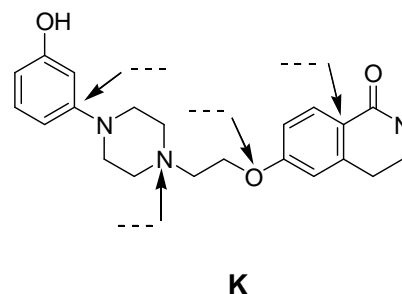
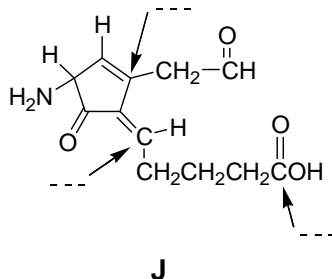
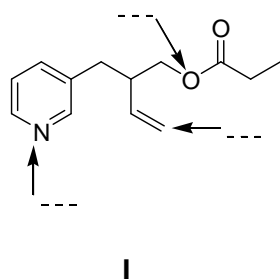
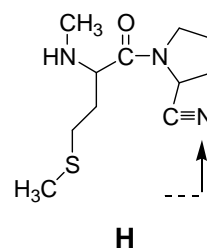
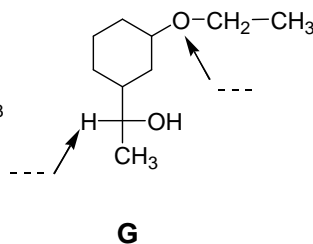
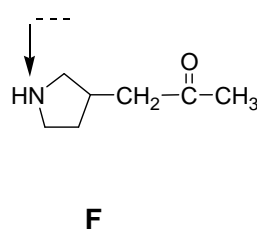
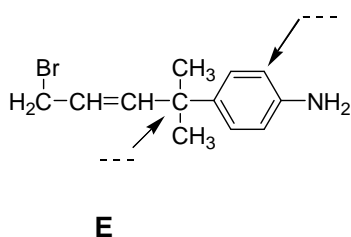
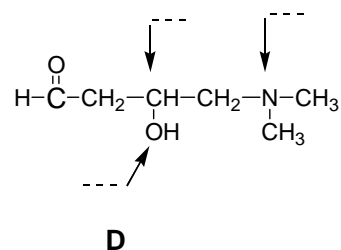
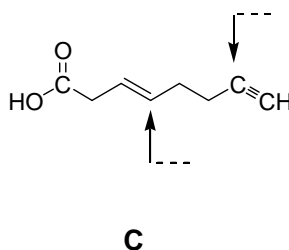
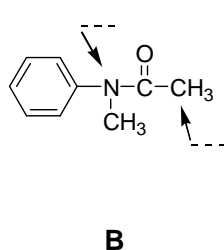
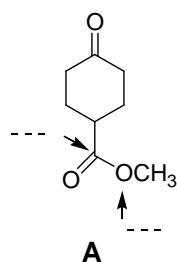
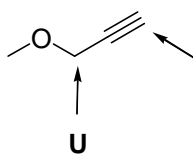
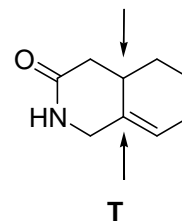
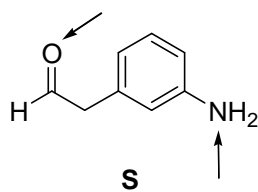
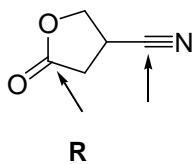
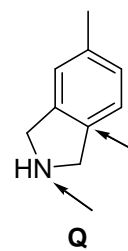
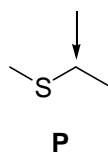
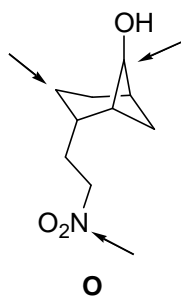
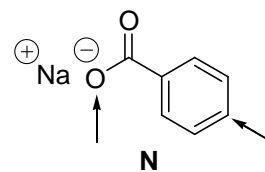
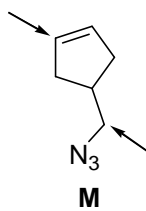
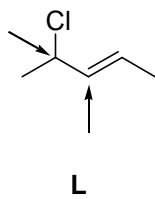


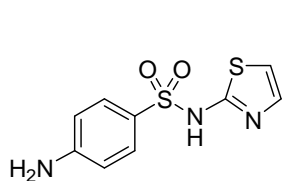
1. Geef de hybridisatie van de aangeduide atomen en benoem de functionele groepen in elke structuur.

Kies tussen: carbonzuur nitril ether azide
 (carboxylaat) nitrogroep thiol
 keton alkyn sulfide
 aldehyde alkeen alkyl halogenide
 ester amide vinyl halogenide
 lacton aromaat amine (prim, sec, tert)
 alcohol zuurchloride (alifatisch, aromatisch, heterocyclisch)

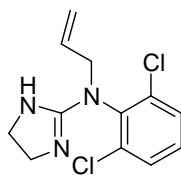




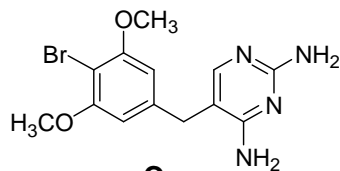
2. Vervolledig volgende Lewis structuren met de ontbrekende vrije elektronenparen:



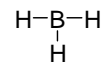
A



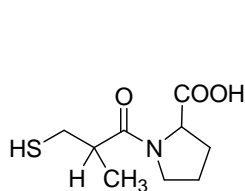
B



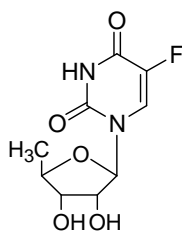
C



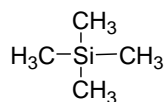
D



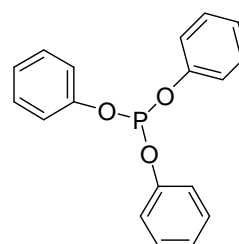
E



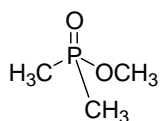
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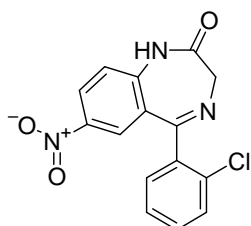
G



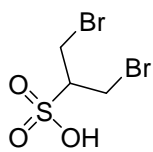
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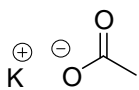
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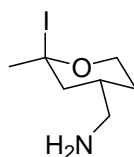
J



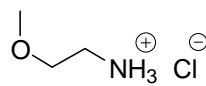
K



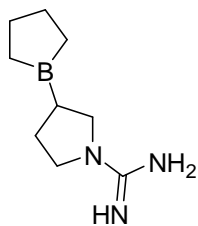
L



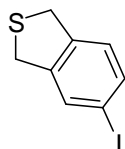
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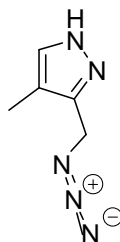
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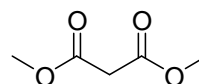
O



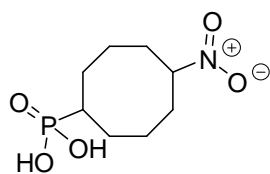
P



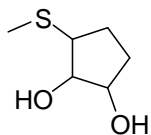
Q



R



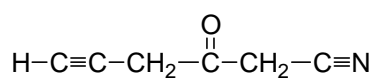
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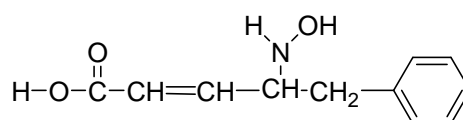
T

3. Benoem elk type binding (a, b, c..) aanwezig in volgende structuren ?

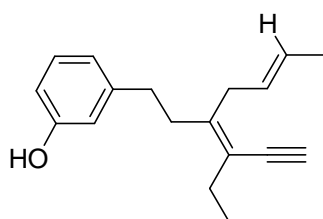
- a. $sp^3-s : \sigma$
- b. $sp^2-s : \sigma$
- c. $sp-s : \sigma$
- d. $p-p : \pi$
- e. $sp^3-sp^3 : \sigma$
- f. $sp^2-sp^2 : \sigma$
- g. $sp-sp : \sigma$
- h. $sp^3-sp^2 : \sigma$
- i. $sp^3-sp : \sigma$
- j. $sp^2-sp : \sigma$



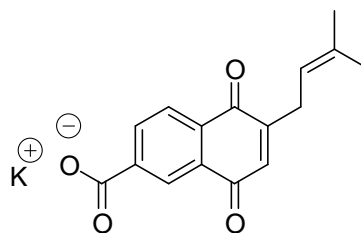
A



B



C



D

4. Teken de Lewisstructuur

- A) CCl_4
- B) CH_3Br
- C) CH_3NH_3^+
- D) H_2NNH_2
- E) PH_3
- F) H_2S
- G) $\text{CH}_3\text{CH}_2\text{OH}$
- H) $\text{O}(\text{CH}_3)_3^+$

