# Operating Parameters

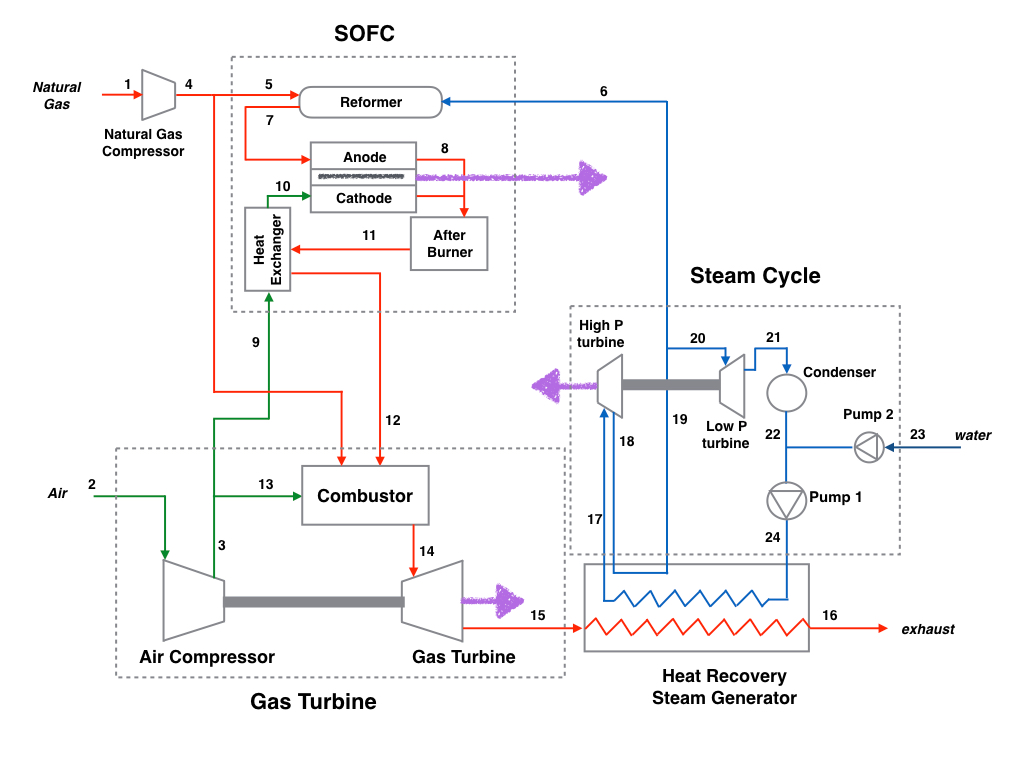


Figure System to be analysed

## Summary

|  |  |  |
| --- | --- | --- |
| Position | Temp (K) | Pressure (bar) |
| 1 | 288.15 | 1.013 |
| 2 | 288.15 | 1.013 |
| 3 | 692 | 23.299 |
| 4 | 536 | 23.299 |
| 5 | 536 | 23.299 |
| 6 | 838.65 | 23.88654 |
| 7 | 1073.1 | 23.299 |
| 8 | 1173.1 | 23.299 |
| 9 | 692 | 23.299 |
| 10 | 1073.15 | 23.299 |
| 11 | 1940.9 | 23.299 |
| 12 | 1737 | 23.299 |
| Position | Temp (K) | Pressure (bar) |
| 13 | 692 | 23.299 |
| 14 | 1852.9 | 21.43508 |
| 15 | 1011.529 | 1.039338 |
| 16 | 428.15 | 1.02313 |
| 17 | 838.65 | 123.586 |
| 18 | 586.7338 | 23.88654 |
| 19 | 838.65 | 23.88654 |
| 20 | 838.65 | 23.88654 |
| 21 | 513.4686 | 2.36029 |
| 22 | 292.15 | 2.006247 |
| 23 | 283.15 | 1.013 |
| 24 | 321.15 | 130.0905 |
| 25 | 536 | 23.299 |

## Analysis

The following parameter information was given by the coursework sheet:

|  |  |  |
| --- | --- | --- |
| Position | Temp (K) | Pressure (bar) |
| 1 | 288.15 | 1.013 |
| 2 | 288.15 | 1.013 |
| 15 | 0 | 1.039338 |
| 16 | 428.15 | 1.02313 |
| 17 | 838.65 | 123.586 |
| 18 | 0 | 23.88654 |
| 19 | 838.65 | 0 |
| 21 | 0 | 2.36029 |
| 22 | 292.15 | 0 |
| 23 | 283.15 | 1.013 |
| 24 | 321.15 | 0 |

|  |  |
| --- | --- |
| Position | Mass (kg/s) |
| 1 | 23.9 |
| 2 | 598.7 |
| 5 | 17.9 |
| 13 | 254.7 |

Mass was calculated from the last 2 digits of CID following the prescribed formula.

Flow 25 was defined as the flow from the NG compressor straight to the combustor.

### Steam cycle

*Assuming all pipes have no pressure loss and all interconnected pipes share intensive properties and composition*. Starting at 22 we determine the pressure of the state from calculating the outlet pressure of the condenser given a pressure loss of 15% (*given by coursework*) and the pressure before condenser at 21. . Then given that pump 2 is isentropic (*given by coursework*) using isentropic relations the temperature at 23 was determined from state 23 with:

Pressure at 24 was calculated by inverting the pressure loss of 5% across the HRSG along flow 24-17. .

Given state 17, state 18 was determined through the polytropic efficiency (*89% given by coursework*). This was converted to isentropic efficiency using a reheat factor of 1.07 (relationship derivation and validity of value chosen for steam turbine given by [ref]) .

Using the isentropic efficiency/enthalpy definition for turbines , the enthalpy of the product was determined. Using the enthalpy determined, coupled with the pressure at 18 (*given by coursework*) the temperature at 18 was found using a steam table.

The pressure at position 19 was determined bycarrying through the pressure at 18 (*assuming no pressure loss as it is superheated vapour and the gas side experienced nearly no pressure loss across a much larger temperature range*).

Using the completed state of 19, state 20 and state 6 were determined to have the same properties (*assuming connected pipes share intensive properties*).

Using state 20, the isentropic efficiency was determined using the reheat factor (again taken as 1.07) and polytropic efficiency (*given as 89%*) and used again to determine output enthalpy and thus determine output temperature, given the pressure of 21 being supplied by the coursework.

### Natural Gas cycle

As natural gas is not a perfect gas ( varies throughout the approximate temperature range, tested through GASEQ) GASEQ was used to adiabatically compress the natural gas (*which was taken as 91.3% Methane, 5.36% Ethane, 2.83% Nitrogen, 0.5% and 0.01% – this was determined through filling the miscellaneous components by looking at major contributions from [ref])* to 23 and thus determine state 4. There was insignificant specific entropy change [fig] and as such GASEQ was used to model the isentropic compression (*given by the coursework*). Given state 4, state 5 and 6 were taken to have the same intensive properties and composition due to being connected. Given state 6 (*calculated as above*) the composition of flow 5 and 6 was put into GASEQ, following a specified molar ratio of steam: carbon (*given as 3 by coursework*) at an isothermal temperature () and equilibrium was calculated for state 7. The pressure of 5 was used as the pressure of 6 was within 2% of the value and as such the resultant pressure would stay approximately the same.

Given state 10 (*calculated as below*) the reaction at the fuel cell enabled deduction of state 8. Using the overall chemical equilibrium equation below, the composition of the flows was changed.

This was achieved by determining an incoming molar flowrate of hydrogen reacting (*mass fraction given by GASEQ, overall mass flowrate of 7 calculated as below and molar weight of hydrogen used to convert*) and multiplying it by the fuel utilisation (*given by coursework*). Given the molar flowrate of hydrogen reacting a molar flowrate of oxygen and water used/produced in the reaction by multiplying by relative moles from above eqn. This was then subtracted (for and ) or added (for ) from the molar flowrate of each respective constituent (obtained as the hydrogen mole flowrate was above but in stream 10 for and stream 7 for and ) coming in. As all composition analysis was done fractionally in GASEQ, molar flowrate fractions were used to make sure all ratios put in the new set of products had the correct ratios. E.g. was used to determine the relative moles of as a multiple of the in the output stream. The fuel cell was assumed to be constant pressure and constant temperature to determine state 8.

It was assumed that there was insufficient time between leaving the fuel cell and entering the combustor for the products to reach equilibrium – thus the molar ratios calculated as above were taken as the incoming composition. GASEQ was used to analyse the composition at 8 as the reactant side of an adiabatic equilibrium at constant pressure (*assuming sufficiently slow burning that it can be modelled as a const. pressure combustion*) at the input state given by position 8. This determined the output temperature and fixed state 11 as well as the composition.

The specific enthalpy of position 11 was determined from GASEQ given pressure and temp. and the enthalpy at state 12 was determined by using the insulated nature of the heat exchanger (*given by coursework*) and thus balancing the rise in actual enthalpy of flow 9 to 10 (*as calculated below*) against flow 11 to 12 (*assuming no work in the heat exchanger*). Given a specific enthalpy (*dividing actual enthalpy by mass flowrate of 11 as determined below*) and pressure at state 12 (*assuming no pressure loss across the heat exchanger*), GASEQ was used to manually iterate state 12 temperature given the composition to achieve the specific enthalpy decrease required.

### Air cycle

Given state 2 (*by coursework*) – state 3 was determined using GASEQ, with an adiabatic compression for a pressure ratio (*given by coursework*) and checking the change in specific entropy is negligible [fig] to confirm the isentropic resultant nature. Although technically a perfect gas it was decided that using GASEQ would be more accurate as ultimately there is no real perfect gas and as such it is was assumed it is better to avoid perfect gas relations unless necessary. The composition of air was assumed to be 78.08% , 20.95% and 0.93% [source].

Given state 3 the intensive properties of state 9 and 13 were assumed to be the same as the share a pipe connection.

The molar composition of flow 13, 25 and 12 was saved into a mixture in GASEQ and then using mass flowrates for each (*as calculated per below*) molar flowrates were determined through division of by mean molecular weight, of each flow mixture and used to determine the ratio of moles of each flow in GASEQ. The flows were then set to their respective temperatures in GASEQ and an adiabatic process at constant pressure was performed to determine the chemical composition and state of 14. This assumes the combustion is slow enough that the process occurs at constant pressure and that the pressure loss due to the combustor happens after the combustion, perhaps due to increased turbulence from combustion and head loss as it travels through pipes.

State 15 was determined by using the polytropic efficiency (*given by coursework*) and the perfect gas polytropic efficiency relationship (*given by Appendix of ME3 TDE notes*). Given the pressure at 15 (*given by coursework*) and the pressure, temperature and gamma of 14 (*given by composition of 14 from GASEQ, and assuming const. gamma*), the formula was rearranged and used:

It is worth noting that this is likely the least accurate assumption of the whole analysis due to the of the combustor products varying largely with temp and pressure differences (drops in order of ~15%) and as such cannot really be considered a perfect gas. However, as reheat relationships for gas turbines could not be found in the literature, the formula was used.

### Mass flowrates

# Fuel Cell Performance

## Volts per Cell

## AC power

## Fuel cell efficiency

## Active area

## Heat losses

# System Parameters

## Power

The specific enthalpy drop was determined by comparing the specific enthalpies across the turbine given by the temperatures, pressure and composition from GASEQ.

## Feed-water

## Efficiency

## Emissions

### CO2

### NOx

# Appendix

# Bibliography

**There are no sources in the current document.**