

Live Lab One

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September 17, 2020

- This Lab Requires:

1. Question
2. Data (including calculations)
3. Conclusion

- $NaHCO_3 + HCl \rightarrow H_2O + NaCl + CO_2$

1. Calculate how $NaCl$ formed
2. Calculate Limiting Factor
3. Determine how much $NaCl$ should form
4. Calculate % yield

- Steps:

1. Weight of dish: 36.72[g]

- Data:

1. Weight of dish: 36.72[g]
2. Weight of $NaCl$: 2.00[g]
3. Molarity of HCl : 6[mol L⁻¹]
4. Moles of HCl : $3[mL] \cdot 6[mol L^{-1}] \cdot \frac{1[L]}{1000[mL]} = .018[mol_{HCl}]$
5. Mass Dish + $NaCl$ after procedure: 38.00[g]