

Responsive Functional Materials

Assist.-Prof. Dr. Heidi A. Schwartz

Photoactive Hybrid Materials

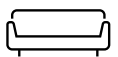
Universität Innsbruck



<https://www.uibk.ac.at/en/aatc/ag-schwartz/>



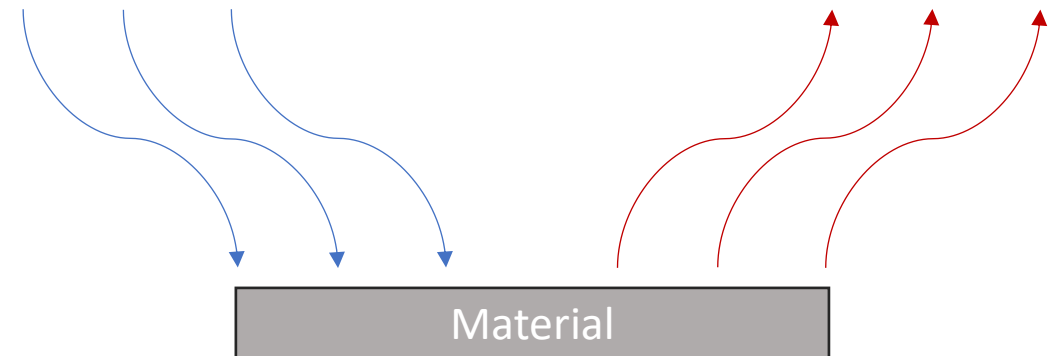
heidi.schwartz@uibk.ac.at



L01.063

External stimuli

Functionality



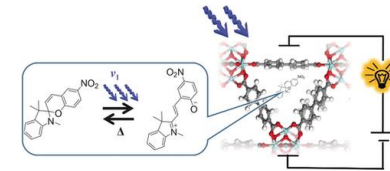
Outline for today's lecture

1. Organizational Points



2. Learning objectives

- Fundamentals and functional materials in everyday life I
- Next time: fundamentals and functional materials in everyday life II

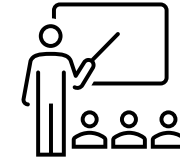


Organizational Points

1. Lecture takes place every Tuesday from 8-10 a.m. (L01.220) and Wednesday (only in March, little complicated) in L03.121 until 29th of April
 2. Powerpoint presentation is available on OLAT after every single lecture
 3. Examination on during easter holiday and 29th April?
-



What you will learn the next weeks

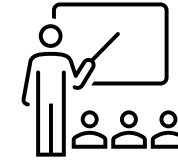


1. Introduction into Material Sciences I
2. Introduction into Material Sciences II
3. Analyzation Tools I
4. Analyzation Tools II
5. Porous Materials I
- 6. Invited Speaker: Artem Mikhailov (Nancy)**
7. Porous Materials II
8. Responsive Materials
9. Photochromism and Luminescence
10. Hybrid Materials





What you will learn the next weeks



1. **Introduction into Material Sciences I**
2. Introduction into Material Sciences II
3. Analyzation Tools I
4. Analyzation Tools II
5. Porous Materials I
6. **Invited Speaker: Artem Mikhailov (Nancy)**
7. Porous Materials II
8. Responsive Materials
9. Photochromism and Luminescence
10. Hybrid Materials



Historical Aspects

Materials are intensively used in our culture:

- Transport
- Housing
- Clothes
- Communication
- Food production, ...



Historical Aspects

Materials are intensively used in our culture:

- Transport
- Housing
- Clothes
- Communication
- Food production, ...



The development and change of societies are always closely linked to the ability to produce and change materials in order to satisfy their needs.



Civilizations were named after predominant material:

Stone age · Bronze age · Iron age



Historical Aspects

Materials are intensively used in our culture:

- Transport
- Housing
- Clothes
- Communication
- Food production, ...



The development and change of societies are always closely linked to the ability to produce and change materials in order to satisfy their needs.



Civilizations were named after predominant material:

Stone age · Bronze age · Iron age



Nowadays - Realization

Structure in strong relation to properties!

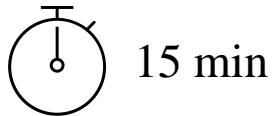
Adjustment of structure – development of numerous materials!

Group-/Tandem-Work

Which materials do you know from everyday life?

What are their properties?

What are they used for?

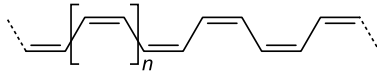


A functional material could be defined as being prepared from a “target-motivated” approach

Ceramics



Polymers



Metals



Composites



Glasses

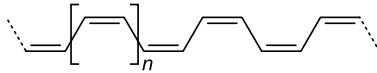


A functional material could be defined as being prepared from a “target-motivated” approach

Ceramics



Polymers



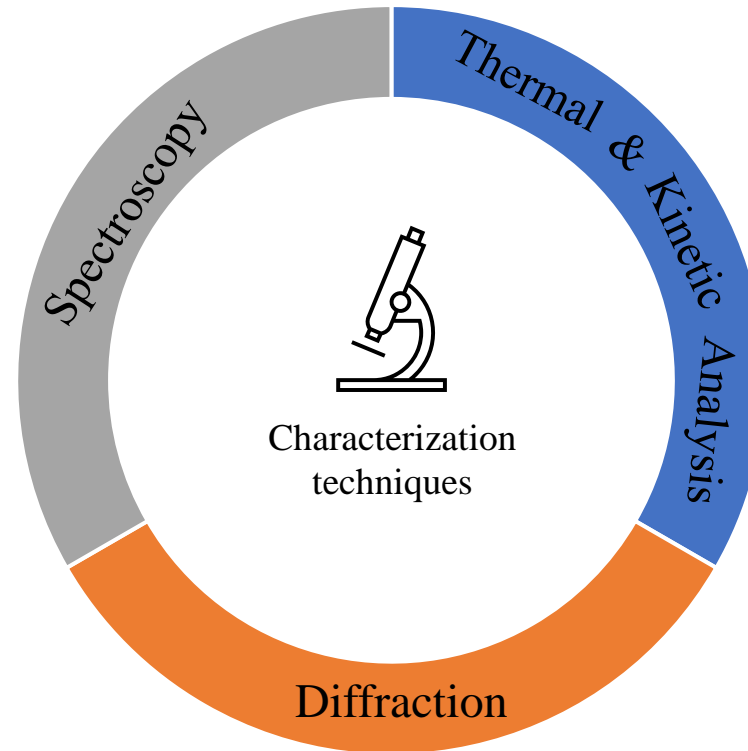
Metals



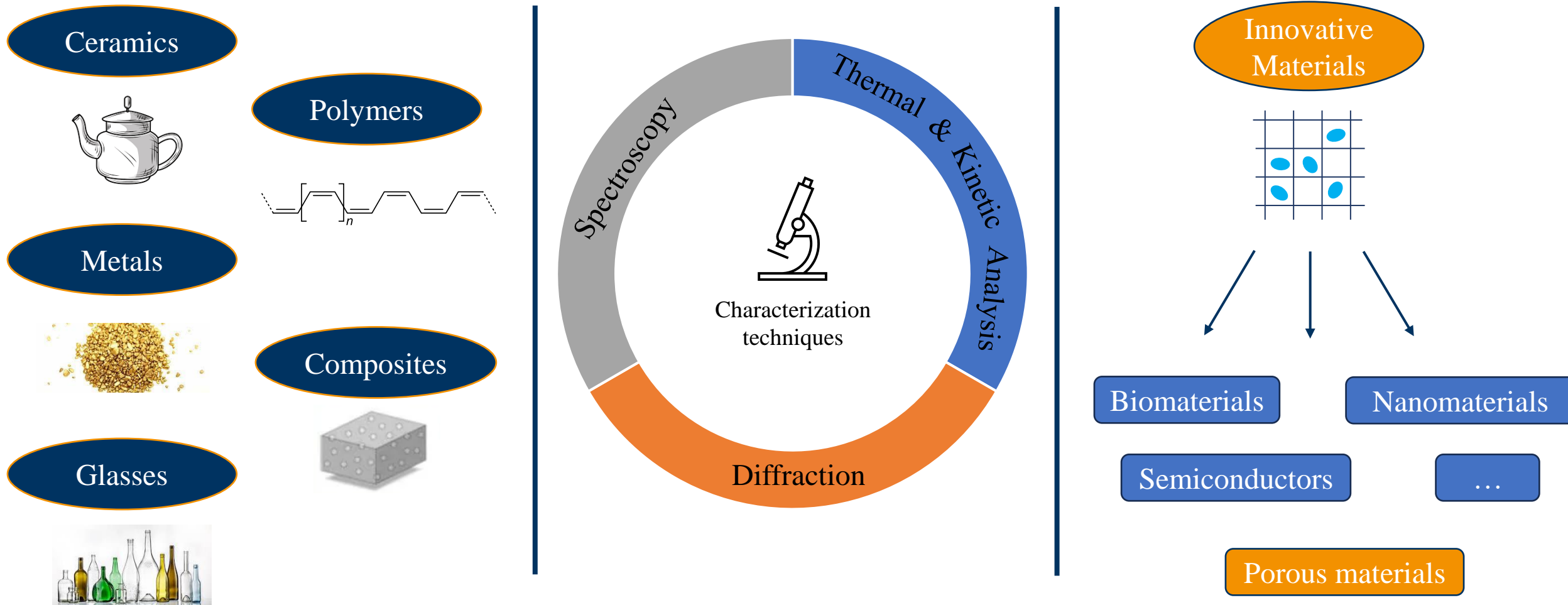
Composites



Glasses



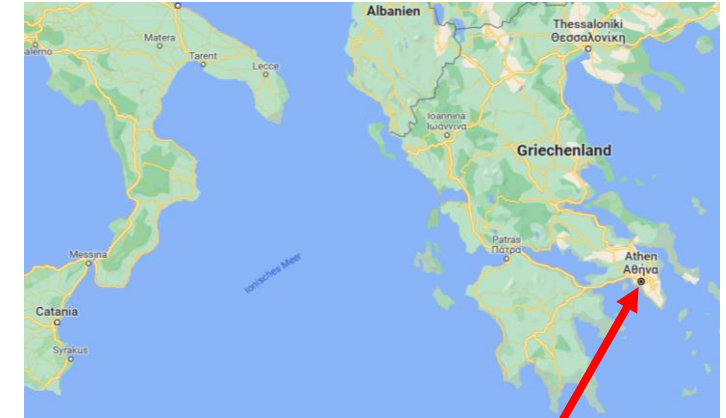
A functional material could be defined as being prepared from a “target-motivated” approach



Ceramics – one of the oldest functional material



Venus from Dolní Věstonice
30.000 years old.



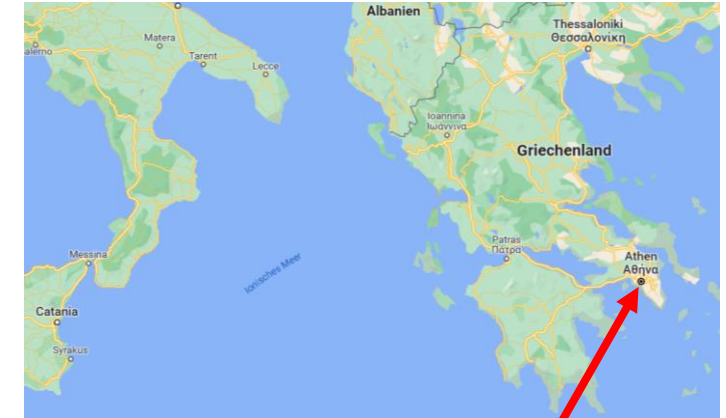
Origin in Kerameikos, a district in
Athens.

Ceramics – one of the oldest functional material

“Ceramic” is the technical term for a variety of inorganic non-metallic materials. They are almost non-soluble in water, and at least of 30% crystallinity.



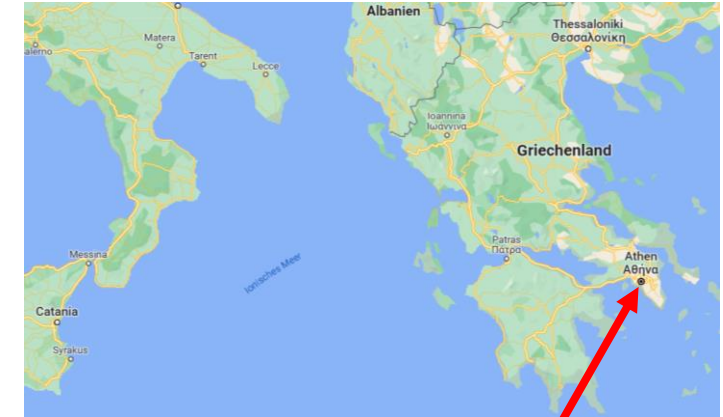
Venus from Dolní Věstonice
30.000 years old.



Origin in Kerameikos, a district in
Athens.

Ceramics – one of the oldest functional material

“Ceramic” is the technical term for a variety of inorganic non-metallic materials. They are almost non-soluble in water, and at least of 30% crystallinity.



Origin in Kerameikos, a district in Athens.



Venus from Dolní Věstonice
30.000 years old.



Synthesis procedure

Raw materials formed at RT



Green body.

e.g.,
Silicates
Oxides
SiC
BN
B₄C

WC
Si₃N₄
AlN
MoSi₂

Sintering, mostly
> 800 °C



Bonding situation and properties

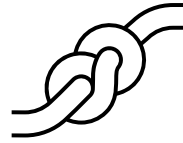
Large range of possible structures/compositions

- Almost all elements
- Almost all type of bonds (mainly covalent & ionic)
- All levels of crystallinity

Classification

- Oxides: alumina, zirconia, ceria, ...
- Non-oxides: carbides, silicides, borides, nitrides
- Composites

Bonding situation



and properties

Large range of possible structures/compositions

- Almost all elements
- Almost all type of bonds (mainly covalent & ionic)
- All levels of crystallinity

Classification

- Oxides: alumina, zirconia, ceria, ...
- Non-oxides: carbides, silicides, borides, nitrides
- Composites

Mechanical properties

- Poor toughness due to bond type
- Pores as stress concentrators

Ceramic matrix composites

nature portfolio

• Nature. 2024 Feb 21;626(8000):779–784. doi: [10.1038/s41586-024-07036-5](https://doi.org/10.1038/s41586-024-07036-5)

Twisted-layer boron nitride ceramic with high deformability and strength

Yingju Wu ^{1,2,✉}, Yang Zhang ^{1,3,✉}, Xiaoyu Wang ^{1,✉}, Wentao Hu ^{1,✉}, Song Zhao ^{1,✉}, Timothy Officer ⁴, Kun Luo ¹, Ke Tong ¹, Congcong Du ⁵, Liqiang Zhang ⁵, Baozhong Li ¹, Zewen Zhuge ¹, Zitai Liang ¹, Mengdong Ma ¹, Anmin Nie ¹, Dongli Yu ¹, Julong He ¹, Zhongyuan Liu ¹, Bo Xu ¹, Yanbin Wang ⁴, Zhisheng Zhao ^{1,✉}, Yongjun Tian ^{1,✉}

• Author information • Article notes • Copyright and License information

[Ceramic – Wikipedia](#) (18th February, 2025)

[The first bulk ceramic that deforms like a metal at room temperature](#) (18th February, 2025)

[Twisted-layer boron nitride ceramic with high deformability and strength | Nature](#) (18th February, 2025)

Bonding situation and properties

Large range of possible structures/compositions

- Almost all elements
- Almost all type of bonds (mainly covalent & ionic)
- All levels of crystallinity

Classification

- Oxides: alumina, zirconia, ceria, ...
- Non-oxides: carbides, silicides, borides, nitrides
- Composites

Mechanical properties

- Poor toughness due to bond type
- Pores as stress concentrators

Ceramic matrix composites

Electrical properties

- Semiconductors
- Superconductors
- Ferroelectricity & superlattices
- Positive thermal coefficient

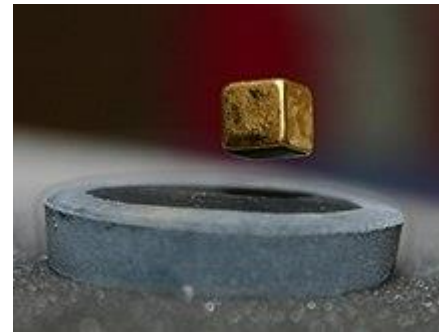
nature portfolio

• Nature. 2024 Feb 21;626(8000):779–784. doi: [10.1038/s41586-024-07036-5](https://doi.org/10.1038/s41586-024-07036-5)

Twisted-layer boron nitride ceramic with high deformability and strength

Yingju Wu ^{1,2,✉}, Yang Zhang ^{1,3,✉}, Xiaoyu Wang ^{1,✉}, Wentao Hu ^{1,✉}, Song Zhao ^{1,✉}, Timothy Officer ⁴, Kun Luo ¹, Ke Tong ¹, Congcong Du ⁵, Liqiang Zhang ⁵, Baozhong Li ¹, Zewen Zhuge ¹, Zitai Liang ¹, Mengdong Ma ¹, Anmin Nie ¹, Dongli Yu ¹, Julong He ¹, Zhongyuan Liu ¹, Bo Xu ¹, Yanbin Wang ⁴, Zhisheng Zhao ^{1,✉}, Yongjun Tian ^{1,✉}

• Author information • Article notes • Copyright and License information



[Ceramic – Wikipedia](#) (18th February, 2025)

[The first bulk ceramic that deforms like a metal at room temperature](#) (18th February, 2025)

[Twisted-layer boron nitride ceramic with high deformability and strength | Nature](#) (18th February, 2025)

Bonding situation and properties

Large range of possible structures/compositions

- Almost all elements
- Almost all type of bonds (mainly covalent & ionic)
- All levels of crystallinity

Classification

- Oxides: alumina, zirconia, ceria, ...
- Non-oxides: carbides, silicides, borides, nitrides
- Composites

Mechanical properties

- Poor toughness due to bond type
- Pores as stress concentrators

Ceramic matrix composites

Electrical properties

- Semiconductors
- Superconductors
- Ferroelectricity & superlattices
- Positive thermal coefficient

Optical properties

- Transmission of light in vis and IR range
- GE manufactured translucent alumina

nature portfolio

• Nature. 2024 Feb 21;626(8000):779–784. doi: [10.1038/s41586-024-07036-5](https://doi.org/10.1038/s41586-024-07036-5)

Twisted-layer boron nitride ceramic with high deformability and strength

Yingju Wu ^{1,2,✉}, Yang Zhang ^{1,3,✉}, Xiaoyu Wang ^{1,✉}, Wentao Hu ^{1,✉}, Song Zhao ^{1,✉}, Timothy Officer ⁴, Kun Luo ¹, Ke Tong ¹, Congcong Du ⁵, Liqiang Zhang ⁵, Baozhong Li ¹, Zewen Zhuge ¹, Zitai Liang ¹, Mengdong Ma ¹, Anmin Nie ¹, Dongli Yu ¹, Julong He ¹, Zhongyuan Liu ¹, Bo Xu ¹, Yanbin Wang ⁴, Zhisheng Zhao ^{1,✉}, Yongjun Tian ^{1,✉}

• Author information • Article notes • Copyright and License information



[Ceramic – Wikipedia](#) (18th February, 2025) [Transparent ceramics – Wikipedia](#) (18th February, 2025)

[The first bulk ceramic that deforms like a metal at room temperature](#) (18th February, 2025)

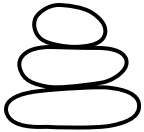
[Twisted-layer boron nitride ceramic with high deformability and strength | Nature](#) (18th February, 2025)

White Board

Technical ceramics

For the resulting properties of ceramic materials, not only the raw materials, but even more the synthesis processes are of outmost importance.

Grain size



+



Firing process

Firing atmospheres

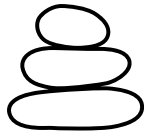
Firing temperatures

Starting materials

Technical ceramics

For the resulting properties of ceramic materials, not only the raw materials, but even more the synthesis processes are of outmost importance.

Grain size



+

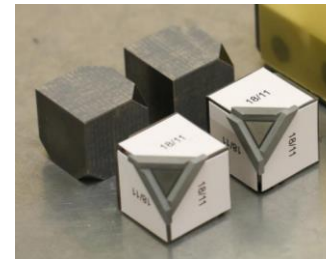


Starting materials

Firing process

Firing atmospheres

Firing temperatures



WC.



ZrO₂.



BN.



Oxidic ceramics.



Ca₅(PO₄)₃(OH).



Siliconized carbon fibre.

Ceramics have several advantages over metals for practical applications, including low density, high compressive strength (hardness) and resistance to corrosion. However, they have a fatal weakness when used as engineering materials — they fracture after undergoing a very small deformation.

Metals – old but gold



First gold found in
Spain, 40.000 B.C..

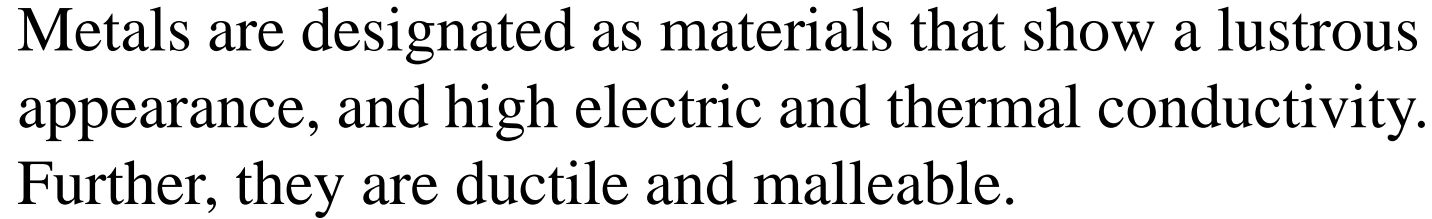
[Metalle – Wikipedia](#) (4th March, 2023)

[CARRS Sterling Silver Cutlery from Lincoln House](#) (4th March, 2023)

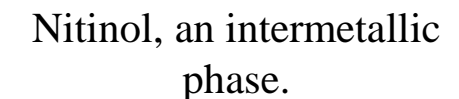
[History of Gold | Gold Eagle \(gold-eagle.com\)](#) (4th March, 2023)

[Nitinol – Wikipedia](#) (4th March, 2023)

[Iron | 7439-89-6 \(chemicalbook.com\)](#) (4th March, 2023)



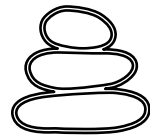
Metallic elements of the periodic table.



Metals are designated as materials that show a lustrous appearance, and high electric and thermal conductivity. Further, they are ductile and malleable.



First gold found in Spain, 40.000 B.C..




Metallurgy

White Board

Ore

Process of metal mining and purification

1. Mining
2. Further processing
- Roasting
 - Extraction with water
 - Electrolytic/-thermic and carbothermic processes
- 



Pure iron.



Sterling silver, an alloy.

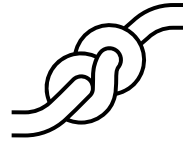


Nitinol, an intermetallic phase.

H																	He	
Li	Be											B	C	N	O	F	Ne	
Na	Mg											Al	Si	P	S	Cl	Ar	
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr	
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe	
Cs	Ba	La	*	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn
Fr	Ra	Ac	**	Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Cn	Nh	Fl	Mc	Lv	Ts	Og
			*	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu	
			**	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr	

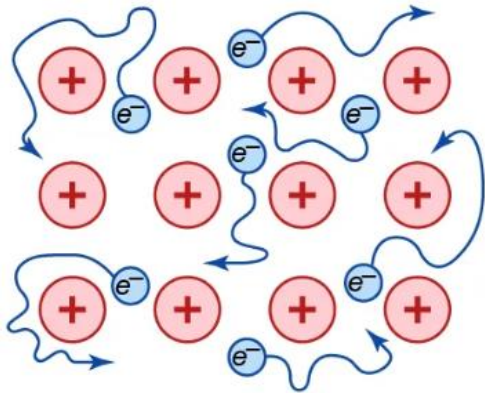
Metallic elements of the periodic table.

Bonding situation

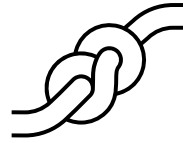


Electron-sea model

Positively charged
atomic bodies
surrounded by freely
mobile negatively
charged valence
electrons.

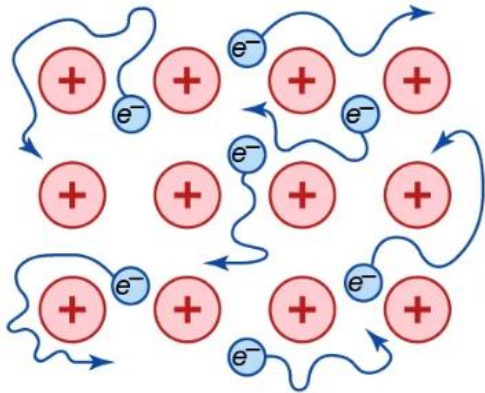


Bonding situation



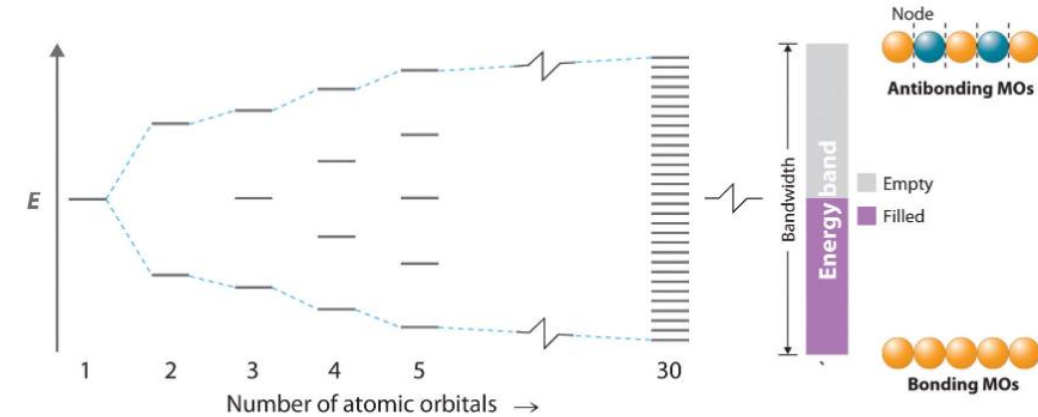
Electron-sea model

Positively charged atomic bodies surrounded by freely mobile negatively charged valence electrons.

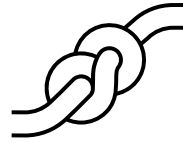


Band theory

Set of molecular orbitals is generated that extends throughout the solid: formation of bands with delocalized electrons.

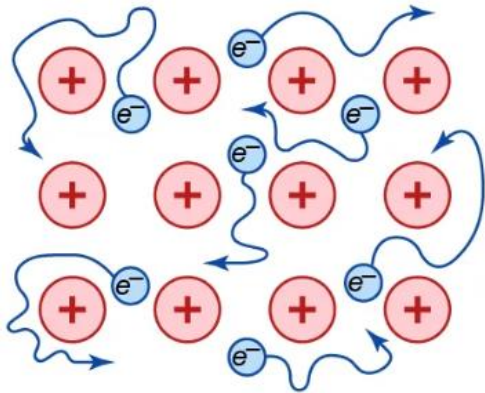


Bonding situation



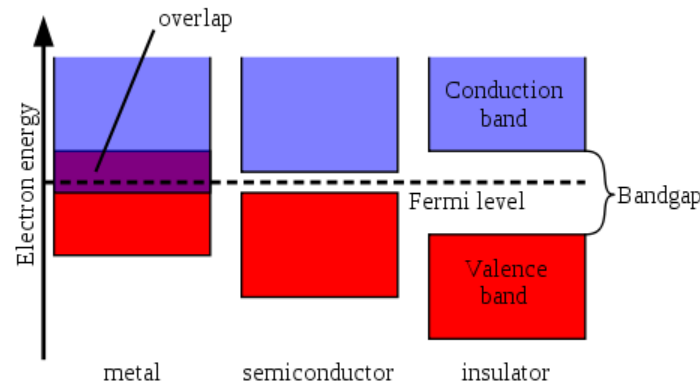
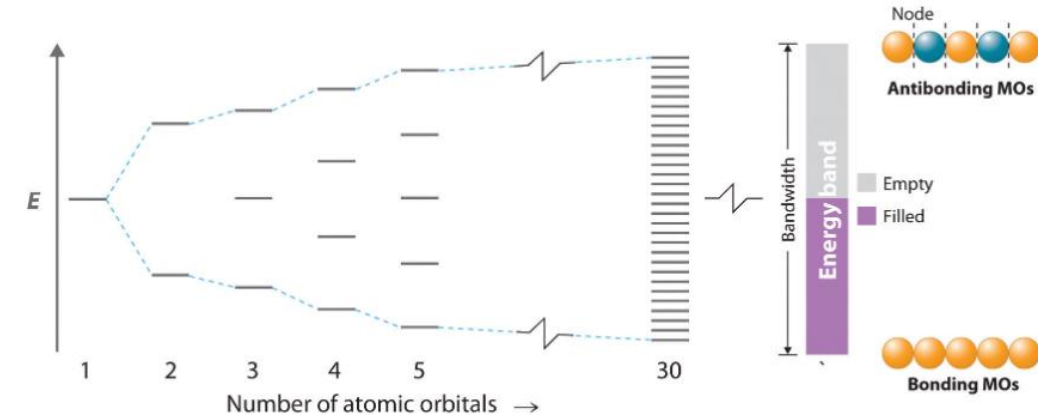
Electron-sea model

Positively charged atomic bodies surrounded by freely mobile negatively charged valence electrons.

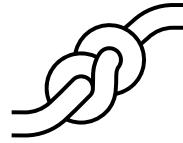


Band theory

Set of molecular orbitals is generated that extends throughout the solid: formation of bands with delocalized electrons.

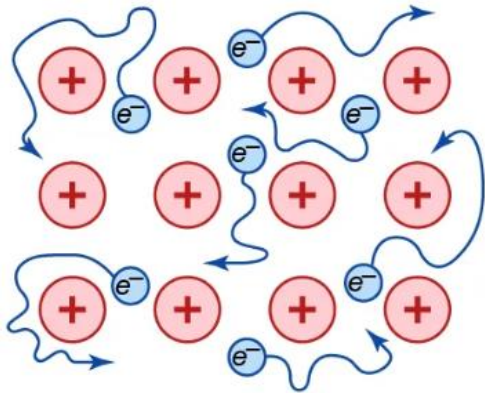


Bonding situation



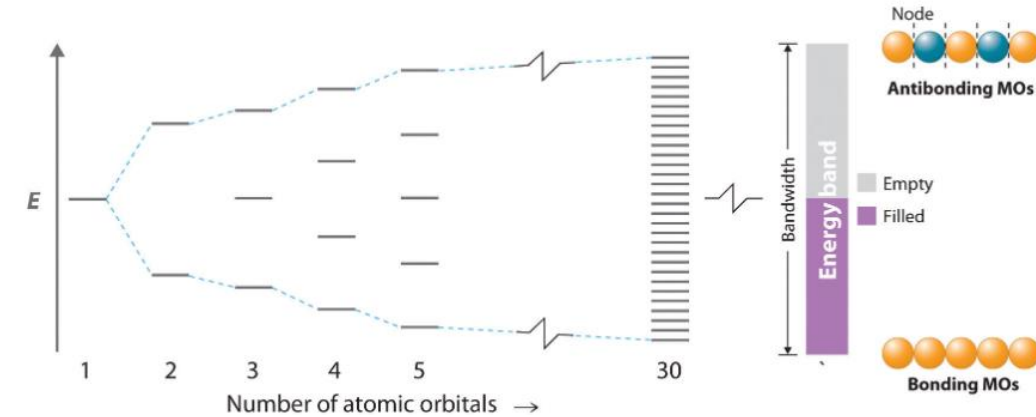
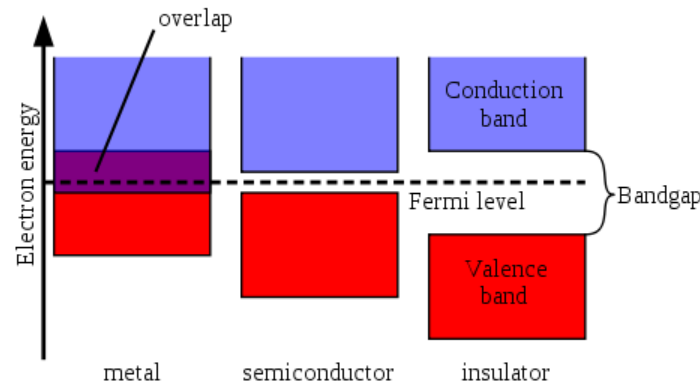
Electron-sea model

Positively charged atomic bodies surrounded by freely mobile negatively charged valence electrons.



Band theory

Set of molecular orbitals is generated that extends throughout the solid: formation of bands with delocalized electrons.



Resulting properties

- Lustrous
- Malleable
- Ductile
- Good conductors of electricity and heat

Technical applications

Pure metals (*e.g.*, copper) used for wires because of their high electric conductivity.
In further applications, metals are applied as alloys or intermetallic phases.

Steel, an iron alloy.



Indium and gallium.



Tungsten wire.



Aluminum.



Alloys – better together

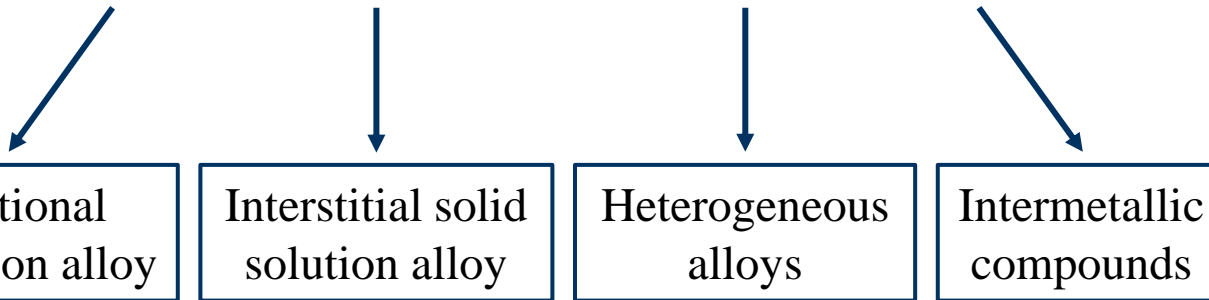


Combination of chemical elements, possessing the characteristic properties of a metal. The production of alloys is of great importance, as it is one of the main processes used to modify the properties of pure metallic elements.

Alloys – better together



Combination of chemical elements, possessing the characteristic properties of a metal. The production of alloys is of great importance, as it is one of the main processes used to modify the properties of pure metallic elements.

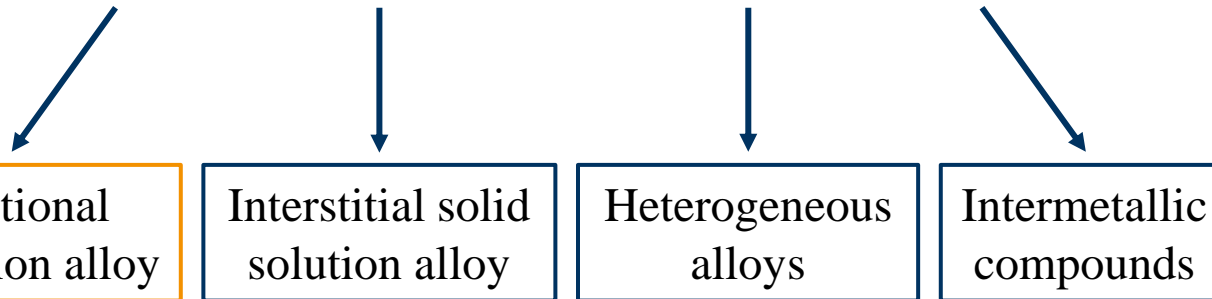


Homogeneous mixtures

Alloys – better together



Combination of chemical elements, possessing the characteristic properties of a metal. The production of alloys is of great importance, as it is one of the main processes used to modify the properties of pure metallic elements.

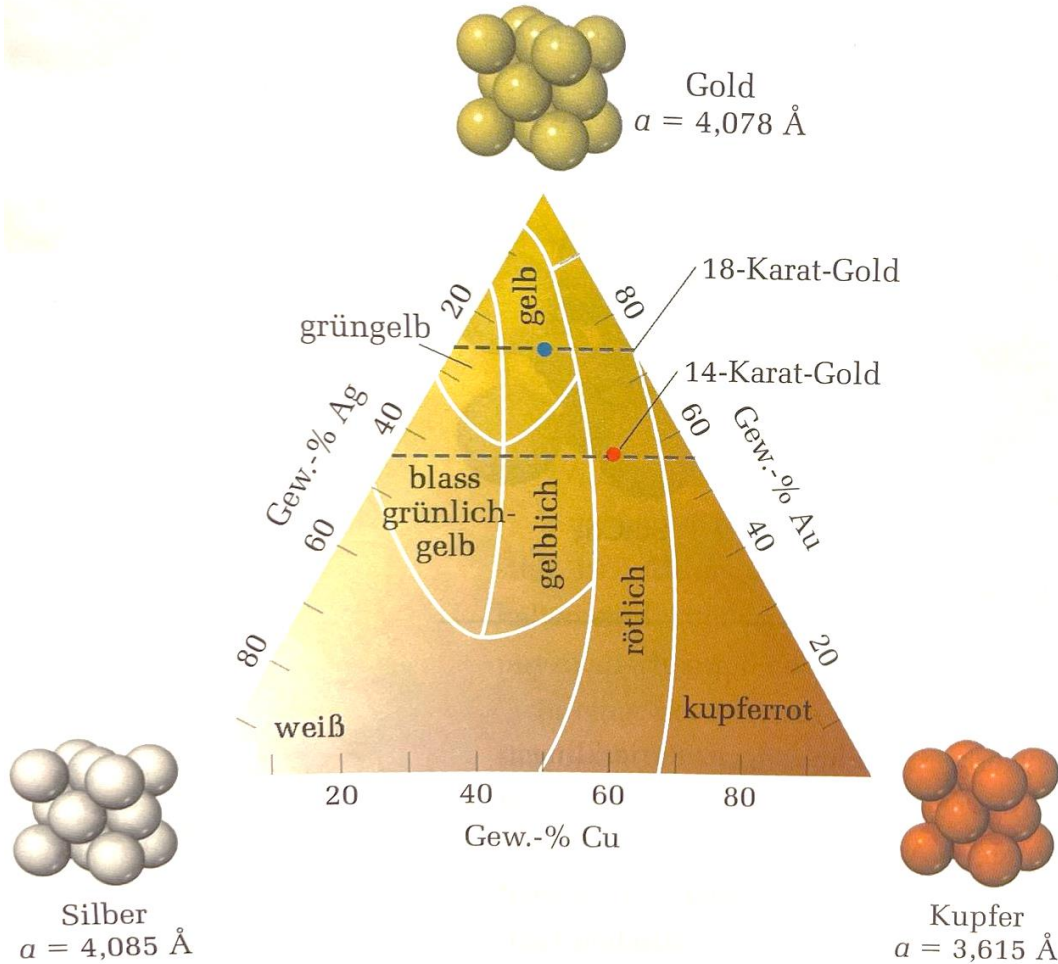


Homogeneous mixtures, **components**
are regularly and randomly distributed
→ Solid solutions



If r_A (metal A) and r_B (metal B) are within 15% deviation

Alloys – better together



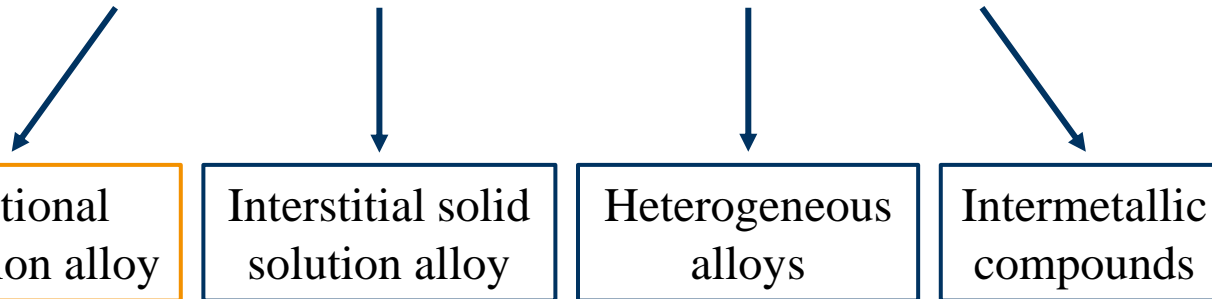
Pure gold too soft – better as alloy!

- Pure gold: 24 Karat
- Ag, Cu and Au: fcc structure and almost same radii
- Same amount Ag and Cu: 18 Karat gold
- Copper-rich alloy: 14 Karat (roségold)
- Silver-rich alloy: white-gold (Karat depending on amount of Ag)

Alloys – better together



Combination of chemical elements, possessing the characteristic properties of a metal. The production of alloys is of great importance, as it is one of the main processes used to modify the properties of pure metallic elements.



Homogeneous mixtures, **components**
are regularly and randomly distributed
→ Solid solutions



If r_A (metal A) and r_B (metal B) are within 15% deviation

Alloys – better together



Combination of chemical elements, possessing the characteristic properties of a metal. The production of alloys is of great importance, as it is one of the main processes used to modify the properties of pure metallic elements.

Substitutional
solid solution alloy

Interstitial solid
solution alloy

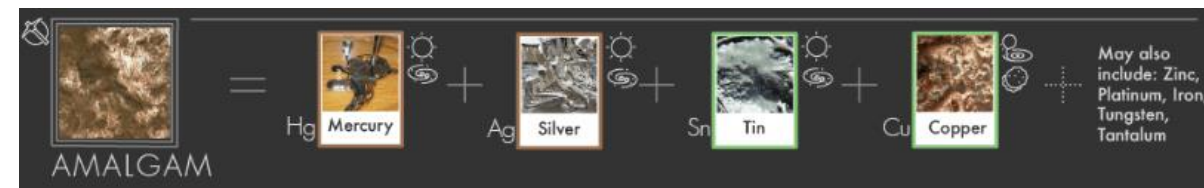
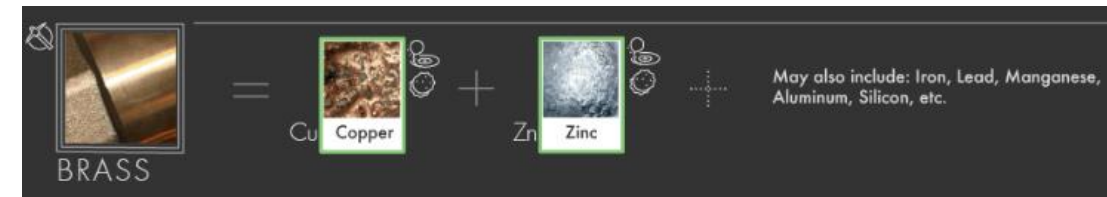
Heterogeneous
alloys

Intermetallic
compounds

Homogeneous mixtures, **components**
are regularly and randomly distributed
→ Solid solutions



If r_A (metal A) and r_B (metal B) are within 15% deviation



[Alloy – Wikipedia](#) (17th February, 2025)

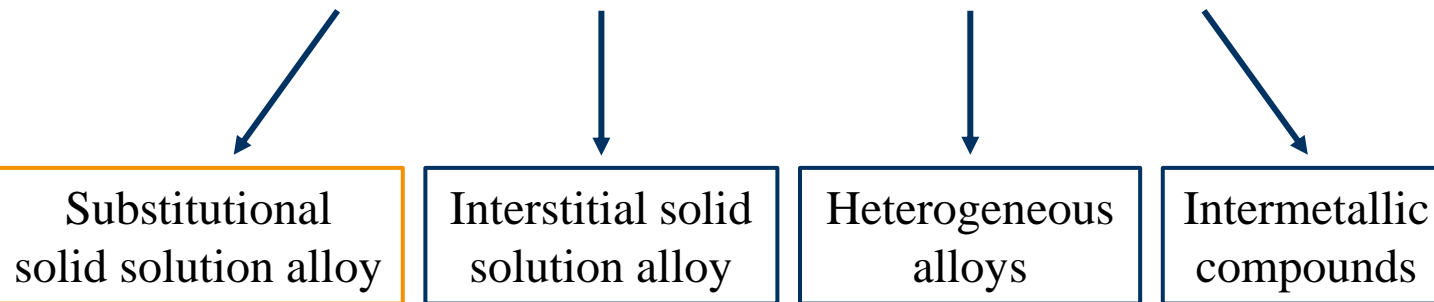
Brown, LeMay et al., Chemie – Studieren kompakt, 14th edition, Pearson Studium, 2018.

[Infographic: 20 Common Metal Alloys and What They're Made Of](#) (17th February, 2025)

Alloys – better together



Combination of chemical elements, possessing the characteristic properties of a metal. The production of alloys is of great importance, as it is one of the main processes used to modify the properties of pure metallic elements.

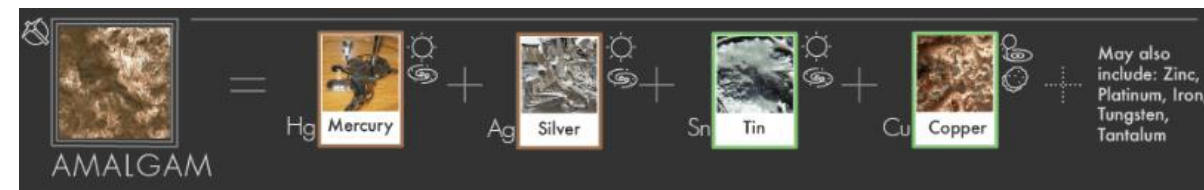
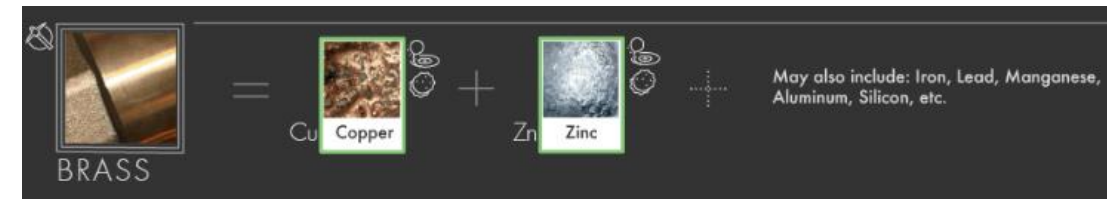


Homogeneous mixtures, **components** are regularly and randomly distributed
→ Solid solutions

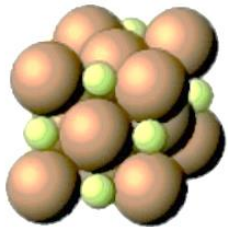
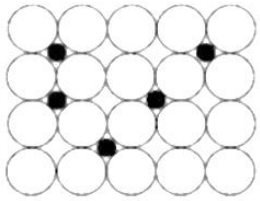


If r_A (metal A) and r_B (metal b) are within 15% deviation

Steel

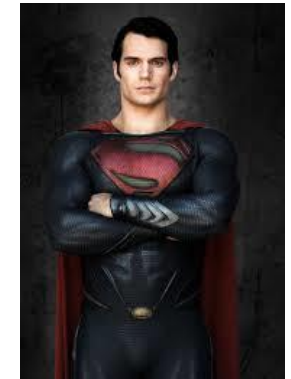
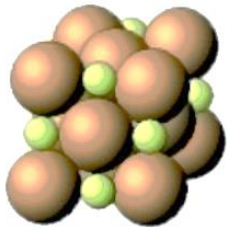
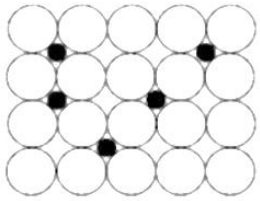


Types of Steel – What doping does!



Types of Steel – What doping does!

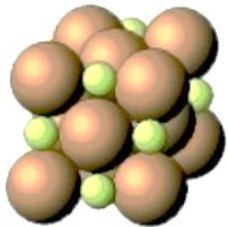
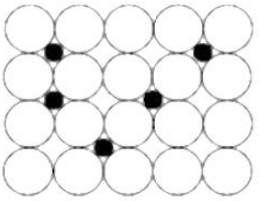
Alloy of iron and max. 2.14% carbon with additional elemental doping resulting in defined properties. Metal lattice becomes harder, stronger and less ductile.



Types of Steel – What doping does!

Alloy of iron and max. 2.14% carbon with additional elemental doping resulting in defined properties. Metal lattice becomes harder, stronger and less ductile.

Most important iron alloy is stainless steel with 0.4% carbon, 18% chromium and 1% nickel.

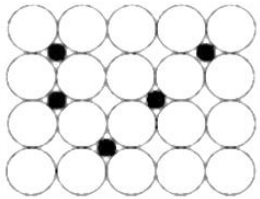


Types of Steel – What doping does!

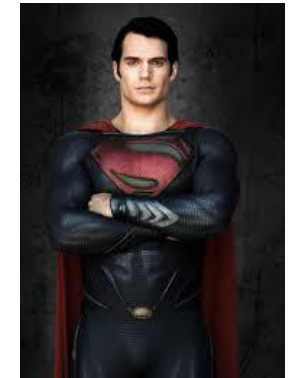
Alloy of iron and max. 2.14% carbon with additional elemental doping resulting in defined properties. Metal lattice becomes harder, stronger and less ductile.



Most important iron alloy is stainless steel with 0.4% carbon, 18% chromium and 1% nickel.



aluminum	Increases desoxygenation and ferrite stabilization
boron	Increases yield strength, strength, brittleness and heat resistance
chromium	Increases cooling rate, wear resistance, heat resistance, tensile strength, hardness
cobalt	Increases strength, heat resistance
molybdenum	Increases hardenability, tensile strength, weldability, ductility, heat resistance
nickel	Increases tensile strength, hardness, strength, yield strength, ferrite stabilization

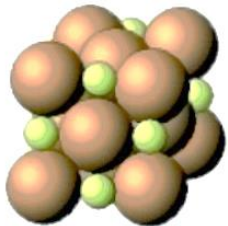
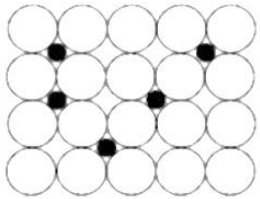


Types of Steel – What doping does!

Alloy of iron and max. 2.14% carbon with additional elemental doping resulting in defined properties. Metal lattice becomes harder, stronger and less ductile.



Most important iron alloy is stainless steel with 0.4% carbon, 18% chromium and 1% nickel.



aluminum	Increases desoxygenation and ferrite stabilization
boron	Increases yield strength, strength, brittleness and heat resistance
chromium	Increases cooling rate, wear resistance, heat resistance, tensile strength, hardness
cobalt	Increases strength, heat resistance
molybdenum	Increases hardenability, tensile strength, weldability, ductility, heat resistance
nickel	Increases tensile strength, hardness, strength, yield strength, ferrite stabilization



Intermetallic compounds

Intermetallic compounds are solid phases with ≥ 2 (semi-) metallic elements possessing an ordered structure with a well-defined stoichiometry.

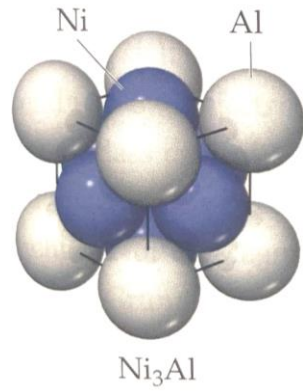


Structure results in higher strength and higher melting points compared to single components
→ very interesting for high temperature applications!

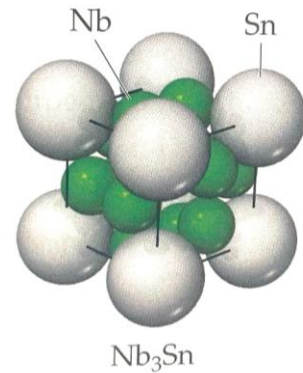
Intermetallic compounds

Intermetallic compounds are solid phases with ≥ 2 (semi-) metallic elements possessing an ordered structure with a well-defined stoichiometry.

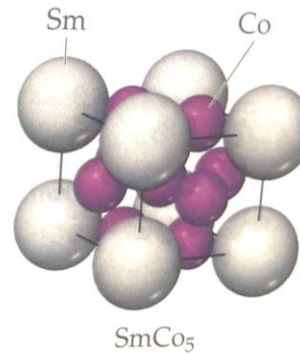
Structure results in higher strength and higher melting points compared to single components
→ very interesting for high temperature applications!



Strength at high temperatures and low density
Jet engines



Superconductor at < 18 K
MRI and NMR

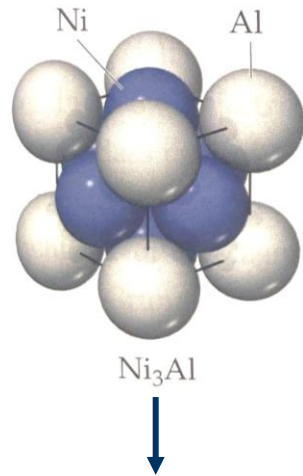


Permanent magnet in light
Headsets or loudspeakers

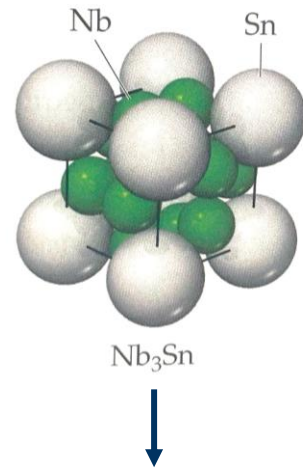
Intermetallic compounds

Intermetallic compounds are solid phases with ≥ 2 (semi-) metallic elements possessing an ordered structure with a well-defined stoichiometry.

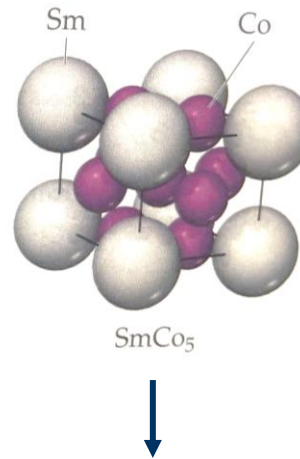
Structure results in higher strength and higher melting points compared to single components
→ very interesting for high temperature applications!



Strength at high temperatures and low density
Jet engines



Superconductor at $< 18 \text{ K}$
MRI and NMR



Permanent magnet in light
Headsets or loudspeakers



No thermal expansion!

What's your wish?



Pyrochlore magnets – no thermal expansion



Eiffel tower: 15 cm
higher in summer
than in winter!

Pyrochlore magnets – no thermal expansion



Alloy from the four metals Zr, Nb, Fe and Co, named pyrochlore-magnet, expands $< 0.0001\%$ in a $400\text{ }^{\circ}\text{C}$ temperature range! Record!



Eiffel tower: 15 cm
higher in summer
than in winter!

Pyrochlore magnets – no thermal expansion



Alloy from the four metals Zr, Nb, Fe and Co, named pyrochlore-magnet, expands $< 0.0001\%$ in a $400\text{ }^{\circ}\text{C}$ temperature range! Record!



Eiffel tower: 15 cm
higher in summer
than in winter!

Note

Pyrochlor: $\text{Ca}_2\text{Nb}_2\text{O}_7$ with ideal cubic structure, natural mineral



Pyrochlore magnets – no thermal expansion



Alloy from the four metals Zr, Nb, Fe and Co, named pyrochlore-magnet, expands $< 0.0001\%$ in a $400\text{ }^{\circ}\text{C}$ temperature range! Record!

History

1896 alloy *Invar* (65% Fe, 35% Ni) shows small TE

Reason: additional magnetic effect

Thermal expansion vs. decreasing magnetic repulsion



Eiffel tower: 15 cm
higher in summer
than in winter!

Note

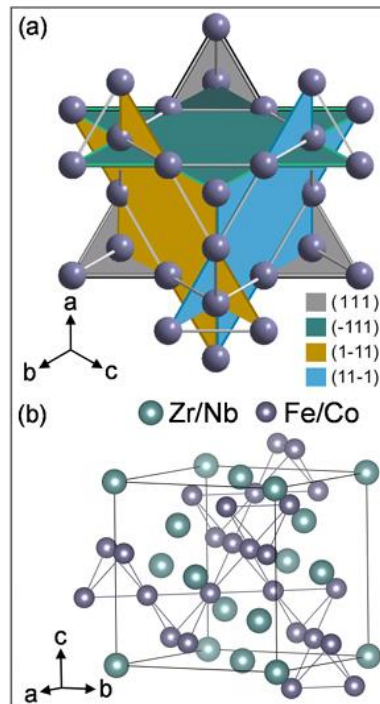
Pyrochlor: $\text{Ca}_2\text{Nb}_2\text{O}_7$ with ideal cubic structure, natural mineral



Pyrochlore magnets – no thermal expansion



Alloy from the four metals Zr, Nb, Fe and Co, named pyrochlore-magnet, expands $< 0.0001\%$ in a $400\text{ }^{\circ}\text{C}$ temperature range! Record!



History

1896 alloy *Invar* (65% Fe, 35% Ni) shows small TE

Reason: additional magnetic effect

Thermal expansion vs. decreasing magnetic repulsion

Reason for almost ZTE

- Heterogeneous lattice structure, not ideally cubic
- Parts within structure react thermally and magnetically differently

Thermal treatment: partial contraction and repulsion with net-balance zero

Note

Pyrochlor: $\text{Ca}_2\text{Nb}_2\text{O}_7$ with ideal cubic structure, natural mineral

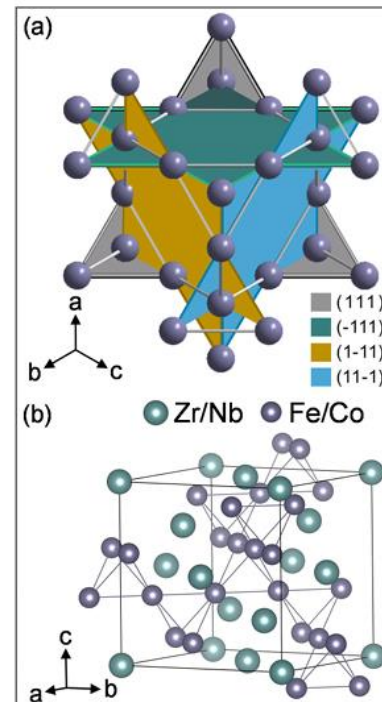
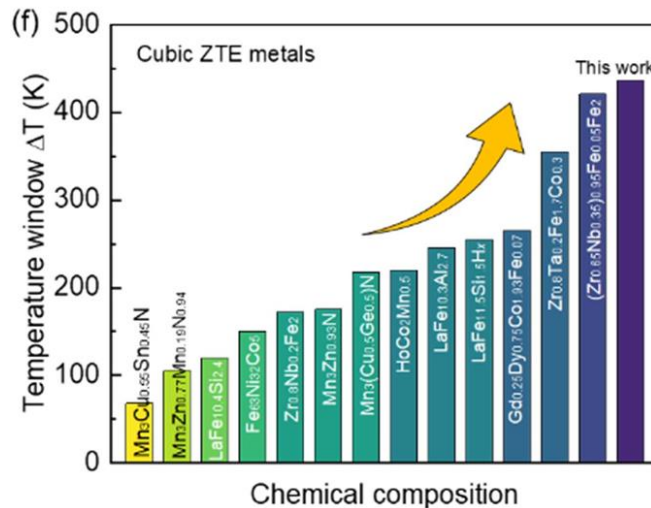


Eiffel tower: 15 cm higher in summer than in winter!

Pyrochlore magnets – no thermal expansion



Alloy from the four metals Zr, Nb, Fe and Co, named pyrochlore-magnet, expands $< 0.0001\%$ in a $400\text{ }^{\circ}\text{C}$ temperature range! Record!



History

1896 alloy *Invar* (65% Fe, 35% Ni) shows small TE

Reason: additional magnetic effect

Thermal expansion vs. decreasing magnetic repulsion

Reason for almost ZTE

- Heterogeneous lattice structure, not ideally cubic
- Parts within structure react thermally and magnetically differently

Thermal treatment: partial contraction and repulsion with net-balance zero

Note

Pyrochlore: $\text{Ca}_2\text{Nb}_2\text{O}_7$ with ideal cubic structure, natural mineral



Eiffel tower: 15 cm higher in summer than in winter!

Excursion: Amorphous metals - Metallic glasses

Usually alloy containing atoms with significant different sizes ($>12\%$) – low free volume!
Higher viscosity in molten state, which prevents atoms to form an ordered lattice.
In contrast to “normal” metals, they are non-crystalline and have a glass-like structure.
They are electrically conductive and lustrous!



Excursion: Amorphous metals - Metallic glasses

Usually alloy containing atoms with significant different sizes ($>12\%$) – low free volume!
Higher viscosity in molten state, which prevents atoms to form an ordered lattice.
In contrast to “normal” metals, they are non-crystalline and have a glass-like structure.
They are electrically conductive and lustrous!

Formation

Extremely rapid cooling
Mechanical alloying
Ion irradiation
...



Properties

Low shrinkage when cooling
Wear resistance
Less corrosion due to missing grain boundaries
Great hardness
Best magnetic soft materials, ...



Excursion: Amorphous metals - Metallic glasses

Usually alloy containing atoms with significant different sizes ($>12\%$) – low free volume!
Higher viscosity in molten state, which prevents atoms to form an ordered lattice.
In contrast to “normal” metals, they are non-crystalline and have a glass-like structure.
They are electrically conductive and lustrous!

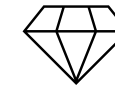


Formation

Extremely rapid cooling
Mechanical alloying
Ion irradiation
...



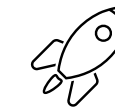
Scalpels, maybe implants



Difficult shapes



Mobile phones, USB sticks



Solar wind collectors (Genesis probe)



Armor-piercing kinetic energy penetrators, aviation

Properties

Low shrinkage when cooling
Wear resistance
Less corrosion due to missing grain boundaries
Great hardness
Best magnetic soft materials, ...

Excursion: Amorphous metals - Metallic glasses

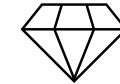
Usually alloy containing atoms with significant different sizes ($>12\%$) – low free volume!
Higher viscosity in molten state, which prevents atoms to form an ordered lattice.
In contrast to “normal” metals, they are non-crystalline and have a glass-like structure.
They are electrically conductive and lustrous!

Formation

Extremely rapid cooling
Mechanical alloying
Ion irradiation
...



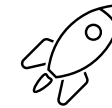
Scalpels, maybe implants



Difficult shapes



Mobile phones, USB sticks



Solar wind collectors (Genesis probe)



Armor-piercing kinetic energy penetrators, aviation

Properties

Low shrinkage when cooling
Wear resistance
Less corrosion due to missing grain boundaries
Great hardness
Best magnetic soft materials, ...



LIQUID METAL GOLF CLUBS

Unusual springiness and a soft feel, how about liquid metal golf clubs?

25/02/2020

One of the first commercial applications of the **Liquid metal** was in the golf field, in order to produce **golf clubs** heads.
Bulk Metallic Glass has several outstanding attributes, such as: high strength-to-weight ratios, high hardness, good forming and shaping qualities, unusual magnetism and, above all, BMG is extremely elastic.

The incomparable mechanical properties of **BMG** can increase the coefficient of restitution at the impact between a golf club and a ball. And, moreover, in addition to all the other characteristics we've already mentioned, liquid metal provides a softer, more solid feel for better control when a golfer strikes the ball.

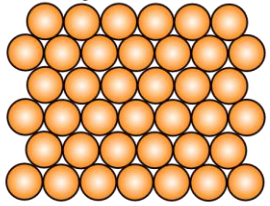
Glasses



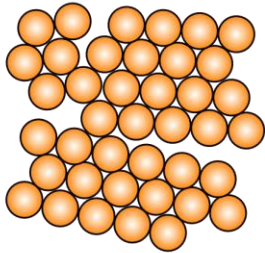
Standard definition glass

- A glass is a non-crystalline solid formed by rapid melt quenching
- It is an amorphous solid and can be understood as a supercooled liquid
- Glasses are arbitrarily shapeable

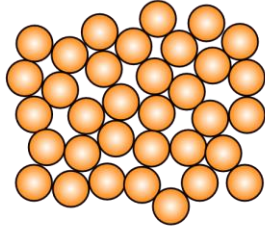
Crystalline



Polycrystalline



Amorphous



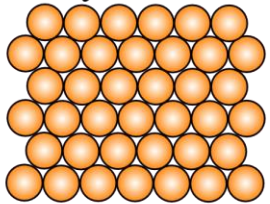
Glasses



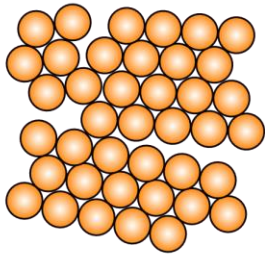
Standard definition glass

- A glass is a non-crystalline solid formed by rapid melt quenching
- It is an amorphous solid and can be understood as a supercooled liquid
- Glasses are arbitrarily shapeable

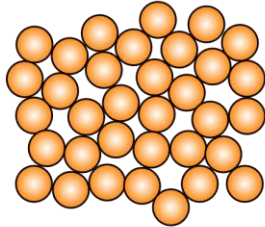
Crystalline



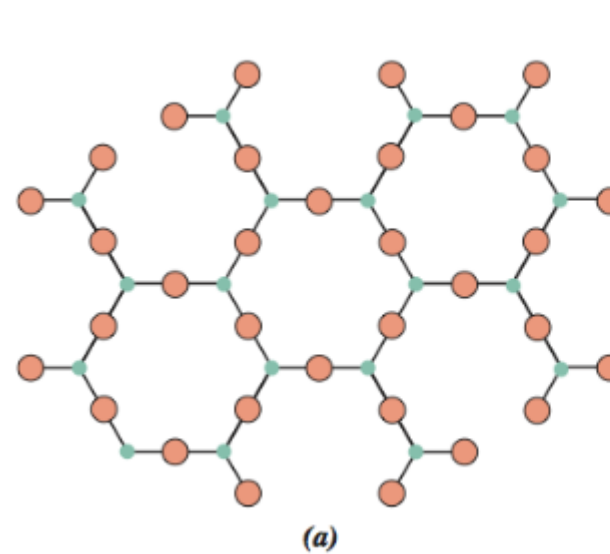
Polycrystalline



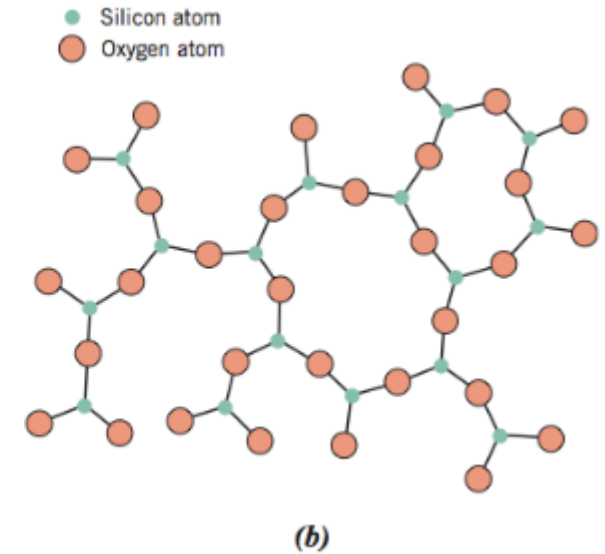
Amorphous



- SiO_2 can also occur as a non-crystalline (amorphous) solid (silica gel or quartz glass)
- Other oxides such as B_2O_3 or GeO_2 can also form glasses
- These materials are called network formers



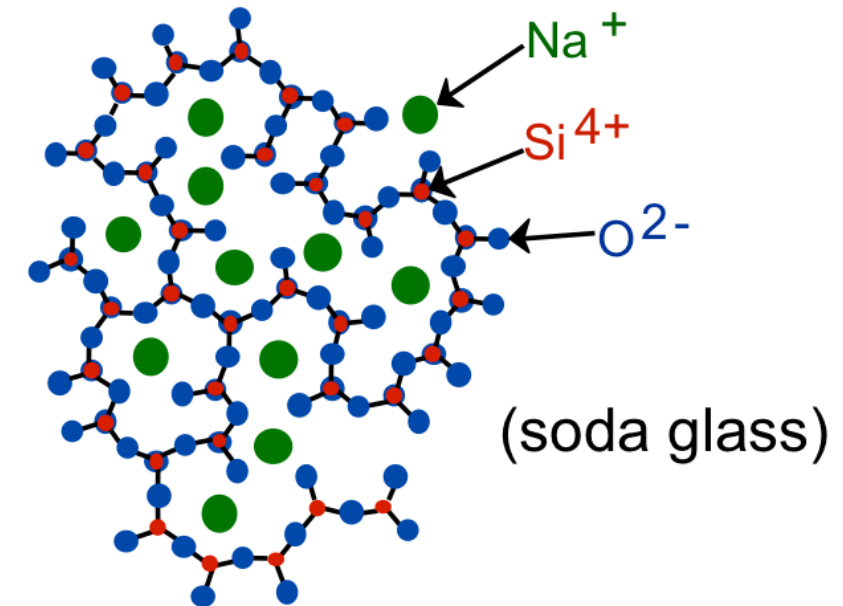
Crystalline SiO_2



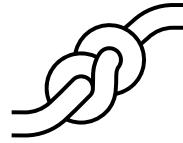
Amorphous SiO_2

Glasses – more than SiO_2

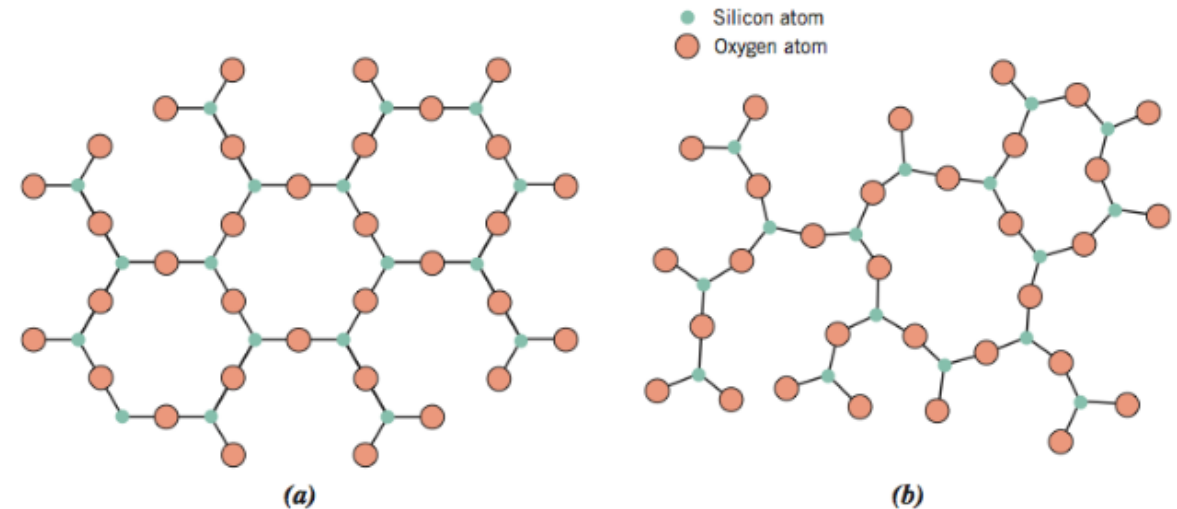
- The most common glasses are silicate glasses, to which additional oxides such as CaO or Na_2O are added
- The cations are incorporated into the tetrahedral network, thereby altering its properties
Therefore, these cations are called network modifiers
- Other oxides such as TiO_2 or Al_2O_3 can replace Si in the network and stabilize it (stabilizers)
- Physically, the addition of network modifiers and stabilizers lowers the melting temperature and the viscosity of the glass



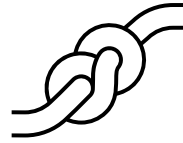
Schematic representation
of the ion positions within
a sodium-silicate glass



In its amorphous structure, the band gap of SiO_2 is broken in a way that no electron transition can occur. As a result, glasses are optically transparent. Further, glass manufacturing is comparably easy.

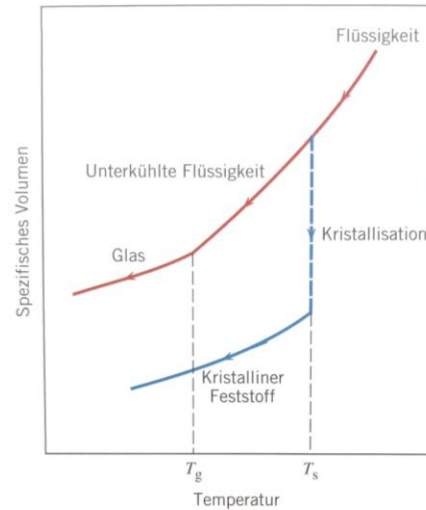
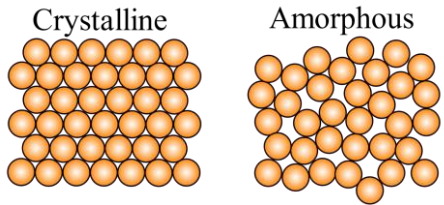


Bonding situation

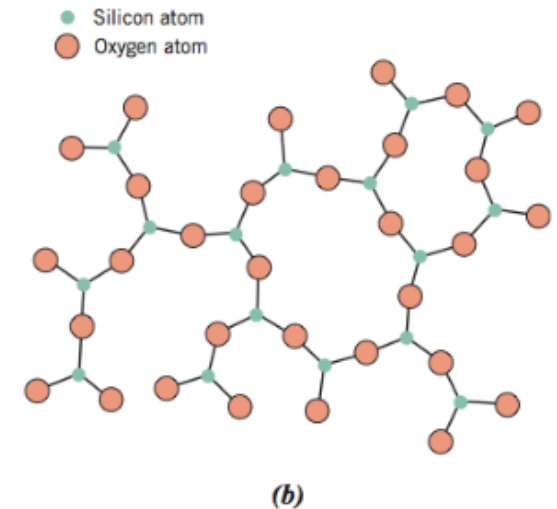
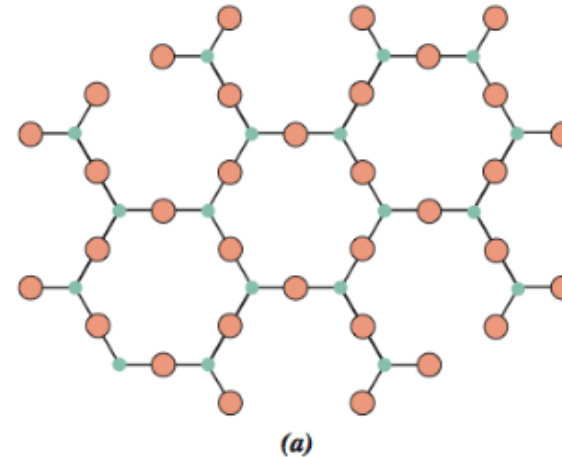


and properties

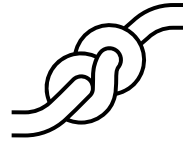
- Crystalline materials: $V \downarrow$ abruptly at T_s
- Glas: $V \downarrow$ continuously at T_g



In its amorphous structure, the band gap of SiO_2 is broken in a way that no electron transition can occur. As a result, glasses are optically transparent. Further, glass manufacturing is comparably easy.

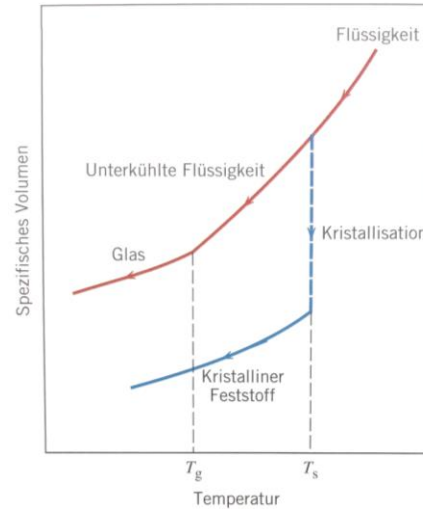
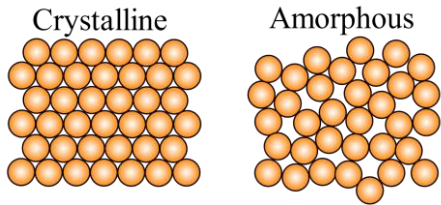


Bonding situation

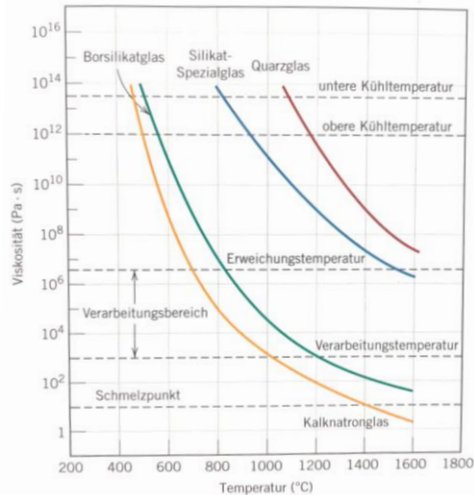


and properties

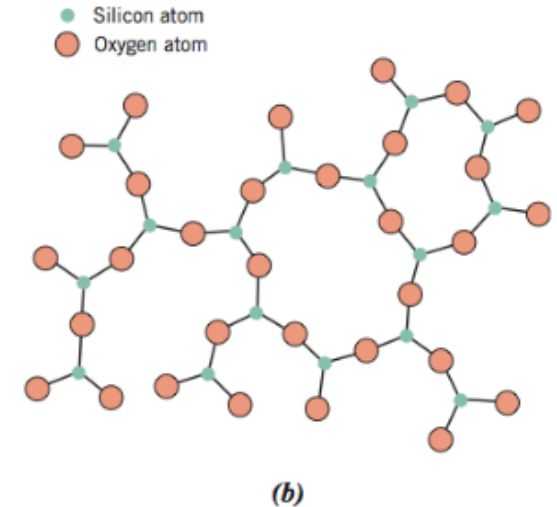
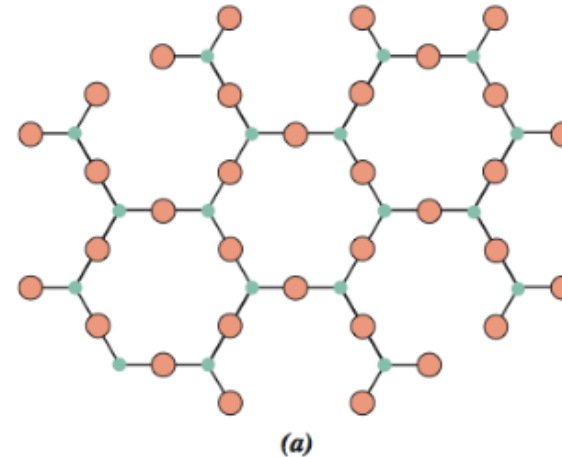
- Crystalline materials: $V \downarrow$ abruptly at T_s
- Glas: $V \downarrow$ continuously at T_s



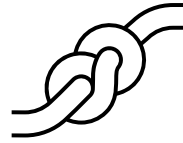
In its amorphous structure, the band gap of SiO_2 is broken in a way that no electron transition can occur. As a result, glasses are optically transparent. Further, glass manufacturing is comparably easy.



- T_{melt} viscosity 10 Pa/s: liquid
- T_{proc} viscosity 10^3 Pa/s: glass can be formed
- T_{soft} viscosity $\sim 10^6$ Pa/s: glass is not deformed
- $T_{\text{up-cool}}$ viscosity 10^{12} Pa/s: present tensions can be reduced
- $T_{\text{down-cool}}$ viscosity $\sim 10^{13}$ Pa/s: breaking before deformation

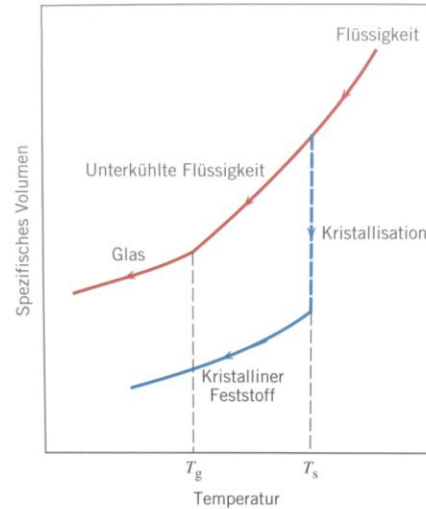
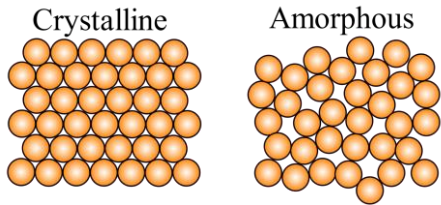


Bonding situation

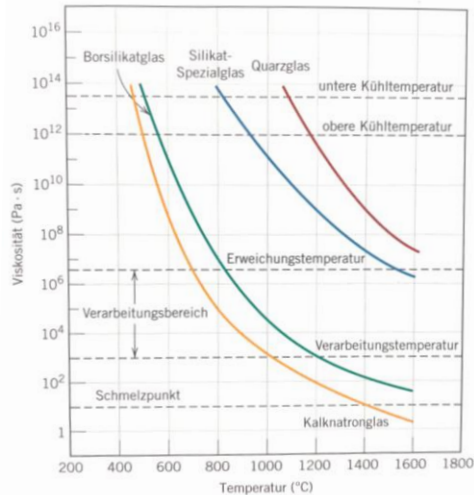


and properties

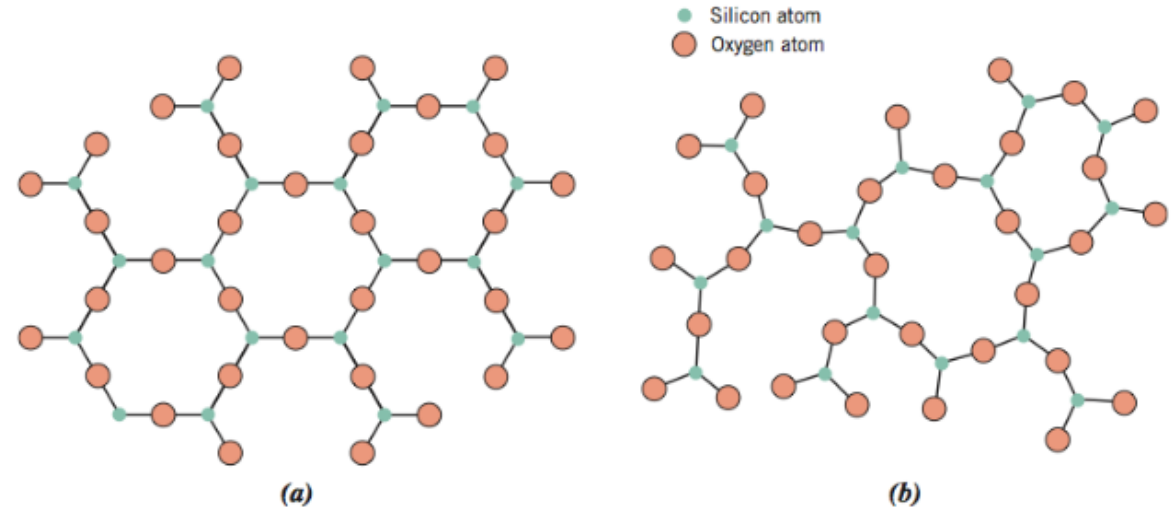
- Crystalline materials: $V \downarrow$ abruptly at T_S
- Glas: $V \downarrow$ continuously at T_S



In its amorphous structure, the band gap of SiO_2 is broken in a way that no electron transition can occur. As a result, glasses are optically transparent. Further, glass manufacturing is comparably easy.



- T_{melt} viscosity 10 Pa/s: liquid
- T_{proc} viscosity 10^3 Pa/s: glass can be formed
- T_{soft} viscosity $\sim 10^6$ Pa/s: glass is not deformed
- $T_{\text{up-cool}}$ viscosity 10^{12} Pa/s: present tensions can be reduced
- $T_{\text{down-cool}}$ viscosity $\sim 10^{13}$ Pa/s: breaking before deformation



More details in glass chemistry lecture – Gunter Heymann!

[Why is Glass Transparent? | MATSE 81: Materials In Today's World](#) (19th February, 2024)

William D. Callister, David G. Rethwisch, Materialwissenschaften und Werkstofftechnik, WILEY-VCH, 2013.

Excursion: Organic glasses



Vesuvius eruption 79 BC

One special victim: Parts of brain and spinal cord remained intact as organic glass.

Herculaneum: Rätsel des gläsernen Gehirns gelöst

Ursache für das einzigartige, verglaste Hirngewebe eines Vulkanopfers identifiziert

28. Februar 2025, Lesezeit: 3 Min.



Teile des Gehirns eines Vulkanopfers von Herculaneum sind bei der Katastrophe verglast – aber wie? © Pier Paolo Petrone

[Unique formation of organic glass from a human brain in the Vesuvius eruption of 79 CE | Scientific Reports](#) (3rd March, 2025)

[Herculaneum: Rätsel des gläsernen Gehirns gelöst - Ursache für das einzigartige, verglaste Hirngewebe eines Vulkanopfers identifiziert - scinexx.de](#) (3rd March, 2025)

Excursion: Organic glasses



Vesuvius eruption 79 BC

One special victim: Parts of brain and spinal cord remained intact as organic glass.

Herculaneum: Rätsel des gläsernen Gehirns gelöst

Ursache für das einzigartige, verglaste Hirngewebe eines Vulkanopfers identifiziert

28. Februar 2025, Lesezeit: 3 Min.



Teile des Gehirns eines Vulkanopfers von Herculaneum sind bei der Katastrophe verglast – aber wie? © Pier Paolo Petrone

For a substance to become glass, a melt must cool down quickly enough to avoid crystallization upon solidification and instead retain an amorphous structure. This requires a significant temperature difference from the surroundings. These special conditions occur very rarely in nature, for example, during certain lightning strikes or volcanic eruptions. The latter can produce the natural rock glass known as obsidian.

[Unique formation of organic glass from a human brain in the Vesuvius eruption of 79 CE | Scientific Reports](#) (3rd March, 2025)

[Herculaneum: Rätsel des gläsernen Gehirns gelöst - Ursache für das einzigartige, verglaste Hirngewebe eines Vulkanopfers identifiziert - scinexx.de](#) (3rd March, 2025)

Excursion: Organic glasses



Vesuvius eruption 79 BC

One special victim: Parts of brain and spinal cord remained intact as organic glass.

Herculaneum: Rätsel des gläsernen Gehirns gelöst

Ursache für das einzigartige, verglaste Hirngewebe eines Vulkanopfers identifiziert

28. Februar 2025, Lesezeit: 3 Min.



Teile des Gehirns eines Vulkanopfers von Herculaneum sind bei der Katastrophe verglast – aber wie? © Pier Paolo Petrone

For a substance to become glass, a melt must cool down quickly enough to avoid crystallization upon solidification and instead retain an amorphous structure. This requires a significant temperature difference from the surroundings. These special conditions occur very rarely in nature, for example, during certain lightning strikes or volcanic eruptions. The latter can produce the natural rock glass known as obsidian.



[Unique formation of organic glass from a human brain in the Vesuvius eruption of 79 CE | Scientific Reports](#) (3rd March, 2025)

[Herculaneum: Rätsel des gläsernen Gehirns gelöst - Ursache für das einzigartige, verglaste Hirngewebe eines Vulkanopfers identifiziert - scinexx.de](#) (3rd March, 2025)

Excursion: Organic glasses



Vesuvius eruption 79 BC

One special victim: Parts of brain and spinal cord remained intact as organic glass.

Herculaneum: Rätsel des gläsernen Gehirns gelöst

Ursache für das einzigartige, verglaste Hirngewebe eines Vulkanopfers identifiziert

28. Februar 2025, Lesezeit: 3 Min.



Teile des Gehirns eines Vulkanopfers von Herculaneum sind bei der Katastrophe verglast – aber wie? © Pier Paolo Petrone

For a substance to become glass, a melt must cool down quickly enough to avoid crystallization upon solidification and instead retain an amorphous structure. This requires a significant temperature difference from the surroundings. These special conditions occur very rarely in nature, for example, during certain lightning strikes or volcanic eruptions. The latter can produce the natural rock glass known as obsidian.

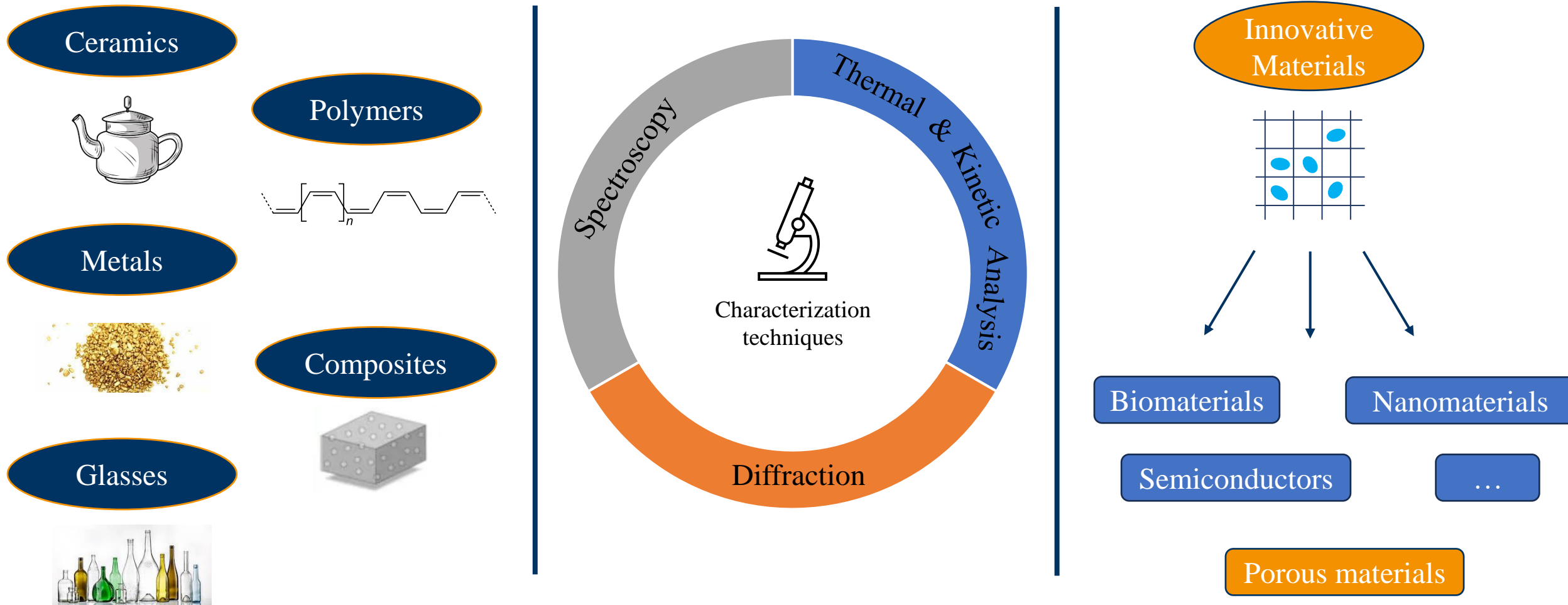


“The cloud must have dissipated just as quickly so that its remnants could cool down rapidly enough to trigger the vitrification process.”

[Unique formation of organic glass from a human brain in the Vesuvius eruption of 79 CE | Scientific Reports](#) (3rd March, 2025)

[Herculaneum: Rätsel des gläsernen Gehirns gelöst - Ursache für das einzigartige, verglaste Hirngewebe eines Vulkanopfers identifiziert - scinexx.de](#) (3rd March, 2025)

A functional material could be defined as being prepared from a “target-motivated” approach



Questions?

