Chapter 15 — Problem Set 1

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1. $PbI_2 \rightleftharpoons Pb^{2+} + 2I^-$

(a)

$$[Pb^{2+}] [I^{-}]^{2} = 10^{-8}$$

$$(x)(2x)^{2} = 10^{-8}$$

$$x = \sqrt[3]{\frac{10^{-8}}{4}}$$

$$= .00136[M]$$
(1)

(b)

$$(x)(.01)^{2} = 10^{-8}$$

$$x = \frac{10^{-8}}{(.01)^{2}}$$

$$= 1 \cdot 10^{-4} [M]$$
(2)

(c)

$$(.02)(2x)^{2} = 10^{-8}$$

$$x = \sqrt{\frac{10^{-8}}{.08}}$$

$$= 3.54 \cdot 10^{-4} [M]$$
(3)

2. $Mg(OH)_2 \rightleftharpoons Mg^{2+} + 2OH^{-}$

(a)

$$[Mg^{2+}] [OH^{-}]^{2} = 1.8 \cdot 10^{-11}$$

$$(x)(2x)^{2} = 1.8 \cdot 10^{-11}$$

$$x = \sqrt[3]{\frac{1.8 \cdot 10^{-11}}{4}}$$

$$= 1.65 \cdot 10^{-4} [M]$$
(4)

(b)

$$(.1)(2x)^{2} = 1.8 \cdot 10^{-11}$$

$$x = \sqrt{\frac{1.8 \cdot 10^{-11}}{.4}}$$

$$= 6.7 \cdot 10^{-6} [M]$$
(5)

(c)

$$(x)(.25)^{2} = 1.8 \cdot 10^{-11}$$

$$x = \frac{1.8 \cdot 10^{-11}}{.125}$$

$$= 2.88 \cdot 10^{-10} [M]$$
(6)

3. (a)

$$.1 \cdot .00045 = .000045 [\text{mol}_{\text{Ag}^{+}}]$$

$$\frac{.000045}{.35} = 1.286 \cdot 10^{-4} [\text{M}]$$
(7)

(b)

$$.25 \cdot .00075 = .0001875 [\text{mol}_{\text{CrO}_4^{2-}}]$$

$$\frac{.0001875}{.35} = 5.36 \cdot 10^{-4} [\text{M}]$$
(8)

(c)

$$(1.286 \cdot 10^{-4})^2 (5.36 \cdot 10^{-4}) = 8.86 \cdot 10^{-12}$$

 $8.86 > 2$ (9)
So solid *is* formed

4.

$$(.00105)^{2}(x) = 6 \cdot 10^{-4}$$

$$x = 544[M]$$
(10)

5. $PbCl_2 \rightleftharpoons Pb^{2+} + 2Cl^{-}$

$$(x)(2x)^{2} = 3.3 \cdot 10^{-3}$$

$$x = .0938[M]$$

$$.0938 \cdot .756 \cdot 278 = 19.71[g] \text{ at } 80[^{\circ}C]$$

$$(x)(2x)^{2} = 1.6 \cdot 10^{-5}$$

$$x = .0159[M]$$

$$.0159 \cdot .756 \cdot 278 = 3.336[g] \text{ at } 25[^{\circ}C]$$

$$19.71 - 3.336 = 16.38[g]$$

$$(11)$$