

## Latent Heat



• (Latent) Heat of fusion – the heat energy needed to melt

(Solid→Liquid) or freeze

(liquid  $\rightarrow$  solid) one gram (or

kg) of a substance.

For water: Hf = 334,000 J/kg or

80 cal/g

(Latent) Heat of vaporization

– the heat energy needed to

vaporize (liquid→gas) or

condense (gas→liquid) one

gram (or kg) of a substance.

For water:  $Hv = 2.26 \times 105$ 

J/kg or 540 cal/g





## The Transfer of Energy as Heat

• Energy moves heat in three ways

- Conduction
- Convection
  Convection is the process of heat transfer by the bulk movement of molecules within fluids such as gases and liquids
- Radiation





The process that moves energy from one object to another when they are touching physically.

- Conductors: materials that transfer energy easily.
- Insulators: materials that do not transfer energy easily.
- Examples: hot cup of cocoa transfers heat energy to cold hands



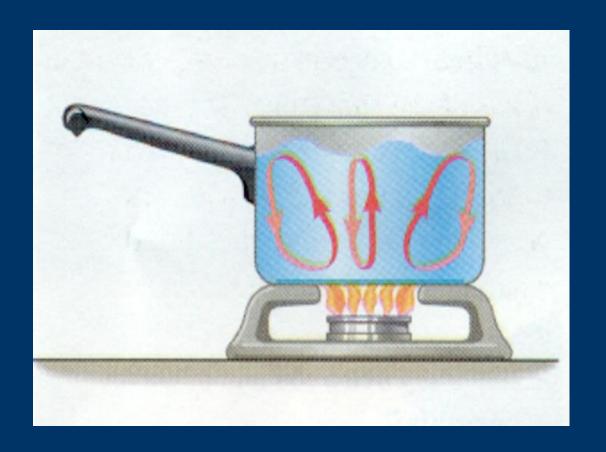
## **Convection-**



The process that transfers energy by the movement of large numbers of particles in the same direction within a liquid or gas.

Cycle in Nature

Boiling water and heating a room

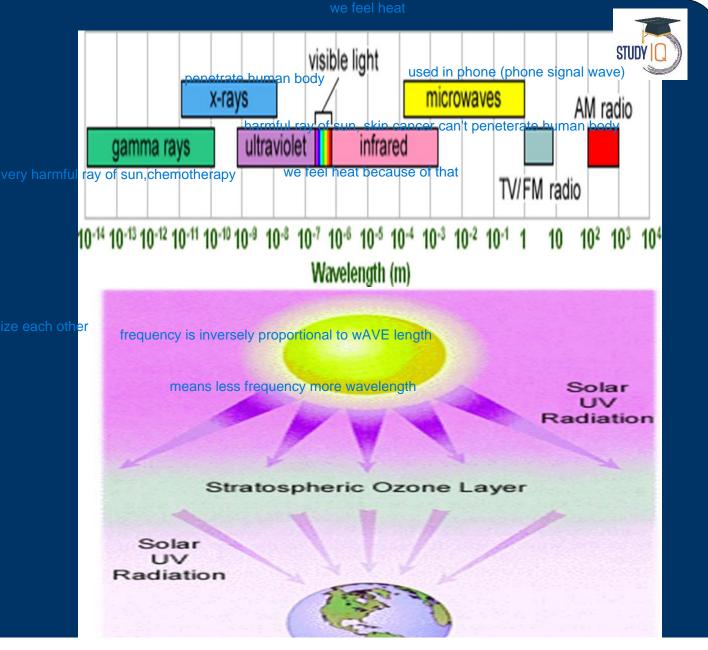


## Radiation-

• The energy that travels by electromagnetic waves (visible light, microwaves, and infrared light)

when electric field meet to magnetic field it not divert it neturalize each other

 Radiation from the sun strikes the atoms in your body and transfers energy







Newtons cooling law-The rate of loss of heat is directly proportional temperature difference between body and atmosphere.





Zeroth law of thermodynamics —If a body A is in thermal equilibrium with body B and body B is in thermal equilibrium with body C then Body A and C will also be in thermal equilibrium.

First law of thermodynamics- Energy neither can be created nor be destroyed it can be transfer from one body to another body.

$$Q = U + W$$





Ask Doubts at : doubts.studyiq@gmail.com

Contact: 9580048004, 7291059476

