

Chapter 7 – Covalent Bonds

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- Lewis Structures:
 1. Sum Valence Electrons
 2. Connect atoms with lines
 3. Arrange remaining e^- to satisfy octet rule
 - Resonance – Being able to draw more than one Lewis structure
 - Halogens will almost never have double bonds
- $\text{C}(-[:0]\text{H})(-[:90]\text{H})(-[:180]\text{H})(-[:270]\text{H})$