1. Transfer 10ml of concentrated NaOH to a 100ml volumetric flask using a 10ml pipette.
2. Fill up the volumetric flask with distilled water and mix the liquid.
3. Transfer 10ml of the diluted NaOH solution to a 250ml conical flask by using a 10ml pipette.
4. Add 1 drop (1ml) of indicator (Phenolphthalein) into the conical flask.
5. Fill the burette up with the unknown acid.
6. Gradually add the acid to the conical flask until a color change has occurred. In this case, the color turns from pink to colorless.
7. Record the volume of the acid added to the conical flask (Vacid).
8. The number of moles of NaOH used is determined by 10mol/dm3 \* 10cm3 /10 = 0.1 mol.
9. Since the acid is monoprotic, same number of moles of acid and base reacts together. So 0.1 mol of unknown acid reacts in this case.
10. The concentration of the acid can be determined by 0.1mol/Vacid.

