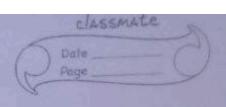


Experiment - 2(a)

- · Aim: Performing and observing the action of exater on quick lime and classifying the reaction
- · Material required: Calcium oxide, water, broker, glass rod, dropper, red litmus paper, testuble, filter paper, bunnel
- · Exchemment:
- Take some calcium exide in beaker a pour some water over quickline slowly stir it with a clean rod & touch the outer surface of the beaker
- >> Take a clear dropper and with its help

 put a drop of liquid from the

 beaker on a red lithus paper
- 37 filter the mixture of beaker and take about 5ml of filtrate in a text tube Blow air through the liquid.
- · Observation:
- hissing sound outer surface of the beaker is not



>> Red litmus paper turns blue

>> on passing to, through the clear filtrate it turns milky.

· Interface:

new substance. As heat is evolved, the reaction of water & quickline is excluded.

>> A new substance bosmed by the reaction
>> ab water & quickline is basic in noture

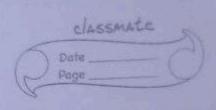
The new substance formed by the reaction by the of quick lime & water is calcium hydroxide.

· Result:

Two compounds, viz, quick lime a water combines to form calcium hydroxide Therefore, this operation is an example of combination reaction.

· Precautions ?

1) The filtrate collected should be clear
2) Quick time can cause severe curas therefore
1+ should be handled with spatula.

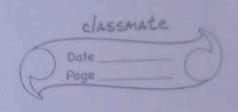


Experiment - 2(6)

- heat on gerrous sulphate crystals & classifying the reaction
- · material required: Ferrous sulphate crystals, test tubes, test tube holder, when litmus paper
- · Eachegiment;
- "> Take about 29 conjustais of gerrous

 Sulphate in a day test tube & note

 the colour of conjustais.
- Heat the test tube containing ferrous
 Sulphate
- : notemation?
- I the crystal are goven in coloux
- The colour of crystals changes to brown a colourless gas with a smell of burning sulphur is evolved
- · guterlace:
- " The colour of ferrous sulphate is green we substance are formed by the



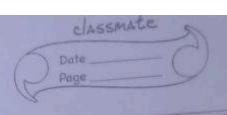
Heating of Berron rolphate

· Result:

on heating, ferrous sulphate decomposes
to give ferric oxide, sulphur dioxide
and sulphur trioxide. This is a decomposes
ses reaction

Precaution:

while heating ferrous sulphate, keep the mouth of test tube away from your goors classimates



Experiment - 2(c)

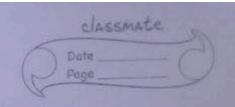
- · Aim? Performing and observing the reaction of iron nails kept in copper sulphate solution and classifying the reaction
- · Material Required: aron nails, copper sulphate solution, test tube, test tube stand, sand paper and thread
- · Procedure:
- >> Take 2 iron nails and clean them by subbing with sandpaper
- >> Take two test tubes and mark them as
- opper sulphate solution
- The one ison hail with a thread and immesse

 this corefully in the copper sulphate

 solution in test tube A for about

 20 min. Keep aside one hail for

 compassion
- After 20min take out the 1800 pail
 grow the of copper sulphate solution
 " compare the intensity of flue cobour
 of copper sulphate solution of both the



test tube 'A' and 'B'

>> Also compare the colour of 1800 nail

dipped in copper sulphate solution with one

kept aside

· Observation:

- The initial colour of copper sulphate

 solution was Glue which after immersing

 iron nails turns to light green the initial

 colour of iron nail was grey. After

 immersing the nail into copper sulphate

 solution, a brown coating developed over

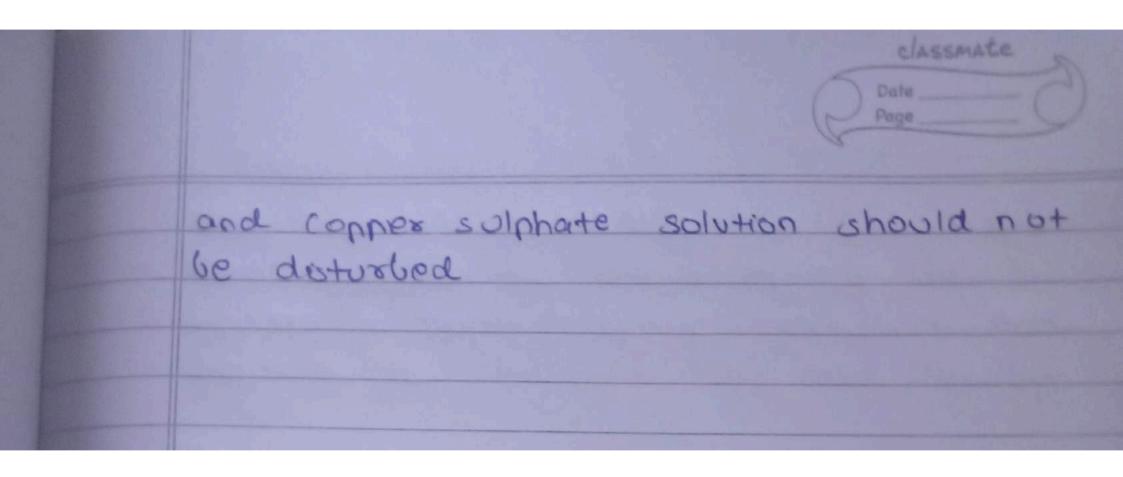
 it because of this coating iron nail

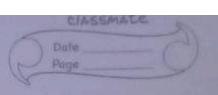
 appears reddish brown.
- Result: On keeping the ison nails in copper sulphate solutions, a displacement reaction takes place. In this displacement reaction, inon displaces copper and two new products between sulphate and copper are somed

· Precautions:

"> The time noils should be cleaned by subbing them a with a sandpaper

The to test tube containing won nails





Experiment (2d)

- Aim & Performing and observing the seartion between sodium sulphote and barium chloride solution and classify their reaction
- · Material required: Sodium sulphate solution barium chloride solution, test tubes beaker
- · Procedure :
- >> Take Sml of rodium sulphate solution in a test tube and mark it as 'A'
- another the test the and mark it as
- Mix the solutions of test tube "A" and
- stir the mixture kept in beaker
- A Becord your observation
- sodium sulphate and barrium chloride, a with precipitate is formed

Result: On mixture the solutions ob sodium sulphate and barium chloride a double duplacement reaction takes on this reaction sodium sulphate and barium chloride exchange ions and now products barium sodi sulphate Cubite ppt) and sodium chloride are gormed

Precautions :

chould be cleaned

& Equal volumes of sodium sulphate and barium Chloride solutions should be used