Data Report

(206) 512-3659 reports@northwestlabs.com 98526 Parada Dr Seattle WA

Pfizer Inc.
235 East 42nd Street NY, NY 10017
(212) 733-2323

Order #2031569
Date Created: 12/6/2017

Below are the results from the compounds you sent. If you have any questions, don't hesitate to reach out to us!

HYDROGEN CARBONATE (LT# 230256)

Assay #8 Results

Biochemical Pharmacology (BP) Test

Reactivity	●●●○○	Compound has moderate reaction with other substances
Covalence	●○○○○	Compound has a low covalence with carbon

PharmaScreen® (PS) Test

Heat Threshold	**FAILED**	Compound's bonds dissolve at high temperatures
Dissolvance Rate	**86 PPM**	Compound dissolves at a high rate

Assay #2 Results

DiscoveryScreen® (DS) Test

RESULTS INCONCLUSIVE

DIHYDROGEN PHOSPHATE (LT# 230257)

Assay #5 Results

ImmunoScreen® (IS) Test

Effectivity	88%	The compound influences symptoms in most test subjects
Catalyst	**26 SECONDS**	This compound reaches a catalysm quickly

Summary Report

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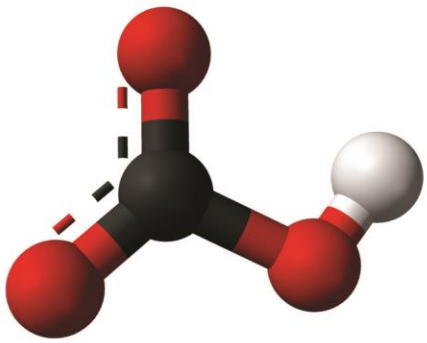
Hydrogen Carbonate (LT# 230256)



A bicarbonate salt forms when a positively charged ion attaches to the negatively charged oxygen atoms of the ion, forming an ionic compound. Many bicarbonates are soluble in water at standard temperature and pressure; in particular, sodium bicarbonate contributes to total dissolved solids, a common parameter for assessing water quality

The most common salt of the bicarbonate ion is sodium bicarbonate, NaHCO3, which is commonly known as baking soda. When heated or exposed to an acid such as acetic acid (vinegar), sodium bicarbonate releases carbon dioxide. This is used as a leavening agent in baking.
The flow of bicarbonate ions from rocks weathered by the carbonic acid in rainwater is an important part of the carbon cycle.

Bicarbonate also serves much in the digestive system. It raises the internal pH of the stomach, after highly acidic digestive juices have finished in their digestion of food. Ammonium bicarbonate is used in digestive biscuit manufacture.

Dihydrogen Phosphate (LT# 230257)

Is formed when a solution of phosphoric acid is added to ammonia until the solution is distinctly acidic. It crystallizes in tetragonal prisms. Monoammonium phosphate is often used in the blending of dry agricultural fertilizers.[3] It supplies soil with the elements nitrogen and phosphorus in a form usable by plants. The compound is also a component of the ABC powder in some dry chemical fire extinguishers. This substance is also supplied in an emerald green, amethyst, or aquamarine crystal growing box kit for children.
Solid monoammonium phosphate shows a dissociation pressure of ammonia of 0.05 mm Hg at 125 °C based on the decomposition reaction as follows:

$NH_4H_2PO_4(s) \rightarrow NH_3(g) + H_3PO_4(l)$

ADP is a widely used crystal in the field of optics due to its birefringence properties. As a result of its tetragonal crystal structure, this material has negative uniaxial optical symmetry with typical refractive indices no = 1.522 and ne = 1.478 at optical wavelengths.