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EXPT. NO.	NAME	M T W T F S S	Page No.	YOUVA
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between a reference electrode and an electrode sensitive to the hydrogen ion activity when they are both immersed in the same aqueous solution. The reference electrode may be a Silver chloride electrode or a calomel electrode. The hydrogen-ion selective electrode is a standard hydrogen electrode.

Procedure :- $XN \text{ HCl} \rightarrow 0.1N \text{ NaOH}$
 $XN \text{ NaOH} \rightarrow 0.1N \text{ HCl}$

you have been given an unknown concⁿ on a 200ml of volumetric flask.

Dilute the given solⁿ to 200ml with distilled water.

Take some of this solⁿ in glass beaker and 20 ml of water into it.

Fill the burette with unknown concentration of HCl and NaOH.

Note down the reading by one ml internal.

Graph plot a graph between.

① $p^H \rightarrow V_{me}$ of titrate added

② $\Delta p^H / \Delta V \rightarrow V_m$ of titrate added

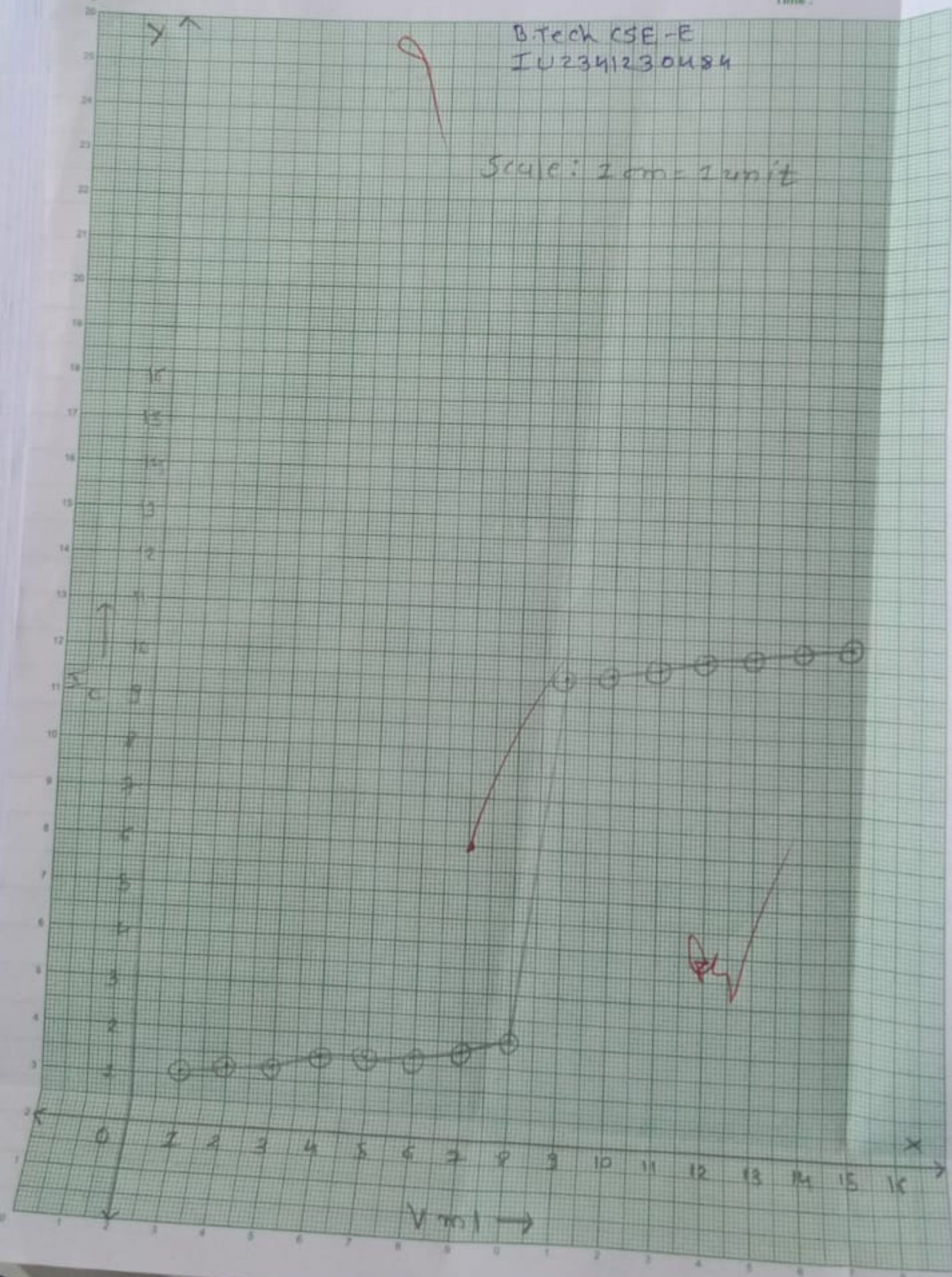
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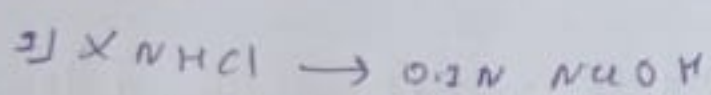
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Scale: 1 cm = 1 unit



→ calculation:

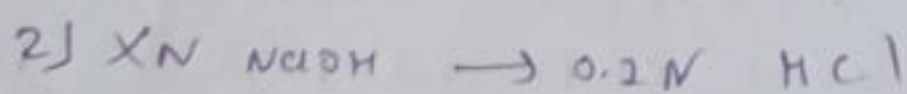


volume of sodium hydroxide $V_1 = \underline{10 \text{ ml}}$

normality of sodium hydroxide $N_1 = \underline{0.2 N}$

volume of HCl $V_2 = \underline{8 \text{ ml}}$

$$\begin{aligned} \text{normality of HCl } N_2 &= \frac{V_1 \times N_1}{V_2} \\ &= \frac{10 \times 0.2}{8} \\ &= 0.125 N \end{aligned}$$



volume of hydrochloric acid $V_1 = \underline{10 \text{ ml}}$

normality of HCl $N_1 = \underline{0.2 N}$

volume of NaOH $V_2 = \underline{8}$

$$\begin{aligned} \text{normality of NaOH } N_2 &= \frac{V_1 \times N_1}{V_2} \\ &= \frac{10 \times 0.2}{8} \\ &= 0.125 N \end{aligned}$$

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