

Abstract

Contents

1	Introduction	1
2	Fundamentals	3
2.1	STM-Imaging	3
2.2	Mathematical Foundation STM	5
2.3	Low Energy Electrons Diffraction (LEED)	7
2.4	Adsorption	8
3	Experimental Setup	9
3.1	Sample Preparation	10
3.1.1	Substrate	10
3.2	Magnesium Oxide (MgO)	10
3.3	Pentacenequinone	11
3.4	2-Hydrogen Phthalocyanine	12
4	Results	13
4.1	Pentacenequinone	13
4.2	2-Hydrogen Phthalocyanin on Ag(100)	14
4.3	Magnesium-Oxide (MgO)	15
4.4	2-Hydrogen Phthalocyanine on MgO	17
5	Discussion	19
	Bibliography	21

Introduction

Since the development of the Scanning Tunneling Microscope (STM) in 1981 it was used to get an insight into electronic structure of molecules and individual atoms in crystals. Especially the adsorption of organic molecules on metal or metal-oxide substrates peaked interest in the STM-Imaging Field in recent years. The reason why there is a lot of interest in studying these organic adsorbents is their promising semiconducting properties and their capacity for self-assembly. Additionally they promise cheap, flexible and tunable alternatives to conventional metal-semiconductors in electronic and optoelectronic Otero et al. (2017).

Fundamentals

2.1 STM-Imaging

The Scanning Tunneling Microscope was introduced in 1981 by Gerd Binning and Heinrich Rohrer. With this measuring technique it is possible to resolve a conductive surface with a precision beyond that of conventional light based Microscopes. In contrast to other electron based microscopy like Scanning Electron Microscopes (SEM) it uses the quantum mechanical phenomenon of tunneling. In classical mechanics, objects cannot overcome a potential if their energy $E < V_0$, as observed in gravitational interactions. This phenomenon is observed for quantum mechanical particles like electrons, which can surpass a potential barrier despite the initial expectation that they should not be able to. The STM uses this effect by precisely positioning a sharp conductive tip close to the surface and applying a bias voltage. Most STM are operated in Ultra-High-Vacuum (UHV), where the distance between the tip and the surface represents the tunneling barrier. By varying the bias voltage, the tunneling probability can be changed, thereby affecting the tunneling current. If the bias voltage, also referred to as the potential difference, is kept constant, the tunneling current is primarily dependent on the distance between tip and surface. The tip is moved in the x,y-plane where a grid is established. There are two modes of operation, the constant-height and the constant-current mode. The latter is especially useful for irregular surfaces, because the tip is moved up and down to keep the tunneling current constant. The movement signal of the piezos is then converted into height. In constant-height mode the position of the tip stays fixed and the tunneling current I_t is measured and converted into height information.

The tunneling current, at a arbitrary gridpoint, is influenced by the electronic structure of the tip and sample. If the tip is in vicinity of the metallic substrate the fermi energies align, resulting in an equal probability of electrons tunneling from the tip to the sample and vice versa. This consequently results in a zero net current. Through the introduction of a electric potential V_{bias} the fermi energies of tip and sample can be shifted relative to each other (Figure 2.1). If the bias voltage is positive the fermi energy of the sample is pushed down and electrons from occupied states in the tip can tunnel into the empty states of the sample. Consequently if the bias voltage is negative electrons from the filled states of the sample tunnel into the tip. The tunneling current is influenced by the distance d between orbitals of the tip and the sample, which makes it possible to gain information about the electronic structure of the sample. This is not really a representation of the real structure of the atoms or molecules, but the Local Density of States (LDOS) of the sample's surface. Utilizing this in Scanning Tunneling Spectroscopy (STS) provides additional information beyond the sample's topography. Such as the chemical composition, bonding, the energy gap and band-bending effects [Bai (1992)].

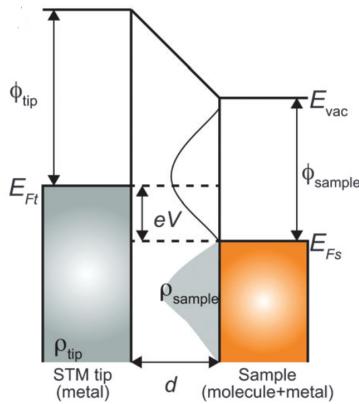


Fig. 2.1: Energy diagram of the tunneling junction with a positive bias voltage applied. Φ_{tip} & Φ_{sample} : working functions of either tip and sample, E_{vac} : vacuum energy level, E_{ft} & E_{fs} : Fermi Energys of the tip and sample, ρ_{tip} & ρ_{sample} : density of states of tip and sample, eV : potential difference caused by applying a bias Voltage V (picture source: Kano et al. (2014))

2.2 Mathematical Foundation STM

To understand the tip sample interaction one must look at the quantum-mechanical Foundation behind it. At its simplest the tip can be approximated as spherical potential well (Figure 2.2). R is in that case the radius of the tip located at position \vec{r}_0 with the distance d from the surface. First order perturbation theory gives the following expression (Eq. 2.1) for the tunneling current of this system [Tersoff and Hamann (1983)]:

$$I = \frac{2\pi e}{\hbar} \sum_{\mu\nu} f(E_\mu)[1 - f(E_\nu + eV_{bias})] \cdot |M_{\mu\nu}|^2 \delta(E_\mu - E_\nu) \quad (2.1)$$

The fermi distribution $f(E) = (\exp((E-E_F)/k_bT)+1)^{-1}$ gives the occupation probability of a fermion (electrons) with the Energy E near the fermi level. In this case it is the occupation probability of the tip states (denoted by the subindex μ) and the occupation probability of the sample states (denoted by the subindex ν). The tunneling matrix $M_{\mu\nu}$ is related to the derivatives of the sample wave functions ψ_ν at the nucleus of the apex atom [Chen (1990)]. Because the STM Imaging is done at low temperatures and with small voltages the Equation 2.1 can be simplified to:

$$I = \frac{2\pi}{\hbar} e^2 V_{bias} \sum_{\mu\nu} |M_{\mu\nu}|^2 \delta(E_\nu - E_F) \delta(E_\mu - E_F) \quad (2.2)$$

With the tunneling matrix $M_{\mu\nu}$, which relates to the probability of an electron transitioning from state μ to state ν [Chen (1990)]:

$$M_{\mu\nu} = \frac{\hbar^2}{2m} \int (\psi_\mu^* \vec{\nabla} \psi_\nu - \psi_\nu \vec{\nabla} \psi_\mu^*) d\vec{S} \quad (2.3)$$

To obtain a solution , the tip wave function can be explicitly chosen as an s-type orbital, this holds true for a tungsten tip as it has a only a s-type orbital occupied in its valence shell. This orbital has a spherical form [Tersoff and Hamann (1983)]:

$$\psi_\mu = \Omega_t^{-1/2} * c_t k R e^{kR} (k|\vec{r} - \vec{r}_0|)^{-1} e^{-k|\vec{r} - \vec{r}_0|} \quad (2.4)$$

$\Omega_t^{-1/2}$ is the volume of the probe, $k = \hbar^{-1} \sqrt{2m\phi}$ is the inverse dacay length for the wave functions in vacuum, c_t is a geometry specific constant on the order of 1 and R is the radius of curvature. The surface wavefunctions can be approximated as

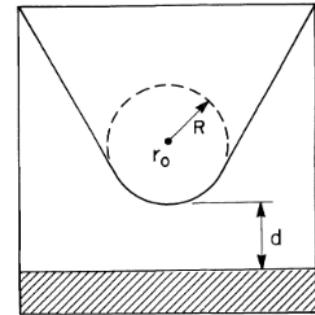


Fig. 2.2: Schematic depiction of the tip geometry [Tersoff and Hamann (1983)]

blochwaves, which are plane waves modulated by a periodic function. In the region of negligible Potential the surface wavefunction can be written as:

$$\psi_\nu = \Omega_s^{-1/2} \sum_G a_G \exp(-(k^2 + |\vec{k}_{||} + \vec{G}|^2)^{1/2} z) \cdot \exp(i(\vec{k}_{||} + \vec{G}) \cdot \vec{x}) \quad (2.5)$$

Here $\Omega_s^{-1/2}$ is the sample volume, k the previously mentioned inverse decay length, $\vec{k}_{||}$ is the surface bloch wave vector and \vec{G} represents the reciprocal surface vector. Substituting into 2.2 as seen in executed in [Tersoff and Hamann (1983)] gives the result:

$$I = 32\pi^3 e^2 V_{bias} \phi^2 D_t(E_F) R^2 k^{-4} e^{2kR} \cdot \sum_\nu |\psi_\nu(\vec{r}_0)|^2 \delta(E_\nu - E_F) \quad (2.6)$$

Where $D_t(E_F)$ is the Density of States of the Tip at the Fermi Level. The sum over the local states of the surface ψ_ν at the Fermi Level ($\delta(E_\nu - E_F)$) can be interpreted as the Local Density of States (LDOS). Using the surface wavefunction one can see that the tunneling current is following an exponential decay, that is proportional to the distance d between the tip and the sample:

$$I \propto e^{-2 \cdot d \sqrt{\frac{2m\phi}{\hbar^2}}} \quad (2.7)$$

2.3 Low Energy Electrons Diffraction (LEED)

A well collimated beam of electrons with the same wavelength is directed at a flat surface. In this case the wavelength of the incoming electrons is determined by the electric field U_a that acts upon them 2.8. To see diffraction the wavelength of the electrons must be in the range of the lattice constant, as is effectively shown by the concept of the Ewald sphere. Typically the kinetic energy of the electrons is up to $E_{kin} = eU_a = 200$ eV.

$$\lambda = \frac{h}{\sqrt{2m_e e U_a}} \quad (2.8)$$

If the surface is periodic this results in a sharp pattern of spots that reflects the k-space of the crystal.

The LEED setup consists of a filament which emits electrons by applying a current. These electrons are then accelerated and collimated in a drift tube which points at the sample. Due to interactions of the electrons with the crystal atoms some of them are not diffracted elastically. To ensure that only the elastically scattered electrons are shown on the fluorescent screen, a grid system is installed in front of the screen. The first grid and the sample are both grounded to ensure that the diffracted electrons propagate undisturbed through the space between the sample and the grid. A filter voltage is applied between the second and third grid that is just below the acceleration voltage. Electrons which are not elastically scattered are completely stopped, which makes it effectively a high pass filter. [Moritz (2022)]

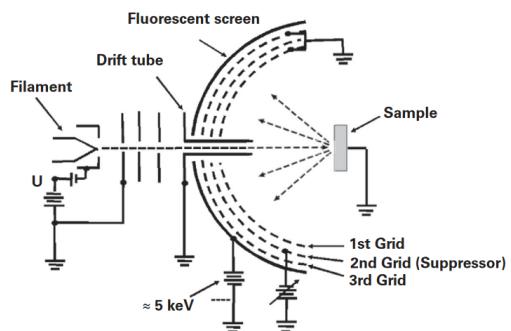


Fig. 2.3: A basic three-grid LEED system which utilizes a fluorescent screen as imaging tool Moritz (2022)

The collimated beam can be represented as a planar wave $e^{i\vec{k}_0 \cdot \vec{r}}$ outside the surface. The wavevector \vec{k}_0 points in the propagation direction and represents the spacial frequency of the wave. This vector can be modeled as:

$$|\vec{k}_0| = \frac{2\pi}{\lambda} = \frac{\sqrt{2m_e E}}{\hbar} \quad (2.9)$$

$$\vec{k}_0 = \begin{pmatrix} |k_0| \cos(\varphi) \sin(\vartheta) \\ |k_0| \sin(\vartheta) \sin(\vartheta) \\ |k_0| \cos(\vartheta) \end{pmatrix} \quad (2.10)$$

A 2D periodic surface structure with a reciprocal basis (b_1, b_2) causes the plane waves to diffract. The diffracted waves have the shape $A_g e^{i\vec{k}_g \cdot \vec{r}}$ with wave vectors $|\vec{k}_g| = |k_0|$. \vec{k}_g transforms according to $\vec{k}_g = \vec{k}_0 + \vec{g}$. The reciprocal lattice vector $\vec{g} = 2\pi(hb_1 + kb_2)$ represents the individual points where constructive interference is present. Usually the indices h and k are integers which characterize the beam, for example $(0,0)$ or $(1,1)$. It can be shown that the wave vector in vacuum outside the surface is [Moritz (2022)]:

$$k_g = -\sqrt{\frac{2m_e E}{\hbar^2} - (k_{0x} + g_x)^2 - (k_{0y} + g_y)^2} \quad (2.11)$$

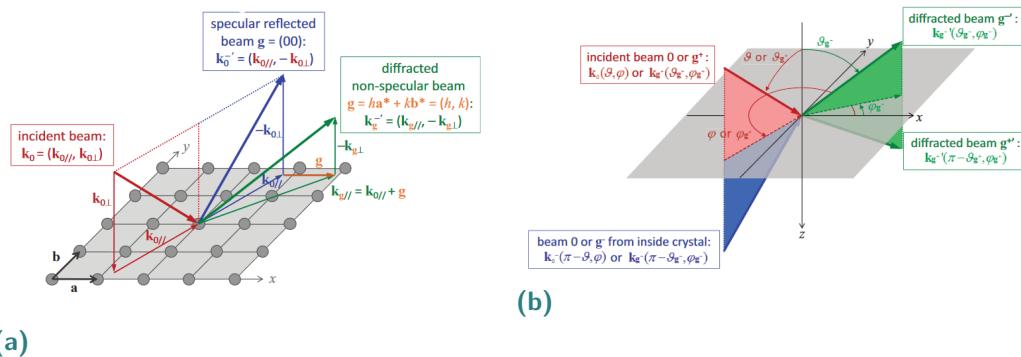


Fig. 2.4: (a): 2D lattice with the reciprocal lattice vectors \vec{a} and \vec{b} . \vec{k}_0 (red): incident beam, \vec{k}_0^- (blue): specular reflected beam, \vec{k}_g (green): non-specular diffracted beam,

2.4 Adsorption

Experimental Setup

The experimental work done in this thesis primarily utilized a Low-Temperature Scanning Tunneling Microscope setup (LT-STM). This STM is composed of two distinct compartments, the preparation chamber (PC) the measurement chamber (MC). The sample is inserted through a airlock into the preparation chamber. In the whole system there is in a Ultra High Vacuum (UHV) at about $10^{-11} - 10^{-10}$ mbar, which is archieved by four individuall pumps. The base Vacuum is achieved with the turbomolecular pump and the scroll pump through the airlock. Additionally there is a titanium sublimation pump and a ion pump in the preparation room.

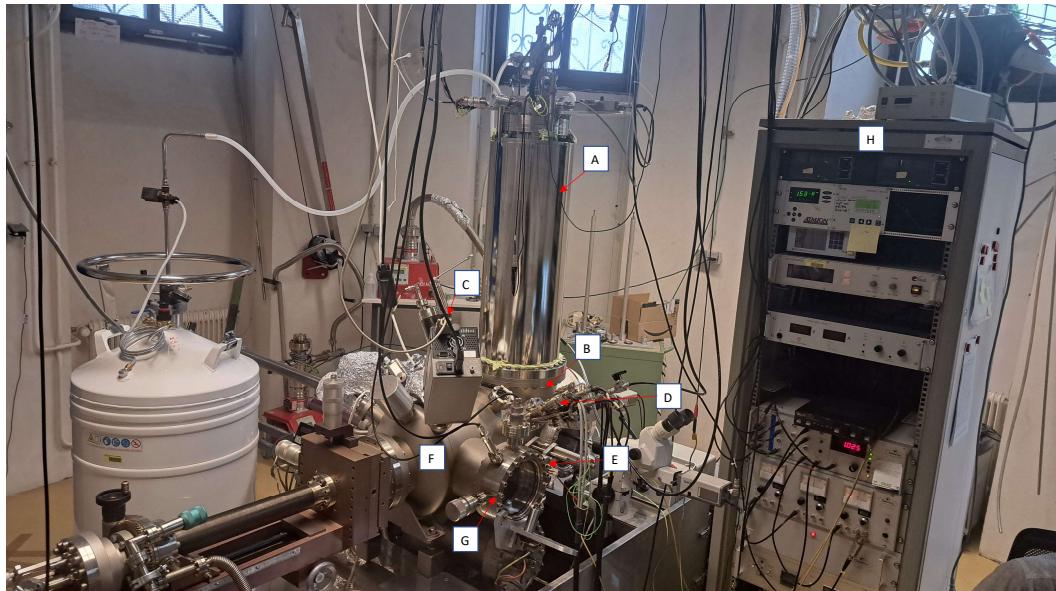


Fig. 3.1: The experimental setup used, A: cooling-chamber filled with Liquid Oxygen, B: measuring-chamber with spring suspended sample holder and tip, C: sputter-gun, D: metal-evaporator, E: molecule-evaporator, F: preparation-chamber with free movable sample holder arm, G: LEED system, H: Electronic used to monitor the function of the STM

The Probe-holder can be moved in each direction and rotated using an integrated arm in the PC. It is also used to insert the Probe-holder into the MC. The PC is

equipped with a LEED system, which fluorescent screen can be extended. Additionally the PC has a metal-evaporator and organic-molecule evaporator (E-beam evaporator with a triple Knudsen cell) which are used to deposit submono- or monolayer structures onto an circular sample (10 mm x 2 mm). This sample is mounted onto the previously mentioned Probe-holder, a Heatwaves Labs Inc. button heater with a resistive heating range of 20 K - 900 K. The sample can also be cooled using LN₂ or LHe through built-in pipe system in the manipulator arm. A Quartz-Crystal Microbalance (QCM) with sub Angström precision is used to monitor the deposition thickness. To clean the sample, an Argon sputter gun is employed, which utilizes an electric field to ionize and accelerate the argon gas. The MC consists of a LT-STM which is surrounded by a two-shell cryostat, to which it is fixed, and two radiation shields to achieve temperatures as low as 7 K. In this thesis the cryostat chamber was filled with LN₂ which means it was operated at 75 K. The primary measuring unit (tip and sample storage place) is vibrationally damped using a coupled spring system.

3.1 Sample Preparation

3.1.1 Substrate

The Substrate used in this thesis is a Silver (Ag) single crystal. Silver is a metal with 47 Protons and two stable isotopes ¹⁰⁷Ag and ¹⁰⁹Ag that are almost equally distributed. It has an atomic radius of 1.44 Å and it naturally crystallizes into a face centered cubic (fcc) structure with a lattice constant of $a = 4.079 \text{ \AA}$. [Davey (1925)] The single crystal is cut along the (100) plane giving it a square surface lattice along the [011] and [01̄1] Direction. To insure that the surface is clean, multiple cycles of Argon-Ion Sputtering with consequent annealing were performed. The first preparation was done with 3x Sputtering for 5 min each and annealing at 500 C° for 3-5 min. Later a longer sputtering time of 20 min proved more effective. Prior to the metal or organic adsorption the substrate is cooled to room temperature.

3.2 Magnesium Oxide (MgO)

The crystal structure of MgO is also fcc with a lattice constant of $a = 4.22 \text{ \AA}$ [Guilliatt and Brett (1969)], which makes it only 3.4 % larger than Ag. This allows the formation of somewhat larger terraces, which often break in rectangular shapes. This comes from the cubic structure of MgO. The preparation on Ag(100) is done using a

E-beam evaporator, in which a filament is heated, causing it to emit electrons. This electrons are than accelerated using a high voltage (250-300V) onto the Magnesium probe. The product of the emission current and acceleration voltage is the heating capacity. To ensure a monolayer the deposition rate is measured using a Quarz-Crystal Microbalance (QCM). A deposition rate of $1.33 \cdot 10^2$ digit/s is established. One digit is than deposited onto the Ag(100) substate, while Oxygen O₂ is being fed into the Preparation Chamber (10^{-6} mbar). After the monolayer is deposited, the sample is being slowly cooled at a cooling rate of 5 K/min. This ensures larger terrases as MgO is known to break into small islands if cooled to fast.

3.3 Pentacenequinone

Pentacenes are a organic molecule group with five linearly organised benzene rings, which share two or four neighboring carbon atoms. The carbon atoms are sp² hybridised, which gives it a hexagonal structure in which the single- and double-bond is alternating. At the double-bond location the π orbital extends orthogonal to the plane in which the benzene rings are aligned. The overlapping of neighbouring π orbitals results in delocalized electrons across the molecule.

Quinone refers to the replacement of two hydrogen atoms with two oxygen atoms in a benzene ring. The two carbonyl bonds (C=O) in the observed adsorbant are located in the central ring. Pentacene is an organic p-type semiconductor, which is used in the domain of organic electronics. Its remarkably stable and has a high charge mobility for an organic molecule of $2.2 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$ [Mota et al. (2018)]. Pentacenequinone exhibits similar properties and is used as a precursor for the synthetization of pentacene.

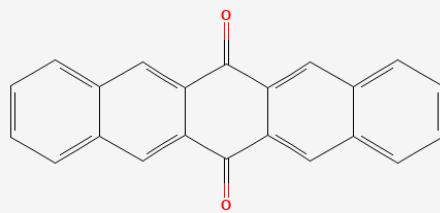


Fig. 3.2: Pentacenequinone PubChem *National Center for Biotechnology Information* (2023). PubChem Compound Summary for CID 76415, 6,13-Pentacenequinone. (N.d.)

3.4 2-Hydrogen Phthalocyanine

Phthalocyanines are a group of organic molecules that exhibit a variety of properties. They are characterized by a symmetric macrocyclic structure of four isoindole units $[(C_6H_4)C_2N]$, which consists of a pyrrole unit fused with a benzene ring. The isoindole units are linked by four nitrogen atoms (aza-bridges) forming a conjugated chain. This allows a delocalization of the π electrons along the alternating double- and single bond macrocycle.

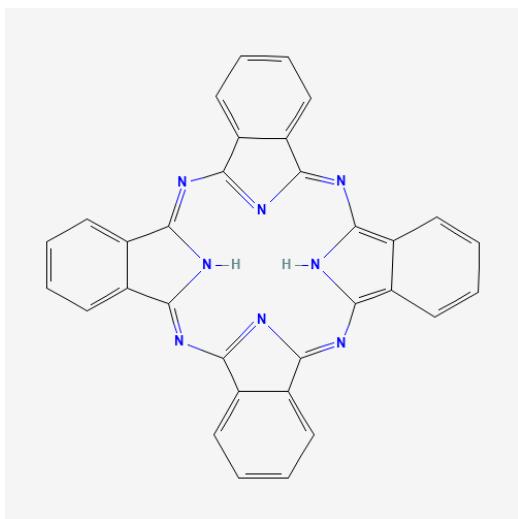


Fig. 3.3: 2-H Phthalocyanine PubChem National Center for Biotechnology Information (2023). PubChem Compound Summary for CID 86280045, Phthalocyanine (n.d.)

There are more than 70 variations of the phthalocyanine, with the key difference in the substitution of different elements or transition metals (e.g. Cu^{2+} and Ni^{2+}) in the middle of the molecule. The physical properties can be strongly altered by the choice of different metal ions, like the adsorption characteristics or the geometry of the molecule itself. In this thesis the 2-Hydrogen Phthalocyanine is observed, which harbors, as the name already tells, two hydrogen atoms in the middle attached to two opposing pyrrole units. Generally Phthalocyanines are remarkably chemically and thermally stable. This can be attributed to its planar geometry and conjugation in the macrocycle. Phthalocyanines are used as dye for textiles and as inks because they strongly absorb light in the

visible range especially in 620 - 750 nm regime, which makes them appear dark green-blue. Its also known for its optical, electronic and photo-electronic properties, all of which makes it a promising sensor chemiresistor. Because of the semiconducting properties of phthalocyanines, its capability as a transistor, its use as liquid crystal display or its solar-cell applications are investigated.

Results

4.1 Pentacenequinone

Unfortunately inspite of three attempts to get a stable evaporation process it was not possible to get sub-monolayer/monolayer growth. This lead to difficulties regarding image resolution, but it was still possible to observe makrostructures forming. Pentacenequinone seems to form round island that are randomly scattered and have no clear periodic structure. If looking closely at the edges of the terasses, it is visible that the sharp edges that are typical for Ag(100) are rough. This could indicate a full monolayer present.

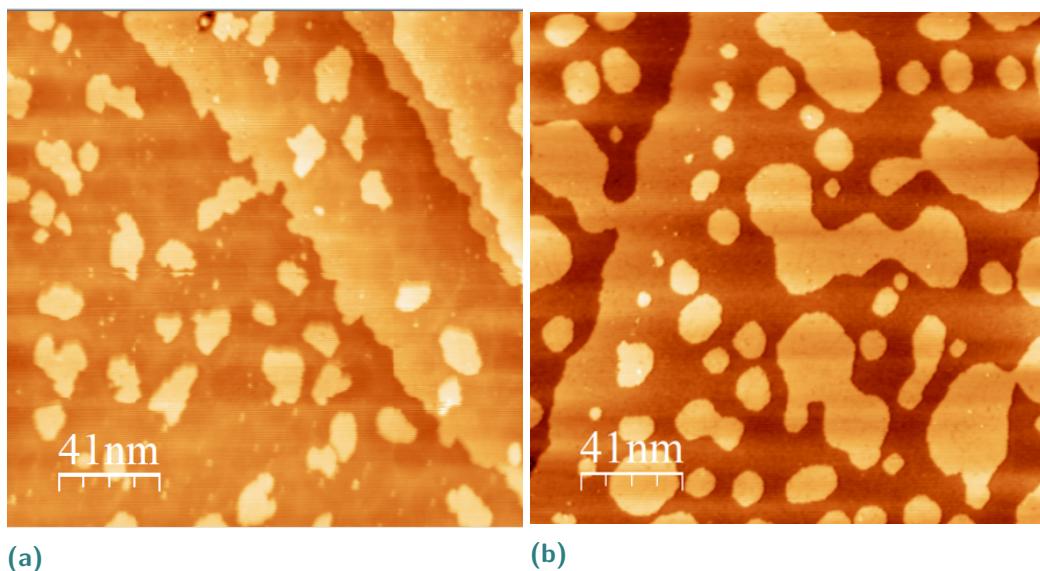


Fig. 4.1: seas

4.2 2-Hydrogen Phthalocyanin on Ag(100)

The STM and LEED Images of 2-Hydrogen Phthalocyanin (2HPc) on Ag (100) show the Formation of an ordered Structure. As seen in the Image 4.2 the 2HPc is most likely in $\sqrt{17} \times \sqrt{17}\text{R}14$ Phase with an relative rotation of the molecular axis (MA) in respect to the [010] direction of 7 ° (see figure ??). Furthermore the STM Images reveal that the distance between the nearest molecules is 15 Å and that the unit cell of the monolayer is also square.

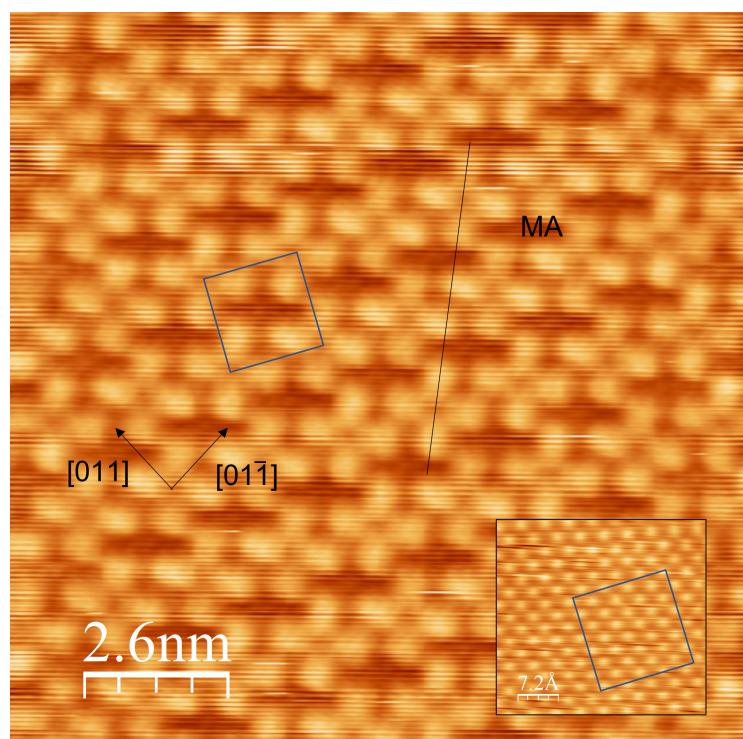


Fig. 4.2: STM image of 2HPc/Ag(100) taken in constant-current mode (monolayer regime).(a) Constant-height Image of Ag(100) as reference . U = 53.5 mV , I = 1 nA. MA: Molecular axis.

The Molecule itself appears as a 4-fold symmetric cross with a dark center in the middle or as 2-fold symmteric cross with two opposing isoindole units seeming brighter than the other two for positive polarity . It oriented so that the benzene ring of one 2HPc molecule is pointed towards the benzene ring of the nearest neighbor. For negative biasvoltages (filled state imaging), near the fermi level a two fold symmetric orbital is seen.

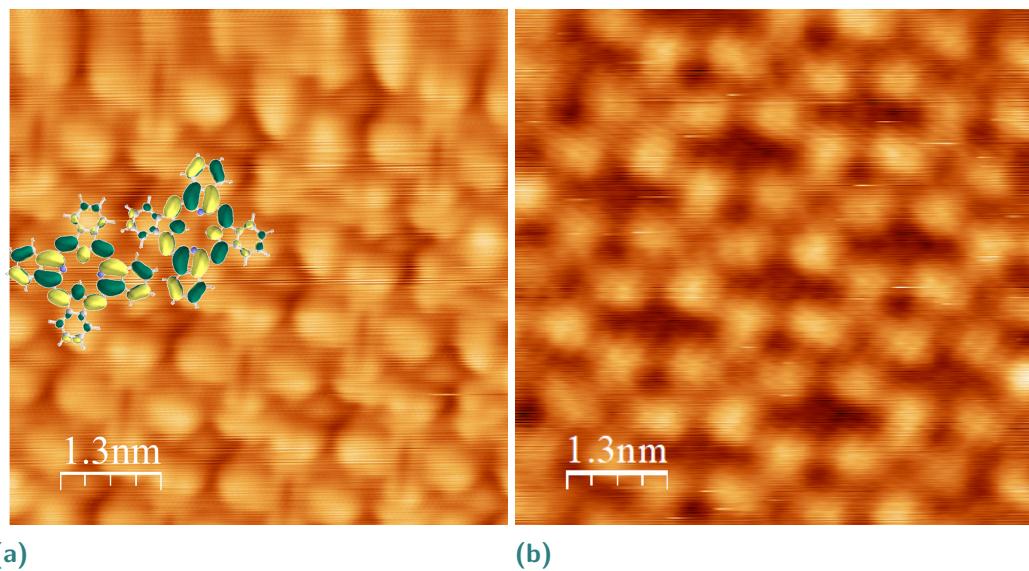


Fig. 4.3: (4.3a) Constant-current of 2Hpc/Ag(100). Two distinct LUMO orientations can be seen and their DFT-Simulation (xxx Puschnig ref) counterparts. $U = -50.0 \text{ mV}$, $I = 0.05 \text{ nA}$. (4.3b) Same spot imaged in constant-current mode at positive Bias-Voltages. $U = 750 \text{ mV}$, 0.02 nA .

This is most likely due to charge transfer from the metal substrate , leading to an occupied LUMO and thus a charged molecule. There are two degenerated LUMOs (same Energy), with the probability density concentrated around two opposing isoindole units. This makes the Molecule reduce its symmetry from 4-fold to 2-fold. One difference of the two degenerated LUMOs is the orientation of the probability density, as it is rotated 90° in respect to eachother. The STM images (see Fig.4.3a) reveal a connection beween two opposing isoindole units, which is in agreement with the simulated orbitals using density functional theory (DFT). There is no favored orientation seen in the STM images, it is seemingly random which suggest that the occupation probability of the two LUMOs is the same. Its worth mentioning that the LUMO is also visible for positive bias which suggest two things: the negatively charged molecules LUMO is singly or fractionaly occupied and therefore it is possible for a electron to tunnel into that state or it tunnels into the other not occupied LUMO. The latter is suported by the

4.3 Magnesium-Oxide (MgO)

Before depositing 2Hpc (2H-phthalocyanine) onto the MgO crystal structure, the quality of the surface was assessed using scanning tunneling microscopy (STM), followed by capturing a low-energy electron diffraction (LEED) image. The MgO forms sharp rectangular Islands with the most edges along the [011] and [01 $\bar{1}$]

directions as expected. Due to a similar lattice constant and an equivalent fcc crystal lattice it can be assumed that there are local coherent phase boundarys. It is shown in the figures 4.4a & 4.4b that most of the islands were 15 - 20 nm wide. The Images show a adsorbtion pattern consistent with literature.

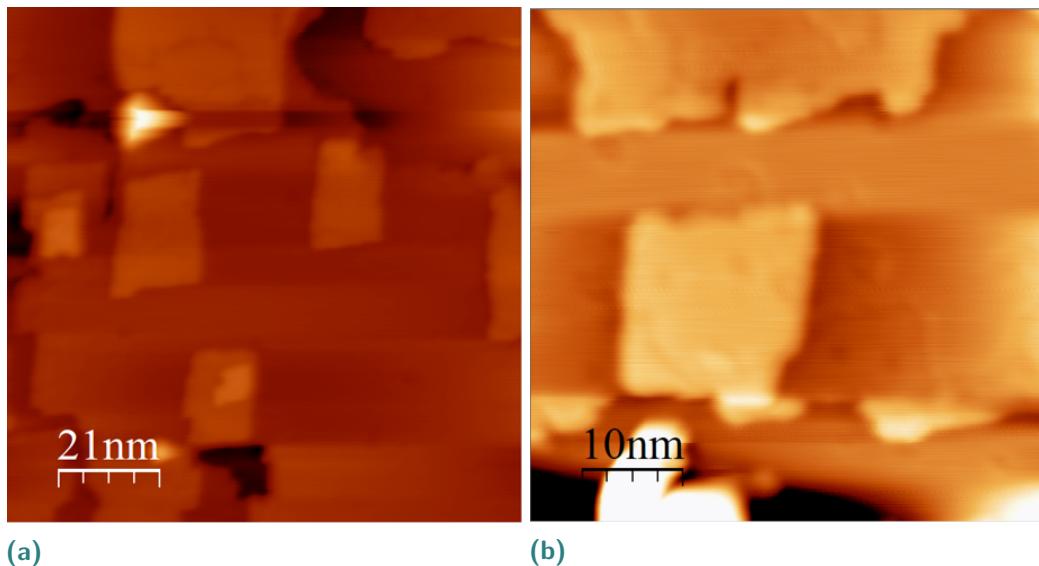


Fig. 4.4: Both (a) and (b) taken at 3.5 V Biasvoltage, most likely sub-monolayer to monolayer regime

4.4 2-Hydrogen Phthalocyanine on MgO

The STM images of 2Hpc on MgO(100)/Ag(100) show again the formation of ordered structure that persists across the the whole crystal lattice. Evident is the presence of linear defects, on which sites the 2Hpc Phase is shifted or even disconnected. This is most likely due to the fact that the underlying MgO long-range order is disturbed by the 3.2 % mismatch of the Ag an MgO lattice. Therefore resulting in incoherent sites, which give rise to intermediate phases with different orientation. In figure 4.5 the phase of the molecular adsorbant was estimated. This was done using the Ag(100) lattice vectors, as it almost identical to the MgO lattice, where Oxygen is adsorbed on top of the Ag atoms. The unit cell of 2HPc on MgO is tilted in direction of the [011] plane and is also larger than on Ag which can be attributed to the the larger lattice constant. The phase was estimated to $\sqrt{22.5} \times \sqrt{22.5}R(-18.4)$ in respect to the Ag(100) lattice, where the angle measurement is in respect to the [010] plane. Note that this was the most common phase and the large scale distribution of different molecular species is not known. In contrast to 2Hpc on Ag the molecular is almost mirrored in respect to the [001] plane with an molecular axis angle of -6°.

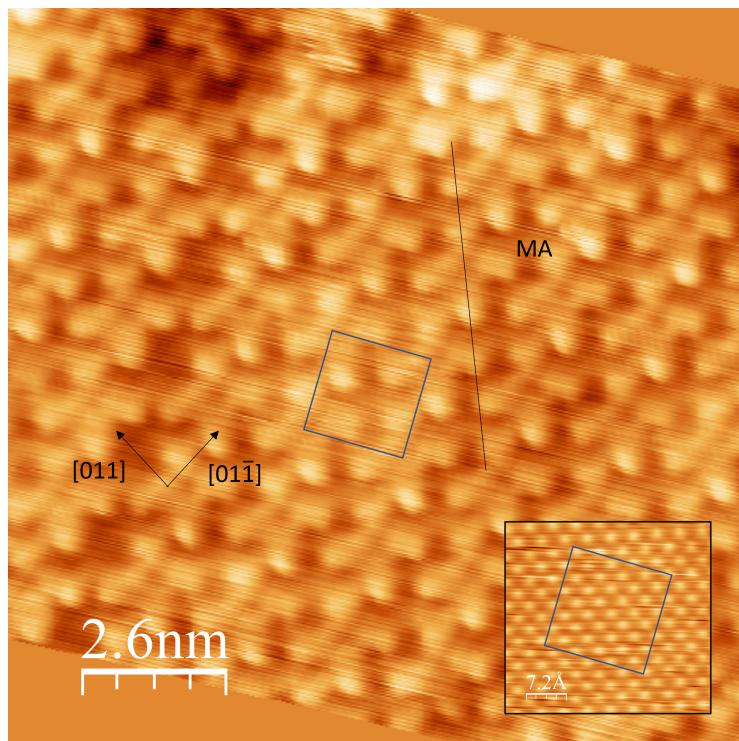


Fig. 4.5: STM image of 2HPC on MgO(100)/Ag(100) taken in constant-current mode (monolayer regime). (a) Constant-height Image of Ag(100) as reference . U = 200 mV , I = 1 nA. MA: Molecular axis.

For further verification of the observed structure a LEED picture at 15 eV was taken after analysis with the STM. The LEED clearly supports the mainly observed phase, but looked diffused which can come from additional phases or additional diffraction caused by the molecular structure. Additionally a 2D-FFT was performed on the the image in figure 4.5 which alignes nicely with the observed diffraction.

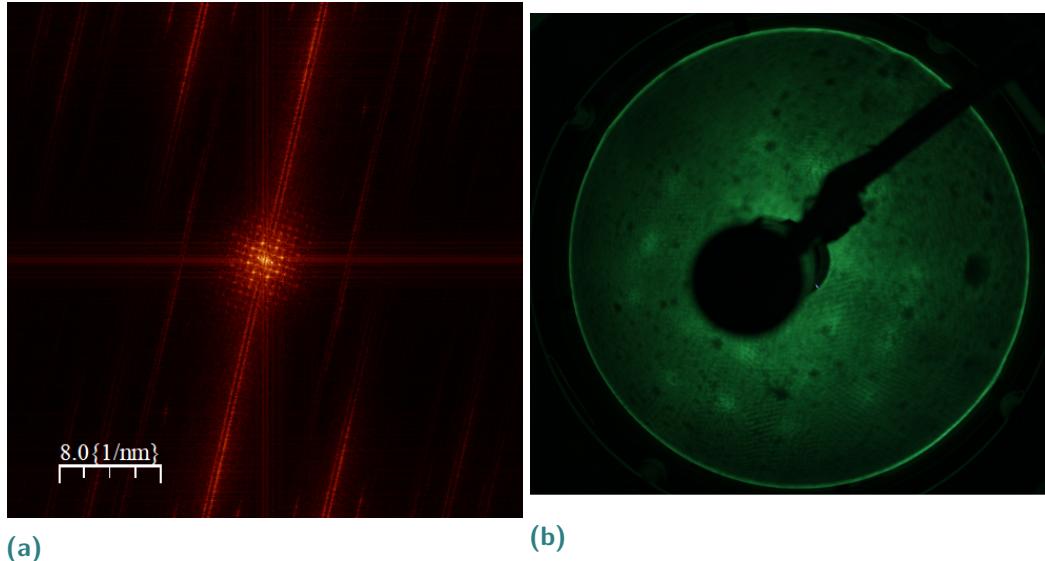


Fig. 4.6: (a) 2D-FFT of image in 4.5, (b) LEED picture taken at $E_{kin} = 15$ eV

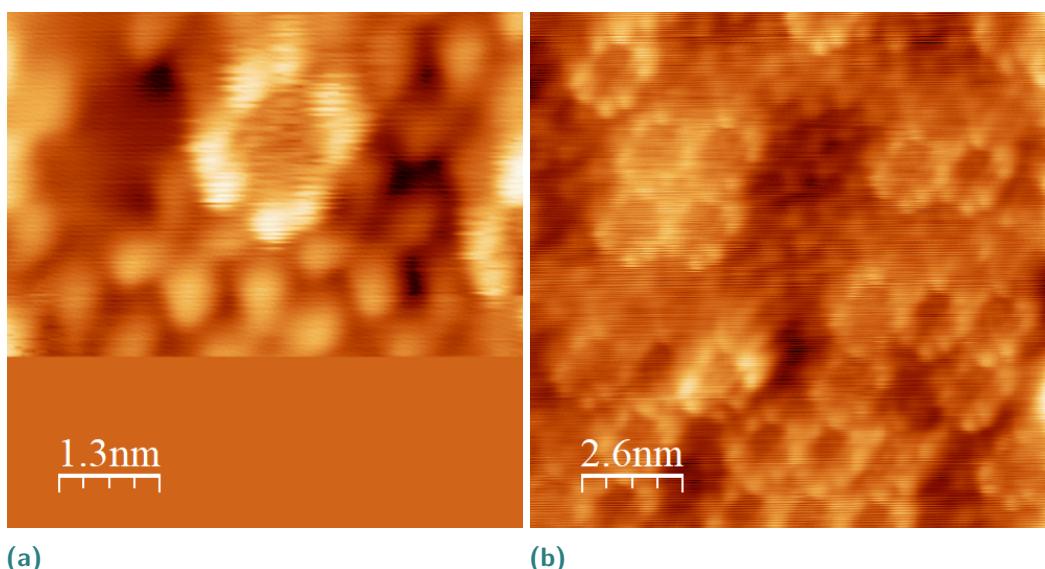


Fig. 4.7: (a)

Discussion

seas

Bibliography

- Bai, Chunli (1992). *Scanning Tunneling Microscopy and Its Application*. Ed. by Springer. Shanghai Scientific and Technical Publishers (cit. on p. 4).
- Chen, Julian (Dec. 1990). „Tunneling matrix elements in three-dimensional space: The derivative rule and the sum rule“. In: *Physical review. B, Condensed matter* 42, pp. 8841–8857. DOI: [10.1103/PhysRevB.42.8841](https://doi.org/10.1103/PhysRevB.42.8841) (cit. on p. 5).
- Davey, Wheeler P. (June 1925). „Precision Measurements of the Lattice Constants of Twelve Common Metals“. In: *Phys. Rev.* 25 (6), pp. 753–761. DOI: [10.1103/PhysRev.25.753](https://doi.org/10.1103/PhysRev.25.753) (cit. on p. 10).
- Guilliatt, IF and NH Brett (1969). „Lattice constant variations in finely divided magnesium oxide“. In: *Transactions of the Faraday Society* 65, pp. 3328–3333 (cit. on p. 10).
- Kano, Shinya, Tsukasa Tada, and Yutaka Majima (Oct. 2014). „Nanoparticle characterization based on STM and STS“. In: *Chem. Soc. Rev.* 44. DOI: [10.1039/C4CS00204K](https://doi.org/10.1039/C4CS00204K) (cit. on p. 4).
- Moritz, Wolfgang (2022). *Surface Structure Determination by LEED and X-rays*. eng. Cambridge: Cambridge University Press (cit. on pp. 7, 8).
- Mota, María L., Bibiana Rodriguez, Amanda Carrillo, et al. (2018). „Fast and inexpensive synthesis of pentacene with high yield using 6,13-pentacenequinone as precursor“. In: *Journal of Molecular Structure* 1154, pp. 511–515. DOI: <https://doi.org/10.1016/j.molstruc.2017.10.070> (cit. on p. 11).
- National Center for Biotechnology Information (2023). PubChem Compound Summary for CID 76415, 6,13-Pentacenequinone. (N.d.). URL: https://pubchem.ncbi.nlm.nih.gov/compound/6_13-Pentacenequinone (cit. on p. 11).
- National Center for Biotechnology Information (2023). PubChem Compound Summary for CID 86280045, Phthalocyanine (n.d.). URL: <https://pubchem.ncbi.nlm.nih.gov/compound/Phthalocyanine> (cit. on p. 12).
- Otero, R., A.L. Vázquez de Parga, and J.M. Gallego (2017). „Electronic, structural and chemical effects of charge-transfer at organic/inorganic interfaces“. In: *Surface Science Reports* 72.3, pp. 105–145. DOI: <https://doi.org/10.1016/j.surfrept.2017.03.001> (cit. on p. 1).
- Tersoff, J. and D. R. Hamann (June 1983). „Theory and Application for the Scanning Tunneling Microscope“. In: *Phys. Rev. Lett.* 50 (25), pp. 1998–2001. DOI: [10.1103/PhysRevLett.50.1998](https://doi.org/10.1103/PhysRevLett.50.1998) (cit. on pp. 5, 6).

