Chapter 8 – Thermochemistry

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- Reactions either absorb heat or release it:
  1. Exothermic Reaction – Releases heat
  2. Endothermic Reaction – Intakes heat
- $q = cm\Delta T$, where c is the specific heat and q is the heat/energy
- Units of specific heat are $\left[\frac{\text{J}}{\text{g C}}\right]$
- Enthalpy (ΔH) – Reaction heat content. If $\Delta H < 0$ the reaction is exothermic, but if $\Delta H > 0$, the reaction is endothermic.
- ΔH for a reaction is equal but opposite in sign for reverse.
- Hess's Law – ΔH for a reaction is same whether it occurs directly or in a series.
  1. The enthalpy is the same for the following reactions:

$$\text{A} \rightarrow \text{D}$$

$$\text{A} \rightarrow \text{B} \rightarrow \text{C} \rightarrow \text{D}$$