, pressure(p) = $\frac{\text{force}}{\text{sa}}$

$$p = \frac{mg}{4\pi r^2}$$

$$986 \times 10^2 = \frac{m \times 9.81}{4 \times \pi \times (6400 \times 10^3)^2}$$

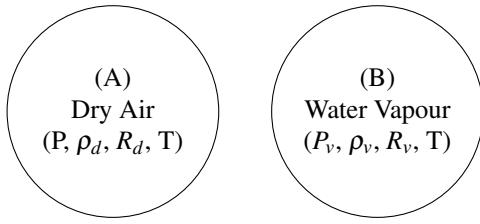
$$m = 5.168172908 \times 10^{18} \text{kg}$$

\therefore The approximate mass of atmosphere is equal to $5.168172908 \times 10^{18} \text{kg}$

5. Lecture 5 21/08/2024

In previous lecture, We got equation for Ideal gas Eq.(52), which only deals with dry air.

We didn't incorporate moisture!



where P_v is vapour pressure and given by

$$P_v = \rho_v R_v T \quad (56)$$

$$R_v = \frac{R^*}{\bar{m}} = \frac{8.314 \times 10^3}{18.01} = 461.63 \text{ JK}^{-1} \text{ kg}^{-1} \quad (57)$$

Note: Water vapour is not same as moisture, mixture is mixture of air and water vapour

5.1 Dalton's law of partial pressure

Dalton's law of partial pressure states that total pressure exerted by the mixture of gas is equal to the sum of partial pressure exerted by individual constituent at a given temperature.

$$P = P_d + e \quad (58)$$

Where

P = Total pressure exerted by all gases in mixture

P_d = Pressure exerted by dry air

$e = P_v$ = Vapour pressure

Substituting Eq.(52) and Eq.(56) in Eq.(58), we get:

$$P = \rho_d R_d T + \rho_v R_v T \quad (59)$$

$$P = (\rho_d R_d + \rho_v R_v) T \quad (60)$$

5.2 Humidity

We define humidity using following parameters:

1. Mixing ratio:

$$\omega = \frac{\text{Mass of water vapour}}{\text{Mass of dry air}} = \frac{M_v}{M_d} \quad (61)$$

$$= \frac{\text{Density of water vapour}}{\text{Density of dry air}} = \frac{\rho_v}{\rho_d} \quad (62)$$

Unit of mixing ratio is g/Kg

2. Specific heat:

$$q = \frac{\text{Mass of water vapour}}{\text{Mass of dry air} + \text{Mass of water vapour}} \quad (63)$$

$$= \frac{M_v}{M_d} \quad (64)$$

$$= \frac{\text{Density of water vapour}}{\text{Density of dry air} + \text{Density of water vapour}} \quad (65)$$

$$= \frac{\rho_v}{\rho_d + \rho_v} \quad (66)$$

$$= \frac{\rho_v}{\rho} \quad (67)$$

$w \approx q$, \because mass of water vapour \ll mass of dry air

$$\omega = \frac{\rho_v}{\rho_d} \quad (68)$$

$$= \frac{e/R_v T}{P_d/R_d T} \quad (69)$$

$$= \frac{e\varepsilon}{P_d} \quad (70)$$

$$= \frac{e\varepsilon}{P - e} \quad (71)$$

$$\approx \frac{e\varepsilon}{P} \quad (72)$$

where $\varepsilon = \frac{R_d}{R_v} = 0.621$
Similiary,

$$\begin{aligned} q &= \frac{\rho_v}{\rho_v + \rho_d} \\ &= \frac{e\varepsilon}{P - (1 - \varepsilon)e} \\ &\approx \frac{e\varepsilon}{P} \\ q &\approx w \end{aligned}$$

5.3 Ideal gas equation for moist gas

Total pressure

$$\begin{aligned} P &= P_d + e \\ &= \rho_d R_d T + \rho_v R_v T \\ &= \rho_d R_d T \left[1 + \frac{\rho_v R_v}{\rho_d R_d} \right] \\ &= \rho_d R_d T \left[1 + \frac{\rho_v}{\rho_d} \cdot \frac{R_v}{R_d} \right] \\ &= \rho_d R_d T \left[1 + \frac{\rho_v}{\rho_d} \cdot \frac{1}{\varepsilon} \right] \\ &= \rho_d R_d T \left[1 - \left(1 - \frac{1}{\varepsilon} \right) \frac{\rho_v}{\rho} \right] \end{aligned}$$

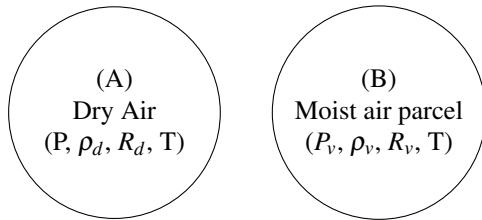
$$\therefore P = \rho_d R_d T \left[1 - \left(1 - \frac{1}{\varepsilon} \right) \cdot q \right] \quad (73)$$

Virtual Temperature

$$\Rightarrow T_v = \frac{P}{\rho_d R_d} = T \left[1 - \left(1 - \frac{1}{\varepsilon} \right) \cdot q \right] \quad (74)$$

Therefore from Eq.(73) and Eq.(74), we get:

$$P = \rho_d R_d T_v \quad (75)$$



$$\rho_d > \rho_m \xrightarrow{\Delta} \rho_d = \rho_m, \text{ where } \Delta \text{ represents heat.}$$

Question 5.1: On a summer day the AC breaks down and the air in the classroom becomes warm and muggy with a vapour pressure of 20 hPa and a temperature of 25°C.

- a. If the volume of the classroom is 40m³. How much water is present in the room in vapour form?
- b. If pressure of the room is 900hPa then what is virtual temerature of the air?

⇒ a. We have,

$$\begin{aligned} P_v &= \rho_v R_v T \\ 20 \times 10^2 &= \rho_v \times 461.62 \times 298 \\ \rho_v &= 0.0145 \text{ kg/m}^3 \\ \frac{m}{V} &= 0.0145 \\ \frac{m}{40} &= 0.0145 \\ m &\approx 0.58149 \text{ kg} \end{aligned}$$

∴ Amount of water vapour in room is 0.58148976 kg.

⇒ b. Given, $P = 900 \text{ hPa} = 90000 \text{ Pa}$

Let T_v be the virtual temperature. We know:

$$T_v = T \left(1 + 0.61 \frac{P_v}{P} \right)$$

Substituting the values:

$$T_v = 298 \left(1 + 0.61 \times \frac{2000}{90000} \right)$$

$$T_v = 298 \times 1.01356$$

$$T_v \approx 302.04 \text{ K}$$

∴ The virtual temperature of the air is approximately 302.04 K.

6. Lecture 6 28/08/2024

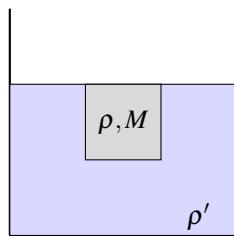
6.1 Archimedes Principle

Archimedes' principle states that any object completely or partially submerged in a fluid (liquid or gas) is buoyed up by a force equal to the weight of the fluid that the object displaces.

Upward force exerted by the fluid = weight of the fluid displaced by the object

6.2 Buoyancy

Buoyancy is the upward force exerted by a fluid (liquid or gas) that opposes the weight of an object submerged in it.



Net force acting on mass(M) and density(ρ) submerged in fluid of density(ρ') is given by:

$$F_B = \rho'Vg - Mg \quad (76)$$

$$= \rho'Vg - \rho Vg \quad (77)$$

$$= (\rho' - \rho)Vg \quad (78)$$

Dividing equation with M on both side,

$$\frac{F_B}{M} = \frac{(\rho' - \rho)Vg}{M} \quad (79)$$

$$f_B = \frac{(\rho' - \rho)Vg}{\rho V} \quad (80)$$

$$= \left(\frac{\rho'}{\rho} - 1 \right) g \quad (81)$$

If Buoyant force per unit mass (f_B),

$$f_B > 0 \rightarrow \text{upward force}$$

$$f_B < 0 \rightarrow \text{downward force}$$

We don't measure density in real case scenario, so we need to convert the equation in the useful form.

Assume pressure inside air parcel and surrounding equal and process to be reversible.

Using Ideal gas equation,

$$P = \rho R_d T_v \quad (82)$$

$$P = \rho' R_d T'_v \quad (83)$$

Substituting equation Eq.(82) & (83) in Eq.(81), we get:

$$f_B = \frac{\left(\frac{P}{R_d T'_v} - \frac{P}{R_d T_v} \right)}{\frac{P}{R_d T'_v}} g \quad (84)$$

$$\Rightarrow f_B = \frac{(T_v - T'_v)}{T'_v} g \quad (85)$$

where T_v and T'_v are virtual temperature of of parcel and fluid respectively.

$$f_B > 0 \rightarrow T_v > T'_v \rightarrow \text{upward force}$$

$$f_B < 0 \rightarrow T_v < T'_v \rightarrow \text{downward force}$$

$$f_B = 0 \rightarrow T_v = T'_v \rightarrow \text{no net force}$$

Question 6.1: A parcel of air has a temperature of 29°C and specific humidity of 24g/kg . It is embedded in the environment having temperature of 30°C and specific humidity of 5g/kg

a. What is vertical acceleration?

b. If there are no forces acting on it, how long would take for the parcel to raise 10m from starting position?

\Rightarrow a. Given $T_{v,a} = 29^\circ\text{C} = 302\text{K}$ and $T_{v,s} = 30^\circ\text{C} = 303\text{K}$,

$$q_a = \text{R.H.}_a = 24\text{g/kg},$$

$$q_s = \text{R.H.}_s = 5\text{g/kg},$$

We know,

$$\begin{aligned} T_{v,a} &= T_a(1 + 0.61q_a) \\ &= 302(1 + 0.61 \times 24 \times 10^{-3}) \\ &= 306.42128\text{K} \end{aligned}$$

$$\begin{aligned} T_{v,s} &= T_s(1 + 0.61q_s) \\ &= 303(1 + 0.61 \times 5 \times 10^{-3}) \\ &= 303.92415\text{K} \end{aligned}$$

Buoyant force per unit mass f_B ,

$$\begin{aligned} f_B &= \left(\frac{T_{v,a} - T_{v,s}}{T_{v,s}} \right) g \\ &= \left(\frac{306.42128 - 303.92415}{303.92415} \right) \times 9.81 \\ &= 0.0806\text{m/s}^2 \end{aligned}$$

\therefore Vertical acceleration due to buoyant force is 0.0806m/s^2 .

\Rightarrow b. Given Height $h = 10\text{m}$,

Vertical acceleration $a = f_B = 0.0806\text{m/s}^2$

Using equation of motion:

$$s = \frac{1}{2}at^2$$

$$10 = \frac{1}{2} \times 0.0806 \times t^2$$

$$t = 15.7524\text{s}$$

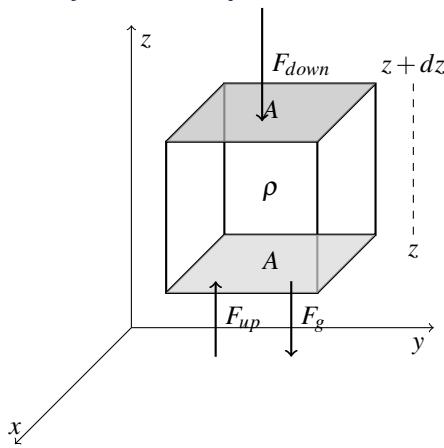
\therefore Time taken by parcel to raise 10m due to buoyant force is 15.75 seconds.

6.3 Hydrostatic equation

\therefore we can rewrite Hydrostatic equation as following Eq.(95):

$$\Rightarrow g = -\frac{1}{\rho} \frac{dP}{dz} \quad (95)$$

This is called **Hydrostatic approximation**.



1. The downward force due to gravity:

$$F_g = mg = \rho(A\delta z)g \quad (86)$$

2. The upward force due to atmosphere force acting on the bottom of the slab is given by:

$$F_{up} = Ap(z) \quad (87)$$

3. The downward force acting on parcel:

$$F_{down} = Ap(z + dz) \quad (88)$$

\therefore Net upward force will be given by Eq.(90)

$$F = F_{up} - F_{down} - F_g \quad (89)$$

$$= Ap(z) - Ap(z + dz) - \rho(A\delta z)g \quad (90)$$

Hence, Upward acceleration will be:

$$a = \frac{F}{\rho A \delta z} \quad (91)$$

$$= \frac{Ap(z) - Ap(z + \delta z) - (A\delta z)\rho g}{\rho A \delta z} \quad (92)$$

$$= -\frac{1}{\rho} \left[\frac{p(z + \delta z) - p(z)}{\delta z} \right] - g \quad (93)$$

Taking $\lim_{\delta z \rightarrow 0}$

$$\Rightarrow a + g = -\frac{1}{\rho} \frac{dp}{dz} \quad (94)$$

Eq.(94) is called Hydrostatic equation.

Question 6.2: Velocity of hyrcene is 10m/s and time taken 10min find acceleration.

$$\Rightarrow \text{Acceleration } a = \frac{v}{t} = \frac{10}{10 \times 60} = \frac{10}{600} \approx 0.0167$$

From the above example question 6.2 we can infer that $a + g \approx g$,

7. Lecture 7 29/08/2024

7.1 Ideal gas equation with Hydrostatic equation

From Ideal gas equation:

$$P = \rho R_d T_v \quad (96)$$

$$\rho = \frac{P}{R_d T_v} \quad (97)$$

Substitute ρ in hydrostatic Eq.(95), we get:

$$g = -\frac{R_d T_v}{P} \frac{\partial P}{\partial z} \quad (98)$$

$$\frac{\partial P}{\partial z} = -\frac{P}{R_d T_v} g \quad (99)$$

$$\frac{1}{P} \frac{\partial P}{\partial z} = -\frac{g}{R_d T_v} \quad (100)$$

$$\frac{\partial \ln P}{\partial z} = -\frac{g}{R_d T_v} \quad (101)$$

The rate of change of logarithm of pressure is Inversely proportional to temperature and does not depend on pressure.

7.2 Geopotential

Geopotential at any point in the atmosphere is defined as the work done against the gravitational field to raise a mass of 1kg from sea level to that point,

Represented by Eq.(102) and has unit $J kg^{-1}$

$$d\phi = gdz \quad (102)$$

From hydrostatic equation Eq.(94),

$$dp = -\rho g dz \quad (103)$$

$$gdz = -\frac{1}{\rho} dp \quad (104)$$

$$gdz = -\alpha dp \quad (105)$$

where α is specific volume.

Integrating Eq.(102), we get:

$$\int_0^{\phi(z)} d\phi = \int_0^z gdz \quad (106)$$

$$\phi(z) = \int_0^z gdz \quad (107)$$

Let g_0 be acceleration due to gravity averaged over the surface.

$$\frac{\phi(z)}{g_0} = \int_0^z \frac{g}{g_0} dz \quad (108)$$

$$\Rightarrow Z = \frac{\phi(z)}{g_0} \quad (109)$$

The Z in Eq(109) is called Geopotential height.

z (km)	Z (km)	g (m/s²)
0	0	9.81
1	1	9.80
10	9.99	9.77
100	98.87	9.50
500	46.36	8.43

Table 2. Deviation of values of g for Geometric Height(z), Geopotential Height(Z)

$$p = \rho R_d T_v \quad (110)$$

$$p\alpha = R_d T_v \quad (111)$$

$$\alpha = \frac{R_d T_v}{p} \quad (112)$$

$$d\phi = gdz = -\alpha dp \quad (113)$$

$$d\phi = -\frac{R_d T_v}{p} dp \quad (114)$$

Integrating from both sides, we get:

$$\int_{\phi_1}^{\phi_2} d\phi = \int_{P_1}^{P_2} -\frac{R_d T_v}{p} dp \quad (115)$$

$$\phi_2 - \phi_1 = -R_d \int_{P_1}^{P_2} T_v \frac{dp}{p} \quad (116)$$

Dividing both side with g_0 , we get:

$$\frac{(\phi_2 - \phi_1)}{g_0} = -\frac{R_d}{g_0} \int_{P_1}^{P_2} T_v \frac{dp}{p} \quad (117)$$

$$(Z_2 - Z_1) = -\frac{R_d}{g_0} \int_{P_1}^{P_2} T_v \frac{dp}{p} \quad (118)$$

By assuming isothermal atmosphere

$$(Z_2 - Z_1) = -\frac{R_d}{g_0} \int_{P_1}^{P_2} \bar{T}_v \frac{dp}{p} \quad (119)$$

$$(Z_2 - Z_1) = -\frac{R_d}{g_0} \bar{T}_v \ln\left(\frac{dp}{p}\right) \quad (120)$$

$$\Rightarrow (Z_2 - Z_1) = -H \ln\left(\frac{dp}{p}\right) \quad (121)$$

Where \bar{T}_v is average temperature of atmosphere taken over geopotential (ϕ_1 & ϕ_2) and H is scale height given by Eq.(122):

$$H = \frac{R_d T_v}{g_0} \quad (122)$$

Scale height H is defined as height at which the pressure reduces to $1/e$ times the surface pressure. It is around 7.8km for Earth's atmosphere.

Eq.(121) is called Hypsometric equation.

Simplifying Eq.(121), we get:

$$P_2 = P_1 e^{-\frac{(Z_2 - Z_1)}{H}} \quad (123)$$

$$\Rightarrow P = P_0 e^{-\frac{(Z_2 - Z_1)}{H}} \quad (124)$$

Question 7.1:

⇒ Solution,

$$\begin{aligned}(Z_2 - Z_1) &= -\frac{R_d}{g_0} \bar{T}_v \ln \left(\frac{dp}{p} \right) \\&= -\frac{287 \times 255}{9.81} \ln \frac{P_2}{P_1} \\&= -7.4 \ln \frac{P_2}{P_1}\end{aligned}$$

8. Lecture 8 30/08/2024

Question 8.1: On May 20 2020 tropical cyclone Ampan of centre pressure at ocean surface dropped to 920hPa. The surrounding region away from influence of centre of cyclone had mean sea level pressure of 1010hPa. The height depression associated with centre of cyclone vanished at height of pressure level of 150hPa. If the mean virtual temperature of the surrounding between the surface at 150hPa was -10°C . What was the corresponding mean virtual temperature in the centre of storm?

⇒ Given data:

$$P_1 = 1010\text{hPa} = 101000\text{Pa}$$

$$P_2 = 920\text{hPa} = 92000\text{Pa}$$

$$P_3 = 150\text{hPa} = 15000\text{Pa}$$

$$T_{v,surr} = -10^{\circ}\text{C} = 263.15\text{K}$$

$$R_d = 287\text{J/(kgK)}$$

$$g_0 = 9.81\text{m/s}^2$$

⇒ Height difference calculation in the surrounding air:

$$\begin{aligned} (Z_2 - Z_1) &= \frac{R_d T_{v,surr}}{g_0} \ln \left(\frac{P_1}{P_3} \right) \\ &= \frac{287 \times 263.15}{9.81} \ln \left(\frac{101000}{15000} \right) \\ &= 7701.43 \times 1.906 \\ &= 14681.92\text{m} \end{aligned}$$

⇒ Solution for mean virtual temperature at the center of the storm:

$$\begin{aligned} (Z_2 - Z_1) &= -\frac{R_d \bar{T}_{v,center}}{g_0} \ln \left(\frac{P_3}{P_2} \right) \\ 14681.92 &= -\frac{287}{9.81} \bar{T}_{v,center} \ln \left(\frac{15000}{92000} \right) \\ \bar{T}_{v,center} &= -\frac{14681.92 \times 9.81}{287 \ln \left(\frac{15000}{92000} \right)} \\ &= -\frac{14681.92 \times 9.81}{287 \times (-1.7749)} \\ &= \frac{144040.6}{-509.4613} \\ &= 282.72\text{K} \end{aligned}$$

$$T_{v,center} \approx 282.72 - 273.15$$

$$T_{v,center} \approx 9.57^{\circ}\text{C}$$

Question 8.2: Calculate the thickness of layer between 1000hPa and 500hPa pressure surface.

- a. At point in tropics where T_v is 15°C
- b. At point in polar where T_v is -40°C

Solution ⇒ a.

$$\begin{aligned} (Z_2 - Z_1) &= -\frac{R_d \bar{T}_v}{g_0} \ln \left(\frac{p_1}{p_2} \right) \\ &= \frac{287 \times 288\text{K}}{9.81\text{m/s}^2} \ln \frac{1000\text{hPa}}{500\text{hPa}} \\ &= 5840.2419\text{km} \end{aligned}$$

⇒ b.

$$\begin{aligned} (Z_2 - Z_1) &= -\frac{R_d \bar{T}_v}{g_0} \ln \left(\frac{p_1}{p_2} \right) \\ &= \frac{287 \times 233\text{K}}{9.81\text{m/s}^2} \ln \frac{1000\text{hPa}}{500\text{hPa}} \\ &= 4724.9179\text{km} \end{aligned}$$

Pressure profiles in the idealized atmosphere

8.1 Constant density atmosphere

Assume atmosphere is at hydrostatic balance and density ρ to be constant.

$$dP = -\rho g dz \quad (125)$$

$$\int_{P(0)}^{P(z)} dP = - \int_0^z \rho g dz \quad (126)$$

$$p(z) - p(0) = -\rho g z \quad (127)$$

Substituting values in above equation, we obtain:

$$(0 - 101.3) = -1.23 \times 9.8 \times z \quad (128)$$

$$z \approx 8.3952\text{km} \quad (129)$$

Using Ideal gas equation, substitute P ,

$$d(\rho RT) = -\rho g dz \quad (130)$$

$$RdT = -gdz \quad (131)$$

$$dT = -\frac{g}{R} dz \quad (132)$$

$$T(z) - T(0) = \frac{g}{R}(z - 0) \quad (133)$$

$$T(z) = T(0) - \frac{g}{R} z \quad (134)$$

$$T(z) = T(0) - 0.0341z \quad (135)$$

$$T(z) = T(0) - 34.1z \quad (136)$$

where 34.1 constant have an unit of $^{\circ}\text{C}/\text{km}$

$$\Rightarrow \Gamma = -\frac{dT}{dz} = -\frac{g}{R} = -34.1^{\circ}\text{C}/\text{km} \quad (137)$$

This is Ideal/theoretical value, but **actual/practical value for Gamma is $6.5^\circ/km$** because of phenomenon called **auto-convective lapse rate** i.e. ρ varies with altitude, and warm air and cold air do vertical circulation.

8.2 Constant temperature atmosphere

Assume atmosphere is at hydrostatic balance and Temperature T to be constant.

$$dP = -\rho g dz \quad (138)$$

$$d\rho RT = -\rho g dz \quad (139)$$

$$p(z) - p(0) = -\rho g dz \quad (140)$$

$$RT d\rho = -\rho g dz \quad (141)$$

$$\frac{d\rho}{\rho} = -\frac{g}{RT} dz \quad (142)$$

$$\ln \rho|_{\rho_1}^{\rho_2} = -\frac{g}{RT} z|_0^z \quad (143)$$

$$\Rightarrow \ln \frac{\rho_2}{\rho_1} = -\frac{g}{RT} z \quad (144)$$

8.3 Constant lapse rate atmosphere

$$T = T_0 - \Gamma_z \quad (145)$$

$$dP = -\rho g dz \quad (146)$$

$$\frac{dP}{dz} = -\frac{Pg}{RT} \quad (147)$$

$$\frac{dP}{dz} = -\frac{Pg}{R(T_0 - \Gamma_z)} \quad (148)$$

$$\frac{1}{P} dP = -\frac{gdz}{R(T_0 - \Gamma_z)} \quad (149)$$

$$\int_{P_1}^{P_2} \frac{1}{P} dP = -\frac{g}{R} \int_0^z \frac{dz}{(T_0 - \Gamma_z)} \quad (150)$$

$$\Rightarrow \ln \frac{P_2}{P_1} = -\frac{g}{R\Gamma_z} \ln \left(\frac{T_0 - \Gamma_z}{T_0} \right) \quad (151)$$

9. Lecture 9 04/09/2024

9.1 1st law of thermodynamics and it's application

Pressure-volume work

Work done δw by any force F to displace object with displacement ds is equal to:

$$\delta w = F \cdot ds \quad (152)$$

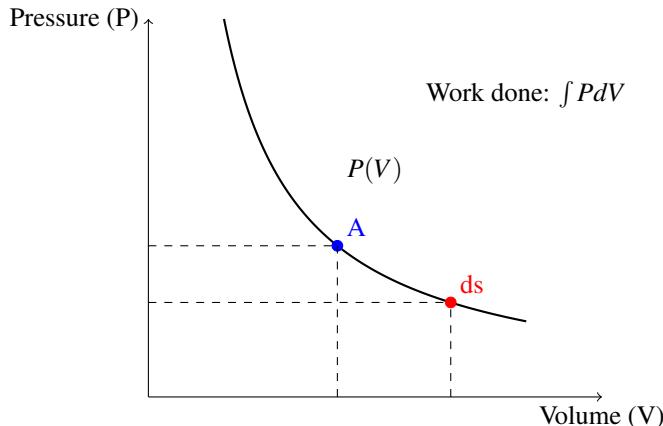
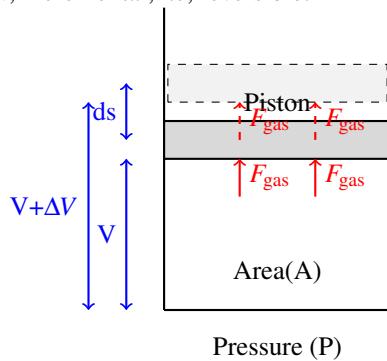
Incremental Work done by Force F to increase the volume will be given as follows:

$$\delta w = F \cdot ds \quad (153)$$

$$= PA \cdot ds \quad (154)$$

$$= PdV \quad (155)$$

Assuming Pressure P constant at each step, and process is slow, incremental, i.e, reversible.



10. Lecture 10 05/09/2024

10.1 . 1st law of thermodynamics and it's application

$$\Delta U = Q - W \quad (156)$$

Mechanical work

$$\delta w = F \cdot ds \quad (157)$$

Pressure-volume work

$$\delta w = PA \cdot ds \quad (158)$$

$$\delta w = PdV \quad (159)$$

$$w = \int_i^f PdV \quad (160)$$

$$\delta w = PAds \quad (171)$$

$$\delta w = F \cdot ds \quad (172)$$

$$\frac{\delta w}{dt} = m \frac{dv}{dt} \frac{ds}{dt} \quad (173)$$

$$\frac{\delta w}{dt} = mv \frac{dv}{dt} \quad (174)$$

$$\frac{\delta w}{dt} = \frac{d}{dt} \left(\frac{1}{2} mv^2 \right) \quad (175)$$

$$\Rightarrow \frac{\delta w}{dt} = \frac{d}{dt} (K.E.) \quad (176)$$

Formulation of 1st law of thermodynamics

Case 1: Heating

$$\Delta U = Q$$

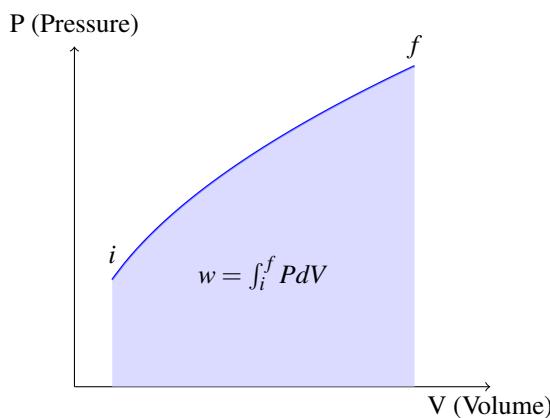
Case 2: By doing work

$$\Delta U = -W$$

From case 1 and 2 for we can write 1st law of thermodynamics as:

$$\Delta U = Q - W \quad (177)$$

$$\delta q = du + \delta w \quad (178)$$



$$w = \int_i^f PdV \quad (161)$$

$$w = \int_i^f PdV + \int_f^i PdV \quad (162)$$

$$w = \left[\int_i^f PdV \right]_1 - \left[\int_i^f PdV \right]_2 \quad (163)$$

$$w \neq 0 \quad (164)$$

$$\Rightarrow \oint_C w = \oint_C PdV \neq 0 \quad (165)$$

$$w = \int_i^f PdV \quad (166)$$

$$w = \int_i^f PdV + \int_f^i PdV \quad (167)$$

$$w = \left[\int_i^f PdV \right]_1 - \left[\int_i^f PdV \right]_2 \quad (168)$$

$$w = 0 \quad (169)$$

$$\Rightarrow \oint_C w = \oint_C PdV = 0 \quad (170)$$

$$\begin{aligned} T_i < T_f \\ T_i \rightarrow T_f \end{aligned}$$

↑ Heat Δ ↑

In terms of Intensive parameters

$$\delta q = du + pd\alpha \quad (179)$$

1. Pressuer-volume wok done by a system = reduction in internal energy + heat supplied by the environment.

2. Pressuer-volume wok done on a system = increase in internal energy + heat transferred to the environment.

10.2 . Heat capacity

$$\frac{\delta q}{dT} = C \quad (180)$$

Unit $Jk^{-1}Kg^{-1}$

Ideal gas equation:

$$P\alpha = R_d T$$

Case 1: Increase in volume (If pressure is kept constant
 → Isobaric process)

Case 2: Increase in pressure (If volume is kept constant
 → Isochoric process)

Case 3: Combination of above the both cases. (i.e. increase in both pressure and volume)

10.3 . Heat capacity at constant volume

From Eq.(179)

$$\delta q = du + pd\alpha \quad (181)$$

$$\delta q = du \quad (182)$$

$$\delta q = C_v dT \quad (183)$$

$$C_v = \left(\frac{\delta q}{dT} \right)_{\alpha=\text{cont}} \quad (184)$$

$$C_v = \left(\frac{du}{dT} \right)_{\alpha=\text{cont}} \quad (185)$$

$$\Rightarrow du = C_v dt \quad (186)$$

From Kinetic theory of gas

$$U = \frac{3}{2} P \alpha = \frac{3}{2} R_d T$$

For monoatomic gas:

$$C_v = \frac{du}{dt} = \frac{3}{2} R_d = 430.5 J k^{-1} kg^{-1}$$

For diatomic gas:

$$C_v = \frac{du}{dt} = \frac{5}{2} R_d = 718 J k^{-1} kg^{-1}$$

11. Lecture 11 06/09/2024

From Ideal gas eqaution

11.1 . Specific heat capacity

$$\text{Specific heat} = \frac{\delta q}{dT}$$

$$\delta q = du + \delta w \quad (187)$$

$$\delta q = du + pd\alpha \quad (188)$$

$$P = R_d T \quad (196)$$

$$dP\alpha = Pd\alpha + \alpha dP \quad (197)$$

$$Pd\alpha + \alpha dP = d(R_d T) \quad (198)$$

$$Pd\alpha + \alpha dP = R_d dT \quad (199)$$

$$Pd\alpha = R_d dT - \alpha dP \quad (200)$$

11.2 . Specific heat at constant volume

Volume(V) is constant ,i.e., specific density(α) = constant

$$\therefore d\alpha = 0$$

$$\delta q = du + pd\alpha \quad (189)$$

$$\delta q = du \quad (190)$$

$$\delta q = C_v dT \quad (191)$$

$$C_v = \left(\frac{\delta q}{dT} \right)_{\alpha=\text{cont}} \quad (192)$$

$$C_v = \left(\frac{du}{dT} \right)_{\alpha=\text{cont}} \quad (193)$$

$$du = C_v dt \quad (194)$$

$$\delta q = du + pd\alpha \quad (201)$$

$$\delta q = du \quad (202)$$

$$\delta q = C_v dT \quad (203)$$

$$C_v = \left(\frac{\delta q}{dT} \right)_{\alpha=\text{cont}} \quad (204)$$

$$C_v = \left(\frac{du}{dT} \right)_{\alpha=\text{cont}} \quad (205)$$

$$du = C_v dt \quad (206)$$

Hence,

$$\Rightarrow \delta q = C_v dt + Pd\alpha \quad (195)$$

From Eq.(195) and Eq.(200)

$$\delta q = C_v dT + R_d dT - \alpha dP \quad (207)$$

$$\delta q = (C_v + R_d) dT - \alpha dP \quad (208)$$

$$\Rightarrow \delta q = C_p dT - \alpha dP \quad (209)$$

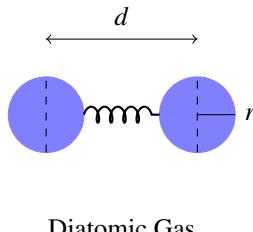
$$C_v = \begin{cases} \frac{3}{2}R_d & \text{for monoatomic gas,} \\ \frac{5}{2}R_d & \text{for diatomic gas.} \end{cases}$$

Rotational K.E. is significant for diaatomic gas but not for monoatomic gas.



Monoatomic Gas

Diatom Gas



11.3 . Specific heat at constant pressure

Pressure(P) is constant $\therefore dP = 0$

$$C_p = \begin{cases} 717.5 \text{ J} \cdot \text{kg}^{-1} \cdot \text{K}^{-1} & \text{for monoatomic gas} \\ 1005 \text{ J} \cdot \text{kg}^{-1} \cdot \text{K}^{-1} & \text{for diatomic gas} \end{cases}$$

11.4 . Special forms of 1st law of thermodynamics

I. Isobaric process

$$\delta q = C_p dT \quad (210)$$

$$= \left(\frac{C_p}{C_v} \right) C_v dT \quad (211)$$

$$= \left(\frac{C_p}{C_v} \right) dU \quad (212)$$

$$= \gamma dU \quad (213)$$

II. Isothermal process

$$\delta q = -\alpha dP = Pd\alpha = \delta w \quad (214)$$

$$q = \int_i^f \alpha dP \quad (215)$$

$$= \int_i^f \frac{R_d T}{P} dP \quad (216)$$

$$= -R_d T \ln P \Big|_i^f \quad (217)$$

$$= -R_d T \ln P \Big|_i^f \quad (218)$$

$$= R_d T \ln \alpha \Big|_i^f = w \quad (219)$$

$$(220)$$

III. Isochoric process

$$\delta q = C_v dT = dud\alpha = 0 \quad (221)$$

$$q = \int_i^f C_v dT \quad (222)$$

$$= C_v(T_f - T_i) \quad (223)$$

$$= C_v \Delta T = U \quad (224)$$

IV. Adiabatic process

$$0 = C_v dT + P\alpha \quad (225)$$

$$0 = C_p dT - P\alpha \quad (226)$$

Question 11.1: For each of the following conditions compute:

- i. Mechanical work done by the sample of air.
- ii. Heat added to the sample.
- a. Isothermal compression to $1/5^{th}$ of its original volume at $15^\circ C$.
- b. Isobaric heating from $0^\circ C$ to $20^\circ C$.
- c. Adiabatic expansion to 5 times its original volume within initial temperature of $20^\circ C$.

⇒ Solution:

a. For Isothermal process

$$\begin{aligned} q = w &= R_d T \ln \left(\frac{1}{V} \right) \Big|_V^{V/5} \\ &= 287 \times 288 \times T \ln(5) \\ &= 133.0297 kJ \end{aligned}$$

b. For Isobaric heating

$$\begin{aligned} q &= C_p dT \\ &= \frac{7}{2} R_d \times \Delta T \\ &= \frac{7}{2} \times 287 \times 20 \\ &= 19740 J \\ &\approx 20 kJ \end{aligned}$$

c. Adiabatic heating

$$\begin{aligned} w &= \frac{P_1 V_1 - P_2 V_2}{\gamma - 1} \\ &= \frac{R_d (T_1 - T_2)}{\gamma - 1} \\ &= \frac{287 \times (293 - \frac{293}{5})}{\frac{7}{5} - 1} \\ &= 168.182 kJ \\ q &= 0 J \end{aligned}$$

12. Lecture 12 11/09/2024

12.1 . Poisson's equation for adiabatic transformation

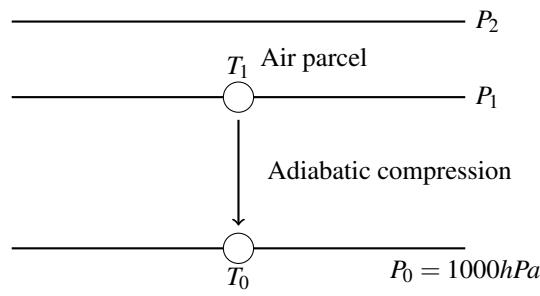
For adiabatic processes:

$$dq = 0 \quad (227)$$

$$c_p dT = \alpha dP \quad (228)$$

$$c_p \frac{dT}{T} = \alpha \frac{dP}{P} \quad (229)$$

$$(230)$$



Let us assume that

From Ideal gas equations:

$$P\alpha = R_d T \quad (231)$$

$$\alpha = \frac{R_d T}{P} \quad (232)$$

$$c_p \frac{dT}{T} = \frac{R_d T}{P} \frac{dP}{T} \quad (233)$$

$$c_p \frac{dT}{T} = R_d \frac{dP}{P} \quad (234)$$

Integrating from an initial temperature T_0 and pressure P_0 to arbitrary temperature and pressure T and P

$$\int_{T_0}^T c_p \frac{dT}{T} = \int_{P_0}^P R_d \frac{dP}{P} \quad (235)$$

$$c_p \ln\left(\frac{T}{T_0}\right) = R_d \ln\left(\frac{P}{P_0}\right) \quad (236)$$

$$\left(\frac{T}{T_0}\right)^{c_p} = \left(\frac{P}{P_0}\right)^{R_d} \quad (237)$$

$$T_0 = T \left(\frac{P}{P_0}\right)^{\frac{R_d}{c_p}} \quad (238)$$

$$\Rightarrow \theta = T \left(\frac{1000}{P_0}\right)^k \quad (239)$$

where constant k is R_d/c_p which is equal to **0.286**, $P_0 = 1000 \text{ hPa}$ which is **near surface pressure** and θ is known as **potential temperature**.

θ known as **potential temperature** Eq.(239) is defined when a adiabatically compressed parcel and brought to 1000 hPa isobar (near surface) and temperature is measure.

taking logarithm on both sides

$$d(\ln \theta) = d(\ln A) + d(\ln T) - kd(\ln P) \quad (241)$$

$$d(\ln \theta) = d(\ln T) - \frac{R_d}{C_p} d(\ln P) \quad (242)$$

$$\text{where } A = (P_0)^k = (1000)^{0.286}$$

Consider 1st law of thermodynamics

$$\delta q = C_p dT - \alpha dP \quad (243)$$

$$\frac{\delta q}{T} = C_p \frac{dT}{T} - \alpha \frac{dP}{T} \quad (244)$$

$$\frac{\delta q}{C_p T} = \frac{dT}{T} - \frac{R_d T dP}{C_p P T} \quad (245)$$

$$\frac{\delta q}{C_p T} = \frac{dT}{T} - \frac{R_d dP}{C_p P} \quad (246)$$

$$\frac{\delta q}{C_p T} = d(\ln T) - \frac{R_d}{C_p} d(\ln P) \quad (247)$$

from Eq.(242) and Eq.(247)

$$d(\ln \theta) = \frac{\delta q}{C_p T} \quad (248)$$

For adiabatic process $\delta q = 0$

$$d(\ln \theta) = 0 \quad (249)$$

θ is constant, conserved for adiabatic process.

Question 12.1: Transcontinental airline flies at an altitude of 12km where the temperature outside is $-55^{\circ}C$ and the pressure is approximately 200hPa

- Compute the potential temperature of air at this altitude.
- Cabin pressure is typically mentioned of 750hPa corresponding to pressure altitude of 2.24km. when outside air is adiabatically compressed to cabin pressure

⇒ Solution:

a.

$$\begin{aligned}\theta &= T \left(\frac{1000}{P} \right)^k \\ &= 218 \left(\frac{1000}{200} \right)^{0.286} \\ &= 218(5)^{0.286} \\ &= 345.4317K \\ &= 72.43^{\circ}C\end{aligned}$$

b.

$$\begin{aligned}\theta &= T \left(\frac{1000}{P} \right)^k \\ &= 218 \left(\frac{750}{200} \right)^{0.286} \\ &= 218 \left(\frac{4}{3} \right)^{0.286} \\ &= 318.14K \\ &= 45^{\circ}C\end{aligned}$$

13. Lecture 13 12/09/2024

13.1 . Adiabatic transformation - Poisson's equation

1. Case I.

$$C_p dT = \alpha dP \quad (250)$$

$$C_p \frac{dT}{T} = R_d \frac{dP}{P} \quad (251)$$

$$\frac{dT}{T} = \left(\frac{C_p - C_v}{C_p} \right) \frac{dP}{P} \quad (252)$$

$$\frac{dT}{T} = \left(1 - \frac{C_v}{C_p} \right) \frac{dP}{P} \quad (253)$$

$$\frac{dT}{T} = \left(1 - \frac{1}{\gamma} \right) \frac{dP}{P} \quad (254)$$

$$\frac{dT}{T} = \left(\frac{\gamma - 1}{\gamma} \right) \frac{dP}{P} \quad (255)$$

$$\ln T = \left(\frac{\gamma - 1}{\gamma} \right) \ln P + \ln C \quad (256)$$

$$T = CP^{\left(\frac{\gamma-1}{\gamma}\right)} \quad (257)$$

$$TP^{\left(\frac{1-\gamma}{\gamma}\right)} = C \quad (258)$$

2. Case II.

$$C_v dT = -P d\alpha \quad (259)$$

$$C_v \frac{dT}{T} = -R_d \frac{d\alpha}{\alpha} \quad (260)$$

$$\frac{dT}{T} = -\left(\frac{C_p - C_v}{C_v} \right) \frac{d\alpha}{\alpha} \quad (261)$$

$$\frac{dT}{T} = \left(1 - \frac{C_p}{C_v} \right) \frac{d\alpha}{\alpha} \quad (262)$$

$$\frac{dT}{T} = \left(1 - \gamma \right) \frac{d\alpha}{\alpha} \quad (263)$$

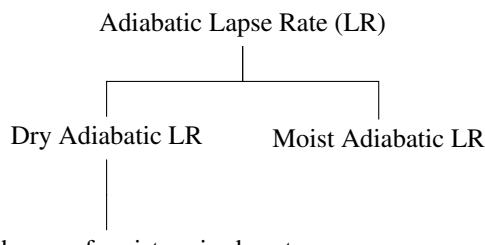
$$\ln T = (1 - \gamma) \ln \alpha + \ln C \quad (264)$$

$$\ln T = \ln \alpha^{(1-\gamma)} + \ln C \quad (265)$$

$$T = C \alpha^{(1-\gamma)} \quad (266)$$

$$T \alpha^{(\gamma-1)} = C \quad (267)$$

13.2 . Adiabatic Lapse Rate



Note: Moisture is present in dry air parcel but it is assumed that it does not show any phase change.

13.3 . Dry Adiabatic Lapse Rate (DALR)

Using the 1st law of thermodynamics

$$\delta q = C_p dT - \alpha dP \quad (268)$$

Adiabatic process

$$\delta q = 0 \quad (269)$$

$$C_p dT = \alpha dP \quad (270)$$

Lapse rate

$$\frac{dT}{dz} = \frac{dP}{dP} \frac{dT}{dz} \quad (271)$$

From Eq.(270) and (271)

$$\frac{dT}{dz} = \frac{R_d T}{C_p P} \frac{dP}{dz} \quad (272)$$

Note: Hydrostatic balance is applied to surrounding not on parcel.

Applying hydrostatic approximation, The pressure of unconfined air parcel is same as that of the environment, we get:

$$\frac{dP}{dz} = \frac{dP'}{dz} = \rho' g = -\frac{P' g}{R_d T'} \quad (273)$$

where P' and T' are ambient pressure and temperature.

From Eq.(270), Eq.(271) and (273)

$$\frac{dT}{dz} = \frac{R_d T}{C_p P} \times \left(-\frac{P' g}{R_d T'} \right) \quad (274)$$

$$\frac{dT}{dz} = \frac{g}{C_p} \times \left(-\frac{T' P'}{T' P} \right) \quad (275)$$

$$\frac{dT}{dz} = -\frac{g}{C_p} \left(\frac{T}{T'} \right) \quad (276)$$

$$\Gamma_d = -\frac{dT}{dz} = \frac{g}{C_p} \left(\frac{T}{T'} \right) \quad (277)$$

$$\Gamma_d = -\frac{dT}{dz} \approx \frac{g}{C_p} \quad (278)$$

$$\Gamma_d = -\frac{dT}{dz} \approx 9.8 \times 10^{-3} {}^\circ C/m \quad (279)$$

$$\Rightarrow \Gamma_d = -\frac{dT}{dz} \approx 9.8 {}^\circ C/km \quad (280)$$

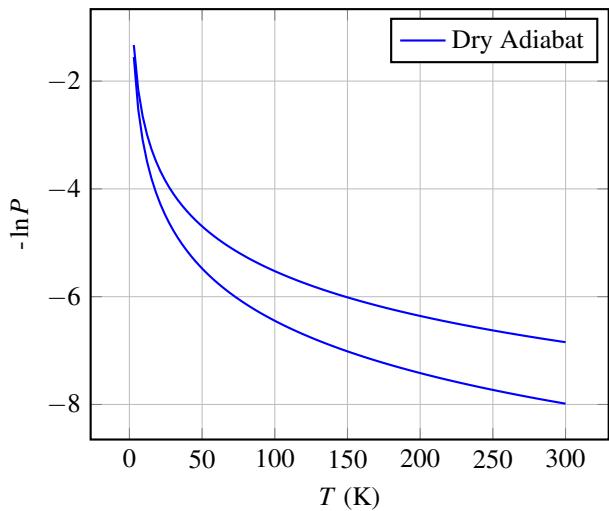
where Γ_d is **Dry Adiabatic Lapse Rate (DALR)** which is approximately equal to $9.8 {}^\circ C/km$, also T and T' are comparably equal.

13.4 . Possion's equations

From Eq.(239)

$$\theta = T \left(\frac{1000}{P} \right)^k \quad (281)$$

$$\Rightarrow P^k = \left(\frac{P_0^k}{\theta} \right) T \quad (282)$$



14. Lecture 14 20/09/2024

14.1 . Heat Engines

- Construct closed cycle of compression and expansion to produce net work.
- Production of work required expenditure of internal energy on heat supplied by the environment.

From 1st law of thermodynamics

$$\delta q = du + \delta w \quad (283)$$

$$\oint \delta q = \oint du + \oint \delta w \quad (284)$$

$$\oint \delta q = \oint^0 du + \oint \delta w \quad (\because u_i = u_f) \quad (285)$$

$$\oint \delta q = \oint \delta w \quad (286)$$

$$\Rightarrow q_{\text{net}} = w_{\text{net}} \quad (\text{Theoretically}) \quad (287)$$

In theory, the equality $q_{\text{net}} = w_{\text{net}}$ arises from the assumption of a perfect heat engine operating in a closed cycle. This idealized scenario implies that all the heat energy supplied to the engine is converted into mechanical work without any losses due to friction, heat dissipation, or other irreversibilities.

However, no physical engine can achieve this ideal efficiency. Real engines inevitably encounter energy losses through various mechanisms, such as:

- Heat Loss:** Part of the input heat is lost to the surroundings, decreasing the effective energy available for work.
- Friction:** Mechanical losses due to friction in moving parts result in energy dissipation as heat, further reducing the net work output.
- Non-ideal Processes:** Real thermodynamic processes often involve irreversible changes, which lead to additional energy losses that are not accounted for in the ideal model.

Thus, while $q_{\text{net}} = w_{\text{net}}$ serves as a theoretical benchmark, practical engines operate with efficiencies below this ideal, governed by real-world constraints and inefficiencies.

14.2 . Efficiency

Efficiency of heat engine,

$$\eta = \frac{q_{\text{in}} - q_{\text{out}}}{q_{\text{in}}} = \frac{w}{q_{\text{in}}} \quad (288)$$

14.3 . Carnot Cycle

The Carnot cycle consists of the following four steps, which can be seen in figure:

- Step 1: Reversible isothermal expansion.

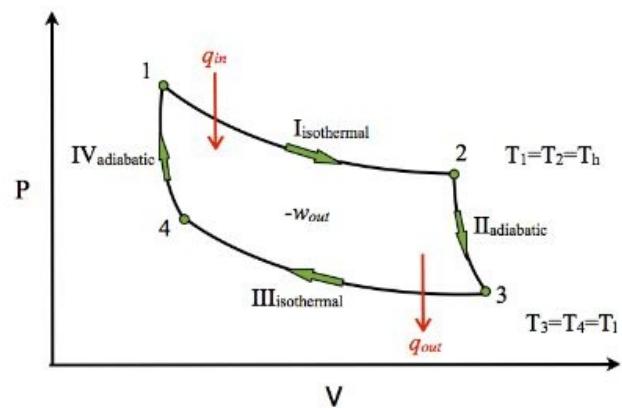


Figure 4. Carnot Cycle

- Step 2: Reversible adiabatic expansion.
- Step 3: Reversible isothermal compression.
- Step 4: Reversible adiabatic compression.

Step 1: Reversible Isothermal Expansion

During this process, the temperature remains constant, i.e., $\Delta T = 0$, which implies that the internal energy change is zero, $\Delta u = 0$.

The work done during this process is given by:

$$w_{12} = \int \delta w = \int_{\alpha_1}^{\alpha_2} p d\alpha \quad (289)$$

$$= \int_{\alpha_1}^{\alpha_2} \frac{R_d T_1}{\alpha} d\alpha \quad (290)$$

$$= R_d T_1 \ln \frac{\alpha_2}{\alpha_1} \quad (291)$$

Since $d\alpha = 0$ (because $dT = 0$), we have:

$$du_{12} = dw_{12} \Rightarrow Q_{12} = W_{12}$$

Step 2: Reversible Adiabatic Expansion

In this step, the gas expands without heat exchange, leading to a change in internal energy. The change in internal energy is given by:

$$\Delta u_{23} = C_v (T_2 - T_1) \quad (292)$$

$$-\Delta u_{23} = w_{23} \quad (293)$$

$$w_{23} = -C_v (T_2 - T_1) \quad (294)$$

Step 3: Reversible Isothermal Compression

In this process, the gas is compressed isothermally, maintaining a constant temperature:

$$w_{34} = \int \delta w = \int_{\alpha_3}^{\alpha_4} p d\alpha \quad (295)$$

$$= \int_{\alpha_3}^{\alpha_4} \frac{R_d T_2}{\alpha} d\alpha \quad (296)$$

$$= R_d T_2 \ln \frac{\alpha_4}{\alpha_3} \quad (297)$$

- Step 1: Reversible isothermal expansion.

Step 4: Reversible Adiabatic Compression

During this step, the gas is compressed adiabatically, resulting in a temperature increase and work done on the gas. The work can be expressed as:

$$w_{41} = C_v(T_2 - T_1) \quad (298)$$

From Poisson's equation (239), we have the following relations for the Carnot cycle:

$$T_1 \alpha_2^{\gamma-1} = T_2 \alpha_3^{\gamma-1} \quad (299)$$

$$\frac{T_1}{T_2} = \left(\frac{\alpha_3}{\alpha_2} \right)^{\gamma-1} \quad (300)$$

Similarly, for the other isentropic process:

$$T_1 \alpha_1^{\gamma-1} = T_2 \alpha_4^{\gamma-1} \quad (301)$$

$$\frac{T_1}{T_2} = \left(\frac{\alpha_4}{\alpha_1} \right)^{\gamma-1} \quad (302)$$

From these two equations, we can equate the temperature ratios:

$$\left(\frac{\alpha_3}{\alpha_2} \right)^{\gamma-1} = \left(\frac{\alpha_4}{\alpha_1} \right)^{\gamma-1} \quad (303)$$

$$\frac{\alpha_3}{\alpha_4} = \frac{\alpha_2}{\alpha_1} \quad (304)$$

Total Work:

From Eq.(291), (294), (297), and (298), the total work done in the cycle is:

$$W_{\text{Total}} = w_{12} + w_{23} + w_{34} + w_{41} \quad (305)$$

$$= R_d T_1 \ln \frac{\alpha_2}{\alpha_1} - C_v(T_2 - T_1) + \quad (306)$$

$$R_d T_2 \ln \frac{\alpha_4}{\alpha_3} + C_v(T_2 - T_1) \quad (307)$$

$$= R_d T_1 \ln \frac{\alpha_2}{\alpha_1} + R_d T_2 \ln \frac{\alpha_4}{\alpha_3} \quad (308)$$

$$= R_d \left[T_1 \ln \frac{\alpha_2}{\alpha_1} + T_2 \ln \frac{\alpha_4}{\alpha_3} \right] \quad (309)$$

$$= R_d (T_1 - T_2) \ln \left(\frac{\alpha_2}{\alpha_1} \right) \quad \text{From Eq.(304)} \quad (310)$$

Eq(310)

$$\eta = \frac{Q_1 + Q_2}{Q_1} \quad (311)$$

$$= 1 + \frac{Q_2}{Q_1} \quad (312)$$

$$= 1 + \frac{R_d T_2 \ln \frac{\alpha_3}{\alpha_4}}{R_d T_1 \ln \frac{\alpha_2}{\alpha_1}} \quad (313)$$

$$= 1 - \frac{T_2 \ln \frac{\alpha_2}{\alpha_1}}{T_1 \ln \frac{\alpha_2}{\alpha_1}} \quad (314)$$

$$\Rightarrow \eta = 1 - \frac{T_2}{T_1} \quad (315)$$

The temperature at which carnot engine becomes 100% efficient (i.e. $\eta = 1$) is called **Absolute temeperature (0K)**

In real world scenarios 0K is not achievable and there is no process which is ideally reversible.

//

Question 14.1: For a potential temperature of 290K compute the corresponding temperature at 700hPa and 500hPa. Sketch the corresponding adiabat in skew-T diagram.

$$\Rightarrow \theta = T \left(\frac{P_0}{P} \right)^k, \text{ where } k = \frac{R_d}{C_p} = \frac{287}{1004} = 0.286$$

$$T = \theta \left(\frac{P_0}{P} \right)^k$$

$$T_1 = \theta \left(\frac{700}{1000} \right)^{0.286} = 261.87K \approx -11.2^\circ C$$

$$T_2 = \theta \left(\frac{500}{1000} \right)^{0.286} = 237.88K \approx -35.15^\circ C$$

14.4 . Skew T - log P diagram

Skew-T plot¹ shows combined and also separate plots for Full Skew-T, Isothermal lines, Isobars, Dry Adiabats, Moist Adiabats, Mixing Ratio and Wind Staff.

Finding the efficiency of carnot cycle using Eq.(288) and

¹<https://www.noaa.gov/jetstream/upperair/skew-t-log-p-diagrams>

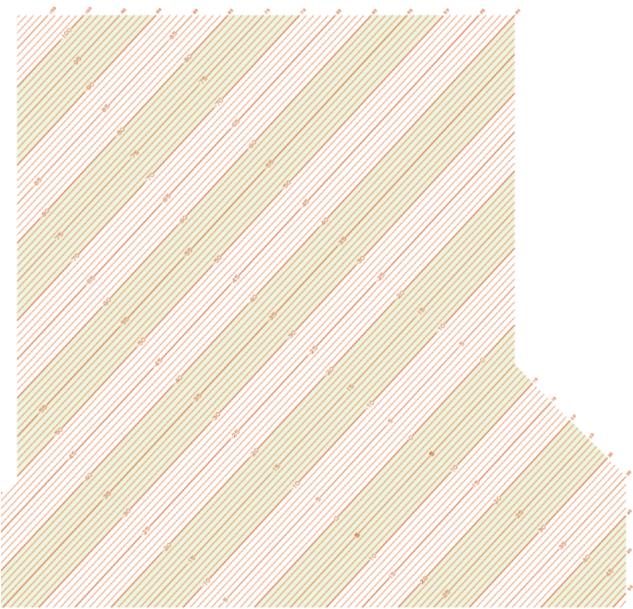


Figure 5. Isothermal lines in Skew-T diagram

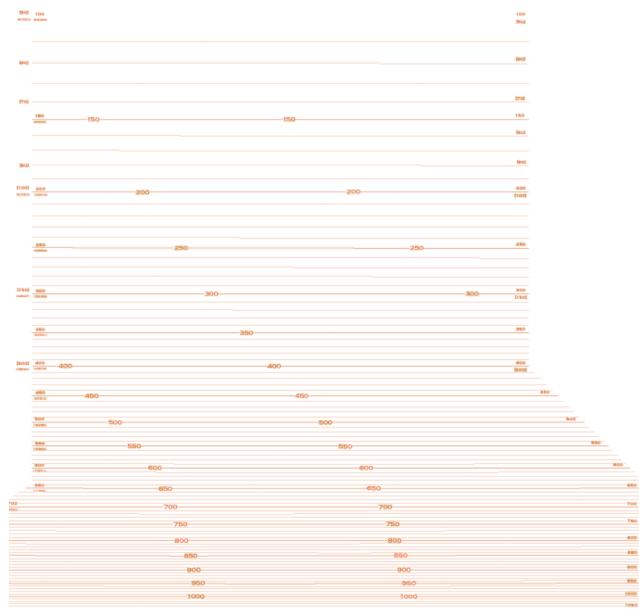


Figure 6. Isobaric lines in Skew-T diagram

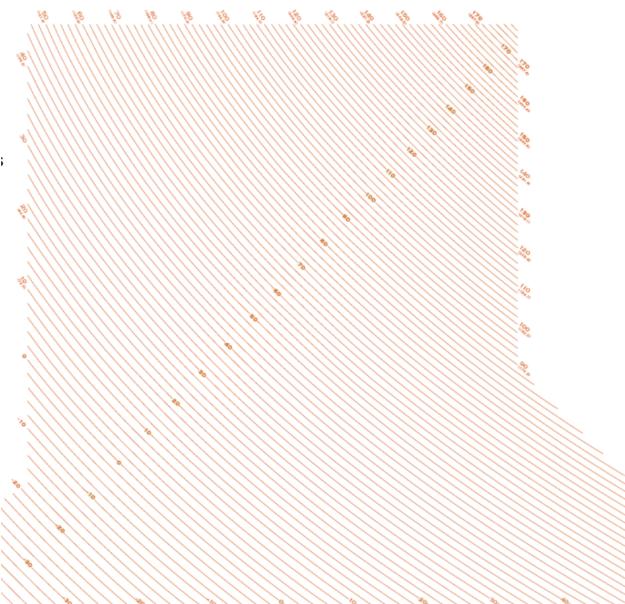


Figure 7. Dry-adiabat lines in Skew-T diagram

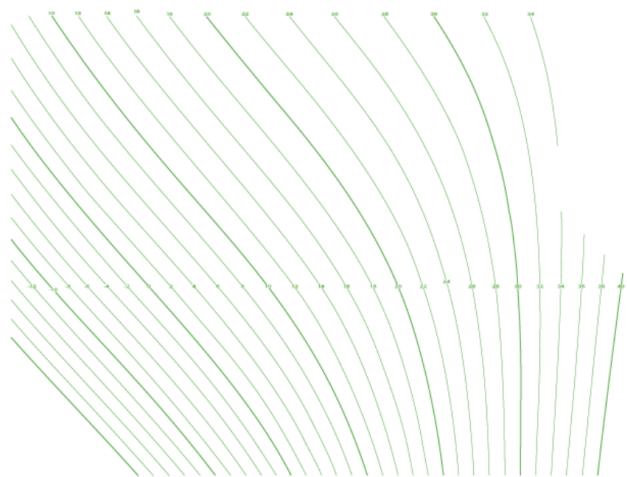


Figure 8. Moist-adiabat lines in Skew-T diagram

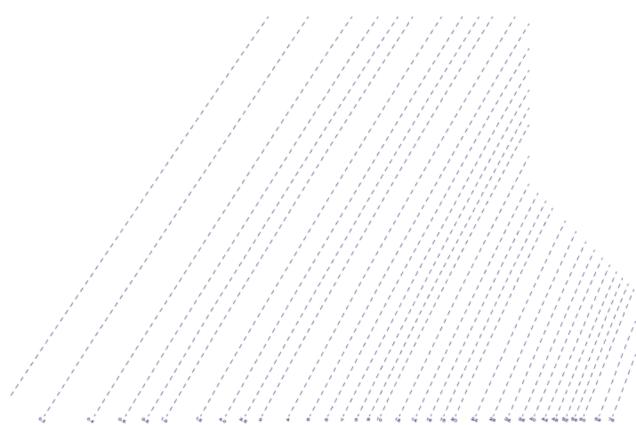


Figure 9. Mixing ratio lines in Skew-T diagram

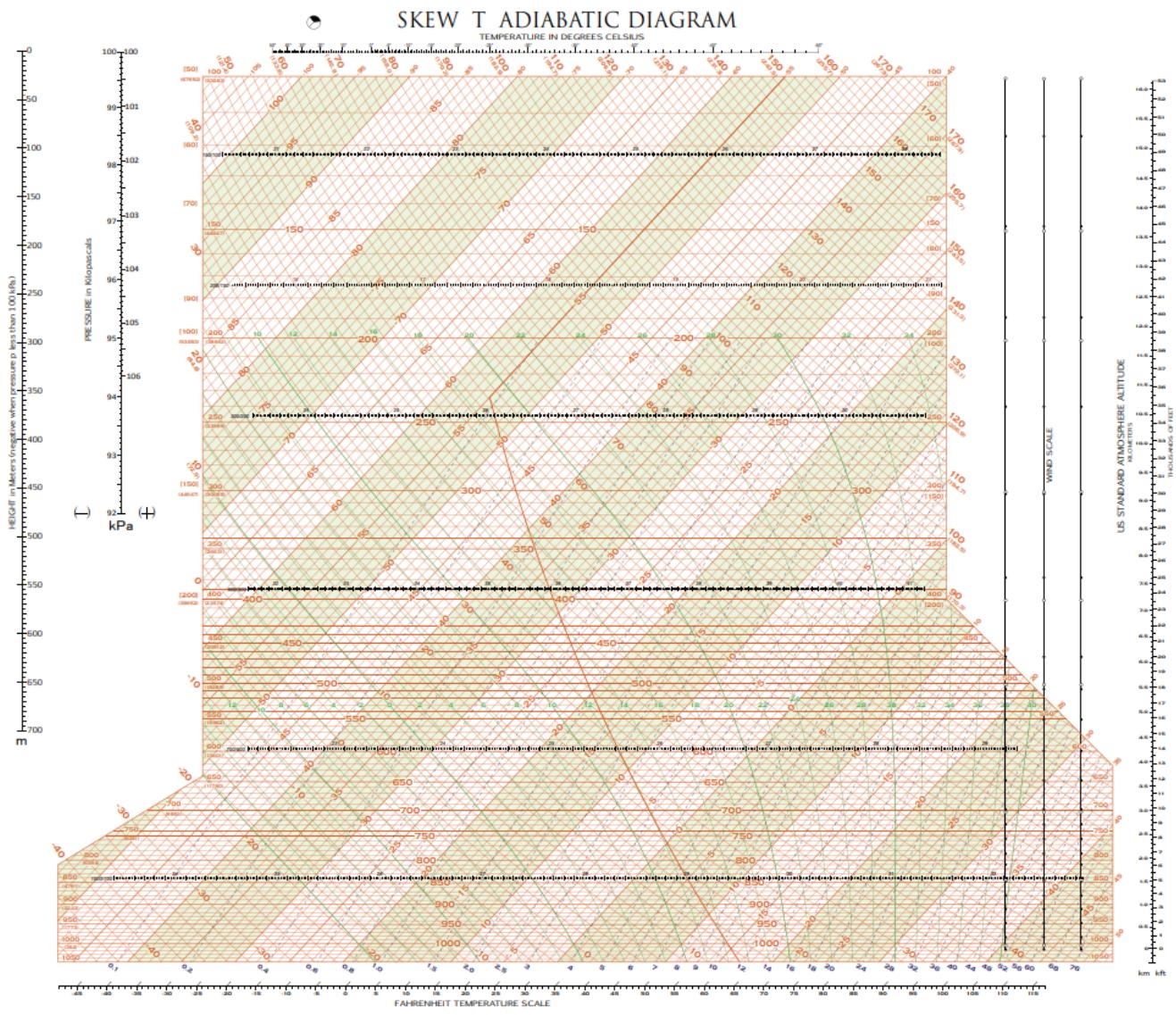


Figure 10. Skew-T diagram

15. Lecture 15 03/10/2024

15.1 . Cyclones

Properties of cyclones

1. Centre of cycle is called Eye of cyclone.
2. Cumulonimbus clouds forms the eye ball of cyclone.
3. Maximum height to where it can reaches is Tropopause

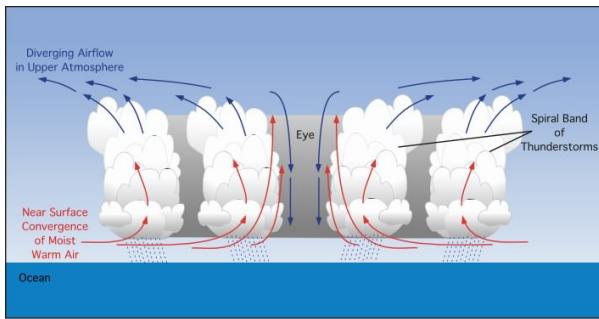


Figure 11. Internal Structure of cyclone

Image source: [Internal Structure of cyclone](#)²

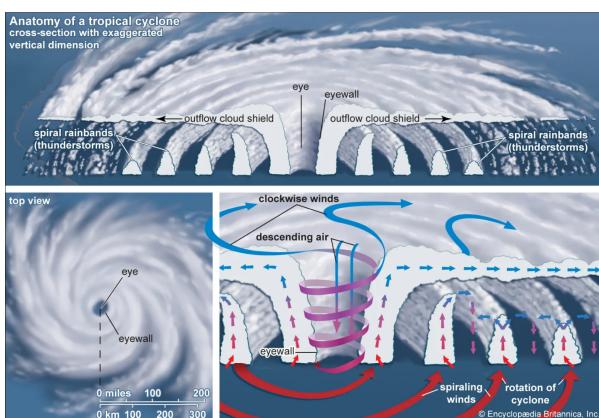


Figure 12. Air circulation in cyclone

Image source: [Air circulation in cyclone](#)³

15.2 . Specific enthalpy

Enthalphy is heat content of state.

$$H = U + PV \quad (316)$$

$$\Delta H = \Delta U + \Delta(PV) \quad (317)$$

$$\Delta H = \Delta U + P\Delta V + V\Delta P \quad (318)$$

$$\Delta H = Q - W + P\Delta V + V\Delta P \quad (319)$$

$$\Delta H = Q - P\Delta V + P\Delta V + V\Delta P \quad (320)$$

$$\Delta H = Q + V\Delta P \quad (321)$$

$$\text{At const. Pressure} \quad (322)$$

$$\Delta H = Q \quad (323)$$

$$h = u + Pv \quad (324)$$

$$dh = du + d(PV) \quad (325)$$

$$dh = du + PdV + VdP \quad (326)$$

$$dh = \delta q + VdP \quad (327)$$

$$\delta q = dh - VdP \text{ also,} \quad (328)$$

$$\delta q = C_p dT - VdP \quad (329)$$

From Eq(328) and Eq(329)

$$dh - VdP = C_p dT - VdP \quad (330)$$

$$\text{At const. Pressure} \quad (331)$$

$$dh = C_p dT \quad (332)$$

Enthalphy is "sensible heat"

Conservative property

For a hydrostatic atmosphere:

$$\frac{dp}{dz} = \rho g \Rightarrow dp = -\rho g dz \quad (333)$$

Substitute Eq(329) in Eq(333)

$$\delta q = C_p dT - \alpha \rho g dz \quad (334)$$

$$\delta q = C_p dT + gdz \quad (335)$$

$$\delta q = dh + d\phi \quad (336)$$

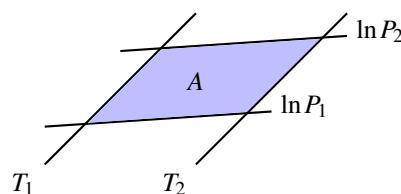
$$\delta q = d(h + \phi) \quad (337)$$

If process is adiabatic $\delta q = 0$

$$h + \phi = \text{const} \quad (338)$$

Eq(338) called dry static equation.

This implies that when air parcel goes up (in geopotential) by the expense of heat content (enthalphy), i.e. $[h \uparrow \phi \downarrow]$.



²https://web.mit.edu/~twcronin/Public/Lupit_Cross_Sections.html

³<https://www.britannica.com/science/tropical-cyclone>

$$\text{Area } A \propto (T_2 - T_1)(\ln P_2 - \ln P_1) \quad (339)$$

$$\propto (T_2 - T_1) \ln \left(\frac{P_2}{P_1} \right) \quad (340)$$

Cycle 1:

$$\begin{aligned} w_1 &= \int P d\alpha \\ &= P \int_{\alpha_1}^{\alpha_2} d\alpha \\ &= P(\alpha_2 - \alpha_1) \\ &= R_d(T_2 - T_1) \end{aligned}$$

Cycle 2:

$$\begin{aligned} w_2 &= \int P d\alpha \\ &= \int_{\alpha_1}^{\alpha_2} \frac{R_d T}{\alpha} d\alpha \\ &= R_d T_2 \ln \left(\frac{\alpha_2}{\alpha_1} \right) \\ &= R_d T_2 \ln \left(\frac{P_1}{P_2} \right) \end{aligned}$$

Cycle 3:

$$\begin{aligned} w_3 &= \int P d\alpha \\ &= P \int_{\alpha_2}^{\alpha_1} d\alpha \\ &= P(\alpha_1 - \alpha_2) \\ &= R_d(T_1 - T_2) \\ &= -R_d(T_2 - T_1) \end{aligned}$$

Cycle 4:

$$\begin{aligned} w_4 &= \int P d\alpha \\ &= \int_{\alpha_2}^{\alpha_1} \frac{R_d T}{\alpha} d\alpha \\ &= R_d T_1 \ln \left(\frac{\alpha_1}{\alpha_2} \right) \\ &= R_d T_1 \ln \left(\frac{P_2}{P_1} \right) \\ &= -R_d T_1 \ln \left(\frac{P_1}{P_2} \right) \end{aligned}$$

$$\begin{aligned} w_{\text{net}} &= w_1 + w_2 + w_3 + w_4 \\ &= R_d(T_2 - T_1) + R_d T_2 \ln \left(\frac{P_1}{P_2} \right) - R_d(T_2 - T_1) - R_d T_1 \ln \left(\frac{P_1}{P_2} \right) \\ &\Rightarrow w_{\text{net}} = 0 \end{aligned}$$

16. Lecture 16 07/10/2024

16.1 . Lesson learnt from carnot cycle

- Thermodynamic efficiency of cyclone depends on sources, given by Eq(315)

When $T_1 = T_2$, carnot cycle does not exist called **Kelvin's Postulate**.

Carnot engine is most efficient at winter because of difference in temperature, causes diverse climate change.

- Transformation of heat is not possible from cold body to hot body called **Clausius postulate**
- Carnot cycle could give us definition for absolute zero temperature.
- All process are irreversible, however, slow process can be considered to be reversible.

16.2 . Entropy

Let's recall that,

$$q_1 = R_d T_1 \ln \left(\frac{\alpha_2}{\alpha_1} \right) \quad (341)$$

$$\frac{q_1}{T_1} = R_d \ln \left(\frac{\alpha_2}{\alpha_1} \right) \quad (342)$$

Similarly,

$$q_2 = -R_d T_2 \ln \left(\frac{\alpha_2}{\alpha_1} \right) \quad (343)$$

$$\frac{q_2}{T_2} = R_d \ln \left(\frac{\alpha_2}{\alpha_1} \right) \quad (344)$$

$$\frac{q_1}{T_1} + \frac{q_2}{T_2} = 0, \quad q_1 > 0, q_2 < 0 \quad (345)$$

Efficiency of heat engine:

$$\eta_{\text{rev}} = 1 - \frac{T_2}{T_1} \quad (346)$$

$$\eta_{\text{rev}} = \frac{T_1 - T_2}{T_1} \quad (347)$$

$$T_1 \eta_{\text{rev}} = T_1 - T_2 \quad (348)$$

Multiply and divide Eq(345) by η_{rev}

$$\frac{q_1}{T_1} \frac{\eta_{\text{rev}}}{\eta_{\text{rev}}} + \frac{q_2}{T_2} = 0 \quad (349)$$

From Eq(348),

$$\frac{q_1}{T_1} \frac{\eta_{\text{rev}}}{(T_1 - T_2)} + \frac{q_2}{T_2} = 0 \quad (350)$$

The above expression is applicable only if process is **perfectly reversible**.

If process is not perfectly reversible, i.e., irreversible

$$\frac{q_1}{T_1} \frac{\eta_{\text{irrev}}}{(T_1 - T_2)} + \frac{q_2}{T_2} < 0 \quad (351)$$

In general,

$$\sum_{i=1}^2 \frac{q_i}{T_i} \leq 0, \begin{cases} = 0 & \text{if perfectly reversible,} \\ < 0 & \text{if irreversible} \end{cases} \quad (352)$$

$$\Rightarrow \sum_{i=1}^N \frac{q_i}{T_i} \leq 0, \quad \text{for } N \text{ number of sources.} \quad (353)$$

If $N \rightarrow \infty$

$$\oint_{\text{cyclic process}} \frac{\delta q}{T} \leq 0 \rightarrow \text{Entropy} \quad (354)$$

$$\delta q = C_v dT + P d\alpha \quad (355)$$

$$\frac{\delta q}{T} = C_v \frac{dT}{T} + P \frac{d\alpha}{T} \quad (356)$$

$$\oint \frac{\delta q}{T} = C_v \oint \frac{dT}{T} + \oint \frac{RT}{\alpha T} d\alpha \quad (357)$$

$$ds = \oint \frac{\delta q}{T} = C_v \oint \frac{dT}{T} + R \oint d \ln \alpha \quad (358)$$

$$ds = \oint \frac{\delta q}{T} = C_v \oint d \ln T + R \oint d \ln \alpha \quad (359)$$

$$\Rightarrow ds = s_f - s_i = \oint_i^f \frac{\delta q}{T} \quad (360)$$

Number of states in which the system can have large disorder hence higher entropy.

(No. of molecule \uparrow) (No. of possible system states \uparrow) \rightarrow (Entropy \uparrow)

//

Question 16.1: Show that entropy of ideal gas depends on initial and final state of temperature and volume.

$$\Rightarrow ds = \oint \frac{dq}{T} = C_v \oint d(\ln T) + R \oint d(\ln \alpha)$$

$$\Delta S = \int_i^f ds = S_f - S_i = C_v \ln \left(\frac{T_f}{T_i} \right) + R \ln \left(\frac{\alpha_f}{\alpha_i} \right)$$

$$S_f = S_i + C_v \ln \left(\frac{T_f}{T_i} \right) + R \ln \left(\frac{\alpha_f}{\alpha_i} \right)$$

$$S_f = S_i + C_v \ln \left(\frac{T_f}{T_i} \right) + C_v \left(\frac{R}{C_v} \right) \ln \left(\frac{\alpha_f}{\alpha_i} \right)$$

$$S_f = S_i + C_v \left[\ln \left(\frac{T_f}{T_i} \right) \times \left(\frac{\alpha_f}{\alpha_i} \right)^{\left(\frac{C_p - C_v}{C_v} \right)} \right]$$

$$S_f = S_i + C_v \left[\ln \left(\frac{T_f}{T_i} \right) \times \left(\frac{\alpha_f}{\alpha_i} \right)^{(\gamma-1)} \right] \text{ where } \gamma = C_p/C_v$$

17. Lecture 17 10/10/2024

17.1 . Entropy

In last lecture we derived expression of entropy, i.e.

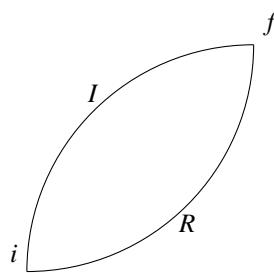
$$\oint \frac{q}{T} \leq 0$$

→ Depends on the final and initial states of temperature and volume.

Let us consider a cyclic process with initial state i and final state f , the path through which process occur be denoted by R and I representing reversible and irreversible processes.

$$i \xrightarrow{R} f \quad \text{Reversible}$$

$$i \xrightarrow{I} f \quad \text{Irreversible}$$



As we know from any cyclic process

$$\oint \frac{q}{T} \leq 0 \quad (361)$$

$$\left[\int_i^f \frac{q}{T} \right]_R + \left[\int_f^i \frac{q}{T} \right]_I \leq 0 \quad (362)$$

Since,

$$ds = \frac{\delta q}{T} \quad (363)$$

$$S_f - S_i + \left[\int_f^i \frac{q}{T} \right]_I \leq 0 \quad (364)$$

$$S_f - S_i \geq \left[\int_f^i \frac{q}{T} \right]_I \quad (365)$$

$$ds \geq \left[\int_f^i \frac{q}{T} \right]_I \quad (366)$$

$$Tds \geq \delta q \quad (367)$$

It indicates that the upperbound of the heat abosorbed by the system during a given changes.

For an isolated system,

$$\delta q = 0 \quad (368)$$

$$S_f - S_i \geq 0 \quad (369)$$

$$S_f \geq S_i \quad (370)$$

For a spontaneous irreversible transformation, occurring in an isolated system, the final entropy is greater then initial entropy.

17.2 . 2nd law of themodynamics

2nd law of themodynamics can be stated as:

1. For reversible transformation, there is no change in entropy of universe.
2. The entropy of universe increase as a result of irreversible transformation.

$$\Delta S_{\text{universe}} = \Delta S_{\text{system}} + \Delta S_{\text{surrounding}}$$

$$\Delta S_{\text{surrounding}} = 0 \quad \text{For reversible transformation}$$

$$\Delta S_{\text{surrounding}} > 0 \quad \text{For irreversible transformation}$$

//

Question 17.1: Calculate the change in air pressure if the specific entropy decrease by $0.05 \text{ J g}^{-1} \text{ K}^{-1}$ and the ait temperature decreases by 5%.

⇒ To calculate the change in air pressure given the specific entropy decrease and temperature change

Given:

- Change in specific entropy: $ds = -0.05 \text{ J g}^{-1} \text{ K}^{-1}$
- Temperature decrease: $dT = -0.05T$ (a decrease of 5%)

Assuming the process to be reversible

Using the equation:

$$ds = C_p d \ln(T) - R d \ln(P)$$

Substituting values:

$$dT = -0.05T \Rightarrow d \ln(T) = \frac{dT}{T} = -0.05$$

$$ds = C_p (-0.05) - R d \ln(P)$$

$$-0.05 = 1005 (-0.05) - 287 d \ln(P)$$

$$-0.05 = -50.25 - 287 d \ln(P)$$

$$287 d \ln(P) = -50.25 + 0.05$$

$$287 d \ln(P) = -50.20$$

$$d \ln(P) = \frac{-50.20}{287} \approx -0.174$$

Integrating:

$$P_f = P_i e^{-0.174}$$

where P_f is the final pressure and P_i is the initial pressure. $ds = C_p d \ln(T) - R d \ln(P)$

Recall that,

$$\theta = T \left(\frac{1000}{P} \right)^{\frac{R_d}{C_p}} \quad (371)$$

$$\ln \theta = \ln T + \frac{R_d}{C_p} \ln(1000) - \frac{R_d}{C_p} \ln P \quad (372)$$

$$C_p d(\ln \theta) = C_p d(\ln T) - R_d d(\ln P) \quad (373)$$

From 1st law of thermodynamic,

$$\delta q = C_p dT - \alpha dP \quad (374)$$

$$ds = \frac{C_p dT}{T} - \frac{\alpha dP}{T} \quad (375)$$

$$ds = \frac{C_p dT}{T} - \frac{R_d dP}{P} \quad (376)$$

$$ds = C_p d \ln(T) - R_d d \ln(P) \quad (377)$$

Comparing Eq(373) and Eq(377) we can find that,

$$ds = C_p d(\ln \theta) \quad (378)$$

$$S = C_p \ln \theta + \text{const.} \quad (379)$$

Therefore, lines of constant potential temperature are lines of constant entropy, but not if process becomes irreversible. Specific entropy is given by logarithm of potential temperature, when θ remains constant \Rightarrow entropy remains constant. For irreversible transformation

$$ds > 0 \quad d\theta = 0$$

All isentropic process are adiabatic but all adiabatic process are not isentropic.

//

Question 17.2: During a process a parcel of dry air decent from 900 to 950hPa and if specific entropy decreases by $30Jkg^{-1}K^{-1}$. If it's initial temperature is 273K; What is it's final temperature and potential temperature?

\Rightarrow The relationship between specific entropy S, temperature T, and pressure P for an ideal gas assuming the process to be reversible can be expressed as:

$$ds = C_p d \ln(T) - R d \ln(P)$$

Given:

$$P_i = 900 \text{ hPa},$$

$$P_f = 950 \text{ hPa},$$

$$T_i = 273 \text{ K},$$

Substituting values:

$$-30 = 1005 \ln \left(\frac{T_f}{273} \right) - 287 \ln \left(\frac{950}{900} \right)$$

$$-14.4827 = 1005 \ln \left(\frac{T_f}{273} \right)$$

$$T_f = 273 \times e^{-0.0144}$$

$$T_f = 269.094 \text{ K}$$

Potential Temperature:

$$\theta = T_f \left(\frac{P_0}{P_f} \right)^{R/C_p}$$

Substituting the values:

$$\theta = 269.7 \times \left(\frac{950}{900} \right)^{0.2855} \approx 273.6 \text{ K}$$

\therefore The final temperature is 269K and potential temperature is 273.6K

18. Lecture 18 10/10/2024

18.1 . Special Forms of 2nd Law of Thermodynamics

1. For finite isothermal transformation

$$\Delta U = 0 \quad (380)$$

$$\Delta S \geq \int \frac{\delta q}{T} \quad (381)$$

$$\Delta S \geq \frac{1}{T} \int \delta q \quad (382)$$

$$\Delta S \geq \frac{q}{T} \quad (383)$$

$$\Delta S \geq \frac{w}{T} \quad (384)$$

2. For adiabatic transformation

$$\Delta S \geq 0 \quad (385)$$

3. For Isochoric transformation

$$\Delta S \geq C_v \frac{dT}{T} \quad (386)$$

$$\Delta S \geq C_v \ln\left(\frac{T_f}{T_i}\right) \quad (387)$$

4. For Isobaric transformation

$$\Delta S \geq C_p \frac{dT}{T} \quad (388)$$

$$\Delta S \geq C_p \ln\left(\frac{T_f}{T_i}\right) \quad (389)$$

5. Combination of 1st and 2nd law

$$\delta q = C_p dT - \alpha dp \quad (390)$$

$$ds \geq \frac{q}{T} \quad (391)$$

$$Tds \geq C_p dT - \alpha dp \quad (392)$$

$$Tds \geq dh - \alpha dp \quad (393)$$

$$(394)$$

similiarly,

$$\delta q = C_v dT - \alpha dp \quad (395)$$

$$ds \geq \frac{q}{T} \quad (396)$$

$$Tds \geq C_v dT - \alpha dp \quad (397)$$

$$Tds \geq du - \alpha dp \quad (398)$$

18.2 . Moist Processes

There are three processes in moist gases:

- Saturation
- Sub-saturation
- Supersaturation

Evaporation: Some water molecules have sufficient kinetic energy to break free from the intermolecular forces of attraction.

Condensation: When water vapor cools down, molecules lose kinetic energy and form liquid droplets due to intermolecular attractions.

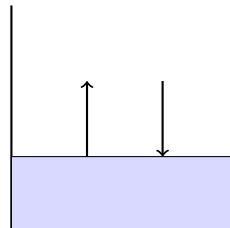


Figure 13. Saturated state

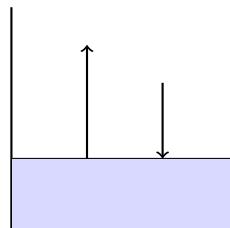


Figure 14. Sub-Saturated state

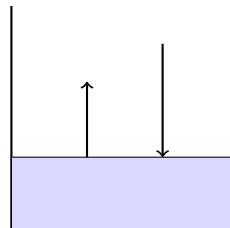


Figure 15. Super-Saturated state

For saturated state rate of evaporation is equals to rate of condensation.

Saturation vapor pressure(e_s) depends only on temperature.

Boiling Point: The temperature at which vapor pressure is equal to atmospheric temperature at the pressure of 1013hPa.

18.3 . Relative humidity

Relative humidity is defined as ratio of vapor pressure to saturation vapor pressure.

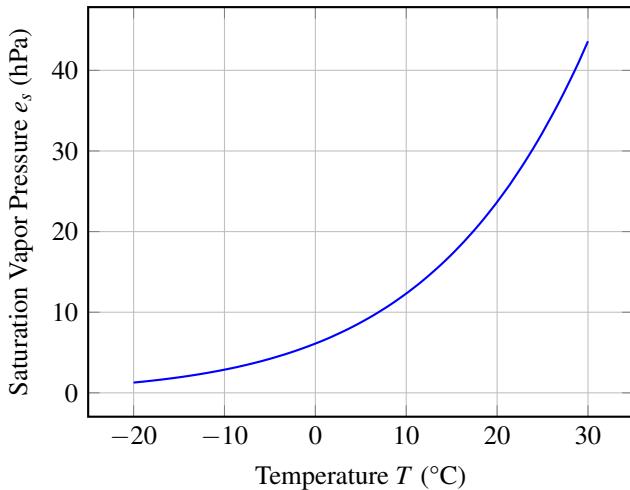
$$\text{Relative humidity (RH)} = \frac{e}{e_s} \times 100\% \quad (399)$$

$$e < e_s(T) \rightarrow \text{Sub-Saturation} \quad (400)$$

$$e = e_s(T) \rightarrow \text{Saturation} \quad (401)$$

$$e > e_s(T) \rightarrow \text{Super-Saturation} \quad (402)$$

Note: Specific humidity is associated with mass

**Figure 16.** Saturation Vapor Pressure e_s vs Temperature T

18.4 . Dew point temperature

There are 2 ways to make sub-saturated to saturated:

1. Reduce temperature
2. Add moisture, so that vapor pressure increase.

If moisture amount remains constant and temperature is reduced, thus saturation is achieved and this temperature is called **Dew point temperature**.

18.5 . Latent heat

Total energy required to convert unit mass from one phase to another.

Specific enthalpy of phase change.

Latent heat of water at STP is $\approx 10^6 \text{ J/Kg}$ (Depends on temperature)

Latent heat of evaporation

Latent heat of evaporation of water at:

- $T = -40^\circ\text{C}$, Latent heat = $2.6 \times 10^6 \text{ J/Kg}$
- $T = 0^\circ\text{C}$, Latent heat = $2.5 \times 10^6 \text{ J/Kg}$
- $T = 40^\circ\text{C}$, Latent heat = $2.26 \times 10^6 \text{ J/Kg}$

Latent heat of fusion

Latent heat of fusion of water at STP = $3.3 \times 10^6 \text{ J/Kg}$

Latent heat of sublimation

Latent heat of sublimation of water at STP = $2.83 \times 10^6 \text{ J/Kg}$

Question 18.1: On a winter day the outside air have temperature of -15°C and relative humidity of 70%.

- a. If the outside air is brought inside and heated to room temperature of 20°C without adding moisture. What is new relative humidity?
- b. If the room volume is 60m^3 then what mass of water must be added to the air by the humidifier to raise the relative humidity to 40%?
- c. How heating is needed to accomplish a. and b.?

\Rightarrow a. The saturation vapor pressure at -15°C can be found using lookup tables.

For $T = 15^\circ\text{C}$, the saturation vapor pressure,

$$e_s(-15^\circ\text{C}) \approx 1.93 \text{ hPa}$$

Given the relative humidity is 70%, the actual vapor pressure at 15°C is:

$$e = \frac{70}{100} \times 1.93 \text{ hPa} \approx 1.35 \text{ hPa}$$

For $T = 20^\circ\text{C}$, the saturation vapor pressure,

$$e_s(20^\circ\text{C}) \approx 23.37 \text{ hPa}$$

Using the actual vapor pressure calculated earlier $e = 1.35 \text{ hPa}$ and the new saturation vapor pressure $e_s(20^\circ\text{C})$:

$$\text{RH at } 20^\circ\text{C} = \frac{e}{e_s} \times 100\% = \left(\frac{1.35 \text{ hPa}}{23.37 \text{ hPa}} \right) \times 100\% \approx 5.8\%$$

\Rightarrow b. To find mass of water needed to raise RH to 40%

$$e = \frac{40}{100} \times 23.37 \text{ hPa} = 9.35 \text{ hPa}$$

The amount of water vapor needed to raise the humidity is proportional to the difference in vapor pressure:

$$\Delta e = 9.35 \text{ hPa} - 1.35 \text{ hPa} = 8.00 \text{ hPa}$$

The mass of water vapor can be calculated using the ideal gas law:

$$m = \frac{\Delta e V}{R_v T}$$

$$m = \frac{800 \text{ Pa} \times 60 \text{ m}^3}{461.5 \text{ J kg}^{-1} \text{ K}^{-1} \times 293 \text{ K}} \approx 0.35 \text{ kg}$$

\Rightarrow c. Heating Needed to Accomplish (a) and (b)

$$m_{\text{air}} = \frac{95000 \text{ Pa} \times 60 \text{ m}^3}{287 \text{ J kg}^{-1} \text{ K}^{-1} \times 258 \text{ K}} \approx 77.7 \text{ kg}$$

$$Q_{\text{air}} = 77.7 \times 1005 \times 35 \approx 2.73 \times 10^6 \text{ J}$$

$$Q_{\text{water}} = 0.35 \times 2.5 \times 10^6 = 0.875 \times 10^6 \text{ J}$$

$$Q_{\text{total}} = 2.73 \times 10^6 + 0.875 \times 10^6 \approx 3.61 \times 10^6 \text{ J}$$