

Answer 2.1

Based on the Law of Mass Action,

E is consumed by forward reaction, and produced by 2 reverse reactions, hence,

$$\frac{d[E]}{dt} = k_2[ES] + k_3[ES] - k_1[E][S] = (k_2 + k_3)[ES] - k_1[E][S]$$

Similarly,

$$\frac{d[S]}{dt} = k_2[ES] - k_1[E][S]$$

$$\frac{d[ES]}{dt} = k_1[E][S] - k_2[ES] - k_3[ES] = k_1[E][S] - (k_2 + k_3)[ES]$$

$$\frac{d[P]}{dt} = k_3[ES]$$

Answer 2.2

For easy writing, let $e=[E]$, $s=[S]$, $c=[ES]$, $p=[P]$, so the 4 equations transform to:

$$\frac{de}{dt} = (k_2 + k_3)c - k_1es$$

$$\frac{ds}{dt} = k_2c - k_1es$$

$$\frac{dc}{dt} = k_1es - (k_2 + k_3)c$$

$$\frac{dp}{dt} = k_3c$$

The formula of the fourth-order Runge-Kutta method is

$$y_{i+1} = y_i + \frac{1}{6}(k_1 + 2k_2 + 2k_3 + k_4) ,$$

$$t_{i+1} = t_i + \Delta t$$

where

$$k_1 = \Delta t \cdot f(y_i, t_i)$$

$$k_2 = \Delta t \cdot f(y_i + \frac{k_1}{2}, t_i + \frac{\Delta t}{2})$$

$$k_3 = \Delta t \cdot f(y_i + \frac{k_2}{2}, t_i + \frac{\Delta t}{2})$$

$$k_4 = \Delta t \cdot f(y_i + k_3, t_i + \Delta t)$$

For the first equation $\frac{de}{dt} = (k_2 + k_3)c - k_1es$

$$e_{i+1} = e_i + \frac{1}{6}(k_{1e} + 2k_{2e} + 2k_{3e} + k_{4e})$$

$$k_{1e} = \Delta t \cdot ((k_2 + k_3)c - k_1es)$$

$$k_{2e} = \Delta t \cdot f(y_i + \frac{k_{1e}}{2}, t_i + \frac{\Delta t}{2})$$

$$k_{3e} = \Delta t \cdot f(y_i + \frac{k_{2e}}{2}, t_i + \frac{\Delta t}{2})$$

$$k_{4e} = \Delta t \cdot f(y_i + k_{3e}, t_i + \Delta t)$$

Similarly, we can work out for s, c and p

Answer 2.3

The Michaelis-Menten equation is

$$v = k_3[E]_0 \frac{[S]}{K_m + [S]}$$

where,

v the rate at which the product P is formed

k_3 the rate constant for dissociation of the enzyme-product complex

$[E]_0$ the enzyme concentration

$[S]$ the substrate concentration

K_m the Michaelis constant which measures the affinity of the substrate for the enzyme.

Here, $k_3 = 150/\text{min}$, $[E]_0 = 1\mu\text{M}$, $[S] = 10\mu\text{M}$, so

$$v = 150 * 1 * \frac{10}{K_m + 10} = \frac{1500}{K_m + 10}, \text{hence}$$

$$V_m = 150$$