

# Chapter 14 — Equilibrium with Acid/Base Reactions

Michael Brodskiy

Instructor: Mr. Morgan

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- Buffered Solutions — Resist pH change. Made of weak acid and concentrated base.
  1. Ex.  $\text{HC}_2\text{H}_3\text{O}_2$  and  $\text{NaC}_2\text{H}_3\text{O}_2$ . Add:  $\text{HCl}$  and  $\text{NaOH}$
- Buffer Capacity — How many ions can be added to destroy the buffers effectiveness