

UCB008 - APPLIED CHEMISTRY



Molecular Spectroscopy Series Lecture - IX

UV-Visible Spectroscopy – Analysis and Applications

by

Prof. Ranjana Prakash

School of Chemistry and Biochemistry
Thapar Institute of Engineering and Technology
Patiala -147004, India

Ranjana Prakash



Learning Outcomes

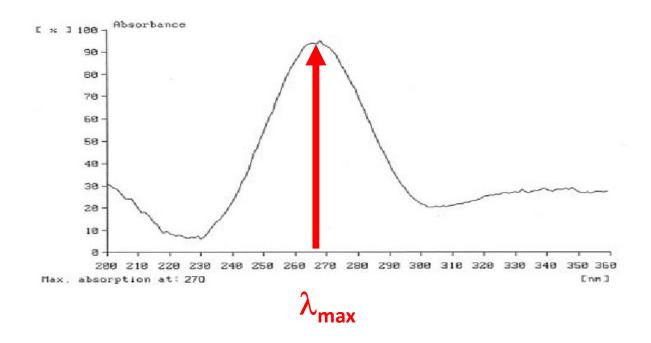
At the end of this session participants should be able to:

- Understand the analytical procedure of qualitative and quantitative analysis using UV-visible spectrophotometer
- Gain knowledge on applications of UV-visible spectroscopy



Qualitative analysis using UV-visible spectrophotometer

- λ_{max} of organic compounds can be determined using UV-visible spectra
- This value can be used to detect the extent of conjugation
- ... and absence/presence of the chromophore in an organic compound.





Qualitative Analysis:

- Extent of conjugation in polyenes can be estimated.
- Shift in the absorption, due to addition of unsaturation in a compound, towards longer wavelength can be detected.
- Presence or absence of a chromophore can be detected.
- If UV-visible spectrum of a compound has no peaks above 200 nm, it shows the absence of conjugation, carbonyl group and aromatic ring.
- It is used for characterizing aromatic compounds and conjugated olefins.
- Distinction between conjugated and non-conjugated compounds

$$(CH_3)_2C=CH-CO-CH_3$$

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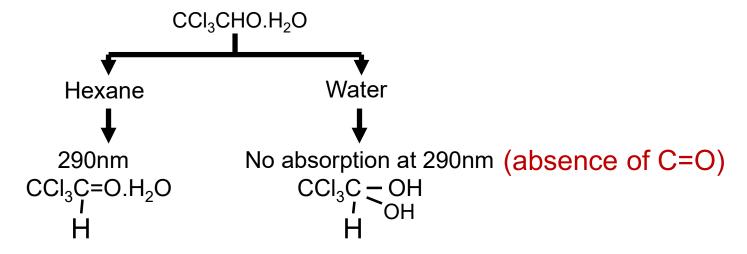
- Qualitative Analysis:
 - Geometric isomers can be distinguished

Trans isomer always absorb at higher wavelength



Qualitative Analysis:

 Identification of a compound in different solvents- If structure of a compound changes in different solvents, the change can be detected by uv-vis spectra.



Sharma YR, 2010, Elementary Organic Spectroscopy, S.Chand & Co.



- Series of standard solutions having known concentration of analyte X are prepared.
- One of these solutions is analyzed by the spectrophotometer and its λ_{max} is determined.
- Absorbance of all the standard solutions are analyzed at the determined $\lambda_{\text{max.}}$
- A calibration curve (absorbance vs concentration) is prepared for the standard samples.
- Signal intensity of unknown sample of analyte X is taken.
- The concentration of the analyte X is then determined based on the absorbance value from the calibration curve.

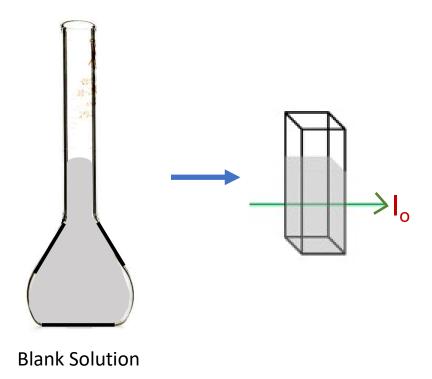


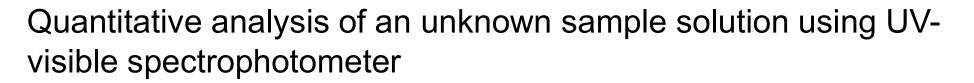
Step 1: Series of standard solutions having known concentration of analyte X are prepared.





Step 2: Solvent is taken as **blank solution** for which absorbance is measured and absorbance value is set to **Zero**.

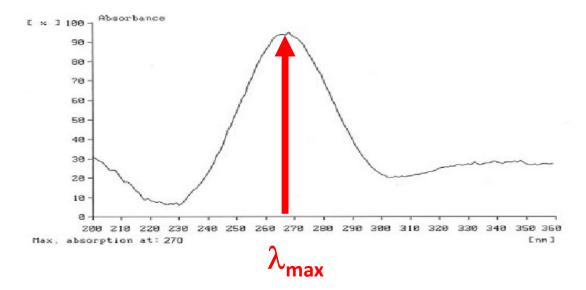




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Step 3: One of the standard solutions is analyzed by the UV-visible spectrophotometer and its λ_{max} is determined.

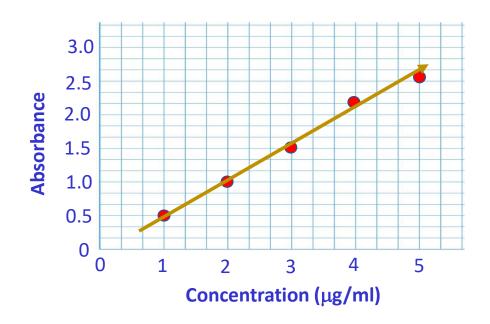


Instrument is calibrated with blank, periodically, when wavelength is changed, in case of single beam spectrophotometer.

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- Step 4: Absorbance of all the standard solutions are analyzed at the determined λ_{max} .
- Step 5: A calibration curve (absorbance vs concentration) is prepared by analyzing the standard samples.

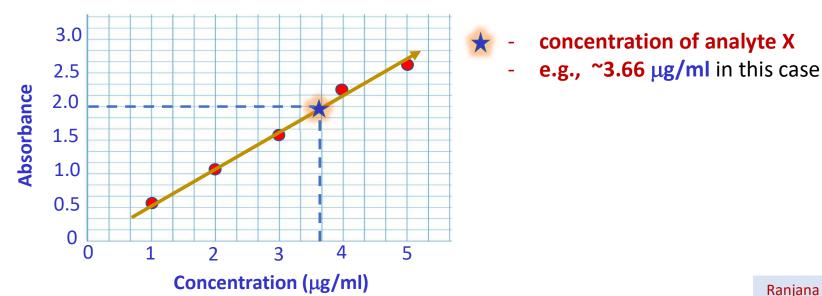




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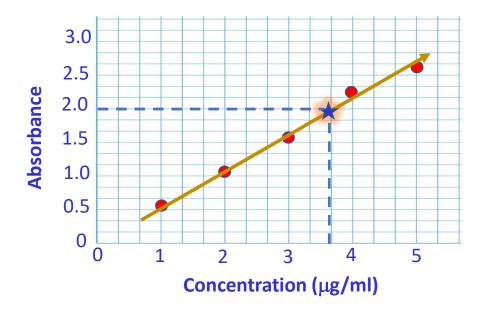
Quantitative analysis of an unknown sample solution using UVvisible spectrophotometer

- Step 6: Absorbance value of the solution of analyte X (unknown) is determined at the same λ_{max} .
- Step 7: The concentration of the analyte X is then determined based on the absorbance value from the calibration curve by extrapolating towards concentration in X axis.





- Quantitative Analysis:
 - It can be used to find out molar concentration of the solute under study.





With this session.....

 We complete the understanding of the fundamental aspects associated with UV-visible spectroscopy

From the next session.....

 We will be exploring the principle, instrumentation and applications of Infra-red spectroscopy