

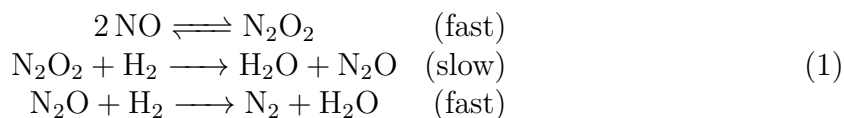
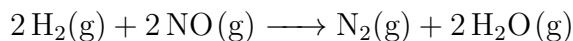
Chapter 11 – Problems 78, 80

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78. For the following reaction, the experimental rate expression is $\text{rate} = k[\text{NO}]^2[\text{H}_2]$. The following mechanism is proposed. Is this mechanism consistent with the rate expression?



Yes, it is consistent, because N_2O_2 may be substituted for 2NO , which, in turn, can be changed into $[\text{NO}]^2$.

80. Two mechanisms are proposed for the following reaction. Show that each of these mechanisms is consistent with the observed rate law: $\text{rate} = k[\text{NO}]^2[\text{O}_2]$

