Chapter 4 – Problem 39-42

Michael Brodskiy
Instructor: Mr. Morgan

October 8, 2020

39. Assign oxidation numbers to each element in:

    (a) $NO \rightarrow N = 2$, $O = -2$
    (b) $NH_3 \rightarrow N = -3$, $H = 1$
    (c) $K_2O_2 \rightarrow K = 1$, $O = -1$
    (d) $ClO_3^- \rightarrow Cl = 5$, $O = -2$

40. Assign oxidation numbers to each element in:

    (a) $CH_4 \rightarrow H = 1$, $C = -4$
    (b) $CO_3^{2-} \rightarrow C = 4$, $O = -2$
    (c) $IO_4^- \rightarrow I = 7$, $O = -2$
    (d) $N_2H_4 \rightarrow H = 1$, $N = -2$

41. Assign oxidation numbers to each element in:

    (a) $ClO_3^- \rightarrow Cl = 5$, $O = -2$
    (b) $H_2SO_3 \rightarrow H = 1$, $O = -2$, $S = 4$
    (c) $K_2O_2 \rightarrow K = 1$, $O = -1$
    (d) $Na_3N \rightarrow Na = 1$, $N = -3$

42. Assign oxidation numbers to each element in:

    (a) $HIO_3 \rightarrow H = 1$, $O = -2$, $I = 5$
    (b) $NaMnO_4 \rightarrow Na = 1$, $O = -2$, $Mn = 7$
    (c) $SnO_2 \rightarrow Sn = 4$, $O = -2$
    (d) $NOF \rightarrow O = -2$, $N = 3$, $F = -1$
    (e) $NaO_2 \rightarrow Na = 1$, $O = -\frac{1}{2}$