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Operando NRIXS and XAFS Investigation of Segregation Phenomena in Fe-Cu and Fe-Ag Nanoparticle Catalysts during CO₂ Electroreduction

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Abstract

Operando nuclear resonant inelastic X-ray scattering (NRIXS) and X-ray absorption fine-structure spectroscopy (XAFS) measurements have been used to gain insight into the structure and surface composition of FeCu and FeAg nanoparticles (NPs) during the CO₂ electrochemical reduction (CO₂RR) and to extract correlations with their catalytic activity and selectivity. The formation of a core-shell structure during CO₂RR for FeAg NPs was inferred from the analysis of the *operando* NRIXS data (phonon density of states, PDOS) and XAFS measurements. Electrochemical analysis of the FeAg NPs revealed a faradaic selectivity of ~ 36 % for CO in 0.1 M KHCO₃ at -1.1 V vs RHE, similar to that of pure Ag NPs. In contrast, a predominant selectivity towards H₂ evolution is obtained in the case of the FeCu NPs, analogous to the results obtained for pure Fe NPs, although small Cu NPs have also been shown to favor H₂ production. These results, together with the findings from NRIXS and XAFS, revealed that Ag is preferentially present on the surface of the FeAg NPs during CO₂RR, while Cu is also likely to be mainly distributed on the surface of the FeCu NPs but with a larger degree of intermixing with Fe.

1. Introduction

One of the most significant challenges in the field of catalysis is the development and optimization of experimental methods that allow the observation of a catalyst while at work.^[1-3] This is especially difficult when nanomaterials are considered under realistic conditions, as for example, when trying to understand atomic segregation phenomena and other structural/chemical modifications at liquid/solid interfaces during an electrochemical process.^[4] A number of established methods are not directly applicable under reaction conditions due to decreased electronic mean free paths, scattering or lack of sufficient spatial resolution or chemical sensitivity. The development of additional techniques suitable for *operando* implementation in an electrochemical environment is therefore a subject of high interest.^[5-6]

Nuclear resonant inelastic X-ray scattering (NRIXS) is a synchrotron radiation technique that is sensitive to vibrational states of the nuclei of Mössbauer-active isotopes, most commonly ^{57}Fe .^[7] It can be used to probe vibrational modes of molecular complexes of iron as well as phonons in solid state materials.^[8] Because the vibrational modes of a crystal lattice depend heavily on its structure, it is possible to relate NRIXS data to structural and thermodynamic properties of a material.^[9] Moreover, due to the isotope-specific detection, one can probe only a specific region of a sample if enriched with the Mössbauer-sensitive element, while the overall structure can be inferred indirectly from the contribution of the partial ^{57}Fe -phonon density of states (PDOS).^[10-11] More importantly, this method can be used to investigate a wide range of materials from bulk materials to thin-films and NPs in experiments under extreme conditions such as high-pressure environments or during thermal or electrochemical catalytic reactions.^[12-15]

In this study, the NRIXS method was used to follow the structural and chemical evolution of small ^{57}Fe -containing bimetallic NPs in a liquid environment under an applied external potential during the electrochemical reduction of CO_2 (CO_2RR). X-ray absorption fine-structure spectroscopy (XAFS), another synchrotron technique with high sensitivity to elemental and local structural composition was also applied here under reaction conditions in an electrochemical environment to complement the NRIXS insight, including providing information on the Ag and Cu components of our electrocatalysts.^[16]

We have used iron-based materials in our study due to their applications as heterogeneous catalysts in the field of sustainable energy conversion. Iron is an abundant metal with little

environmental impact, though it is unsuitable for CO₂RR in its bulk form, since it favors the parasitic hydrogen evolution reaction (HER).^[17] Nevertheless, if present in the NP form at the core of nanostructures with a CO₂RR-active thin shell, it could contribute to a decrease in the catalyst price. To date, research on iron-based materials for CO₂RR focuses mainly on molecular complexes, as for example Fe-N-C materials in porphyrin-like structures.^[18-19] Here, we combine iron with copper and silver within a micellar nanoreactor since these elements are promising for the selective conversion of CO₂ to C₁-C₃ products (Cu) and CO (Ag).^[17] In this work, the structure and composition of the FeCu and FeAg catalysts during the reaction will be extracted from a synergistic combination of *operando* NRIXS and XAFS measurements and correlated with the selectivity trends of monometallic Cu, Ag and Fe NPs of similar size.

2. Experimental

Synthesis

Inverse micelle encapsulation was used to prepare colloidal solutions of size-selected, isolated Fe, FeCu, FeAg, Ag and Cu NPs.^[20-21] Two sets of NPs were synthesized by loading a nonpolar/polar diblock copolymer [poly(styrene)-block-poly-(2-vinylpyridine), Polymer Source Inc.] dissolved in toluene with metal salts (⁵⁷FeCl₂ x 2 H₂O, AgNO₃, CuCl₂ x 2H₂O, FeCl₃). One of the NP sets synthesized contained the isotopically enriched ⁵⁷FeCl₂-salt needed for the NRIXS measurements, **Table 1**. The enriched iron salt was synthesized from an iron foil with 95 % ⁵⁷Fe isotopic enrichment by adapting a procedure described in the literature.^[22] A 0.2 metal loading (metal/P2VP molecular weight ratio) was used in the synthesis of the micellar samples. Samples for NRIXS and XAFS measurements (⁵⁷Fe, ⁵⁷FeCu, ⁵⁷FeAg) were prepared by impregnating the NP solution on carbon black powder (5 wt%). A N₂-plasma treatment was used for polymer removal on the impregnated powders (300 mTorr, 20 W, 5 cycles of 10 min duration) as well as on samples dip-coated on silicon wafers used for the AFM analysis (300 mTorr, 20 W, 20 min). The NP powder (5 mg – 10 mg) was then dispersed in an ethanol/naion solution. The former solution was deposited on a low-porosity carbon paper disc (Sigracet® SGL 24 AA, SGL Carbon GmbH) by filtration, during which the catalyst powder stays on one side, bonded by naion, while the other side is not modified. These carbon paper samples were used in order to minimize the signal attenuation for *operando* NRIXS and XAFS measurements, since our cell is designed such that the sample is irradiated from the back, and the subsequent fluorescence signal is also collected

from the back. Finally, Fe, FeCu, FeAg, Ag and Cu samples were prepared for electrochemical characterization in our laboratory by two successive cycles of dip-coating the bimetallic NP solution (nafion-free) on fresh glassy carbon electrodes and subsequent a N₂ plasma treatment (300 mTorr, 20 W, 2x20 min) was employed for polymer removal.

Sample Characterization

Tapping Mode atomic force microscopy (AFM) (Bruker, Multimode 8) images were acquired on samples supported on SiO₂/Si(100) after polymer removal. The apparent NP height was used to determine the NP size and distribution. We used the Gwyddion software package to analyze the height maps.^[23] STEM-EDX measurements of supported micellar NPs, dispersed in ethanol and drop-casted on 5nm amorphous Si TEM windows, were carried out in a JEOL ARM 200F microscope. We employed our XPS setup (SPECS GmbH, a non-monochromatic source, Al K-edge @ 1486.6 eV) to determine the different metal ratios in the bimetallic NPs.

Electrochemistry

The laboratory-based CO₂RR characterization was carried out in a gas-tight H-type cell. A Selemion ion exchange membrane (AGC Engineering Co., Ltd.) separated the two compartments, which were filled with 40 mL 0.1 M KHCO₃ (Sigma Aldrich, 99.7 %). A three-electrode setup with a platinum mesh counter electrode (MaTeck, 3600 cm²), leak-free Ag/AgCl/3.4 M KCl reference (Innovative Instruments Inc.) and the glassy carbon sample acting as the working electrode (4.4 cm² exposed area) were employed. A CO₂ flow of 20 mL/min was used during the measurements and for purging. The potential was controlled by a Metrohm-Autolab (M204) potentiostat with corrected iR drop. Gas analysis was done online by a Shimadzu 2014 Gas-Chromatograph with HayesSepQ + HayeSepR packed columns and a flame ionization detector (FID) for hydrocarbon separation and detection. An MS-13X column and thermal conductivity detector (TCD) were used for the detection of H₂, N₂ and O₂, while CO and CO₂ were methanized before detection by the FID. Liquid products such as formate were analyzed with a high-performance liquid chromatograph (Shimadzu Prominence HPLC) with a NUCLEOGEL SUGAR 810 column and a refractive index detector (RID). Alcohols were analyzed with a liquid GC (Shimadzu 2010 plus) with silica column and FID. Product selectivities were calculated with product distribution and current data obtained after 1h of reaction.

Synchrotron Measurements

Operando XAFS and NRIXS measurements were performed in a home-made electrochemical cell (see Suppl. Inf. for details, **Fig. S1**). During both measurements, CO₂ was dosed to the electrolyte (0.1 M KHCO₃) solution (20 mL/min) and a potential of -1.1 V vs RHE was applied. The same three-electrode setup as described above was employed. A potentiostat EmStat 3 from PalmSense was used to control the applied potential. The XAFS experiments were performed at the SAMBA beamline at the SOLEIL synchrotron. Data at the K-edges of Cu, Fe and Ag were collected, using a Si(220) monochromator for energy selection. The measurements were performed in fluorescence mode using a 13-channel Ge detector. The NRIXS experiments were performed at the 3-ID-B beamline at the Advanced Photon Source (APS) of the Argonne National Laboratory. The beamline features a monochromatic beam at 14.41 keV with an energy resolution of 1 meV. Analysis of the NRIXS data was done with the PHOENIX software from W. Sturhahn.^[24] XAFS data were processed and analyzed with the ATHENA and ARTEMIS modules of the Demeter software package, using a FEFF6 code.^[25]

Results and Discussion

AFM images of the ⁵⁷Fe, ⁵⁷FeCu and ⁵⁷FeAg NPs deposited on SiO₂/Si(100) are shown in **Fig. 1**. Additional images and height histograms can be found in the suppl. documents, **Figs. S2, S3**. Well dispersed, size-selected NPs with average size (NP height) under 8 nm were obtained with the micellar synthesis.

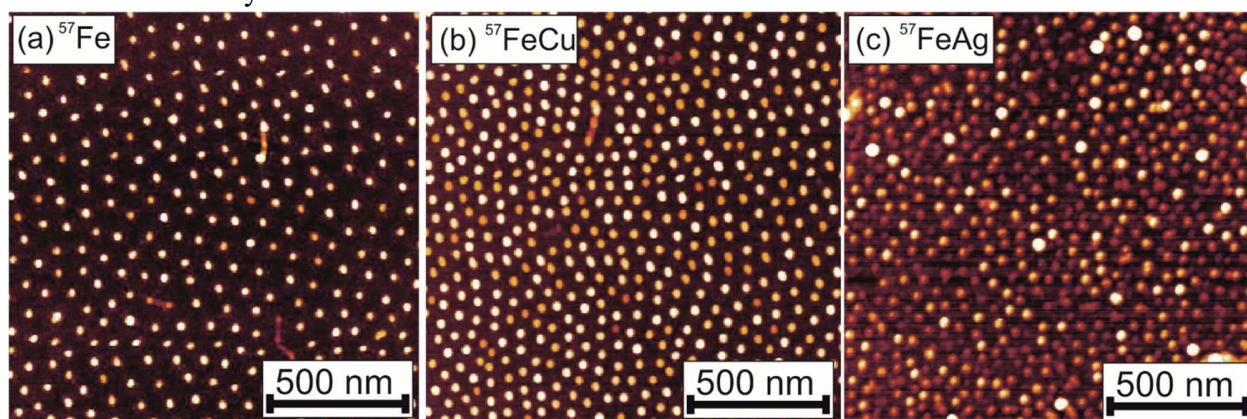


Figure 1 Tapping mode AFM images of: (a) ⁵⁷Fe, (b) ⁵⁷FeCu, (c) ⁵⁷FeAg NPs deposited on SiO₂/Si(100) after a N₂ plasma treatment for polymer removal. Average particle sizes are: 7.8 nm for ⁵⁷Fe, 5.2 nm for ⁵⁷FeCu and 4.2 nm for ⁵⁷FeAg.

STEM-EDX measurements show that both metals are indeed present in our NPs, **Fig. S4**.

XPS spectra of the Fe-2p region of the as-prepared bimetallic NPs deposited on a SiO₂/Si(100) substrate and the corresponding fits displaying the different Fe oxidation states are shown in **Fig. 2**. Additional XPS data from the Cu-2p and Ag-3d regions are shown in **Fig. S5**. It is evident that the iron in the NPs after synthesis, N₂-plasma treatment and subsequent air exposure is almost completely cationic (Fe²⁺ or Fe³⁺, see **Table S1**). A strong Fe-2p satellite is apparent in the FeCu spectrum, which has previously been reported for oxidized iron in the presence of copper.^[26] The Fe:Cu ratio was calculated as 55:45. The Ag-3s peak in FeAg overlaps with the Fe-2p region, which we compensated for, when we calculated the Fe:Ag ratio of 64:36. Due to the small size of the nanoparticles, we expect the entire particle volume to be probed by XPS.

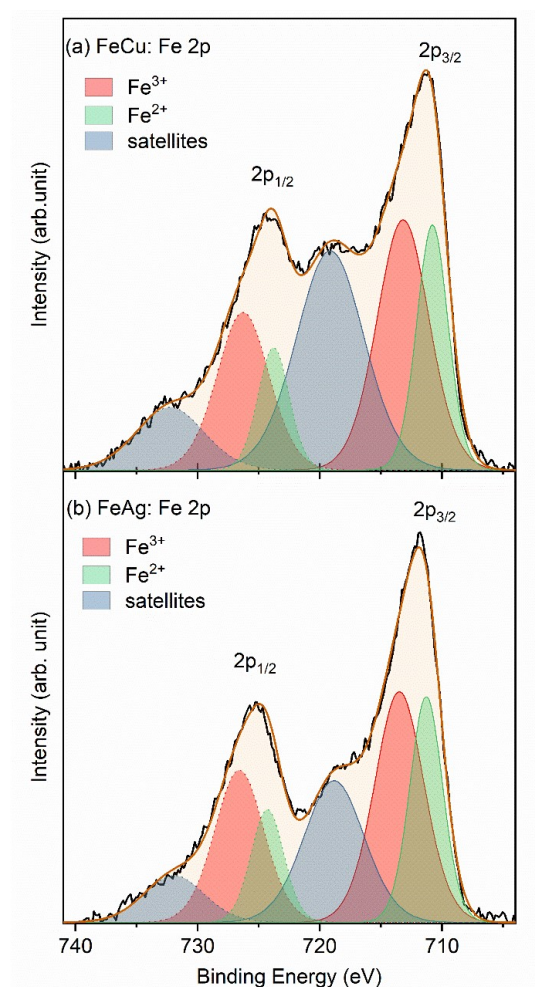


Figure 2: Background-subtracted XPS spectra and corresponding fits of the Fe-2p core level region of as-prepared (a) FeCu and (b) FeAg NPs supported on SiO₂/Si(100).

The product selectivities and current densities of our Fe, FeCu, FeAg, Ag and Cu NP samples are shown in **Fig. 3**. While hydrogen is the main product for all samples, the selectivities for CO of the FeAg and Ag NPs are similar (36 % and 39 %, respectively) at -1.1 V vs RHE. This similarity is consistent with a core-shell rearrangement of the FeAg NPs under reaction conditions, with Ag at the NP surface and Fe at the core. The FeCu NPs also produce CO, but with a lower selectivity of under 2%. In contrast to FeCu, both CO and formic acid selectivity are higher in the pure Cu NP sample with 4.2% and 5.5%, respectively. This is likely indicative of a more Fe-rich surface in the FeCu NPs. As expected, no higher hydrocarbons or alcohols were detected in either of the FeCu or Cu samples, contrary to the case of bulk Cu, due to the enhanced number of low-coordinated surface sites on small NPs that are known to favor the HER.^[27-29] Methane however is a product for all iron-containing samples here, albeit with very low selectivity (< 1.7 %), but not for pure Ag or Cu NPs. Such production might be also affected by the NP size^[30], since bulk iron is not known to yield methane^[17]. The methane production in the case of the FeAg samples might be explained due to the presence of an incomplete (non-uniform) Ag-shell, leaving exposed small areas of the Fe core which might lead to a synergistic interaction between Fe and Ag. We observed only low amounts of C₁ products typical of bulk Cu electrodes (CH₄, CO, HCOO⁻)^[17] for FeCu. The pure iron sample produced overwhelmingly hydrogen as its main product, with only traces of formic acid and methane. Besides CO and hydrogen, the silver NPs were also found to produce formic acid (4.3 % selectivity), similar to the FeAg NPs with about 4.4 %. The current densities normalized to the geometric area of the sample for the Fe, FeCu, FeAg, Ag and Cu NP samples are: 0.71 mA/cm², 0.49 mA/cm², 0.35 mA/cm², 0.79 mA/cm² and 0.99 mA/cm² respectively. A possible explanation for the lower current density, as well as slightly lower CO selectivity of the FeAg sample as compared to pure Ag, is the compressive strain induced by the smaller iron core on the silver overlayer. A shift of the d-band center away from the Fermi level, caused by compressive strain, is expected to influence the CO bonding strength negatively, leading to a more favorable H₂ production.

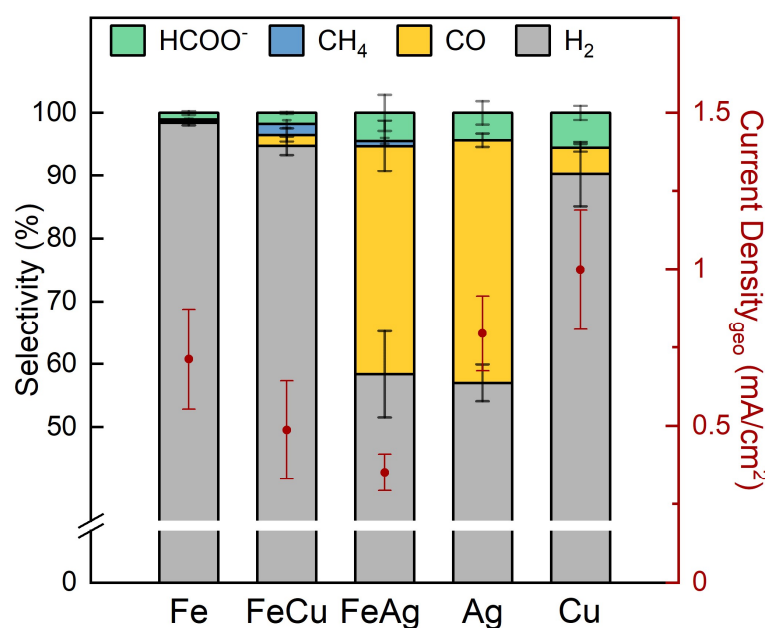


Figure 3 Selectivity for the reaction products of CO₂RR obtained from Fe (6.7 nm), FeCu (5.3 nm), FeAg (4.1 nm), Ag (4.1 nm) and Cu (5.6 nm) NPs after 1h of reaction at -1.1 V vs RHE in CO₂-saturated 0.1 M KHCO₃ and a CO₂ flow of 20 mL/min. Plotted in red are the geometric current densities of each sample (right-hand side red scale).

NRIXS measurements were carried out before (⁵⁷Fe, ⁵⁷FeCu in air) and *operando* under CO₂RR conditions in the electrolyte and under potential control (⁵⁷Fe, ⁵⁷FeCu and ⁵⁷FeAg). **Fig. 4** shows the corresponding Fe-partial PDOS of these samples. The spectrum from a bulk bcc-Fe foil is also shown for reference. The spectra recorded in air indicate enhanced atomic disorder and oxidic Fe features at 41 meV - 43 meV that are at least partially reduced under reaction conditions, Fig. 4(a). The longitudinal acoustic (LA) phonon peak near 36 meV observed in bulk bcc-Fe becomes also visible in the NP phonon spectra under reaction conditions. Nevertheless, it appears shifted to lower phonon energies (by 1.3 meV to 2 meV, Table S2) in both of the bimetallic samples due to the presence of the phononically softer Ag and Cu metals in the Fe environment. Furthermore, significant damping is observed in the LA peak of the three NP samples as compared to that of bulk bcc-Fe due to size-effects.

Interestingly, the PDOS of the ⁵⁷FeAg NP sample shows the most clear resemblance to that of bulk bcc-Fe, since for that sample the LA peak is more prominent and we also see the transverse acoustic (TA) phonon modes, enhanced ordering (as indicated by sharper features), and the absence of oxidic species under the reducing conditions available during CO₂RR. Such

observations point towards a significant reduction of the FeO_x species when the potential is applied in parallel to the segregation of Ag to the NP surface and the formation of a protective Ag shell and a “pure” bcc-Fe NP core.

For the pure ^{57}Fe NPs we still observe hints of the incomplete FeO_x reduction under reaction conditions and enhanced disorder (smeared lines). The $^{57}\text{FeCu}$ NPs look more disordered and the Fe-partial phonon DOS has a larger deviation with respect to that of pure bcc-Fe (**Fig. 4b**), indicating some degree of Fe-Cu intermixing. Nevertheless, a Cu-rich surface is still expected since the PDOS of this sample is still in close agreement with that of pure bcc-Fe aside from the size-dependent phonon damping and enhanced disorder. Such spectra are also seen for thin Fe layers in the proximity of metals with a lower phonon cut-off energy, as is the case here for both, Cu and Ag.^[14] Overall, the PDOS corresponding to the FeAg and FeCu NP samples acquired under CO_2RR approaches a bcc-Fe structure, not fcc-Fe as could be expected if Fe would be embedded in the fcc-Cu or fcc-Ag matrixes.^[31] It should be noted that the relative content of Fe in the as prepared FeCu and FeAg NPs is similar according to XPS, namely 55 % and 64 %.

The Fe/Ag is a layered system known for a low interfacial intermixing. The PDOS of the NPs in this sample displays analogous characteristics to that of a Fe/Ag multilayer interface which was investigated by Roldan *et al.*^[12] **Fig. 4(c)** shows the Fe-partial PDOS data of this FeAg NP sample alongside PDOS data from Ref^[12]. Apart from the overall shape and correspondence to a bcc structure, the position of the LA peak around 35.5 meV and the TA peaks (20 meV - 30 meV) of the FeAg NPs are similar to those of the sandwiched Fe/Ag layers with Fe thicknesses of 2 nm - 4 nm. This is in accordance with the average particle size of the FeAg sample of about 4.2 nm (the real size of the iron part is smaller than this due to the bimetallic composition).

An analogous comparison of the $^{57}\text{FeCu}$ NP sample under operando conditions with respect to Fe(1.5 nm)/Cu(4 nm) multilayers from Roldan *et al.*^[12] is displayed in **Fig. S6**. While the LA peaks of the multilayer sample and the $^{57}\text{FeCu}$ NPs under *operando* conditions are similarly positioned (34.1 meV vs 33.7 meV) and shaped, corresponding to the bcc-Fe structure, there is a lack of unambiguous structural features in the TA region, indicating high disorder due to the small NP size and phonon softening caused by the Fe-Cu interaction.

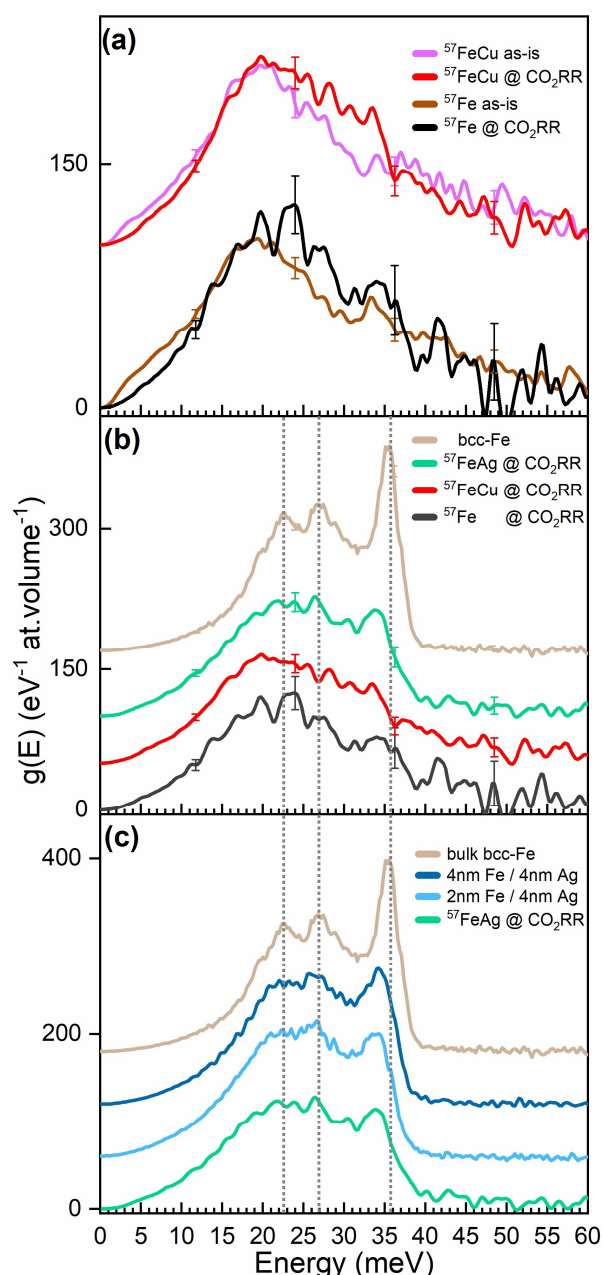


Figure 4 (a) ^{57}Fe -partial PDOS, $g(E)$, obtained from raw NRIXS spectra of ^{57}Fe and $^{57}\text{FeCu}$ NPs in air and under *operando* CO_2RR . The spectra in (b) display the *operando* CO_2RR PDOS data of ^{57}Fe , $^{57}\text{FeCu}$ and $^{57}\text{FeAg}$ NPs together with a bulk bcc-Fe reference foil for comparison.²² (c) ^{57}Fe -partial PDOS of $^{57}\text{FeAg}$ NPs plotted together with thin ^{57}Fe layers deposited on a Ag film (4 nm) reproduced from Ref.^[12] Representative error bars in different regions of the spectra are also shown. The spectra have been vertically offset for better visibility and the vertical lines indicate the position of the two TA and the LA peaks of bulk bcc-Fe.

XAFS data were also acquired under CO_2RR conditions using the same cell design and samples as in the NRIXS experiments in order to extract complementary information about the chemical state and structure of the samples, including their changes under reaction conditions. In this case, we could gain access not only to the Fe-component of the bimetallic systems, but also to the Ag and Cu constituents.

It is evident from the XANES spectra in **Fig. 5 (a)** and **(c)** that the particles are in an oxidized state from the beginning, as they show typical oxidic features. For comparison with our oxidized

particles, we chose here to show the spectrum for an iron mineral (lepidocrocite) that consists of a naturally occurring hydroxide/oxide mixture, since the initial particle composition after air exposure is likely a complex mixture as well. From **Fig. 5 (b) and (d)** we can conclude that the first coordination shell of our oxidized NP samples is similar to that in the reference mineral, but the mean Fe-O distance of $1.81 \pm 0.06 \text{ \AA}$ in our samples is smaller than the reference value of 2.00 \AA .^[32] **Table 1** summarizes further relevant parameters extracted from the EXAFS measurements, both for as-prepared samples and samples under CO₂RR conditions.

Upon applying a potential of -1.1 V vs RHE, the particles start reducing to a more metallic state. However, we still see a significant contribution from oxidized iron species even under *operando* conditions. Some of these cationic species might also be assigned to iron compounds dissolved into the electrolyte.

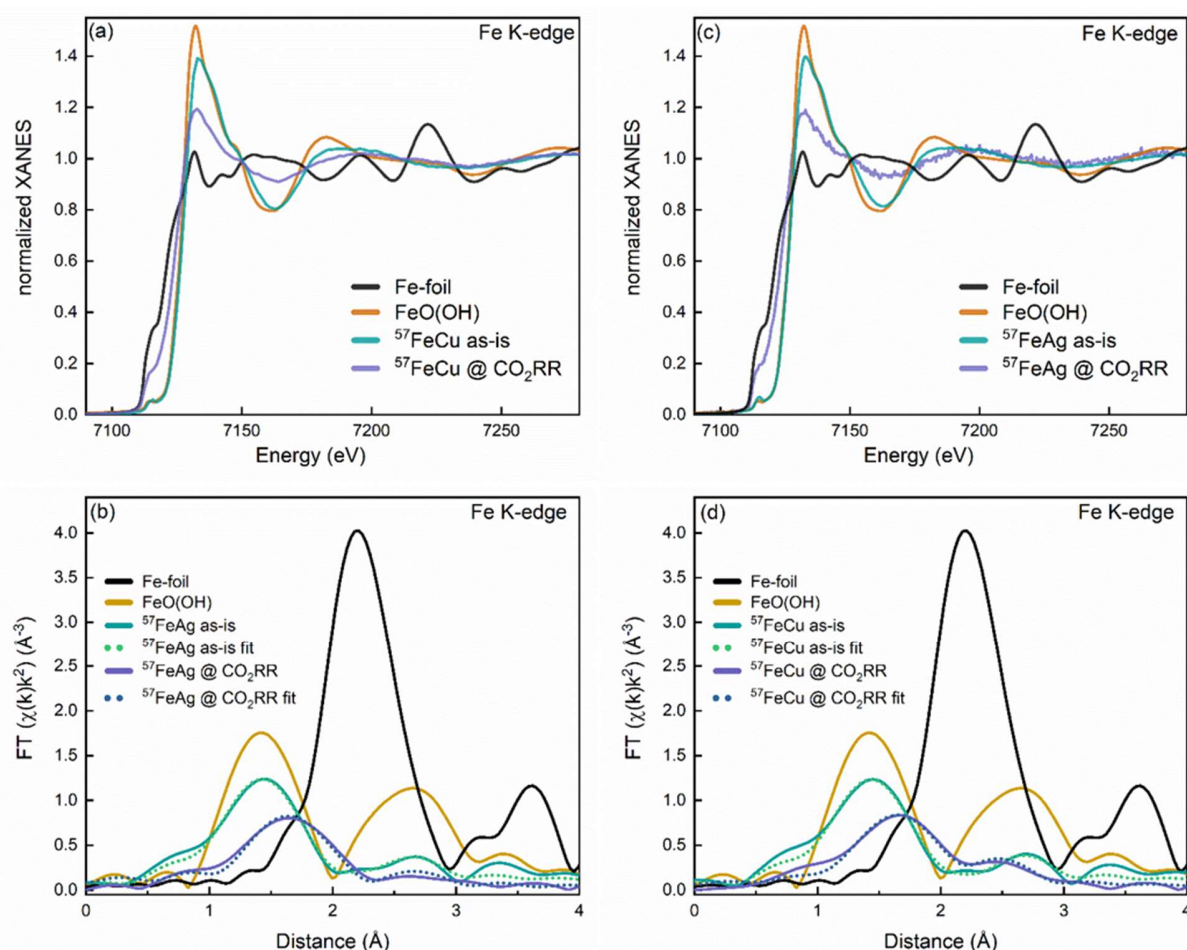


Figure 5 Fe K-edge XANES (a,c) and Fourier-transforms of k^2 -weighted EXAFS data (phase-uncorrected) and the corresponding first-shell fits (dotted lines) (b,d) of ⁵⁷FeCu and ⁵⁷FeAg NPs as prepared (as-is) and under CO₂RR conditions at -1.1 V vs RHE in 0.1 M KHCO_3 after 3.5 h

($^{57}\text{FeCu}$) and 1.5 h ($^{57}\text{FeAg}$). Reference spectra from lepidocrocite $\text{FeO}(\text{OH})$ ^[25] and a bulk iron foil are also shown for comparison.

We can show a reduction and increase in crystalline order in the iron parts of the NP samples under reaction conditions that are in agreement with our previous NRIXS results. The operando XAFS measurements were acquired over periods of 1.5 h - 3.5 h of the reaction, while the NRIXS data were acquired over periods of 11 h – 18 h, which might also explain the higher apparent content of iron oxides in the XAFS data. The long acquisition times were needed due to the low count rates in the NRIXS experiment.

Fig. 6 displays Cu K-edge (a,b) and Ag K-edge (c,d) XANES and EXAFS data of the $^{57}\text{FeCu}$ and $^{57}\text{FeAg}$ NP samples. The presence of oxidic copper compounds similar to CuO was observed, and a reduction of copper to the metallic state under *operando* conditions. The average interatomic distances for Cu-O in the as-prepared sample are in line with bulk CuO (Table 5). The time dependence of the reduction process (**Fig. S7**) reveals a fast initial CuO reduction, followed by a slower increase over the next 1.5 h of the peak in the Fourier-transformed (FT) EXAFS at ca. 2.2 Å that corresponds to a contribution of the metallic Cu phase. For the as-synthesized samples, both XANES and EXAFS data collected at the Ag K-edge and shown in **Fig. 6(c,d)**, indicate the presence of stable AgCl compounds due to residual Cl from the synthesis procedure. Upon applying the CO₂RR potential, a clear increase of the metallic silver contribution (peak in the FT-EXAFS at ca. 2.7 Å) is seen, and no other components could be fitted. There is no contribution from Fe-Ag scattering paths, and the Fe-Fe and Ag-Ag distances do not hint towards intermixing of Fe and Ag either, since they are in line with those for the corresponding bulk materials. In the case of the $^{57}\text{FeCu}$ sample, the Cu-Cu distance agree with those in bulk Cu, while the Fe-Fe distance of 2.48 ± 0.05 Å is in line with that of bulk Fe at 2.47 ± 0.01 Å.^[33]

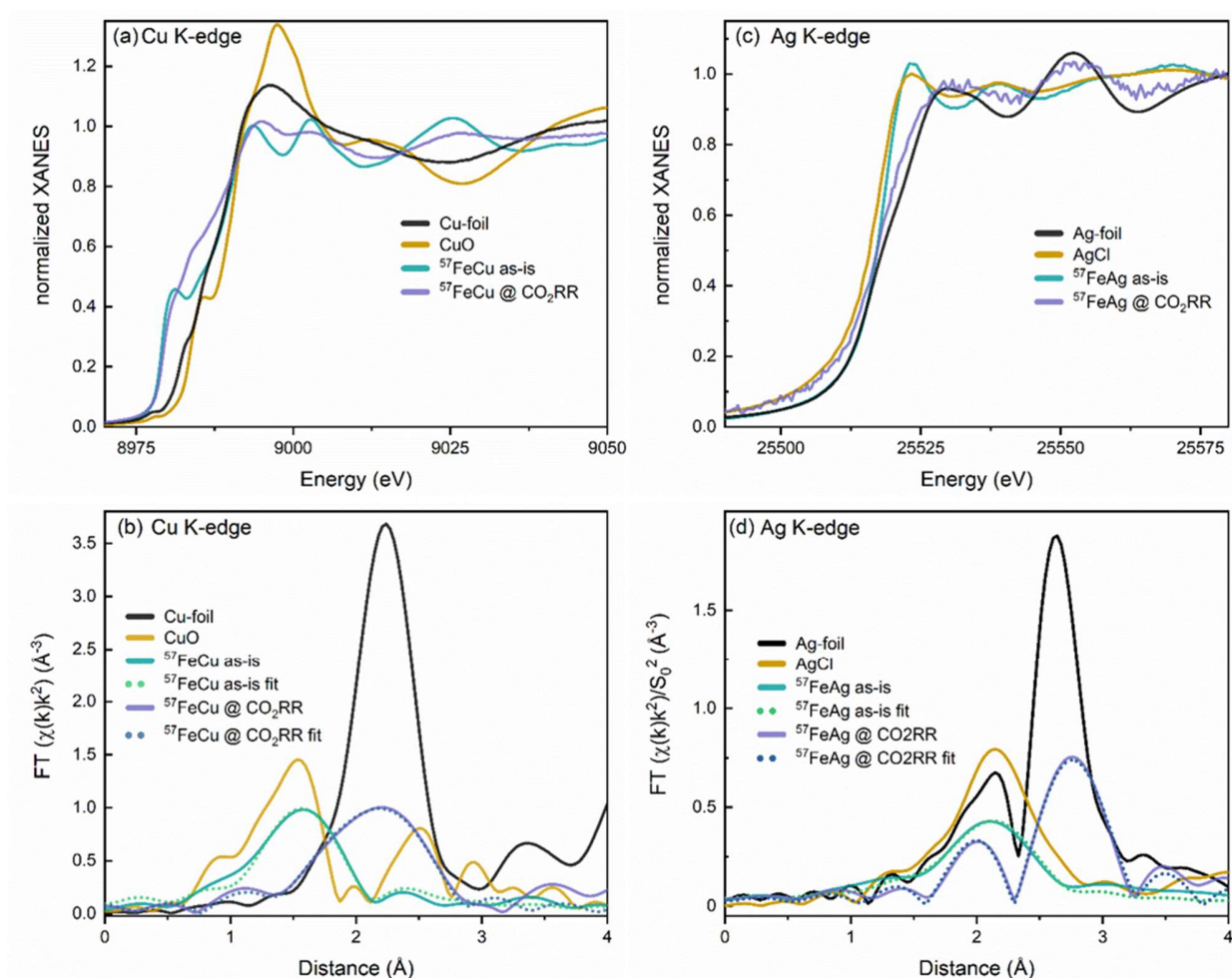


Figure 6 Cu K-edge (a,b) and Ag K-edge (c,d) XANES (a,c) and Fourier-transforms of k^2 -weighted EXAFS data (phase-uncorrected) (b,d) of $^{57}\text{FeCu}$ and $^{57}\text{FeAg}$ NP samples in as-prepared state (as-is) and under CO_2RR conditions (-1.1 V vs RHE in 0.1 M KHCO_3 after 1.5 h ($^{57}\text{FeCu}$) and 2.5 h ($^{57}\text{FeAg}$)). Reference spectra from CuO, Cu-foil, AgCl^[25] and a Ag-foil are shown for comparison.

After careful evaluation of the *operando* structural and chemical information extracted from two complementary synchrotron X-ray methods (XAFS and NRIXS), we demonstrate a significant reduction of iron oxide to metallic species under reaction conditions, accompanied by a structural transformation from an amorphous or atomically disordered phase to a crystalline structure. We observed a similar transformation for the secondary metals in the bimetallic NP samples via XAFS; a clear reduction of CuO to Cu in $^{57}\text{FeCu}$, and even more pronounced, AgCl to Ag transformation in $^{57}\text{FeAg}$. NRIXS served to distinguish segregated Fe-Ag and Fe-Cu phases (the present case) from alloyed structures, since we only detected a bcc-like Fe structure typical for a segregated Ag

shell with bcc-Fe NP core. A small degree of Fe-Cu intermixing is obtained for the $^{57}\text{FeCu}$ sample from the XAFS data analysis. The bcc-Fe phase also predominates the NRIXS signal. Both these findings suggest the predominant segregation of both metals in our samples under CO_2RR conditions.

It should be noted that the X-ray-based spectroscopic methods employed here for the characterization of our nanosized electrocatalysts (NRIXS and XAS) also have limitations that need to be considered for the interpretation of the data obtained. For instance, both methods are bulk-sensitive ensemble-averaging techniques, i.e. the measured signal is an average of contributions from all metal species, residing both at the surface and in the core regions of all NPs within the X-ray irradiated sample area. Therefore, meaningful composition information can only be obtained when applied to multimetallic samples with an homogeneous inter- and intraparticle composition and chemical state. In addition, the interpretation of structural characteristics extracted from X-ray spectroscopy methods will be further hindered when wide nanoparticle size and shape distributions are available in the as prepared samples.^[34-35] This is even more relevant when *operando* catalysis studies are undertaken, since in that case the coexistence of particles of different size and shape evolving under reaction conditions is common. Therefore, the interpretation of structural information extracted from XAS and NRIXS in terms of 3D structural motifs for inhomogeneous systems should be carried out with extreme care, and must rely on complementary insight from additional techniques such as microscopy methods and theoretical modelling.

In the present study, we combined AFM, TEM, XPS, NRIXS and XAS to gain insight into the evolution of the structure and composition of homogeneously dispersed size and shape-controlled FeCu and FeAg NPs during CO_2RR . We remark here that for the FeAg NPs, the EDX line profiles acquired after reaction (**Fig. S4**) extend further outwards as compared to the Fe profiles, which suggests some degree of Ag segregation to the surface. For the FeCu NPs, the line profiles indicate more even mixing of the two elements. These STEM-EDX results are however limited because of the difficulty of obtaining high quality and contaminant-free images in the presence of the Nafion NP binder and carbon support, which strongly interacts with the electron beam, resulting in the need of employing short acquisition times. Nevertheless, the microscopy data are consistent with the X-ray spectroscopy and selectivity data regarding the segregation behavior in the FeCu and FeAg NPs.

According to the respective surface energies^[36-38] and atomic sizes, the formation of an either Cu or Ag shell and a bcc-Fe core is expected. For instance, we have previously shown the segregation of copper to the surface of CuNi NPs under reducing conditions (CO₂+CO+H₂).^[39] This is in good agreement with the electrochemical behavior observed here for the FeAg NPs, which was found to be reminiscent to pure silver. The trend was not as clear for the FeCu sample, where size effects already demonstrated for pure Cu NPs will also lead to an increase of the H₂ production with decreasing NP size at the expense of CO and hydrocarbons. Therefore, for the Fe-Cu system, we cannot discard the possibility of having Cu-rich and Fe-rich regions coexisting at the NP surface, which would also explain the increased H₂ production (taking place over the exposed Fe component) as compared to the selectivity obtained in similarly sized monometallic Cu NPs.

Table 1 Coordination numbers (CN) and interatomic distances (R), derived from the fit of EXAFS data of samples ⁵⁷FeCu and ⁵⁷FeAg at the Fe, Cu and Ag K-edges. The data were measured in air (as-is) and *operando* during CO₂RR at -1.1 V vs RHE in 0.1 M KHCO₃ after 1.5 h-3.5 h. The uncertainty of the last digit is given in parentheses. Additional parameters and fit results are shown in Tables S3 and S4. A single Fe-Fe scattering path was used to fit the overlapping contributions from the first two coordination shells in bcc Fe due to the limited resolution in R-space as a result of the short Fe K-edge spectrum.

Sample	Treatment	CN	R(Å)	CN	R(Å)	CN	R(Å)	CN	R(Å)
		Cu-O		Cu-Cu		Fe-O		Fe-Fe	
⁵⁷ FeCu	As is	2.3(3)	1.95(1)			6.9(2.1)	1.81(6)		
	Operando	0.7(6)	1.95(3)	4.2(1.2)	2.56(1)	2.3(1.2)	1.81(4)	1.5(0.9)	2.48(5)
		Ag-Cl		Ag-Ag		Fe-O		Fe-Fe	
⁵⁷ FeAg	As is	4.6(6)	2.81(1)			6.7(1.6)	1.81(2)		
	Operando			5.4(7)	2.941(1)	2 (1)	1.81(4)	1(1)	2.48(6)
Bulk	Cu foil ^[40]			12	2.566				
	CuO ^[41]	4	1.95(1)						
	Ag foil ^[42]			12	2.883(7)				
	AgCl ^[43]	6	2.79(1)						
	FeO(OH) ₂ ^[3]					6	2.00		
	Fe foil ^[33]							8+6	2.47(1)

Although the present nanosized catalysts do not display an outstanding activity or selectivity for CO₂RR, in line with previous studies for small NPs which also favor H₂-evolution,^[27-29] our study serves to provide fundamental insight into the dynamic behavior of electrocatalysts under reaction conditions. Furthermore, we illustrate a powerful combination of X-ray based synchrotron techniques for the characterization of the structural and chemical evolution of electrocatalysts that, when available hand in hand with reactivity data, allows one to gain in depth understanding of the structural motifs and chemical compositions responsible for specific selectivity trends.

Summary

This study demonstrates the dynamic transformations undergone by FeCu and FeAg NPs during the electrochemical reduction of CO₂. In particular, *operando* NRIXS and XAFS data revealed the formation of an Ag shell surrounding a bcc-Fe core. On the other hand, more intermixing was observed for the FeCu NPs, and the presence of separated Fe and Cu regions on the NP surface under reaction conditions could not be excluded. Interestingly, a similar CO production was observed for the thin Ag-shell in the FeAg NPs as compared to pure Ag NPs, which indicates the optimization of the use of the noble metal. Overall, and due to the small NP size, the production of H₂ was, however, favored, especially in the Fe and FeCu samples.

Finally, our work emphasizes that *operando* experiments are a very valuable tool to link catalytic properties to structure and composition of electrocatalysts under realistic working conditions.

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