Chemistry 4-1 - definitions

Main points of that video:

* We won’t be dealing with ionic bonds like the one in sodium chloride
* We won’t be dealing with ionic bonds that involve electron theft
* We won’t be dealing with ionic bonds that involve differences in electronegativity
* We will be dealing with ionic bonds between full positive and full negative charges
* We won’t be dealing with ionic bonds with partial charges
* Unlike charges attract
* Solid lines indicate covalent bonds only
* Dotted, squiggly, other than solid lines indicate non-covalent bonds
* Covalent bonds hold the parts of a molecule together
* Van Der Waals bonds are extremely weak
* VDW are the weakest bonds
* VDW bonds can form between any two atoms
* VDW bonds are formed by transient polarities in atoms that allow them to attract each other
* VDW are always possible but usually only important when no other bonds possible
* VDW are the only forces involved in hydrogen being a liquid