# GILES CHEMICAL CORPORATION COMPANY PROCEDURE Standard Operating Procedure Page: 1 of 2 Revision: Date: 3/27/06 Reviewed: Carl Mooney Title: Determination of % MgSO<sub>4</sub> by Specific Gravity

QA-LAB-06

**Safety:** Wear safety glass and/or goggles when working in the lab.

Purpose: Determination of % MgSO<sub>4</sub> by Specific Gravity

**Procedure:** 

### **Background Information:**

Specific gravity is usually, and most easily, determined by immersing a suitable hydrometer in the subject solution, with adjustment for temperature as necessary. In order to check any solution thus obtained with a somewhat greater degree of accuracy the following procedure is used.

#### **Procedure:**

A measured volume of the subject solution is weighed and the weight divided by the volume. This produces a value for the specific gravity of the solution being measured and the % MgSO<sub>4</sub> is determined by reference to standard charts.

## **Equipment:**

25 - mL laboratory volumetric flask Weighing Balance -- B440 Sartorius Laboratory Thermometer -- 0 - 100° C

Set of standard charts for conversion of Specific Gravity and Temperature determination to  $\%\ MgSO_4$ 

#### **Procedures:**

- 1 A dry 25 mL volumetric flask is placed on the weighing balance and tarred to zero
- 2 Approximately 25 mL of subject sample is added to the graduate and the weight recorded.
- 3. The temperature of the sample is taken and recorded
- 4. Specific gravity is determined using the following formula

## <u>Weight of sample (g)</u> = Specific Gravity (g/mL) Volume of sample (mL)

5. % MgSO4 is determined by referring the above mentioned charts, using the specific gravity reading and the temperature of the solution

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# TRAINING DOCUMENTATION

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