KQS COACHING CENTER

|Date: _____| |Class: XI | | Paper: Chemistry | | Time: 60 minutes | |Max. Marks: 15 | |Test # 1|

|NAME:| _____ | |F.NAME:| _____|

Q1: Solve the following

i. 15.5m, 651.8cm, and 4291m ii. $\frac{56\ x\ 725\ x\ 273}{760\ x\ 298}$

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Q2: Define (i) Molecular Mass (ii) Molecular Formula (iii) Avogadro's number (iv) Mole.

Q3: Calculate the mass of Carbondioxide (CO_2) that can be obtained by heating 30g of limestone ($CaCO_3$) and also calculate the mass of Calcium Oxide (CaO) . $CaCO_3$ - - - - - - - - - - > CaO + CO_2

Q4: Calculate the volume of CO_2 gas produced at Standard Temperature and pressure by the combustion of 40g of CH_4

$CH_4 + 2O_2$ - - - - - - - > $CO_2 + 2H_2O$

Q5: What volume of O_2 at S.T.P is required to burn 500dm^3 of Ethene C_2H_2 gas ? What volume of CO_2 will be formed?

$C_2H_4 + 3O_2$ - - - - - - - - > $2\ CO_2 + 2H_2O$

Q6: i. Calculate the number of atoms in 9.2g of Na (Na = 23 a.m.u.).

ii. Calculate the mass in grams of 3.01×10^{20} molecules of glucose ($C_6H_{12}O_6$).

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