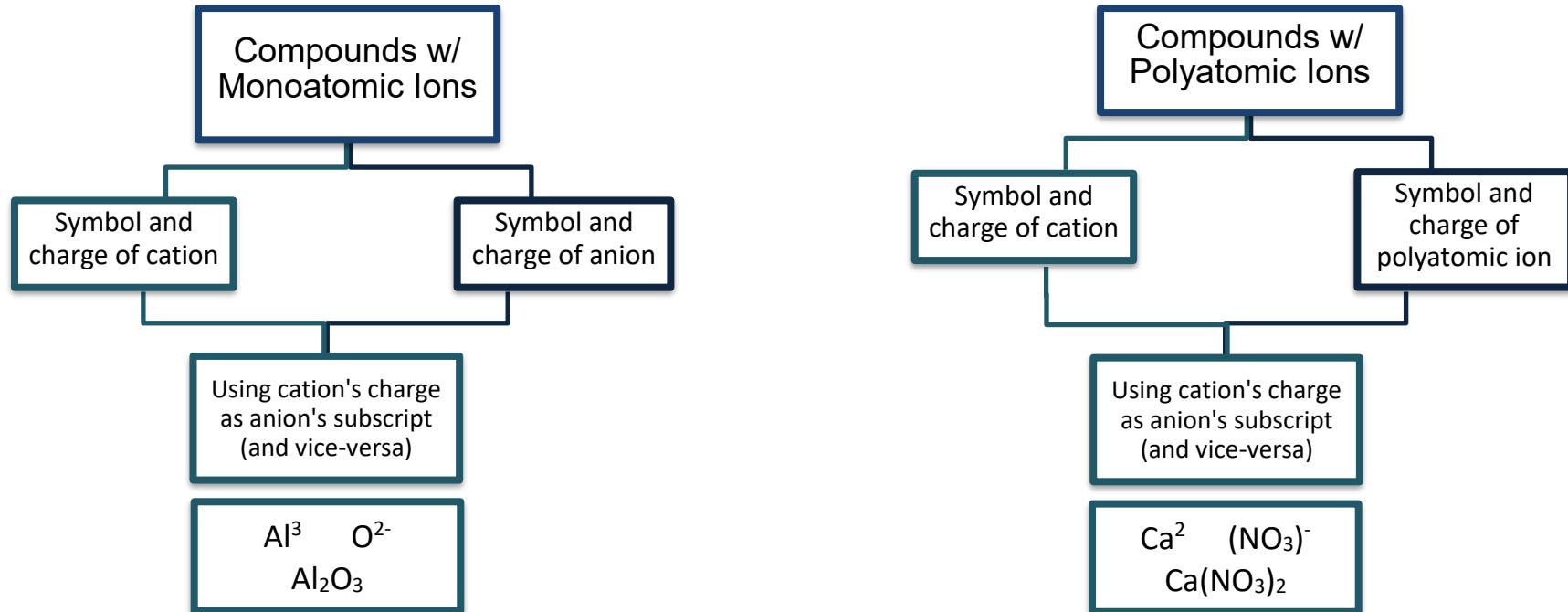
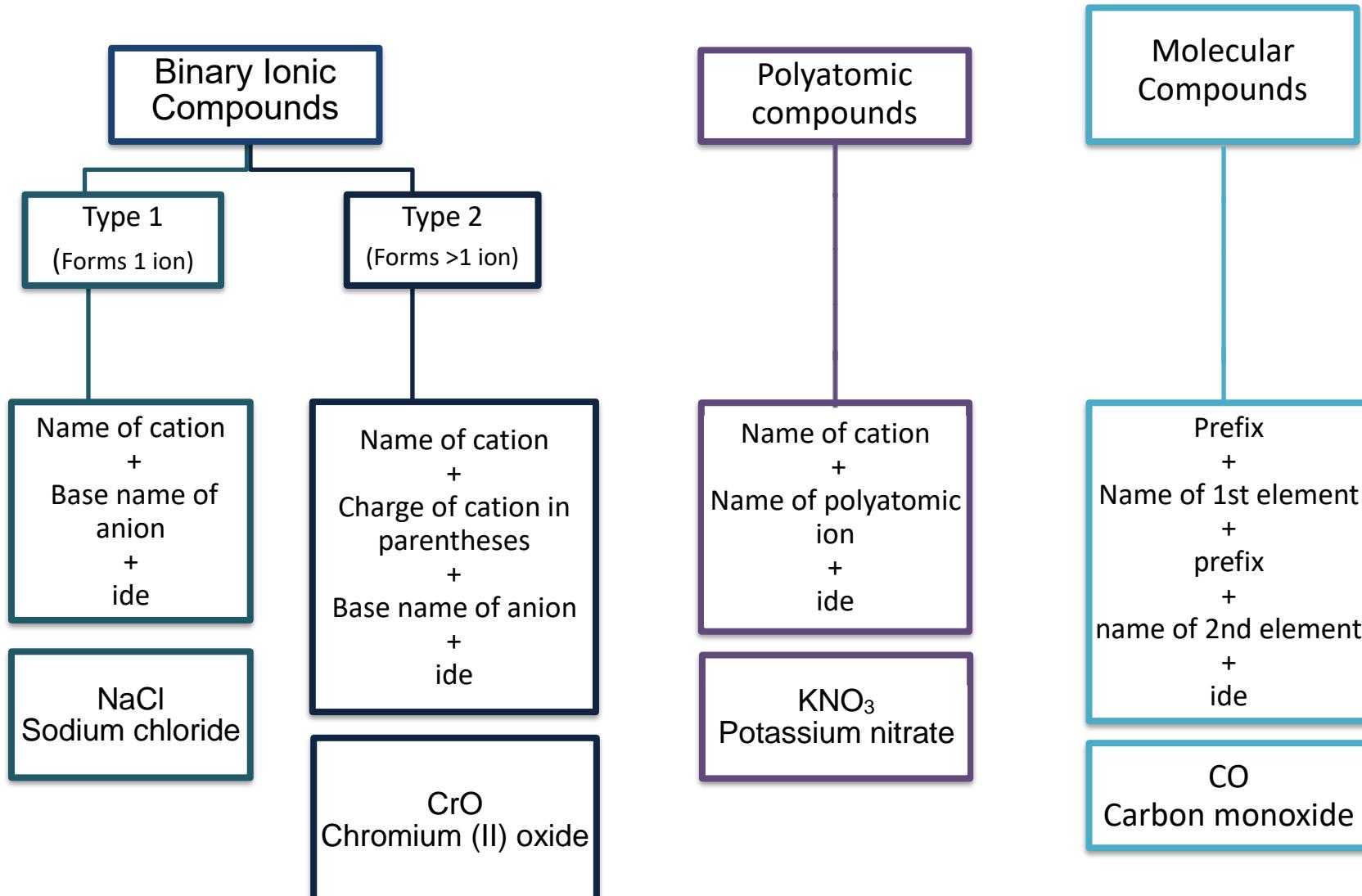


## FLOWCHART FOR WRITING MOLECULAR FORMULAS OF COMPOUNDS



Name	Formula	Name	Formula	Name	Formula	Name	Formula
Phosphate	PO <sub>4</sub> <sup>3-</sup>	Permanganate	MnO <sub>4</sub> <sup>-</sup>	Sulfate	SO <sub>4</sub> <sup>2-</sup>	Ammonium	NH <sub>4</sub> <sup>+</sup>
Carbonate	CO <sub>3</sub> <sup>2-</sup>	Perchlorate	ClO <sub>4</sub> <sup>-</sup>	Sulfite	SO <sub>3</sub> <sup>2-</sup>	Acetate	C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> <sup>-</sup>
Hydroxide	OH <sup>-</sup>	Chlorate	ClO <sub>3</sub> <sup>-</sup>	Hydrogen sulfite	HSO <sub>3</sub> <sup>-</sup>	Chromate	CrO <sub>4</sub> <sup>2-</sup>
Nitrate	NO <sub>3</sub> <sup>-</sup>	Chlorite	ClO <sub>2</sub> <sup>-</sup>	Hydrogen sulfate	HSO <sub>4</sub> <sup>-</sup>	Dichromate	Cr <sub>2</sub> O <sub>7</sub> <sup>2-</sup>
Nitrite	NO <sub>2</sub> <sup>-</sup>	Hypochlorite	ClO <sup>-</sup>	Cyanide	CN	Peroxide	O <sub>2</sub> <sup>2-</sup>

## FLOWCHART FOR NOMENCLATURE OF COMPOUNDS



### Common Metals That Form More Than One (1) Type of Ion

$\text{Cr}^2$	Chromium (II)
$\text{Cr}^3$	Chromium (III)
$\text{Fe}^2$	Iron (II)
$\text{Fe}^3$	Iron (III)
$\text{Hg}_2^2$	Mercury (I)
$\text{Hg}^2$	Mercury (II)
$\text{Co}^2$	Cobalt (II)
$\text{Co}^3$	Cobalt (III)
$\text{Cu}^+$	Copper (I)
$\text{Cu}^2$	Copper (II)
$\text{Pb}^2$	Lead (II)
$\text{Pb}^4$	Lead (IV)
$\text{Sn}^2$	Tin (II)
$\text{Sn}^4$	Tin (IV)

### Greek Prefixes

1	Mono-
2	Di-
3	Tri-
4	Tetra-
5	Penta-
6	Hexa-
7	Hepta-
8	Octa-
9	Nona-
10	Deca-

### Prefixes and Suffixes for Oxyanions (Oxygen containing compounds)

Hypo-  
Hyper-  
Per-  
-ite  
-ate

### References

- Hoefnagels, M. (2016). *General biology books I and II*. Quezon City: McGraw-Hill Education.  
 Mason, K. A., Losos, J. B., & Singer, S. R. (2017). *Biology* (11th ed.). New York: McGraw-Hill Education.  
 Tro, N. J., & Au-Yeung, H. Y. (2015). *Introductory chemistry*. Singapore: Pearson Education Limited.