

KQS COACHING CENTER

| Date: _____ | | Class: XI | | Paper: Chemistry |

| Time: 60 minutes | | Max. Marks: 15 | | Test # 1 |

| NAME:| _____ | | F.NAME:| _____ |

Q1: Solve the following

i. 15.5m, 651.8cm, and 4291m ii. $\frac{56 \times 725 \times 273}{760 \times 298}$

Q2: Define (i) Molecular Mass (ii) Molecular Formula (iii) Avogadro's number (iv) Mole.

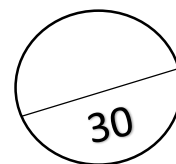
Q3: Calculate the mass of Carbondioxide (CO₂) that can be obtained by heating 30g of limestone (CaCO₃) and also calculate the mass of Calcium Oxide (CaO). $\text{CaCO}_3 \longrightarrow \text{CaO} + \text{CO}_2$

Q4: Calculate the volume of CO₂ gas produced at Standard Temperature and pressure by the combustion of 40g of CH₄
 $\text{CH}_4 + 2\text{O}_2 \longrightarrow \text{CO}_2 + 2\text{H}_2\text{O}$

Q5: What volume of O₂ at S.T.P is required to burn 500dm³ of Ethene C₂H₂ gas ? What volume of CO₂ will be formed?
 $\text{C}_2\text{H}_4 + 3\text{O}_2 \longrightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$

Q6: i. Calculate the number of atoms in 9.2g of Na (Na = 23 a.m.u.).

ii. Calculate the mass in grams of 3.01x10²⁰ molecules of glucose (C₆ H₁₂ O₆).



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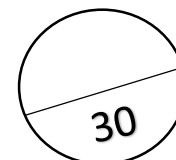
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