

Tarea 4

4.1, 4.3, 4.4, 4.5, 4.7

6.1 y 6.2.

4.1.)

Del apéndice D.3.

$$\frac{\mu}{\rho} = 5.27 \text{ E-2}$$

$$\frac{\mu_{en}}{\rho} = 2.43 \text{ E-2}$$

$$\rho = 8.86 \text{ g/cm}^3$$

$$\mu = 4.69 \text{ E-2 m}^{-1}$$

$$\mu_{en} = 2.14 \text{ E-2 m}^{-1}$$

$$B_0 \approx 1.06.$$

$$\frac{N_L}{N_0} = B e^{-\mu L} \Rightarrow L = -\frac{1}{\mu} \ln\left(\frac{N_L}{N_0 B}\right)$$

$$L(\mu) = 1.46 \text{ m}$$

$$L(\mu_{en}) = 3.14 \text{ m}$$

4.3

$$\bar{E} = \sum Q$$

$$\bar{D} = \frac{d\bar{E}}{dm}$$

$$\bar{E} = \sum Q = (10^{-16} \text{ m}^2 \text{ s}^{-2}) \times 0.40$$

$$d\bar{E}/dm = c^2 10^{-16} \cdot 0.4 = 3.595 \text{ Gy}$$

4.4.

$$a) (K_e)_{air} = X \cdot \left(\frac{\bar{W}}{e} \right)_{air}$$

$$= X \cdot 33.97 \text{ J/C}$$

$$= X \cdot 0.876 \text{ Gy/100 R}$$

$$= X \cdot 8.76 \text{ Gy/R}$$

$$= 2.409$$

$$b) D_{air} = (K_e)_{air} = 2.409$$

c) Si CPE existe en P.

4.5

$$\frac{D_{\text{cooper}}}{D_{\text{air}}} = \frac{(K_c)_{\text{cooper}}}{(K_c)_{\text{air}}} = \frac{(M_{\text{en}}/\rho)_{\text{cooper}}}{(M_{\text{en}}/\rho)_{\text{air}}}$$

$$D_{\text{cooper}} = \frac{(M_{\text{en}}/\rho)_{\text{cooper}}}{(M_{\text{en}}/\rho)_{\text{air}}} D_{\text{air}} = \frac{2142}{526.2} 2.47 \text{ Gy} = 9.81.$$

4.7

$$E_{\text{min}} = 300 \text{ s.}$$

$$a) \Psi \cdot 300 \text{ s} = 0.111 \text{ J/cm}^2 = \Psi$$

$$K_c = \Psi \left(\frac{M_{\text{en}}}{\rho} \right).$$

Dice "baja energía", lo que tomamos como 0.05 MeV.

$$K_c = 0.111 \text{ J/cm}^2 \cdot 0.184 \text{ cm}^2/\text{g} = 0.020424 \text{ J/g} \cdot 1000 \text{ g/kg}$$

$$b) K_c = 20.424 \text{ Gy}$$

$$c) D^{\text{CPE}} = K_c = 20.424 \text{ Gy.}$$

$$d) \text{ La mitad. } 10.212 \text{ Gy.}$$

6.1)

1.

a)

$$\lambda = \frac{1}{\tau} = \frac{0.6931}{\tau_{1/2}} = \frac{0.6931}{1602 \cdot 365 \cdot 24 \cdot 60 \cdot 60 \text{ s}} = 1.37 \text{ E-11 s}^{-1}$$

b) $\tau = \frac{1}{\lambda} = 7.29 \text{ E10 s.}$

c) $226000 \text{ mg} = 1 \text{ mol.}$

$$\left(\frac{5 \text{ mg}}{226000 \text{ mg}} \right) \text{ mol.} \cdot \frac{1 \text{ Na}}{1 \text{ mol.}} = 1.332 \text{ E20 atoms}$$

d) $= 1.333 \text{ E20 atoms}$

e)

$$N = N_0 e^{-\lambda t}$$

$$= 1.333 \text{ E20 } e^{-1.37 \text{ E-11 s}^{-1} \cdot t}$$

$$49.4 \text{ mCi}, 8.71 \text{ E14.}$$

6.2)

a) $\lambda = \frac{1}{\tau} = \frac{0.6931}{\tau_{1/2}} = 4.48 \text{ E-7 s}^{-1}$

b) $\lambda_1 = 3.05 \text{ E-7 s}^{-1}$

$$\lambda_2 = 1.43 \text{ E-7 s}^{-1}$$

c) $2.42 \text{ Ci}, 8.92 \text{ E10 Bq}$

d) $2.87 \text{ E10 s}^{-1}, 4.65 \text{ E9 s}^{-1}$