

<div>FormulaBa Ti O3</div> <div>NameBarium Titanium Oxide</div> <div>Name (mineral)</div> <div>Name (common)</div>		d	2θ	l	h	k	l
		4.04000	21.9836	14	1	0	0
		2.85000	31.3619	100	1	1	0
		2.32800	38.6450	30	1	1	1
		2.01600	44.9268	35	2	0	0
		1.80300	50.5841	10	2	1	0
		1.64450	55.8622	25	2	1	1
		1.42500	65.4437	15	2	2	0
		1.34400	69.9389	5	3	0	0
		1.27400	74.4046	15	3	1	0
<div>Lattice: Cubic</div> <div>S.G.:</div>		Mol. weight = 233.23					
		Volume [CD] = 65.5					
		Dx =					
<div>a = 4.03100</div> <div>b =</div> <div>c =</div> <div>a/b = 1.00000</div> <div>c/b = 1.00000</div>		Dm =					
		I/Icor = -1.000					
		alpha =					
		beta =					
		gamma =					
		Z =					
		<div>Sample Preparation: Sample was prepared by hydrolysis of titanium tetraisopropoxide in aqueous solution of barium hydroxide and has submicron size of particles. Absorbs about 6% of OH and alcoholic radicals. Metastable form. It changes to stable tetragonal form above 800 C</div>					
				<div>Primary Reference</div> <div>Publication: Bull. Chem. Soc. Jpn.</div> <div>Detail: volume 47, page 1168 (1974)</div> <div>Authors: Naka, S. et al.</div>			
<div>Radiation: CuKa</div> <div>Wavelength: 1.54060</div> <div>SS/FOM: 39.9 (0.025,11)</div>						<div>Filter: F</div> <div>d-spacing:</div>	