Quiz Chapter 4 — Question One

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$$\begin{split} 3\,\mathrm{Ba(NO_3)_2} + 2\,\mathrm{(NH_4)_3PO_4} &\longrightarrow 6\,\mathrm{NH_4NO_3} + \mathrm{Ba_3(PO_4)_2} \\ 29.5[\mathrm{mL}] &= .0295[\mathrm{L}] \\ .631 \cdot .0295 &= .0186[\mathrm{mol_{(NH_4)_3PO_4}}] \rightarrow .0279[\mathrm{mol_{Ba(NO_3)_2}}] \\ &\frac{.0279}{.773} = .0361[\mathrm{L_{Ba(NO_3)_2}}] \end{split} \tag{1}$$